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Studying the cellular distribution of highly phototoxic platinumated metalloporphyrins using isotope labelling†

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Novel tetraplatinated metalloporphyrin-based photosensitizers (PSs) are reported, which show excellent phototoxic indexes (PIs) up to 5800 against HeLa cells, which is, to the best of our knowledge, the highest value reported for any porphyrin so far. Furthermore, ⁶⁷Zn isotope labelling allowed the determination of the ratio of zinc to platinum inside the cells using ICP-MS.

PDT is a clinically approved and minimally invasive therapeutic treatment for dermatological, ophthalmic, and infectious diseases, as well as cancer, which involves the application of a photosensitizer (PS) that ideally localizes disease-specifically. The application is followed by low energy irradiation to activate the PS leading to the generation of reactive oxygen species (ROS) which in turn induce cell death.¹ A number of publications have reported about PSs conjugated with a chemotoxic unit, *e.g.*, a cisplatin analogue.² While this approach normally renders the conjugate more toxic,³ it might diminish one of the main advantages of PDT, namely the special selectivity by the applied light source.

Previously, we have reported a series of easily accessible tetrapyrrolyl-porphyrins (**tPt-H₂4PyP**, **cPt-H₂4PyP**, **dPt-H₂4PyP**) that were coordinated by four platinum(II) complexes to yield highly phototoxic agents (Fig. 1). These compounds displayed a

light toxicity down to 19 nM and at the same time low dark toxicities of around 45 μM in human cancer cell lines.⁴ The ratio of dark to light toxicity is referred to as the phototoxic index (PI), and it should be as high as possible. Our systems displayed PI values higher than 1000 in HeLa cells after being excited with light of a wavelength of 420 nm. These PI values are better than currently clinically used second generation PSs that display a phototoxicity of > 260.⁵

Based on our previous studies,^{1c,4} we investigated a set of important follow-up questions:

(1) How does the insertion of either zinc or copper^{1c} modulate the phototoxicity of **tPt-H₂4PyP**, **cPt-H₂4PyP** and **dPt-H₂4PyP**? The singlet oxygen quantum yield of porphyrins was reported to increase⁶ or not to change at all⁷ after insertion of zinc into the macrocycle. In another case, the metal-free porphyrin generated mainly free hydroxy radicals, which were *in vitro* more cytotoxic than the singlet oxygen produced by the zinc porphyrin.⁸

(2) What is the fate of the complex **tPt-H₂4PyP** in the cell?

(3) What is the reason for the low dark toxicities of our platinum-porphyrin conjugates?

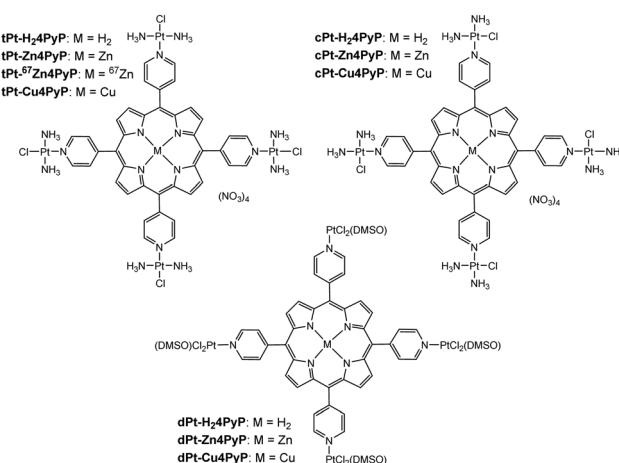


Fig. 1 Studied tetraplatinated porphyrins.

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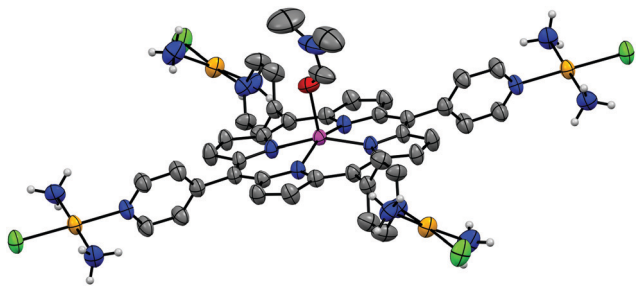


Fig. 2 Ellipsoidal plot of the crystal structure of $[trans\text{-PtCl}(\text{NH}_3)_2]_4\text{-5,10,15,20-tetra}(4'\text{-pyridyl})\text{-zinc(II)porphyrin tetraphenylborate}\cdot(\text{DMF})_7$. Hydrogen atoms attached to a carbon atom, non-coordinated anions and DMF molecules not coordinated to the zinc metal center were omitted for clarity.

(4) What is the origin of the slow reaction of **tPt-H₂4PyP** with guanosine?

To study the effect of a metal inside the center of the porphyrins **tPt-H₂4PyP**, **cPt-H₂4PyP** and **dPt-H₂4PyP**, the novel complexes **tPt-Zn4PyP**, **tPt-Cu4PyP**, **cPt-Zn4PyP**, **cPt-Cu4PyP**, **dPt-Zn4PyP** and **dPt-Cu4PyP** were synthesized following our previously published procedures.⁴ Subsequently, these complexes were characterized using ¹H- and ¹⁹⁵Pt-NMR, IR, UV-Vis, MS and elemental analysis (ESI⁺). Additionally, the cationic unit of **tPt-Zn4PyP** could be crystallized as the tetraphenylborate salt using the under-oil technique (ESI⁺) and was studied using X-ray analysis (Fig. 2). We were particularly pleased to obtain this crystal structure, as it is the first example of expanding our robotic, aqueous crystallization trials⁹ towards organic solvents. The zinc metal center is in a square pyramidal geometry with one DMF molecule as the fifth ligand. This is only the third crystallographic report about a porphyrin with an exocyclic, non-organometallically bound platinum within the porphyrin plane.¹⁰

The compounds were then tested in the cancerous HeLa cell line and the non-cancerous MRC-5 cell line composed of fibroblasts derived from lung tissue to assess their anti-proliferative properties. The results of the cytotoxicity studies of the compounds are presented in Table 1. The insertion of zinc into the complex **tPt-H₂4PyP** to yield the zinc complex **tPt-Zn4PyP** improves the phototoxicity and lowers the dark toxicity at the same time. Together these two effects improve the PI by almost a factor of five. These results can be attributed to the higher singlet oxygen quantum yield (determined as described in ref. 11) of **tPt-Zn4PyP** compared to **tPt-H₂4PyP** (Table S3, ESI⁺) and are in line with previous reports about other (metal-)porphyrin containing photosensitizers.^{6a,b,d}

However, the zinc complexes **cPt-Zn4PyP** and **dPt-Zn4PyP** are less phototoxic than their corresponding metal-free porphyrins **cPt-H₂4PyP** and **dPt-H₂4PyP**. This difference in behaviour can be explained for **cPt-Zn4PyP**. After the insertion of zinc, the complex becomes rather photosensitive, which leads to photobleaching under irradiation and therefore to a lower singlet oxygen quantum yield (Table S3, ESI⁺). The higher phototoxic indexes of complexes **tPt-H₂4PyP**, **tPt-Zn4PyP** and **tPt-Cu4PyP** compared with **cPt-H₂4PyP**, **cPt-Zn4PyP** and **cPt-Cu4PyP**, respectively, are further explained below. Furthermore, **tPt-Cu4PyP** is, to the

Table 1 Anti-proliferative effects of the various compounds on non-cancerous MRC5 and cancerous HeLa cells in the dark and upon irradiation

Compound	IC ₅₀ MRC5 dark (μM)	IC ₅₀ HeLa dark (μM)	IC ₅₀ HeLa 420 nm (μM)	PI (HeLa)	Ref.
Photofrin	n. d.	> 41	4.3 ^{±0.2}	> 9.5	t. w.
Transplatin (tPt)	86 ^{±8}	~ 130	n. d.	n. d.	12/13
tPt-H₂4PyP	93.4 ^{±6.5}	44.9 ^{±8.0}	0.037 ^{±0.02}	1210	4
tPt-Zn4PyP	> 100	> 100	0.017 ^{±0.004}	> 5882	t. w.
tPt-Cu4PyP	82.5 ^{±17.5}	56.4 ^{±6.5}	14.1 ^{±6.3}	4.0	t. w.
Cisplatin (cPt)	7.9 ^{±1.2}	11.5 ^{±2.9}	22.3 ^{±5.7}	0.52	4
cPt + H₂4PyP (1 : 1)	> 100	> 100	5.78 ^{±1.9}	> 17	4
cPt-H₂4PyP	50.2 ^{±0.6}	35.4 ^{±4.4}	0.054 ^{±0.01}	655	4
cPt-Zn4PyP	65.0 ^{±0.6}	17.1 ^{±5.1}	0.25 ^{±0.02}	68.4	t. w.
cPt-Cu4PyP	> 100	13.9 ^{±1.6}	> 100	< 0.14	t. w.
cisPtDMSO (dPt)	> 100	> 100	> 100	n. a.	t. w.
dPt-H₂4PyP	> 100	> 100	0.15 ^{±0.02}	> 680	4
dPt-Zn4PyP	> 100	> 100	1.28 ^{±0.4}	> 78	t. w.
dPt-Cu4PyP	> 100	61.4 ^{±1.5}	> 100	< 0.61	t. w.

PI = phototoxic index, n. d. = not determined, t. w. = this work, n. a. = not applicable.

best of our knowledge, only the second ever reported copper containing PS,^{1c} which is phototoxic, as normally copper porphyrinoids are non-phototoxic. The DMSO containing complexes **dPt-H₂4PyP** and **dPt-Zn4PyP** have comparable PIs to the cisplatin derived complexes **cPt-H₂4PyP** and **cPt-Zn4PyP**. For a careful discussion on the influence of DMSO upon any platinum containing drug, the reader is directed to the excellent study by Gottesman.¹⁴ The singlet oxygen quantum yields (Φ_{Δ}) of **tPt-H₂4PyP**, **cPt-H₂4PyP**, **tPt-Zn4PyP** and **cPt-Zn4PyP** are shown in Table S3 (ESI⁺).

As previously reported by us, **tPt-H₂4PyP** reacts extremely slowly with guanosine, as the platination of N7 progressed to less than 50% completion within 10 days.⁴ We speculate that the high charge (4+) of the cation of **tPt-H₂4PyP** might be responsible for this slow reaction, since the platinum chloride bond must be hydrolysed before the platinum can react with N7 of guanosine. The hydrolysis yields an intermediate that possesses an even higher charge of +5 (Fig. S1, ESI⁺). In order to test this hypothesis, we synthesized *trans*-[Pt(NH₃)₂(pyridine)Cl]₂NO₃ (**tPt-Py**)¹⁵ and repeated the earlier mentioned reaction with guanosine (Fig. S2, ESI⁺). The reaction of **tPt-Py** and one equivalent of guanosine was monitored using ¹H-NMR spectroscopy in a 1 : 1 DMF-d₇ : D₂O mixture (see ESI⁺), showing that 50% of the binding occurred after nearly three days. Compared to previous results of **tPt-H₂4PyP**, the reaction of **tPt-Py** is at least three times faster. This supports the assumption that the rather slow reaction rate of **tPt-H₂4PyP** with guanosine is caused by unfavourable electrostatic interactions. Since guanosine and **tPt-Py** react with a 1 : 1 stoichiometry, the assumption was made that the reaction can be treated as a second order reaction (see ESI⁺). The determined slope *S* is equal to a second order reaction constant *k* of 0.0034 ± 0.0001 M⁻¹ s⁻¹. This value compares well with the pseudo-first order reaction rate (*k*₁ of 0.00036 ± 0.00002 s⁻¹) for the reaction of the mono aquated phenanthriplatin with 9-methylguanine.¹⁶ If one takes the concentration of the excess 9-methylguanine into account, one can convert the pseudo-first order reaction rate into a second order rate (*k*₂ of 0.012 ± 0.007 M⁻¹ s⁻¹). This rate is



about 3.5 times faster than the determined rate of **tPt-Py** with guanosine, which is reasonable since **tPt-Py** must first hydrolyse before it can react with guanosine.

Next, we investigated the uptake and distribution of the Zn–Pt complex in HeLa cells using ICP-MS. The expected ratio of zinc to platinum is 1:4 in **tPt-Zn4PyP**, provided the compound does not disintegrate. To study the Zn and Pt uptake, we first synthesized a highly enriched isotope ^{67}Zn complex **tPt- $^{67}\text{Zn}4\text{PyP}$** .¹⁷ Using the highly enriched zinc isotope makes it possible to differentiate between naturally occurring zinc and ^{67}Zn ¹⁸ that was introduced into the cells with compound **tPt- $^{67}\text{Zn}4\text{PyP}$** . As ^{67}Zn is expensive, initially, a careful and lengthy series of optimizations of the synthesis of zinc-5,10,15,20-tetra-(4'-pyridyl)-porphyrin (**Zn4PyP**) starting from elemental zinc with a natural distribution of isotopes was performed (Table S2, ESI†). Using these optimized conditions, the synthesis of **tPt-Zn4PyP** was then repeated starting from elemental ^{67}Zn , which delivered complex **tPt- $^{67}\text{Zn}4\text{PyP}$** (Fig. S5–S10, ESI†). The uptake study shows that only one quarter of all applied platinum atoms entered the HeLa cells, whereas three quarters of ^{67}Zn atoms entered the cells (Fig. S13 and S14, ESI†). Subsequently, we studied the distribution of platinum and ^{67}Zn in the different cell compartments (Table S4, ESI†). In the nucleus, the ratio of ^{67}Zn to Pt was found to be 3:1 instead of 1:4. These results clearly demonstrate that the complex **tPt- $^{67}\text{Zn}4\text{PyP}$** is disintegrated into the ^{67}Zn containing porphyrin and the relatively non-toxic mono-activated transplatin.¹⁹ The analogous complex **tPt-Zn4PyP** with the natural Zn isotope distribution is expected to behave in the same way as **tPt- $^{67}\text{Zn}4\text{PyP}$** . The transplatin moiety helps to improve solubility and transport the PS into the cell while not being toxic after the cleavage of the conjugate. Transplatin is known to be effluxed faster from the cell than cisplatin.²⁰ This explains why there is less platinum found inside of the cells than ^{67}Zn .

We also investigated the differences between the dark cytotoxicities of the transplatinated series of compounds, **tPt-H₂4PyP**, **tPt-Zn4PyP** and **tPt-Cu4PyP** and the cisplatinated series, **cPt-H₂4PyP**, **cPt-Zn4PyP** and **cPt-Cu4PyP**. The transplatinated series generates, as described before, the *trans*-diamminemonochloroplatinum(II) moiety²¹ after breakage of the N(pyridine)–platinum bond. This fragment has a much lower cytotoxicity than *cis*-diamminemonochloroplatinum(II), which is the first intermediate of cisplatin that must be generated in order to achieve any biological impact of cisplatin. Both, fragments of *cis*- and *trans*-diamminechloridoplatinum(II) in the conjugates **c/tPt-H₂4PyP**, **c/tPt-Zn4PyP** and **c/tPt-Cu4PyP** serve as a cancer-cell penetrating vector. However, the *trans* moiety has a much lower dark toxicity after hydrolysis. The *cis*-diamminemonochloroplatinum(II) fragment has been previously used as a chemotoxic part of a dual acting reagent.^{2b} Our results imply that the *trans*-diamminemonochloroplatinum(II) conjugate would be even more interesting to study as well, maybe yielding a photosensitizer with an improved PI. Next, we examined whether the compounds induce damage to the DNA. It is established, that platinum drugs form drug-DNA adducts,

leading to DNA damage, which triggers cell cycle arrest and DNA repair.²² The drug-induced DNA damage is followed by the phosphorylation of the histone H2AX, which is involved in the recruiting and localizing of the DNA damage repair proteins.

Therefore, γH2AX is considered an important marker to observe DNA damage, in particular double strand breaks.²³ To determine the effectiveness of light-triggered DNA damage by the metal porphyrin conjugates, we treated HeLa cells with 500 nM **tPt-H₂4PyP** and **tPt-Zn4PyP** for 14 h and subsequent light irradiation. Cells were then stained with γH2AX -specific Alexafluor 594 antibody and visualized by confocal microscopy.

As shown in the immunofluorescent images (Fig. 3 and Fig. S15, ESI†), we could observe that **tPt-H₂4PyP** and **tPt-Zn4PyP** both induced severe DNA damage upon light irradiation, as observed by the increased accumulation of γH2AX . This means that **tPt-H₂4PyP** and **tPt-Zn4PyP** generate significantly more DNA damage after light irradiation compared with the experiments in the dark.

To assess the intracellular localization of the photosensitizers, the compounds were co-stained with a marker for the Golgi organelle. **tPt-Zn4PyP** co-localized in the Golgi apparatus (Fig. 4). This experimental finding is pretty inspiring, as some reported drugs directed against the Golgi complex have been shown to be effective in both androgen-dependent and androgen-independent prostate cancer, through targeting abnormal glycosylation.²⁴

In summary, the insertion of a zinc cation into the phototoxic photosensitizer **tPt-H₂4PyP** yielded an even superior PS **tPt-Zn4PyP** with a phototoxic index of higher than 5880. This is likely due to the higher singlet oxygen quantum yield (Φ_{Δ}) of the zinc porphyrin platinum complex **tPt-Zn4PyP** compared to the free-base porphyrin platinum complex **tPt-H₂4PyP**. The ¹H-NMR experiments to follow the kinetics of the reaction of *trans*-[Pt(NH₃)₂(pyridine)Cl]NO₃ (**tPt-Py**) with guanosine revealed that the monocation *trans*-[Pt(NH₃)₂(pyridine)Cl]⁺ reacts faster than the tetracation **tPt-H₂4PyP**. The tetracation has to overcome the barrier of generating an even five-fold charged state during the reaction with N7 of the guanine base. ICP-MS

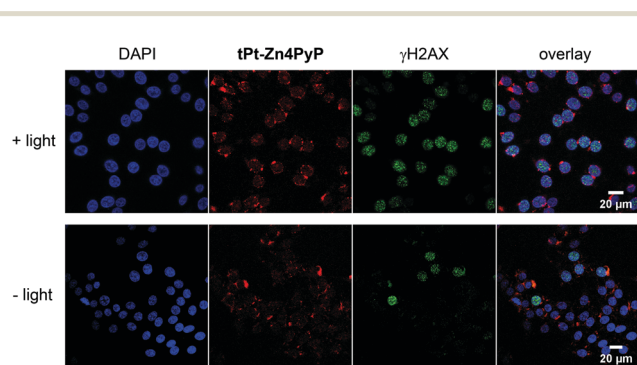


Fig. 3 DNA damage induced by light activated photosensitizer **tPt-Zn4PyP**. Immunofluorescent images of HeLa cells treated with 500 nM **tPt-Zn4PyP** for 14 h, then treated with light compared to the non-irradiated control. Cells were stained with γH2AX -Alexafluor 594 antibody (green) and DAPI. **tPt-Zn4PyP**: 458 nm ex., 630–750 nm em.; γH2AX -Alexafluor 594: 594 nm ex., 610–630 nm.



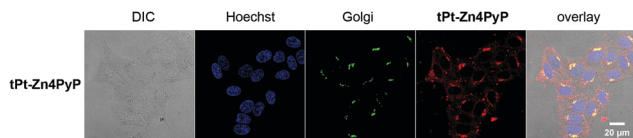


Fig. 4 Co-localization of **tPt-Zn4PyP** with the Golgi apparatus. Representative images of live-cell imaging of HeLa cells treated with 50 μM **tPt-Zn4PyP** and stained with GOLGI-ID[®] Green dye (ENZO ENZ-51028).

studies of the ^{67}Zn labelled complex **tPt- $^{67}\text{Zn4PyP}$** showed that the *trans*-diamminemonochloroplatinum(II) units of the complex **tPt-Zn4PyP** are dissociated from the conjugate inside the HeLa cells. For the first time, such a ^{67}Zn labelled conjugate was employed to monitor the fate of the zinc conjugate within cells. The *trans*-diamminemonochloroplatinum(II) unit is thought to act as a cell-penetrating vector for the photosensitizer, which permits its activation inside the cancer cell by light. We could demonstrate that within the cell, the PS co-localizes with the Golgi apparatus and additionally, induces DNA damage in the nucleus. Unlike conjugates with a cisplatin, transplatin-photosensitizer conjugates have no increased dark toxicity. Therefore, we propose the further investigation of transplatin-drug conjugates for PDT and other oncological treatments.

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Conflicts of interest

There are no conflicts to declare.

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- The maybe obvious alternative to ^{67}Zn -4PyP would have been Pd-4PyP. We have tried many times to synthesize this compound, however the reactions products were always insoluble precipitates. Most likely, the metal coordinates to the exocyclic pyridyl nitrogen atoms, forming an infinite 2D-network.
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