Investigation of the structure and performance of Li\([\text{Li}_{0.13}\text{Ni}_{0.305}\text{Mn}_{0.565}]\text{O}_2\) Li-rich cathode materials derived from eco-friendly and simple coating techniques†

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Constructing uniform nanoceramic coating layers is a well-known challenge in the field of coating materials. Herein, Al2O3-coated Li[Li0.13Ni0.305Mn0.565]O2 (LLNM) Li-rich cathode materials are successfully prepared through a dry prilling coating (DPC) method. The structures and electrochemical performances of the Al2O3-coated products are systematically examined. Typically, the cycling stability is enhanced and voltage degradation upon cycling is reduced, benefitting from the unique and controllable nano-sized Al2O3 coating layer. Moreover, metal ion dissolution is avoided when using the DPC method, which is eco-friendly and suitable for large scale production.

Wet chemical synthesis is a common method that is used to form a coating layer on cathode particles. Wang et al.30 reported a freeze-drying method for coating with Al2O3, and the long-term cycling performance was enhanced and capacity decay was effectively suppressed in the meantime. Zhang et al.31 reported TiO2(B)-coated Li[Li0.2Mn0.54Co0.13Ni0.13]O2 composites obtained via a hydrothermal method, and these showed higher reversible capacity and better cycling stability and rate capabilities. Yang et al.32 synthesized Al2O3/SiO2-coated Li1.2Mn0.54Ni0.13Co0.13O2 via a sol–gel method, and the initial coulombic efficiency and capacity retention were improved, while voltage decay was suppressed effectively during charge/discharge cycling. Zheng et al.33 used MgO and TiO2 to coat 0.5Li1.2MnO3•0.5LiNi0.5Co0.2Mn0.3O2 cathode materials. The results showed that the surface coating materials could improve the discharge voltage retention, decrease the Li2CO3 content on the surface of the materials, and suppress side reactions. However, non-uniform coating and metal-ion leaching were observed during the wet chemical synthesis process. Also, the thickness of the coating layer is usually more than 100 nm, which hinders the diffusion of Li ions. Moreover, water pollution caused by wet chemical reactions is also a signification problem that is concerning.

Physical deposition methods, including atomic layer deposition (ALD) and magnetron sputtering, are considered ideal surface coating techniques. Uniformity and atomic-level control of films can be achieved using these techniques, making them useful for a wide range of applications. Dannehl et al.34 prepared an Al2O3-coated Li1.2Mn0.55Ni0.15Co0.15O3 material via an ALD method. The results showed that the capacity retention was improved and the initial capacity was increased. Xia et al.35

1. Introduction

Lithium ion batteries, which are considered a high-performance and low-cost motive power source, are required for boosting the development and popularity of electric vehicles.1 Li-rich cathode materials (Li[Ni,Mn]O2) have attracted extensive research attention due to their high specific capacity and low-cost.2,3 However, Mn/Ni ion dissolution, structural degradation, polarization effects, side reactions with electrolytes, and electrode decomposition at higher potentials result in poor cycling stability at high cut-off voltages.4,5 In the meantime, non-negligible voltage degradation during cycling and poor rate performance6–11 severely hinder the performance and potential application of these materials on a large scale. Generally, surface coating,12–13 interface modification14–17 and ion-doping18–26 are widely used and have been proven to be effective strategies to enhance the electrochemical performance of Li-rich cathode materials.27,28 Among these methods, surface coating with a ceramic layer, such as Al2O3, TiO2 or SiO2, is a facile and valid technique for improving the cycling stability of Li-rich cathode materials.

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fabricated ZnO-modified 0.3Li2MnO3·0.7LiNi5/21Co5/21Mn11/21O2 cathode materials via reactive magnetron sputtering. The experimental data clearly exhibited a much higher initial discharge capacity and coulombic efficiency. Unfortunately, these physical deposition methods rely heavily on sophisticated equipment, and they are not suitable for large-scale industrial production at present.

The dry prilling coating (DPC) method is a facile strategy for material surface modification. As a kind of excellent surface treatment technology, DPC can form uniform coating layers on materials and the thickness of the coating layer can be expeditiously controlled at the nanoscale. Meanwhile, the procedure is simple and eco-friendly and does not require the use of water, and the cost of production is cheap. The DPC method can satisfactorily meet the demands of whole-scale industrialization, and it has been widely recognized and applied for obtaining powder materials in recent years. Nowadays, it is becoming a prospective strategy for constructing uniform thin films, especially in cases where high reproducibility and excellent conformity are required.

As far as we know, the surface coating of Li-rich cathode materials using DPC has been rarely reported previously. This work will validate a novel and simple method for Li-rich cathode materials and provide further understanding of Al2O3 coatings obtained with DPC technology. Herein, Li:[Li0.13Ni0.305Mn0.565]O2 (LLNM) Li-rich cathode materials with an Al2O3 coating are successfully prepared through a dry prilling coating method. The optimized coated samples obtained via DPC coating are investigated. The effects of Al2O3 coating via DPC on the surface morphology before and after cycling, the discharge capacity, the voltage decay, the cycling stability, and Mn/Ni ion dissolution are systematically investigated. The results indicate that appropriate Al2O3 coating dosages could effectively improve the electrochemical cycling stability and decrease capacity fading.

2. Experimental section

Detailed experimental information is included in the ESL†

3. Results and discussion

The simple and efficient DPC strategy includes two processes, mixing and sintering, as shown in Fig. 1a. Fig. 1b shows XRD patterns of samples with different Al2O3 coating dosages. The α-NaFeO2 structure with an R3m space group, which is ordinarily taken as the signature of a LiMO2 (M = Ni, Mn) phase, is indexed using the diffraction peaks.24 The layered structure can be confirmed via the splitting of the (006)/(102) and (108)/(110) peaks.39 Moreover, the peaks at around 23° originate from superlattice ordering of Li+ and Mn2+ in the layered structure of the transition metal oxide, proving the existence of Li2MnO3 (C2/m). No other impure phases are detected, which indicates that phase-pure derivatives with layered structures are achieved. Furthermore, lattice parameters of the samples calculated from the XRD patterns are given in Table 1. There are no significant differences in the lattice parameters between the pristine and coated samples, which suggests that this amount of Al2O3 coating does not change the crystal structure. The integrated I(003)/I(104) intensity ratio in the XRD patterns of the layered oxides is a measurement of the cation mixing between Li+ and Ni2+. The I(003)/I(104) values of all the samples are greater than 1.2, illustrating that these materials effectively restrain cation mixing. Also, the I(003)/I(104) values of the coated samples are greater than that of the pristine sample. This suggests that the LLNM structure is more stabilized after Al2O3 coating.

Table 1 The calculated lattice parameters of the pristine and coated Li-rich cathode materials

<table>
<thead>
<tr>
<th>Sample</th>
<th>a (Å)</th>
<th>c (Å)</th>
<th>c/a</th>
<th>I(003)/I(004)</th>
</tr>
</thead>
<tbody>
<tr>
<td>LLNM-A0</td>
<td>2.8579</td>
<td>14.2215</td>
<td>4.9762</td>
<td>1.3271</td>
</tr>
<tr>
<td>LLNM-A5</td>
<td>2.8568</td>
<td>14.2187</td>
<td>4.9771</td>
<td>1.4270</td>
</tr>
<tr>
<td>LLNM-A10</td>
<td>2.8575</td>
<td>14.2213</td>
<td>4.9768</td>
<td>1.4435</td>
</tr>
<tr>
<td>LLNM-A15</td>
<td>2.8580</td>
<td>14.2226</td>
<td>4.9764</td>
<td>1.4255</td>
</tr>
</tbody>
</table>
The morphologies of Al₂O₃, a pristine sample, and the series of Al₂O₃-coated samples were confirmed via FESEM (Fig. S1† and 2). Both the pristine and coated materials contain fine and uniform particles ranging from 8 to 12 μm in diameter. The surfaces of the pristine spherical particles are smooth, with slight extensions resulting from primary particles (Fig. 2a, e and i); this is very different from the coated samples. The surfaces of LLNM samples coated with nanoscale Al₂O₃ particles via DPC are fuzzy. In particular, with an increase in the Al₂O₃ content, the roughness of the coated samples increased significantly. When the Al₂O₃ content is increased to 1.5 wt%, the LLNM sample is totally covered and the surface is rougher than the other Al₂O₃-coated samples, which may imply that the amount of Al₂O₃ exceeds requirements.

Similarly, SEM images and elemental mapping of an LLNM-A10 sample are also provided (Fig. S2†). Mn, Ni and trace Al

![Fig. 2](image) FESEM images of samples: LLNM-A0 (a, e and i); LLNM-A5 (b, f and j); LLNM-A10 (c, g and k); and LLNM-A15 (d, h and l).

![Fig. 3](image) TEM images (a and c) and HRTEM (b and d) images of LLNM-A0 and LLNM-A10, respectively. (e) An enlargement of the red rectangle area in (d). (f) An enlargement of the red rectangle area in (e).
elements exist on the particles of LLNM. The uniform distribution of Mn, Ni and Al clearly illustrates that their distribution is similar to what is seen in Fig. S2a. This also verifies the homogeneous nature of the LLNM-A10 particles. Combining the elemental distribution mapping and SEM images, it can be proved that Al₂O₃ successfully coated the obtained materials.

Fig. 3 shows HRTEM images of LLNM-A0 and LLNM-A10. The images clearly show the bulk particles of the pristine and coated samples. The pristine material LLNM-A0 displays smooth boundaries without any coating layer on the surface (Fig. 3a and b). For the coated sample of LLNM-A10 (Fig. 3c), some Al₂O₃ nanoparticles are distributed on the LLNM material; the thickness of the coating layer is 28 nm (Fig. 3d), which illustrates that a coating layer has been obtained. Fig. 3e and f further demonstrate the presence of Al₂O₃ (Fig. 3e and f show partial enlargements of Fig. 3d and e). As displayed in Fig. 3f, the interplanar spacing of 0.35 nm corresponds to the (210) plane of Al₂O₃. A uniform modification layer is helpful to suppress the dissolution of transition metal ions and maintain structural stability during the charge–discharge process. The Al₂O₃ coating layer is further probed via Ni, Mn, O, and Al elemental mapping using HRTEM (Fig. S3†). The Al elements are uniformly distributed in the LLNM-A10 sample, which shows that Al₂O₃ has coated the obtained materials successfully.

X-ray photoelectron spectroscopy (XPS) analysis was performed in order to confirm the oxidation states of Ni, Mn and Al. The C 1s peak (284.4 eV) is used to calibrate the binding energy (BE). The elements Ni, Mn and Al in LLNM-A0 and LLNM-A10 are detected in the survey spectrum, as shown in Fig. 4a. The XPS spectra of Ni 2p, Mn 2p and Al 2p species are further illustrated in Fig. 4b–d, respectively. In the Ni 2p spectra, the peak located at 872.3 eV with a satellite peak at 879.1 eV is attributed to Ni 2p½. The Ni 2p½ peaks in the spectra located at 854.1 and 855.4 eV, with a satellite peak at 861.0 eV, are assigned to Ni²⁺ and Ni³⁺, respectively. Ni exists as a mixture of +2 and +3 valence states. The Mn 2p½ and Mn 2p¾ XPS peaks are located at about 653.9 eV and 642.1 eV, respectively, which demonstrates that the main valence state of Mn is +4. The peak at 73.8 eV in the Al 2p spectrum is assigned to Al³⁺ in Al₂O₃. Moreover, the Al 2p spectrum of the DPC derived sample is in accordance with previously reported results from Al₂O₃ coatings. The XPS results indicate that Al₂O₃ is successfully coated on the surface of LLNM.

Fig. 5 and S4–S8 (ESI†) present the cycling performance and charge–discharge curves of the pristine and coated materials at 0.1C (activated for an initial three cycles, 1C = 300 mA g⁻¹) and 1C at room temperature. The LLNM-A0 sample manifests excellent discharge capacities (254.4 mA h g⁻¹ and 186.5 mA h g⁻¹ for the first cycle at 0.1C and 1C, respectively). Al₂O₃ is inactive during the charge–discharge process, which results in distinct initial polarization and specific capacity loss being shown by the coated samples. This can probably be attributed to increases in the contact resistance and charge
transfer resistance. However, the Al$_2$O$_3$-coated samples present more stable long-term cycling performance than the pristine sample. As the cycling goes on, the LLNM-A0 sample capacity retention is approximately level with the LLNM-A10 sample after 200 cycles at 1C. And after 500 cycles, LLNM-A10 displays a higher capacity retention ratio (85.87%) than the pristine sample (72.92%). Appropriate coating with Al$_2$O$_3$ could effectively decrease polarization, which is beneficial for decreasing capacity fading and improving cycling stability performance.

To better describe and understand the effects of Al$_2$O$_3$ coating via DPC on the electrochemical performance, Fig. 6 shows the discharge behaviors of pristine and Al$_2$O$_3$-coated samples at 1C. Obviously, there are remarkable differences between the discharge voltage plateaus of LLNM-A0 and LLNM-A10. The voltage decay of the discharge plateaus is suppressed by the Al$_2$O$_3$ coating layer. Upon cycling, the discharge capacity of LLNM-A0 (Fig. 6a) decreases at a faster rate than LLNM-A10 (Fig. 6b). It can be found that the LLNM-A0 sample undergoes rapid voltage decay (0.31 V) from the 4th to the 500th cycle at 1C, which is much higher than that undergone by LLNM-A10 (0.22 V). Benefiting from the effective coating technology and synergistic effects arising from the Al$_2$O$_3$ coating thickness, using the DPC method to coat LLNM could restrain discharge voltage fading upon cycling. As illustrated in Fig. 6, the Al$_2$O$_3$ coating layer plays a central role in slowing voltage fading polarization. These results indicate that this coating strategy may suppress interfacial reactions between the cathode material and the electrolyte, reducing interfacial resistance.

To demonstrate the suppression of transition metal ion dissolution by the Al$_2$O$_3$ coating layer on the LLNM material, ICP analyses of electrodes were carried out after 500 cycles. Fig. 6c illustrates that transition metal ion dissolution levels in the LLNM-A10 sample were obviously less than in the LLNM-A0 sample. This is a result of the appropriate Al$_2$O$_3$ coating.
strategy. This indicates that Al2O3 coating via DPC technology could stabilize the framework and decrease Mn/Ni ion dissolution from the LLNM cathode material, improving the structural stability over numerous battery cycles.

FESEM and EDS analysis of typical pristine and Al2O3-coated samples after 500 cycles was carried out in order to further figure out how Al2O3 coating via DPC technology affected the cycling stability. LLNM-A0 and LLNM-A10 electrode samples were soaked and washed after 500 charge-discharge cycles before observations were made. As shown in Fig. 7a–c, the images manifest that a number of cracks appeared and the structure deteriorated in the case of LLNM-A0. By contrast, these phenomena are not found in the case of LLNM-A10. Also, the Al2O3 coating layer maintains its morphological and structural

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**Fig. 7** FESEM images of LLNM-A0 (a–c) and LLNM-A10 (d–f) samples after 500 cycles.

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**Fig. 8** AC impedance spectra from LLNM-A0 and LLNM-A10 samples before cycling (a), after 300 cycles (b), and after 500 cycles (c), and the equivalent circuit diagram (d).
completeness, indicating reduced volume change (Fig. 7d–f). The homogeneous distribution of Mn, Ni and Al elements on the surface of LLNM, as shown in the element mapping in Fig. S9, reveals that the coating layer remains nearly intact after 500 cycles, thus effectively protecting LLNM from HF attack.\(^{49}\)

Electrochemical impedance spectra (EIS) from LLNM-A0 and LLNM-A10 before and after different numbers of cycles and the corresponding equivalent circuit diagram are shown in Fig. 8. The Nyquist plots are formed from a semicircle in the high frequency region and an oblique line in the low frequency region. The large intercept with the real axis demonstrates the large ohmic resistance of the cell.\(^7\) In addition, the depressed semicircle is caused by reaction processes at the electrolyte/electrode interface.\(^7\) The oblique line in the lower frequency range is attributed to Warburg impedance, which was associated with lithium ion diffusion through the cathode electrode.

The LLNM-A0 sample manifests excellent \(R_c\) and \(Z_w\) values before cycling. However, as the number of cycles increases, it is obviously seen that the \(R_c\) and \(Z_w\) values of LLNM-A0 are visibly increased compared with LLNM-A10. It is indicated that this is consistent with the battery discharge capacities of LLNM-A0 and LLNM-A10. Furthermore, the \(R_c\) and \(Z_w\) values of LLNM-A0 and LLNM-A10 are increased, but the \(R_c\) values remain approximately similar. It can be inferred that the use of the same conductive agent is the primary cause. The results indicate that the coating dosage maintains a compromise between impedance and polarization: the surface Al\(_2\)O\(_3\) coating slows down the Li mobility in LLNM-A10, while it is beneficial for improving the structural stability of the active material and decreasing polarization.

4. Conclusions

In this work, an Al\(_2\)O\(_3\) thin coating layer on LLNM cathode materials has been successfully obtained through DPC technology. Samples with different Al\(_2\)O\(_3\) coating dosages obtained via a DPC strategy are investigated. An optimized Al\(_2\)O\(_3\) coating layer thickness is beneficial for enhancing the cycling stability of the LLNM cathode materials. At 25 °C and 1C, LLNM-A10 displays a higher capacity retention ratio (85.87%) than a pristine sample (72.92%) after 500 cycles. Meanwhile, the LLNM-A10 sample effectively restrains voltage degradation to about 0.22 V from the 4th to the 500th cycle at 1C. Typically, LLNM-A10 exhibits better material structural stability and cycling performance in batteries compared to the pristine sample. The excellent structural properties and cycling stability can be attributed to the Al\(_2\)O\(_3\) coating layer, which can effectively prevent direct contact between the electrolyte and active material, decrease polarization, and suppress the dissolution of transition metal ions upon cycling. Also, the coating layer will influence the initial discharge efficiency and electrochemical impedance. This study intends to illustrate the characteristics of Al\(_2\)O\(_3\) coating methods and provide some enlightenment regarding further scientific research into lithium ion battery cathode materials and their industrial application. Meanwhile, we hope that these study results can offer and assist the choice of this simple and eco-friendly Al\(_2\)O\(_3\) coating technology for LLNM cathode materials.

Conflicts of interest

There are no conflicts to declare.

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