

Cite this: *Chem. Sci.*, 2019, **10**, 10765 All publication charges for this article have been paid for by the Royal Society of ChemistryReceived 6th August 2019
Accepted 1st October 2019

DOI: 10.1039/c9sc03917a

rsc.li/chemical-science

Near-infrared light-responsive hydrogels *via* peroxide-decorated MXene-initiated polymerization[†]

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Two-dimensional MXene $Ti_3C_2T_x$ nanosheets with peroxide decoration ($p-Ti_3C_2T_x$) are synthesized by a sonication-assisted MILD etching method. The obtained MXenes can generate hydroxyl radical species and act as an initiator for free-radical polymerization of a series of acrylic monomers without the use of light illumination or co-initiators. The monomers analyzed include acrylamide, *N*-isopropylacrylamide (NIPAM), *N,N*-dimethylacrylamide, methyl methacrylate, and hydroxyethyl methacrylate. By simply mixing *N*-isopropylacrylamide monomers and $p-Ti_3C_2T_x$ nanosheets under deoxygenated conditions, PNIPAM-based nanocomposite hydrogels are synthesized using a high concentration of the monomer. The nanocomposite hydrogels have a photothermal conversion efficiency of 34.7% and photothermal stability superior to that of pristine $Ti_3C_2T_x$. Taking advantage of the thermal responsive behavior of PNIPAM, the nanocomposite hydrogels are successfully exploited as remotely near-infrared light controlled "smart" windows, fluidic valves and photodetectors.

Introduction

"Smart" responsive hydrogels are now being widely studied in various fields such as soft robotics,^{1,2} sensors,^{3,4} microdevices,⁵ drug delivery,^{6,7} cell/tissue engineering scaffolds,^{8,9} and reaction vessels.^{10,11} Exposure to different stimuli (for example light, temperature, pH, and electric field) may cause these hydrogels to swell/shrink, bend/unbend, or swim/crawl, mimicking natural creatures with responsive behavior.^{12–15} In particular, light is a distinct stimulus to tune the properties of a hydrogel spatiotemporally in a remote controlled manner with great ease and convenience.^{6,16–19} During the past decade, increasing attention has been paid to near-infrared (NIR) light-absorbing materials. As opposed to UV-Vis light, NIR light has advantages of deep penetration, low energy, and little damage to structures. Additionally, some NIR-absorbing materials have strong photothermal conversion ability resulting in a rapid temperature increase at the irradiation site.^{20–22} Moreover, NIR can trigger phase separation or phase transition of thermo-

responsive polymer networks, causing various physical and morphological changes. Hence, NIR-responsive hydrogels are being exploited for different applications, including smart windows, microfluidic valves, 3D microfabrication of complex objects, and drug delivery.^{18,20,23–26}

MXenes have been an emerging class of graphene-like two-dimensional nanomaterials since their first discovery in 2011.²⁷ Owing to their properties of a functionalized surface, small band gaps, and excellent photothermal properties and conductivity, they have diverse applications in batteries, supercapacitors, photocatalysis, electrocatalysis, and biomedicine.^{28–33} $Ti_3C_2T_x$, a representative MXenes, has abundant functional groups (*i.e.* $-OH$, $-O$ or $-F$) and is well dispersible in aqueous solution. $Ti_3C_2T_x$ has outstanding photothermal properties, and $Ti_3C_2T_x$ -based hydrogels have been applied in the remote light-controlled microfluidic pipeline.^{34,35} $Ti_3C_2T_x$ -incorporated hydrogels were also recently applied as strain sensors and exhibited outstanding tensile strain sensitivity, with a gauge factor one order of magnitude higher than that of pristine hydrogels.³⁶ Notably, $Ti_3C_2T_x$ participates in many chemical processes (such as CO oxidation, oxygen evolution, and oxygen reduction), by acting as catalysts, cocatalysts or supporting substrates. $Ti_3C_2T_x$ itself can catalyze *in situ* polymerization of aromatic EDOT or pyrrole using the mechanism of charge-transfer-induced polymerization for the polymerization process.³⁷ To the best of our knowledge, the catalytic role of MXenes in radical polymerization has not been reported yet.^{37–39}

Herein, we aim to design MXene-based NIR-responsive nanocomposite (NC) hydrogels towards remote-controlled

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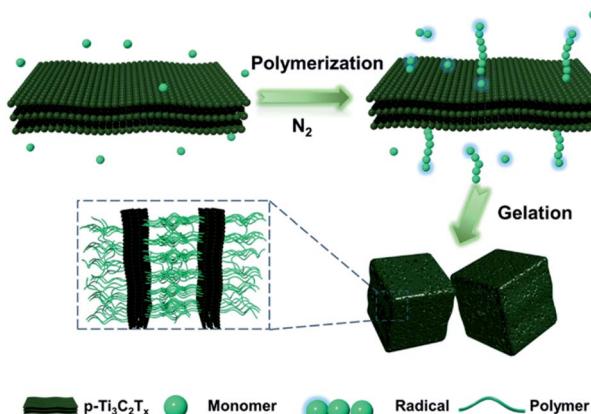
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† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9sc03917a

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Scheme 1 Schematic illustration of p-Ti₃C₂T_x-initiated polymerization and subsequent gelation.

devices. To be specific, we synthesized a novel kind of Ti₃C₂T_x nanosheet with peroxide decoration (p-Ti₃C₂T_x), which itself acts as an initiator in the polymerization of acrylic monomers without any light irradiation or co-initiators. The preparation procedure for the MXene-based hydrogels is quite simple and efficient (Scheme 1). Briefly, after adding monomers into a freshly prepared p-Ti₃C₂T_x nanosheet suspension in an ice bath, polymerization occurs within several minutes after the N₂ purging process. Gelation is observed at a high concentration of the monomer. In view of the widespread use of inorganic semiconductor nanoparticles (such as V₂C, perovskite nanocrystals, graphitic carbon nitride, TiO₂ nanoparticles, and titania nanosheets) as photo-catalysts in polymerization,^{40–45} it is surprising to observe this rapid p-Ti₃C₂T_x-initiated polymerization in the absence of light illumination.

Results and discussion

In our work, p-Ti₃C₂T_x nanosheets are synthesized by a modified MILD etching method (LiF/HCl) at room temperature.⁴⁶ The modified etching process is performed under bath sonication, defined as a sonication-assisted MILD etching method. As a result, delamination is simultaneously completed along with the etching process. The p-Ti₃C₂T_x nanosheets exhibit a new peak at $2\theta \approx 6^\circ$ in the XRD pattern (Fig. 1a), while the 2θ peak at 39° (characteristic of the precursor Ti₃AlC₂) almost disappears in the p-Ti₃C₂T_x sample. The appearance of a low angle peak is attributed to a typical (002) plane for most reported MXenes, in agreement with previous studies.⁴⁶ This result indicates the high etching efficiency in the sonication-assisted MILD etching process. These obtained nanosheets are highly water dispersible. The polymerization is first examined with *N*-isopropylacrylamide (NIPAM). By adding NIPAM monomers into the nanosheet suspension, the homogeneous mixture is further purged with N₂ for 5 minutes. Polymerization is easily detected from the increase of viscosity within minutes. To be specific, the reaction vessel is well protected against light during the whole N₂ purging and polymerization process. Using a high concentration of the monomer, a hydrogel is formed. Typically, in the

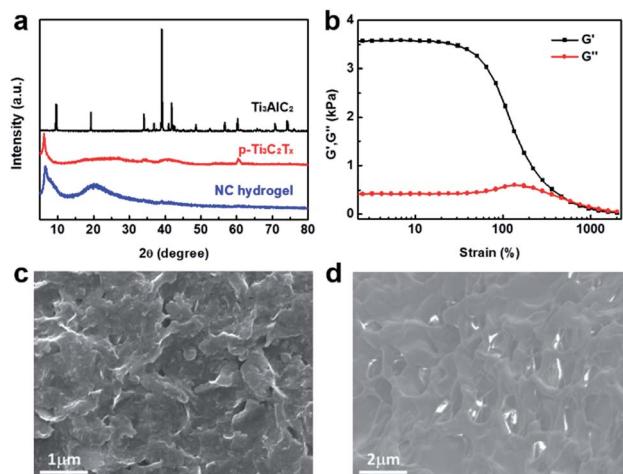


Fig. 1 (a) XRD patterns of Ti₃AlC₂, p-Ti₃C₂T_x and PNIPAM-based NC hydrogels. (b) The storage and loss moduli of PNIPAM-based NC hydrogels obtained by continuous strain sweep at 25 °C. SEM images of the p-Ti₃C₂T_x (c) and the PNIPAM-based NC hydrogels (d).

presence of p-Ti₃C₂T_x, gelation of NIPAM takes place at a concentration above 8 wt%.

The gelation of NC hydrogels was established by rheology tests. The storage modulus (G') is dominant over the loss modulus (G'') over a broad strain range, demonstrating that NC hydrogels have a good linear viscoelastic region (Fig. 1b). Nevertheless, G' of hydrogels decreases dramatically when the strain is more than 40%. The G'' surpasses G' when the strain is over 600%, indicating the destruction of hydrogels. Nevertheless, alternate-step strain sweeps (30 s at a strain of 800% for destruction, 3 min at a strain of 0.1% for recovery) reveal the good recovery properties of NC hydrogels for 6 cycles (Fig. S1†). This self-recovery is a characteristic of physical crosslinking. For poly(*N*-isopropylacrylamide) (PNIPAM) hydrogels, hydrogen bonds, hydrophobic interaction and chain entanglement cooperatively induce gelation.⁴⁷

Before polymerization, pristine p-Ti₃C₂T_x has a layered sheet structure (Fig. 1c). In contrast, the NC hydrogel sample obtained after polymerization displays a distinct interconnected network structure (Fig. 1d). This indicates that the layered p-Ti₃C₂T_x nanosheets are wrapped by polymer chains. Although the nanosheets are embedded in the polymer matrix, the NC hydrogel still exhibits periodic XRD peaks of p-Ti₃C₂T_x (Fig. 1a).

The polymerization is confirmed by ¹H-nuclear magnetic resonance (¹H-NMR) measurement. The ¹H-NMR spectrum of the NC hydrogel has new broad signals around 1.5 ppm and 2.0 ppm, attributed to the protons of polymer backbone chains. On the other hand, peaks around 5.5–6.3 ppm, attributed to the protons of the monomeric double bond, are greatly decreased after polymerization (Fig. 2a).

Besides NIPAM described above, other monomers including acrylamide (AAm), *N,N*-dimethylacrylamide (DMA), methyl methacrylate (MMA) and hydroxyethyl methacrylate (HEMA) are also successfully polymerized in the presence of Ti₃C₂T_x (Fig. 2b). Like NIPAM, AAm and DMA are able to form hydrogels



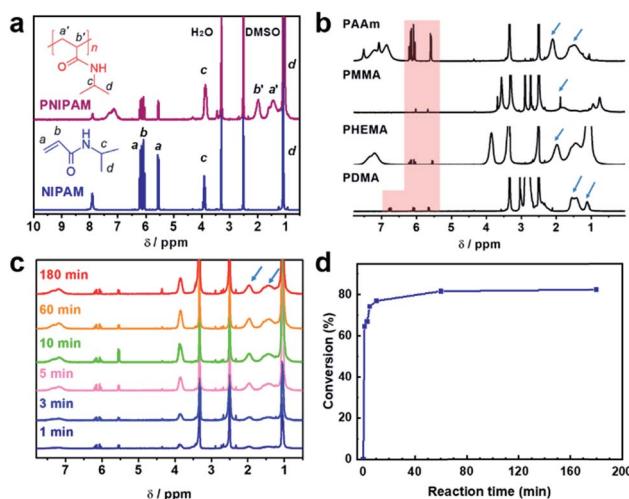


Fig. 2 (a) ¹H-NMR spectra of the NIPAM/p-Ti₃C₂T_x mixture (before N₂ purging) and NC hydrogel (after N₂ purging and polymerization). (b) Typical ¹H-NMR spectra of the different polymers initiated by p-Ti₃C₂T_x. PAAm: poly(acrylamide); PMMA: poly(methyl methacrylate); PHEMA: poly(hydroxyethyl methacrylate); PDMA: poly(*N,N*-dimethylacrylamide); solvent, DMSO-d₆. The arrows denote the proton signals in the polymer backbone. The shaded areas denote the proton signals in the monomeric double bonds. (c) ¹H-NMR spectra of PNIPAM initiated by p-Ti₃C₂T_x and (d) conversion of NIPAM to its polymer at different reaction times. All samples were polymerized using a 1.0 wt% NIPAM monomer concentration and 0.35 mg mL⁻¹ Ti₃C₂T_x nanosheet concentration.

in the presence of p-Ti₃C₂T_x (Fig. S2†). ¹H-NMR measurement also confirmed the polymerization of different monomers (Fig. 2b and Table S1†). In this work, we selected NIPAM for the preparation of the NC polymer or hydrogel unless specified.

The kinetics of the reaction was explored by monitoring the proton signals as a function of reaction time using NMR measurement. As shown in Fig. 2c, the chemical shift signals from the backbone protons increase with polymerization time. Meanwhile, the signals from protons of the monomeric double bond decrease with the polymerization time. The monomer conversion of NIPAM increases with time gradually (Fig. 2d). These results are similar to the features of classical free-radical polymerization. However, the conversion of NIPAM to PNIPAM was decreased to 73.5% in the NC hydrogel prepared using a monomer concentration of 8 wt% (Fig. S3†). This decrease is caused by the slow diffusion of monomers in the hydrogels.

To verify the mechanism of initiation, electron paramagnetic resonance (EPR) was used to analyze the radical species. In addition, another kind of Ti₃C₂T_x nanosheet was prepared by the conventional MILD etching method as a control.⁴⁶ Using DMPO as the spin trapping agent, p-Ti₃C₂T_x from the sonication-assisted process displays an obvious ·OH signal (Fig. 3a), while conventional Ti₃C₂T_x shows no EPR signals. Although several studies reported highly reactive hydroxyl radicals could be generated by photocatalysis as Ti₃C₂T_x is partially oxidized to TiO₂,^{48–50} the experiments for the MXene-initiated radical polymerization in this work are performed in the darkness and the absence of light illumination.

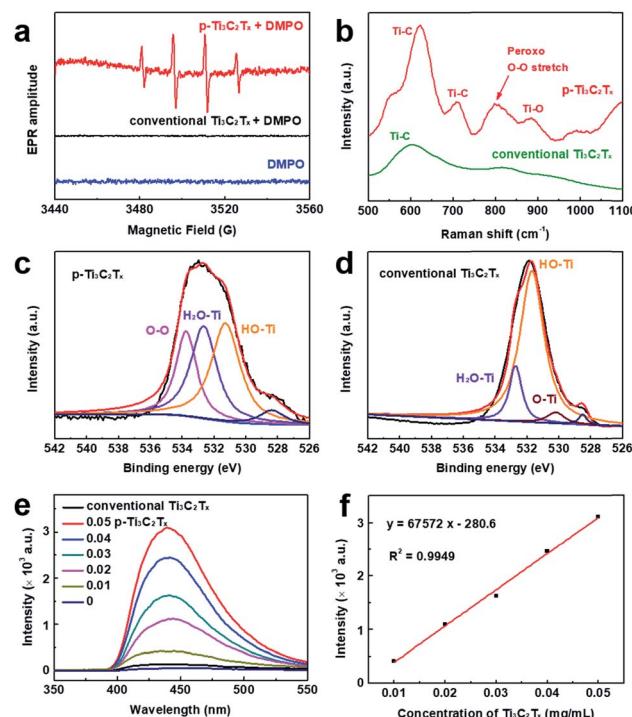


Fig. 3 (a) EPR spectra of p-Ti₃C₂T_x and conventional Ti₃C₂T_x obtained by DMPO trapping for ·OH characterization. (b) Raman spectra of p-Ti₃C₂T_x and conventional Ti₃C₂T_x. (c) High-resolution XPS spectra of p-Ti₃C₂T_x nanosheets prepared by the sonication-assisted MILD etching method in the O 1s region. (d) High-resolution XPS spectra in the O 1s region of Ti₃C₂T_x nanosheets prepared by the conventional method. (e) Fluorescence spectra of luminol (0.025 M) upon addition of different concentrations of p-Ti₃C₂T_x nanosheets. The final concentration of p-Ti₃C₂T_x was from 0.01 to 0.05 mg mL⁻¹; $\lambda_{\text{ex}} = 290$ nm. (f) Calibration curve of the luminol assay for p-Ti₃C₂T_x.

Interestingly, the catalytic activity of p-Ti₃C₂T_x in the polymerization disappeared after being stored over one week at room temperature. Conventional Ti₃C₂T_x is unable to catalyze the polymerization, either for the as-prepared samples or for stored samples. This phenomenon demonstrates that p-Ti₃C₂T_x nanosheets themselves generate the reactive species for initiation.

To gain insight into the structure of p-Ti₃C₂T_x nanosheets, Raman and X-ray photoelectron spectroscopy (XPS) were employed. The peroxy O–O stretching vibration in Raman spectroscopy is found at 796 cm⁻¹ for p-Ti₃C₂T_x instead of conventional Ti₃C₂T_x (Fig. 3b),⁵¹ which strongly supports the fact that peroxy is decorated on the surface of Ti₃C₂T_x. The XPS survey spectrum demonstrated the presence of Ti, C, F and O in both kinds of Ti₃C₂T_x (Fig. S4†). The high-resolution XPS O 1s spectrum of p-Ti₃C₂T_x can be split into 4 peaks including absorbed oxygen, HO-Ti, H₂O-Ti (absorbed H₂O) and O–O at 528.2, 531.4, 532.8, and 533.8 eV, respectively (Fig. 3c).^{52,53} The p-Ti₃C₂T_x nanosheets exhibit an obvious O–O peak (Fig. 3c), differing from conventional Ti₃C₂T_x (Fig. 3d). In addition, this peak disappears after storage for over one week at room temperature. In Fig. S5,† an O-Ti/TiO₂ peak becomes dominant in inactive p-Ti₃C₂T_x, suggesting a conversion to TiO₂.⁵⁴ Thus, we deduce that the surface decoration of peroxy O–O in the p-

$\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets is ascribed to the generation of reactive radical species.

A further understanding of the occurrence of peroxy species in $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets was gained when we found that the sonication process in the etching method has a profound effect on the formation of peroxide-decorated $\text{Ti}_3\text{C}_2\text{T}_x$. It was reported that titanium peroxide can be formed if TiO_2 reacts with H_2O_2 .^{55–57} Taking into account the fact that H_2O_2 is produced from cavitation bubbles during sonication^{55,58} and $-\text{OH}$ groups are anchored onto the surface of $\text{Ti}_3\text{C}_2\text{T}_x$ during the etching process, it is well-reasoned that p- $\text{Ti}_3\text{C}_2\text{T}_x$ is formed in the sonication-assisted MILD etching process. Conventional $\text{Ti}_3\text{C}_2\text{T}_x$ is inactive to initiate polymerization because no sonication treatment is involved in the preparation process.

In order to quantify p- $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets, the luminol assay was used. Luminol is an extensively used chemiluminescence reagent which generates a blue glow in the presence of an appropriate oxidizing agent. Peroxide, $\text{Fe}^{2+}/\text{Fe}^{3+}$ and various biochemical substrates like ALP and glucose were detected successfully by using the luminol system. In this work, different concentrations of p- $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets are incubated with a luminol solution. In the case of luminol mixed with p- $\text{Ti}_3\text{C}_2\text{T}_x$, the fluorescence intensity increases with the increase of the concentration of p- $\text{Ti}_3\text{C}_2\text{T}_x$ (Fig. 3e). A good linear relationship between the fluorescence intensity and the concentration of p- $\text{Ti}_3\text{C}_2\text{T}_x$ ranging from 0.01 to 0.05 mg mL^{−1} was obtained (Fig. 3f). The optimal concentration of $\text{Ti}_3\text{C}_2\text{T}_x$ to initiate polymerization is above 0.05 mg mL^{−1}.

The proposed mechanism of polymerization is shown in Fig. 4a. After the initiation by p- $\text{Ti}_3\text{C}_2\text{T}_x$, two kinds of radicals (hydroxyl and peroxy radicals) may be generated. After chain propagation, two kinds of polymers (grafted polymer and free polymer) are therefore created. Although some peroxide salts (such as ammonium persulfate) can initiate the polymerization

of acrylic monomers with a similar mechanism (EPR spectrum in Fig. S6†), the p- $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets themselves are functional materials and have the ability to trigger a surface-initiated “grafting-from” polymerization. The resulting graft polymers were examined using acidic sediment. Pristine p- $\text{Ti}_3\text{C}_2\text{T}_x$ is not stable to be dispersed in an acidic solution due to the decreased negative charges at lower pH (Fig. 4b). In addition, p- $\text{Ti}_3\text{C}_2\text{T}_x$ is also not stable at acidic pH when physically mixed with free PNIPAM. In contrast, the NC polymer dispersion remains well dispersed at acidic pH. This result demonstrates that the $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets of the NC polymer dispersion are covalently grafted with PNIPAM. The steric stability of grafted PNIPAM ensures the high stability of the NC polymer dispersion.

Furthermore, the grafted polymer can be observed by TEM and AFM. The size of p- $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets is about 200 nm with a single or multilayered structure (Fig. 4c). The morphology of the NC polymer dispersion is different from that of pristine p- $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets. Instead, clusters of thin nanosheets of p- $\text{Ti}_3\text{C}_2\text{T}_x$ are found in the polymer dispersion (Fig. S7†). After being stained with phosphotungstic acid, NC polymer shows a dark core with a grey shell coating in the TEM image (Fig. 4d). Meanwhile, the height of p- $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets in the NC polymer is larger than that of pristine p- $\text{Ti}_3\text{C}_2\text{T}_x$ as shown in the AFM image, suggesting the occurrence of polymer grafting on the surface of nanosheets (Fig. S8†). Unfortunately, the molecular weight of the polymers in the NC polymer dispersion is not determined in this work and remains a challenge due to the difficulty of the separation and/or detachment of the grafted polymer from the MXenes in the NC polymer dispersion.

As is well-known, PNIPAM is a typical thermosensitive polymer and shows temperature-dependent optical transmittance. In this work, the PNIPAM-based NC hydrogels are synthesized using p- $\text{Ti}_3\text{C}_2\text{T}_x$ as the initiator. By monitoring the transmittance as a function of time, the NC hydrogel has a LCST of 34 °C, close to the transition temperature of the pure PNIPAM hydrogel.⁵⁹ The little influence of p- $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets on the LCST is probably because of the low content of p- $\text{Ti}_3\text{C}_2\text{T}_x$ (0.1 mg mL^{−1}) in the hydrogel matrix. Below the LCST, the NC hydrogel is transparent, whereas above the LCST, the hydrogels become turbid indicating the dehydration of the PNIPAM networks (Fig. 5a).

To verify the photothermal performance of PNIPAM-based NC hydrogels, the absorbance spectra were measured. The NC hydrogels with different concentrations of p- $\text{Ti}_3\text{C}_2\text{T}_x$ nanosheets show strong NIR absorption (650–900 nm, Fig. 5b). The photothermal conversion of NC hydrogels is proportional to the power density of irradiation (Fig. 5c). Under exposure to 808 nm laser irradiation, the temperature of NC hydrogels gradually increase as a function of irradiation time and power. The photothermal-conversion efficiency (η) of NC hydrogels was calculated to be 34.7% at 808 nm irradiation (Fig. S9†). In addition, the NC hydrogels retain similar photothermal effects after 5 cycles (Fig. S10†). The photothermal properties of the NC hydrogel under light beyond the infrared region were also studied. Compared to other light wavelengths (405, 532 and 648

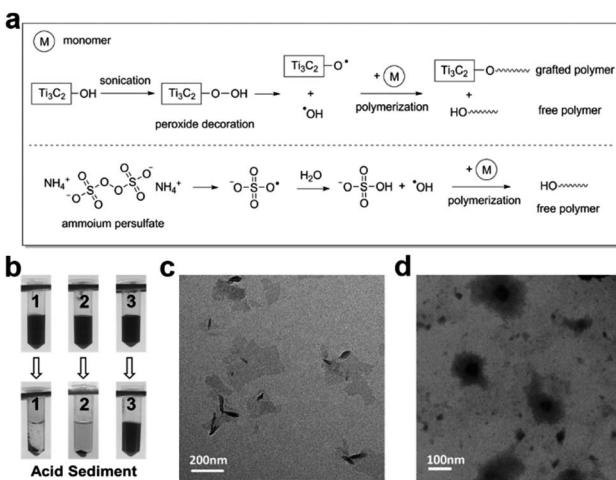


Fig. 4 (a) The proposed mechanism of polymerization initiated by p- $\text{Ti}_3\text{C}_2\text{T}_x$ or ammonium persulfate. (b) Acid stability of three p- $\text{Ti}_3\text{C}_2\text{T}_x$ -containing dispersions. Sample 1: pristine p- $\text{Ti}_3\text{C}_2\text{T}_x$; Sample 2: physical mixture of PNIPAM and p- $\text{Ti}_3\text{C}_2\text{T}_x$; Sample 3: PNIPAM-based NC polymer. The TEM image of p- $\text{Ti}_3\text{C}_2\text{T}_x$ (c) and the PNIPAM-based NC polymer (d).



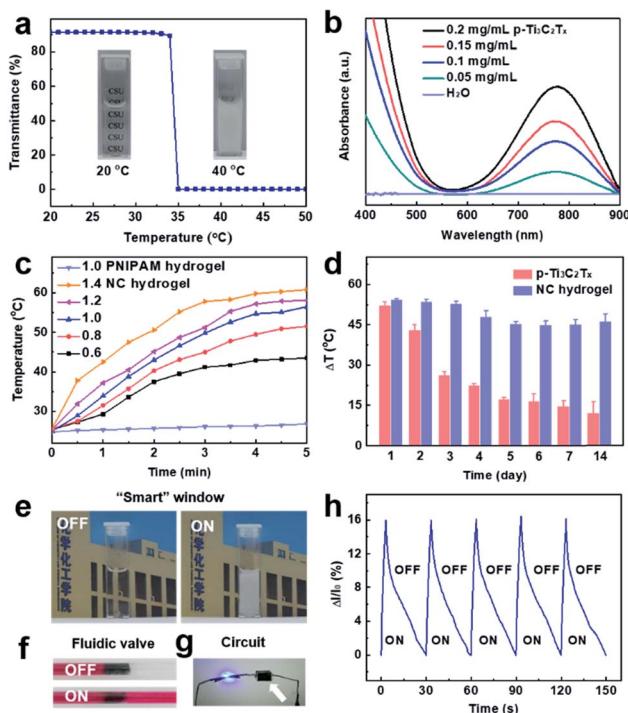


Fig. 5 (a) Transmittance of the PNIPAM-based NC hydrogel as a function of temperature. The inset shows the thermo-responsive-ness of the hydrogels (80 mg mL⁻¹ NIPAM monomer, 0.1 mg mL⁻¹ p-Ti₃C₂T_x). (b) Normalized Vis-NIR absorbance of the NC hydrogel prepared using different p-Ti₃C₂T_x concentrations. (c) Temperature-evolution of the PNIPAM hydrogel (control) and NC hydrogel under 808 nm irradiation at different power (unit: W cm⁻², 0.35 mg mL⁻¹ p-Ti₃C₂T_x). (d) Photothermal stability of the p-Ti₃C₂T_x nanosheet suspension and the PNIPAM-based NC hydrogel (0.35 mg mL⁻¹ p-Ti₃C₂T_x) as a function of storage time. (e) Application of the PNIPAM-based NC hydrogel as a simulated "smart" window. Left: before irradiation. Right: 808 nm irradiation (0.375 W cm⁻²) for 5 min. (f) Application of the NC hydrogel as a fluidic valve. Top: before irradiation. Bottom: 808 nm irradiation (1.0 W cm⁻²) for 30 s. (g) Electrical conductivity of the PNIPAM-based NC hydrogel. The LED of a demo circuit connected to the PNIPAM-based NC hydrogel (indicated by a white arrow) glows. (h) Application of the PNIPAM-based NC hydrogel as a photodetector under 808 nm irradiation for five cycles (On: 3 s, 1.0 W cm⁻²; Off: 27 s; ΔI/I₀ indicates the NIR-dependent current change behavior).

nm), NIR (808 nm) triggers a pronounced photothermal effect in the NC hydrogel (Fig. S11†).

We evaluated the photothermal stability of Ti₃C₂T_x nanosheets and PNIPAM-based NC hydrogels under 808 nm irradiation (1.0 W cm⁻²) for 5 min. The results indicated that the pristine p-Ti₃C₂T_x nanosheets were almost oxidized and degraded gradually at low temperature (4 °C) in one week (Fig. 5d). In contrast, p-Ti₃C₂T_x embedded in the hydrogel matrix shows remarkable photothermal stability. The oxidation kinetics of p-Ti₃C₂T_x is quite slow in the polymer hydrogel matrix, which allows the MXenes to have a prolonged shelf life and reproducible functional properties in its applications.⁵⁴ The long-term stability of the NC hydrogel depends on the inherent oxidation potential and environmental factors. The strategy of slowing down the kinetics in the semi-solid

hydrogel media could greatly improve the stability of MXenes against oxidation.

In this work, the excellent photothermal performance and stability endow the PNIPAM-based NC hydrogel with potential applications such as NIR-responsive "smart" windows and fluidic valves based on its opaque character and volume shrinkage. As a proof of concept, simple models of a "smart" window and fluidic valve are rationally designed (Fig. 5e and f). The NC hydrogels also exhibit higher electrical conductivity than the pristine PNIPAM hydrogel. For example, when the p-Ti₃C₂T_x concentration is about 1.0 mg mL⁻¹, the conductivity of the NC hydrogel increases from 0.0098 S m⁻¹ to 0.019 S m⁻¹. The enhancement in conductivity of the hydrogels depends on the content of p-Ti₃C₂T_x.⁶⁰ The LED of a demo circuit glows throughout the NC hydrogels (Fig. 5g). Therefore, the NC hydrogel is designed to be a NIR detector. Upon irradiation of NIR for 3 s, the current increases abruptly (Fig. 5h). The reason for the increase of current may be ascribed to photoconductive effects and the accelerated electron mobility with a temperature rise.⁶¹⁻⁶³ A decay of current to the original value is observed after turning off the laser. Repeatable signal patterns are observed in response to NIR irradiation, suggesting the potential application of NC hydrogels as infrared photodetectors and photovoltaics. Additionally, since both MXenes and PNIPAM are biocompatible,^{32,33,64} the NC hydrogel may be further designed as a photothermal agent for localized anticancer treatment or enzyme bioreactors.

Conclusions

In summary, peroxide-decorated Ti₃C₂T_x nanosheets are synthesized *via* the sonication-assisted MILD etching method. The nanosheets are able to generate reactive radical species, which initiate polymerization of acrylic monomers in a simple N₂ purging procedure. Our approach to initiate the acrylic monomers has good universality, as confirmed by over 5 monomers. Physically crosslinked PNIPAM-based NC hydrogels are facilely formed using a high monomer concentration. The NC hydrogels are sensitive to temperature and NIR, which can serve as a means to adjust transmittance, volume, and conductivity. Their excellent photothermal performance and electrical conductivity guarantee their future application as remotely light controlled "smart" windows, fluidic valves, and photodetectors.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was financially supported by the National Natural Science Foundation of China (No. 21574147 and 21374133), Hunan Provincial Natural Science Foundation of China (No. 2018JJ2483), and Innovation-Driven Project of Central South University (No. 2017CX020).



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