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Herein, we present a new strategy to design metal–organic frameworks (MOFs) as adsorbents for ammonia (NH₃) vapour. The linking ligand is functionalized with a sterically hindered Lewis acidic boron (B) centre, allowing efficient capture of NH₃ and easy recycling of the MOF by simply heating at low temperature. The recycled MOF material can be used for NH₃ capture for at least 5 cycles without losing its crystallinity or its luminescence properties.

Toxic gases, such as NH₃, CO, or H₂S, cause an immediate danger to life even at ppm concentrations. For example, the exposure limit of NH₃ in industrial settings recommended by the US Occupational Safety and Health Administration is 25 ppm. Therefore, developing adsorption materials capable of capturing these gases could offer a strategy of great importance for the protection of the environment, workplace safety and public health.

Porous materials with relatively high surface areas such as zeolites and activated carbons have been the traditional adsorbents for toxic gases. During the last decade, metal–organic frameworks (MOFs) and covalent-organic frameworks (COFs), which often possess high porosity and pore-surface tunability, have emerged as adsorbents that are superior to zeolites and activated carbons. It is well known that the sole reliance on the materials’ porosity is not sufficient for effective capture of gas/vapour. The adsorption rate and capacity for gas molecules can be increased when the pore surface incorporates sites for coordination, acid–base interactions, electrostatic interactions, π–π stacking, or H-bonding. The incorporation of these sites either on the pore surface or at the nodes within the scaffold of MOFs and COFs has been widely examined. A large number of MOFs and COFs have integrated Bronsted and Lewis acidic groups, e.g. –COOH, –SO₃H, –PO₃H₂, –OH, open metal sites, or boroxine rings have been reported as promising NH₃ sorbent due to the strong acid–base interaction.

Strong interactions of the gas molecules with the metal nodes or organic linkers of the MOF and COF adsorbents could cause collapse of the materials’ frameworks or difficulties in recycling the adsorbents. For example, in the case of the boronic acid-based COF-10, adsorbed NH₃ can only completely release the adsorbed NH₃ upon heating, affording a stable, recyclable MOF for NH₃ adsorption.

Targeting recyclable NH₃ binding, the highly stable MOF, SION105-Eu with the chemical formula of [[Eu(tctb)₃(H₂O)]] (Fig. 1, left) was selected. This MOF contains the linking ligand tctb with a central sterically hindered Lewis acidic B centre (Fig. 1, right) and has been shown to adsorb CO₂, with an uptake capacity of ~1.9 mmol g⁻¹ at 195 K and 1 bar, and a BET surface area of 216 m² g⁻¹ (Fig. S4, ESI†), and has been utilized for fluoride ion detection in drinking water, and as a catalyst for CO₂ transformation. Herein, we demonstrate this MOF also captures NH₃ vapour but at the same time, releases it upon heating, affording a stable, recyclable MOF for NH₃ adsorption.
SION105-Eu was synthesized based on a solvothermally synthetic procedure, by heating Eu(NO₃)₃·6H₂O (1 equivalent) and tris(p-carboxylic acid)tridurylborane (H₃tctb) (1 equivalent) in a 2 : 1 mixture of DMF : H₂O at 120 °C for 72 hours. The crystalline powder was filtered, soaked in MeOH for three days, filtered again, and then dried in oven at 120 °C. The purity of the sample was confirmed by PXRD, in which the experimental pattern matched well with the simulation based on the single-crystal structure (Fig. S1, ESI†). The MOF floats in water (Fig. S3, ESI†), The MOF shows negligible adsorption, suggesting a good interaction between NH₃ and the MOF.

The activated sample of SION105-Eu was first subjected to NH₃ gas adsorption under dry conditions, with the isotherms measured at 303 and 313 K (Fig. S5, ESI†). At ~1 bar and 303 K, the MOF adsorbs 5.7 mmol g⁻¹ NH₃. The isosteric heat of NH₃ adsorption calculated based on the Clausius–Clapeyron relation is ~28.7 kJ mol⁻¹ for the coverage of 1.5 mmol g⁻¹ and slightly decreases at higher adsorption amounts (Fig. S6, ESI†), suggesting a good interaction between NH₃ and the MOF.

The ability of the material to capture NH₃ vapour from an aqueous NH₃ solution was tested at room temperature (Fig. S7, ESI†). 100 mg of the powder in a 2 mL vial was placed inside a 50 mL vessel. A second vial containing 3.0 mL of aqueous NH₃ solution (28 wt%) was also placed into the larger vessel. With the cap loose, the capture of NH₃ by the MOF was assessed at ambient pressure while higher pressures were achieved with the cap tightly closed. The amount of NH₃ vapour adsorbed by the sample was gravimetrically recorded as a function of time (Fig. 2). The MOF adsorbs up to ~10 wt% after 6 hours, and ~36 wt% after 66 hours when the vessel is tightly closed. Adsorption is slower and reaches saturation of ~10 wt% after 12 hours when the cap is loose. These findings are consistent with the adsorption of NH₃ being pressure dependent. The thermal gravimetric analysis (TGA) of the sample collected after 6 hours of NH₃ adsorption in a closed vessel (NH₃@SION105-Eu) shown in Fig. S8 (ESI†) also illustrates the weight loss of around 10 wt% at below 100 °C, corresponding to the release of NH₃. As the MOF unit molecular weight is 709.43 g mol⁻¹, 10 wt% (5.9 mmol g⁻¹) is similar to the value obtained with dry NH₃ gas, and indicates an average molar ratio of NH₃ : B of ca. 4 : 1.

This uptake is similar to that seen for Zn(INA)₂ (~6 mmol g⁻¹)⁶ and Al-PMOF (~7.6 mmol g⁻¹) at room temperature and 1 bar. In the present case, this suggests the presence of the Lewis acidic B centres prompts non-stoichiometric capture of NH₃ molecules, presumably facilitated by hydrogen bonding interactions among NH₃ molecules. In contrast, exposure of SION105-Eu to pure water showed negligible adsorption, suggesting that capture of more basic NH₃ is enhanced by the electrophilic nature of the MOF.

A suspension containing ~2 mg of the ground MOF powder in 200 µL of THF was deposited onto a filter-paper plate and allowed to dry (Fig. S9, ESI†). The luminescence emission measurements of this sample were made in the presence and absence of NH₃ vapour. After 20 minutes of exposure to NH₃ vapour, the luminescence peak at ~615 nm, characteristic for the transition ⁵D₀ → ⁷F₂ of Eu³⁺, decreased in intensity by ~30% (Fig. S10, ESI†). Since lanthanide luminescence emission is induced by the antenna effect, this partial quenching is consistent with a rather weak electrostatic interaction between the B centre of the antenna ligand tctb⁻ and NH₃.

The PXRD pattern of the MOF sample collected after 6 hours of NH₃ adsorption (NH₃@SION105-Eu) showed the preservation of crystallinity, although the two original peaks in the regions of 2θ = 6.5–6.9° and 10.4–10.9° were split, and their intensities were altered (Fig. 2, bottom). This also suggests only weak electrostatic...
or van der Waals interactions between the MOF and NH₃ as the formation of the B-NH₃ adduct would be expected to lead to quaternization of B and a drastic perturbation of the PXRD pattern. It is noteworthy that van der Waals interactions have been detected for frustrated Lewis acidic boron–olefin interactions.¹⁴ The B environment in trimesitylborane (Mes₃B) is similar to that in SION105-Eu and indeed, exposure of Mes₃B to NH₃ showed no evidence of adduct formation by¹⁵¹¹B-NMR spectra (Fig. S15, ESI†). The FTIR spectrum of the NH₃@SION105-Eu (Fig. S16, ESI†) shows the appearance of N–H stretching bands at 3300–3500 cm⁻¹, confirming the retention of NH₃ within the MOF. In contrast, the MOF MIL103-Eu which is highly porous, but does not contain B shows no evidence of NH₃ adsorption in the FTIR spectrum (Fig. S18, ESI†). This further supports the notion that the electrophilic pores in SION105-Eu facilitates NH₃ adsorption.

To assess the stability of SION105-Eu over the course of adsorption, the powdered MOF was exposed to NH₃ vapour (generated by an aqueous NH₃ solution (28 wt%) in a tightly closed vessel) for 6, 12, and 66 hours. The samples were subsequently heated at 75 °C in an oven for 30 min. In each case, the original mass of the sample was retrieved. PXRD measurements showed that the crystallinity was unaltered after 6 hours, slightly decreased after 12 hours and fully degraded after 66 hours of exposure to NH₃ (Fig. S19, ESI†). Similar experiments with the well-known MOFs, HKUST-1 (Cu) and MOF-5 (Zn) showed that these MOFs completely lost crystallinity in less than 1 hour of exposure to aqueous NH₃ (Fig. S20 and S21, ESI†). Indeed, while a number of transition metal-based MOFs are reported to be stable in the presence of dry NH₃ gas, they decomposed in “wet” NH₃ vapour, thus precluding recycling.⁹ The present data demonstrate that SION105-Eu offers superior stability in this regard.

SION105-Eu was subjected to adsorption of NH₃ for 6 hours in a closed cap vessel and the NH₃ subsequently liberated by heating at 75 °C for 30 min; this process was repeated 5 times. The NH₃ adsorption capacity remains unchanged. The PXRD pattern, FTIR spectrum and luminescence emission of the MOF were unaltered after the 5th cycle of NH₃ capture, being essentially the same as those derived from the original material (Fig. 3). These observations affirm that SION105-Eu is recyclable for NH₃ capture.

Density functional theory (DFT) calculations were performed to confirm the electrostatic mechanism of SION105-Eu (see details in ESI†). It was revealed that no covalent bonds were formed between NH₃ molecules and the MOF scaffold as expected, and the computed heat of NH₃ adsorption was determined to be −40.9 kJ mol⁻¹ at zero coverage, which is close to the value of −28.7 kJ mol⁻¹ derived from the Clausius–Clapeyron equation.

The above results demonstrate that the MOF SION105-Eu has several key features that may be further exploited in the design of stable and recyclable materials for toxic gas capture. Firstly, the use of lanthanide ions in the +3 oxidation state provides hard Lewis acids that associate strongly with hard donor atoms such as O from carboxylate ligands, providing stability in the presence of substrate molecules, in the present case, NH₃. Similarly, the incorporation of the bulky duryl groups on the Lewis acidic B centre within tetrabenzobarrelene ligands precludes strong acid–base B–N interactions, again making the MOF structure robust in the...
presence of NH$_3$. However, the presence of the B on the linkers also makes the pores electrophilic, prompting electrostatic attraction of NH$_3$ and thus its capture.

In conclusion, this work demonstrates a strategy for the design of materials for toxic gas adsorption based on the incorporation of sterically encumbered, electrophilic B sites in SION105-Eu. The resulting highly stable MOF is shown to capture NH$_3$ vapour. Moreover, this binding is reversible with simple heating to 75 $^\circ$C, affording a recyclable material for NH$_3$ capture. Efforts to tune the Lewis acidity of the boron centers for improved uptake capacity and recyclability of related MOFs will be undertaken. In addition, practical applications of such materials are also being targeted via the shaping of the MOF powder into pellets or beads.

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Conflicts of interest

There are no conflicts to declare.

Notes and references