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## A designed and potentially decadentate ligand for use in lanthanide(III) catalysed biomass transformations: targeting diastereoselective *trans*-4,5-diaminocyclopentenone derivatives†

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The goal of this study was to design a ligand system which can accommodate single lanthanide(III)-ions and investigate the properties of the resulting complexes. The complexes of all the accessible lanthanides and yttrium with the new ligand **LH<sub>6</sub>** = *N,N'*-dimethyl-*N,N'*-ethylene-bis(5-bromo-3-(1*H*-benzimidazol-2-yl)hydrazinylidene)-2-hydroxybenzylamine) were obtained in high yield at room temperature under aerobic reaction conditions. The corresponding compounds were characterised using X-ray diffraction, FT-IR, elemental analysis and the optical properties of all complexes were investigated using UV-vis and fluorescence spectroscopy. The air stable complexes efficiently transform biomass furfural to *trans*-4,5-cyclopentenones in high yield.

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## Introduction

Recently the importance of using lanthanide complexes as catalysts in organic reactions<sup>1–4</sup> and enzymatic processes has been recognised.<sup>5</sup> The properties of lanthanide coordination complexes include (i) electronic and magnetic properties (ii) usually small ligand field but large spin-orbit coupling effects (iii) large coordination numbers and adaptable coordination geometry, both of which are desirable for catalytic applications (iv) a preference for hard donor ligands for oxidation state III as well as high Lewis acidity. This latter point is illustrated by the investigations carried out in natural synthesis reactions for *e.g.* prostaglandins, prostanoid derivatives, and a range of bioactive molecules.<sup>6–12</sup> A particular example of interest to us, is the transformation of furfural to asymmetric bifunctionalized

*trans*-4,5-diaminocyclopentenones through addition of appropriate amines with the help of promoters (Fig. 1). In such a way, furfural, which is readily available from biomass, produced from corncobs, bagasse, *etc.*<sup>13</sup> These structural motifs are versatile building blocks for the retrosynthetic pathway of (–)-agelastatin A,<sup>12</sup> an architecturally unique cytotoxic tetracyclic alkaloid first isolated from the axinellid sponge Agelas dendromorpha.<sup>14,15</sup>

Over the past 20 years the transformation of furfural available from biomass to diastereoselective cyclopentenones has been intensively developed and investigated. Batey and co-workers was first reported the utilisation of Dy(OTf)<sub>3</sub> as Lewis acid to catalyse a domino condensation/ring-opening/electrocyclization process.<sup>6</sup> This pioneer study led to an increasing number of alternative strategies for the formation of *trans*-4,5-diamino-cyclopent-2-enone derivatives such as the employment of the ionic liquid 1-methylimidazolium tetrafluoroborate,<sup>16</sup> tosylamine metal free catalyst,<sup>17</sup> Al(Cl)<sub>3</sub>,<sup>18</sup> 3d/4f-clusters,<sup>19,20</sup> ErCl<sub>3</sub><sup>10</sup> or ErCl<sub>3</sub> immobilized on silica,<sup>21</sup> under microwave irradiation<sup>22</sup> and Cu(OTf)<sub>2</sub> in water as a reusable catalyst.<sup>23</sup> A recent report from Batey and co-workers revealed the efficiency of the Dy(OTf)<sub>3</sub> catalyst for the transformation of

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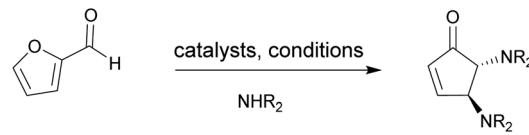


Fig. 1 Transformation of furfural to a *trans*-4,5-diaminocyclopentenone.



the substituted furfuraldehydes to 3-substituted-*trans*-4,5-cyclopentenone in high stereoselectivity and yield.<sup>11</sup> Herein a new family of lanthanide(III)-complexes was prepared and the resulting compounds were characterised using X-ray diffraction, FT-IR, elemental analysis and spectroscopy and their catalytic properties were investigated. We designed the ligand **LH<sub>6</sub>** = (*N,N'*-dimethyl-*N,N'*-ethylene-bis(5-bromo-3-(1*H*-benzimidazol-2-yl)hydrazineylidene)-2-hydroxybenzyl amine) which contains the previously used acetato-diphenoxoamine building block (**1**).<sup>24</sup> The complexations of the corresponding ligand with all available  $\text{Ln}^{3+}$  and  $\text{Y}^{3+}$  ions were performed. In addition, the catalytic properties of these complexes for transformation of furfuran to *trans*-4,5-diamino-cyclopent-2-enone under various reaction conditions were investigated.

## Results and discussion

### Synthesis

The novel ligand **LH<sub>6</sub>** was prepared by the condensation reaction of *N,N'*-dimethyl-*N,N'*-ethylenedi(5-bromo-3-formyl-2-hydroxybenzyl amine (**1**) and 2-hydrazino-1*H*-1,3-benzimidazole (**2**) (Fig. 2).

The  $\text{Ln}^{(III)}$ -complexes were obtained in high yield by the reaction of ligand **LH<sub>6</sub>** with appropriate  $\text{Ln}(\text{NO}_3)_3 \cdot \text{H}_2\text{O}$  salts in the ratio 1 : 1 in a mixture of acetonitrile and methanol (1 : 1, %v/v) at room temperature in air for 30 min. After standing for 7 days, crystal products were collected by filtration. The product could be also directly obtained in the form of powder in higher yield by rapid and prolonged stirring for up to 2 h. The powder adducts were characterised using elemental analysis, IR, UV-vis, fluorescence spectroscopy and Powder X-ray Diffraction (PXRD) (see ESI<sup>†</sup>).

### Crystal structures

The single crystal X-ray data shows that two classes of lanthanide complexes were obtained using the same reaction conditions. The first family (lanthanum to holmium) was found to have 10-fold coordination and the second family (erbium to lutetium) has 8-fold coordination (Fig. 3). This reflects the fact that the atomic radii of lanthanides decrease from lanthanum to lutetium. Additionally, the complexation of yttrium with the same ligand was also investigated. As expected the  $\text{Y}^{(III)}$  has 10-fold coordination given that the atomic radius of  $\text{Y}^{(III)}$  lies

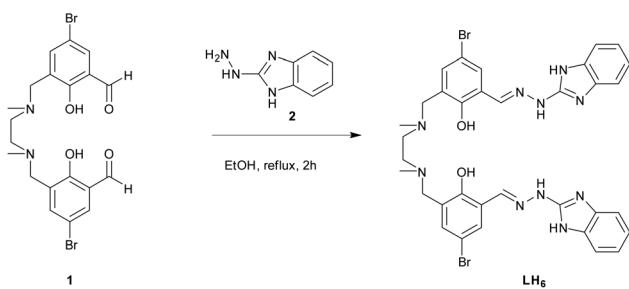


Fig. 2 Preparation of Schiff-base ligand **LH<sub>6</sub>**.

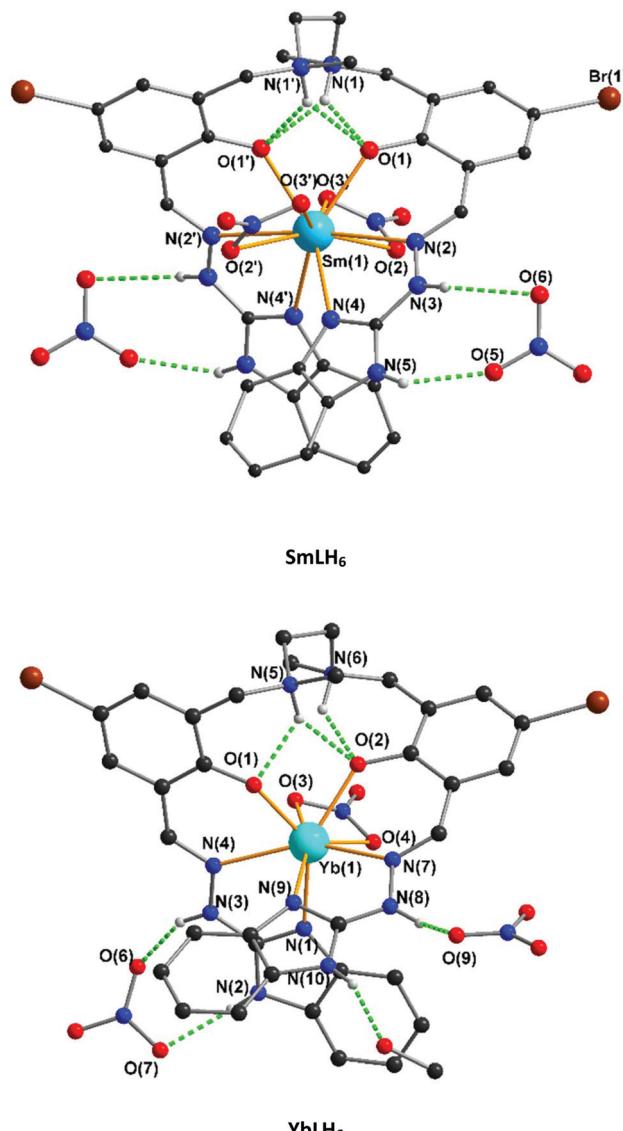


Fig. 3 The structure of **SmLH<sub>6</sub>** and **YbLH<sub>6</sub>**. Colour codes for the atoms: light blue (Sm and Yb), blue (N), red (O) black (C) and brown (Br). The hydrogen bonds are highlighted in green dashed line. Solvent molecules (MeCN) in lattice and H atoms are omitted for clarity.

between that of  $\text{Dy}^{(III)}$  and  $\text{Ho}^{(III)}$  (Fig. S34<sup>†</sup>). The characterization of the powder sample using PXRD confirmed the isostructural nature for family one by comparison with the **SmLH<sub>6</sub>** or family two with **YbLH<sub>6</sub>** (Fig. 4). Regarding the first class of this system **SmLH<sub>6</sub>** was chosen as representative for the lanthanides La, Ce, Pr, Nd, Sm, Eu, Gd, Tb, Dy and Ho. From the X-ray crystallographic analysis, **SmLH<sub>6</sub>** is composed of a  $\text{Ln}^{(III)}$  which is centred on a two fold axis. Thus the asymmetric unit is half of the complete molecule. The  $\text{Sm}^{(III)}$ -complex crystallizes in the space group  $P6_{1}22$  and the analysis using SHAPE<sup>25</sup> shows that the best coordination geometry is Sphenocorona J87 with a deviation value 1.921. The coordination sphere consists of four-coordinating N atoms and two phenoxo O atoms of the ligand and four coordinating O atoms of two bidentate nitrates. Intramolecular hydrogen bonds



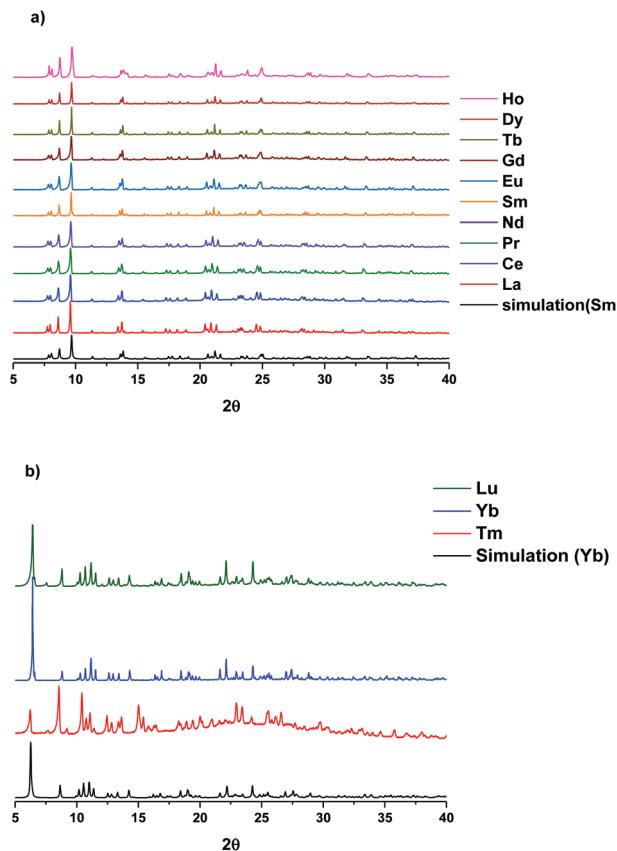


Fig. 4 PXRD of measured  $\text{Ln}(\text{III})\text{LH}_6$  compare to simulation (a) 10-fold coordination and (b) 8-fold coordination.

between the phenolate O1 and the nitrogen back bone amine N1 form through the formal transfer of the phenolate hydrogen to N1. The charge balance is provided by two nitrates, which coordinate to the  $\text{Sm}^{3+}$  ion and one nitrate which form intermolecular hydrogen bonds with the N3 and N5 of the ligand. The Sm–O distance ranges from 2.334(2) to 2.584(2) Å, whereas the Sm–N distance is between 2.538(3) and 3.000(3) Å.

On the other hand, smaller sized lanthanide metal ions from  $\text{Tm}^{3+}$  to  $\text{Lu}^{3+}$  represented by the  $\text{Yb}(\text{III})$ -complex, which crystallizes in the space group  $P\bar{1}$  adopt a triangular dodecahedron geometry with a deviation value 1.911 (SHAPE) through the coordination of four N atoms and two phenoxy O atoms of the ligand and two O atoms from a bidentate nitrate. The charge balance is provided by one nitrate coordinated to  $\text{Yb}^{3+}$  and two nitrate counterions which form intermolecular hydrogen bonds with N2, N3 and N8 of the ligand. The Yb–O distance ranges from 2.141(3) to 2.463(4) Å, whereas the Yb–N distance is between 2.345(4) and 2.854(4) Å.

The crystallographic parameters and refinement details for  $\text{SmLH}_6$  and  $\text{YbLH}_6$  are summarized in Table S1.† The  $\text{ErLH}_6$  is the only complex which crystallizes in the space group  $P2_1/c$  and adopts triangular dodecahedron geometry with a deviation value 2.258 (SHAPE) (Fig. 5). The coordination environment of  $\text{Er}(\text{III})$ -complex is similar to that of the second family with 8-fold coordination but now the two coordinating oxygen

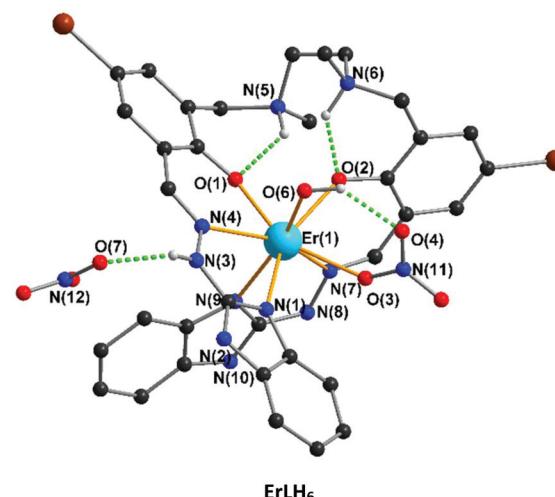


Fig. 5 The structure of  $\text{ErLH}_6$ . Colour codes for the atoms: light blue (Er), blue (N), red (O) black (C) and brown (Br). The hydrogen bond is highlighted in green dashed line. Solvent molecules (MeCN) in lattice and H atoms are omitted for clarity.

atoms are from O6 of a water molecule and O3 of a nitrate. This is reflected by the different space group and powder X-ray pattern (Fig. 6). The crystallographic data for  $\text{ErLH}_6$  is summarized in Table S2.†

### Spectroscopic properties

IR spectra of  $\text{LH}_6$  and the corresponding  $\text{Ln}(\text{III})$ -complexes were measured (see the ESI†). The broad band of  $\nu(\text{OH})$  of phenolic alcohol of the free ligand around 2863–2767  $\text{cm}^{-1}$  disappears when it coordinates to the metal ion. In addition, the sharp bands in the region 1656–1599  $\text{cm}^{-1}$  which can be assigned to the azomethine  $\nu(\text{C}=\text{N})$  stretching frequency of the coordinated ligands were found to be shifted to lower frequency in comparison to that of the free ligands (1687–1619  $\text{cm}^{-1}$ ). This indicates the coordination of the imine nitrogen atom to the metal centre. The other  $\nu(\text{C–H})$ ,  $\nu(\text{C}=\text{C})$  and  $\nu(\text{C–O})$  vibrations are found in the expected ranges from 1465–735  $\text{cm}^{-1}$  (Fig. S5†).

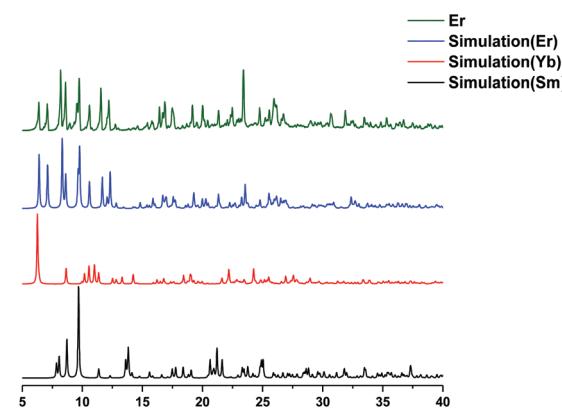


Fig. 6 Comparison of PXRD  $\text{ErLH}_6$  with the simulation (Er, Sm, and Yb).

## UV-vis and fluorescence

The optical properties were investigated using UV-vis and fluorescence spectroscopy. For UV-vis spectroscopy, the absorption maxima at 368 nm (**LaLH<sub>6</sub>**), 369 nm (**CeLH<sub>6</sub>**), 375 nm (**GdLH<sub>6</sub>**), 381 nm (**YbLH<sub>6</sub>**) and 381 nm (**LuLH<sub>6</sub>**) were found to be slightly red shifted from 368 nm of the free ligand. In addition, the strong absorption peaks at 286 and 316 nm (**LaLH<sub>6</sub>**), 284 and 311 nm (**CeLH<sub>6</sub>**), 286 and 320 nm (**GdLH<sub>6</sub>**), 280 and 321 nm (**YbLH<sub>6</sub>**) and 282 and 321 nm (**LuLH<sub>6</sub>**) are related to metal to ligand charge transfer (MLCT) which, of course, are not found in the free ligand. The fluorescence characteristics of these complexes in comparison to **LH<sub>6</sub>** were measured using an excitation wavelength of 390 nm. The emission maxima at 481 nm (**LaLH<sub>6</sub>**), 480 (**CeLH<sub>6</sub>**), 479 (**GdLH<sub>6</sub>**), 484 (**YbLH<sub>6</sub>**), and 480 nm (**LuLH<sub>6</sub>**) were found to be slightly red shifted in comparison to the free ligand at 479 nm. The fluorescence intensity of these complexes were found to be quenched when the free ligand coordinates to lanthanide ions. This optical behaviour reveals the stronger the metal to ligand charge transfer, the weaker the intensity of fluorescence. It is clearly seen that the diamagnetic **La(III)**-complex shows only 4% decrease in fluorescence intensity in comparison to the 55% in the **Yb(III)**-complex (Fig. 7).

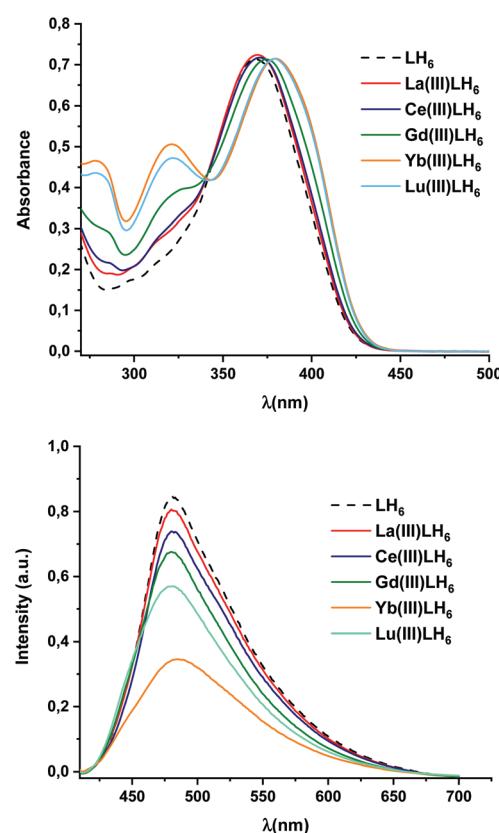


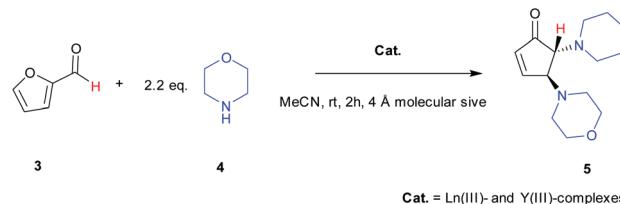
Fig. 7 UV-vis and fluorescence spectra of **LH<sub>6</sub>** compared with **LaLH<sub>6</sub>**, **GdLH<sub>6</sub>**, **YbLH<sub>6</sub>** and **LuLH<sub>6</sub>** complexes in 50  $\mu$ M in DMSO ( $\lambda_{\text{exc}} = 390$  nm).

## Catalytic studies

The catalytic efficiencies of the series of **Ln(III)**-complexes and **Y(III)**-complex (**Cat.**) for transformation of furfural (**3**) to *trans*-4,5-dimopholin-2-pentenone (**5**) were first investigated (Scheme 1). All complexes were used at 1 mol% catalytic loading. The reactions were carried out in acetonitrile at room temperature for 2 h. The results reveal that the mononuclear complexes are highly efficient in converting furfural to the target products in high conversion yield as seen from the ratio of <sup>1</sup>H signal at 9.67 ppm (aldehyde proton) to the product at 3.29 ppm (*trans* proton) (Table 1).

The **Dy(III)-Cat.** shows the maximum conversion yield in comparison to others. Thus, the next step was to evaluate the catalytic properties of the **DyLH<sub>6</sub>** complex under different reaction conditions. To this end, the time dependency and catalytic loading were investigated. The results show that **DyLH<sub>6</sub>** at low catalytic loading is highly efficient in transforming biomass furfural to the desired *trans*-4,5-diaminocyclopentenone adduct in high to quantitative conversion yield (Table 2).

Quantitative yield is achieved after 24 h. However, 98% can be achieved after 4 h thus, the condition for 4 h reaction was used for the extension of catalysis scope (Table 3) for the formation of *trans*-4,5-diaminocyclopentenone using amine derivatives.



Scheme 1 Catalysis of **Ln(III)**- and **Y(III)**-complexes (**Cat.**) for transformation of furfural (**3**). The protons followed using NMR are shown in red.

Table 1 Catalytic study of **Ln(III)-LH<sub>6</sub>** under air exposure for 2 h

Catalyst	(Yield) <sup>a</sup>
<b>LaLH<sub>6</sub></b>	80
<b>CeLH<sub>6</sub></b>	83
<b>PrLH<sub>6</sub></b>	90
<b>NdLH<sub>6</sub></b>	88
<b>SmLH<sub>6</sub></b>	82
<b>EuLH<sub>6</sub></b>	81
<b>GdLH<sub>6</sub></b>	85
<b>TbLH<sub>6</sub></b>	81
<b>DyLH<sub>6</sub></b>	95 <sup>b</sup>
<b>HoLH<sub>6</sub></b>	90
<b>ErLH<sub>6</sub></b>	87
<b>TmLH<sub>6</sub></b>	79
<b>YbLH<sub>6</sub></b>	89
<b>LuLH<sub>6</sub></b>	81
<b>YLH<sub>6</sub></b>	87

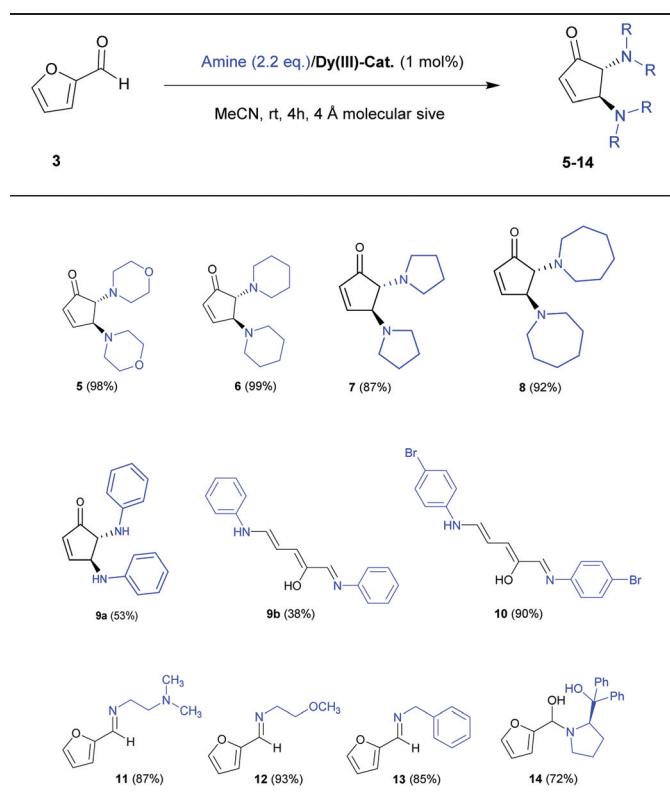
<sup>a</sup> The conversion yields were determined by the ratio of <sup>1</sup>H NMR signal at 9.67 ppm (aldehyde proton) to the product at 3.29 ppm (*trans* proton). <sup>b</sup> Highlighting the maximum conversion yield for the **Dy(III)-Cat.**



**Table 2** Catalytic study of **DyLH<sub>6</sub>** under air exposure

Time (h)	0.5 mol% (Yield %) <sup>a</sup>	1.0 mol% (Yield) <sup>a</sup>
0.5	87	91
1	93	95
2	95	96
4	97	98
24	Quantitative	Quantitative

<sup>a</sup> Followed by <sup>1</sup>H NMR (see ESI Fig. S38 and S39†).

**Table 3** Preparation of *trans*-4,5-diaminocyclopentenone catalysed by **Dy(III)-Cat.** under air exposure

Primary and secondary amine derivatives and also aromatic amines were employed in this study. It is clearly seen that the reaction of furfural with secondary amines gives the *trans*-4,5-diaminocyclopentenone product in high yield (5–8). In addition, when the primary aromatic amine derivatives were used, a mixture of the open and closed form was observed (9a and 9b). This can be ascribed to the more stable Stenhouse intermediate formation in the first step (9b) thus it is less likely that the cyclopentenone product will form in the cyclization step. However, protic acid conditions used in the purification step (solvents and silica gel), the desired closed form (9a) product can be obtained. In addition, the reaction of primary aliphatic amine derivatives with furfural under the same reaction conditions was investigated. The expected *trans*-4,5-diaminocyclopentenone products were not achieved unless the imine intermediates were obtained in high yield (com-

pounds 11–14). Nevertheless, this catalytic system can be employed for the transformation of biomass furfural to *trans*-4,5-diaminocyclopentenones in high yield.

This system is attractive because the synthesis of the complexes can be done under mild aerobic conditions in high yields. Furthermore, the almost quantitative catalysis of furfural to the building blocks for natural products synthesis is also achieved under mild aerobic conditions. A further attractive aspect is the fluorescence behaviour of the complexes which could be used in lanthanide sensing.

## Conclusions

The potentially decadentate ligand **LH<sub>6</sub>** can accommodate 8-fold and 10-fold coordinate single ion lanthanum to lutetium and yttrium complexes. Crystalline products were obtained in high yield after 7 days under aerobic conditions. Moreover, the powder product can be directly obtained in higher yield after vigorously stirring for 2 h. The optical behaviour of these complexes, for instance the metal to ligand charge transfer process, was reflected by the increase in absorption intensity at 280 nm and 321 nm from UV-vis and the decrease in fluorescence intensities in comparison to the free ligand, making this a candidate for metal fluorescence sensors. The catalytic study reveals high efficiency and the amount of catalysts can be reduced to 0.5 mol% loading without loss of activity. Furthermore, the efficient 4 h reaction time could be applied to a range of amines with over 90% yield. This is thus a promising catalytic system for transforming furfural which also shows fluorescence properties.

## Conflicts of interest

The authors declare no conflict of interest.

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## References

- 1 T. Y. Chiu, W. Chin, J. R. Guo, C. F. Liang and P. H. Lin, *Chem. – Asian J.*, 2019, **14**, 627–633.
- 2 X. Xia, C. Lu, B. Zhao and Y. Yao, *RSC Adv.*, 2019, **9**, 13749–13756.
- 3 Y. Xie, C. Lu, B. Zhao, Q. Wang and Y. Yao, *J. Org. Chem.*, 2019, **84**, 1951–1958.
- 4 C. Lu, L. Hu, B. Zhao and Y. Yao, *Organometallics*, 2019, **38**, 2167–2173.
- 5 L. Daumann, *Angew. Chem., Int. Ed.*, 2019, **58**, 12795–12802.



6 S. W. Li and R. A. Batey, *Chem. Commun.*, 2007, **8**, 3759–3761.

7 L. Sanzhong, Z. Baolian, X. Ming, A. Jañczuk, X. Wenhua, C. Jinpei and P. G. Wang, *Chin. Sci. Bull.*, 2008, **46**, 1673–1681.

8 G. K. Veits, D. R. Wenz and J. Read De Alaniz, *Angew. Chem., Int. Ed.*, 2010, **49**, 9484–9487.

9 L. I. Palmer and J. Read De Alaniz, *Angew. Chem., Int. Ed.*, 2011, **50**, 7167–7170.

10 A. Procopio, P. Costanzo, M. Curini, M. Nardi, M. Oliverio and G. Sindona, *ACS Sustainable Chem. Eng.*, 2013, **1**, 541–544.

11 A. Hiscox, K. Ribeiro and R. A. Batey, *Org. Lett.*, 2018, **20**, 6668–6672.

12 P. A. Duspara and R. A. Batey, *Angew. Chem., Int. Ed.*, 2013, **52**, 10862–10866.

13 J. J. Bozell and G. R. Petersen, *Green Chem.*, 2010, **12**, 539–554.

14 M. M. Domostoj, E. Irving, F. Scheinmann and K. J. Hale, *Org. Lett.*, 2004, **6**, 2615–2618.

15 F. A. Davis and J. Deng, *Org. Lett.*, 2005, **7**, 621–623.

16 D. Ramesh, T. S. Reddy, M. Narasimhulu, S. Rajaram, N. Suryakiran, K. C. Mahesh and Y. Venkateswarlu, *Chem. Lett.*, 2009, **38**, 586–587.

17 J. Liu, J. Yu, M. Zhu, J. Li, X. Zheng and L. Wang, *Synthesis*, 2013, **45**, 2165–2170.

18 J. P. M. Nunes, C. A. M. Afonso and S. Caddick, *RSC Adv.*, 2013, **3**, 14975–14978.

19 K. Griffiths, C. W. D. Gallop, A. Abdul-Sada, A. Vargas, O. Navarro and G. E. Kostakis, *Chem. -Eur. J.*, 2015, **21**, 6358–6361.

20 K. Griffiths, P. Kumar, J. D. Mattock, A. Abdul-Sada, M. B. Pitak, S. J. Coles, O. Navarro, A. Vargas and G. E. Kostakis, *Inorg. Chem.*, 2016, **55**, 6988–6994.

21 M. S. Estevão and C. A. M. Afonso, *Tetrahedron Lett.*, 2017, **58**, 302–304.

22 M. Nardi, P. Costanzo, A. De Nino, M. L. Di Gioia, F. Olivito, G. Sindona and A. Procopio, *Green Chem.*, 2017, **19**, 5403–5411.

23 R. F. A. Gomes, N. R. Esteves, J. A. S. Coelho and C. A. M. Afonso, *J. Org. Chem.*, 2018, **83**, 7509–7513.

24 Y. Masami, M. Yukiko, O. Masaaki and O. Hisashi, *Chem. Lett.*, 1996, **25**, 601–602.

25 P. Llunell, M. Casanova, D. Cirera, J. Alvarez and S. Alemany, *SHAPE: Program for the Stereochemical Analysis of Molecular Fragments by Means of Continuous Shape Measures and Associated Tools, version 2.1*, Universitat de Barcelona, Barcelona, 2013.

