Near-infrared light-responsive hydrogels via peroxide-decorated MXene-initiated polymerization†

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Two-dimensional MXene Ti3C2Tx nanosheets with peroxide decoration (p-Ti3C2Tx) are synthesized by a sonication-assisted MILD etching method. The obtained MXenes can generate hydroxyl radical species and act as an initiator for free-radical polymerization of a series of acrylic monomers without the use of light illumination or co-initiators. The monomers analyzed include acrylamide, N-isopropylacrylamide (NIPAM), N,N-dimethylacrylamide, methyl methacrylate, and hydroxyethyl methacrylate. By simply mixing N-isopropylacrylamide monomers and p-Ti3C2Tx nanosheets under deoxygenated conditions, PNIPAM-based nanocomposite hydrogels are synthesized using a high concentration of the monomer. The nanocomposite hydrogels have a photothermal conversion efficiency of 34.7% and photothermal stability superior to that of pristine Ti3C2Tx. Taking advantage of the thermal responsive behavior of PNIPAM, the nanocomposite hydrogels are successfully exploited as remotely near-infrared light controlled “smart” windows, fluidic valves and photodetectors.

Introduction

“Smart” responsive hydrogels are now being widely studied in various fields such as soft robotics,1,2 sensors,3,4 microdevices,5 drug delivery,6,7 cell/tissue engineering scaffolds,8,9 and reaction vessels.10,11 Exposure to different stimuli (for example light, temperature, pH, and electric field) may cause these hydrogels to swell/shrink, bend/unbend, or swim/crawl, mimicking natural creatures with responsive behavior.12–15 In particular, light is a distinct stimulus to tune the properties of a hydrogel spatiotemporally in a remote controlled manner with great ease and convenience.5,16–19 During the past decade, increasing attention has been paid to near-infrared (NIR) light-absorbing materials. As opposed to UV-Vis light, NIR light has advantages of deep penetration, low energy, and little damage to structures. Additionally, some NIR-absorbing materials have strong photothermal conversion ability resulting in a rapid temperature increase at the irradiation site.20–22 Moreover, NIR can trigger phase separation or phase transition of thermo-responsive polymer networks, causing various physical and morphological changes. Hence, NIR-responsive hydrogels are being exploited for different applications, including smart windows, microfluidic valves, 3D microfabrication of complex objects, and drug delivery.18,20,21–23

MXenes have been an emerging class of graphene-like two-dimensional nanomaterials since their first discovery in 2011.27 Owing to their properties of a functionalized surface, small band gaps, and excellent photothermal properties and conductivity, they have diverse applications in batteries, supercapacitors, photocatalysis, electrocatalysis, and biomedicine.28–31 Ti3C2Tx, a representative MXenes, has abundant functional groups (i.e. –OH, –O or –F) and is well dispersible in aqueous solution. Ti3C2Tx has outstanding photothermal properties, and Ti3C2Tx-based hydrogels have been applied in the remote light-controlled microfluidic pipeline.34,35 Ti3C2Tx incorporated hydrogels were also recently applied as strain sensors and exhibited outstanding tensile strain sensitivity, with a gauge factor one order of magnitude higher than that of pristine hydrogels.36 Notably, Ti3C2Tx participates in many chemical processes (such as CO oxidation, oxygen evolution, and oxygen reduction), by acting as catalysts, cocatalysts or supporting substrates. Ti3C2Tx itself can catalyze in situ polymerization of aromatic EDOT or pyrrole using the mechanism of charge-transfer-induced polymerization for the polymerization process.37 To the best of our knowledge, the catalytic role of MXenes in radical polymerization has not been reported yet.37–39

Herein, we aim to design MXene-based NIR-responsive nanocomposite (NC) hydrogels towards remote-controlled...
To be specific, we synthesized a novel kind of Ti₃C₂Tₓ nanosheet with peroxide decoration (p-Ti₃C₂Tₓ), which itself acts as an initiator in the polymerization of acrylic monomers without any light irradiation or co-initiators. The preparation procedure for the MXene-based hydrogels is quite simple and efficient (Scheme 1). Briefly, after adding monomers into a freshly prepared p-Ti₃C₂Tₓ nanosheet suspension in an ice bath, polymerization occurs within several minutes after the N₂ purging process. Gellation is observed at a high concentration of the monomer. In view of the widespread use of inorganic semiconductor nanoparticles (such as V₂C, perovskite nanocrystals, graphitic carbon nitride, TiO₂ nanoparticles, and titania nanosheets) as photo-catalysts in polymerization, it is surprising to observe this rapid p-Ti₃C₂Tₓ-initiated polymerization in the absence of light illumination.

Results and discussion

In our work, p-Ti₃C₂Tₓ nanosheets are synthesized by a modified MILD etching method [LiF/HCl] at room temperature. The modified etching process is performed under bath sonication, defined as a sonication-assisted MILD etching method. As a result, delamination is simultaneously completed along with the etching process. The p-Ti₃C₂Tₓ nanosheets exhibit a new peak at 2θ ≈ 6° in the XRD pattern (Fig. 1a), while the 2θ peak at 39° (characteristic of the precursor Ti₃AlC₂) almost disappears in the p-Ti₃C₂Tₓ sample. The appearance of a low angle peak is attributed to a typical [002] plane for most reported MXenes, in agreement with previous studies. This result indicates the high etching efficiency in the sonication-assisted MILD etching process. These obtained nanosheets are highly water dispersible. The polymerization is first examined with N-isopropylacrylamide (NIPAM). By adding NIPAM monomers into the nanosheet suspension, the homogeneous mixture is further purged with N₂ for 5 minutes. Polymerization is easily detected from the increase of viscosity within minutes. To be specific, the reaction vessel is well protected against light during the whole N₂ purging and polymerization process. Using a high concentration of the monomer, a hydrogel is formed. Typically, in the presence of p-Ti₃C₂Tₓ, gelation of NIPAM takes place at a concentration above 8 wt%.

The gelation of NC hydrogels was established by rheology tests. The storage modulus (G') is dominant over the loss modulus (G'') over a broad strain range, demonstrating that NC hydrogels have a good linear viscoelastic region (Fig. 1b). Nevertheless, G' of hydrogels decreases dramatically when the strain is more than 40%. The G'' surpasses G' when the strain is over 600%, indicating the destruction of hydrogels. Nevertheless, alternate-step strain sweeps (30 s at a strain of 800% for destruction, 3 min at a strain of 0.1% for recovery) reveal the good recovery properties of NC hydrogels for 6 cycles (Fig. S1†). This self-recovery is a characteristic of physical crosslinking. For poly(N-isopropylacrylamide) (PNIPAM) hydrogels, hydrogen bonds, hydrophobic interaction and chain entanglement cooperatively induce gelation.⁴⁷

Before polymerization, pristine p-Ti₃C₂Tₓ has a layered sheet structure (Fig. 1c). In contrast, the NC hydrogel sample obtained after polymerization displays a distinct interconnected network structure (Fig. 1d). This indicates that the layered p-Ti₃C₂Tₓ nanosheets are wrapped by polymer chains. Although the nanosheets are embedded in the polymer matrix, the NC hydrogel still exhibits periodic XRD peaks of p-Ti₃C₂Tₓ (Fig. 1a).

The polymerization is confirmed by 1H-nuclear magnetic resonance (1H-NMR) measurement. The 1H-NMR spectrum of the NC hydrogel has new broad signals around 1.5 ppm and 2.0 ppm, attributed to the protons of polymer backbone chains. On the other hand, peaks around 5.5–6.3 ppm, attributed to the protons of the monomeric double bond, are greatly decreased after polymerization (Fig. 2a).

Besides NIPAM described above, other monomers including acrylamide (AAm), N,N-dimethylacrylamide (DMA), methyl methacrylate (MMA) and hydroxyethyl methacrylate (HEMA) are also successfully polymerized in the presence of Ti₃C₂Tₓ (Fig. 2b). Like NIPAM, AAm and DMA are able to form hydrogels.
in the presence of p-Ti$_3$C$_2$T$_x$ (Fig. S2†). $^1$H-NMR measurement also confirmed the polymerization of different monomers (Fig. 2b and Table S1†). In this work, we selected NIPAM for the preparation of the NC polymer or hydrogel unless specified.

The kinetics of the reaction was explored by monitoring the proton signals as a function of reaction time using NMR measurement. As shown in Fig. 2c, the chemical shift signals from the backbone protons increase with polymerization time. Meanwhile, the signals from protons of the monomeric double bond decrease with the polymerization time. The monomer conversion of NIPAM increases with time gradually (Fig. 2d). These results are similar to the features of classical free-radical polymerization. However, the conversion of NIPAM to PNIPAM decreased to 73.5% in the NC hydrogel prepared using a monomer concentration of 8 wt% (Fig. S3†). This decrease is caused by the slow diffusion of monomers in the hydrogels.

To verify the mechanism of initiation, electron paramagnetic resonance (EPR) was used to analyze the radical species. In addition, another kind of Ti$_3$C$_2$T$_x$ nanosheet was prepared by the conventional MILD etching method as a control. Using DMPO as the spin trapping agent, p-Ti$_3$C$_2$T$_x$ from the sonication-assisted process displays an obvious OH signal (Fig. 3a), while conventional Ti$_3$C$_2$T$_x$ shows no EPR signals. Although several studies reported highly reactive hydroxyl radicals could be generated by photocatalysis as Ti$_3$C$_2$T$_x$ is partially oxidized to TiO$_2$, the experiments for the MXene-initiated radical polymerization in this work are performed in the darkness and the absence of light illumination. Interestingly, the catalytic activity of p-Ti$_3$C$_2$T$_x$ in the polymerization disappeared after being stored over one week at room temperature. Conventional Ti$_3$C$_2$T$_x$ is unable to catalyze the polymerization, either for the as-prepared samples or for stored samples. This phenomenon demonstrates that p-Ti$_3$C$_2$T$_x$ nanosheets themselves generate the reactive species for initiation.

To gain insight into the structure of p-Ti$_3$C$_2$T$_x$ nanosheets, Raman and X-ray photoelectron spectroscopy (XPS) were employed. The peroxy O–O stretching vibration in Raman spectroscopy is found at 796 cm$^{-1}$ for p-Ti$_3$C$_2$T$_x$ instead of conventional Ti$_3$C$_2$T$_x$ (Fig. 3b), which strongly supports the fact that peroxy is decorated on the surface of Ti$_3$C$_2$T$_x$. The XPS survey spectrum demonstrated the presence of Ti, C and O in both kinds of Ti$_3$C$_2$T$_x$ (Fig. S4†). The high-resolution XPS O 1s spectrum of p-Ti$_3$C$_2$T$_x$ can be split into 4 peaks including absorbed oxygen, HO–Ti, H$_2$O–Ti (absorbed H$_2$O) and O–O at 528.2, 531.4, 532.8, and 533.8 eV, respectively (Fig. 3c). The p-Ti$_3$C$_2$T$_x$ nanosheets exhibit an obvious O–O peak (Fig. 3c), differing from conventional Ti$_3$C$_2$T$_x$ (Fig. 3d). In addition, this peak disappears after storage for over one week at room temperature. In Fig. S5, an O–Ti/O=O peak becomes dominant in inactive p-Ti$_3$C$_2$T$_x$, suggesting a conversion to TiO$_2$. Thus, we deduce that the surface decoration of peroxy O–O in the p-
Ti$_3$C$_2$T$_x$ nanosheets is ascribed to the generation of reactive radical species.

A further understanding of the occurrence of peroxo species in Ti$_3$C$_2$T$_x$ nanosheets was gained when we found that the sonication process in the etching method has a profound effect on the formation of peroxide-decorated Ti$_3$C$_2$T$_x$. It was reported that titanium peroxide can be formed if TiO$_2$ reacts with H$_2$O$_2$. Taking into account the fact that H$_2$O$_2$ is produced from cavitation bubbles during sonication and -OH groups are anchored onto the surface of Ti$_3$C$_2$T$_x$ during the etching process, it is well-reasoned that p-Ti$_3$C$_2$T$_x$ is formed in the sonication-assisted MILD etching process. Conventional Ti$_3$C$_2$T$_x$ is inactive to initiate polymerization because no sono-cation treatment is involved in the preparation process.

In order to quantify p-Ti$_3$C$_2$T$_x$ nanosheets, the luminol assay was used. Luminol is an extensively used chemiluminescence reagent which generates a blue glow in the presence of an appropriate oxidizing agent. Peroxide, Fe$^{2+}$/Fe$^{3+}$ and various biochemical substrates like ALP and glucose were detected successfully by using the luminol system. In this work, different concentrations of p-Ti$_3$C$_2$T$_x$ nanosheets are incubated with a luminol solution. In the case of luminol mixed with p-Ti$_3$C$_2$T$_x$, the fluorescence intensity increases with the increase of the concentration of p-Ti$_3$C$_2$T$_x$ (Fig. 3e). A good linear relationship between the fluorescence intensity and the concentration of p-Ti$_3$C$_2$T$_x$ ranging from 0.01 to 0.05 mg mL$^{-1}$ was obtained (Fig. 3f). The optimal concentration of p-Ti$_3$C$_2$T$_x$ to initiate polymerization is above 0.05 mg mL$^{-1}$.

The proposed mechanism of polymerization is shown in Fig. 4a. After the initiation by p-Ti$_3$C$_2$T$_x$, two kinds of radicals (hydroxyl and peroxyl radicals) may be generated. After chain propagation, two kinds of polymers (grafted polymer and free polymer) are therefore created. Although some peroxide salts (such as ammonium persulfate) can initiate the polymerization of acrylic monomers with a similar mechanism (EPR spectrum in Fig. S6†), the p-Ti$_3$C$_2$T$_x$ nanosheets themselves are functional materials and have the ability to trigger a surface-initiated “grafting-from” polymerization. The resulting graft polymers were examined using acidic sediment. Pristine p-Ti$_3$C$_2$T$_x$ is not stable to be dispersed in an acidic solution due to the decreased negative charges at lower pH (Fig. 4b). In addition, p-Ti$_3$C$_2$T$_x$ is also not stable at acidic pH when physically mixed with free PNIPAM. In contrast, the NC polymer dispersion remains well dispersed at acidic pH. This result demonstrates that the Ti$_3$C$_2$T$_x$ nanosheets of the NC polymer dispersion are covalently grafted with PNIPAM. The steric stability of grafted PNIPAM ensures the high stability of the NC polymer dispersion.

Furthermore, the grafted polymer can be observed by TEM and AFM. The size of p-Ti$_3$C$_2$T$_x$ nanosheets is about 200 nm with a single or multilayered structure (Fig. 4c). The morphology of the NC polymer dispersion is different from that of pristine p-Ti$_3$C$_2$T$_x$ nanosheets. Instead, clusters of thin nanosheets of p-Ti$_3$C$_2$T$_x$ are found in the polymer dispersion (Fig. S7†). After being stained with phosphotungstic acid, NC polymer shows a dark core with a grey shell coating in the TEM image (Fig. 4d). Meanwhile, the height of p-Ti$_3$C$_2$T$_x$ nanosheets in the NC polymer is larger than that of pristine p-Ti$_3$C$_2$T$_x$ as shown in the AFM image, suggesting the occurrence of polymer grafting on the surface of nanosheets (Fig. S8†). Unfortunately, the molecular weight of the polymers in the NC polymer dispersion is not determined in this work and remains a challenge due to the difficulty of the separation and/or detachment of the grafted polymer from the MXenes in the NC polymer dispersion.

As is well-known, PNIPAM is a typical thermosensitive polymer and shows temperature-dependent optical transmittance. In this work, the PNIPAM-based NC hydrogels are synthesized using p-Ti$_3$C$_2$T$_x$ as the initiator. By monitoring the transmittance as a function of time, the NC hydrogel has a LCST of 34 °C, close to the transition temperature of the pure PNIPAM hydrogel. The little influence of p-Ti$_3$C$_2$T$_x$ nanosheets on the LCST is probably because of the low content of p-Ti$_3$C$_2$T$_x$ (0.1 mg mL$^{-1}$) in the hydrogel matrix. Below the LCST, the NC hydrogel is transparent, whereas above the LCST, the hydrogels become turbid indicating the dehydration of the PNIPAM networks (Fig. 5a).

To verify the photothermal performance of PNIPAM-based NC hydrogels, the absorbance spectra were measured. The NC hydrogels with different concentrations of p-Ti$_3$C$_2$T$_x$ nanosheets show strong NIR absorption (650–900 nm, Fig. 5b). The photothermal conversion of NC hydrogels is proportional to the power density of irradiation (Fig. 5c). Under exposure to 808 nm laser irradiation, the temperature of NC hydrogels gradually increase as a function of irradiation time and power. The photothermal-conversion efficiency (η) of NC hydrogels was calculated to be 34.7% at 808 nm irradiation (Fig. S9†). In addition, the NC hydrogels retain similar photothermal effects after 5 cycles (Fig. S10†). The photothermal properties of the NC hydrogel under light beyond the infrared region were also studied. Compared to other light wavelengths (405, 532 and 648
nm), NIR (808 nm) triggers a pronounced photothermal effect in the NC hydrogel (Fig. S11†).

We evaluated the photothermal stability of Ti3C2T x nanosheets and PNIPAM-based NC hydrogels under 808 nm irradiation (1.0 W cm\(^{-2}\)) for 5 min. The results indicated that the pristine p-Ti3C2T x nanosheets were almost oxidized and degraded gradually at low temperature (4 °C) in one week (Fig. 5d). In contrast, p-Ti3C2T x embedded in the hydrogel matrix shows remarkable photothermal stability. The oxidation kinetics of p-Ti3C2T x is quite slow in the polymer hydrogel matrix, which allows the MXenes to have a prolonged shelf life and reproducible functional properties in its applications.\(^{34}\) The long-term stability of the NC hydrogel depends on the inherent oxidation potential and environmental factors. The strategy of slowing down the kinetics in the semi-solid hydrogel media could greatly improve the stability of MXenes against oxidation.

In this work, the excellent photothermal performance and stability endow the PNIPAM-based NC hydrogel with potential applications such as NIR-responsive “smart” windows and fluidic valves based on its opaque character and volume shrinkage. As a proof of concept, simple models of a “smart” window and fluidic valve are rationally designed (Fig. 5e and f). The NC hydrogels also exhibit higher electrical conductivity than the pristine PNIPAM hydrogel. For example, when the p-Ti3C2T x concentration is about 1.0 mg mL\(^{-1}\), the conductivity of the NC hydrogel increases from 0.0098 S m\(^{-1}\) to 0.019 S m\(^{-1}\). The enhancement in conductivity of the hydrogels depends on the content of p-Ti3C2T x.\(^{60}\) The LED of a demo circuit glows throughout the NC hydrogels (Fig. 5g). Therefore, the NC hydrogel is designed to be a NIR detector. Upon irradiation of NIR for 3 s, the current increases abruptly (Fig. 5h). The reason for the increase of current may be ascribed to photoconductive effects and the accelerated electron mobility with a temperature rise.\(^{61–63}\) A decay of current to the original value is observed after turning off the laser. Repeatable signal patterns are observed in response to NIR irradiation, suggesting the potential application of NC hydrogels as infrared photodetectors and photovoltaics. Additionally, since both MXenes and PNIPAM are biocompatible,\(^{32,33,64}\) the NC hydrogel may be further designed as a photothermal agent for localized anticancer treatment or enzyme bioreactors.

**Conclusions**

In summary, peroxide-decorated Ti3C2T x nanosheets are synthesized via the sonication-assisted MILD etching method. The nanosheets are able to generate reactive radical species, which initiate polymerization of acrylic monomers in a simple N\(_2\) purging procedure. Our approach to initiate the acrylic monomers has good universality, as confirmed by over 5 monomers. Physically crosslinked PNIPAM-based NC hydrogels are facilely formed using a high monomer concentration. The NC hydrogels are sensitive to temperature and NIR, which can serve as a means to adjust transmittance, volume, and conductivity. Their excellent photothermal performance and electrical conductivity guarantee their future application as remotely light controlled “smart” windows, fluidic valves, and photodetectors.

**Conflicts of interest**

There are no conflicts to declare.

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Notes and references


