



# Hydrophosphination using $[\text{GeCl}\{\text{N}(\text{SiMe}_3)_2\}_3]$ as a pre-catalyst†

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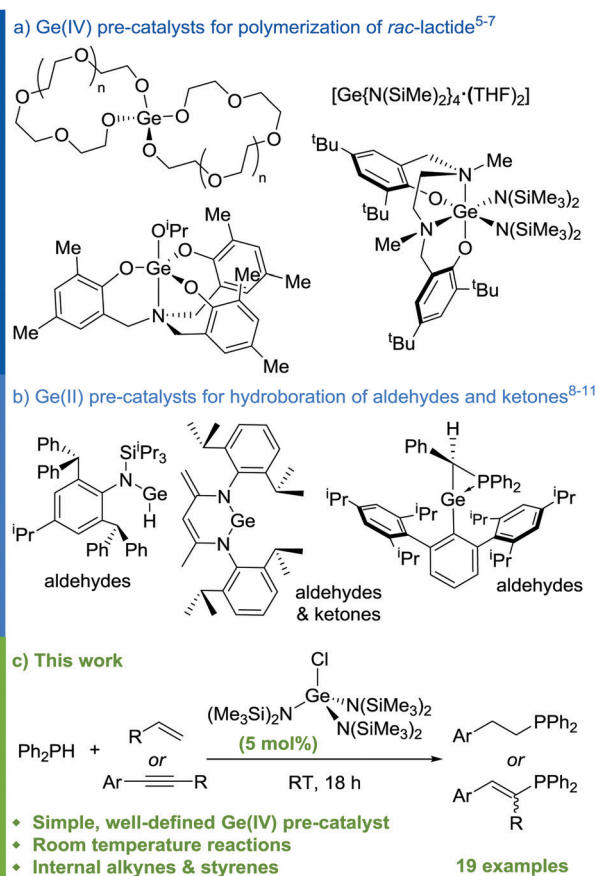
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Transformations catalyzed by germanium are scarce, with examples mainly limited to widely catalyzed processes such as polymerisation of lactide and hydroboration of carbonyls. Reported is the first example of hydrophosphination using a germanium pre-catalyst, yielding anti-Markovnikov products when diphenylphosphine is reacted with styrenes or internal alkynes at room temperature.

Over the last two decades, the use of main group compounds as catalysts for organic reactions has seen a vast increase in focus.<sup>1–3</sup> Amongst this work, significant progress has been made in the field of germanium catalysis. A fundamental challenge lies in unlocking the catalytic potential of this group 14 element, which is far behind tin in terms of its exploitation in catalysis, and such work may prove instrumental to furthering understanding and harbouring new discoveries in main group catalysis. It is not surprising, based on the ability of Sn(IV) complexes to act as polymerization catalysts,<sup>4</sup> that initial disclosures in the literature demonstrated the ability of Ge(IV) complexes to act as initiators for the ring-opening polymerization of lactide (Scheme 1a).<sup>5–7</sup> Subsequently, there have been numerous reports on low-coordinate Ge(II) species acting as efficient pre-catalysts for the hydroboration of carbonyls (Scheme 1b).<sup>8–11</sup> This reactivity has also been extended to the catalytic cyanosilylation of aldehydes.<sup>12,13</sup>

Despite this progress, lactide polymerization and carbonyl hydrofunctionalization are commonly reported catalytic transformations,<sup>14,15</sup> therefore a desire remains to uncover use for germanium (both Ge(II) and Ge(IV)) in the catalysis of a wider range of transformations.<sup>16</sup> It is clear from the current literature that, in the presence of at least one labile co-ligand, Ge(IV) complexes can catalyze non-redox,  $\sigma$ -bond metathesis type transformations such as lactide polymerization. This indicates



Scheme 1 Examples of germanium catalysis reported in the literature to date.

that opportunities exist for Ge(IV) catalysts, particularly in hydrofunctionalization reactions with pertinent substrates.

The hydrophosphination (HP) of unsaturated hydrocarbons is an atom economical process of forming P–C bonds.<sup>17</sup> Using a main group element to transform another is an area of growing interest, amongst which examples of main group catalyzed HP

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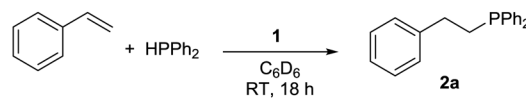


have emerged.<sup>18</sup> Reaction mechanisms for these transformations involving main group, and lanthanide,<sup>19–23</sup> complexes are often postulated to proceed *via*  $\sigma$ -bond metathesis type steps at the catalyst center,<sup>24</sup> leading us to form the hypothesis that a Ge catalyst may have success in this area. To the best of our knowledge, no examples of Ge-catalyzed HP have been reported. However, the heavier group 14 element, Sn, has been shown to be active catalytically in the HP of alkenes and alkynes by Waterman and co-workers, who, in 2014, reported a modest scope of HP reactions using PhPH<sub>2</sub>.<sup>25</sup> Subsequent studies on the same Sn(IV) system using Ph<sub>2</sub>PH in 2016 expanded the scope to further substrates.<sup>26</sup> However, these transformations required an atmosphere of H<sub>2</sub> due to the catalyst's propensity to undergo homodehydrocoupling of the phosphine substrate over the desired HP reaction. This, along with the elevated temperatures necessary (65 °C) and high catalyst loading (10 mol%), makes the reaction only marginally more competitive than the analogous catalyst-free, thermal process.<sup>27,28</sup> We were therefore keen to investigate if a more active catalyst could be found within group 14 in the form of a Ge(IV) species. Rather than start our studies with an over-engineered catalyst structure, we looked to initiate our studies with an operationally simple motif; understand the catalytic process and thus be in a strong position to build in ligand design at a later stage, based on understanding. We herein present the first HP of alkenes using a structurally simple Ge(IV) pre-catalyst under mild conditions (Scheme 1c).

Reaction of GeCl<sub>4</sub> with 10 equiv. of Na(HMDS) yields [GeCl{N(SiMe<sub>3</sub>)<sub>2</sub>}<sub>3</sub>], **1**, in good yield (63%, Scheme 2), the structure of which was confirmed by X-ray diffraction.<sup>29</sup> In an analogous reaction, Thomas and co-workers propose the formation of tetra(amido) THF adduct [Ge(HMDS)<sub>4</sub>(THF)<sub>2</sub>].<sup>7</sup> in our synthesis <sup>1</sup>H DOSY analysis indicates lack of persistent THF coordination in solution.<sup>29</sup>

Pleasingly, **1** is an efficient pre-catalyst for the HP of styrene with diphenylphosphine (Ph<sub>2</sub>PH) and provides 91% conversion

**Table 1** Reaction optimization for the hydrophosphination of styrene using pre-catalyst **1**<sup>a</sup>



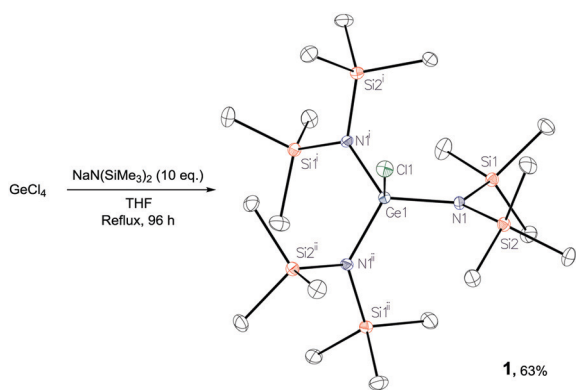
Entry	Catalyst loading/mol%	Solvent	Spectroscopic yield <sup>b</sup> /%
1	5	C <sub>6</sub> D <sub>6</sub>	91
2	2.5	C <sub>6</sub> D <sub>6</sub>	81
3	1	C <sub>6</sub> D <sub>6</sub>	38
4 <sup>c</sup>	5	THF	99
5 <sup>c</sup>	5	C <sub>6</sub> D <sub>6</sub>	94

<sup>a</sup> Conditions: styrene (0.5 mmol, 57  $\mu$ L), Ph<sub>2</sub>PH (0.5 mmol, 87  $\mu$ L), **1** (5 mol%, 0.025 mmol, 14.7 mg), C<sub>6</sub>D<sub>6</sub> (0.6 mL). <sup>b</sup> Spectroscopic yield calculated *via* integration against 1,3,5-trimethoxybenzene internal standard. <sup>c</sup> 0.6 mmol styrene.

to the anti-Markovnikov HP product **2a** after 18 hours at RT with a catalyst loading of 5 mol%, using C<sub>6</sub>D<sub>6</sub> as solvent (Table 1, entry 1). Only trace quantities of homodehydrocoupling by-product P<sub>2</sub>Ph<sub>4</sub> is formed, therefore any concerns over a competing reaction of this type prove insignificant. Markovnikov product is never observed. Catalyst loadings of 2.5 mol% and 1 mol% result in the reduction of spectroscopic yield to 81% and 38% respectively (entries 2 and 3). Notably, no reaction is observed when GeCl<sub>4</sub> is used as the catalyst. Change of solvent to THF gives similar conversion compared to benzene (entry 4). Other solvents (MeCN (*vide infra*), DCM, CDCl<sub>3</sub>) show reactivity with **1** and are therefore not suitable for the HP reaction.<sup>29</sup> A modest excess of styrene (1.2 equiv.) gives 94% conversion in an inexpensive NMR solvent (entry 5).

With the optimized conditions in hand (Table 1, entry 5) we then set out to investigate the substrate scope for the reaction. The reaction proceeds with a wide range of styrenes (Table 2). Biphenyl-substituted phosphine **2b** is formed in excellent yield (entry 2). *p*-Me-, *p*-MeO- and *p*-CF<sub>3</sub>-substituted phosphines (**2c–2e**) are formed in reduced conversions (entries 3 to 5). Halide substituents are well tolerated, with phosphine products **2f** to **2h** formed in good to moderate conversion (entries 6 to 8). *m*- and *o*-substituted phosphines (**2i** to **2m**) are also formed in good yield (entries 9 to 13).

Naphthyl substituted phosphine **2n** is formed in excellent yield (entry 14) and 1,1-diphenylethylene performs comparably well, albeit under slightly more forcing conditions (50 °C, 24 h, entry 15). Phenyl acetylene does not react, but internal alkynes (entries 16 to 19) react at RT: conditions that are highly competitive compared to the leading literature, for example Schmidt's RT, La-catalyzed process in pyridine solvent<sup>30</sup> and Cui's Ca- and Y-mediated chemistry that uses temperatures ranging from RT–75 °C.<sup>31</sup> Unfortunately, reactions of 2- and 4-vinylpyridine and acrylonitrile with **1** all result in the formation of unidentified polymers irrespective of order of addition of Ph<sub>2</sub>PH. Methyl acrylate, butyl acrylate, *cis*- and *trans*-stilbene, along with  $\beta$ -methyl styrene do not react. HP reactions using unactivated alkenes (*e.g.* cyclohexene, allyl benzene) result in no conversion while no stoichiometric reaction of **1**



**Scheme 2** Synthesis and single crystal X-ray structure of **1** displaying the inherent C<sub>3v</sub> symmetry of the molecule. Ellipsoids are represented at 30% probability and hydrogen atoms have been omitted for clarity. Symmetry operations: <sup>i</sup> 1 + y – x, 1 – x, z; <sup>ii</sup> –1 – y, x – y, z. Bond distance (Å) and angles (°): Ge1–Cl1, 2.2154(11); Ge1–N1, 1.8542(16); Cl–Ge–N1 102.12(6); N1–Ge–N1<sup>i</sup> 115.71(4).



Table 2 Substrate scope for HP using pre-catalyst **1**<sup>a</sup>

Entry	Alkene	Product	Spec. yield <sup>b</sup> (%)
1	R = H	<b>2a</b>	94
2	4-Ph	<b>2b</b>	98
3	4-Me	<b>2c</b>	94 <sup>c</sup>
4	4-OMe	<b>2d</b>	56 <sup>c</sup>
5	4-CF <sub>3</sub>	<b>2e</b>	83 <sup>c</sup>
6	4-F	<b>2f</b>	76 <sup>c</sup>
7	4-Cl	<b>2g</b>	94
8	4-Br	<b>2h</b>	64
9	3-Me	<b>2i</b>	92
10	3-Br	<b>2j</b>	82
11	2-Me	<b>2k</b>	89
12	2-OMe	<b>2l</b>	89
13	2-Br	<b>2m</b>	84
14			<b>2n</b> 90
15			<b>2o</b> 70 <sup>c</sup>
16			<b>2p</b> 84 <sup>d</sup>
17			<b>2q</b> 3.9:1, E:Z
18			<b>2r</b> 81 <sup>e</sup>
19			<b>2s</b> 41 <sup>f</sup>

<sup>a</sup> Conditions: styrene (0.6 mmol), Ph<sub>2</sub>PH (0.5 mmol, 87 μL), **1** (5 mol%, 0.025 mmol, 14.7 mg), C<sub>6</sub>D<sub>6</sub> (0.6 mL), 18 h, RT. <sup>b</sup> Spectroscopic yield calculated *via* integration of product peak against a trimethoxybenzene internal standard. <sup>c</sup> 50 °C, 24 h. <sup>d</sup> RT, 96 h. <sup>e</sup> RT, 48 h. Spec. yield and E:Z calculated *via* <sup>31</sup>P inverse gated NMR spectroscopy against a P(OEt)<sub>3</sub> internal standard. <sup>f</sup> RT, 72 h. Spec. yield and E:Z ratio calculated *via* <sup>31</sup>P inverse gated NMR spectroscopy against a P(OEt)<sub>3</sub> internal standard.

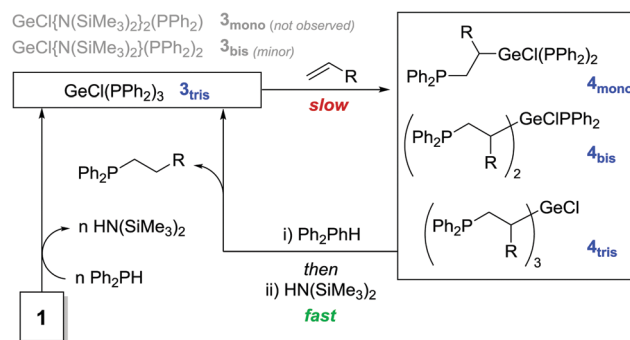
and HPCy<sub>2</sub>, or catalytic HP of styrene using **1** and HPCy<sub>2</sub>, is observed.<sup>29</sup> However, scale-up synthesis of **2a** is successful: using 5 mmol Ph<sub>2</sub>PH (0.93 g/0.87 mL), 6 mmol styrene, 2.5 mol% Ge, at RT gives 94% after 24 h. Catalyst robustness was tested through an iterative addition protocol: 0.5 mmol Ph<sub>2</sub>PH, 0.6 mmol styrene, 5 mol% **1**, were reacted at RT in C<sub>6</sub>D<sub>6</sub> until complete conversion of Ph<sub>2</sub>PH was observed. Further equivalents of Ph<sub>2</sub>PH and styrene were added sequentially until six additions had been achieved (3 mmol Ph<sub>2</sub>PH added) giving 87% spectroscopic yield.<sup>29</sup>

Focus was then turned to investigating the mechanism of the reaction. A kinetic study of the reaction to form **2a** was undertaken *via* variable time normalization analysis,<sup>32</sup> implicates the reaction is approximately first order with respect to **1** and Ph<sub>2</sub>PH and a more complex relationship with styrene (≤1, ≥0.5), in-line with a complex reaction where multiple

species can propagate from one Ge-center.<sup>29</sup> There is no reaction when a stoichiometric quantity of styrene is added to **1**. Conversely, reaction of Ph<sub>2</sub>PH with **1** results in the formation of a bright yellow color, which is also observed during catalysis. <sup>1</sup>H DOSY analysis of this species shows fast diffusing HN(SiMe<sub>3</sub>)<sub>2</sub> peaks and slower diffusing aryl resonances, consistent with protonation of Ge-bound N(SiMe<sub>3</sub>)<sub>2</sub> by HPPH<sub>2</sub>.<sup>29</sup> Monitoring catalysis (15 mol% **1**) as a function of time by <sup>31</sup>P NMR spectroscopy shows the Ph<sub>2</sub>PH doublet decreasing in intensity as product **2a**, and a singlet at -36 ppm, increase. The signal at -36 ppm reaches a maximum when all Ph<sub>2</sub>PH is consumed (~60% **2a** formed), and then shows a slow decrease of this signal into **2a** for the remainder of the reaction. There is also a very small singlet at -27 ppm that decreases in intensity slowly over the course of the reaction. Our data is consistent with bis- and tris-phosphido complexes (-27 and -36 ppm respectively) due to the loss of P-H, the fact that integration of the -36 ppm signal is consistent with 3 equiv. Ph<sub>2</sub>PH being sequestered from solution, alongside <sup>31</sup>P NMR signals being more shielded than [GeCl{N(SiMe<sub>3</sub>)<sub>2</sub>}<sub>2</sub>(PPh<sub>2</sub>)] (not observed).<sup>33</sup> Noteworthy is that a stabilized radical, (\*Ge{N(SiMe<sub>3</sub>)<sub>2</sub>}<sub>3</sub>), from photochemical disproportionation of [Ge{N(SiMe<sub>3</sub>)<sub>2</sub>}<sub>2</sub>] has been reported. However, we have conclusively ruled-out radicals from this reaction.<sup>29</sup>

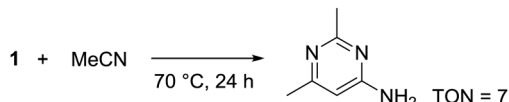
Combining these experimental observations, we propose that the tris-phosphido species (**3<sub>tris</sub>**) is the principal active catalyst for this transformation, and therefore postulate a mechanism based upon this (Scheme 3). After the formation of on-cycle **3<sub>tris</sub>**, styrene may then insert into the Ge-P bond in a regioselective manner to form an alkyl intermediate (**4<sub>mono/bis/tris</sub>**). We have discounted this step occurring *via* attack of the β-position of styrene by the P-atom as no telomerization side products are observed.<sup>24</sup> We propose 1,2-insertion to take place in a non-stepwise manner in accordance with the reaction being non-integer order with respect to styrene. From short-lived alkyl intermediates **4**, stepwise protonolysis accounts for the approximate integer relationship in Ph<sub>2</sub>PH. This step liberates the product and reforms active species **3**.

During solvent screening, we also observed that **1** could catalytically induce the trimerization of MeCN to 2,6-dimethylpyrimidin-4-amine (Scheme 4). Base-mediated



Scheme 3 Proposed method by which **1** catalyzes the HP of styrenes. **3<sub>bis</sub>** is only observed in trace quantities.



Scheme 4 Trimerization of acetonitrile catalyzed by **1**.

cascade reactions<sup>34,35</sup> and an iridium hydride catalyst<sup>36</sup> have previously been used to obtain the same product under more forcing conditions (*e.g.* Murahashi's Ir-catalyzed method uses IrH<sub>5</sub>(P<sup>t</sup>Pr<sub>3</sub>)<sub>2</sub> at 140 °C for 12 h giving TON = 12). This reactivity represents a new avenue of reactions this germanium species is capable of catalyzing and will be explored in due course.

In summary, we have disclosed a rare example of a Ge(IV)-catalyzed (non-polymerization) reaction. An operationally simple Ge(IV) tris(amido) complex has shown excellent ability as a pre-catalyst for the hydrophosphination of a range of styrenes and internal alkynes. The transformations proceed at room temperature over a time period competitive with transition metal catalyzed analogues. Preliminary mechanistic studies implicate the formation of a Ge-(tris)phosphido complex as the active catalyst and a redox-neutral mechanistic pathway was postulated. We hope that this work, using such a simple pre-catalyst, serves as a platform for future research in germanium-catalyzed hydrofunctionalization reactions.

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## Conflicts of interest

There are no conflicts to declare.

## Notes and references

- 1 P. P. Power, *Nature*, 2010, **463**, 171–177.
- 2 T. J. Hadlington, M. Driess and C. Jones, *Chem. Soc. Rev.*, 2018, **47**, 4176–4197.
- 3 C. Weetman and S. Inoue, *ChemCatChem*, 2018, **10**, 4213–4228.
- 4 J. Wu, T.-L. Yu, C.-T. Chen and C.-C. Lin, *Coord. Chem. Rev.*, 2006, **250**, 602–626.
- 5 A. Finne, Reema and A.-C. Albertsson, *J. Polym. Sci., Part A: Polym. Chem.*, 2003, **41**, 3074–3082.
- 6 A. J. Chmura, C. J. Chuck, M. G. Davidson, M. D. Jones, M. D. Lunn, S. D. Bull and M. F. Mahon, *Angew. Chem. Int. Ed.*, 2007, **46**, 2280–2283.
- 7 J. Guo, P. Haquette, J. Martin, K. Salim and C. M. Thomas, *Angew. Chem. Int. Ed.*, 2013, **52**, 13584–13587.
- 8 T. J. Hadlington, M. Hermann, G. Frenking and C. Jones, *J. Am. Chem. Soc.*, 2014, **136**, 3028–3031.
- 9 Y. Wu, C. Shan, Y. Sun, P. Chen, J. Ying, J. Zhu, L. Liu and Y. Zhao, *Chem. Commun.*, 2016, **52**, 13799–13802.
- 10 J. Schneider, C. P. Sindlinger, S. M. Freitag, H. Schubert and L. Wesemann, *Angew. Chem. Int. Ed.*, 2017, **56**, 333–337.
- 11 S. Sinhababu, D. Singh, M. K. Sharma, R. K. Siwath, P. Mahawar and S. Nagendran, *Dalton Trans.*, 2019, **48**, 4094–4100.
- 12 R. K. Siwath and S. Nagendran, *Chem. – Eur. J.*, 2014, **20**, 13551–13556.
- 13 R. Dasgupta, S. Das, S. Hiwase, S. K. Pati and S. Khan, *Organometallics*, 2019, **38**, 1429–1435.
- 14 O. Dechy-Cabaret, B. Martin-Vaca and D. Bourissou, *Chem. Rev.*, 2004, **104**, 6147–6176.
- 15 C. C. Chong and R. Kinjo, *ACS Catal.*, 2015, **5**, 3238–3259.
- 16 T. Sugahara, J. D. Guo, T. Sasamori, S. Nagase and N. Tokitoh, *Angew. Chem. Int. Ed.*, 2018, **57**, 3499–3503.
- 17 I. Wauters, W. Debrouwer and C. V. Stevens, *Beilstein J. Org. Chem.*, 2014, **10**, 1064–1096.
- 18 M. S. Hill, D. J. Liptrot and C. Weetman, *Chem. Soc. Rev.*, 2016, **45**, 972–988.
- 19 M. R. Douglass, C. L. Stern and T. J. Marks, *J. Am. Chem. Soc.*, 2001, **123**, 10221–10238.
- 20 A. M. Kawaoka, M. R. Douglass and T. J. Marks, *Organometallics*, 2003, **22**, 4630–4632.
- 21 I. V. Basalov, S. C. Roşca, D. M. Lyubov, A. N. Selikhov, G. K. Fukin, Y. Sarazin, J.-F. Carpentier and A. A. Trifonov, *Inorg. Chem.*, 2014, **53**, 1654–1661.
- 22 I. V. Basalov, V. Dorcet, G. K. Fukin, J.-F. Carpentier, Y. Sarazin and A. A. Trifonov, *Chem. – Eur. J.*, 2015, **21**, 6033–6036.
- 23 J. Yuan, H. Hu and C. Cui, *Chem. – Eur. J.*, 2016, **22**, 5778–5785.
- 24 L. Rosenberg, *ACS Catal.*, 2013, **3**, 2845–2855.
- 25 K. A. Erickson, L. S. H. Dixon, D. S. Wright and R. Waterman, *Inorg. Chim. Acta*, 2014, **422**, 141–145.
- 26 J. P. W. Stelmach, C. A. Bange and R. Waterman, *Dalton Trans.*, 2016, **45**, 6204–6209.
- 27 F. Alonso, Y. Moglie, G. Radivoy and M. Yus, *Green Chem.*, 2012, **14**, 2699–2702.
- 28 Y. Moglie, M. J. Gonzalez-Soria, I. Martin-Garcia, G. Radivoy and F. Alonso, *Green Chem.*, 2016, **18**, 4896–4907.
- 29 See ESI†.
- 30 M. M. I. Basiouny, D. A. Dollard and J. A. R. Schmidt, *ACS Catal.*, 2019, **9**, 7143–7153.
- 31 H. Hu and C. Cui, *Organometallics*, 2012, **31**, 1208–1211.
- 32 J. Burés, *Angew. Chem. Int. Ed.*, 2016, **55**, 16084–16087.
- 33 J. K. West and L. Stahl, *Organometallics*, 2012, **31**, 2042–2052.
- 34 R. Bossio, S. Marcaccini, V. Parrini and R. Pepino, *J. Heterocycl. Chem.*, 1986, **23**, 889–891.
- 35 A. Verma, S. Jana, C. Durga Prasad, A. Yadav and S. Kumar, *Chem. Commun.*, 2016, **52**, 4179–4182.
- 36 H. Takaya, T. Naota and S.-I. Murahashi, *J. Am. Chem. Soc.*, 1998, **120**, 4244–4245.

