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## Synthesis of benzyl sulfides *via* substitution reaction at the sulfur of phosphinic acid thioesters†

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**An ambident electrophilicity of phosphinic acid thioesters is disclosed. Unexpected carbon–sulfur bond formation took place in the reaction between phosphinic acid thioesters and benzyl Grignard reagents. The developed method for benzyl sulfides has a wide substrate scope and was applicable for the synthesis of a drug analog.**

Sulfides play significant roles in various fields such as medicinal chemistry, agrochemistry, and materials science, and thus many synthetic methods for sulfides have been developed.<sup>1,2</sup> Especially, a group of benzyl sulfides have attracted synthetic chemists because of their utility. Benzyl sulfides are used not only as protected thiols<sup>3</sup> or precursors of sulfonyl chlorides,<sup>4</sup> but also as important structures of bioactive compounds such as sulconazole<sup>5</sup> or dosulepin.<sup>6</sup> However, the synthesis of benzyl sulfides depends almost only on the nucleophilic attack of thiols to alkyl halides, often suffering from the limited scope as well as the unpleasant odor of thiols, especially of benzyl mercaptans. Herein, we disclose a new synthetic method for benzyl sulfides *via* the reaction between benzyl Grignard reagents and phosphinic acid thioesters, which have a phosphorus–sulfur bond.

Previously, we found that Grignard reagents attacked the phosphorus atoms<sup>7,8</sup> of phosphonic acid dithioesters to give phosphinic acid thioesters, which also reacted with Grignard reagents at higher temperature by further substitution reactions on the phosphorus atoms to provide a wide range of phosphine oxides.<sup>9</sup> In the course of further investigation for the reactivity of phosphinic acid thioesters, we unexpectedly found that these compounds behaved as two-faced “ambident” electrophiles.<sup>10,11</sup> Thus, while *S*-(4-tolyl) diphenylphosphinothioate (**1a**) reacted with phenylmagnesium bromide at the phosphorus atom to smoothly

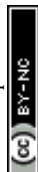
give triphenylphosphine oxide (**2a**), treatment of **1a** with a benzyl Grignard reagent furnished the target phosphine oxide **2b** only in low yield, along with benzyl sulfide **3a** as the major product (Fig. 1A). This result suggested a possibility that benzyl sulfide **3a** was obtained directly by the reaction of the benzyl Grignard reagent at the sulfur atom of phosphinic acid thioester **1a**,<sup>12</sup> contrary to our previous report that C–P bond formation took place using alkyl and aryl Grignard reagents.

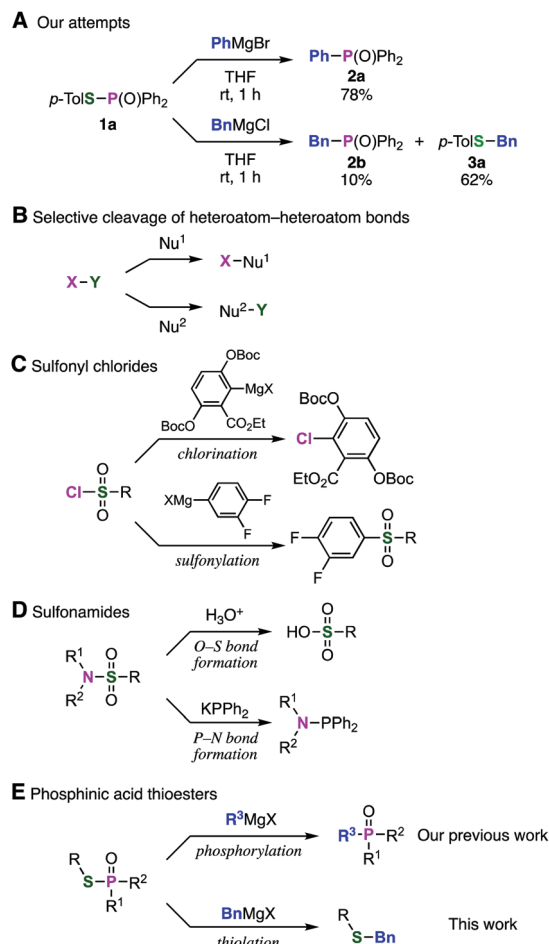
To date, several reports switching the bond-forming site have been reported in the substitution reaction using electrophiles bearing a heteroatom–heteroatom bond depending on the nucleophiles examined (Fig. 1B). For instance, reactions of sulfonyl chlorides with aryl Grignard reagents result in C–Cl bond formation to afford aryl chlorides<sup>10a</sup> or C–S bond formation to afford sulfones<sup>10b</sup> according to the type of Grignard reagents (Fig. 1C). Two types of substitution reactions of sulfonamides have been also developed; P–N bond formation occurred by the reaction with phosphide anion,<sup>10c</sup> while various nucleophiles such as water attacked selectively on the sulfur atom (Fig. 1D). However, limited numbers of the ambident electrophiles were reported so far and controlling the ambident electrophilicity is not easy.<sup>11</sup> These limited but significant site-selective reactions motivated us to examine the C–S bond-forming reaction of phosphinic acid thioesters with benzyl Grignard reagents in terms of the site-selectivity for efficient synthesis of benzyl sulfides (Fig. 1E).

To improve the site-selectivity and to clarify the generality of the reaction, a variety of phosphinic acid thioesters **1**, which have different substituents on the phosphorus atoms, were subjected to the reaction with a benzyl Grignard reagent (Table 1). Changing the substituents on the phosphorus atom from a phenyl group to alkyl groups such as butyl or cyclohexyl group decreased the selectivities, affording the desired sulfide **3a** in low yields (entries 2 and 3). On the other hand, *S*-benzylation of *S*-(4-tolyl) di(2-tolyl)phosphinothioate (**1e**) proceeded smoothly to afford benzyl sulfide **3a** in the highest yield among tested, accompanying with a small amount of phosphine oxide **2f** (entry 4). Thus, we successfully controlled the ambident

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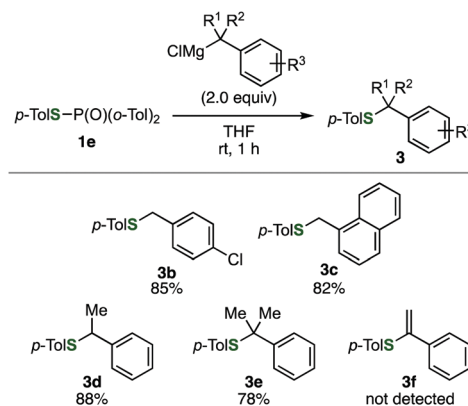
**Fig. 1** (A) Abnormal reactivity of a benzyl Grignard reagent toward a phosphinic acid thioester. (B) Nucleophilic substitution reaction of heteroatom–heteroatom compounds. (C) Sulfonyl chlorides. (D) Sulfonamides. (E) Phosphinic acid thioesters.

**Table 1** Effect of substituents on the phosphorus atom upon the selectivity of nucleophilic attack

Entry	R	Yield <sup>a</sup> (%)			
		<b>1</b>	<b>2</b>	<b>3a</b>	<b>2</b>
1	Ph	<b>1a</b>	<b>2b</b>	62	7
2	<i>n</i> -Bu	<b>1b</b>	<b>2c</b>	12	30
3	<i>c</i> -Hex	<b>1c</b>	<b>2d</b>	33	7
4	<i>o</i> -Tol	<b>1d</b>	<b>2e</b>	86 (88) <sup>b</sup>	5 (6) <sup>b</sup>
5	Mes	<b>1e</b>	<b>2f</b>	69	10
6	OEt	<b>1f</b>	<b>2g</b>	79	ND <sup>c</sup>

<sup>a</sup> NMR yields unless otherwise noted. <sup>b</sup> Isolated yields in parentheses. <sup>c</sup> Not detected.

electrophilicity of phosphinic acid thioesters by changing the substrates, enabling efficient *S*-benzylation in terms of the yield and selectivity. With a bulkier mesityl group, the selectivity rather degraded (entry 5). In addition, when using thiophosphoric



**Fig. 2** Reactions of a phosphinic acid thioester with various benzyl Grignard reagents.

acid ester **1f**, benzyl sulfide **3a** was obtained with perfect selectivity (entry 6).

Benzylation of phosphinic acid thioester **1e** using other benzyl Grignard reagents allowed for efficient preparation of various benzyl sulfides (Fig. 2). Indeed, 4-chlorobenzylation and 1-naphthylmethylation of **1e** proceeded smoothly to give **3b** and **3c**, respectively. A benzyl Grignard reagent bearing a methyl group at the benzylic position was also applicable to the C–S bond formation, affording benzyl sulfide **3d** in high yield. Furthermore, bulky (2-phenylpropan-2-yl)magnesium chloride successfully reacted with **1e** to furnish sulfide **3e**, which is not easy to synthesize by the conventional method using thiols and alkyl halides. In addition, a 1-phenylvinyl Grignard reagent did not participate in the C–S bond forming reaction.

The use of readily available thiophosphoric acid **4** enabled easy access to various alkyl di(2-tolyl)phosphinothioates **1**, which were also benzylation by treating with a benzyl Grignard reagent (Table 2). Alkylation of **4** with a variety of alkyl halides afforded phosphinothioates **1g–k** without damaging functional

**Table 2** Alkylation of thiophosphoric acid and benzylation of phosphinic acid thioesters

Entry	R–X	Yield <sup>a</sup> (%)		Product	Yield <sup>a</sup> (%)
		<b>1</b>	<b>3</b>		
1	MeI	<b>1g</b>	97	<b>3g</b>	82
2	TsO–CH <sub>2</sub> –CH <sub>2</sub> –OTs	<b>1h</b>	11	<b>3h</b>	78
3	BrH–Et <sub>2</sub> N–CH <sub>2</sub> –Br	<b>1i</b>	96	<b>3i</b>	80
4		<b>1j</b>	Quant	<b>3j</b>	40
5		<b>1k</b>	83	<b>3k</b>	ND <sup>b</sup>

<sup>a</sup> Isolated yields. <sup>b</sup> Not detected.



groups such as amino, bromo, and tosyloxy groups (entries 1–4). When the synthesized phosphinic acid thioesters **1g–j** were treated with benzylmagnesium chloride, *S*-benzylation proceeded smoothly to give benzyl sulfides **3g–j**. Notably, a bromo group and a tosyloxy group were tolerated under the benzylation conditions. On the other hand, *S*-benzylation of bulky  $\alpha$ -methylbenzyl ester prepared from **1k** did not occur (entry 5).

Not only alkyl phosphinothioates, an array of aryl phosphinothioates were also prepared *via* the Chan–Lam–Evans-type deborylthiolation (Table 3).<sup>13</sup> Treatment of the corresponding boronic acid with thiophosphinic acid **4** in the presence of triethylamine and a catalytic amount of copper(II) triflate and bipyridyl afforded a variety of phosphinic acid thioesters, such as substituted phenyl thioesters **1l–p** (entries 1–5), 2-naphthyl thioester **1q** (entry 6), 3-thienyl thioester **1r** (entry 7), and styryl thioester **1s** (entry 8). Benzylation of these phosphinic acid thioesters proceeded smoothly to give corresponding benzyl sulfides **3l–s** (entries 1–8). It is noteworthy that benzyl 3-thienyl sulfide (**3r**) and benzyl styryl sulfide (**3s**) were generally difficult to synthesize because of the lack of the availability of the corresponding thiols. Thus, these results clearly demonstrated the utility of this method *via* the Chan–Lam–Evans-type deborylthiolation.

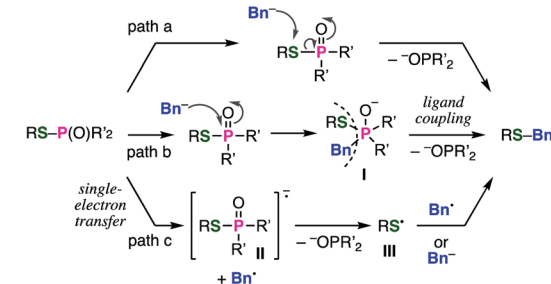
We next turned our attention to the mechanism of this unusual *S*-benzylation (Fig. 3). Plausible reaction mechanism

**Table 3** Chan–Lam–Evans-type synthesis and benzylation of phosphinic acid thioesters

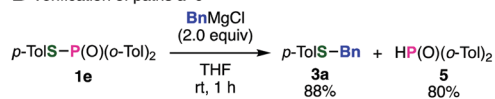
Entry	R	Step A		Step B	
		<b>1</b>	Yield <sup>a</sup> (%)	<b>3</b>	Yield <sup>a</sup> (%)
1		<b>1l</b>	86	<b>3l</b>	97
2		<b>1m</b>	83	<b>3m</b>	88
3		<b>1n</b>	59	<b>3n</b>	61
4		<b>1o</b>	82	<b>3o</b>	77
5		<b>1p</b>	74	<b>3p</b>	57
6		<b>1q</b>	83	<b>3q</b>	67
7		<b>1r</b>	41	<b>3r</b>	87
8		<b>1s</b>	53	<b>3s</b>	61

<sup>a</sup> Isolated yields. bpy = 2,2'-bipyridyl.

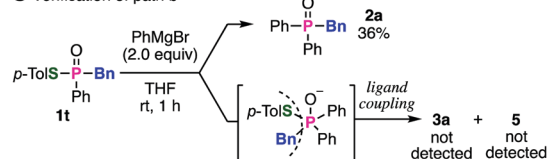
### A Possible mechanisms



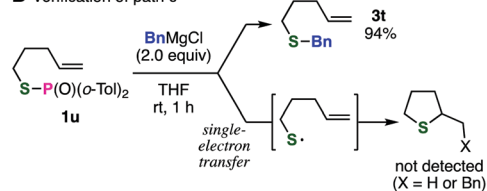
### B Verification of paths a–c



### C Verification of path b



### D Verification of path c

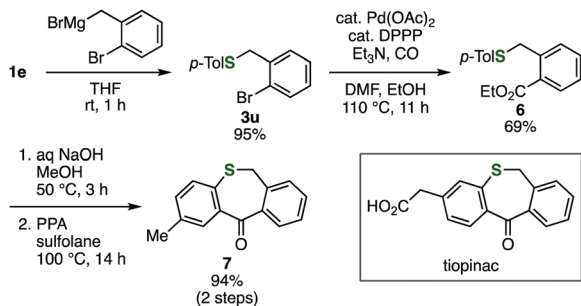


**Fig. 3** Reactions of a phosphinic acid thioester with various benzyl Grignard reagents.

is shown in Fig. 3A including direct substitution reaction of phosphinic acid thioesters **1** with a benzyl Grignard reagent on the sulfur atom (path a). Other reaction mechanisms involving the ligand coupling of pentavalent organophosphorus intermediate **I** (path b) or the single-electron transfer (SET) (path c) could not be excluded at this stage. Considering that C–P bond formation took place using various alkyl or aryl Grignard reagents probably through pentavalent phosphorus intermediates, the *S*-benzylation *via* *P*-benzylation and subsequent ligand coupling<sup>14</sup> of the resulting pentavalent phosphorus intermediate **I** to give the benzyl sulfide is also possible (path b). Another possibility is shown as path c through the SET mechanism, followed by *S*-benzylation between thiyl radical **III** and benzyl radical to give the benzyl sulfide.

To gain insight into the mechanism, we at first examined the products of the reaction between di(2-tolyl)phosphinic acid thioester **1e** and a benzyl Grignard reagent (Fig. 3B). The result showed that di(2-tolyl)phosphine oxide (**5**) was obtained after the protonation of the phosphinate anion in paths a–c as a leaving group,<sup>15</sup> along with benzyl sulfide **3a**. Then, we next treated phosphinic acid thioester **1t**, equipped with a benzyl group in advance, with phenylmagnesium bromide to verify the possibility of path b involving the pentavalent phosphorus intermediate having benzyl and sulfanyl groups (Fig. 3C). As a result, phosphine oxide **2a** was obtained similar to our previous report<sup>9</sup> without generation of benzyl sulfide **3a** and





**Scheme 1** Synthesis of tiopinac analog. DPPP = 1,3-bis(diphenylphosphino)propane. PPA = polyphosphoric acid.

diarylphosphine oxide **5**, suggesting that path b involving reductive elimination of the phosphorus is not plausible. In order to verify the possibility of the SET mechanism, *S*-(4-penten-1-yl) di(2-tolyl)phosphinic acid thioester (**1u**) was subjected to the *S*-benzylation reaction (Fig. 3D). While benzyl 4-penten-1-yl sulfide (**3t**) was obtained in high yield, any radical-cyclization products<sup>16</sup> were not detected in the reaction.<sup>17</sup> From these results, benzyl sulfides would be generated *via* direct attack of benzyl Grignard reagents on the sulfur atoms of phosphinic acid thioesters.

By using the developed benzyl sulfide synthesis, we synthesized an analog of tiopinac, which is a highly potent anti-inflammatory, analgesic, and anti-pyretic agent.<sup>18</sup> Treatment of phosphinic acid thioester **1e** with 2-bromo-5-methoxybenzylmagnesium bromide gave benzyl sulfide **3u** without damaging the bromo group (Scheme 1). Palladium-catalyzed carbonylation of phenyl bromide in methanol afforded methyl ester **6**. Hydrolysis of the ester moiety, and following intramolecular Friedel–Crafts cyclization<sup>19</sup> of the resulting carboxylic acid provided tiopinac analog **7** in shorter steps.

In summary, we found an ambident electrophilicity of phosphinic acid thioesters. While the reactions using various alkyl and aryl Grignard reagents results in C–P bond formations, *S*-benzylation took place by treatment of phosphinic acid thioesters with benzyl Grignard reagents, probably *via* the direct nucleophilic *S*-benzylation reaction at the sulfur atom. The developed *S*-benzylation reaction allowed for synthesizing a variety of benzyl sulfides involving an intermediate of a drug analog from phosphonic acid thioesters and Grignard reagents. Further studies of ambident electrophiles based on the heteroatom chemistry are now underway.

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## Conflicts of interest

There are no conflicts to declare.

## Notes and references

- (a) T. Yamamoto and K. Takimiya, *J. Am. Chem. Soc.*, 2007, **129**, 2224; (b) M. Feng, B. Tang, S. H. Liang and X. Jiang, *Curr. Top. Med. Chem.*, 2016, **16**, 1200; (c) N. Wang, P. Saidharedy and X. Jiang, *Nat. Prod. Rep.*, 2020, **37**, 246.
- (a) S. V. Ley and A. W. Thomas, *Angew. Chem., Int. Ed.*, 2003, **42**, 5400; (b) A. Ghaderi, *Tetrahedron*, 2016, **72**, 4758; (c) Z. Qiao and X. Jiang, *Org. Biomol. Chem.*, 2017, **15**, 1942; (d) T. Matsuzawa, S. Yoshida and T. Hosoya, *Tetrahedron Lett.*, 2018, **59**, 4197.
- J. L. Wood and V. du Vigneaud, *J. Biol. Chem.*, 1939, **131**, 267.
- N. Kharasch and R. B. Langford, *Org. Synth.*, 1964, **44**, 47.
- F. C. Odds, C. E. Webster and A. B. Abbott, *J. Antimicrob. Chemother.*, 1984, **14**, 105.
- E. Richelson and M. Pfenning, *Eur. J. Pharmacol.*, 1984, **104**, 277.
- For substitution reactions of compounds bearing phosphorus–heteroatom bonds, see: (a) C. Lopin, G. Gouhier, A. Gautier and S. R. Piettre, *J. Org. Chem.*, 2003, **68**, 9916; (b) A. Sato, H. Yorimitsu and K. Oshima, *Angew. Chem., Int. Ed.*, 2005, **44**, 1694; (c) T. Wada, A. Kondoh, H. Yorimitsu and K. Oshima, *Org. Lett.*, 2008, **10**, 1155; (d) M.-C. Lamas and A. Studer, *Org. Lett.*, 2011, **13**, 2236; (e) Y. Sato, S.-i. Kawaguchi, A. Nomoto and A. Ogawa, *Angew. Chem., Int. Ed.*, 2016, **55**, 9700; (f) T. Murai, Y. Maekawa, Y. Hirai, K. Kuwabara and M. Minoura, *RSC Adv.*, 2016, **6**, 15180; (g) Y. Maekawa, T. Maruyama and T. Murai, *Org. Lett.*, 2016, **18**, 5264; (h) Y. Maekawa, K. Kuwabara, A. Sugiyama, K. Iwata, T. Maruyama and T. Murai, *Chem. Lett.*, 2017, **46**, 1068; (i) Y. Okugawa, Y. Hayashi, S. Kawachi, K. Hirano and M. Miura, *Org. Lett.*, 2018, **20**, 3670; (j) R. J. Reddy, A. Shankar and A. H. Kumari, *Asian J. Org. Chem.*, 2019, **8**, 2269.
- D. J. Jones, E. M. O’Leary and T. P. O’Sullivan, *Tetrahedron Lett.*, 2018, **59**, 4279.
- Y. Nishiyama, Y. Hazama, S. Yoshida and T. Hosoya, *Org. Lett.*, 2017, **19**, 3899.
- (a) N. Hendricks, P. Himmelberg and E. Fossum, *Polym. Prepr. (Am. Chem. Soc., Div. Polym. Chem.)*, 2007, **48**, 1004; (b) W. Lin, O. Baron and P. Knochel, *Org. Lett.*, 2006, **8**, 5673; (c) S. Yoshida, K. Igawa and K. Tomooka, *J. Am. Chem. Soc.*, 2012, **134**, 2269.
- For examples of studies on ambident electrophiles, see: (a) C. C. P. Wagener, A. M. Modro and T. A. Modro, *J. Phys. Org. Chem.*, 1991, **4**, 516; (b) P. Tundo, L. Rossi and A. Loris, *J. Org. Chem.*, 2005, **70**, 2219.
- For recent reviews of masked thiols as electrophilic sulfur surrogates, see: (a) H. Liu and X. Jiang, *Chem. – Asian J.*, 2013, **8**, 2546; (b) Z. Qiao and X. Jiang, *Org. Biomol. Chem.*, 2017, **15**, 1942; (c) P. Mampuy, C. R. McElroy, J. H. Clark, R. V. A. Orru and B. U. W. Maes, *Adv. Synth. Catal.*, 2020, **362**, 3.
- To our knowledge, this is the first example of Chan–Lam–Evans-type coupling between a boronic acid and a thiophosphinic acid. For a related reaction, Tang and coworkers reported one-pot synthesis of *S*-aryl phosphorothioates probably including a Chan–Lam–Evans-type process: J. Xu, L. Zhang, X. Li, Y. Gao, G. Tang and Y. Zhao, *Org. Lett.*, 2016, **18**, 1266.
- Y. Uchida, K. Onoue, N. Tada, F. Nagao and S. Oae, *Tetrahedron Lett.*, 1989, **30**, 567.
- (a) C. Tejo, J. H. Pang, D. Y. Ong, M. Oi, M. Uchiyama, R. Takita and S. Chiba, *Chem. Commun.*, 2018, **54**, 1782; (b) T. Yano, T. Kawasaki, T. Yuhki, N. Ishida and M. Murakami, *Org. Lett.*, 2018, **20**, 1224.
- M. P. Winters, C. A. Teleha, F.-A. Kang, D. McComsey, J. C. O’Neill, C. Hou, T. Kirchner, P. Wang, D. Johnson and Z. Sui, *Bioorg. Med. Chem. Lett.*, 2014, **24**, 2137.
- The possibility of the reaction pathway involving the benzyl radical generated only from benzylmagnesium chloride by the dissociation was not excluded.
- J. Ackrell, Y. Antonio, F. Franco, R. Landeros, A. Leon, J. M. Muchowski, M. L. Maddox, P. H. Nelson, W. H. Rooks, A. P. Roszkowski and J. M. Muchowski, *J. Med. Chem.*, 1978, **21**, 1035.
- S. A. Laufer, G. M. Ahrens, S. C. Karcher, J. S. Hering and R. Niess, *J. Med. Chem.*, 2006, **49**, 7912.

