



Cite this: *Sustainable Energy Fuels*, 2019, 3, 888

Received 27th November 2018
Accepted 6th February 2019

DOI: 10.1039/c8se00581h

rs.c.li/sustainable-energy

Solid-state p-type dye-sensitized solar cells: progress, potential applications and challenges

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The fabrication of solid-state p-type dye-sensitized solar cells (p-ssDSCs) using electron transport materials instead of the conventional I^-/I_3^- redox couple in liquid devices can completely eliminate the liquid phase and enhance the photovoltage. The performance of p-ssDSCs has been improved by optimizing the dyes and electron-transport materials. Moreover, p-ssDSCs provide a possibility to fabricate solid-state tandem dye-sensitized solar cells and show promising application in dye-sensitized solar fuel devices as well. Herein, the development of p-ssDSCs has been overviewed, their potential applications have been discussed, and the challenges remaining in p-ssDSCs have been highlighted.

Introduction

The p-type dye-sensitized solar cell (p-DSC) was first reported in 1999 by Lindquist and co-workers,¹ in which a porphyrin dye was used as a photosensitizer, nanoporous NiO acted as a large band gap p-type semiconductor to load the dye, and a liquid I^-/I_3^- electrolyte was adopted as a redox mediator to undertake electron transport between the counter electrode and photocathode. Subsequently, p-DSCs have been attracting scientists' interest due to their potential applications in the preparation of tandem dye-sensitized solar cells^{2,3} in cooperation with conventional n-type DSCs (n-DSCs)⁴ as well as in the fabrication of dye-sensitized solar fuel devices (DSSFDS).⁵⁻⁹ However, the p-DSCs with NiO as a p-type semiconductor and the I^-/I_3^- redox couple always suffer from low photovoltages due to small energy difference (ΔE) between the valence band (VB) of NiO and the redox potential of the I^-/I_3^- redox couple.¹⁰ Some studies have been focused on the development of new p-type semiconductors with respect to NiO;¹¹⁻¹³ however, NiO still provides a higher photocurrent and has more facile preparation process than other materials. Researchers have also tried to use other redox couples, such as tris(1,2-diaminoethane)cobalt(III)/(II)¹⁴ and tris(acetylacetonato)iron(III)/(II),¹⁵ with more negative reduction potentials than I^-/I_3^- to enlarge the ΔE . This kind of strategy indeed can improve the photovoltage without scarifying the photocurrent of the p-DSCs. However, such redox couples are always sensitive to oxygen, which makes the solar cell fabrication process more complex. Also, the system still contains organic solvents, with a potential issue of electrolyte leaking from the device. On the basis of this background, we proposed a concept of solid-state p-type dye-sensitized solar cells (p-ssDSCs) in 2016¹⁶ in order to eliminate the liquid phase

existing in the conventional p-DSCs as well as to improve the photovoltage of p-DSCs. Some follow-up work further proved the concept of the p-ssDSCs and made progress toward improving the efficiency of this kind of device from different angles.¹⁷⁻²⁰ The p-ssDSCs also showed potential applications in DSSFDS and tandem solar cells. In this perspective, I give an overview of the basic concept and progress of p-ssDSC, discuss their potential applications, and point out the challenges still left to be overcome.

Basic concept of p-ssDSCs

Fig. 1 shows the configurations of a conventional liquid p-DSC and a p-ssDSC. In a p-ssDSC, a solid-state electron-transport material (ETM) is used instead of the liquid electrolyte, such as I^-/I_3^- , in a liquid device to undertake dye regeneration and electron transport between the dye-sensitized photocathode and back contact. The ETM should be able to infiltrate in the nanoporous p-type semiconductor (p-SC) film, such as NiO film, with a good contact with the dye. This requires that the ETM should have good solubility in the solvent for the solution-process or it may be *in situ* prepared in the p-SC film. In order to obtain a satisfactory photovoltage from a p-ssDSC, the ΔE between the VB of the p-SC and reduction potential of the ETM should be as large as possible, but the reduction potential of the ETM has to meet the need for dye regeneration. The ETM should also hold good electron mobility to efficiently transport electrons to the back contact.

Proof of concept

In our proof-of-concept work for the p-ssDSCs,¹⁶ we chose [6,6]-phenyl-C61-butyric acid methyl ester (PCBM) as the ETM and a standard organic dye P1²¹ as the photosensitizer (Fig. 2a) to fabricate our p-ssDSCs. PCBM is a benchmark organic ETM

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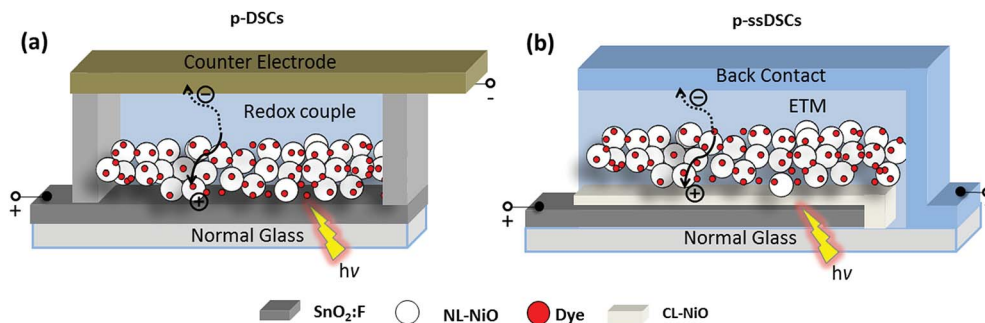


Fig. 1 Schematic of the configurations of a conventional liquid p-DSC (a) and a solid-state device p-ssDSC (b).

used in organic solar cells,²² showing electron mobility of $ca. 2 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$.²³ It also has excellent thermal stability with a crystallization temperature and melting point of 195 °C and 290 °C, respectively.²³ Its good solubility in many solvents makes it possible to fabricate solar cell by a low energy-consumption method, *i.e.*, a *solution process*. Moreover, PCBM holds its first reduction potential at -0.4 V vs. a Normal Hydrogen Electrode (NHE), which is more negative than that of the I^-/I_3^- redox couple, 0.3 V vs. NHE.¹⁶ The dye generation is also thermodynamically feasible, as the driving force for PCBM to regenerate P1 dye is 0.4 V . Theoretically, PCBM as the ETM in NiO-based p-ssDSCs should render a much higher photovoltage (V_{OC} , 0.9 V) than that of the liquid device using the I^-/I_3^- redox couple (0.2 V), as shown in Fig. 2b.

Different from liquid p-DSCs, in the p-ssDSCs, a compact NiO layer (see Fig. 1) should be placed between the FTO ($\text{SnO}_2:\text{F}$) and nanoporous NiO layer, in order to block the physical contact between the FTO and the penetrated PCBM for suppressing the serious hole and electron recombination. With 100 nm Al as the back contact, a device based on NiO/P1/PCBM rendered a photocurrent density (J_{SC}) of $50 \mu\text{A cm}^{-2}$ and a photovoltage (V_{OC}) of 620 mV (Fig. 3a). Without the dye, the NiO/PCBM-based device also showed a photovoltaic effect from a pure p-n junction, but the performance was much lower than that obtained from the NiO/P1/PCBM-based device. With help from transient absorption spectroscopy (TAS), we were able to monitor the kinetics of dye bleaching with and without PCBM (Fig. 3b). The result showed the P1 dye could be regenerated by PCBM. We also detected $\text{PCBM}^{\cdot-}$ around 1000 nm , which

survived for a time period of μs , indicating a slow charge recombination between holes in NiO and electrons in $\text{PCBM}^{\cdot-}$. In an efficient organic solar cell, the thickness of the active layer containing PCBM is normally kept at $ca. 100 \text{ nm}$.^{24,25} Too thick an active layer will limit the obtainable photocurrent due to the low electron mobility of PCBM. In p-ssDSCs, the thickness of PCBM is mainly dominated by the thickness of the nanoporous NiO film, as an ideal case is one where all pores of the NiO film can be filled with PCBM to enhance regeneration of all the dyes. However, too thin a NiO film will result in poor dye loading, leading to an unsatisfactory light-harvesting efficiency. Therefore, in our first PCBM-based p-ssDSCs, the thickness of the NiO was optimized to be 700 nm . The unsatisfactory photovoltaic performance of the first p-ssDSC was therefore mainly due to the thick PCBM layer ($>700 \text{ nm}$). Seeking new alternative ETMs with high electron mobility therefore became an important way to optimize the p-ssDSCs.

Optimization of p-ssDSCs

Photosensitizer

The photosensitizer is one of the key components in p-ssDSCs undertaking light absorbing and charge generation. A high extinction coefficient and broad absorption spectrum are general requirements for the photosensitizer used in p-ssDSCs. Moreover, a long excited state and reduced state of the photosensitizer are beneficial for charge generation in p-ssDSCs.

The extinction coefficient of P1 dye is $5.8 \times 10^4 \text{ M}^{-1} \text{ cm}^{-1}$ at 468 nm in ethanol.²¹ If the value could be increased to double, the thickness of the NiO film used in p-ssDSCs, in principle,

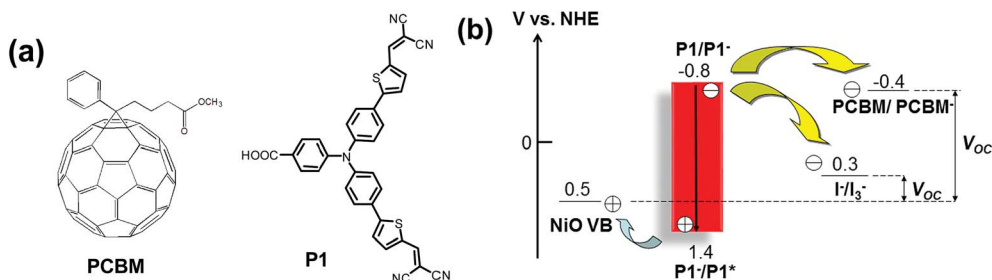


Fig. 2 The molecular structures of PCBM and P1 dye (a) and potential diagram of p-ssDSCs based on a PCBM ETM and liquid p-DSCs with the I^-/I_3^- redox couple (b).



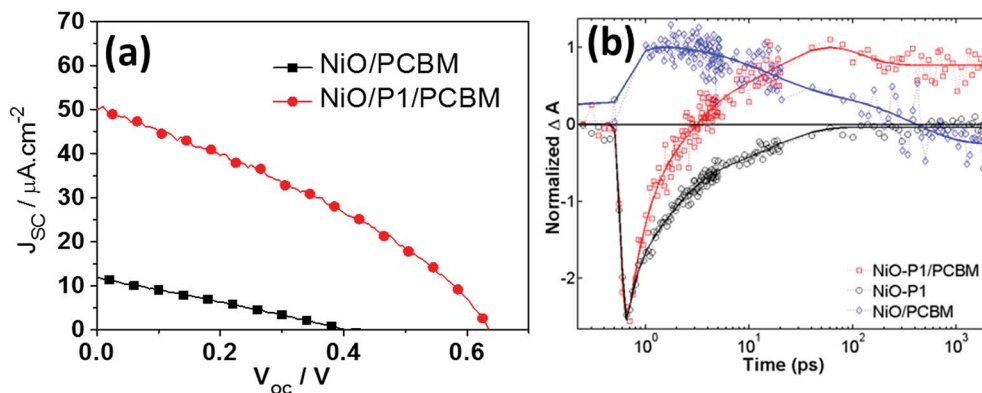


Fig. 3 J - V curves of the first p-ssDSC (a) and comparison kinetics traces of samples probed at 480 nm for dye bleaching (b), the blue diamonds are NiO/PCBM, red squares are NiO-P1/PCBM, and black cycles are NiO-P1, the solid line is a multi-exponential fit for each kinetic trace (ref. 16).

could be reduced to half without sacrificing the light-harvesting efficiency. In this case, the photocurrent of the device should be improved due to the thickness of PCBM layer becoming thin accordingly. On the other hand, the reduced P1 dye has a very short lifetime, *ca.* 20 ps,²⁶ indicating a fast charge recombination between the injected holes in NiO VB and the reduced dye. To address this issue, there are two solutions. One is to accelerate the dye regeneration process, making the ETM efficiently obtain electrons from the reduced dyes to compete with the unwanted recombination process. This relies on the ETM materials if we don't change the dye, which will be discussed in the following section. The other is to design new dyes, which should have long lifetimes of the excited and reduced states, thus facilitating the charge generation from the dye. Bouclé, Odobel, and co-workers developed two diketopyrrolopyrrole (DPP)-based dyes: **DPP-PYRO** and **DPP-Br**.¹⁷ The structures of both dyes are shown in Fig. 4. DPP-based dyes have been proved to have a long-lived intramolecular charge-separated state as well as a long-lived reduced state up to 205 μ s.^{27,28}

DPP-PYRO and **DPP-Br** also have more negative reduction potentials, -1.18 and -1.26 V vs. NHE, respectively, than P1 dye, rendering a large driving force to reduce an ETM, such as PCBM. To prove this, the dye was sensitized on the large band gap semiconductor Al_2O_3 , on which the dye was not expected to inject either holes or electrons in Al_2O_3 . By monitoring the photoluminescence (PL) quenching experiment of the dye on

Al_2O_3 upon adding PCBM, PL quenching values of 85% and 70% were found for **DPP-Br** and **DPP-PYRO**, respectively; while that of the P1 dye was limited to around 30%. This proved that these two DPP dyes indeed had more efficient charge transfer to PCBM than the P1 dye.

With PCBM as the ETM, both **DPP-Br** and **DPP-PYRO** dyes showed higher V_{OC} and J_{SC} values than the P1 dye. Notably, the V_{OC} obtained from P1 dye in this work, *ca.* 120 mV, was much lower than the value obtained in our previous work, 620 mV.¹⁶ The authors attributed this drop of V_{OC} in their work to the poor quality of the NiO blocking layer according to dark current measurement, which indicated serious current leaking from the device. This implied that the performance of the DPP dye-based p-ssDSCs could be further improved by optimizing the blocking layer. The higher J_{SC} values from these two dyes than that of P1 should be due to their efficient charge transfer reactions, since the P1 dye has a much larger extinction coefficient (by a factor of 2) than **DPP-Br** and **DPP-PYRO**. A **DPP-Br** based device rendered the highest performance among these dyes (Fig. 5), showing a J_{SC} of 0.45 mA cm^{-2} , a V_{OC} of 198 mV, a fill factor (ff) of 0.32, and an overall efficiency of 0.028%. This is an important study to show the possibility of improving the p-ssDSCs performance by finely tuning the dye structure.

Electron transport materials

The poor electron mobility of PCBM significantly limits the maximally obtainable efficiency of the p-ssDSCs. Reducing the

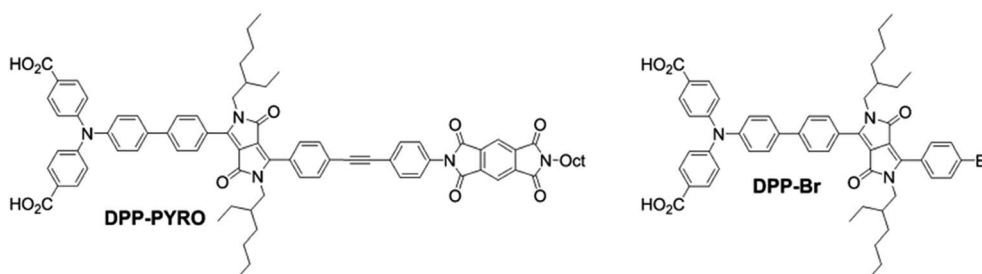


Fig. 4 The molecular structures of DPP-PYRO and DPP-Br dyes.



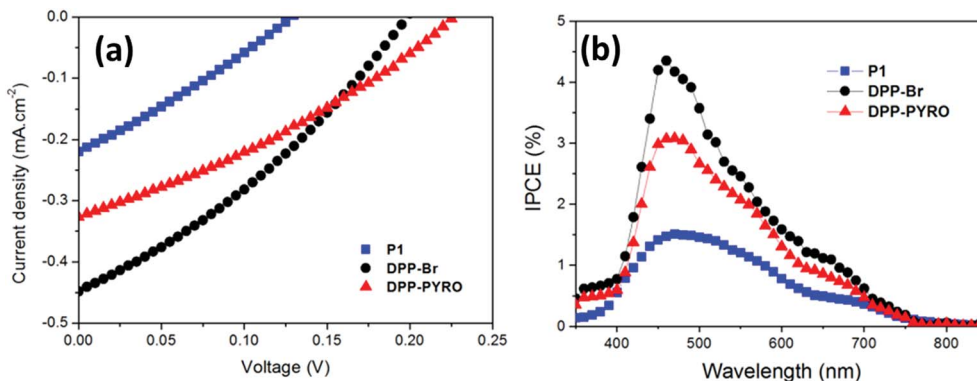


Fig. 5 The J - V curves (a) and IPCE spectra (b) of p-ssDSCs using DPP-PYRO, DPP-Br, and P1 dyes (adapted with permission from ref. 17. Copyright 2017 American Chemical Society).

PCBM layer thickness is not a practical solution to this problem, since a certain thick NiO film is always required in order to load enough dyes for efficient light harvesting. Alternatively, using an ETM with a high electron mobility instead of PCBM could enable great progress toward improving the performance of the p-ssDSCs. TiO₂ has been successfully used as an n-type semiconductor (n-SC) in conventional n-DSCs, showing excellent performance even with a thickness of up to 10 μm.²⁹ TiO₂ film has shown an electron mobility of 0.01–0.1 cm² V⁻¹ s⁻¹,^{30,31} which is 1–2 orders of magnitude higher than that of PCBM film. Moreover, TiO₂ CB can have a potential of *ca.* −0.5 V *vs.* NHE at pH 7,³² which is similar to the reduction potential of PCBM, −0.4 V *vs.* NHE. Therefore, we proposed that TiO₂ should be a good alternative ETM to PCBM in p-ssDSCs. However, considering the need for good pore filling in dye-sensitized NiO films, how to prepare TiO₂ in an NiO film was a challenge to us. Solution-processing was the first method considered by us. Unfortunately, the commonly used solution-processable TiO₂ precursor, titanium tetrachloride (TiCl₄), is a strong acid. The dye became severely decomposed in our initial attempt when the dye-sensitized film was immersed into these compounds solutions for long time at a high temperature. Under these circumstances, we instead chose to use atomic layer deposition (ALD) to grow the TiO₂ layer directly in the dye-sensitized NiO film¹⁸ as the ALD technique can use a short pulse of TiO₂ precursor TiCl₄ for the experiment and the temperature of formation of TiO₂ can be mild. Still, to make sure the dye was stable during the ALD process, we abandoned the traditional p-type dyes containing a malonitrile acceptor, as we found such dyes showed a decrease in absorbance after ALD measurement more or less, probably due to the degradation of cyano units by the precursor of ALD TiO₂ used in our experiment. Moreover, the conventional triphenylamine dyes with a malonitrile acceptor, such as P1-type dyes,⁹ have hardly distinguishable absorption spectra of its reduced and oxidized states, which is not a good sign for us to study charge-transfer processes in the device. Therefore, we synthesized PB6 dye (inset in Fig. 6a) using two carboxylic acid units as anchoring groups and perylene-monoimide (PMI) as an acceptor, which is thermally stable and has distinguishable absorption spectra of its reduced

(620 nm) and oxidized state (920 nm). PB6 dye can thermodynamically inject holes and electrons into NiO VB and TiO₂ CB, respectively (Fig. 6b). Eventually, we had success preparing a PB6-sensitized nanoporous NiO film, in which a thin TiO₂ layer was deposited by ALD to cover the dye on the NiO particles to maintain the nanoporous structure. The TiO₂ layer as a shell could physically contact with the dye in this configuration. We therefore named it as dye-sensitized NiO-dye-TiO₂ core-shell film.

From scanning electron microscopy (SEM) images of a cross-section of the prepared NiO-dye-TiO₂ film (Fig. 7), we could clearly distinguish FTO glass, the NiO blocking layer, and the nanoporous NiO layer. With help from energy-dispersive X-ray spectroscopy (EDX) mapping images, as shown in Fig. 7, we were able to observe the presence of Ti element across the entire NiO film. However, it is also obvious that there are more Ti element nearby the top region than close to the FTO surface. This implies that it is possible for ALD TiO₂ to penetrate inside of the NiO nanoporous film, but the nanopores of the NiO film still limit the uniformity of the TiO₂ layer formed. Unfortunately, we did not observe any crystalline structure of the TiO₂ prepared by ALD in the NiO film, probably indicating that the prepared TiO₂ on NiO particles in fact holds an amorphous structure (confirmed by TEM), or only a small portion are crystalline which could not be clearly observed.

In order to investigate if the PB6 dye can kinetically perform charge separation into two different semiconductors upon light illumination, we used photon-induced absorption (PIA) spectroscopy and TAS to monitor the charge-separation processes or intermediators. Using ZrO instead of NiO, we were able to catch the oxidized PB6 in the presence of the ALD TiO₂ layer from the PIA spectrum, indicating that the dye could inject electrons to TiO₂ just through physical contact. In the NiO-dye-TiO₂ system, formation of the reduced PB6 after hole injection into NiO from the excited dye could be clearly seen within 200 fs (Fig. 8a) in the TAS measurement. TAS produced pronounced spectra induced by the Stark effect due to an electrical field built between NiO and TiO₂ after electron injection from the reduced dye into TiO₂ (Fig. 8b). This process happened within 500 fs and lasted for more than 2 ns, suggesting that an ultrafast dye regeneration



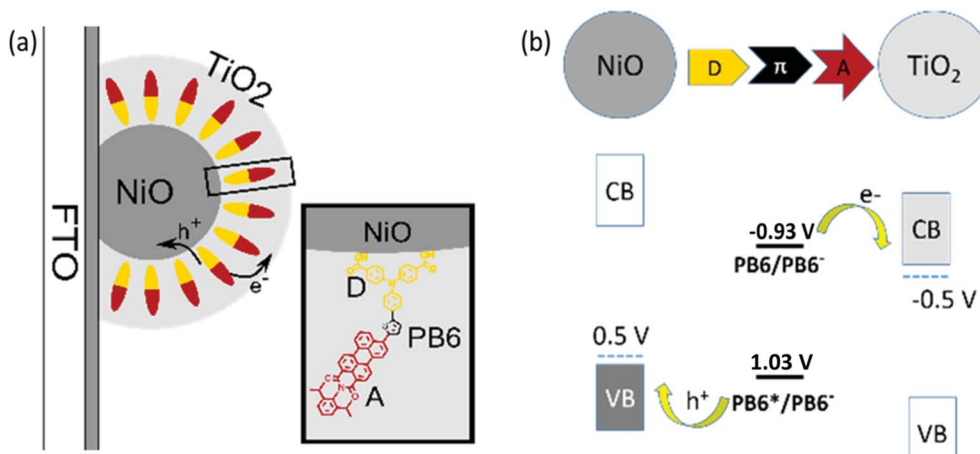


Fig. 6 (a) Schematic drawing of the well-orientated mesoporous NiO-dye-TiO₂ core-shell structure and the donor-acceptor dye (PB6). (D) Electron-donor part, triphenylamine (TPA); (A) electron acceptor, perylene monoimide (PMI). (b) Potential diagram vs. NHE of PB6 between NiO and TiO₂. (CB) Conduction band; (VB) valence band; (π) conjugated linker (ref. 18).

process and long-lived charge separation existed between NiO and TiO₂.

With these positive results, there was no doubt for us to move one step further to build p-ssDSCs by simply evaporating an Au layer as a back contact directly on the TiO₂ layer in the core-shell film to collect electrons.¹⁹ The p-ssDSCs based on NiO-PB6-TiO₂ indeed showed a photovoltaic effect (Fig. 9a). It rendered a V_{OC} of 400 mV. However, the J_{SC} was still low, only 12 $\mu\text{A cm}^{-2}$. Considering that charge recombination was probably caused by direct contact between NiO and TiO₂, we introduced a very thin (*ca.* 1 nm) Al₂O₃ layer after dye sensitization on NiO to isolate them and at the same time to protect the dye. Both the J_{SC} and V_{OC} of the p-ssDSCs based on NiO-PB6-Al₂O₃-TiO₂ were significantly improved as compared to that of the NiO-PB6-TiO₂-based devices. These improvements were attributed to the prolongation of the charge-separation lifetime in the presence of the inner barrier layer-Al₂O₃, as shown in Fig. 9b.

Interestingly, varying the thickness of the Al₂O₃ layer significantly influenced the solar cell performance (Fig. 10). The thickness of the Al₂O₃ layer was estimated from the ALD cycles. A 1 nm Al₂O₃ layer showed the optimal performance of the solar cell. From the molecular structure simulated from ChemDraw 3D, we concluded that an avg. 1 nm Al₂O₃ almost covers the donor moiety of the PB6 dye if the dye is assumed to stand on

the NiO surface. Too thin an Al₂O₃ layer can not effectively suppress the charge recombination and too thick an Al₂O₃ layer completely buries the dye and makes it impossible to communicate with the TiO₂ ETM layer, thus giving unsatisfactory performance.

Although TiO₂ as the ETM worked in p-ssDSCs, the performance of the device was still far behind what we expected. There were two main reasons for this: one is that the TiO₂ layer prepared in NiO film was not crystalline, at least as proved from the TEM images, which limited the electron transport in the TiO₂ layer; the other is that the TiO₂ layer was not uniform, in which case the discontinuity of the TiO₂ layer made electrons become trapped in the isolated TiO₂ islands, resulting in a poor electron-collection efficiency. In order to address these issues, we attempted to use ZnO instead of TiO₂ as the ETM in p-ssDSCs.²⁰ ZnO can form a crystalline structure from a molecular complex at a much lower temperature than that needed for formation of the crystalline TiO₂.³³ Also, ZnO can be prepared by solution-processing from an organic dye-friendly compound, such as Zn(II) acetylacetonate (Zn(acac)₂).³³ The solution process is beneficial for improving the poor filling of NiO film and for having a continuous ZnO layer on the NiO surface. More important, ZnO film has shown excellent electron mobility of 1–10 $\text{cm}^2 \text{V}^{-1} \text{s}^{-1}$,^{34,35} which is 2–3 orders of magnitude higher

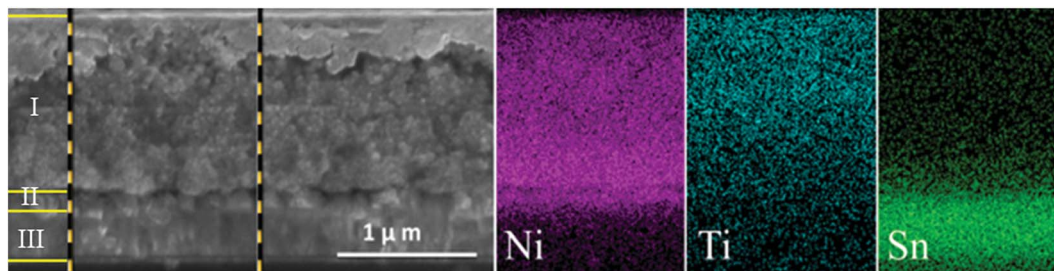


Fig. 7 SEM and EDX images of a cross-section of the NiO-PB6-TiO₂ film: (I) mesoporous NiO layer coated by TiO₂; (II) NiO blocking layer; (III) FTO layer, and the corresponding EDX element mapping analysis (ref. 18).



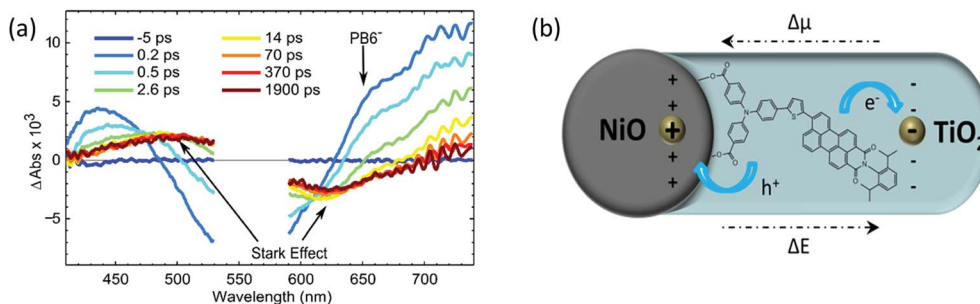


Fig. 8 (a) Transient absorption spectra of NiO-PB6-TiO₂ film from -5 ps to 1.9 ns excited by a pulse laser of 560 nm, 300 nJ per pulse; (b) schematic drawing of the electric field formed after charge separation, $\Delta\mu$ is the difference in dipole moment and ΔE is the difference in the electrical field (ref. 18).

than that of TiO₂ film. Bearing these in mind, we fabricated the p-ssDSCs by using ZnO ETM prepared from a thermal method. The fabrication process is shown in Fig. 11. Notably, we could grow the ZnO layer in the NiO film *in situ* by sintering the spin-coated Zn(acac)₂ at 200 °C. Thanks to the excellent thermal stability of the PB6 dye, there was no obvious degradation of the dye after thermally growing the ZnO layer. However, we had to use Al instead of Au as the back contact in the solution-processed ZnO-based p-ssDSCs, because we found the Au easily caused device short-circuit. The reason for this is still unclear at this moment. One hypothesis is that the solution-processed ZnO could not be as compact as the TiO₂ layer prepared by ALD, thus giving a possibility for the heavy Au atoms to pass through and then contact with NiO, and maybe the FTO as well.

From the SEM and EDX mapping images of a cross-section of a NiO-PB6-ZnO solar cell (Fig. 12), we can clearly see different layers in the solar cell. In particular, we can see that the Zn element is uniformly spread across the entire NiO film, confirming that good infiltration of the ZnO precursor in the pores of NiO film can be obtained by the solution-processing method. Moreover, the formed ZnO was proved to be crystalline structure by TEM and X-ray powder diffraction (XRD) measurements. Excitingly, the p-ssDSCs using ZnO as the ETM showed an enhanced J_{SC} of up to 680 $\mu\text{A cm}^{-2}$, which was 60-fold higher than that of the previously reported device using TiO₂ ETM with

a similar architecture. Although the J_{SC} of the NiO-PB6-ZnO solar cell was still lower than that of the liquid device using an identical NiO film, the V_{OC} of the NiO-PB6-ZnO solar cell, 440 mV, was much higher than that of the liquid device, 120 mV. Eventually, the solid-state device rendered a higher efficiency of 0.135% than that of the liquid device, 0.106%. The p-ssDSCs with this configuration also showed good stability due to the sandwich structure in which the dye is well protected by two inorganic materials (Fig. 13).

Potential applications

In addition to the fabrication of p-ssDSCs, the dye-sensitized photocathodes with (p-SC)-dye-(n-SC) configuration have two more potential applications: dye-sensitized solar fuel devices (DSSFs) and tandem solid-state dye-sensitized solar cells (t-ssDSCs).

(a) Dye-sensitized solar fuel device

Using the concept and configuration of DSCs to build up DSSFs⁵⁻⁸ has been attracting more interest from scientists. In DSSFs, the dye-sensitized photoanode is used to work with a catalyst for oxidation reactions, such as water oxidation, and the photocathode is adopted to work with a catalyst for reduction reactions, such as proton reduction and CO₂ reduction. For

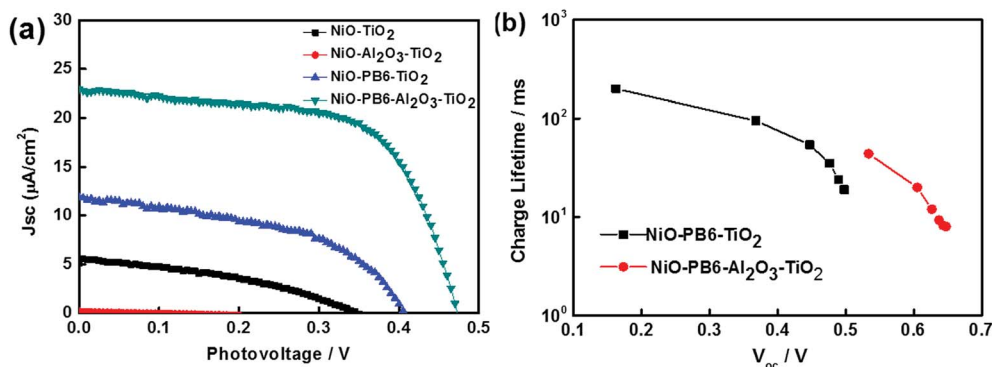


Fig. 9 J - V curves of different NiO solar cells (a) and electron lifetime as a function of V_{OC} density of NiO-PB6-Al₂O₃-TiO₂- and NiO-PB6-TiO₂-based p-ssDSCs (b) (ref. 19).



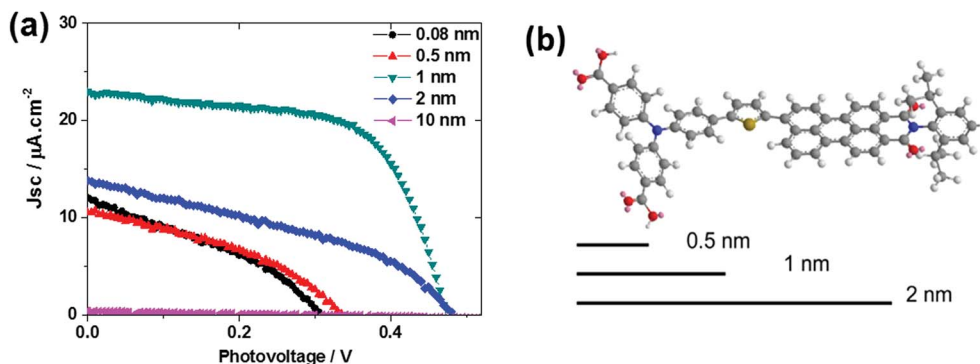


Fig. 10 (a) J - V curves for NiO-PB6-Al₂O₃-TiO₂ solar cells with different Al₂O₃ thickness layers; (b) 3D molecular structure of PB6 with 0.5 nm, 1 nm, and 2 nm length scales as a comparison (ref. 19).

the photoanode, it has reached 1.7 mA cm⁻² photocurrent at a bias potential of 0.2 V vs. NHE for water oxidation.³⁶ The performance of the photocathode for the reduction reaction is still too low to match the photoanode. In order to optimize the photocathode, catalysts and photosensitizers have been arranged on photocathodes by co-sensitization,³⁷⁻³⁹ using a supramolecule,^{40,41} and covalent linkage.^{9,42,43} However, few of them could significantly improve the performance for proton or CO₂ reduction. Since hole injection from the dye into NiO has been proved as fast enough (*ca.* 200 fs) and the catalyst reduction from the reduced dye can be efficient and fast as well (*ca.* 10 ps),⁴⁴ the low performance of the photocathode is in fact caused by the serious charge recombination, which is a consequence of slow hole diffusion in NiO and a slow catalytic cycle from the catalyst. However, the liquid p-DSCs using NiO still can generate a mA-scale photocurrent with a redox couple, which is much higher than the photocurrent obtained from the photocathodes for solar fuel production. Therefore, the main obstacle of the low photocatalytic current from photocathodes is due to

the fast charge recombination between the reduced catalyst and the holes in NiO. Either prolongation of the charge separation between the reduced catalyst and holes in the NiO or the development of a more efficient catalyst could address the issue.

With the (p-SC)-dye-(n-SC) structure, we provide a possibility to build up an efficient photocathode for solar fuel production, as shown in Fig. 14. As the charges generated in dyes are initially separated by p-SC and n-SC and the separate holes and electrons have a long lifetime (>ns), this creates an opportunity to immobilize reduction catalysts directly on n-SC to receive electrons and perform catalytic reactions.²⁰ Since the catalyst cannot touch the p-SC, the charge-recombination process, as we mentioned above, can be therefore suppressed. We proved the concept by putting Pt nanoparticles on our NiO-dye-ZnO photocathode for light-driven proton reduction. A thin Al₂O₃ layer and a compact TiO₂ layer prepared by ALD were introduced to protect the dye and ZnO in acidic condition, respectively. We were able to obtain *ca.* 100 μA cm⁻² photocurrent for

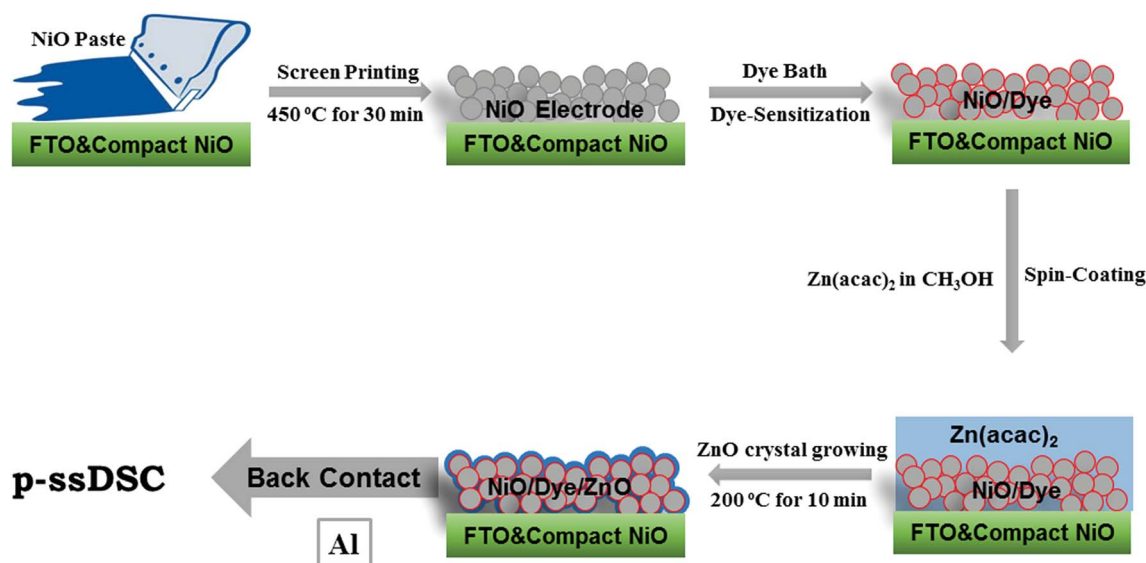


Fig. 11 Preparation process of p-ssDSCs based on NiO-dye-ZnO structure (reprinted and modified from ref. 20, with permission from Elsevier).



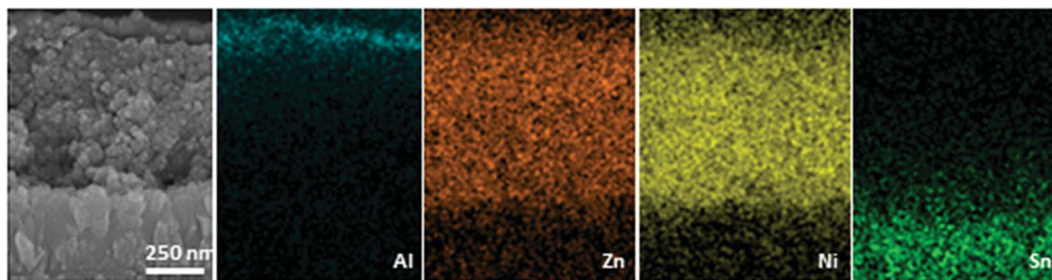


Fig. 12 Corresponding EDX element mapping analysis in the selected SEM cross-section range of the NiO–PB6–ZnO/Al-based solar cell device (reprinted and modified from ref. 20, with permission from Elsevier).

hydrogen evolution at a mild bias potential of 0.05 V vs. reversible hydrogen electrode (RHE). The photocurrent only dropped by 20% after 9 h photoelectrocatalysis, rendering a faradaic efficiency of hydrogen evolution up to 100%. Although this proof-of-concept work used a precious metal catalyst, it also provides a possibility to integrate suitable low-cost molecular catalysts in this system.

(b) Tandem solid-state dye-sensitized solar cells

Tandem dye-sensitized solar cells (t-DSCs) with liquid redox couples have been reported.^{2,3} The theoretical V_{OC} (V_{OC}^*) of the tandem solar cells is the sum of the photovoltage from both individual solar cells. Since the photoanode and photocathode in a liquid t-DSC share one redox couple, the V_{OC}^* is in principle determined by the energy difference between p-SC and n-SC (eqn (1)).

$$\begin{aligned} V_{OC}^* &= V_{OC,p} + V_{OC,n} = (E_{VB,p} - E_{redox}) + (E_{redox} - E_{CB,n}) \\ &= E_{VB,p-sc} - E_{CB,n-sc} \end{aligned} \quad (1)$$

where $V_{OC,p}$ is the open-circuit voltage of a p-type device; $V_{OC,n}$ is the open-circuit voltage of an n-type device; $E_{VB,p}$ is the VB level of p-SC; $E_{CB,n}$ is the CB level of n-SC, and E_{redox} is the redox potential of the redox couple.

Using NiO as the p-SC and TiO₂ as the n-SC in a liquid t-DSC, respectively, the V_{OC}^* is ca. 1 V. With the p-ssDSCs, it is possible to make t-ssDSCs as shown in Fig. 15a, in which each

photoelectrode has its own charge-transport material to collect charges and both photoelectrodes are connected by a transparent and conductive layer to perform inner charge transfer. With the transparent and conductive layer, this type of tandem device has a loss-free connection and in fact can be considered as a series connection of two individual solar cells obeying Kirchhoff's law. This kind of connection has been used in many types of solar cells to build up tandem devices.^{45–47}

$$\begin{aligned} V_{OC}^* &= V_{OC,p} + V_{OC,n} = (E_{VB,p} - E_{ETM}) + (E_{HTM} - E_{CB,n}) \\ &= E_{VB,p-sc} - E_{CB,n-sc} + (E_{HTM} - E_{ETM}) \end{aligned} \quad (2)$$

where E_{ETM} is the reduction potential of the electron-transport material for a p-type device and E_{HTM} is the oxidation potential of the hole-transport material for an n-type device. In this case, the V_{OC}^* of the t-ssDSC can be more than 2 V with suitable charge-transport materials (eqn (2) and Fig. 15b). Making a rough calculation based on dyes on both photoelectrodes having complementary light absorption up to 850 nm, an efficiency of 18% can be obtainable from a t-ssDSC with a V_{OC} of 2 V, a J_{SC} of 12 mA cm⁻², and an ff of 0.75.

Challenges

To make efficient t-ssDSCs, improvement of the performance of the p-ssDSCs is desirable, since the photocurrent and efficiency of the p-ssDSCs are still far behind that of the conventional n-type

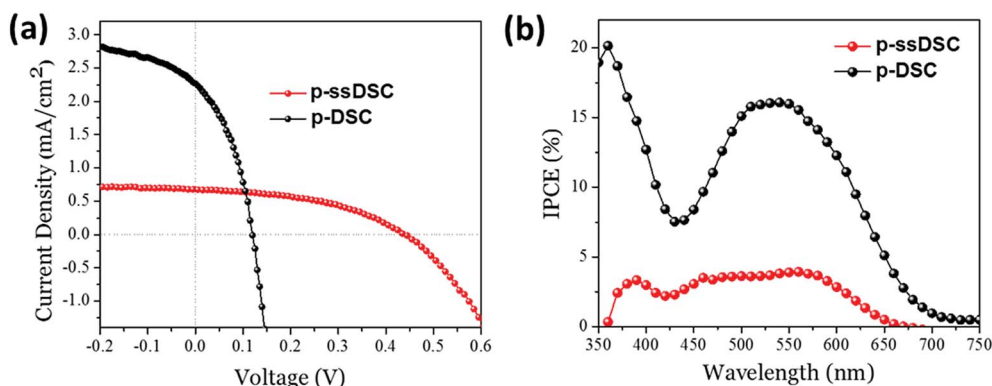


Fig. 13 J – V characteristics of the champion p-ssDSC and liquid p-DSC devices under 100 mW cm⁻², AM 1.5G simulated sunlight illumination (a) and corresponding IPCE spectra of the p-ssDSC and liquid p-DSC (b) (reprinted and modified from ref. 20, with permission from Elsevier).



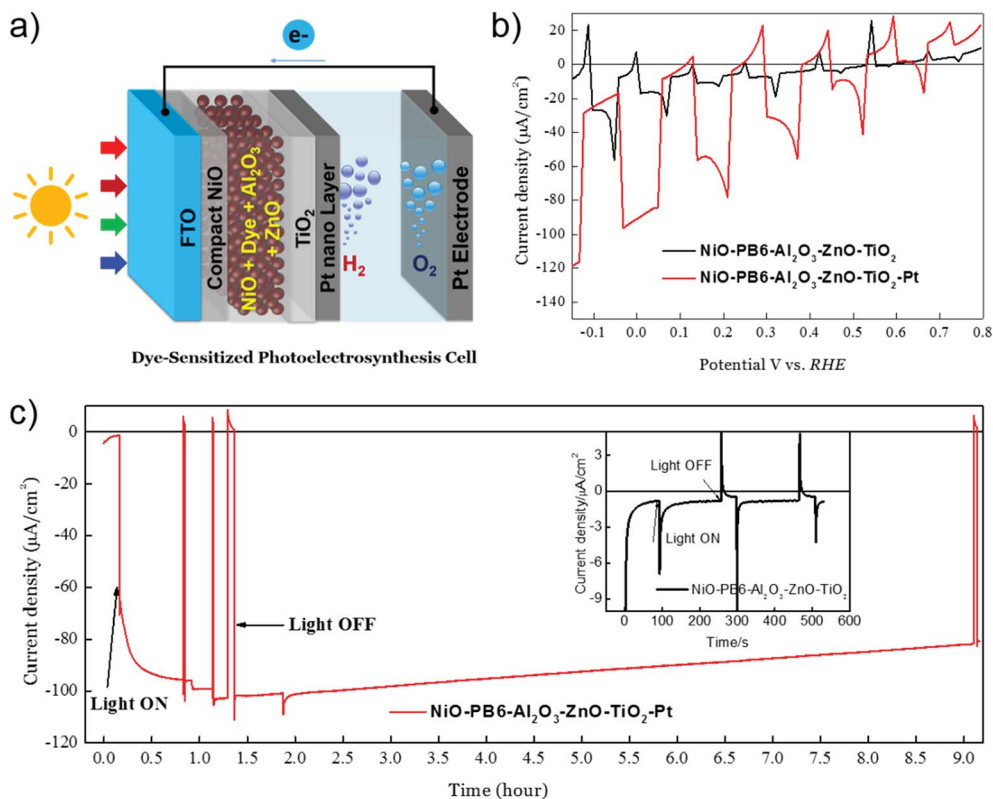


Fig. 14 (a) Device structure of a p-DSPEC; (b) linear sweep voltammetry (LSV) measurements of NiO–PB6–Al₂O₃–ZnO–TiO₂ with Pt (red curve) and without Pt catalyst (black curve); (c) chronoamperometry measurement in a 0.1 M acetate buffer (pH = 5.0) under bias of 0.05 V vs. RHE: NiO–PB6–Al₂O₃–ZnO–TiO₂–Pt photocathode (red curve); NiO–PB6–Al₂O₃–ZnO–TiO₂ photocathode (the inserted image, black curve), LED white light illumination (420–750 nm), 50 mW cm⁻², equivalent 1 sun condition from 420 nm to 750 nm (reprinted and modified from ref. 20, with permission from Elsevier).

ssDSCs (n-ssDSCs). One of major challenges is to seek new ETMs having high electron mobility and that can be prepared from solution-processing to giving good pore filling in the nanoporous p-SC film as well as to efficiently regenerate the dye. Non-fullerene organic ETMs can be considered in this scenario, since they have shown excellent performance as acceptors in organic solar cells and they can contribute photocurrent from light absorbing as well.⁴⁸ Moreover, the development of new p-SCs with fast hole-transport properties and advanced dyes with broad absorption

spectra and a high extinction coefficient can also help with improving the efficiency. Once the performance of p-ssDSC is significantly enhanced, the challenge left for t-ssDSCs is then how to connect the photocathode and the photoanode with a transparent and conductive layer without scarifying the individual photovoltaic parameters. For DSSFs based on this kind of photocathode using molecular catalysts, matching the energy levels between the n-type semiconductor and the catalyst is crucial.

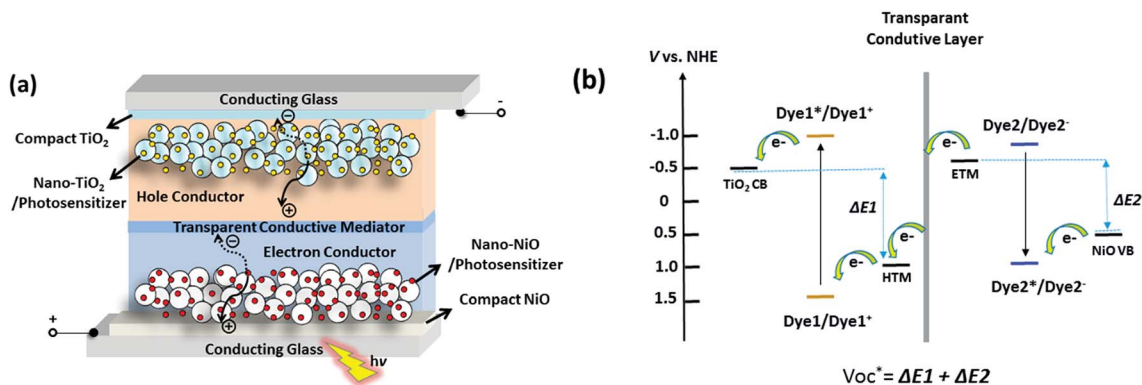


Fig. 15 Configuration (a) and potential diagram (b) of the proposed t-ssDSCs.



Conclusions

p-ssDSCs have been proposed and proven as functional devices for a few years. The photocurrent of the p-ssDSCs has been improved from an initial $50 \mu\text{A cm}^{-2}$ to ca. 1 mA cm^{-2} . This implies that there is a possibility to further improve the performance of p-ssDSCs by tuning different components as well as by optimizing their configuration. Together with n-ssDSCs, p-ssDSCs can be used to develop solid-state tandem dye-sensitized solar cells (t-ssDSCs), which can render much higher photovoltage than that from conventional liquid tandem dye-sensitized solar cells. This has the potential to dramatically improve the efficiency of the solid-state dye-sensitized solar cells. A dye-sensitized photocathode using ZnO as the ETM has shown potential to contribute an efficient and stable device for light-driven hydrogen production. The preliminary results suggest that this type of photocathode is also good to improve the performance of dye-sensitized solar fuel devices (DSSFs). Low-cost molecular catalysts are encouraged to be used in this kind of photocathode in follow up work. Still, there are many challenges still left behind in order to further improve the efficiency for both solar cells and solar fuel devices, as mentioned in this perspective.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The author acknowledges the foundation support from the Swedish Energy Agency (43237-1). The author also thanks Sina Wrede (UU) for valuable discussion and comments and Lei Tian (UU) for providing some figures.

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