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CORRECTION

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Correction: Interplay between H₂O and CO₂ coadsorption and space-charge on Y-doped BaZrO₃ surfaces

Jonathan M. Polfus, (10) *ab Jing Yang (10) b and Bilge Yildiz*bc

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Correction for 'Interplay between H_2O and CO_2 coadsorption and space-charge on Y-doped BaZrO₃ surfaces' by Jonathan M. Polfus *et al.*, *J. Mater. Chem. A*, 2018, **6**, 24823–24830.

The calculated entropies of H_2O and CO_2 chemisorption presented in Table 2 in the published article, -6.37×10^{-4} eV K⁻¹ and -8.78×10^{-4} eV K⁻¹, respectively, were not correct due to an error in the reference value used for gaseous H_2O and CO_2 . The correct adsorption entropies are -1.39×10^{-3} eV K⁻¹ for H_2O and -1.68×10^{-3} eV K⁻¹ for CO_2 , respectively.

The less favorable adsorption entropies lead to somewhat lower surface coverages, especially as the temperature increases (Fig. 4). Overall, surface protons become relatively more favored and predominate over carbonate adsorbates across a wider range of temperature and pressure. Hydroxide species become less relevant, and consequently, the maximum space-charge potential increases from about 0.3 V at 500–700 K to about 0.4 V at 500–600 K.

Fig. 7 illustrates that taking into account coadsorption and space-charge formation remains crucial for the calculated surface coverages. Despite the current numerical corrections, the overall findings and conclusions in the original manuscript remain the same.



Fig. 4 Surface coverages of adsorbates and defects (protons, oxygen vacancies, hydroxide and carbonate species) as well as free surface sites (top) and the corresponding space-charge potential (bottom) on $BaZr_{0.8}Y_{0.2}O_3$ (0 0 1) as function of temperature under 1 bar H_2O , 400 ppm CO_2 (a), 0.01 bar H_2O , 400 ppm CO_2 (b) and 1 bar CO_2 , 1 bar H_2O (c) and 0.1 bar H_2O , 1 bar CO_2 .

"SINTEF Industry, Sustainable Energy Technology, PO Box 124 Blindern, NO-0314 Oslo, Norway. E-mail: jonathan.polfus@sintef.no

^bDepartment of Materials Science and Engineering, Massachusetts Institute of Technology, 77 Massachusetts Avenue, Cambridge, MA 02139, USA

Department of Nuclear Science and Engineering, Massachusetts Institute of Technology, 77 Massachusetts Avenue, Cambridge, MA 02139, USA. E-mail: byildiz@mit.edu



Fig. 7 Surface coverages of adsorbates and defects (protons, oxygen vacancies, hydroxide and carbonate species) as well as free surface sites and space-charge potential as function of temperature calculated without space-charge (a) and without competitive CO_2 adsorption (b) compared to the complete surface equilibria for $BaZr_{0.8}Y_{0.2}O_3$ (0 0 1) (c).

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.