

Cite this: *Chem. Sci.*, 2019, 10, 10018

All publication charges for this article have been paid for by the Royal Society of Chemistry

# CO<sub>2</sub>-induced single-crystal to single-crystal transformations of an interpenetrated flexible MOF explained by *in situ* crystallographic analysis and molecular modeling†

Arpan Hazra,<sup>1</sup> Dewald P. van Heerden,<sup>1</sup> Somananda Sanyal,<sup>1</sup> Prem Lama,<sup>1</sup> Catharine Esterhuysen<sup>1</sup> and Leonard J. Barbour<sup>1\*</sup>

A molecular-level investigation is reported on breathing behaviour of a metal–organic framework (**1**) in response to CO<sub>2</sub> gas pressure. High-pressure gas adsorption shows a pronounced step corresponding to a gate-opening phase transformation from a closed (**1**<sub>cp</sub>) to a large-pore (**1**<sub>lp</sub>) form. A plateau is observed upon desorption corresponding to narrow-pore intermediate form **1**<sub>np</sub> which does not occur during adsorption. These events are corroborated by pressure-gradient differential scanning calorimetry and *in situ* single-crystal X-ray diffraction analysis under controlled CO<sub>2</sub> gas pressure. Complete crystallographic characterisation facilitated a rationalisation of each phase transformation in the series **1**<sub>cp</sub> → **1**<sub>lp</sub> → **1**<sub>np</sub> → **1**<sub>cp</sub> during adsorption and subsequent desorption. Metropolis grand-canonical Monte Carlo simulations and DFT-PBE-D3 interaction energy calculations strongly underpin this first detailed structural investigation of an intermediate phase encountered upon desorption.

Received 13th August 2019  
Accepted 6th September 2019

DOI: 10.1039/c9sc04043a

rsc.li/chemical-science

## Introduction

Metal–organic frameworks (MOFs) have emerged as a distinct new class of porous materials of potential importance for various applications.<sup>1</sup> Flexible MOFs offer the possibility of stimulus-responsive dynamic behaviour,<sup>2</sup> and thus hold distinct advantages over conventional rigid porous materials such as activated carbons and zeolites. The flexibility of a MOF depends on several factors, such as the nature of the organic ligands, cleavage and regeneration of coordination bonds, and interaction between interpenetrated subnets.<sup>3</sup> The dynamic behaviour of such “soft porous crystals” (as described by Kitagawa)<sup>4</sup> is often due to transformations between multiple stable states of the flexible framework. Therefore, understanding the multiple structural changes in a dynamic MOF in response to external stimuli (*e.g.* guest removal or exchange, gas pressure, mechanical force, temperature)<sup>1b,3a,5</sup> is not only interesting but also highly desirable for the design of materials with potential for storage,<sup>6</sup> separation,<sup>7</sup> sensing,<sup>8</sup> and other related applications.<sup>9</sup> It should be noted that, among the plethora of reported MOFs, only a limited number exhibit dynamic behaviour such

as expansion/shrinkage of the framework<sup>10</sup> or opening/closing of pores,<sup>11</sup> giving rise to reversible changes in their physico-chemical properties.<sup>12</sup> This is because tailoring framework flexibility is much more challenging than controlling static structural features such as framework topology and pore size/shape.<sup>13</sup> Appreciating the obvious relationship between framework flexibility and host structure, we have successfully employed the strategy of using flexible organic linkers to imbue the framework with dynamic behaviour, both locally and overall.<sup>14</sup>

It should be noted that MOFs exhibiting dynamic behaviour pose a great challenge in terms of extracting atomic-level structural information on their dynamics.<sup>15</sup> Owing to the loss of crystal singularity that commonly occurs upon transformation between multiple stable states, single crystal X-ray diffraction (SCD) analysis is often not an option. The most credible structural information must therefore be derived from powder X-ray diffraction (PXRD) and other supporting analyses.<sup>16</sup>

Here we report structural analysis of a flexible MOF constructed from 1,2-bis-(4-pyridyl)ethane (bpa, with a flexible ethylene bridge) and 1,4-naphthalene dicarboxylic acid (ndcH<sub>2</sub>, with a rotatable aromatic ring). Solvothermal reaction of Zn(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O, ndcH<sub>2</sub> and bpa in a 1 : 1 : 1 molar ratio, following the procedure reported by Li *et al.*,<sup>17</sup> afforded the twofold interpenetrated 3D framework [Zn<sub>2</sub>(ndc)<sub>2</sub>(bpa)]<sub>n</sub>, hereafter referred to as **1**. Interestingly, the as-synthesised structure (**1**<sub>cp</sub>; cp = closed pore) does not possess any guest-accessible

Department of Chemistry and Polymer Science, University of Stellenbosch, Matieland, 7600, South Africa. E-mail: ljb@sun.ac.za; Fax: +27-21-808-3360

† Electronic supplementary information (ESI) available: Synthetic procedure, thermogravimetric analysis, additional figures, single-crystal and powder X-ray diffraction results, and computational details. CCDC: 1944820–1944823. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9sc04043a



space, as evidenced by the absence of solvent molecules. Nevertheless, high pressure gas sorption analysis reveals that **1<sub>cp</sub>** undergoes stimulus-driven breathing behaviour to an open phase under CO<sub>2</sub> gas pressure at 298 K. Moreover, the structural transformation occurs in single-crystal to single-crystal fashion. We note that Kanoo *et al.* recently reported the related 3D coordination polymer  $\{[\text{Zn}_2(\text{ndc})_2(\text{bpa})] \cdot \text{solvent}\}_n$  which, upon activation, yields a framework with the same topology as that of **1<sub>cp</sub>**.<sup>18</sup> However, the activated form of this framework is crystallographically different from **1<sub>cp</sub>** and the authors also reported that it does not exhibit any CO<sub>2</sub> adsorption at 293 K.

In an attempt to visualise the structural transformation of the host, as well as the mode of interaction between CO<sub>2</sub> guest molecules and the host framework, SCD analysis of the same crystal was carried out under controlled CO<sub>2</sub> gas pressure at 298 K using an environmental gas cell. The sorption and SCD results are also well supported by pressure-gradient differential scanning calorimetry (PGDSC) and variable-pressure powder X-ray diffraction (VP-PXRD) analysis. Theoretical calculations provide further insight into the dynamic behaviour.

## Results and discussion

**1<sub>cp</sub>** crystallises in the orthorhombic space group *Fmmm* and consists of  $\{\text{Zn}_2(\text{COO})_4\}$  paddlewheel secondary building units (SBU) with square pyramidal zinc centres coordinated to four oxygen atoms from four different ndc ligands at the basal positions, and one nitrogen atom from a bpa molecule at the apical position (Fig. 1a and Table S1†). The SBUs are linked in

the *ac* plane by ndc ligands to generate rhombic 2D nets (Fig. S1a†) that are pillared by bpa along the *b* axis to generate a 3D framework with  $\alpha$ -polonium net topology. Each net is large enough to accommodate another net, thus leading to the formation of a twofold interpenetrated framework (Fig. 1b). Each of the ndc naphthalene moieties in **1<sub>cp</sub>** is fourfold disordered (Fig. S2b†) while bpa displays twofold disorder (Fig. S2c†). As **1<sub>cp</sub>** is a dense structure, the interpenetrated networks are in close proximity and interact with each other by means of C–H $\cdots$ O ( $d_{\text{C}\cdots\text{O}} = 3.437(12)$  Å) and C–H $\cdots$  $\pi$  ( $d_{\text{H}\cdots\pi} = 3.321(9)$  Å) interactions (Fig. S2d†).

Although the as-synthesised framework does not possess any cavity or open channel, it is known that carboxylate-linked paddlewheel SBUs are conducive to dynamic behaviour.<sup>26,19</sup> Such phase transitions are generally accompanied by “kneecap” rotation about the O $\cdots$ O axis of the carboxylate groups, which facilitates reorientation of the SBUs while maintaining their interconnectivity.<sup>2a,20</sup> Furthermore, phenylene spacers can rotate in the presence of suitable external stimuli, thus allowing an increase in accessible space.

High-pressure sorption/desorption analyses with N<sub>2</sub>, CH<sub>4</sub>, and CO<sub>2</sub> were carried out at 298 K using a single sample of **1<sub>cp</sub>** (Fig. 1c). The isotherms reveal negligible uptake of N<sub>2</sub> and CH<sub>4</sub> at 50 bar. However, the CO<sub>2</sub> sorption profile shows gated uptake with a large hysteresis loop in the desorption curve. Modest uptake occurs until a threshold pressure of 32 bar is reached (point B in Fig. 1c), after which there is a sharp increase in uptake with a saturation plateau at 50 bar, corresponding to a capacity of 11.5 wt%, or two CO<sub>2</sub> molecules per host formula

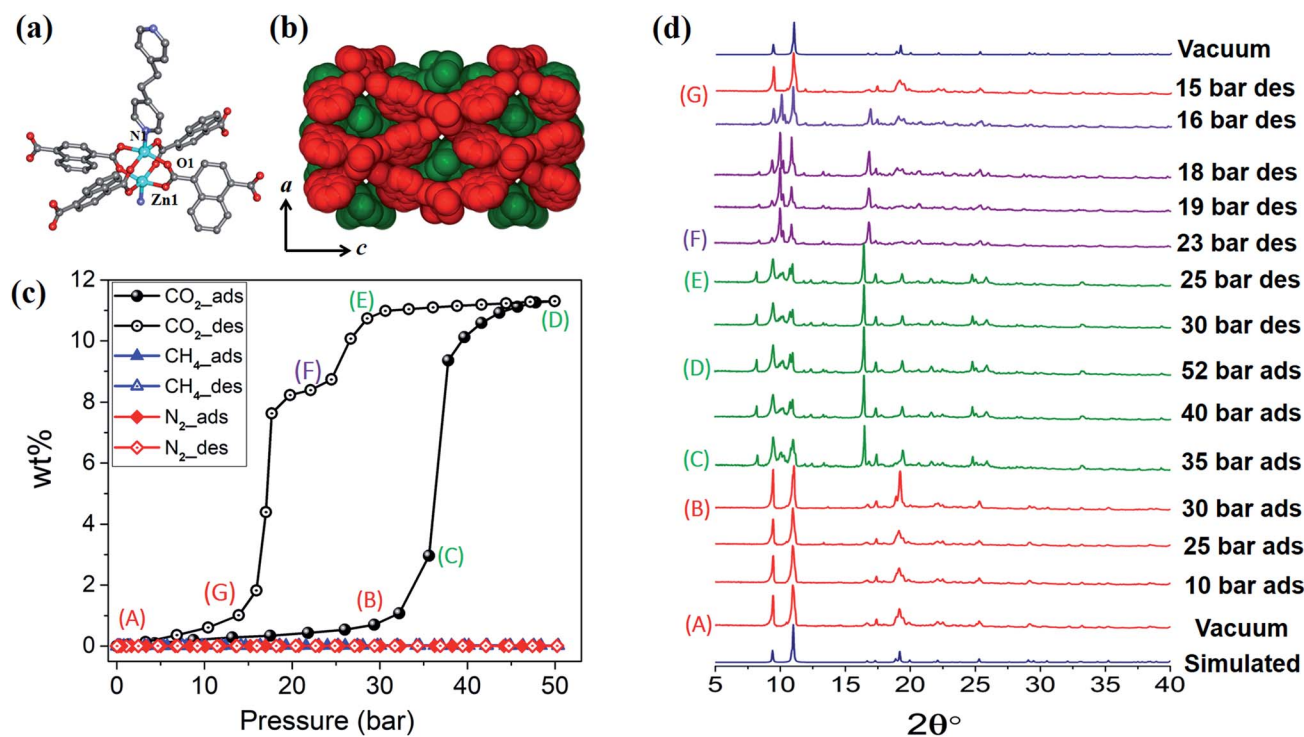


Fig. 1 (a) Coordination environment of Zn(II) in **1<sub>cp</sub>**. (b) Spacefilling view along [010] showing the twofold entanglement of **1<sub>cp</sub>** and its lack of guest-accessible space. (c) Adsorption (filled symbols) and desorption (open symbols) isotherms for **1** recorded at 298 K with test gases CO<sub>2</sub>, CH<sub>4</sub> and N<sub>2</sub>. (d) PXRD patterns of **1** under variable CO<sub>2</sub> gas pressure at 298 K.



unit (point D). The sudden uptake after 32 bar hints at a phase change from a nonporous to a porous phase. The desorption profile is significantly different from the adsorption profile. Pronounced hysteresis suggests strong interaction of the channel surface with trapped CO<sub>2</sub> molecules, which are retained down to 28 bar during desorption (point E). Moreover, the desorption isotherm features two steps with a plateau that implies a metastable phase with a capacity of 8.2 wt%, or 1.5 CO<sub>2</sub> molecules per formula unit (point F). This is evidence for a CO<sub>2</sub>-induced structural transformation of one open phase to a different open phase. All of the CO<sub>2</sub> is released upon further decreasing the pressure and the salient features of the isotherms are also present in a subsequent adsorption/desorption cycle (Fig. 2a).

PGDSC measurements were also carried out for **1<sub>cp</sub>** exposed to CO<sub>2</sub> in the pressure range 1–50 bar at 298 K (Fig. 2b).<sup>14c,21</sup> During adsorption an exothermic peak occurs with an onset pressure of 35 bar, corresponding to the sudden uptake of CO<sub>2</sub> observed in the adsorption isotherm. The integrated heat of  $-23.69 \text{ kJ mol}^{-1}$  for this event includes the heat of adsorption along with a small contribution from the associated structural change. During desorption, endothermic events occur with onset pressures of 23 and 15 bar, with associated energies of 13.07 and 37.66 kJ mol<sup>-1</sup>, respectively, spanning the plateau observed in the desorption isotherm.

The “breathing” effect was further investigated using VP-PXRD. As shown in Fig. 1d, no change is observed in the diffractograms until 35 bar of CO<sub>2</sub> pressure, after which a distinct

change is observed that corresponds to point C in the sorption isotherm (Fig. 1c). The diffractograms then remain unchanged until 52 bar is reached and, upon decreasing the pressure, a change is observed at 23 bar, which corresponds to point F on the desorption profile. This PXRD pattern does not resemble that of either the fully loaded or the as-synthesised structures, thus indicating a new phase. This phase persists until 15 bar (point G), after which the diffractogram resembles that of the as-synthesised **1<sub>cp</sub>** (down to vacuum).

In order to obtain detailed structural information regarding the CO<sub>2</sub> sorption sites and loading-induced changes to the framework, *in situ* VP-SCD studies were carried out on **1<sub>cp</sub>** under controlled CO<sub>2</sub> gas pressures using an environmental gas cell.<sup>22</sup> Although the onset of the first structural change occurs at 35 bar (as evident from VP-PXRD analysis), the saturation pressure of 50 bar was used. The crystal retains its singularity at this pressure and structural characterisation reveals a single-crystal to single-crystal phase transition from **1<sub>cp</sub>** to a large pore phase **1<sub>lp</sub>** with the formula  $\{[\text{Zn}_2(\text{ndc})_2(\text{bpa})] \cdot 2\text{CO}_2\}_n$  (Table S1†). Significant relative movement of the two interpenetrated nets is accompanied by a decrease in space group symmetry from orthorhombic *Fm $\bar{3}m$*  to monoclinic *C2/c*. The phase transition **1<sub>cp</sub>** → **1<sub>lp</sub>** resembles the unfolding of an articulated wine-rack.<sup>3b</sup> Although the SBU and the framework topology remain unchanged, the acute angle ( $\theta_1$ ) between ndc ligands increases from 59.747(8)° to 81.621(7)° (Fig. S6 and Table S2†). The movement of the interpenetrated nets relative to each other with concomitant changes in the conformation of the linkers results in the formation of guest-accessible space in **1<sub>lp</sub>**. Two crystallographically unique CO<sub>2</sub> molecules (labelled A and B, Fig. 3) are located in the 3D guest-accessible space of **1<sub>lp</sub>** and like types repeat along [001]. The  $d_{C\dots C}$  separations between the pseudo-slipped-parallel arranged CO<sub>2</sub> molecules are in the range 3.928(1) to 4.927(2) Å (Fig. 3c). The prefix “pseudo” distinguishes the arrangements in **1<sub>lp</sub>** from the idealised arrangements in dry ice.<sup>23</sup> The CO<sub>2</sub> molecules form strong electrostatic interactions with the framework. The nearest contacts between carbon atoms of CO<sub>2</sub> molecules and oxygen atoms of the carboxylate groups are 3.465(9) and 3.63(1) Å for types A and B, respectively (Fig. S8†). Close contacts also occur between CO<sub>2</sub> molecules and the hydrogen atoms of the naphthalene groups, as well as those of the ethylene bridges (Table S3†).

To investigate the phase changes occurring upon desorption, the pressure was reduced from 50 to 23 bar. *In situ* VP-SCD analysis was again carried out using the same crystal. Very few reports exist that describe desorption VP-PXRD analysis of flexible MOFs.<sup>24,16</sup> However, to the best of our knowledge, this is the first detailed SCD investigation on an intermediate phase obtained upon desorption. At 23 bar the space group changes to *P2<sub>1</sub>/m* and further adjustment of the two nets (Fig. S6 and Table S2†) yields the narrow pore (**1<sub>np</sub>**) form  $\{[\text{Zn}_2(\text{ndc})_2(\text{bpa})] \cdot 1.5\text{CO}_2\}_n$ .<sup>28</sup> The 3D accessible space in **1<sub>lp</sub>** transforms into discrete voids (Fig. 4). Although it was only possible to model one CO<sub>2</sub> molecule, the difference map reveals the presence of residual electron density consistent with the presence of unmodelled diffuse CO<sub>2</sub>. Metropolis<sup>24</sup> grand-canonical Monte

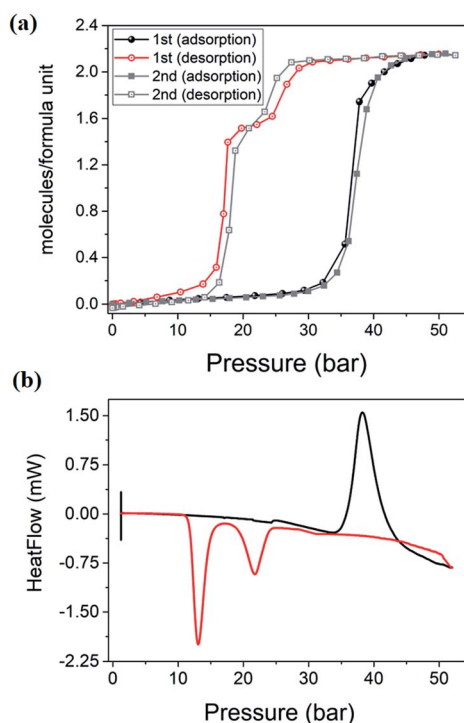


Fig. 2 (a) Two consecutive cycles of CO<sub>2</sub> adsorption/desorption for **1** at 298 K. (b) PGDSC trace showing thermal events corresponding to inflections observed during adsorption (black) and desorption (red) in the isotherm.



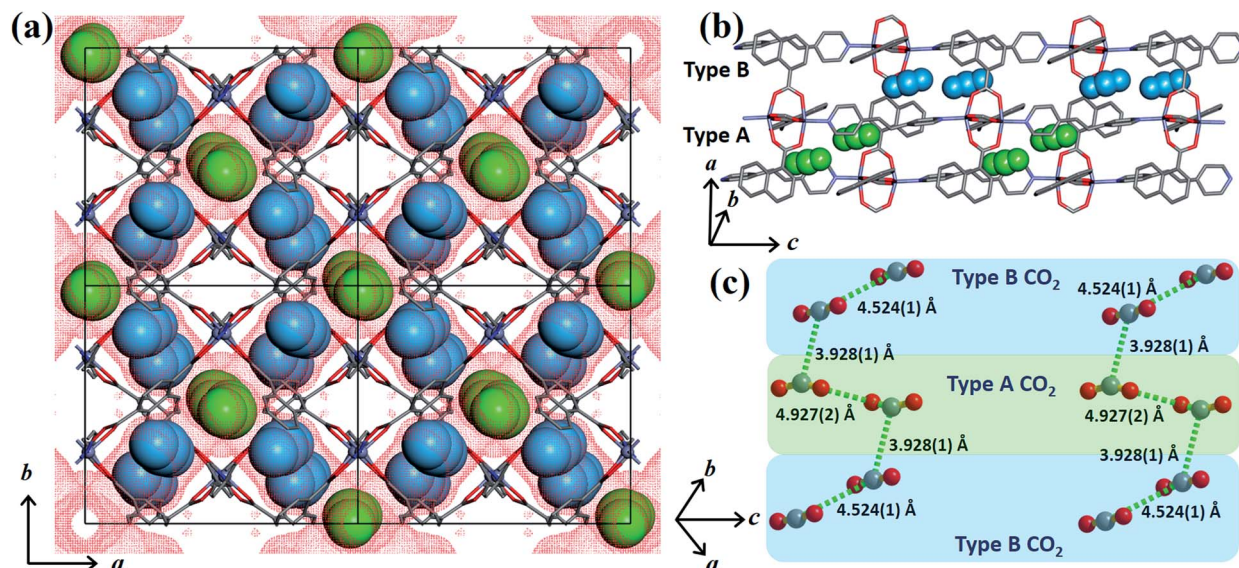


Fig. 3 (a) View of  $1_{Ip}$  along [001]. The guest-accessible space<sup>27</sup> is shown as red dots while type A and B CO<sub>2</sub> molecules (in spacefilling representation) are coloured green and blue, respectively. (b) and (c) Perspective views of the pseudo-slipped-parallel arrangement of the CO<sub>2</sub> molecules with  $d_{C...C}$  separations indicated. Hydrogen atoms are omitted for clarity.

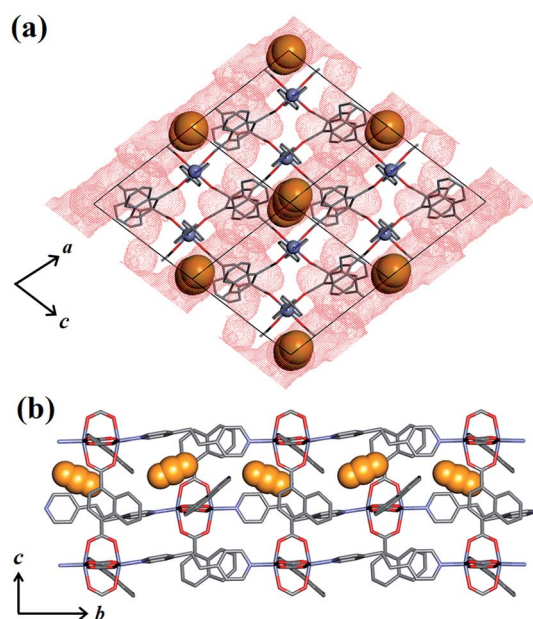


Fig. 4 (a) View of  $1_{nIp}$  along [010] with the guest-accessible void surfaces<sup>27</sup> shown as red dots and CO<sub>2</sub> molecules in spacefilling representation. (b) View of  $1_{nIp}$  along [100] showing propagation of CO<sub>2</sub> molecules in the crystallographic  $b$  direction. Hydrogen atoms are omitted for clarity.

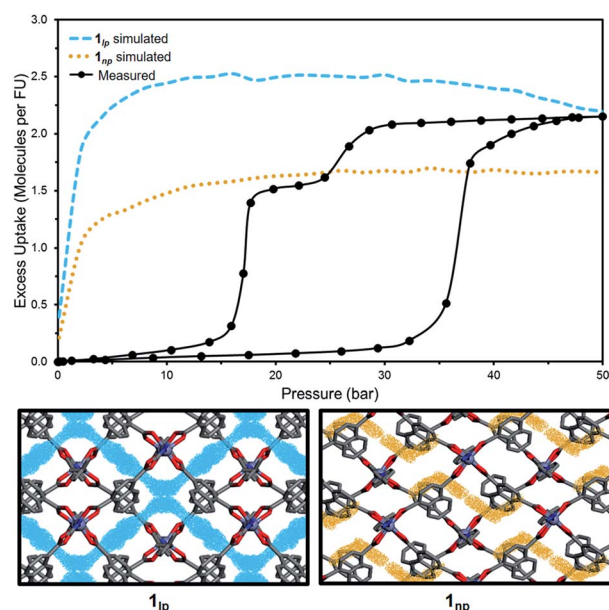


Fig. 5 Comparison of experimental and simulated excess CO<sub>2</sub> uptake at 298 K for  $1_{Ip}$  (blue) and  $1_{nIp}$  (orange) with corresponding simulated probability density maps at 50 and 23 bar, respectively.

Carlo simulations corroborate the measured sorption capacity of 1.5 CO<sub>2</sub> molecules per formula unit of  $1_{nIp}$  (Fig. 5).

The breathing mechanism is derived from two principal deformations of the framework (Fig. 6a). The dihedral angle  $\theta_2$  between the naphthalene ring and the plane passing through the four Zn atoms (as shown in Fig. 6b–d) changes owing to linker rotation. The O...O vector of the carboxylate group acts as

a “kneecap” around which the SBUs can rotate during the phase transformation. This hinge-like motion changes the dihedral angle  $\theta_3$  between the plane passing through all four oxygen atoms of ndc and that between two oxygen and two zinc atoms of the SBU (blue and pink plane in Fig. 6e–g, respectively).

The host–guest and guest–guest interactions involving the two types of CO<sub>2</sub> molecules in  $1_{Ip}$  were investigated theoretically in order to rationalise the occurrence of the intermediate phase  $1_{nIp}$  during desorption. Periodic unit cell and geometry



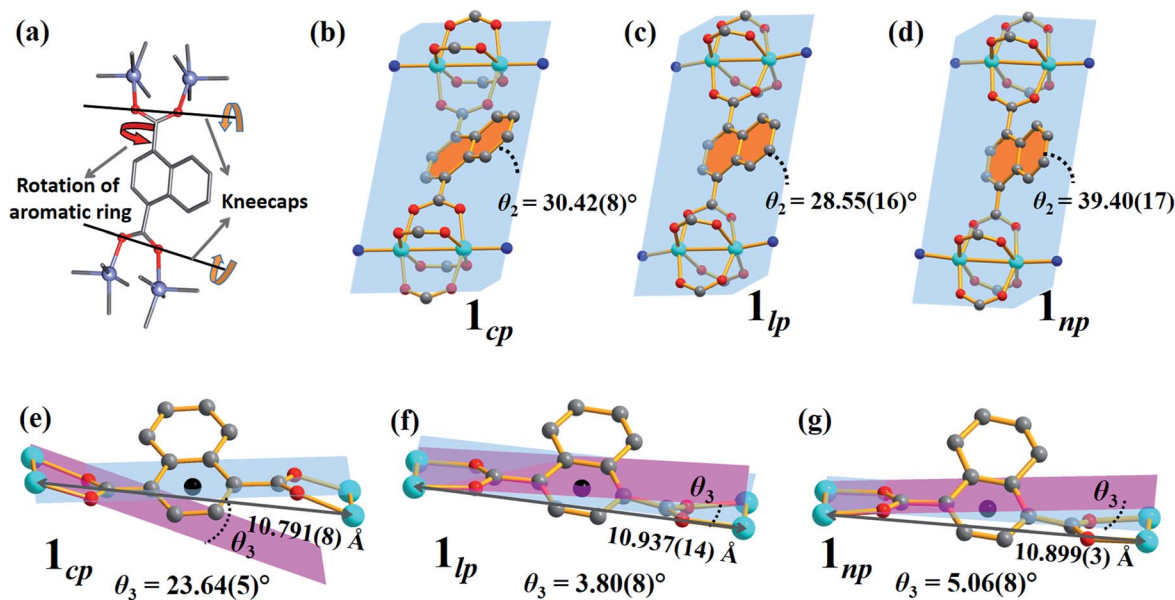


Fig. 6 (a) Schematic representations of the two factors governing the breathing behaviour in **1**: (b–d) rotation of the naphthalene rings (orange) relative to the plane between the four Zn ions (blue) and (e–g) kneecap hinge motion between planes passing through the ndc oxygen atoms (blue) and the  $\text{Zn}_2\text{O}_2$  moiety (pink) of the SBU.

optimisations of a  $1 \times 1 \times 1$  representation for  $1_{lp}$  and  $1_{np}$  showed negligible changes in cell parameters and atomic positions, confirming the suitability of the DFT-PBE-D3/double-zeta level of theory (Table S5, Fig. S12 and S13<sup>†</sup>).<sup>25</sup> Counterpoise-corrected<sup>26</sup> host-guest interaction energies of 40.09 and 34.35  $\text{kJ mol}^{-1}$  were calculated for the type A and B  $\text{CO}_2$  molecules of  $1_{lp}$ , respectively (Table S6<sup>†</sup>). Favourable binary guest-guest interaction energies between two type A, two type B and between type A and type B  $\text{CO}_2$  molecules amount to 1.06, 2.35 and 3.36  $\text{kJ mol}^{-1}$ , respectively (Fig. 7a). Computed electron density difference maps (EDDMs) for  $1_{lp}$  reveal interactions between the  $\pi$  systems of type A  $\text{CO}_2$  molecules and aromatic C–H bonds of ndc, while type B  $\text{CO}_2$  molecules interact strongly with carboxylate oxygen atoms (Fig. 7b; see ESI for more details and Fig. S14<sup>†</sup> for geometry-optimised  $d_{C\dots C}$  separations).

Collectively these and other factors such as channel shape and surface potential differences explain why  $\text{CO}_2$  molecules are retained until 28 bar during desorption, thus accounting for the hysteresis observed in Fig. 1c. The host-guest interaction energy of 40.67  $\text{kJ mol}^{-1}$  calculated for  $1_{np}$  (Table S6<sup>†</sup>) is consistent with the observation that only a fraction of the  $\text{CO}_2$  molecules are removed during the transition from  $1_{lp}$  to  $1_{np}$ . The latter persists until 15 bar, after which the VP-PXRD diffractogram again resembles that of  $1_{cp}$  (Fig. 1d).

## Conclusion

We report three  $\text{CO}_2$  pressure-induced forms of a flexible, twofold interpenetrated pillared-layered MOF that reversibly switches between closed-, narrow- and large-pore forms. All of these forms have been characterised structurally using *in situ* SCD analysis. The transformation is brought about by  $\text{CO}_2$  but not by  $\text{N}_2$  or  $\text{CH}_4$  in the pressure range 0–50 bar at 298 K. Inflections observed in the sorption isotherm are consistent with PGDSC and VP-PXRD analyses. This study describes the first SCD analysis of an intermediate phase that is only accessible during desorption. Using complementary theoretical evaluations, we provide a molecular-level explanation for the plateau observed in the desorption isotherm. Not only do the results presented here provide considerable new insight into the breathing behaviour of a flexible MOF, but they also offer opportunities to develop a new class of materials where the structure–property relationships can be exploited in molecular recognition.

## Conflicts of interest

There are no conflicts to declare.

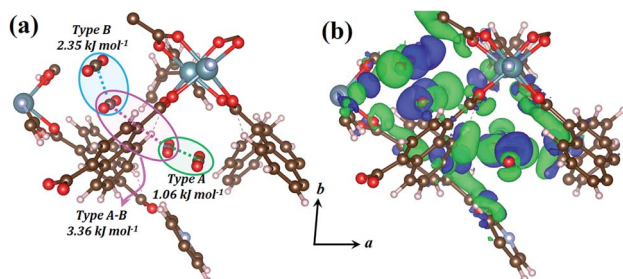


Fig. 7 (a) Perspective view showing  $\text{CO}_2$  positions within the geometry-optimised structure of  $1_{lp}$  with guest-guest interaction energies between different combinations of two  $\text{CO}_2$  molecules. (b) EDDM showing areas of enhanced (green) and depleted (blue) electron density of the geometry-optimised structure of  $1_{lp}$  with respect to its isolated MOF and  $\text{CO}_2$  components. The EDDM was generated for a periodic  $1 \times 1 \times 1$  unit cell and only a fragment is shown for clarity.



## Acknowledgements

We thank the National Research Foundation of South Africa and Stellenbosch University (Subcommittee B) for financial support and the Centre for High Performance Computing (CHPC) in Cape Town for the use of their computational resources.

## Notes and references

- (a) H. Furukawa, K. E. Cordova, M. O'Keeffe and O. M. Yaghi, *Science*, 2013, **341**, 1230444; (b) S. Kitagawa, R. Kitaura and S.-i. Noro, *Angew. Chem., Int. Ed.*, 2004, **43**, 2334.
- (a) A.-X. Zhu, Q.-Y. Yang, A. Kumar, C. Crowley, S. Mukherjee, K.-J. Chen, S.-Q. Wang, D. O'Nolan, M. Shivanna and M. J. Zaworotko, *J. Am. Chem. Soc.*, 2018, **140**, 15572; (b) J. Rabone, Y. F. Yue, S. Y. Chong, K. C. Stylianou, J. Bacsá, D. Bradshaw, G. R. Darling, N. G. Berry, Y. Z. Khimiyak, A. Y. Ganin, P. Wiper, J. B. Claridge and M. J. Rosseinsky, *Science*, 2010, **329**, 1053; (c) A. Schneemann, V. Bon, I. Schwedler, I. Senkovska, S. Kaskel and R. A. Fischer, *Chem. Soc. Rev.*, 2014, **43**, 6062; (d) R. Haldar, N. Sikdar and T. K. Maji, *Mater. Today*, 2015, **18**, 97; (e) T. K. Maji, R. Matsuda and S. Kitagawa, *Nat. Mater.*, 2007, **6**, 142; (f) X.-N. Wang, P. Zhang, A. Kirchon, J.-L. Li, W.-M. Chen, Y.-M. Zhao, B. Li and H.-C. Zhou, *J. Am. Chem. Soc.*, 2019, **141**, 13654.
- (a) G. Férey and C. Serre, *Chem. Soc. Rev.*, 2009, **38**, 1380; (b) E. J. Carrington, C. A. McAnally, A. J. Fletcher, S. P. Thompson, M. Warren and L. Brammer, *Nat. Chem.*, 2017, **9**, 882; (c) R. E. Morris and L. Brammer, *Chem. Soc. Rev.*, 2017, **46**, 5444; (d) H. S. Scott, N. Ogiwara, K.-J. Chen, D. G. Madden, T. Pham, K. Forrest, B. Space, S. Horike, J. J. Perry IV, S. Kitagawa and M. J. Zaworotko, *Chem. Sci.*, 2016, **7**, 5470.
- S. Kitagawa, *Angew. Chem., Int. Ed.*, 2015, **54**, 10686.
- (a) K. Uemura, R. Matsuda and S. Kitagawa, *J. Solid State Chem.*, 2005, **178**, 2420; (b) S. Henke, M. T. Wharmby, G. Kieslich, I. Hante, A. Schneemann, Y. Wu, D. Daisenberger and A. K. Cheetham, *Chem. Sci.*, 2018, **9**, 1654; (c) G. A. Craig, P. Larpent, S. Kusaka, R. Matsuda, S. Kitagawa and S. Furukawa, *Chem. Sci.*, 2018, **9**, 6463; (d) A. Hazra, S. Bonakala, K. K. Bejagam, S. Balasubramanian and T. K. Maji, *Chem.–Eur. J.*, 2016, **22**, 7792; (e) S. Horike, S. Shimomura and S. Kitagawa, *Nat. Chem.*, 2009, **1**, 695; (f) W. Cai, A. Gładysiak, M. Anioła, V. J. Smith, L. J. Barbour and A. Katrusiak, *J. Am. Chem. Soc.*, 2015, **137**, 9296; (g) C. Serre, F. Millange, C. Thouvenot, M. Noguès, G. Marsolier, D. Louër and G. Férey, *J. Am. Chem. Soc.*, 2002, **124**, 13519; (h) S. Henke, W. Li and A. K. Cheetham, *Chem. Sci.*, 2014, **5**, 2392.
- J. A. Mason, J. Oktawiec, M. K. Taylor, M. R. Hudson, J. Rodriguez, J. E. Bachman, M. I. Gonzalez, A. Cervellino, A. Guagliardi, C. M. Brown, P. L. Llewellyn, N. Masciocchi and J. R. Long, *Nature*, 2015, **527**, 357.
- (a) R. K. Motkuri, P. K. Thallapally, H. V. R. Annapureddy, L. X. Dang, R. Krishna, S. K. Nune, C. A. Fernandez, J. Liu and B. P. McGrail, *Chem. Commun.*, 2015, **51**, 8421; (b) C.-T. He, Z.-M. Ye, Y.-T. Xu, D.-D. Zhou, H.-L. Zhou, D. Chen, J.-P. Zhang and X.-M. Chen, *Chem. Sci.*, 2017, **8**, 7560; (c) F. Cacho-Bailo, I. Matito-Martos, J. Perez-Carbajo, M. Etxeberria-Benavides, O. Karvan, V. Sebastián, S. Calero, C. Téllez and J. Coronas, *Chem. Sci.*, 2017, **8**, 325.
- M. D. Allendorf, R. J. T. Houk, L. Andruszkiewicz, A. A. Talin, J. Pikarsky, A. Choudhury, K. A. Gall and P. J. Hesketh, *J. Am. Chem. Soc.*, 2008, **130**, 14404.
- R. K. Motkuri, P. K. Thallapally, S. K. Nune, C. A. Fernandez, B. P. McGrail and J. L. Atwood, *Chem. Commun.*, 2011, **47**, 7077.
- (a) Y.-X. Tan, F. Wang, Y. Kang and J. Zhang, *Chem. Commun.*, 2011, **47**, 770; (b) S. Krause, V. Bon, I. Senkovska, U. Stoeck, D. Wallacher, D. M. Többsen, S. Zander, R. S. Pillai, G. Maurin, F.-X. Coudert and S. Kaskel, *Nature*, 2016, **532**, 348.
- (a) F. Salles, G. Maurin, C. Serre, P. L. Llewellyn, C. Knöfel, H. J. Choi, Y. Filinchuk, L. Oliviero, A. Vimont, J. R. Long and G. Férey, *J. Am. Chem. Soc.*, 2010, **132**, 13782; (b) P. Deria, D. A. Gómez-Gualdrón, W. Bury, H. T. Schaefer, T. C. Wang, P. K. Thallapally, A. A. Sarjeant, R. Q. Snurr, J. T. Hupp and O. K. Farha, *J. Am. Chem. Soc.*, 2015, **137**, 13183.
- (a) X.-L. Qi, R.-B. Lin, Q. Chen, J.-B. Lin, J.-P. Zhang and X.-M. Chen, *Chem. Sci.*, 2011, **2**, 2214; (b) T. K. Maji, G. Mostafa, R. Matsuda and S. Kitagawa, *J. Am. Chem. Soc.*, 2005, **127**, 17152; (c) M. Souto, J. Romero, J. Calbo, I. J. Vitorica-Yrezabal, J. L. Zafra, J. Casado, E. Ortí, A. Walsh and G. Mínguez Espallargas, *J. Am. Chem. Soc.*, 2018, **140**, 10562; (d) J. Su, S. Yuan, H.-Y. Wang, L. Huang, J.-Y. Ge, E. Joseph, J. Qin, T. Cagin, J.-L. Zuo and H.-C. Zhou, *Nat. Commun.*, 2017, **8**, 2008.
- (a) Z.-J. Lin, J. Lü, M. Hong and R. Cao, *Chem. Soc. Rev.*, 2014, **43**, 5867; (b) Y. Zhang, X. Zhang, J. Lyu, K.-i. Otake, X. Wang, L. R. Redfern, C. D. Malliakas, Z. Li, T. Islamoglu, B. Wang and O. K. Farha, *J. Am. Chem. Soc.*, 2018, **140**, 11179.
- (a) E. R. Engel, A. Jouaiti, C. X. Bezuidenhout, M. W. Hosseini and L. J. Barbour, *Angew. Chem., Int. Ed.*, 2017, **56**, 8874; (b) P. Lama, H. Aggarwal, C. X. Bezuidenhout and L. J. Barbour, *Angew. Chem., Int. Ed.*, 2016, **55**, 13271; (c) P. M. Bhatt, E. Batisai, V. J. Smith and L. J. Barbour, *Chem. Commun.*, 2016, **52**, 11374; (d) H. Aggarwal, R. K. Das, P. M. Bhatt and L. J. Barbour, *Chem. Sci.*, 2015, **6**, 4986.
- Y.-X. Shi, W.-X. Li, W.-H. Zhang and J.-P. Lang, *Inorg. Chem.*, 2018, **57**, 8627.
- S. Jeoung, S. Lee, J. H. Lee, S. Lee, W. Choe, D. Moon and H. R. Moon, *Chem. Commun.*, 2019, **55**, 8832.
- N.-Y. Li, Y. Ge, T. Wang, S.-J. Wang, X.-Y. Ji and D. Liu, *CrystEngComm*, 2014, **16**, 2168.
- P. Kanoo, R. Haldar, S. K. Reddy, A. Hazra, S. Bonakala, R. Matsuda, S. Kitagawa, S. Balasubramanian and T. K. Maji, *Chem.–Eur. J.*, 2016, **22**, 15864.
- (a) S. K. Elsaidi, M. H. Mohamed, D. Banerjee and P. K. Thallapally, *Coord. Chem. Rev.*, 2018, **358**, 125; (b) M. Shivanna, Q.-Y. Yang, A. Bajpai, E. Patyk-Kazmierczak and M. J. Zaworotko, *Nat. Commun.*, 2018, **9**, 3080; (c)



- J. H. Lee, S. Jeoung, Y. G. Chung and H. R. Moon, *Coord. Chem. Rev.*, 2019, **389**, 161.
- 20 C. Serre, C. Mellot-Draznieks, S. Surblé, N. Audebrand, Y. Filinchuk and G. Férey, *Science*, 2007, **315**, 1828.
- 21 P. Lama and L. J. Barbour, *J. Am. Chem. Soc.*, 2018, **140**, 2145.
- 22 T. Jacobs, G. O. Lloyd, J.-A. Gertenbach, K. K. Müller-Nedebock, C. Esterhuysen and L. J. Barbour, *Angew. Chem., Int. Ed.*, 2012, **51**, 4913.
- 23 C. X. Bezuidenhout, V. J. Smith, P. M. Bhatt, C. Esterhuysen and L. J. Barbour, *Angew. Chem., Int. Ed.*, 2015, **54**, 2079.
- 24 N. Metropolis, A. W. Rosenbluth, M. N. Rosenbluth, A. H. Teller and E. Teller, *J. Chem. Phys.*, 1953, **21**, 1087.
- 25 (a) J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865; (b) S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- 26 (a) B. Liu and A. D. McLean, *J. Chem. Phys.*, 1973, **59**, 4557; (b) H. B. Jansen and P. Ros, *Chem. Phys. Lett.*, 1969, **3**, 140.
- 27 We note that descriptions of guest-accessible space are susceptible to a number of subjective factors such as probe radii, grid spacings and models of disorder. While it seems reasonable to use a probe radius equal to the van der Waals radii of guest molecule terminal atoms, the volume calculations utilise equilibrium atomic coordinates and do not take structural dynamics into account. To determine guest-accessible space in the present study we have used a probe radius of 1.6 Å, which is slightly larger than the van der Waals radius of an oxygen atom.
- 28 The host:guest stoichiometry was determined from the sorption data.

