



Cite this: *Phys. Chem. Chem. Phys.*, 2020, 22, 11474

The viscosity of dilute carbon nanotube (1D) and graphene oxide (2D) nanofluids†

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Controlling the physicochemical properties of nanoparticles in fluids directly impacts on their liquid phase processing and applications in nanofluidics, thermal engineering, biomedicine and printed electronics. In this work, the temperature dependent viscosity of various aqueous nanofluids containing carbon nanotubes (CNTs) or graphene oxide (GO), *i.e.* 1D and 2D nanoparticles with extreme aspect ratios, is analyzed by empirical and predictive physical models. The focus is to understand how the nanoparticle shape, concentration, motion degrees and surface chemistry affect the viscosity of diluted dispersions. To this end, experimental results from capillary viscosimeters are first examined in terms of the energy of viscous flow and the maximum packing fraction applying the Maron–Pierce model. Next, a comparison of the experimental data with predictive physical models is carried out in terms of nanoparticle characteristics that affect the viscosity of the fluid, mostly their aspect ratio. The analysis of intrinsic viscosity data leads to a general understanding of motion modes for carbon nanoparticles, including those with extreme aspect ratios, in a flowing liquid. The resulting universal curve might be extended to the prediction of the viscosity for any kind of 1D and 2D nanoparticles in dilute suspensions.

Received 28th January 2020,
Accepted 24th April 2020

DOI: 10.1039/d0cp00468e

rsc.li/pccp

Introduction

Carbon nanotubes (CNTs) and graphene have found a number of applications in electronics, sensors, energy devices, catalysis and biomedicine.^{1–7} Very often, these nanostructures need to be dispersed in liquid media. A high quality dispersion, in which nanoparticles are stable and individualized, facilitates processing by different techniques, such as spray-coating, spin-coating, inkjet printing, coagulation spinning, and dip-coating. Also, the liquid suspensions, usually called nanofluids, can be directly put into heat exchange, microfluidic or biological systems. In all the cases, the control of viscosity is a key issue, since it determines the conditions of processing, such as pumping, injection, stirring, permeation and painting, and lastly, the final device operation.

According to the ideas of green chemistry, inks and nanofluids should be most preferably prepared using water as the base liquid. Both CNTs and graphene are hydrophobic and cannot be directly dispersed in water. Therefore, CNTs have to be stabilized in suspension through either the addition of surfactants,^{8,9} or suitable chemical functionalization.¹⁰ Besides, the oxidized form of graphene, namely graphene oxide (GO), is relatively stable in water due to the high ratios of oxygen functional groups.¹¹

The rheology of aqueous carbon nanofluids has been extensively studied for thermal transfer and lubricant applications.^{12–15} Experimental works include measurements on different kinds of multi-walled carbon nanotubes (MWCNTs),^{16–20} single-walled carbon nanotubes (SWCNTs),^{15,21–23} graphene nanoplatelets,²⁴ and GO.²⁵ Also, measurements on covalently functionalized CNTs (f-CNTs) have been reported,^{26,27} and it has been observed that GO reduction by chemical methods leads to a decrease in the viscosity of the nanofluid.²⁵ Both CNT and GO dispersions in water are non-Newtonian liquids at relatively high concentrations, and typically form isotropic (nematic) phases.^{28,29} However, they approach the Newtonian behavior in the low concentration region.^{18–20,30,31}

From a fundamental viewpoint, CNTs and GO are nanoparticles with extreme aspect ratios, by far exceeding those of micrometric rods, cylinders and discs made of metals, ceramics and polymer materials.^{32–34} The viscosity of dilute CNT suspensions

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† Electronic supplementary information (ESI) available: Elemental analysis of powder samples; analysis of dispersion quality; kinematic viscosity; density of the CNT and GO dispersions; dependence of relative viscosity with temperature; the Maron and Pierce equations for non-Newtonian suspensions; fitting parameters of eqn (1); accuracy of viscosity measurements; calculation of the particle density; electron microscopy images and the log-normal distribution; XRD experiments for the determination of n_L and d_{LL} in GO; chemical functionalization of MWCNTs. See DOI: 10.1039/d0cp00468e



has been correlated with the CNT length,^{35,36} as well as with the bundling degree.^{37,38} Sedimentation studies by analytical ultracentrifugation reveal that CNTs in stable dispersions can be regarded as rigid cylinders.^{39,40} Also, analytical ultracentrifugation has been applied to the determination of lateral dimensions of GO, considering GO flakes as flat discs.⁴¹

Most theoretical studies on the CNT viscosity are based on different expansions of classical two phase flow models,⁴² which are here the starting point for the discussion. Molecular dynamics simulations have been applied to the CNT viscosity,¹⁵ but are out of the scope of the present study. Also, recent mathematical approaches searching for improved algorithms to numerically fit experimental data are not discussed.^{43,44} Our focus lies on a general conceptual understanding by first principles theories. Since both CNTs and GO are nanostructures based on sheets of hexagonally bonded carbon atoms, the search for a common description of their properties might be natural.

We performed viscosity measurements on dilute dispersions of various CNT samples, f-CNTs and GO. All the GO samples and one type of f-CNT sample were directly dispersed in water. The other CNT and f-CNT samples needed the addition of surfactants. The complete set of measurements was first analyzed as a function of the temperature and the nanoparticle concentration, according to the empirical equation of Maron and Pierce.⁴⁵ Next, size distributions for each sample were determined, and the viscosity was evaluated in terms of the aspect ratio (r_p) using predictive physical models. The r_p parameter is defined identically for both CNTs and GOs as the ratio between the particle dimension along a symmetry rotation axis and the perpendicular diameter. For CNTs, r_p is the ratio between the nanotube length and diameter, and for GOs, it is the ratio between the thickness and the average flake width. In this novel way, $r_p \gg 1$ for CNTs, while $r_p \ll 1$ for GO particles and we interpret the viscosity of 1D and 2D nanostructures only as a function of a unique common r_p parameter. Further analysis allows the discussion of the nanoparticle dynamics in the fluid according to the two employed predictive models.

Dependence of viscosity on temperature and concentration

Viscosity data and choice of the fitting model

Dilute aqueous dispersions were prepared with various samples of GO, CNTs and f-CNTs (Tables 1–3). Three GO materials were home made from a graphite precursor applying different oxidation times, while the other GO sample is a commercial product. Three commercial CNT powders were considered,

including two SWCNTs and a MWCNT sample. The MWCNT material was functionalized by four procedures providing different chemical groups on the nanotube surface (Table 3). The complete chemical analysis of all the powder materials is provided in the ESI† (Section S1).

Most of the CNT samples were dispersed with the help of a surfactant, either sodium dodecyl benzene sulfonate (SDBS) or Solsperse™ 27000, which is an acrylic commercial component for inks. One of the f-CNT samples, the HSO₃-MWCNT sample, and the GOs were directly dispersed in water. Dispersions of the nanomaterials were prepared at initial concentrations in the range of 0.4–4 mg mL⁻¹ with the help of ultrasound. Most of the dispersions were centrifuged at a moderate speed of 4500 rpm to avoid large aggregates. Only the GO-16 and GO-G dispersions were not centrifuged since they did not contain large aggregates. Further details on the quality of the dispersions are provided in the ESI† (Section S2).

The viscosity (η) of a fluid is defined as the ratio between the shear stress (F) and the shear rate (G): $\eta = F/G$. For a suspension of solid particles in a liquid, several models aim to describe the dependence of η on factors such as temperature and particle concentration and shape. The relative viscosity (η_r), which is defined as the ratio between the dispersion viscosity (η) and the viscosity of the base fluid (η_0), is mainly controlled by the particle shape and size, as well as the volume fraction (ϕ). Einstein's pioneering work on dilute suspensions of hard spheres,⁴⁶ for which $\eta_r = 1 + 2.5\phi$, has been complemented by successive approaches taking into account the non-sphericity of the particles and the effect of higher concentrations. In some of those approaches, η_r is a polynomial function of ϕ , whose coefficients model the dependence on size and shape.^{47–50} A drawback of the mentioned models is their prediction of a finite viscosity regardless of the ϕ value, which is clearly unphysical. Viscosity must diverge as ϕ tends to the maximum packing fraction (ϕ_m), which depends on the particle geometry. This behaviour is predicted by models such as Mooney,⁵¹ Krieger–Dougherty,⁵² and Maron–Pierce.⁴⁵ While Mooney and Krieger–Dougherty models have two fitting parameters, the Maron–Pierce model considers just one, as far as regards η_r . Fitting our experimental data with two parameters results in a slightly modified value of ϕ_m , but brings out no significant improvement in the standard deviation. Thus, the Maron–Pierce model is chosen here for simplicity.

The model of Maron and Pierce leads to the following equation for the absolute viscosity:⁴⁵

$$\eta = \frac{K_a^0}{(1 - \varepsilon C)^2} e^{\Delta H_1^*/RT} \quad (1)$$

Table 1 Synthesis of GO materials: t_o = time in the oxidant liquid mixture

Sample	Manufacturer	Synthesis method	t_o [h]
GO-2	GCNN, ICB-CSIC	Modified hummers	2
GO-4	GCNN, ICB-CSIC	Modified hummers	4
GO-16	GCNN, ICB-CSIC	Modified hummers	Overnight
GO-G	Graphenea (San Sebastián, Spain)	—	—



Table 2 Pristine CNT materials: specifications from the provider. D = average nanotube diameter; L = nanotube length

Label ^a	Manufacturer (commercial name)	Synthesis method	D [nm]	L [μm]	Carbon purity ^b [%]
MWCNT	Nanocyl (NC7000)	CCVD	9.5	1.5	90
CVD-SWCNT	Chasm Advanced Materials (SG76)	CVD (CoMoCAT [®])	0.83	1	≥ 90
AD-SWCNT	Carbon Solutions Inc. (AP-SWNT)	Arc discharge	1.4	> 1	60–70

^a The abbreviation in SWCNTs indicates the synthesis method. ^b From TGA residues.

Table 3 Functionalized MWCNT materials: summary of functionalization protocols and resulting functional groups

Label	Treatment	Functional groups
c-HNO ₃ -MWCNT	9.5 M HNO ₃ , 18 h reflux	-COOH, -CHO, -OH
d-HNO ₃ -MWCNT	1.5 M HNO ₃ , 2 h reflux	-COOH, -CHO, -OH
F-MWCNT	Isopentyl nitrite, H ₂ N-C ₆ H ₅ -(CF ₂) ₇ -CF ₃ , 12 h, 60 °C	-C ₆ H ₅ -(CF ₂) ₇ -CF ₃
HSO ₃ -MWCNT	Isopentyl nitrite, H ₂ N-C ₆ H ₅ -SO ₃ H, 1 h, 80 °C	-C ₆ H ₅ -SO ₃ H

where K_a^0 is a constant, C is the concentration of nanoparticles in mg mL^{-1} , and ε modulates the influence of C . The enthalpy of activation for viscous flow (ΔH_1^*) indicates the change with T , and is expected to depend only on the liquid medium. The viscosity of the liquid medium (η_0) is thus given by

$$\eta_0 = K_a^0 e^{\Delta H_1^*/RT} \quad (2)$$

Therefore, the relative viscosity (η_r) is defined as

$$\eta_r = \frac{\eta}{\eta_0} = \frac{1}{(1 - \varepsilon C)^2} \quad (3)$$

Actually, eqn (1) and (3) are simplified equations that are only valid for Newtonian fluids. Eqn (3) has been previously applied to suspensions of MWCNTs in aqueous solutions of sodium dodecyl benzene sulfonate (SDBS) surfactant.¹⁸

For the present work, viscosity measurements were performed on dilute aqueous suspensions. The solid concentration (C) in the liquid was always kept below 4 mg mL^{-1} , so it can be reasonably assumed that dispersions behave as Newtonian fluids. Accordingly, viscosities were measured in Ubbelohde capillary viscosimeters, which provide very precise measurements in the low-viscosity regime. For instance, an Ubbelohde viscosimeter has been utilized for the assessment of the exfoliation degree of SWCNT dispersions.³⁷ In Ubbelohde viscosimeters, a suspended column of the liquid flows inside the capillary under laminar regime, by gravity and in a unidirectional way, while the flowing time is registered. The direct measurement from Ubbelohde viscosimeters is the kinematic viscosity (ν), which is related to the dynamic viscosity (η) through the fluid density (ρ): $\eta = \nu \cdot \rho$. The complete set of ν and ρ measurements is provided in the ESI† (Sections S3 and S4).

The kinematic viscosity of GO dispersions at temperatures between 298.15 and 318.15 K is in the range of 4.0 to $0.7 \text{ mm}^2 \text{ s}^{-1}$. For CNTs, the kinematic viscosity is between 2.8 and $0.6 \text{ mm}^2 \text{ s}^{-1}$, except in the Solspense medium, where it reaches $6.1 \text{ mm}^2 \text{ s}^{-1}$ at 298.15 K. The ρ values of all the measured CNT and GO dispersions are nearly that of pure water (Table S4, ESI†).

Before continuing with the analysis of viscosity at different nanoparticle concentrations, it was confirmed that the relative viscosity (η_r) does not substantially change with temperature.²¹

This condition, which indicates a Newtonian behavior, was clearly fulfilled by all the GO and CNT systems in water and SDBS. Deviations of 10% in the relative viscosity were found for MWCNT dispersions in Solspense at the highest temperatures and concentrations. Solspense leads to viscosities of $> 4 \text{ cP}$ at room temperature, and thus the data interpretation has to be performed carefully. However, for most of the analyzed dispersions, it is evident that eqn (1) and (3) can be applied instead of the complete model for non-Newtonian suspensions (ESI† Sections S5 and S6).

Outcome of the experimental data fitting

The simplified Maron–Pierce model fits reasonably well our viscosity measurements. Fig. 1 and 2 show the plots of η_r vs. C for GO-16 in water and MWCNTs in 0.5% SDBS and 20% Solspense. The difference between measured and fitted values is $< 1\%$, except in the MWCNT/Solspense system. The trend for GO (Fig. 1) is linear at concentrations below 2 mg mL^{-1} , and

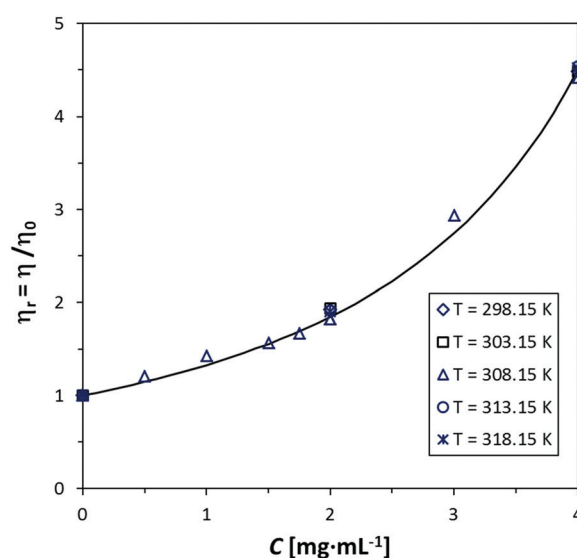


Fig. 1 Relative viscosity (η_r) as a function of concentration (C) of GO-16 in water at 5 temperatures. The fitting line was calculated with the Maron–Pierce model (eqn (1)).



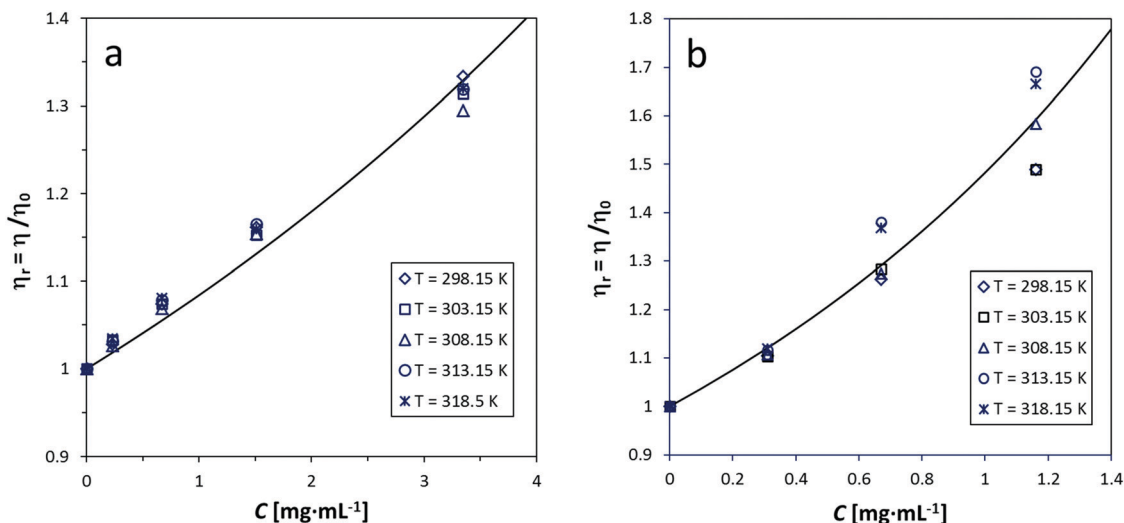


Fig. 2 Relative viscosity (η_r) as a function of concentration (C) of MWCNTs at 5 temperatures in (a) 0.5% SDBS and (b) 20% Solsperse. The fitting line was calculated with the Maron–Pierce model (eqn (1)).

it curves upward at higher concentrations. This fact most probably indicates the very beginning of a semi-dilute regime and thus the Newtonian limit. For CNTs (Fig. 2), the trend is nearly linear up to concentrations around 3 mg mL^{-1} . Although plots for the other GO, SWCNT and f-CNT dispersions are not shown here, our experimental data suggest trends similar to Fig. 1 and 2. The deviation of values in Solsperse (Fig. 2b) is higher, compared to the other dispersions, probably because of the increase in the viscosity of the medium. The fitting parameters for all the samples are listed in the ESI† (Section S7). A comment on the accuracy of measurements in the Solsperse medium is also included in Section S8 (ESI†).

It has been commented above that ΔH_1^* is expected to depend only on the liquid medium. Fig. 3 presents the calculated ΔH_1^* values for all the samples. In fact, ΔH_1^* is nearly constant in almost all the dispersions, and in pure water and the 0.5% SDBS medium as well. The values of $\Delta H_1^*/R$ are typically around 1864 K with a deviation of $\pm 54 \text{ K}$. Apparently, the SDBS surfactant does not lead to any substantial change with respect to

pure water, and thus both media can be regarded as equivalent in terms of ΔH_1^* . However, certain deviations from water exist, which should be consequently interpreted as induced changes in the molecular structure of the liquid.

The addition of HSO_3 -MWCNTs to pure water causes a strong decrease in ΔH_1^* . The presence of benzene sulfonic acid groups on MWCNT walls allows the preparation of a stable dispersion in water, but also disrupts the liquid structure. The creation of hydrogen bonds between benzene sulfonic acid groups and water might be an explanation. Curiously, the effect is not observed in the presence of SDBS surfactant. The ΔH_1^* value decreases by 5% after the addition of the CVD-SWCNT sample to the 0.5% SDBS medium. Even though the change is relatively small, it might be related to the extraordinarily high aspect ratio of CVD-SWCNTs (Table 2), which leads to a correction in the flowing energetics of the liquid. An increase of 6% in ΔH_1^* was calculated for the GO-G dispersion in water. The change in ΔH_1^* might reflect the presence of a stabilizing additive. A strong increase in ΔH_1^* up to 2330 K occurs in the 20% Solsperse medium. Acrylic chains of Solsperse greatly influence the associated structure of water, hindering their molecular mobility. The effect is evident despite a decrease in the accuracy of viscosity measurements for the Solsperse medium.

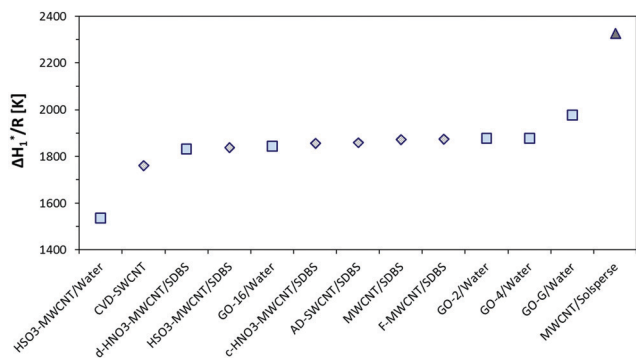


Fig. 3 Enthalpy of activation for viscous flow (ΔH_1^*) for all the CNT and GO systems that are considered in the present work.

Maximum packing fraction

In this section, the viscosity is related to the degree of nanoparticle packing in the fluid. First, the experimental data set is expressed in terms of nanoparticle volume fractions. For that reason, geometrical parameters of CNT and GO units are experimentally measured, so that the particle density (ρ_p) and aspect ratio (r_p) can be calculated for each sample. Later, experimental data are compared with predictive theories that consider suspended particles as ellipsoids of varying r_p .



In terms of the particle volume fraction (ϕ), the equation of Maron and Pierce (eqn (1)) is written as

$$\eta = \frac{K_a^0}{\left(1 - \frac{\phi}{\phi_m}\right)^2} e^{\Delta H_1^*/RT} \quad (4)$$

where ϕ_m is the maximum packing fraction. Mathematically, ϕ_m represents the particle concentration at which the viscosity becomes infinite. The ϕ_m coefficient is related to ε through the particle volumetric density:

$$\phi_m = \frac{1}{\varepsilon \rho_p} L \quad (5)$$

The ρ_p value can be calculated from structural parameters such as the carbon-carbon distance in a graphene sheet, the mass of the carbon atom, the CNT diameter and the number of layers. The ρ_p value for graphene, SWCNTs and MWCNTs is thus calculated taking into account geometrical considerations (ESI,† Section S9).⁵³

The CNTs are considered as rigid cylinders, which is a reasonable assumption according to the literature.^{39,40} An average diameter (D) for each sample is taken from product specifications (Table 2). The presence of chemical groups on MWCNT walls is neglected in terms of D . Length distributions for each CNT and f-CNT sample were analyzed from scanning electron microscopy (SEM) images. The length of a number of CNTs ($N = 290-480$) was directly measured on SEM images (ESI,† Section S10). The length determination is assumed to be valid, although the morphology of the solid on the SEM substrate might be substantially different from that of free nanostructures in the liquid dispersion.⁵⁴ According to the literature,^{38,55} the length histogram was fitted to a log-normal distribution (Section S10, ESI†). The calculated average and median lengths (L_a and L_m respectively) from log-normal distributions are listed in Table 4. The median length was used for the calculation of ρ_p in each case. In order to complete the geometrical description of MWCNTs, the number of layers was deduced from an experimental measurement of the specific surface area ($SSA_{\text{MWCNT}} = 230 \text{ m}^2 \text{ g}^{-1}$) by nitrogen adsorption at 77 K.

Surface area distributions of GO flakes were also determined from SEM images. The calculated average and median surface areas (S_a and S_m respectively) of the four GO samples are listed

Table 4 Characterization details of CNT samples: L_m = median length and L_a = average length

Sample	L_m [nm]	L_a [nm]
MWCNT	439	506
CVD-SWCNT	426	515
AD-SWCNT	315	384
c-HNO ₃ -MWCNT	293	375
d-HNO ₃ -MWCNT	463	585
F-MWCNT	289	389
HSO ₃ -MWCNT	339	453

Table 5 Characterization details of GO flakes: S_m = median surface area; S_a = average surface area; n_L = number of layers; and d_{LL} = distance between layers

Sample	S_m [μm^2]	S_a [μm^2]	n_L	d_{LL} [nm]
GO-2	0.33	0.61	7	0.821
GO-4	0.63	1.17	7	0.833
GO-16	0.86	1.56	9	0.833
GO-G	3.84	7.27	5	0.879

in Table 5. Assuming an approximately elliptic shape, the complete description of flake dimensions requires determining its thickness. Equivalently, the number of layers (n_L) and the distance between layers (d_{LL}) in the flake were calculated from X-ray diffraction (XRD) patterns, applying the Bragg and Scherrer equations (ESI,† Section S11). The calculated d_{LL} and n_L dimensions are listed in Table 5. The flake thickness can be directly calculated as $t_f = (n_L - 1) \cdot d_{LL}$.

Once the ρ_p value was calculated from geometric parameters for all the CNT and GO samples, a correction due to the presence of heteroatoms and impurities was considered. The correction was performed with a parameter that took into account the weight of atoms different from carbon in the structures. However, this additional parameter only causes relatively small changes in the calculated values, which do not alter the qualitative discussion of the results. Similarly, another correction factor was considered for the weight change undergoing if the inner cavity of CNTs was filled by the liquid medium. The filling factor of CNTs, which was not finally applied, does not change the qualitative discussion that is presented next.

Now, theoretical models to predict the relationship of ϕ_m vs. r_p are discussed. In the first approximation, both CNTs and GO are considered ellipsoidal particles with extreme aspect ratios. Ellipsoidal particles are defined by the semi-axes a , b and c . The CNTs may be described as elongated (prolate, $a \gg b = c$) ellipsoids, while GO flakes look like flattened (oblate, $a \ll b = c$) ellipsoids. The coefficient ϕ_m can be related to the extra power or rate of work (\dot{W}_p) required to shear a volume of the nanofluid due to the presence of the particle:⁵⁶

$$\phi_m = \frac{2\eta_0\gamma^2 V_p}{\dot{W}_p} \quad (6)$$

where γ is the gradient of velocity in the fluid (the shear rate) and V_p is the particle volume. In the particular case of spheroids with $b = c$, Jeffery deduced an integral equation for \dot{W}_p , and solved it analytically.⁵⁷ That equation takes into account the rotational period, initial orientation and aspect ratio (r_p) of a spheroidal particle in a flowing medium. The final model equation does not depend on the particle size, but just on the particle shape through the aspect ratio.

In its movement inside the fluid, the particle is free to rotate along its axis of symmetry (in the a direction), which is a natural easy movement. The ϕ_m coefficient has a maximum when all the ellipsoids are oriented in the flow direction. However, a gradient of velocity in the fluid may exist that



additionally forces the particle to tumble end over end. In fact, the velocity of a fluid in a capillary is maximum at the center and minimum at the capillary wall. The end-over-end rotation is much more energy consuming than the simple rotation along the symmetry axis. Moreover, the particle requires a lot of extra space to rotate, which hinders particle packing and leads to a minimum ϕ_m coefficient (Fig. 4). A number of intermediate situations can be considered depending on the angle between the symmetry axis of the spheroid and the flow direction.

Therefore, the theoretical value of ϕ_m for spheroidal particles can be calculated as a function of r_p for different orientations of the particle in the flowing liquid. GO flakes are considered as oblate spheroids ($r_p \ll 1$) and CNTs as prolate spheroids ($r_p \gg 1$). Three hypothetical orientations of the spheroidal particle are here taken into account, giving three ϕ_m functions: maximum, minimum and average. The average ϕ_m coefficient is calculated accounting for all the possible orientations of the particle, which are assumed to be equally probable.

Intrinsic viscosity

The relationship between η_r and ϕ is, according to the model of Maron and Pierce (eqn (2) and (4)), as follows:

$$\eta_r = \left(1 - \frac{\phi}{\phi_m}\right)^{-2} \quad (7)$$

In the dilute regime, eqn (7) can be approximated by the first term of a series expansion:

$$\eta \approx \eta_0(1 + \chi\phi) \quad (8)$$

where χ is known as the intrinsic viscosity. Equivalently,

$$\chi = \lim_{\phi \rightarrow 0} \frac{\eta - \eta_0}{\phi\eta_0} \quad (9)$$

For spherical particles, Einstein demonstrated that $\chi = 5/2$,⁴⁶ while the model of Maron and Pierce leads to $\chi = 2/\phi_m$.⁴⁵ As discussed in the previous section, the dependence of χ on r_p for spheroidal particles can be solved analytically. Since χ is inversely proportional to ϕ_m , the maximum ϕ_m coefficient leads to a minimum χ branch, while the minimum ϕ_m leads to a maximum viscosity, which is associated with the end-over-end rotation of the ellipsoid. When several particles in suspension are considered, the intrinsic viscosity χ is the addition of their individual contributions. Thus, an average was calculated taking into account all the possible orientations of the particles, most specifically averaging contributions from a million ellipsoids. The calculations are presented in Fig. 5, together with our experimental χ values for GO and CNT samples. The viscosities of GO dispersions agree with the maximum calculated for oblate ellipsoids ($r_p < 1$). Therefore, GO flakes, rather than being randomly oriented, rotate end over end, perpendicularly to the fluid moving direction, in their way through the

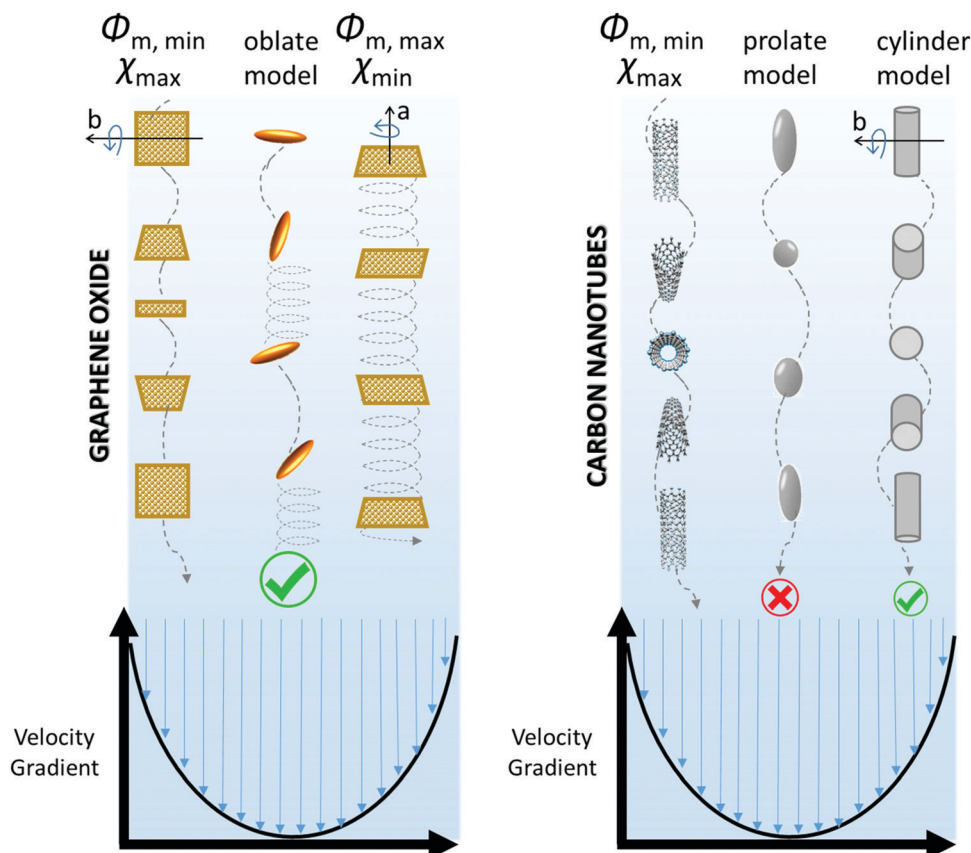


Fig. 4 Schematic explanation of GO and CNT nanoparticle dynamics, according to their viscous behavior in a velocity gradient.



capillary (Fig. 4). On the contrary, experimental data of CNT samples are quite far from values that are predicted for prolate ellipsoids ($r_p > 1$).

The integral equation of \dot{W}_p can be solved analytically for spheroids, but not for cylinders. Therefore, Aragon and Flamik proposed a solution by numerical methods, more specifically applying the boundary element method.⁵⁸ Cylindrical particles are characterized by their aspect ratio $r_p = L/D$. The model of Aragon and Flamik assumes $r_p > 1$ and takes into account both rotational and translational movements of cylindrical particles in the fluid. They considered three possible configurations for cylinder edges: open, closed by spherical caps, and closed by a circle in the rectangular configuration. The results for the three configurations converge at high enough r_p values. The final equation of Aragon and Flamik for χ totally agrees with a solution that was given by Mansfield and Douglas applying another numerical method, namely, the path integration method.⁵⁹

Experimental values of χ for CNTs match quite well the numerical equation for cylindrical particles (Fig. 5). The numerical model of cylinders suitably predicts the behavior of CNTs in the dilute regime. Since the χ vs. r_p branch (solid line in Fig. 5) was first calculated under the assumption that all nanotubes have an identical length L_m , a question remains about the reliability of such a simplifying approximation. Therefore, another calculation was performed taking into account the range of L_m values that comes out from considering a log-normal distribution (dashed line in Fig. 5). The expected deviation in terms of length variability is not qualitatively relevant. Regarding the applicability of numerical methods to solve Eqn 6 for the case of GO as a flat cylinder with $r_p \ll 1$, the solution is not trivial and falls out of the scope of the present work.

From the theoretical $\chi = f(r_p)$ function, the variation of χ with the geometrical parameters of length, diameter and thickness can be calculated for CNT and GO dispersions (Fig. 6).

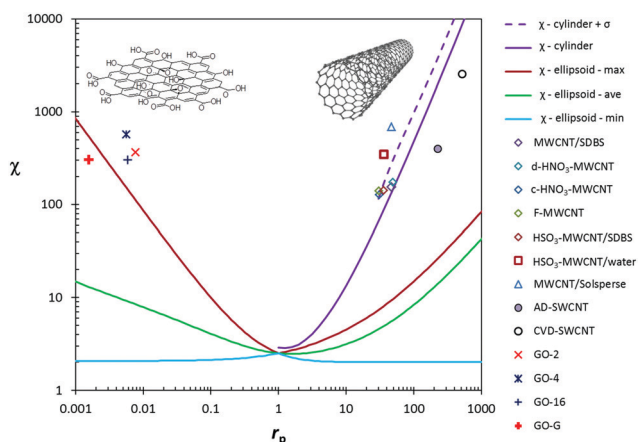


Fig. 5 Intrinsic viscosity (χ) as a function of the aspect ratio (r_p). Experimental data are superimposed to trend lines that were calculated for spheroidal particles⁵⁶ and for cylinders.⁵⁸ The dashed line indicates the correction arising from the statistical dispersion in the CNT length.

Experimental data of GO in water and MWCNTs in SDBS agree well with the theoretical function. Chemical functionalization of MWCNTs, which causes certain changes in the nanotube length, has a minor influence on viscosity. The theoretical χ value for SWCNT dispersions is overestimated with respect to experimental data. The deviation might be associated with a certain degree of bundling and, indirectly, with the presence of impurities in the starting SWCNT powder material.

A step forward in the assessment of the theoretical model is the calculation of the relative viscosity (η_r) as a function of geometrical parameters and the particle concentration. According to the Maron–Pierce model, η_r is related to χ by:

$$\eta_r = \left(1 - \frac{\chi}{2}\phi\right)^{-2} \quad (10)$$

Provided that χ is given by the theoretical model, η_r can be immediately calculated as a function of ϕ . As an example, Fig. 7 shows the result for MWCNT/SDBS and GO-16/water dispersions. For MWCNTs, the calculated curves correspond to $D = 9.5$ nm and log-normal distributions of nanotube lengths with two different median values (L_m). The solid line was obtained with $L_m = 439$ nm, which was determined from SEM images. The dashed line is the best fitting to experimental η_r values, and is associated with $L_m = 364$ nm; thus, the relative deviation is 17% in terms of the

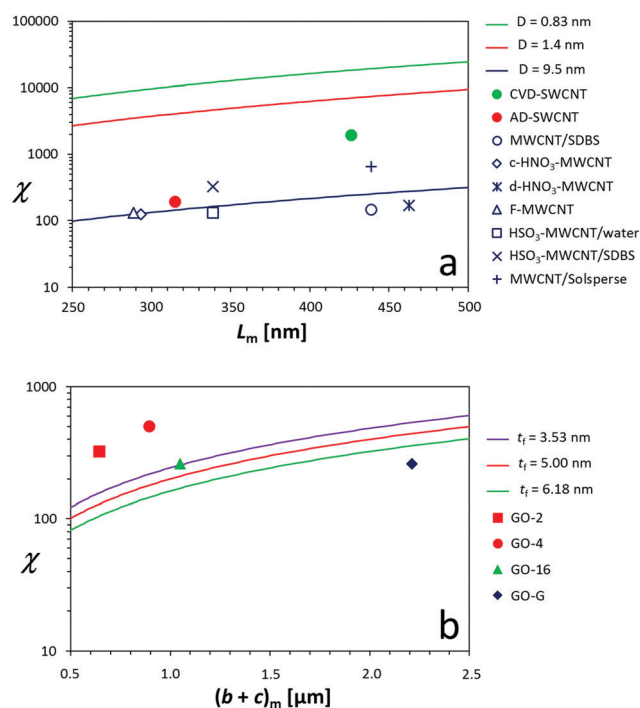


Fig. 6 Experimental and theoretical intrinsic viscosity (χ) as a function of particle shape parameters: (a) experimental data for CNT dispersions as a function of their length (L) and theoretical trends for cylinders,⁵⁸ with diameters $D = 0.83, 1.49$ and 9.5 nm; (b) experimental data for GO dispersions as a function of their flake diameter ($b + c$) and theoretical trends for oblate ellipsoids,⁵⁶ whose minor axis ($2-a$) is equal to the flake thickness $t_f = 3.53, 5.00$ and 6.18 nm. Theoretical calculations correspond to the distributions of multi-disperse particles whose length (CNT) or surface (GO) follows a log-normal distribution.



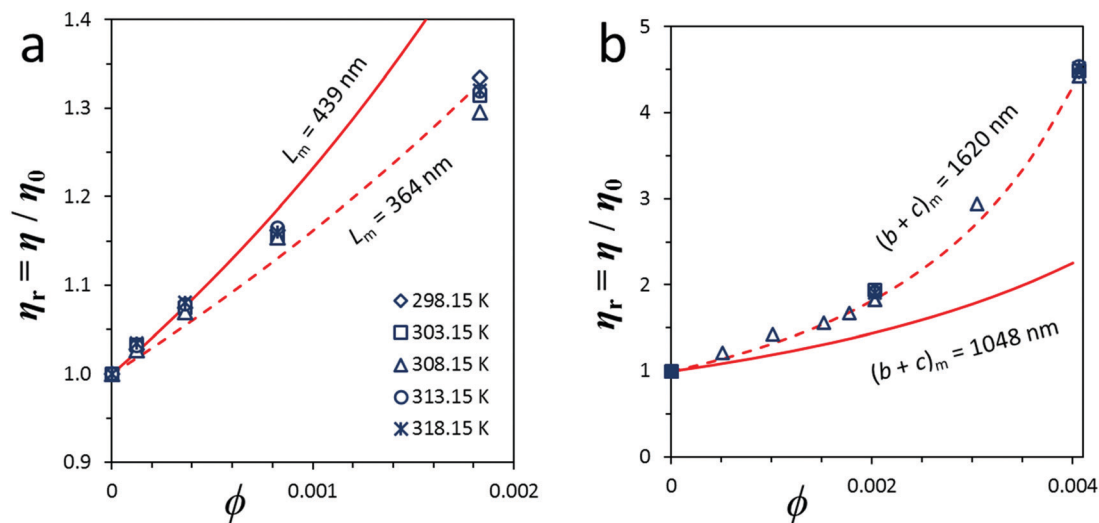


Fig. 7 Relative viscosity (η_r) as a function of the volume fraction (ϕ) for (a) MWCNT/SDBS and (b) GO-16/water dispersions. Experimental data are shown together with theoretical trends for multi-disperse distributions of particles.

experimental L_m parameter. For GO-16, calculated trends come from $t_r = 6.66$ nm and χ values corresponding to log-normal distributions with two different $(b+c)_m$ parameters. The solid line is for the experimental $(b+c)_m$ from SEM images, while the dashed line is the best fitting to experimental η_r data; the relative deviation is 55%. In both MWCNT and GO dispersions, the theory reflects qualitatively the experimental trend.

Finally, it is noteworthy that the absolute viscosity (η) depends on the particle size,¹³ while the intrinsic viscosity (χ value) does not theoretically depend on size. Regarding the intrinsic viscosity, let us recall the universal value $\chi = 5/2$ deduced by Einstein for spheres. Moreover, the more general expressions reported for spheroids and rods depend only on the shape and the aspect ratio.^{56–59} Thus, the relationship $\chi = f(r_p)$ can be regarded as a universal prediction of the viscosity for 1D and 2D nanoparticle dispersions in the dilute regime.

Conclusions

It is demonstrated that the experimental viscosity of dilute CNT and GO dispersions can be fitted to the empirical Maron–Pierce model, as a function of the particle concentration and temperature. The fitting allows the calculation of two parameters, namely, the maximum packing fraction (ϕ_m) and the enthalpy of activation for viscous flow (ΔH_1^*); the ΔH_1^* values are nearly constant for most of the dispersions. Experimental data for GOs and CNTs agree with theoretical predictions by first principles, which evaluate ϕ_m and the intrinsic viscosity (χ) as a function of the aspect ratio (r_p). It is deduced that r_p is the most relevant intrinsic parameter in the studied experimental range. The surface functionality only implies a small correction, as far as it does not induce substantial changes in the r_p value. The comparison between experimental values and theoretical trends indicates that GO flakes behave as ellipsoids that can rotate perpendicularly to the moving direction of the fluid, while the CNT dispersion can be envisaged as cylindrical

particles that rotate in all the possible orientations. The data set collected in the present work is relevant for applications using dilute CNT and GO inks, and contributes to a better understanding of mass transport phenomena and particle dynamics in nanofluids. The viscosity of CNT and GO dispersions can be conceptualized as a function of the aspect ratio, which becomes the most relevant parameter to predict in a universal way the viscosity of dilute nanofluids of any kind of 1D and 2D nanoparticles in dilute suspensions.

Experimental section

Materials and dispersions

MWCNTs (NC7000 grade) were purchased from Nanocyl (Sambreville, Belgium, EU). The chemical functionalization protocols are based on previous works,^{10,60,61} and are detailed in the ESI† (Section S12); different characterization techniques, including infrared spectroscopy, Raman spectroscopy and thermogravimetric analysis, were applied to confirm the functionalization efficiency in those previous works. CVD-SWCNTs (S76 grade) were provided by Sigma Aldrich (Ref. 704121); the material was manufactured by SouthWest Nanotechnologies Inc., now Chasm Advanced Materials (Norman, OK, USA), by the CoMoCAT process. Pristine arc-discharge SWCNTs (AD-SWCNTs) were purchased from Carbon Solutions Inc. (Riverside, CA, USA). Graphite flakes were purchased from Aldrich (ref. 332461).

The GO-2, GO-4 and GO-16 samples were synthesized from graphite flakes by a modified Hummers' method including the following stages:¹¹ (i) treatment in $\text{NaNO}_3/\text{H}_2\text{SO}_4/\text{KMnO}_4$ at 35 °C for 2, 4 or 16 h, respectively; (ii) cooling in 30% H_2O_2 ; (iii) filtration, washing with diluted HCl and drying; and (iv) resuspension in water at 0.5 mg mL⁻¹ and exfoliation in an ultrasound bath for 2 h. The GO-G sample was purchased from Graphenea (San Sebastian, Spain).



The SDBS surfactant was purchased from Sigma Aldrich and acrylic Solsperse™ 27000 was provided by Lubrizol Advanced Materials (Sant Cugat del Vallés, Barcelona, Spain). For the preparation of CNT dispersions, SDBS was dissolved at 0.5 mass/vol% and Solsperse was diluted at 20 vol% in water. The solutions were added to CNT powders and the mixtures were treated with a 400 W Hielscher ultrasonic probe for 1 h at 60% amplitude and a frequency of 0.5 cycles. Dispersions were centrifuged at 4500 rpm for 30 min and the sediment was discarded.

Viscosity measurements

Viscosity measurements were performed in Ubbelohde viscosimeters (types 0a and I) (SI Analytics). The correction of kinematic energy (Hagenbach correction) was applied when needed. The kinematic viscosity was calculated as $\nu = K \cdot t$, where K is the constant of the viscosimeter ($0.004955 \text{ mm}^2 \text{ s}^{-2}$ for the 0a and $0.009596 \text{ mm}^2 \text{ s}^{-2}$ for the I viscosimeter). The temperature was controlled at ± 0.01 degree with a Julabo thermostatic bath provided with an external temperature probe. The temperature calibration was performed by measuring the viscosity of deionized water. Each viscosity determination was repeated at least three times, achieving deviations of $< 1\%$, most typically around 0.1% .

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work has been funded by the Spanish MINEICO under the project ENE 2016-79282-C5-1-R (AEI/FEDER, UE), the Government of Aragón (Grupo reconocido T03-20R) and associated EU Regional Development Funds (DGA/FEDER, UE). Special thanks are directed to Ms A. Santidrián for the preparation of HSO₃-MWCNT, Mr D. Velilla for zeta-size measurements, Dr E. Istif for GO-G sample characterization, Mr S. Escuriet from Lubrizol Advanced Materials for providing the Solsperse™ 27000 surfactant, the Analysis Service at Instituto de Carboquímica ICB-CSIC, the Advanced Microscopy Laboratory (LMA) at Instituto de Nanociencia de Aragón, and the Electron Microscopy Service at the University of Zaragoza. SVR thanks Spanish MINEICO for her PhD grant (BES2014-068727 and associated EU Social Funds). J. M. G.-D. greatly acknowledges the Spanish Ministry of Science, Innovation and Universities (MICINN, formerly MINEICO) for his “Juan de la Cierva – incorporación” grant (Ref. UCI-2016-27789).

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