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A double-dysprosocenium single-molecule magnet bound together with neutral ligands†

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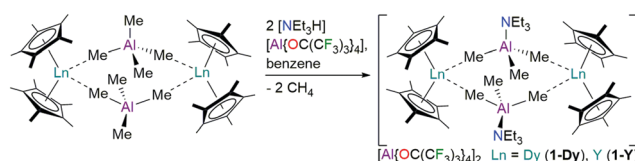
A dinuclear dysprosocenium dication has been synthesised that is bound together by weak interactions between $\{\text{Dy}(\text{Cp}^*)_2\}^+$ fragments and neutral $\text{NEt}_3\text{AlMe}_3$ molecules. The axiality of the Dy^{3+} crystal fields are perturbed by these equatorial interactions but a relatively large effective barrier to magnetisation reversal of $860(60) \text{ cm}^{-1}$ and magnetic hysteresis up to 12 K are observed.

The potential use of single-molecule magnets (SMMs) in high-density data storage devices requires them to retain magnetisation on long timescales (*i.e.* months to years) at liquid nitrogen temperatures, but the best SMMs currently only have relaxation times of a few hours at liquid helium temperatures.^{1,2} The first lanthanide (Ln) SMM was reported in 2003³ and this field has grown rapidly⁴ since design criteria to build Ln SMMs with large effective barriers to magnetisation reversal (U_{eff}) were popularised in 2011.⁵ For Dy^{3+} SMMs an axial ligand field is required to maximise U_{eff} ,⁶ but a report of magnetic hysteresis at $T_{\text{H}} = 60 \text{ K}$ in the isolated dysprosocenium cation $[\text{Dy}(\text{C}_5\text{H}_2^t\text{Bu}_3-1,2,4)_2]^+$ in 2017 showed that a combination of charge-dense π -aromatic rings and an absence of equatorial anion interactions can provide high temperature magnetic remanence.^{2a} This report also predicted that the removal of substituted cyclopentadienyl ring C–H bonds would provide even higher hysteresis temperatures, and this hypothesis has now been proved correct, with T_{H} up to 80 K being reported in the last two years in peralkylated $\{\text{Ln}(\text{Cp}^R)_2\}$ -based SMMs.^{2b–d}

Prior to the synthesis of $[\text{Dy}(\text{C}_5\text{H}_2^t\text{Bu}_3-1,2,4)_2]^+$, much dysprosocene-based SMM research focused on the less bulky $\{\text{DyCp}^*_2\}$ ($\text{Cp}^* = \text{C}_5\text{Me}_5$) framework⁷ and the magnetic properties of a hypothetical $[\text{DyCp}^*_2]^+$ cation were calculated.⁸ Two strategies were investigated in parallel to increase U_{eff} values for $\{\text{DyCp}^*_2\}$ systems; in one approach the equatorial ligand field strength was

progressively decreased,⁹ and a second tactic employed anionic radical ligands to promote magnetic exchange coupling between $\{\text{DyCp}^*_2\}$ units.¹⁰ By utilising weakly coordinating equatorial anions $U_{\text{eff}} = 662 \text{ cm}^{-1}$ and $T_{\text{H}} < 1.8 \text{ K}$ was found for $[\text{Dy}(\text{Cp}^*)_2\{\mu\text{-Fe}(\text{Cp})(\text{CO})_2\}_2]^{9b}$ and by blocking the equatorial sites with neutral Lewis base ligands a U_{eff} of 546 cm^{-1} and a T_{H} of 5.2 K was found for $[\text{Dy}(\text{Cp}^*)_2(\text{NH}_3)_2][\text{BPh}_4]^{9c}$. Given the improvement of the magnetic properties with weakening equatorial ligand fields, we set out to probe the magnetic behaviour of $\{\text{Dy}(\text{Cp}^*)_2\}^+$ fragments with even weaker equatorial interactions; we report our results herein.

The dinuclear Ln^{3+} complexes $[\{\text{Ln}(\text{Cp}^*)_2(\mu\text{-Me}_3\text{AlNEt}_3)_2\}][\text{Al}\{\text{OC}(\text{CF}_3)_3\}_4]_2$ (**1-Ln**, Ln = Dy, Y) were prepared by the protonolysis reactions of parent $[\{\text{Ln}(\text{Cp}^*)_2(\mu\text{-Me}_4\text{Al})\}_2]^{11}$ with two equivalents of $[\text{NEt}_3\text{H}][\text{Al}\{\text{OC}(\text{CF}_3)_3\}_4]^{12}$ in benzene at room temperature; the elimination of two equivalents of methane provides an entropic driving force (Scheme 1). After stirring reaction mixtures overnight the solvent was removed *in vacuo* and crystals of **1-Ln** were grown from chlorobenzene in good yields (**1-Dy**: 76%; **1-Y**: 75%). A doped sample **1-Dy@Y** containing a 5:95 ratio of Dy:Y (composition confirmed by ICP-OES) was synthesised by recrystallizing a mixture of **1-Dy** and **1-Y**. $[\{\text{Ln}(\text{Cp}^*)_2(\mu\text{-Me}_4\text{Al})\}_2]^{11}$ and $[\text{NEt}_3\text{H}][\text{Al}\{\text{OC}(\text{CF}_3)_3\}_4]^{12}$ were prepared by adapting literature procedures; see the ESI† for full details. We selected $[\text{NEt}_3\text{H}][\text{Al}\{\text{OC}(\text{CF}_3)_3\}_4]$ as an abstraction agent as it was recently used to synthesise a bis-phospholyl dysprosium cation,¹² and $[\text{B}(\text{C}_6\text{F}_5)_4]^-$ anions have previously been shown to coordinate to a $\{\text{Sc}(\text{Cp}^*)_2\}^+$ fragment.¹³ Elemental analysis results for **1-Ln-3C₆H₅Cl** gave reproducibly low carbon values due to carbide formation (see ESI,† Section 1).



Scheme 1 Synthesis of **1-Ln**.

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† Electronic supplementary information (ESI) available: Full synthetic details, crystallography, NMR and FTIR spectra, magnetism and *ab initio* calculations. CCDC 1964114–1964119. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9cc08945d



Complexes **1-Ln** are not soluble in non-polar solvents and decompose rapidly in CD_2Cl_2 ; as we avoided the use of coordinating solvents that could displace weakly-bound $\{\mu\text{-Me}_3\text{AlNEt}_3\}$ moieties our NMR characterisation of **1-Ln** was limited to $\text{C}_6\text{D}_5\text{Cl}$ solutions, though these also decomposed into unidentified species over several hours (see ESI†, Section 4). This precluded more detailed analysis of these equilibria by diffusion ordered ^1H NMR spectroscopy and variable temperature (VT) experiments. However, a monomer–dimer equilibrium was previously established for C_7D_8 solutions of $[\{\text{Y}(\text{Cp}^*)_2(\mu\text{-Me}_4\text{Al})\}_2]$ by VT ^1H NMR spectroscopy,^{11a} thus it is reasonable to assume that similar processes are in operation here for **1-Ln**. The paramagnetism of **1-Dy** thwarted its assignment by NMR spectroscopy apart from the presence of two signals in the $^{19}\text{F}\{^1\text{H}\}$ NMR spectrum at -81.40 and -80.28 ppm.

The solid state structures of **1-Ln** were determined by single crystal XRD; **1-Dy-3C₆H₅Cl** is depicted in Fig. 1 and isomorphous **1-Y-3C₆H₅Cl** is shown in the ESI† together with $[\{\text{Ln}(\text{Cp}^*)_2(\mu\text{-Me}_4\text{Al})\}_2]$ ($\text{Ln} = \text{Y}, \text{Dy}$). Each $[\{\text{Dy}(\text{Cp}^*)_2(\mu\text{-Me}_3\text{AlNEt}_3)\}_2]^{2+}$ dication in **1-Dy** is associated with two $[\text{Al}\{\text{OC}(\text{CF}_3)_3\}_4]^-$ counterions and three co-crystallized chlorobenzene molecules in the lattice; the structural parameters of these latter components are unremarkable and there are no significant interactions with the dication. The dimeric structures in **1-Ln** and $[\{\text{Ln}(\text{Cp}^*)_2(\mu\text{-Me}_4\text{Al})\}_2]$ lie about inversion centres, with the two Ln^{3+} cations in each example exhibiting bent metallocene motifs, e.g. **1-Dy**: $\text{Cp}_{\text{centroid}} \cdots \text{Dy} \cdots \text{Cp}_{\text{centroid}} = 138.03(13)^\circ$; $\text{Dy} \cdots \text{Cp}_{\text{centroid}}(\text{av.}) = 2.343 \text{ \AA}$; $\text{Dy} \cdots \text{C}_{\text{ring}} = 2.580(8)\text{--}2.665(7) \text{ \AA}$ (mean 2.634 \AA). The Dy^{3+} coordination spheres are each completed by two Me groups of two separate $\{\mu\text{-Me}_3\text{AlNEt}_3\text{-}\kappa^1\text{-}\kappa^1\text{-C,C'}\}$ moieties (mean $\text{Dy} \cdots \text{C}_{\text{Me}} = 2.794(8) \text{ \AA}$) to give distorted *pseudo*-tetrahedral Dy^{3+} geometries ($\text{C}_{\text{Me}} \cdots \text{Dy} \cdots \text{C}_{\text{Me}} = 82.8(2)^\circ$; range $\text{Cp}_{\text{centroid}} \cdots \text{Dy} \cdots \text{C}_{\text{Me}} = 103.6(2)\text{--}107.8(2)^\circ$) and a diamond-shaped $\text{Dy}_2\text{Al}_2\text{Me}_4$ core, with two Dy and two Al vertices and Me groups on each face; the two

$\{\text{Dy}(\text{Cp}^*)_2\}^+$ units are orthogonal to the Dy_2Al_2 mean plane and the intramolecular $\text{Dy} \cdots \text{Dy}$ distance is $7.365(2) \text{ \AA}$. As expected, the core Al–C bond lengths (mean $2.011(8) \text{ \AA}$) are longer than the external Al–C distances ($1.952(6) \text{ \AA}$) as the latter do not interact with the Dy^{3+} centres.

The equatorial interactions in **1-Dy** likely arise from the electrostatic attraction of the Dy^{3+} centres with $\text{C}^{\delta-}\text{--Al}^{\delta+}$ partial dipoles. The two shortest equatorial $\text{Dy} \cdots \text{H}$ distances that result from these interactions are 2.407 and 2.482 \AA , which are shorter than the two shortest equatorial $\text{Dy} \cdots \text{H}$ distances in $[\text{Dy}(\text{C}_5\text{H}_2\text{Bu}_3\text{-1,2,4})_2][\text{B}(\text{C}_6\text{F}_5)_4]$ (mean 2.499 \AA).^{2a} Each $\{\mu\text{-Me}_3\text{AlNEt}_3\}$ unit in **1-Dy** also exhibits *ca.* $10 \text{ C} \cdots \text{C}$ distances $< 5 \text{ \AA}$ between its Me and Et groups and the Me groups of the Cp^* ligands; it is postulated that the combined strength of these inter-ligand London dipole interactions is also crucial for binding the dication together as such forces are of particular importance in low-coordinate metal complexes.¹⁴ Five-coordinate C atoms have been invoked in related Ln complexes,¹⁵ but we do not comment on this aspect here. The bulk features of **1-Dy** are most reminiscent of the dinuclear Dy^{3+} complex $[\text{Dy}(\text{Cp}^*)_2\{\mu\text{-Fe}(\text{Cp})(\text{CO})_2\}_2]$ (mean $\text{Dy} \cdots \text{Cp}_{\text{centroid}} = 2.342(10) \text{ \AA}$; $\text{Cp}_{\text{centroid}} \cdots \text{Dy} \cdots \text{Cp}_{\text{centroid}} = 141.5(3)^\circ$; $\text{Dy} \cdots \text{Dy} = 7.14(2) \text{ \AA}$),^{9b} but this contains bridging $\{\mu\text{-Fe}(\text{Cp})(\text{CO})_2\text{-}\kappa^1\text{:}\kappa^1\text{-O,O'}\}^-$ anions rather than the neutral $\{\mu\text{-Me}_3\text{AlNEt}_3\}$ units in **1-Dy**. The metrical parameters about the Dy^{3+} centres of **1-Dy** are similar to those seen for $[\text{Dy}(\text{Cp}^*)_2(\text{NH}_3)_2][\text{BPh}_4]$ (mean $\text{Dy} \cdots \text{Cp}_{\text{centroid}} = 2.350(3) \text{ \AA}$; $\text{Cp}_{\text{centroid}} \cdots \text{Dy} \cdots \text{Cp}_{\text{centroid}} = 140.25(6)^\circ$),^{9c} which also contains neutral donors in the equatorial plane, though the dative Dy–N distances in the literature complex (mean $2.47(1) \text{ \AA}$)^{9c} are shorter than the two equatorial $\text{Dy} \cdots \text{C}_{\text{Me}}$ distances in **1-Dy**.

Although the Dy^{3+} centres of **1-Dy** have less axial geometries than several decamethyldysprosocene-based complexes in the literature,^{9,10} we envisioned that the combination of peralkylated Cp ligands and neutral equatorial ligands bound only through electrostatic interactions could endow favourable SMM properties for **1-Dy**. The magnetic susceptibility temperature product ($\chi_{\text{M}}T$) of a microcrystalline sample of **1-Dy** suspended in eicosane is $27.72 \text{ cm}^3 \text{ K mol}^{-1}$ at 300 K (Fig. S22, ESI†), as determined by SQUID magnetometry. This high-temperature value is in excellent agreement with twice the Dy^{3+} free-ion Curie value ($14.17 \text{ cm}^3 \text{ K mol}^{-1}$)¹⁶ and the predicted value from *ab initio* calculations ($13.76 \text{ cm}^3 \text{ K mol}^{-1}$ at 300 K , see below). As expected, the experimental $\chi_{\text{M}}T$ value decreases monotonically down to $24.16 \text{ cm}^3 \text{ K mol}^{-1}$ at 10 K , as excited CF states are thermally depopulated (*ab initio* predicted $12.34 \text{ cm}^3 \text{ K mol}^{-1}$), followed by a sharp decrease below 10 K due to the onset of slow dynamics; this is below the temperature at which we observe bifurcation of the field-cooled/zero-field-cooled (FC/ZFC) data (22 K , Fig. S34, ESI†), but is very close to the local ZFC maximum (8 K , Fig. S31, ESI†).

The dynamic magnetic behaviour of **1-Dy** was probed by low temperature ac susceptibility measurements in zero-field. These data show slow relaxation of magnetisation, with clear in-phase and out-of-phase maxima between 25 and 68 K (Fig. S39, ESI†). We have used CC-FIT2¹⁷ to fit the data to a generalised Debye model, where we can obtain estimated standard deviations

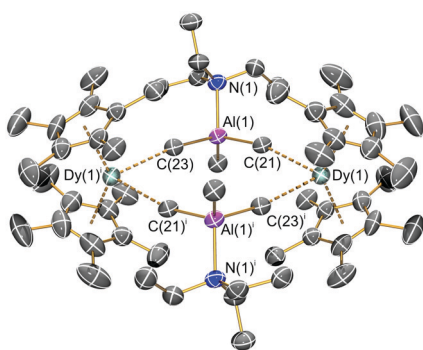


Fig. 1 Molecular structure of the dication of **1-Dy-3C₆H₅Cl**, shown at the 50% probability level; $[\text{Al}\{\text{OC}(\text{CF}_3)_3\}_4]^-$ counterions and lattice solvent are omitted for clarity. Symmetry operation used to generate equivalent atoms: $i = 2 - x, 1 - y, -z$. Selected bond lengths (Å) and angles ($^\circ$): $\text{Dy}(1) \cdots \text{Cp}_{\text{centroid}1} = 2.388(3)$; $\text{Dy}(1) \cdots \text{Cp}_{\text{centroid}2} = 2.348(3)$; $\text{Dy}(1) \cdots \text{C}(21) = 2.800(6)$; $\text{Dy}(1) \cdots \text{C}(23)^i = 2.788(6)$; range $\text{Dy}(1) \cdots \text{C}_{\text{ring}} = 2.580(8)\text{--}2.665(7)$; $\text{Cp}_{\text{centroid}1} \cdots \text{Dy}(1) \cdots \text{Cp}_{\text{centroid}2} = 138.03(13)$; $\text{Cp}_{\text{centroid}1} \cdots \text{Dy}(1) \cdots \text{C}(21) = 103.74(14)$; $\text{Cp}_{\text{centroid}1} \cdots \text{Dy}(1) \cdots \text{C}(23)^i = 107.77(15)$; $\text{Cp}_{\text{centroid}2} \cdots \text{Dy}(1) \cdots \text{C}(21) = 107.29(14)$; $\text{Cp}_{\text{centroid}2} \cdots \text{Dy}(1) \cdots \text{C}(23)^i = 103.56(15)$; $\text{C}(21) \cdots \text{Dy}(1) \cdots \text{C}(23)^i = 82.80(17)$; $\text{Dy}(1) \cdots \text{C}(21) \cdots \text{Al}(1) = 169.9(3)$.



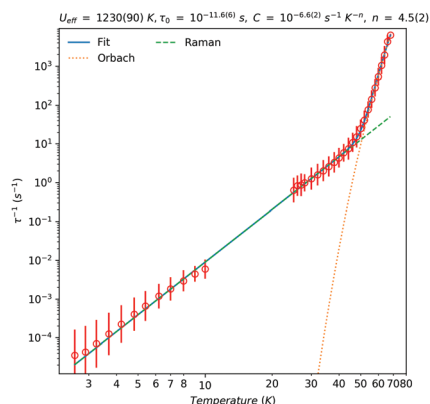


Fig. 2 Temperature dependence of the magnetic relaxation rate of **1-Dy**. Red circles are the relaxation rates extracted from AC susceptibility data (high temperature) and DC magnetisation decay data (low temperature); solid red lines are estimated standard deviations. The solid green line is given by $\tau^{-1} = \tau_0^{-1} \exp[-U_{\text{eff}}/kT] + CT^n$, the dashed green line is given by $\tau^{-1} = \tau_0^{-1} \exp[-U_{\text{eff}}/kT]$ and the dotted orange line is given by $\tau^{-1} = CT^n$ with $U_{\text{eff}} = 860 \text{ cm}^{-1}$ (1230 K), $\tau_0 = 2.51 \times 10^{-12} \text{ s}$, $C = 2.51 \times 10^{-7} \text{ s}^{-1} \text{ K}^{-n}$ and $n = 4.5$.

for the relaxation times (Fig. 2). Below 25 K, the maxima in the ac data fall out of the range of accessible frequencies and we turn to magnetisation decay measurements to access the lowest temperature regime. The DC magnetisation decay data were fitted with stretched exponentials (Table S4, ESI†), where we have used a similar model to that employed by CC-FIT2 to obtain estimated standard deviations for the relaxation times.¹²

In the high-temperature region ($T > 46 \text{ K}$), an Orbach process dominates ($\tau^{-1} = \tau_0^{-1} \exp[-U_{\text{eff}}/kT]$) with an associated effective barrier of $U_{\text{eff}} = 860(60) \text{ cm}^{-1}$ (1230(90) K) and $\tau_0 = 10^{-11.6(6)} \text{ s}$ (ca. $2.5 \times 10^{-12} \text{ s}$); cf. structurally similar $[\text{Dy}(\text{Cp}^*)_2(\mu\text{-Fe}(\text{Cp})(\text{CO}))_2]$ which has $U_{\text{eff}} = 662 \text{ cm}^{-1}$ (952 K),^{9b} though the current record U_{eff} value is 1541 cm^{-1} (2217 K) for $[\text{Dy}(\text{C}_5\text{Pr}_5)(\text{Cp}^*)][\text{B}(\text{C}_6\text{F}_5)_4]$.^{2c} Below 46 K the relaxation rate is governed by a Raman process ($\tau^{-1} = CT^n$) with $C = 10^{-6.6(2)} \text{ s}^{-1} \text{ K}^{-n}$ (ca. $2.5 \times 10^{-7} \text{ s}^{-1} \text{ K}^{-n}$) and $n = 4.5(2)$. The value of n is higher than typical values extracted for isolated Ln metalocenium cations (e.g. $[\text{Dy}(\text{Cp}^{\text{ttt}})_2]^+$, $n = 2.15$;^{2a} $[\text{Dy}(\text{C}_5\text{Pr}_5)(\text{Cp}^*)]^+$, $n = 3.0$;^{2c}), likely due to the influence of the equatorial ligand interactions in **1-Dy**.¹⁸ From these data we can directly measure a 100 s blocking temperature of 10.5 K.

Open magnetic hysteresis loops were observed for **1-Dy** between 2 and 12 K at a sweep rate of ca. 20 Oe s^{-1} around the zero-field region (Fig. 3 and Fig. S27, ESI†); the waist-restricted “butterfly” shapes of these loops are typical for Ln^{3+} SMMs and arise due to strongly field-dependent relaxation near zero field.⁴ The highest temperature at which we observe an open loop is $T_H = 12 \text{ K}$; cf. 6.5 K at ca. 20 Oe s^{-1} for $[\{\text{Dy}(\text{Cp}^*)_2\}_2(\mu\text{-2,2'-bipyrimidine}^*)][\text{BPh}_4]^{10a}$ and 5.2 K at ca. 20 Oe s^{-1} for $[\text{Dy}(\text{Cp}^*)_2(\text{NH}_3)_2][\text{BPh}_4]^{9c}$. The current record value of $T_H = 80 \text{ K}$ at ca. 25 Oe s^{-1} is seen for $[\text{Dy}(\text{C}_5\text{Pr}_5)(\text{Cp}^*)][\text{B}(\text{C}_6\text{F}_5)_4]^{2c}$ thus it is evident that even the very weak neutral equatorial contacts in **1-Dy** promote faster magnetic relaxation than is observed for isolated $[\text{Dy}(\text{Cp}^*)_2]^+$ cations.²

To confirm our implicit assumption that stronger equatorial interactions due to the anionic bridges in the precursor

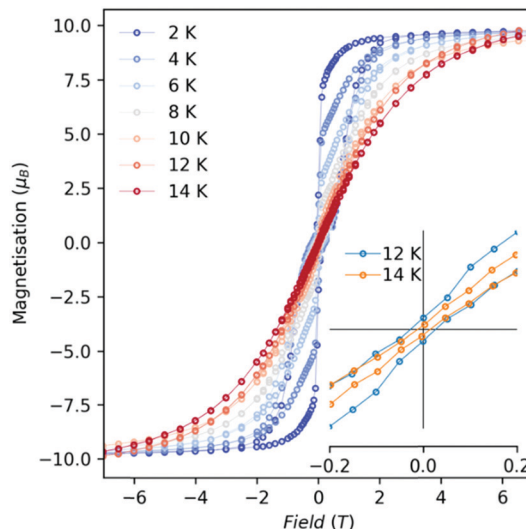


Fig. 3 Magnetic hysteresis of solid **1-Dy**, measured with a mean field sweep rate of $21(9) \text{ Oe s}^{-1}$ for $|H| < 10 \text{ kOe}$, $49(12) \text{ Oe s}^{-1}$ for $10 < |H| < 20 \text{ kOe}$, and $88(17) \text{ Oe s}^{-1}$ for $20 < |H| < 70 \text{ kOe}$. Hysteresis loops recorded from 2–14 K in 2 K steps; inset shows zoomed zero-field region.

complex $[\{\text{Dy}(\text{Cp}^*)_2(\mu\text{-Me}_4\text{Al})\}_2]$ result in a much poorer SMM, we have collected hysteresis and AC relaxation data for this compound. Hysteresis is only open up to 2 K (Fig. S24 and S28, ESI†) and relaxation rates are approximately an order of magnitude faster than for **1-Dy** (Fig. S41 and S42, ESI†). To probe the effect of intramolecular exchange interactions on magnetic relaxation in **1-Dy**, we have examined two doped samples prepared by different routes: **1-Dy@Y** (scrambling of **1-Dy** and **1-Y** in a 5 : 95 ratio in chlorobenzene) and **1-Dy@Y_2** (reaction of a mixture of $[\{\text{Ln}(\text{Cp}^*)_2(\mu\text{-Me}_4\text{Al})\}_2]$ ($\text{Ln} = 5\text{Dy}:95\text{Y}$) with $[\text{NEt}_3\text{H}][\text{Al}(\text{OC}(\text{CF}_3)_3)_4]$). Both samples give practically identical data to one-another (Fig. S28, S29, S32, S33, S35 and S36, ESI†), ruling out any substantial contributions from pure $\{\text{Dy}_2\}$ dimers in either case. The doped samples show similar, yet slightly faster, relaxation rates in zero-field compared to **1-Dy** (Fig. S43 (ESI†); note that we cannot collect AC susceptibility data on these doped samples due to low magnetic moments at high temperatures),¹⁹ and also show open hysteresis up to at least 8 K (Fig. S28, ESI†; we cannot reliably measure the low-moments of these samples near zero-field at higher temperatures). The doped samples also show slightly sharper steps near zero-field in the 2 K hysteresis loops compared to **1-Dy** (Fig. S30, ESI†), where the drop occurs much closer to zero-field; this, along with faster zero-field relaxation rates, suggests that the intramolecular interactions in **1-Dy** provide a small exchange-bias that reduces relaxation in zero-field.

An immobilised powdered sample of **1-Dy** was found to be EPR silent at X-band at 5 K, thus we cannot measure the Dy ··· Dy interaction directly here.²⁰ However, given the core-like nature of 4f electrons and the large separation of the two ions, we suggest that this interaction is purely dipolar. Indeed, simple dipolar calculations based on g -values obtained from complete active space self-consistent field spin-orbit (CASSCF-SO) calculations (see below) suggest a zero-field exchange splitting of ca. 0.42 cm^{-1} in **1-Dy**, which would require a field of ca. 0.09 T to bring the states



into resonance. This is in good agreement with the significant loss of magnetisation for **1-Dy** between 0.05 and 0.1 T (ESI†, Section S7).¹⁷ Using a purely electrostatic model, equatorial ligands in Ln metallocene frameworks tend to increase Dy···Cp_{centroid} distances and decrease Cp_{centroid}···Ln···Cp_{centroid} angles, which disrupts the axially of the ligand field, lowering the U_{eff} value and the purity of the $m_J = \pm 15/2$ ground state. The effect of molecular geometry and ligand vibrational modes on T_H is a more complex relationship to fully disentangle. To probe the electronic structure of **1-Dy** we performed CASSCF-SO calculations on a model of **1-Dy**, where one of the Dy³⁺ centres was replaced by Lu³⁺ for computational efficiency (Fig. S42; see ESI† for details). The predicted U_{eff} for this model was 842 cm⁻¹ (1211 K), in excellent agreement with experimentally obtained values, thus we confidently assign the Orbach relaxation pathway as proceeding *via* the fifth excited doublet.

To conclude, the double-decamethyldysprosocenium dication $[\{\text{Dy}(\text{Cp}^*)_2(\mu\text{-Me}_3\text{AlNEt}_3)\}_2]^{2+}$ is held together by multiple weak dipolar interactions between the cationic $\{\text{Dy}(\text{Cp}^*)_2\}^+$ fragments and neutral $\{\mu\text{-Me}_3\text{AlNEt}_3\}$ moieties. This dication fulfils the design criteria for high-performance dysprosocenium SMMs of an axial ligand field, rigid charge-dense π -aromatic rings, an absence of ring C–H bonds and a lack of equatorial anion interactions.^{2a} However, magnetic relaxation promoted by the equatorial electrostatic contacts of the $\{\mu\text{-NEt}_3\text{AlMe}_3\}$ adducts have limited both the effective barrier to magnetisation reversal and maximum hysteresis temperature compared to dysprosocenium cations that lack such interactions. An alternative synthetic route is required to prepare an isolated decamethyldysprosocenium cation, which could represent the pinnacle of isolated $[\text{Dy}(\text{Cp}^R)_2]^+$ SMMs as it contains the smallest peralkylated Cp^R rings. Although the $[\text{Dy}(\text{Cp}^*)_2]^+$ cation would exhibit short Dy···Cp_{centroid} distances, the Cp_{centroid}···Dy···Cp_{centroid} angle is more difficult to predict and is also crucial for determining its magnetic parameters.⁸

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Conflicts of interest

There are no conflicts to declare.

Notes and references

- D. Gatteschi, R. Sessoli and J. Villain, *Molecular Nanomagnets*, Oxford University Press, 2006.
- (a) C. A. P. Goodwin, F. Ortu, D. Reta, N. F. Chilton and D. P. Mills, *Nature*, 2017, **548**, 439; (b) K. R. McClain, C. A. Gould, K. Chakarawet, S. J. Teat, T. J. Groshens, J. R. Long and B. G. Harvey, *Chem. Sci.*, 2018, **9**, 8492; (c) F.-S. Guo, B. M. Day, Y.-C. Chen, M.-L. Tong, A. Mansikkamäki and R. A. Layfield, *Science*, 2018, **362**, 1400; (d) C. A. Gould, K. R. McClain, J. Yu, T. J. Groshens, F. Furché, J. R. Long and B. G. Harvey, *J. Am. Chem. Soc.*, 2019, **141**, 12967.
- N. Ishikawa, M. Sugita, T. Ishikawa, S. Koshihara and Y. Kaizu, *J. Am. Chem. Soc.*, 2003, **125**, 8694.
- R. A. Layfield and M. Murugesu, *Lanthanides and Actinides in Molecular Magnetism*, Wiley-VCH, 2015.
- J. D. Rinehart and J. R. Long, *Chem. Sci.*, 2011, **2**, 2078.
- See, for example: Y.-S. Ding, N. F. Chilton, R. E. P. Winpenny and Y.-Z. Zheng, *Angew. Chem., Int. Ed.*, 2016, **55**, 16071.
- B. M. Day, F.-S. Guo and R. A. Layfield, *Acc. Chem. Res.*, 2018, **51**, 1880.
- Y.-S. Meng, Y.-Q. Zhang, Z.-M. Wang, B.-W. Wang and S. Gao, *Chem. – Eur. J.*, 2016, **22**, 12724.
- (a) S. Demir, J. M. Zadrozny and J. R. Long, *Chem. – Eur. J.*, 2014, **20**, 9524; (b) T. Pugh, N. F. Chilton and R. A. Layfield, *Angew. Chem., Int. Ed.*, 2016, **55**, 11082; (c) S. Demir, M. D. Boshart, J. F. Corbey, D. H. Woen, M. I. Gonzalez, J. W. Ziller, K. R. Meihaus, J. R. Long and W. J. Evans, *Inorg. Chem.*, 2017, **56**, 15049; (d) R. Collins, M. J. Heras Ojea, A. Mansikkamäki, J. Tang and R. A. Layfield, *Inorg. Chem.*, 2020, **59**, 642.
- (a) S. Demir, J. M. Zadrozny, M. Nippe and J. R. Long, *J. Am. Chem. Soc.*, 2012, **134**, 18546; (b) S. Demir, M. Nippe, M. I. Gonzalez and J. R. Long, *Chem. Sci.*, 2014, **5**, 4701; (c) F.-S. Guo and R. A. Layfield, *Chem. Commun.*, 2017, **53**, 3130; (d) C. A. Gould, L. E. Darago, M. I. Gonzalez, S. Demir and J. R. Long, *Angew. Chem., Int. Ed.*, 2017, **56**, 10103.
- (a) M. A. Busch, R. Harlow and P. L. Watson, *Inorg. Chim. Acta*, 1987, **140**, 15; (b) P. L. Watson and T. Herskovitz, *Am. Chem. Soc. Symp. Ser.*, 1983, **212**, 459; (c) P. L. Watson, *J. Am. Chem. Soc.*, 1983, **105**, 6491.
- P. Evans, D. Reta, G. F. S. Whitehead, N. F. Chilton and D. P. Mills, *J. Am. Chem. Soc.*, 2019, **141**, 19935.
- A. Berkefeld, W. E. Piers, M. Parvez, L. Castro, L. Maron and O. Eisenstein, *J. Am. Chem. Soc.*, 2014, **134**, 10843.
- D. J. Liptrout and P. P. Power, *Nat. Rev. Chem.*, 2017, **1**, 0004.
- See, for example: (a) C. J. Burns and R. A. Andersen, *J. Am. Chem. Soc.*, 1987, **109**, 5853; (b) W. T. Klooster, R. S. Lu, R. Anwander, W. J. Evans, T. F. Koetzle and R. Bau, *Angew. Chem., Int. Ed.*, 1998, **37**, 1268, and references cited therein.
- D. A. Atwood, *The Rare Earth Elements: Fundamentals and Applications*, John Wiley and Sons Ltd., 2012.
- D. Reta and N. F. Chilton, *Phys. Chem. Chem. Phys.*, 2019, **21**, 23567.
- (a) C. A. P. Goodwin, D. Reta, F. Ortu, N. F. Chilton and D. P. Mills, *J. Am. Chem. Soc.*, 2017, **139**, 18714; (b) C. A. P. Goodwin, D. Reta, F. Ortu, J. Liu, N. F. Chilton and D. P. Mills, *Chem. Commun.*, 2018, **54**, 9182.
- (a) F. Habib, P.-H. Lin, J. Long, I. Korobkov, W. Wernsdorfer and M. Murugesu, *J. Am. Chem. Soc.*, 2011, **133**, 8830; (b) Y.-N. Guo, G.-F. Xu, W. Wernsdorfer, L. Ungur, Y. Guo, J. Tang, H.-J. Zhang, L. F. Chibotaru and A. K. Powell, *J. Am. Chem. Soc.*, 2011, **133**, 11948; (c) J. Xiong, H.-Y. Ding, Y.-S. Meng, C. Gao, X.-J. Zhang, Z.-S. Meng, Y.-Q. Zhang, W. Shi, B.-W. Wang and S. Gao, *Chem. Sci.*, 2017, **8**, 1288; (d) E. Moreno Pineda, G. Taran, W. Wernsdorfer and M. Ruben, *Chem. Sci.*, 2019, **10**, 5138.
- (a) E. Moreno Pineda, N. F. Chilton, R. Marx, M. Dörfel, D. O. Sells, P. Neugebauer, S.-D. Jiang, D. Collison, J. van Slageren, E. J. L. McInnes and R. E. P. Winpenny, *Nat. Commun.*, 2014, **5**, 5243; (b) M. J. Giansiracusa, E. Moreno-Pineda, R. Hussain, R. Marx, M. Martinez Prada, P. Neugebauer, S. Al-Badran, D. Collison, F. Tuna, J. van Slageren, S. Carretta, T. Guidi, E. J. L. McInnes, R. E. P. Winpenny and N. F. Chilton, *J. Am. Chem. Soc.*, 2018, **140**, 2504.

