

PAPER

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View Journal | View IssueCite this: *J. Mater. Chem. A*, 2019, 7, 25593Niobium disulphide (NbS₂)-based (heterogeneous) electrocatalysts for an efficient hydrogen evolution reaction†Leyla Najafi,^{‡a} Sebastiano Bellani,^{‡a} Reinier Oropesa-Nuñez,^{Ⓜb} Beatriz Martín-García,^{Ⓜa} Mirko Prato,^c Vlastimil Mazánek,^d Doriana Debellis,^e Simone Lauciello,^e Rosaria Brescia,^{Ⓜe} Zdeněk Sofer^{Ⓜd} and Francesco Bonaccorso^{Ⓜ*ab}

The design of efficient and cost-effective catalysts for the hydrogen evolution reaction (HER) is the key for molecular hydrogen (H₂) production from electrochemical water splitting. Transition metal dichalcogenides (MX₂), most notably group-6 MX₂ (e.g., MoS₂ and WS₂), are appealing catalysts for the HER alternative to the best, but highly expensive, Pt-group elements. However, their HER activity is typically restricted to their edge sites rather than their basal plane. Furthermore, their semiconducting properties hinder an efficient electron transfer to the catalytic sites, which impedes a high rate of H₂ production. Herein, we exploit liquid-phase exfoliation-produced metallic (1H, 2H and 3R) NbS₂ nanoflakes, belonging to the class of metallic layered group-5 MX₂, to overcome the abovementioned limitations. Both chemical treatment with hygroscopic Li salt and electrochemical *in operando* self-nanostructuring are exploited to improve the NbS₂ nanoflake HER activity. The combination of NbS₂ with other MX₂, in our case MoSe₂, also provides heterogeneous catalysts accelerating the HER kinetics of the individual counterparts. The designed NbS₂-based catalysts exhibit an overpotential at a cathodic current of 10 mA cm⁻² (η_{10}) as low as 0.10 and 0.22 V vs. RHE in 0.5 M H₂SO₄ and 1 M KOH, respectively. In 0.5 M H₂SO₄, the HER activity of the NbS₂-based catalysts is also superior to those of the Pt/C benchmark at current densities higher than 80 mA cm⁻². Our work provides general guidelines for a scalable and cost-effective exploitation of NbS₂, as well as the entire MX₂ portfolio, for attaining a viable H₂ production through electrochemical routes.

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Introduction

Molecular hydrogen (H₂) represents an ideal high-energy density source (between 120 and 140 MJ kg⁻¹),^{1,2} since it can be produced *via* electrochemical water splitting from renewable sources^{3,4} and its consumption is sustainable and environmentally friendly.^{5,6} To speed up the rate of the hydrogen evolution reaction (HER) (4H⁺ + 4e⁻ → 2H₂ in acidic media; 4H₂O + 4e⁻ → 2H₂ + 4OH⁻ in alkaline media), commercial electrolyzers demand effective catalysts based on Pt-group

elements.⁷⁻⁹ However, the cost¹⁰ and the scarcity¹¹ of these noble metal-based electrocatalysts hinder their massive use in a move toward a *hydrogen economy*.^{12,13} Therefore, great efforts have been made towards the development of cost effective noble-metal-free alternatives, including transition metal alloys,^{14,15} nitrides,^{16,17} chalcogenides,¹⁸⁻²³ phosphides,^{24,25} carbides,²⁶⁻²⁸ and carbonaceous nanomaterials.²⁹⁻³¹ In this context, transition metal dichalcogenides –MX₂– (M = transition metal; X = S, Se, Te), composed of covalently bonded X–M–X blocks,^{32,33} are appealing alternative catalysts for the HER.³⁴⁻³⁸ Theoretical³⁹⁻⁴¹ and experimental studies⁴¹⁻⁴³ for the most investigated group-6 MX₂ (e.g., H-MoS₂ and H-WS₂) have shown that the scarce unsaturated X-edges are the HER-active sites, since they have a near zero Gibbs free energy of adsorbed atomic H (ΔG_{H}^0), while the more abundant basal plane is catalytically inert. Furthermore, the semiconducting properties of group-6 MX₂ limit the electron transfer to the catalytic edge sites,^{44,45} thus hindering the high H₂ production rate required in electrolyzer systems. The design of nanostructured MX₂ (ref. 42 and 46–49) has been pursued to exploit their high *per site* activity, achieving overpotential *versus* the reversible hydrogen electrode (RHE) at 10 mA cm⁻² cathodic current density (η_{10}),

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approaching those of noble metal-based electrocatalysts (*i.e.*, <0.1 V).^{28,50,51} However, MX₂ nanostructuring inevitably raises scalability and cost issues, impeding their facile implementation. Recently, metallic MX₂ based on group-5 metals (*i.e.*, vanadium (V), niobium (Nb), tantalum (Ta)) attracted utmost interest for the HER process due to the predicted HER activity of their basal planes (especially for sulphides),^{52–58} beyond that of the chalcogen and metal edges.^{55,56} Additionally, group-5 MX₂ have been theoretically proposed as ideal supports for anchoring single metal atom catalysts (*e.g.*, Pt, Ni and Pd), boosting the HER activity of the latter.⁵⁹ Recent experimental studies have validated the theoretical predictions, producing chemical vapour deposition (CVD)-synthesized 2H-TaS₂ and 2H-NbS₂ nanoplatelets displaying an η_{10} of ~ 50 – 60 mV with a low catalyst loading of only 10 – 55 $\mu\text{g cm}^{-2}$.⁵⁴ Similar results have also been achieved by another NbS₂ polytype, *i.e.*, chemically exfoliated 3R-NbS₂ flakes.⁵⁷ However, the aforementioned performances were always reached after an electrochemical pre-conditioning of the electrodes consisting of thousands of cyclic voltammetry (CV) scans.^{54,57} The latter cause a progressive self-optimizing evolution of the NbS₂ morphology, *i.e.*, a reduction in the thickness of the nanoplatelets.^{54,57} This morphology change speeds up the electron transport towards the HER-active sites by shortening the interlayer electron-transfer pathways and facilitating the access of the aqueous protons (H_3O^+) to the catalytic films.^{54,57} Meanwhile, it increases the electrochemically accessible surface area of the electrode films.^{54,55,57} Parallel investigations also claimed that during the cycling process, the oxides naturally formed on the ambient-exposed H-TaS₂ surface are peeled off by the H₂ bubbles and the real HER activity of H-TaS₂ is therefore exhibited subsequently.⁵⁵ Despite these breakthrough performances, such self-optimizing fragmentation of the catalyst films could negatively affect their adhesion to the electrode substrates, raising severe doubts about their reliability for high-rate H₂ production.⁶⁰ Therefore, novel insight into the processing of metallic group-5 MX₂ is urgently required for their practical validation for the HER. Prospectively, optimized electrode morphologies obtained during the deposition of the electrode films could be “frozen” by using catalyst binders, such as sulfonated tetrafluoroethylene-based fluoropolymer copolymers (*e.g.*, Nafion).

In this work, we report the HER activity of single/few-layer NbS₂ flakes produced by an environmentally friendly liquid phase exfoliation (LPE) of synthesized material crystals (mixture of 3R- and 2H-polytypes). In contrast to CVD growth⁶¹ and mechanical exfoliation,⁶² LPE is promptly scalable^{63–68} and does not require expensive nanostructuring of NbS₂ materials. Moreover, LPE does not use hazardous chemicals,⁶⁹ which could affect the physico-chemical properties of the chemically exfoliated materials.^{70–72} Finally, LPE does not face the issues related to electrochemical exfoliation methods, *e.g.*, the material oxidation during anodic exfoliation in aqueous electrolytes and the ion intercalation-induced morphology degradation effects.^{73,74} To improve the catalytic activity of the NbS₂ flakes, a solution-based chemical treatment with hygroscopic

bis(trifluoromethane)sulfonimide lithium salt (Li-TFSI) was used to enhance the electrolyte accessibility to the HER-active surface of NbS₂ flakes, overcoming the electrically insulating oxidation of the flake surfaces.^{75,76} *In operando* electrochemical cycling treatment, as recently reported in the literature^{54,55} and licensed document,⁷⁷ was also applied to the LPE-produced NbS₂ flakes. Furthermore, the chemical and the electrochemical treatments were combined to positively benefit from possible synergistic effects on the HER activity of the electrodes. Finally, on the basis of published density functional theory (DFT) calculations and *ab initio* molecular dynamics (AIMD) simulations,⁵⁶ NbS₂ flakes were hybridized with MoSe₂ flakes to tune the ΔG_{H}^0 of the basal planes of NbS₂ flakes to optimal thermo-neutral values (*i.e.*, 0 eV) in the resulting heterogeneous configurations for both edge and basal sites. In acidic solution (0.5 M H₂SO₄), the designed heterogeneous group-5 TMD catalysts (named NbS₂:MoSe₂) indeed outperform their individual components (η_{10} of 0.10 V for NbS₂:MoSe₂, 0.42 V for NbS₂ and 0.28 V for MoSe₂). For the first time, the HER activity was also investigated in alkaline (1 M KOH) solutions, showing a promising η_{10} of 0.22 V. The HER activity of our best electrocatalysts was tested over 12 h of continuous operation at a fixed potential corresponding to an initial cathodic current density of 80 mA cm^{−2}, validating the electrochemical durability of NbS₂-based electrocatalysts. Our results aim to provide key guidelines to efficiently exploit two-dimensional (2D) NbS₂ and, more generally, 2D metallic group-5 MX₂ for the HER *via* scalable material and electrode processing.

Results and discussion

Synthesis and exfoliation of NbS₂ crystals

NbS₂ crystals were synthesized by direct reaction from Nb and S elements.⁷⁸ In more detail, Nb powder and S granules at a *ca.* $1 : 2$ elemental stoichiometry (1 wt% excess of S to avoid Nb-intercalated structures,⁷⁹) were loaded in a quartz glass ampoule were heated at 450 °C for 12 h and subsequently at 600 °C for 48 h. The products were then treated at 900 °C for 48 h and cooled down to room temperature over a period of 24 h. The as-produced NbS₂ crystals were characterized by scanning electron microscopy (SEM) coupled with energy-dispersive X-ray spectroscopy (EDS) (Fig. 1a–c), revealing a near-ideal stoichiometric composition of the NbS₂ crystals (S : Nb atomic ratio ~ 1.8 , see ESI, Table S1†), which is in agreement with previous studies.⁷⁸ The layered structure of the NbS₂ crystals is clearly visible on the edges, as shown by the SEM image in Fig. 1d. The NbS₂ flakes were produced by LPE^{63,80} of the synthesized crystals in 2-propanol (IPA) followed by sedimentation-based separation (SBS)^{81,82} to remove un-exfoliated crystals by ultracentrifugation (see the Experimental section). The use of IPA as a solvent is effective to remove possible S impurities from the sample during the LPE, SBS and filtration processes in the supernatant, since S is slightly soluble in alcohols (solubility > 0.05 mg mL^{−1}).⁸³ Moreover, starting from cost-effective artificial crystals, the LPE method does not require complex material processing or time-consuming bottom-up nanomaterial synthesis (*e.g.*, CVD).⁸⁴ The morphology of the as-produced NbS₂ flakes was characterized



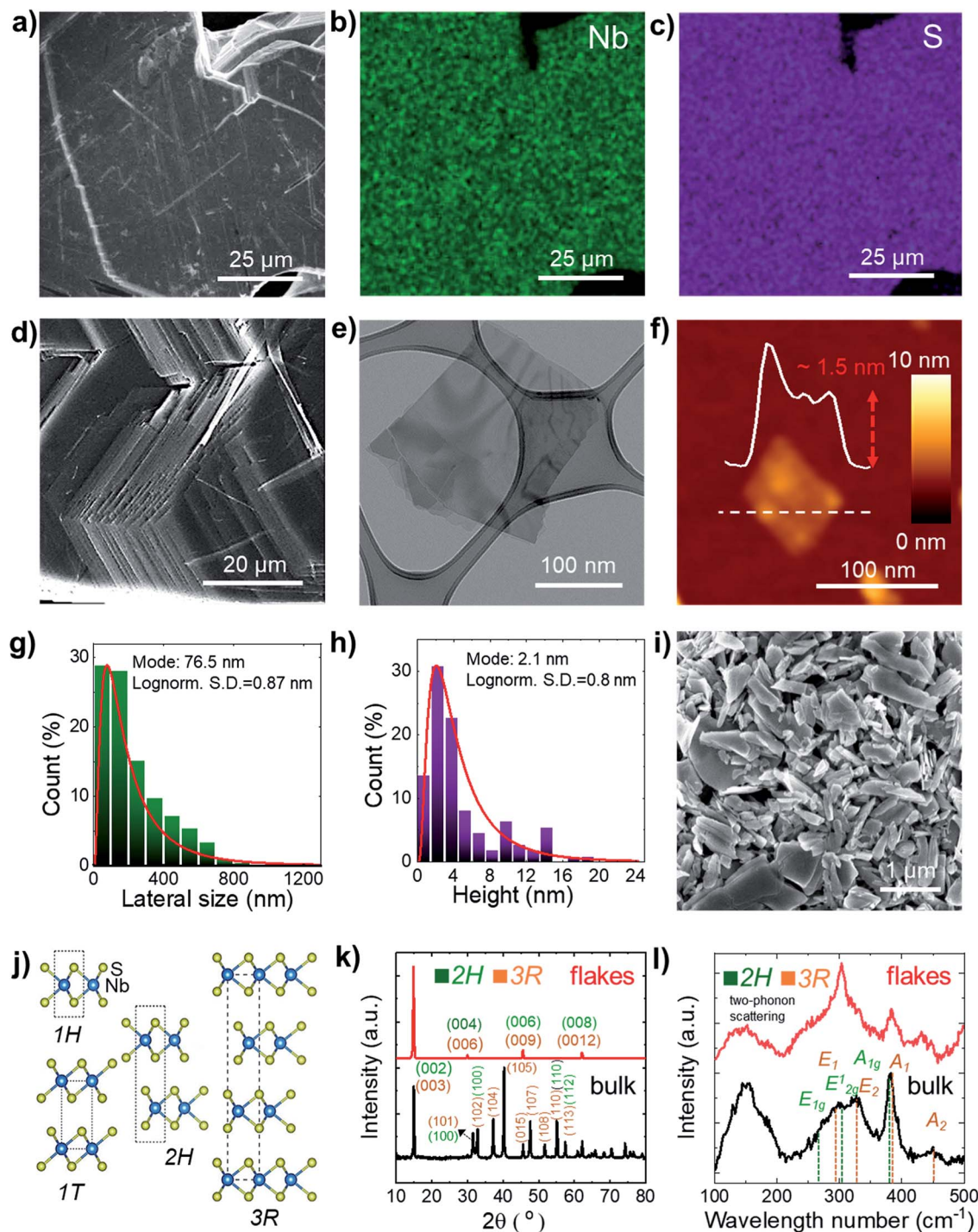
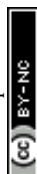


Fig. 1 (a) SEM image of the as-synthesized NbS₂ crystals and the corresponding quantitative EDS maps for (b) Nb (L α peak at 2.17 keV, in green) and (c) S (K α peak at 2.31 keV, in violet). (d) SEM image of a NbS₂ crystal edge showing its layered structure. (e) BF-TEM image of LPE-produced NbS₂ flakes. (f) AFM image of LPE-produced NbS₂ flakes. (g) BF-TEM-based statistical analysis of the lateral dimension of the NbS₂ flakes. (h) AFM statistical analysis of the NbS₂ flakes. (i) SEM image of a film of NbS₂ flakes filtered through a microporous nylon membrane, showing the presence of NbS₂ flakes with a lateral size >1 μ m. (j) Structure of the NbS₂ polytypes experimentally observed in the literature.⁷⁹ (k) XRD and (l) Raman spectra of the synthesized NbS₂ crystals and the LPE-produced NbS₂ flakes. The XRD and Raman peaks assigned to 2H- and 3R-NbS₂ are also shown.

by bright-field transmission electron microscopy (BF-TEM) and atomic force microscopy (AFM) in order to evaluate their lateral dimension and thickness, respectively. Fig. 1e shows the BF-TEM image of representative NbS₂ flakes, displaying wrinkled

surfaces with irregular shapes, but sharp edges. Fig. 1f shows an AFM image of a representative NbS₂ flake, together with its height profile showing a step edge of ~ 1.5 nm.



Statistical BF-TEM-based analysis (Fig. 1g) indicates that the lateral size of the flakes follows a log-normal distribution peaked at ~ 76 nm, with maximum values above $1\ \mu\text{m}$. The statistical AFM analysis (Fig. 1h) shows that the sample is mainly composed of few-layer NbS₂ flakes (the AFM thickness of a NbS₂ monolayer lies generally between 0.6 nm and 0.8 nm, depending on the substrate and the AFM instrumentation,^{85–87} the NbS₂ interlayer distance is ~ 0.6 nm (ref. 88–91)), with a log-normal distribution peaked at ~ 2.1 nm.

Additional SEM measurements were carried out on NbS₂ flakes filtered through a microporous nylon membrane (Fig. 1i). Such a material deposition method produce films exhibiting NbS₂ flakes with a lateral size larger than $1\ \mu\text{m}$. The crystal structure of the NbS₂ crystals and flakes was evaluated by X-ray diffraction measurements (XRD). Fig. 1j shows the structure of the NbS₂ polytypes experimentally observed in the literature.⁷⁹ Typically, NbS₂ crystals are found in two possible hexagonal polytypes with a distinctive stacking order of the NbS₆ prisms.^{92,93} The first one, 2H-NbS₂, is composed of two NbS₂ layer per unit cell (space group: $P6_3/mmc$), whereas the second one, 3R-NbS₂, consists of three layers per unit cell (space group: $R3m$).⁹² Other phases of NbS₂ are thermodynamically unfavoured (trigonal 1T-NbS₂; space group: $P3m1$, derived under special deposition conditions in a monolayer or thin film form),^{94,95} or only theoretically predicted (Haeckelite 1S-NbS₂; space group: P_4/mbm ; at pressure >20 GPa, and tetragonal NbS₂; space group: $I4/mmm$ (ref. 93)). The XRD data (Fig. 1k) show that the as-synthesized NbS₂ crystals consist of a mixture of both 2H and 3R polytypes (indexed to PDF cards no. 04-005-8447 (ref. 97 and 98) and no. 04-004-7343,^{91,98} respectively). After exfoliation, the sample shows XRD peaks corresponding to the diffractions on NbS₂ crystals aligned along the (001) plane^{99,100} (or to 1H-NbS₂, *i.e.*, the monolayer form of both 2H and 3R-NbS₂). This indicates that the NbS₂ flakes are arranged parallel to the *c*-axis perpendicular to the substrate and retain their native crystal structure.^{48,101} Raman spectroscopy analysis (Fig. 1l) further confirms the crystallinity of the exfoliated samples, well displaying the nondegenerate Raman active modes of the native crystals predicted by group theory for 2H-NbS₂ (space group: D_{6h}^4)^{102,103} and 3R-NbS₂ (space group: C_{3v}^5).^{102,104} In more detail, 2H-NbS₂ exhibits the peaks attributed to E_{1g}, E_{2g} and A_{1g} modes at ~ 265 , $\sim 305\ \text{cm}^{-1}$ and ~ 380 , respectively,^{102,103} while 3R-NbS₂ shows the ones associated to E₁, E₂, A₁ and A₂, at ~ 290 , ~ 330 , ~ 385 , and $\sim 450\ \text{cm}^{-1}$, respectively.^{102,104–106} The broad peaks observed at ~ 160 and $\sim 180\ \text{cm}^{-1}$ originate from two-phonon scattering processes in the presence of defects.^{86,91,102,104} Notably, NbS₂ flakes show a prevalence of the characteristic Raman peaks assigned to 2H-NbS₂. This suggests that the exfoliation of 3R-NbS₂ crystals intrinsically causes a 3R-to-H phase conversion toward bilayer and monolayer flakes (*i.e.*, 2H-NbS₂ and 1H-NbS₂), in agreement with previous observations.⁸⁶ In addition, the A₂ mode of 3R-NbS₂ redshifts from $\sim 450\ \text{cm}^{-1}$ in the NbS₂ crystal to $\sim 435\ \text{cm}^{-1}$ in the NbS₂ flakes, as a consequence of the relaxation of the interlayer van der Waals forces with decreasing the NbS₂ thickness.^{86,107} The prevalence of the 2H-phase in the NbS₂ flakes was further confirmed by high resolution TEM (HRTEM)

analysis. Fig. 2a shows a BF-TEM image of a partially suspended NbS₂ flake. The corresponding HRTEM image of a portion of the suspended region is shown in Fig. 2b, which shows a single-crystal structure. Fig. 2c shows the corresponding fast Fourier transform (FFT), which matches with that of [001]-oriented 2H-NbS₂ (ICSD card no. 250595 (ref. 108)). Fig. 2d shows a magnified portion of the HRTEM image, indicating lattice planes. Scanning transmission electron microscopy coupled with EDS (STEM-EDS) analyses were carried out to evaluate the chemical quality of the NbS₂ flakes. Fig. 2e shows the high-angle annular dark-field STEM (HAADF-STEM) image of a partially suspended NbS₂ flake. Fig. 2f and g show the corresponding STEM-EDS maps of Nb and S. The quantitative elemental analysis results in a S/Nb atomic ratio of ~ 1.8 and a low atomic content of O (O/Nb atomic ratio ~ 0.1), which excludes a predominant presence of oxide domains on the flakes. X-ray photoelectron spectroscopy (XPS) measurements (Fig. S1†) were performed to further analyse the element oxidation states on the surface of both the NbS₂ crystal and flakes, revealing the presence multiple NbS₂ phases, *i.e.*, 2H (or 1H) and 3R ones. In particular, the Nb 3d spectra show the presence of three doublets. The first doublet (peaks at $203.4 \pm 0.2\ \text{eV}$ and $206.1 \pm 0.2\ \text{eV}$) and the second doublet (peaks at 204.0 ± 0.2 and $206.7 \pm 0.2\ \text{eV}$) may be both assigned to Nb⁴⁺ states in NbS₂.^{107,109–111} The origin of two doublets for NbS₂ could be ascribed to the presence of multiple NbS₂ phases, *i.e.*, 2H (or 1H) and 3R ones. The peaks located at binding energies of $207.7 \pm 0.2\ \text{eV}$ and $210.4 \pm 0.2\ \text{eV}$ are assigned to the Nb(5+) state in Nb₂O₅.^{112–116} Both the 3R phase and oxidization are attributed to the surface or the edges of the flakes because they were not detected using HRTEM and STEM-EDS (oxygen map not shown here, due to the low oxygen content), which are less surface-sensitive than XPS.

Overall, the morphological, structural and elemental analyses discussed above have shown that NbS₂ flakes were successfully produced by LPE in IPA starting from synthesized NbS₂ crystals. The NbS₂ flakes optimally retained the crystallinity of the native crystal. However, their surface oxidation has to be considered for exploring effectively their electrocatalytic properties for the HER,⁷⁸ in agreement with theoretical expectations.^{52–56}

Fabrication and electrochemical characterization of NbS₂-based electrodes

The direct exploitation of NbS₂ as an electrocatalytic material for the HER has been reported only recently exclusively under acidic conditions.^{58,78,109,117,118} However, just two studies on 2H-NbS₂ nanoplatelets⁵⁴ and mechanically exfoliated tens of nm-thick 3R-NbS₂ flakes⁵⁷ have shown the possibility to approach Pt-like performance after thousands of electrochemical CV cycles, in agreement with theoretical predictions.^{52–57} The unusual self-optimizing/nanostructuring behaviour of NbS₂ nanoplatelets inevitably raises questions about their reliability for a long-term HER (see the discussion in the Introduction section), pointing out the importance to provide new understanding for the fully exploitation of group-5 MX₂.⁶⁰

The electrodes were advantageously produced through a room temperature vacuum filtration of the as-produced NbS₂



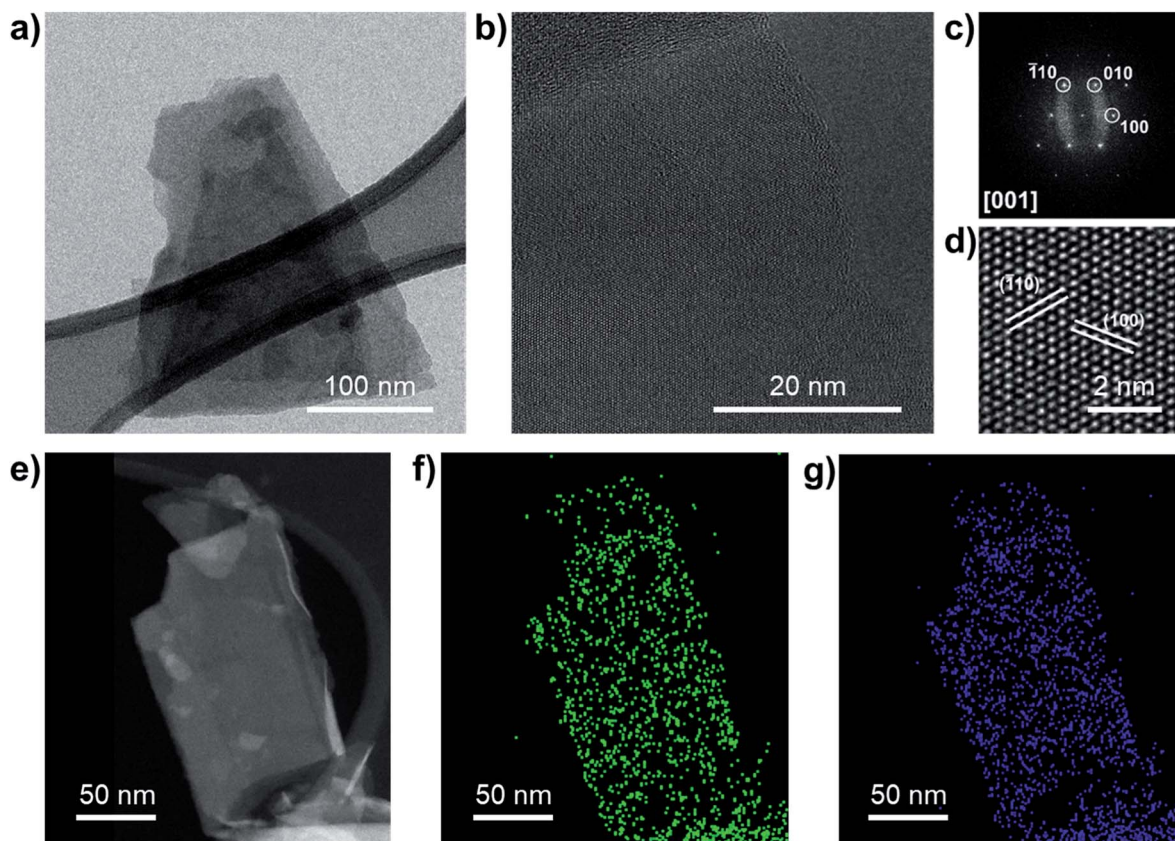


Fig. 2 (a) Elastically filtered BF-TEM image of a NbS₂ flake, partially suspended on a hole in the carbon support film. (b) HRTEM image of a portion of the suspended region of the NbS₂ flake, with (c) the corresponding FFT matching with that of [001]-oriented 2H-NbS₂. (d) Magnified portion of the HRTEM image. The lattice planes are also indicated. (e) HAADF-STEM image of a partially suspended NbS₂ flake and (f and g) corresponding quantitative STEM-EDS maps of Nb (K α , in green) and S (K α , in blue), respectively.

flake dispersions (catalyst mass loading, $\sim 0.20 \text{ mg cm}^{-2}$) on top of a single-walled carbon nanotube (SWCNT) buckypaper (SWCNT mass loading, $\sim 1.31 \text{ mg cm}^{-2}$). In fact, by taking advantage of the material topologies, the NbS₂ flakes cannot pass through the SWCNT tangle.¹¹⁹ Consequently, the SWCNTs effectively filter the NbS₂ flakes, which then form a homogeneous film onto the buckypaper. The preparation of the SWCNT dispersion and the details of the fabrication of the electrodes are described in the Experimental section, following similar protocols adopted in our previous studies on LPE-produced transition metal (di or mono)chalcogenide-based electrocatalysts.^{21,28,48,50,51} Fig. 3a shows a photograph of a representative NbS₂ electrode, which was manually bent to show its flexibility. The surface morphology of the as-prepared electrodes was characterized by SEM measurements. The NbS₂ flakes uniformly cover the mesoporous network of SWCNTs (Fig. 3b), which form a bundle-like morphology (Fig. S2†). The top-view SEM enlargement (Fig. 3c) shows a similar electrode surface to that displayed by NbS₂ flakes directly filtered through a nylon membrane without SWCNTs (see Fig. 1i). To improve the catalytic properties of the NbS₂ electrodes, a solution-based chemical treatment with hygroscopic Li-TFSI salt was used to enhance the electrolyte accessibility to the HER-active surface of the NbS₂ flakes (Fig. 3d). In more detail, the salt ions penetrate

into the porous electrodes, occupying *inter*-flakes sites. The Li⁺ can also intercalate into both surface niobium oxides (*i.e.*, Nb₂O₅ and NbO)^{120–123} and NbS₂,^{98,124–126} thus filling material interstitial sites. Concurrently, the Li-TFSI salt is highly hygroscopic, which means that it strongly adsorbs water, as shown in singular aqueous lithium-ion chemistry.¹²⁷ The water uptake was confirmed by four-probe resistivity measurement of films of NbS₂ flakes, showing a sheet resistance of $\sim 5.1 \times 10^2 \text{ k}\Omega \square^{-1}$ after Li-TFSI treatment, whereas the untreated films exhibited a sheet resistance higher than $10^3 \text{ k}\Omega \square^{-1}$ due to the surface oxidation-induced passivation of the *inter*-flake electrical contact. During the air ambient exposure of the NbS₂-based electrodes, Li-TFSI can also diffuse outward, as also observed in Li-TFSI-doped organic semiconducting films commonly exploited in perovskite solar cells.^{128,129} In addition, Li₂O or LiOH can also be formed by exposing Li-TFSI-treated electrodes to air or water,^{130,131} resulting in H₂ gas evolution. Therefore, the water dragging, the outward diffusion of Li-TFSI and the H₂ evolution may drive the reorientation and nanostructuring (*i.e.*, exfoliation and lateral dimension reduction) of the NbS₂ flakes (Fig. S3†). These effects facilitate the water accessibility to the catalytic surface of the NbS₂ flakes. Notably, the effect of the Li-TFSI treatment resembles the H₂ evolution-aided nanostructuring of the group-5 MX₂-based electrodes occurring



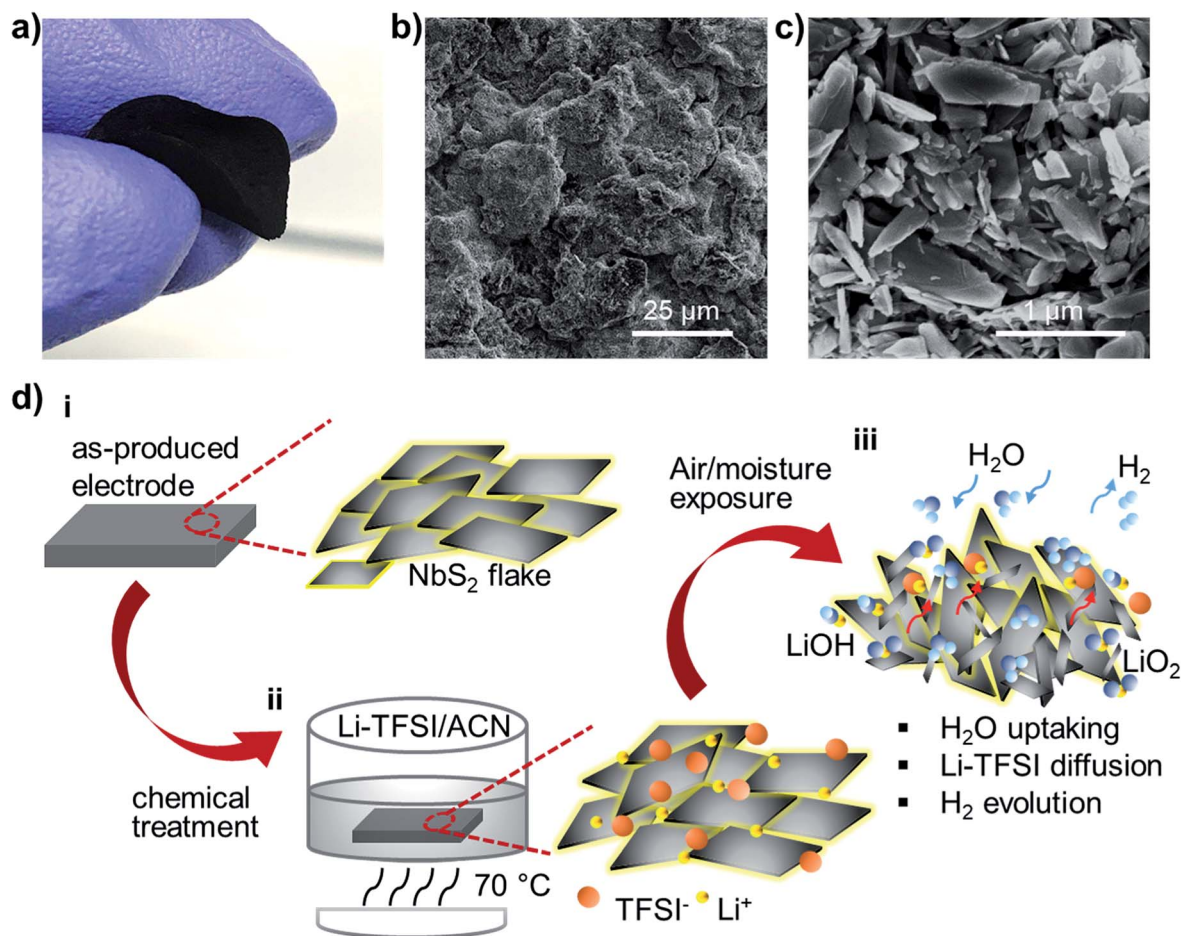


Fig. 3 (a) A photograph of a flexible self-standing NbS₂-based electrode produced by room-temperature sequential vacuum filtration of SWCNT and NbS₂ flake dispersions onto a nylon membrane. (b and c) Top-view SEM images of a NbS₂-based electrode. (d) Sketch of the chemical treatment of the NbS₂-based electrodes: the as-produced NbS₂-based electrode (panel i) and that immersed in a Li-TFSI solution in ACN and heated at 70 °C (panel ii). The salt ions (Li⁺ and TFSI⁻) penetrate between the flakes and/or intercalate into surface oxide or NbS₂, forming a hygroscopic electrode surface. The latter takes up the water, Li-TFSI diffuses towards the surface, and Li⁺ species react with air and water to form Li₂O and LiOH and evolve H₂ (panel iii), reorienting the NbS₂ flakes to increase the water accessibility to the electrode.

during electrochemical treatments (e.g., CV scans).^{54,55} As reported in previous studies,^{54,55,57} the H₂ evolution during CV scans causes a progressive self-optimizing evolution of the NbS₂ morphology from thick-to-thin flakes. This change of the NbS₂ morphology speeds up the electron transport towards the HER-active sites by shortening the interlayer electron-transfer pathways and facilitating the access of the aqueous protons (H₃O⁺) to the catalytic films.^{54,55,57} Meanwhile, it increases the electrochemically accessible surface area of the electrode films.^{54,55,57} However, in contrast to our chemical approach, the electrochemical preconditioning of the electrodes is time-consuming and does not allow any reliable control of the electrode morphology from the start of the HER operation.

The modification of the catalyst film morphology (deposited on flat glassy carbon substrates) induced by both chemical and electrochemical treatments was verified by measuring the double layer capacitances (C_{dl}) of the films (Fig. S4a†). The C_{dl} is proportional to the electrochemically accessible surface area of the films. These results show that the C_{dl} of the Li-TFSI-treated

catalyst film ($\sim 10.8 \text{ mF cm}^{-2}$) is more than three orders of magnitude higher than that of the untreated film. After the electrochemical treatment (i.e., 1000 CV scans, see details in the Experimental section) in 0.5 M H₂SO₄, the catalyst film lost a significant amount of material, which was visible to the naked eye (Fig. S4b†). This effect could be a consequence of the catalyst film fragmentation originated from the H₂ bubbling,^{54,55,57} as well as by the dissolution of possible surface oxides, which can form Nb(OH)₄⁻ in acidic media.¹³² However, the C_{dl} of the electrochemically treated NbS₂ film is similar to that of the untreated film, indicating that the CV scans increase the specific electrochemically accessible surface area (defined by the ratio of the electrochemically accessible surface area and the mass loading) of the as-produced electrodes. X-ray photoelectron spectroscopy measurements (Fig. S5†) do not reveal significant changes in the chemical composition of the NbS₂ flakes after both the chemical and the electrochemical treatments, indicating that the latter mainly affect the morphology of the catalysts. Beyond the aforementioned morphology



changes, it is worth noticing that the presence of alkali-metal cations, including Li^+ , can control the catalytic activity of a catalyst in alkaline media.^{133–135} In fact, it has been suggested that Li^+ induces steric and/or electronic effects on the interfacial water structure and reactivity, enhancing the HER activity of catalysts anchoring Li^+ in the form of complexes.¹³³ Noteworthy, the effective interaction between the Li-TFSI-treated electrode and the electrolyte is also supported by C_{dl} measurements (Fig. S4†), which show an efficient electrochemical double layer formation (Fig. S5†).

Electrochemical measurements of the electrodes

The HER activity of the as-produced electrodes was evaluated in both acidic (0.5 M H_2SO_4) and alkaline (1 M KOH) N_2 -purged solutions at room temperature. To the best of our knowledge, no previous study investigated the HER activity of NbS_2 in alkaline solutions, either theoretically or experimentally.

Fig. 4a and b show the iR -corrected linear sweep voltammetry (LSV) curves in 0.5 M H_2SO_4 and 1 M KOH, respectively, for the as-produced electrodes before and after Li-TFSI treatment (samples hereafter named NbS_2 and Li-TFSI-treated NbS_2 , respectively). In addition, the LSV curves obtained for the NbS_2 electrode after 1000 CV cycles (samples named NbS_2 -CV @ 1000 cycles), the commercial Pt/C benchmark and the SWCNTs (catalyst support) are also plotted. In 0.5 M H_2SO_4 , Li-TFSI-treated NbS_2 exhibits a HER activity ($\eta_{10} = 0.31$ V) significantly higher than that of as-produced NbS_2 ($\eta_{10} = 0.42$ V). The electrochemical cycling also improves the HER activity of NbS_2 , and NbS_2 -CV @ 1000 cycles shows an η_{10} of 0.20 V. Similar results were also obtained in 1 M KOH, where Li-TFSI-treated NbS_2 shows an η_{10} of 0.33 V, similar to that of NbS_2 -CV @ 1000 (0.31 V), whereas the untreated as-produced NbS_2 displays an η_{10} of 0.49 V. Notably, the electrochemical treatment is also effective to improve the HER activity of Li-TFSI-treated NbS_2 (see Fig. S6†) ($\eta_{10} = 0.27$ V in 0.5 M H_2SO_4 ; $\eta_{10} = 0.28$ V in 1 M KOH). As discussed in the previous section,^{54,55,57} the electrochemical treatments cause a progressive self-optimizing evolution of the NbS_2 morphology from thick-to-thin flakes. This change of the flake morphology speeds up the electron transport towards the HER-active sites by shortening the interlayer electron-transfer pathways and facilitating the access of the aqueous protons (H_3O^+) to the catalytic films.^{54,55,57} Moreover, in the absence of catalyst film detachment/dissolution, this morphology change progressively increases the electrochemically accessible surface area of the electrodes, as demonstrated by C_{dl} measurement (Fig. S5a†). In this context, it is important to evidence that the SWCNT substrate is effective in maintaining the mechanical stability of the catalyst films during electrochemical tests.^{68,119} In fact, the electrode using SWCNTs as a support did not show any macroscopic material loss, while the catalyst film deposited onto glassy carbon partially precipitated on the bottom of the cell (see Fig. S4b†).

A rigorous analysis of the HER kinetics, including the extrapolation of the Tafel slope and the exchange current (j_0), was not carried out in this work because ambiguous results can originate from the presence of the SWCNTs. In fact, SWCNTs have a high surface area that leads to a significant capacitive

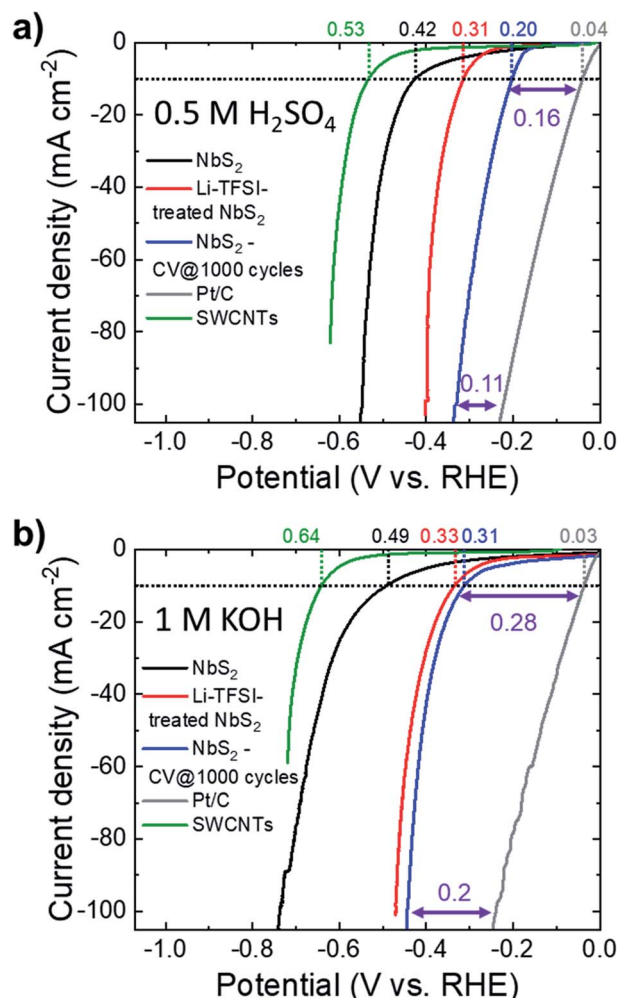


Fig. 4 (a and b) iR -corrected LSV curves for NbS_2 , Li-TFSI-treated NbS_2 , NbS_2 -CV @ 1000 cycles in acidic (0.5 M H_2SO_4) and alkaline (1 M KOH) solutions, respectively. The LSV curves of the Pt/C benchmark and SWCNTs (catalyst support) are also shown for comparison. The η_{10} values measured for the electrodes are also shown. For NbS_2 -CV @ 1000 cycles, the HER overpotential vs. the Pt/C electrode at 10 mA cm^{-2} and 100 mA cm^{-2} is also indicated.

current density (in the range of mA cm^{-2}) even at a low LSV sweep voltage rate (*i.e.*, $\leq 5 \text{ mV s}^{-1}$).⁵¹ This can be misleading during the interpretation of HER kinetic parameters.^{136,137} Interestingly, the overpotentials vs. Pt/C of our electrodes at 100 mA cm^{-2} are significantly lower (*e.g.*, by $\sim 31\%$ and 29% for NbS_2 -CV @ 1000 in 0.5 M H_2SO_4 and 1 M KOH, respectively) than those at a current density of 10 mA cm^{-2} . This agrees with a HER activity also relying on highly abundant basal planes^{52–57} beyond that of the chalcogen and metal edges.^{55,56} Noteworthy, this is substantially different from group-6 MX_2 , in which the HER activity is restricted to only edge/defective sites.^{39–43}

Heterogeneous NbS_2 : MoSe_2 catalysts

To further exploit the potential of NbS_2 flakes for the HER by taking advantage of the theoretically calculated database for the catalytic properties of 2D MX_2 (ref. 39 and 56) (and combination



thereof),⁵⁶ heterogeneous electrocatalysts were produced by simply mixing LPE-produced NbS₂ and MoSe₂ flakes (electrocatalyst hereinafter referred to as NbS₂:MoSe₂). In fact, recent DFT calculations and AIMD simulations⁵⁶ have shown that the heterogeneous stacking of NbS₂ and (2H) MoSe₂ flakes tunes the ΔG_{H}^0 of the resulting heterogeneous configurations in acidic media to optimal thermo-neutral values (*i.e.*, 0 eV) for both edge and basal sites. In more detail, the stacking process initiates an electron transfer from the MoSe₂ to NbS₂ flakes (Fig. 5a), decreasing the positive ΔG_{H}^0 of the NbS₂ basal plane (theoretical values: 0.061 eV for H-coverage = 8.3% and 0.11 for H-coverage = 16.6% from ref. 138; 0.31 eV from ref. 56; 0.12 eV from ref. 39; 0.15 eV from ref. 139) to nearly 0 eV (Fig. 5b).⁵⁶

In addition, we recently demonstrated that MoSe₂ flakes are also efficient HER catalysts,^{28,50,51} showing (after appropriate physical or chemical treatments) an $\eta_{10} < 0.10$ V in both acidic and alkaline solutions.^{28,50,51} Theoretical calculations have proposed that the HER activity of MoSe₂ flakes relies on edge

sites (although it is still up for debate whether they are Mo or Se-edge sites, or both),^{39,56} whereas basal planes are inert ($\Delta G_{\text{H}}^0 > 2$ eV),^{39,56} also when stacked onto NbS₂ flakes ($\Delta G_{\text{H}}^0 \sim 0.98$ eV). As shown in Fig. 5c, the alleged HER activity of the Se-edge sites in MoSe₂ flakes makes the latter advantageous over MoS₂ flakes, whose S-edges strongly bind the H, resulting in negative ΔG_{H}^0 (~ -0.45 eV).^{39,56} The ΔG_{H}^0 of the chalcogen edges of MoSe₂ flakes (~ -0.05 eV from ref. 39 and 56) is slightly varied around the thermoneutral value when MoSe₂ flakes are stacked onto the NbS₂ flakes ($\Delta G_{\text{H}}^0 \sim 0.06$ eV). In addition, by stacking MoSe₂ flakes onto the NbS₂ flakes, the positive ΔG_{H}^0 of metallic edges of MoSe₂ (*e.g.*, 0.31 eV from ref. 56) are low-shifted toward slightly negative near zero values (-0.01 eV from ref. 56). Similar effects have been also theoretically predicted and experimentally confirmed for MoS₂-coated NbS₂ catalysts.¹⁰⁹ Noteworthy, the preparation of heterogeneous catalysts from LPE-produced catalysts is a scalable approach, which does not resort to any morphological/structural and/or chemical modifications of the

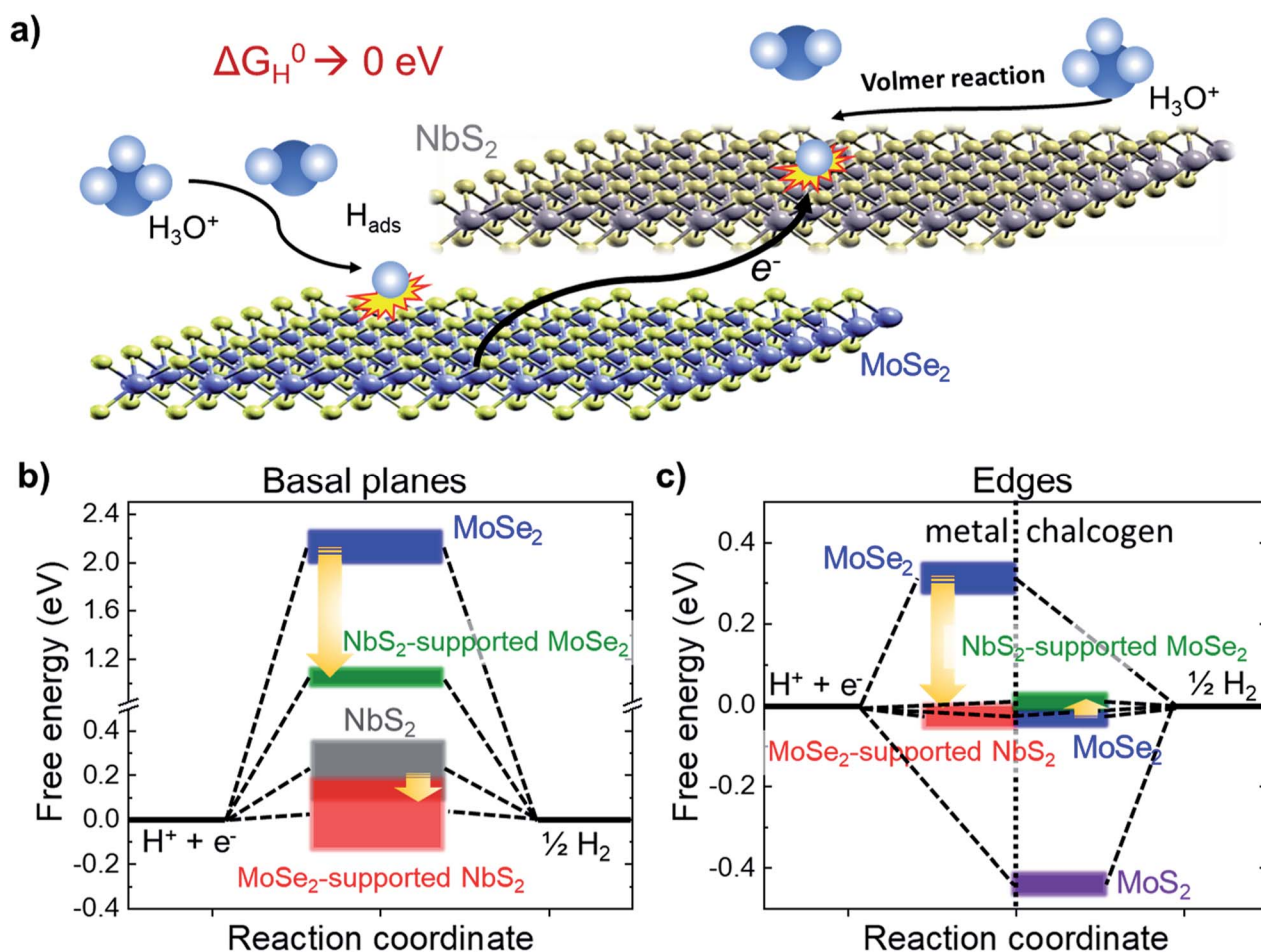


Fig. 5 (a) Schematic illustration of the HER activation of the basal plane of the NbS₂ flakes and metallic edges of MoSe₂ through stacking the NbS₂ and MoSe₂ flakes. The interaction between the flakes promotes an electron transfer from the MoSe₂ flakes towards the NbS₂ flakes, decreasing the positive ΔG_{H}^0 values of the basal plane of the NbS₂ flakes and the metallic edges of MoSe₂ toward 0 eV, promoting the Volmer reaction ($\text{H}_3\text{O}^+ + \text{e}^- \rightleftharpoons \text{H}_{\text{ads}} + \text{H}_2\text{O}$ in acidic condition), *i.e.*, the first step of the HER. (b) Evolution of ΔG_{H}^0 for the basal planes of the NbS₂ and MoSe₂ flakes, and the NbS₂:MoSe₂ hybrid catalysts. The data ranges have been extrapolated from literature databases (ref. 21, 39, 103 and 104). (c) Evolution of ΔG_{H}^0 for the metallic (left side) and chalcogen (right side) edges of the MoSe₂ flakes, and the NbS₂:MoSe₂ hybrid catalysts. The data ranges have been extrapolated from literature databases (ref. 21 and 39).



LPE-produced flakes, such as the introduction of artificial defects, chemical doping of heteroatoms or strain constraints.⁵⁶

The details regarding the production of MoSe₂ flakes are reported in the Experimental section. The morphological, structural, and chemical characterization of the LPE-produced MoSe₂ flakes has been recently reported by our studies.^{28,50,51} In particular, the MoSe₂ flakes have a lateral size (measured by TEM) and thickness (measured by AFM) following lognormal distributions peaked at ~ 30.0 nm and ~ 2.9 nm, respectively.^{28,50,51} Optical absorbance, Raman spectroscopy and XRD measurements have shown that MoSe₂ flakes optimally preserve the MoSe₂ bulk crystallinity.^{28,50,51} In addition, the analysis of HRTEM images of the MoSe₂ flakes and the corresponding FFTs confirmed that the flakes display a hexagonal phase,²⁸ in agreement with XRD and Raman analyses.^{28,50,51}

Following the protocols previously adopted for NbS₂ electrodes, MoSe₂ and NbS₂:MoSe₂ (material mass ratio of 1 : 1) electrodes were produced by vacuum filtration deposition of their dispersions on top of the SWNTs. Fig. S7† shows the top-view SEM images of the MoSe₂ and NbS₂:MoSe₂ electrodes, respectively. The as-produced electrodes display a surface uniformly covered by the flakes. In particular, the heterogeneous electrode (*i.e.*, NbS₂:MoSe₂) shows a morphology resembling those of the MoSe₂ electrode, whose flakes are smaller than NbS₂ flakes, in agreement with the characterization of the materials (see Fig. 1g for NbS₂ flakes and ref. 28 and 51 for MoSe₂ flakes). Although it is reasonable to assume that the vacuum filtration of a mixture of NbS₂ and MoSe₂ flakes in dispersion intrinsically promotes the stacking between flakes of different materials, SEM-EDS measurements were carried out to prove at least the absence of single material domains. The cross-sectional SEM-EDS analysis of the heterogeneous films composed of NbS₂ and MoSe₂ flakes (Fig. S8a†) shows homogeneous distributions for both Nb and Mo elements, which are compatible with an optimal mixing in the dispersion used for the preparation of the electrodes. The Li-TFSI treatment of the MoSe₂ and NbS₂:MoSe₂ electrodes is expected to cause the same effects observed for the NbS₂ electrodes (*i.e.*, hygroscopicity and nanostructuring of the electrodes). The SEM-EDS analysis of the Li-TFSI-treated heterogeneous films (Fig. S8b†) leads to similar results to those of the untreated film (*i.e.*, a homogeneous distribution of both Nb and Mo elements), suggesting a similar packing between NbS₂ and MoSe₂ flakes. The top-view SEM-EDS analyses of the untreated and the Li-TFSI heterogeneous films (Fig. S8c and d,† respectively) also show homogeneous distributions of the composing elements, further confirming the absence of single material domains. Fig. 6a and b show the LSV curves in 0.5 M H₂SO₄ and 1 M KOH, respectively, for the heterogeneous electrodes before and after the Li-TFSI treatment (electrodes named NbS₂:MoSe₂ and Li-TFSI-treated NbS₂:MoSe₂). In addition, we also plotted the LSV curves obtained for the untreated heterogeneous electrode after 1000 CV cycles (NbS₂:MoSe₂-CV @ 1000 cycles), the commercial Pt/C benchmark and the references NbS₂ and MoSe₂. The electrochemical characterization of the electrode based on solely MoSe₂ before and after Li-TFSI treatment, as well as after 1000 CV cycles, is reported in Fig. S9,† showing an improvement of the initial HER

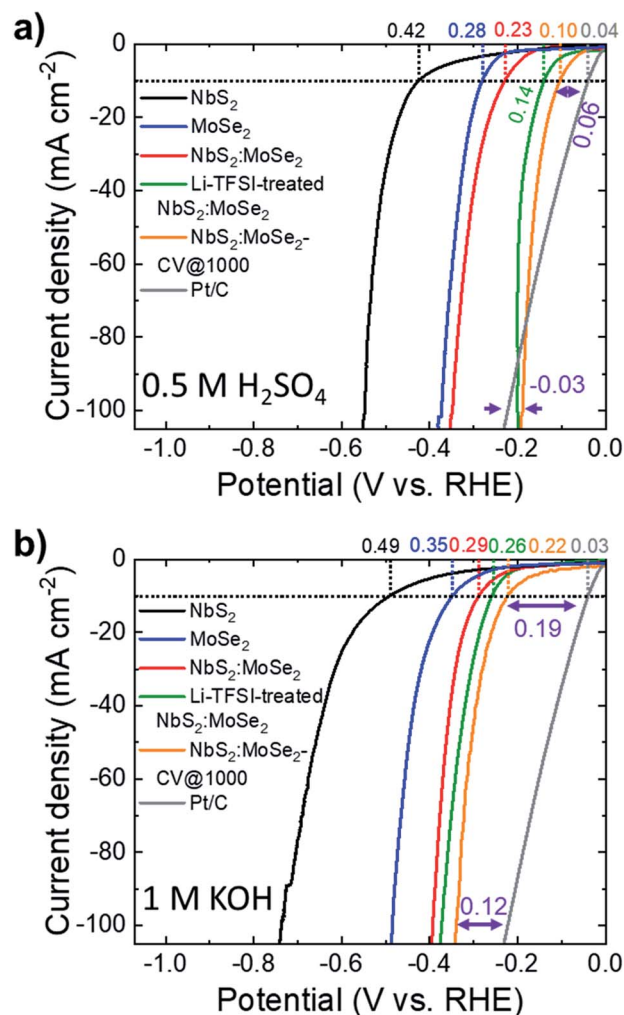


Fig. 6 (a and b) *iR*-corrected LSV curves for NbS₂:MoSe₂, Li-TFSI-treated NbS₂:MoSe₂ and NbS₂-CV @ 1000 cycles in acidic (0.5 M H₂SO₄) and alkaline (1 M KOH) solutions, respectively. The LSV curves of the Pt/C benchmark and NbS₂ and MoSe₂ references are also shown for comparison. The η_{10} values measured for the electrodes are also shown. For NbS₂:MoSe₂-CV @ 1000 cycles, the HER overpotential vs. the Pt/C electrode at 10 mA cm⁻² and 100 mA cm⁻² is indicated.

activity ($\eta_{10} = 0.28$ and 0.35 V in 0.5 M H₂SO₄ and 1 M KOH, respectively) after both Li-TFSI treatment ($\eta_{10} = 0.27$ and 0.30 V in 0.5 M H₂SO₄ and 1 M KOH, respectively) and electrochemical cycling (η_{10} equal to 0.18 and 0.3 V in 0.5 M H₂SO₄ and 1 M KOH, respectively). However, the beneficial effects of the electrode treatments are less pronounced compared to the case of NbS₂ electrodes. The heterogeneous NbS₂:MoSe₂ electrodes show a HER activity ($\eta_{10} = 0.23$ and 0.29 V in 0.5 M H₂SO₄ and 1 M KOH, respectively) which is higher than that of both NbS₂ and MoSe₂ electrodes. As reported for the individual counterpart-based electrodes, both the Li-TFSI treatment and the electrochemical cycling increase the HER activity of the heterogeneous electrodes. In 0.5 M H₂SO₄, the Li-TFSI-treated NbS₂:MoSe₂ electrode reaches an η_{10} of 0.14 V, approaching that of NbS₂:MoSe₂-CV @ 1000 cycles ($\eta_{10} = 0.10$ V). A similar trend was observed in 1 M KOH, in which the Li-TFSI-treated NbS₂:MoSe₂ electrode shows an η_{10} of 0.26 V, whereas



NbS₂:MoSe₂-CV @ 1000 cycles achieves an η_{10} of 0.22 V. At a current density of 100 mA cm⁻², NbS₂:MoSe₂-CV @ 1000 cycles displays a HER activity higher than that of the Pt/C benchmark in an acidic environment. Under alkaline conditions, the HER overpotential vs. the Pt/C benchmark (0.12 V) at 100 mA cm⁻² is reduced by 0.03 V compared to that measured at 10 mA cm⁻² (0.19 V).

Beyond the electrocatalytic activity, the long-term steady-state durability of the catalysts is another important criterion for their practical utilization. Fig. 7a shows the chronoamperometric measurements for the Li-TFSI-treated NbS₂:MoSe₂ electrodes at a fixed potential corresponding to an initial cathodic current density of 80 mA cm⁻². In 0.5 M H₂SO₄, the electrodes retained 97.3% of the initial current densities after 12 h. In 1 M KOH, the electrode improves the HER activity over time, increasing its initial current density by +22.2%. Fig. 7b shows the *iR*-corrected LSV curves of the electrodes measured

after the stability test, showing an η_{10} of 0.11 and 0.12 V in 0.5 M H₂SO₄ and 1 M KOH, respectively.

Under acidic conditions, the stability of the electrode is further verified by XPS analysis (Fig. S10a†), which does not show relevant chemical alterations of the NbS₂ and MoSe₂ flakes after the long-term test. Noteworthy, these results have been achieved without using any binder, such as Nafion, which could prospectively improve further the mechanical stability of the electrodes during H₂ evolution.^{140,141} In fact, although the gas evolution, causing the nanostructuring of the electrodes, has been correlated with the self-optimization behaviour of catalytic group-5 MX₂ in acidic media,^{54,55,57,77} it may imply a loss of catalytic materials, which should be controlled for practical purposes. In this context, the Li-TFSI-treated electrodes perform similarly to the electrochemically cycled electrodes, but do not require *in operando* pre-conditioning of the electrodes. Therefore, our approach to chemically treat the NbS₂-based electrodes could be beneficial for controlling the design of the group-5 MX₂ catalyst, as well as for their long-term durability. Under alkaline conditions, the remarkable decrease of η_{10} (from 0.26 to 0.12 V) is still under investigation. In fact, the XPS (Fig. S10†) data reveal a significant decrease of the NbS₂-related components (*i.e.*, those attributed to Nb(4+) and S 2p). These observations indicate a progressive oxidation and a possible dissolution of the NbS₂ flakes. The decrease of the Nb atomic content relative to that of Mo is further verified by SEM-EDS analysis with a clear intensity decrease of the Nb (K α) peak compared to the Mo (K α) peak (Fig. S11†). Moreover, the peaks of the doublet assigned to such Nb(5+) (\sim 207.4 and \sim 210.1 eV) are located at binding energy slightly lower than those of the doublet assigned to Nb₂O₅ in the as-produced NbS₂ flakes (typically at \sim 207.7 eV and \sim 210.4 eV,^{112–116} see Fig. S1†). Therefore, such XPS peaks could be ascribed to either sub-stoichiometric Nb₂O_{5–x} or hydr(oxy)oxide species. Noteworthy, the latter can synergistically interact with MoSe₂ flakes (which are chemically stable, see Fig. S10† and ref. 28 and 51), increasing the HER activity compared to that of the initial electrode (see Fig. 7a). In fact, it has been recently demonstrated that the presence of transition metal oxides (or hydroxides) on an MX₂ surface can also increase the HER activity of the pristine MX₂ under alkaline conditions,^{28,51,142} similarly to what was observed in noble metal-based electrocatalysts.^{143–145}

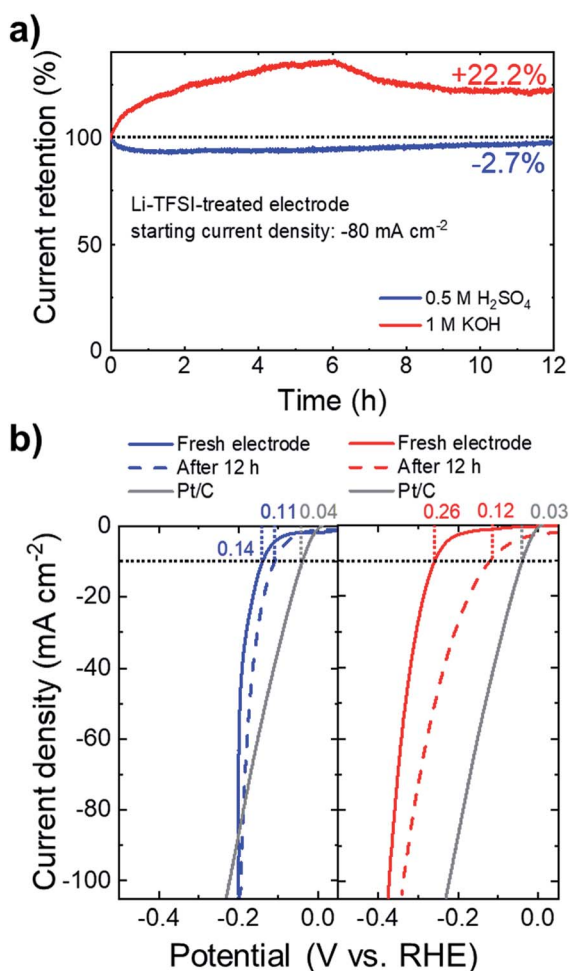


Fig. 7 (a) Chronoamperometry measurement of Li-TFSI-treated NbS₂:MoSe₂ in 0.5 M H₂SO₄ (blue line) and 1 M KOH (red line). The percentage increase of the starting cathodic current density (80 mA cm⁻²) after 12 h is also indicated. (b) The LSV curves of Li-TFSI-treated NbS₂:MoSe₂ in 0.5 M H₂SO₄ (left panel) and 1 M KOH (red panel). The LSV curves of the Pt/C benchmark are shown for comparison. The η_{10} values measured for the electrodes are also indicated.

Conclusion

In summary, we have produced phase-mixed NbS₂ flakes (*i.e.*, 1H-, 2H- and 3R-NbS₂ flakes) by an environmentally friendly liquid phase exfoliation (LPE) of synthesized NbS₂ crystals in 2-propanol (IPA). On the basis of a literature database for theoretically calculated Gibbs free energy of adsorbed atomic H (ΔG_{H}^0), NbS₂ flakes have been investigated as efficient HER catalyst candidates. Vacuum filtration has been used as a scalable approach to manufacture electrodes based on NbS₂ flakes. The as-produced electrodes were also chemically treated with a Li-TFSI bath in order to vertically orient the NbS₂ flakes and to induce the nanostructuring of their morphology, increasing the water accessibility to their surface. The full potential of NbS₂



flakes was exploited in a heterogeneous electrode obtained by mixing NbS₂ and MoSe₂ flakes. In fact, the flake stacking increases the HER activity of both the basal plane of the NbS₂ flakes and the metallic edge of MoSe₂, without affecting that of the catalytic sites of the single counterparts. In acidic solution (0.5 M H₂SO₄), the designed NbS₂-based electrocatalysts reach an overpotential at a cathodic current density of 10 mA cm⁻² (η_{10}) as low as 0.14 and 0.10 V after chemical Li-TFSI treatment and electrochemical cycling, respectively. At current density ≥ 100 mA cm⁻² the HER activity of the electrodes is higher than that of the Pt/C benchmark (overpotential reduced by 0.03 V). For the first time, the HER activity of NbS₂-based electrocatalysts is investigated in alkaline solutions (1 M KOH), showing a promising η_{10} of 0.26 and 0.22 V after chemical Li-TFSI treatment and electrochemical cycling, respectively. The durability of the electrodes is validated over 12 h of continuous operation at a fixed potential corresponding to an initial cathodic current density of 80 mA cm⁻². The results show an optimal performance retention of the chemically treated electrodes in 0.5 M H₂SO₄ (97.3%). In 1 M KOH, the electrodes significantly improve their initial cathodic current density (+22.2%), displaying an η_{10} of only 0.12 V after 12 h. However, chemical changes affect the NbS₂ flakes during HER operation, and further studies are needed to validate practical applications. Our results provide new insight into the exploitation of metallic group-5 TMDs for the HER under pH-universal conditions *via* scalable material preparation and electrode processing. The chemical treatment of the electrodes may be advantageous to their *in operando* electrochemical cycling, because it can be prospectively coupled with the use of proton conducting binders to avoid catalyst dissolution/loss.

Experimental

Materials

Nb (99.9%, <100 μ m) and S (99.999%, <6 mm) powders were purchased from Strem, USA. H₂SO₄ (99.999%), KOH ($\geq 85\%$ purity, ACS reagent, pellets), Pt/C (10 wt% loading) and Nafion solution (5 wt%) were purchased from Sigma Aldrich. The SWCNTs (>90% purity) were purchased from Cheap Tubes. The MoSe₂ crystal was purchased from HQ graphene.

Crystals' synthesis and exfoliation

NbS₂ crystals were made by direct synthesis from elements. An amount of Nb and S powders with a Nb : S stoichiometry of *ca.* 1 : 2 (1 wt% excess of S) and corresponding to a total mass of 10 g of NbS₂ was placed in a quartz glass ampoule (20 mm \times 120 mm) and sealed under high vacuum (1×10^{-3} Pa). The ampoule was heated at 450 $^{\circ}$ C for 12 h and subsequently at 600 $^{\circ}$ C for 48 h. Finally the derived product was treated at 900 $^{\circ}$ C for 48 h and cooled down to room temperature over a period of 24 h, obtaining NbS₂ crystal powder. The NbS₂ and MoSe₂ flakes were produced by LPE in IPA from the as-synthesized NbS₂ crystal powder and purchased MoSe₂ crystal,^{63,80} followed by SBS to remove the unexfoliated material by ultracentrifugation.^{81,82} Experimentally, 50 mg of bulk crystals were added to 50 mL of

anhydrous IPA and ultrasonicated in a bath sonicator (Branson® 5800 cleaner, Branson Ultrasonics) for 6 h. The resulting dispersions were ultracentrifuged at 2700 g (Optima™ XE-90 with a SW32Ti rotor, Beckman Coulter) for 20 min at 15 $^{\circ}$ C in order to separate un-exfoliated bulk crystals (collected as sediment) from the exfoliated materials that remained in the supernatant. Then, 80% of the supernatant was collected by pipetting, obtaining an exfoliated material dispersion. The concentration of the NbS₂ and MoSe₂ flake dispersions was 0.30 g L⁻¹ and 0.25 g L⁻¹, respectively.

Material dispersions' preparation

The dispersions of NbS₂ and MoSe₂ flakes were used as-produced. The heterogeneous dispersions of the NbS₂ and MoSe₂ flakes were obtained by mixing the single material component dispersions with a material weight ratio (w/w) of 1 : 1. The SWCNT dispersions were produced by dispersing SWCNTs in *N*-methyl-2-pyrrolidone (NMP) with a concentration of 0.2 g L⁻¹ by means of ultrasonication-based de-bundling.^{146–148} In particular, 10 mg of SWCNT powder was added to 50 mL of NMP. The dispersion was then ultrasonicated for 30 min by using a sonic tip (Vibra-cell 75185, Sonics) with the vibration amplitude set to 45%. The sonic tip was pulsed for 5 s on and 2 s off to reduce the solvent heating, which was also mitigated by using an ice bath around the vessel. The dispersion of Pt/C was produced by dissolving 4 mg of Pt/C and 80 μ L of Nafion solution in 1 mL of 1 : 4 v/v ethanol/water.

Materials' characterization

The SEM analysis of the as-synthesized crystals was carried out using a JEOL JSM-7500FA equipped with a cold FEG, operated at 10 kV accelerating voltage. The EDS analyses were performed using an Oxford X-Max 80 system equipped with an 80 mm² silicon drift detector (SDD) and using Oxford's AZtec TruMap software. The displayed SEM-EDS maps were obtained by background subtraction and peak deconvolution in order to correct for the overlapping S (K α peak) and Nb (L β peak) features in the EDS spectrum. The surface of the samples was cleaned by mechanical exfoliation and imaged without any conductive coating. The BF-TEM images of the NbS₂ flakes were acquired with a JEM 1011 (JEOL) TEM (thermionic W filament), operated at 100 kV. The morphological and statistical analysis was performed by using ImageJ software (NIH) and OriginPro 9.1 software (OriginLab), respectively. The samples for the TEM measurements were prepared by drop casting the as-prepared exfoliated material dispersions onto ultrathin C-on-hole C-coated Cu grids and rinsed with deionized water and subsequently dried overnight under vacuum. The HRTEM, HAADF-STEM imaging and STEM-EDS analyses were carried out on an image-Cs-corrected JEOL JEM-2200FS TEM (Schottky emitter), operated at 200 kV, equipped with an in-column image filter (Ω -type) and a Bruker XFlash 5060 EDS detector. For these analyses, the NbS₂ flake dispersion was drop cast onto a holey-carbon-coated Cu grid. The displayed STEM-EDS maps are "quantitative", in analogy to what was done for SEM-EDS maps of the crystals.



The AFM images were acquired with a Nanowizard III (JPK Instruments, Germany), mounted on an Axio Observer D1 (Carl Zeiss, Germany) inverted optical microscope. The AFM measurements were carried out by means of PPP-NCHR cantilevers (Nanosensors, USA) having a nominal tip diameter of 10 nm. We used a drive frequency of ~ 295 kHz. We collected intermittent contact mode AFM images (512×512 data points) of $2.5 \times 2.5 \mu\text{m}^2$ by keeping the working set point above 70% of the free oscillation amplitude. We used a scan rate of 0.7 Hz for the acquisition of the images. JPK Data Processing software (JPK Instruments, Germany) was exploited to process the height profiles, while the data were analysed by using OriginPro 9.1 software. The latter was also used to carry out the statistical analysis on multiple AFM images for all the tested samples. The samples were prepared by drop-casting the as-prepared exfoliated material dispersions onto mica sheets (G250-1, Agar Scientific Ltd.) and dried overnight under vacuum. The XRD measurements were acquired with a PANalytical Empyrean using Cu K_α radiation. The samples for XRD were prepared by drop-casting the as-prepared exfoliated material dispersions onto Si/SiO₂ substrates and dried overnight under vacuum. Raman spectroscopy measurements were performed by using a Renishaw microRaman inVia 1000 mounted with a $50\times$ objective, with an excitation wavelength of 532 nm and an incident power of 1 mW on the samples. For each sample, we collected 50 spectra. The samples were prepared by drop casting the as-prepared exfoliated material dispersions onto Si/SiO₂ substrates and subsequently dried under vacuum. The XPS characterization was performed with a Kratos Axis UltraDLD spectrometer, having a monochromatic Al K_α source (15 kV, 20 mA). The spectra were acquired on a $300 \times 700 \mu\text{m}^2$ area. A constant pass energy of 160 eV and an energy step of 1 eV were used to collect wide scans. High-resolution spectra were acquired at a constant pass energy of 10 eV with an energy step of 0.1 eV. The binding energy scale was referenced to the C 1s peak at 284.8 eV. The spectra were analysed using CasaXPS software (version 2.3.17). The samples were prepared by drop casting the NbS₂ flake dispersions onto a Si/SiO₂ substrate (LDB Technologies Ltd), followed by a drying process under vacuum.

Electrodes' fabrication

The electrodes were produced by sequentially depositing the SWCNT and exfoliated catalytic material (NbS₂, MoSe₂ and heterogeneous NbS₂:MoSe₂) dispersions onto nylon membranes (Whatman® nylon membrane filters, 0.2 μm pore size) through a vacuum filtration process.⁴⁸ The material mass loading was 1.31 mg cm^{-2} for SWCNTs and 0.2 mg cm^{-2} for the exfoliated catalytic materials. The electrode area was 3.8 cm^2 . Additional electrodes were produced by depositing the as-produced NbS₂ flake dispersion onto glassy carbon by a drop casting method (NbS₂ flake mass loading = 0.5 mg cm^{-2}). Before the electrochemical characterization, the as-prepared electrodes were dried overnight at room temperature. The electrodes were chemically treated with Li-TFSI inside a glove box. Experimentally, 2 mg of Li-TFSI (Sigma-Aldrich) were dissolved in 10 mL of anhydrous acetonitrile (ACN) (Sigma-Aldrich)

and stirred for 15 min at room temperature to make a 0.2 g L^{-1} solution. Subsequently, the electrodes were immersed in the Li-TFSI solution in a closed vial for 20 min at a temperature of 70°C . The electrodes were removed (without rinsing) and subsequently annealed at 100°C for 5 min in air in order to remove the residuals of ACN. Electrodes made entirely of SWCNTs were also produced as a reference with a mass loading of 1.31 mg cm^{-2} . A Pt/C electrode was produced as a benchmark for the HER by depositing the corresponding dispersion onto glassy carbon substrates (Sigma Aldrich). The active material mass loading of the electrodes was 0.262 mg cm^{-2} for Pt/C, in agreement with previously reported protocols.^{28,51}

Electrode characterization

The XPS analyses of the as-produced electrodes and the electrodes after chemical and electrochemical treatments were performed with the same set-up and parameters used for the characterization of the materials. The detailed SEM images of the electrodes were collected on a Helios Nanolab® 600 Dual-Beam microscope (FEI Company) and 5 kV and 0.2 nA as measurement conditions. The SEM-EDS measurements were performed on a JEOL JSM-6490LA SEM at 30 kV. The electrochemical measurements were performed at room temperature in a flat-bottom fused silica cell using the three-electrode configuration of a potentiostat/galvanostat station (VMP3, Bio-logic), controlled *via* own software. A glassy carbon rod and saturated KCl Ag/AgCl were employed as the counter electrode and reference electrode, respectively. The measurements were carried out in 200 mL of 0.5 M H₂SO₄ or 1 M KOH. Before starting the measurements, the oxygen was purged from the electrolyte by flowing N₂ gas throughout the liquid volume using a porous frit. The Nernst equation: $E_{\text{RHE}} = E_{\text{Ag/AgCl}} + 0.059 \times \text{pH} + E_{\text{Ag/AgCl}}^0$, where E_{RHE} is the converted potential *vs.* RHE, $E_{\text{Ag/AgCl}}$ is the experimental potential measured against the Ag/AgCl reference electrode, and $E_{\text{Ag/AgCl}}^0$ is the standard potential of Ag/AgCl at 25°C (0.1976 V vs. RHE), was used to convert the potential difference between the working electrode and the Ag/AgCl reference electrode to the reversible hydrogen electrode (RHE) scale. The double-layer capacitances (C_{dl}) of the untreated and treated NbS₂ films deposited onto glassy carbon were estimated by cyclic voltammetry (CV) measurements in a non-faradaic region of potential (between 0.2 and 0.4 V *vs.* RHE) at various potential scan rates (ranging from 20 to 600 mV s⁻¹). Cyclic voltammetry was also exploited for the electrochemical treatment of the electrodes. Experimentally, 1000 CV scans were carried out between 0.25 and 0.75 V *vs.* RHE at a potential scan rate of 100 mV s⁻¹. The LSV curves of the electrodes using a SWCNT substrate were acquired at a 5 mV s⁻¹ scan rate and were *iR*-corrected by considering *i* as the measured working electrode current and *R* as the series resistance arising from the working electrode substrate and electrolyte resistances. Electrochemical impedance spectroscopy (EIS) measurements were performed on the electrodes at open circuit potential and at a frequency of 10 kHz to measure *R* (in agreement with methods reported in previous studies).^{48,50,51} The stability tests were performed by chronoamperometry



measurements, *i.e.*, by measuring the current in potentiostatic mode over time, at a fixed overpotential corresponding to a cathodic current density of 80 mA cm^{-2} .

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

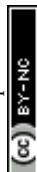
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