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Interplay between spin-crossover and luminescence in a multifunctional single crystal iron(II) complex: towards a new generation of molecular sensors†

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We present a new example of a mononuclear iron(II) complex exhibiting a correlated spin-crossover (SCO) transition and strong fluorescence, whose coordination sphere is saturated, for the first time, by six phosphorescent ligands. The interplay between SCO and light emission properties in the thermal region of the spin transition was investigated by means of magnetic, fluorescence, optical absorption and optical microscopy measurements on a single crystal. Overall, the results show an excellent correlation between fluorescence and magnetic data of the present gradual transition, indicating an extreme sensitivity of the optical activity of the ligand to the spin state of the active iron(II) ions. These results open the way for conceiving new prototypes of pressure and temperature sensors based on this synergy between SCO and luminescence properties. In particular, the fact that cooperative SCO material is not a prerequisite for obtaining such synergetic effects, is useful for the design of thin films or nanoparticles, in which the cooperativity is reduced, for appropriate implementation in nanosized devices to enhance the sensing properties at the nanoscale.

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Introduction

Among the switchable molecular solids, spin crossover (SCO) complexes are interesting specifically because of their multifunctional properties at the molecular scale and they are used as nano-switches due to many assets compared to classic silicon technology. Hence, SCO solids have potential applications in the development of new generations of electronic devices. Indeed, iron(II)-based SCO complexes were selected due to their variety of nano-structuration possibilities¹ which will be useful for implementation into devices for a set of currently considered applications, *e.g.* as displays and memory devices, molecular sensors, and rapid switches,² in addition to their

barocaloric properties for refrigeration which have been recently demonstrated.³ The SCO phenomenon, mostly in iron(II) systems, is an entropy-driven process involving the switching between the diamagnetic low-spin (LS, $S = 0$, $^1A_{1g}$) and the paramagnetic high-spin (HS, $S = 2$, $^5T_{2g}$) states, as a response to an external stimulus as a variation of temperature, pressure, light or a magnetic field,^{1a} as well as by an electric polarization leading to magneto-electric effects.⁴ As a result, different optical, magnetic, vibrational and structural properties can be recorded by a variety of physical techniques. Among the challenging quests in multi-functional materials, the combination of the SCO phenomenon with luminescence in a synergetic way will extend the innovative nature of the SCO materials to new potential applications such as magneto-optical switches or new generation of multi-sensors.^{5–8}

The attempts to achieve this goal are based on two strategies: the first one consists in the synthesis of non molecular composites combining the suited properties such for SCO-luminescent core-shell nanoparticles⁹ or electroluminescent thin films doped by SCO complexes⁷ while the second strategy is based on the design of fluorescent SCO complexes in which the fluorophore groups act as metal complexes or organic counter anions^{5a,7a,10} or are organic molecules covalently linked to the SCO complexes.^{6,8a,11} While some materials, prepared using these strategies exhibit the two desired properties,^{5a,6–11} the

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correlation between the SCO transition and the fluorescence has been rarely observed.^{6a-d,8,9,11} Among which those based on covalently linked fluorophore groups seem to be more promising multifunctional systems, since the corresponding, more rigorous synthetic approach, allows obtaining single crystals suitable for structural characterizations. Within this class of molecular systems, the first fluorescent SCO material was reported in 2011 by Garcia *et al.*,^{6a} as a dinuclear complex based on the *N*-salicylidene-4-amino-1,2,4-triazole ligand which is well-known for its strong fluorescence at the solid state. This material exhibits an abrupt SCO transition that can be tracked through the variation of the fluorescence signal. Using an original synthetic approach, J. Tao *et al.*^{6b} have reported in 2015 two fluorescent one-dimensional SCO coordination polymers, based on 3-amino-1,2,4-triazole ligand for which some of the triazole motifs were statistically grafted by the 1-pyrenecarboxaldehyde or rhodamine fluorophores. The two systems showed correlation between the thermally-induced SCO transition and associated fluorescence variation. Months later, the same group has reported another dinuclear Fe(II), almost similar to that described by Garcia group, which displayed three-step SCO behavior with multiple correlations with the fluorescence.^{6c} More recently, B. Weber *et al.*^{11a} and H.-Z. Kou *et al.*^{11b} have separately designed, respectively, a SCO Fe(II) coordination polymer and a SCO discrete Fe(II) mononuclear complex, with thermal hysteresis loops above room temperature (RT) both exhibiting fluorescent properties with strong correlations between the light emission and the spin state properties. Nevertheless, besides the two dinuclear Fe(II) complexes which have been crystallographically characterized,^{6b,c} the molecular structures of all these systems were assessed with essentially elemental analysis and X-ray powder diffraction. In addition, in none of these materials the coordination sphere was saturated by six fluorescent ligands. In this context, here we report on the mononuclear Fe(II) SCO and fluorescent complex [Fe(naph-trz)₆](tcnsme)₂·4CH₃CN, **C1** (naph-trz = **L1** = *N*-(1,2,4-triazol-4-yl)-1,8-naphthalimide; (tcnsme)[−] = 1,1,3,3-tetracyano-2-thiomethylpropenide anion), whose coordination sphere is saturated, for the first time, by six phosphorescent ligands showing a coupling between the two properties. Due to this interaction, the emerging response is more than the sum of the of the individual responses of the ligand and the SCO metal. The fluorescence and the magnetic responses reveal the interplay between the SCO and the light emission properties in the thermal region of the SCO transition. In order to identify the interaction between the phosphorescent ligand and the spin transition of the Fe(II) active center, fluorescence studies have been also performed on the ligand naph-trz (**L1**). Furthermore, to assess the influence of the spin transition on the light emission properties in complex **C1**, the homologous compound [Cu(naph-trz)₆](tcnsme)₂·4CH₃CN (**C2**) was synthesized and its thermal properties were investigated.

Results and discussion

The naph-trz molecule (**L1**) and the potassium salt K(tcnsme) were prepared according to modified synthetic strategies (see

ESI†).^{12,13} Single crystals of [M(naph-trz)₆](tcnsme)₂·4CH₃CN (M = Fe^{II} (**C1**); M = Cu^{II} (**C2**)) have been prepared under aerobic conditions by refluxing acetonitrile solutions containing [M(BF₄)₂]·xH₂O (M = Fe^{II} (**C1**); M = Cu^{II} (**C2**)), naph-trz and K(tcnsme) in a 1 : 6 : 2 molar ratio. Details of the synthesis, NMR spectra (Fig. S1–7†), Infrared spectra (Fig. S8–12†), as well as the relevant crystallographic parameters and data (Fig. S13–15, Tables S1 and 2) are reported in the ESI.†

The thermal variation of the product of the molar magnetic susceptibility per Fe(II) ion times the temperature ($\chi_m T$) shows at RT a value of 3.9 cm³ K mol^{−1} (Fig. 1a). This value is close to the expected one for an isolated HS (*S* = 2) hexacoordinated Fe(II) ion with *g* ≈ 2.28. Upon cooling, the $\chi_m T$ product remains constant down to 180 K and below this temperature, it shows a gradual decrease to reach a plateau of ca. 0.45 cm³ K mol^{−1} below ca. 70 K, indicating the presence of a gradual and almost complete HS to LS SCO with a *T*_{1/2} = 114 K, measured as the maximum of the $\delta(\chi_m T)/\delta T$ vs. *T* plot (inset in Fig. 1a). At low temperatures the $\chi_m T$ value indicates that there is a fraction of

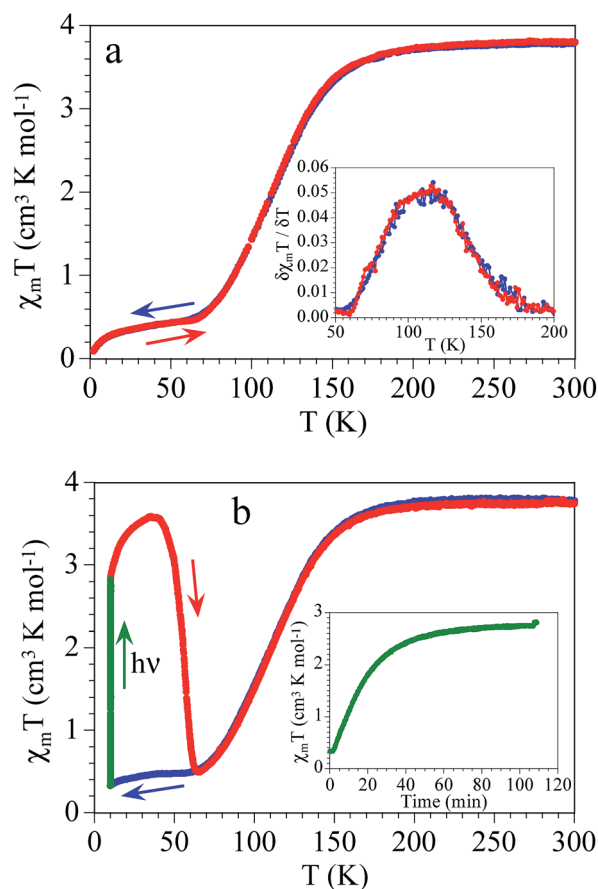


Fig. 1 (a) Thermal variation of the $\chi_m T$ product of [Fe(naph-trz)₆](tcnsme)₂·4CH₃CN (**C1**) in a cooling and heating scan in the thermal interval 2–300 K. Inset shows the thermal variation of $\frac{\delta(\chi_m T)}{\delta T}$. (b) Photo-magnetic measurements of **C1**: dark cooling (blue points), irradiation at 10 K (green points) and dark heating after irradiation (red points). Inset shows the time evolution of the signal at 10 K under irradiation. Temperature scan rate is 0.4 K min^{−1}.



ca. 10% of the molecules that remains in the HS state. This fraction shows the expected zero field splitting, responsible of the more abrupt decrease of $\chi_m T$ at lower temperatures (below 10 K).

The photomagnetic measurements performed at 10 K show a continuous increase of the magnetization (Fig. 1b) when the sample is irradiated with a green laser (532 nm) to reach saturation after ca. 2 hours (inset of Fig. 1b). The saturation value is $2.8 \text{ cm}^3 \text{ K mol}^{-1}$, indicating an almost complete recovery of the HS state. Indeed, when the sample is heated in dark, $\chi_m T$ shows a gradual increase to reach a maximum value of $3.7 \text{ cm}^3 \text{ K mol}^{-1}$ at 40 K, characteristic of a magnetic response with zero a field splitting (magnetic anisotropy) which lifts the degeneracy of the spin state $S = 2$, confirming the total photo-conversion to the HS state of C1. At higher temperature, the HS fraction relaxes back to the LS state at $T_{\text{LIESST}} = 55 \text{ K}$ (Fig. 1b). On further heating, the sample recovers the equilibrium curve with a gradual SCO temperature at $T_{1/2} = 114 \text{ K}$, as observed in Fig. 1a.

Based on the previous magnetic measurements, the crystal structure of C1 has been determined at 296 K and 100 K (see

Table 1 Fe–N distances (Å), distortion parameters ($^\circ$) in the HS and LS states. Σ is the sum of the deviation from 90° of the 12 *cis*-angles of the FeN_6 octahedron¹⁴

<i>T</i> (K)	296 (HS)	100 (LS)
Fe1–N1	2.200 (4)	1.999 (4)
Fe1–N2	2.190 (4)	1.993 (4)
Fe1–N3	2.176 (4)	1.994 (4)
(Fe–N)	2.189 (4)	1.995 (4)
Σ ($^\circ$)	17.3	16.8

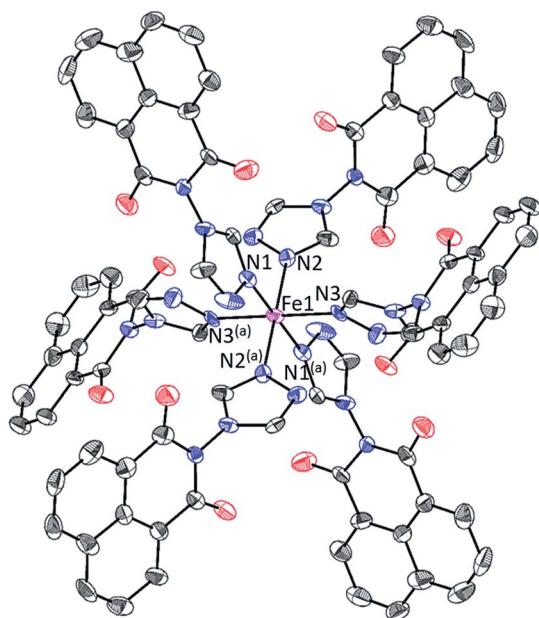


Fig. 2 Perspective ORTEP view of the molecular structure of the $[\text{Fe}(\text{naph-trz})_6]^{2+}$ complex. Hydrogen atoms are omitted for the sake of clarity. Symmetry transformation used to generate equivalent atoms: (a) $-x, -y, -z$.

Table S1†). Selected Fe–N bond lengths and N–Fe–N bond angles at both temperatures are gathered in Table S2.† Average Fe–N distances and distortion parameters for FeN_6 environment are listed in Table 1 for the two spin states.

The structure of C1 is built from one Fe(II) ion located on (0 0 0) special position, three naph-trz molecules acting as unidentate ligands, one $(\text{tcnsme})^-$ anion and two solvent CH_3CN molecules, all located on general positions. This asymmetric unit leads to the general formula $[\text{Fe}(\text{naph-trz})_6](\text{tcnsme})_2 \cdot 4\text{CH}_3\text{CN}$ (C1), in which the molecular structure consists of mononuclear complex of formula $[\text{Fe}(\text{naph-trz})_6]^{2+}$ (Fig. 2), with the two $(\text{tcnsme})^-$ anions acting as counterions to balance the charge of the complex. Examination of the Fe–N

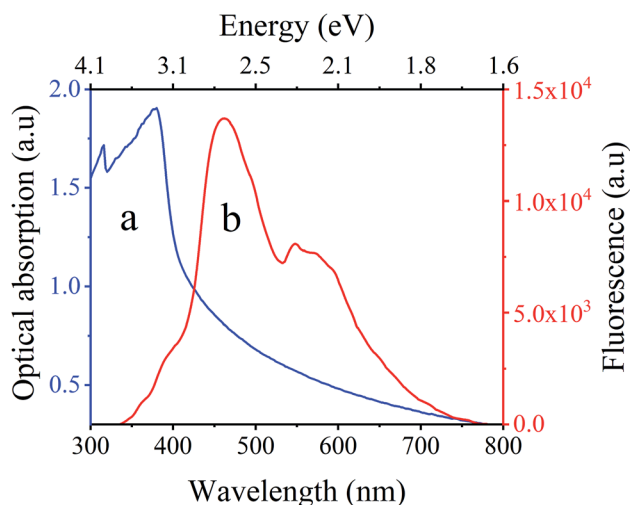


Fig. 3 Room temperature (a) optical absorption and (b) Fluorescence spectra of C1.

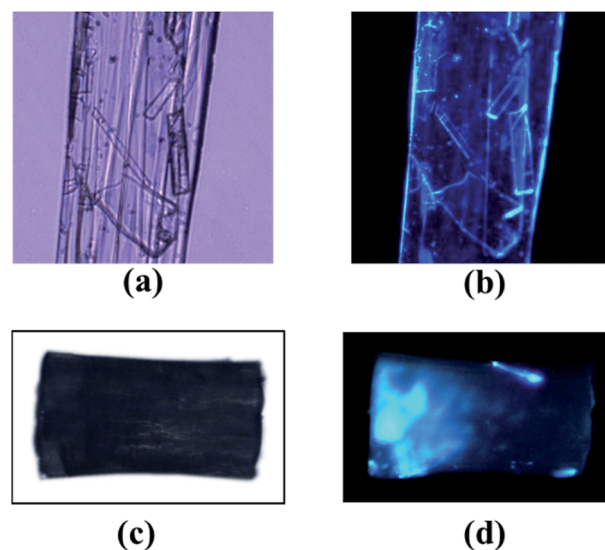


Fig. 4 Optical microscopy images at RT of (a) L1; (b) L1 under 320 nm-laser irradiation. (c) and (d) show the crystal C1 before and under irradiation with a 320 nm-laser, respectively. The fluorescence of L1 is clearly stronger than that of C1.



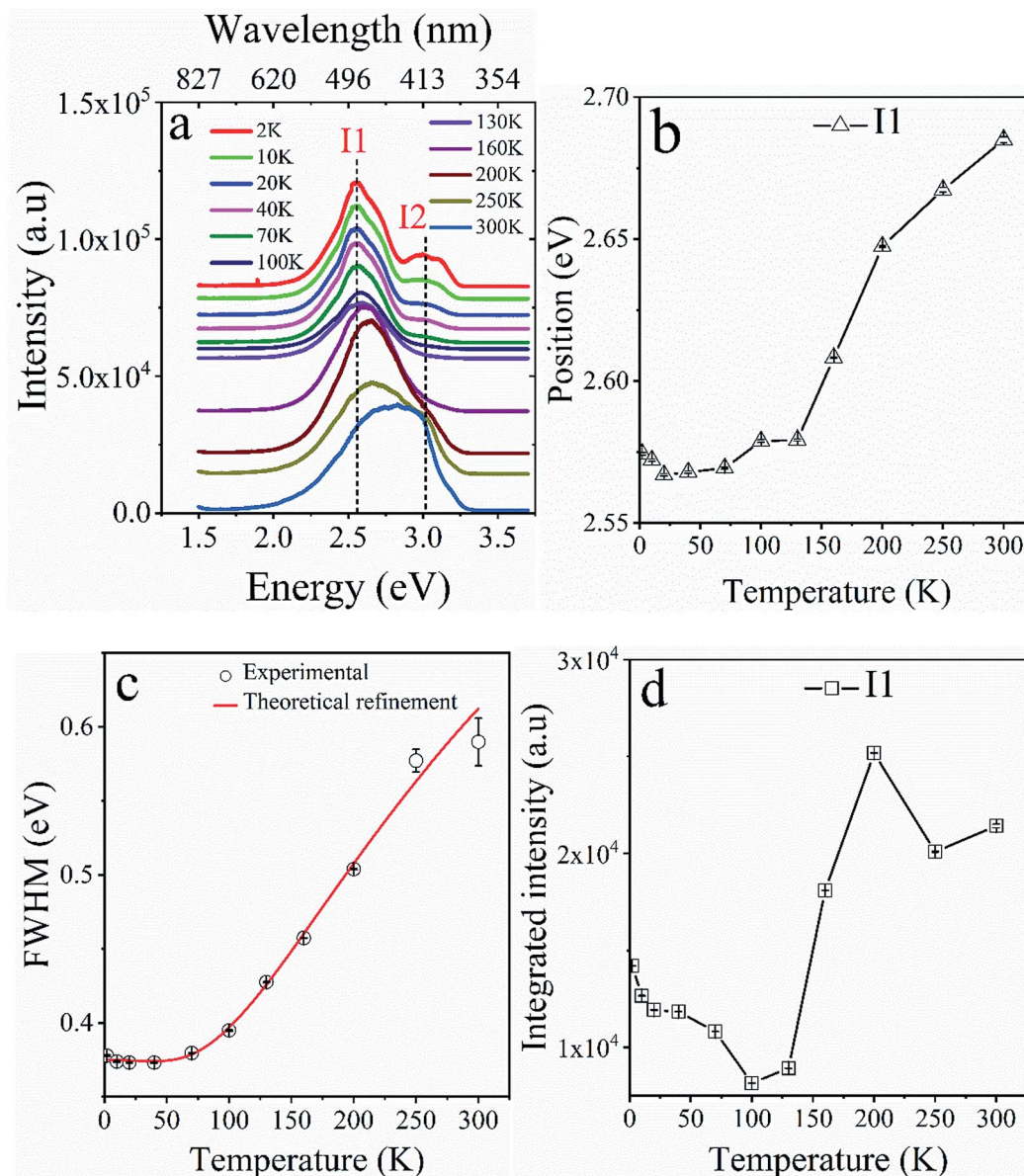


Fig. 5 (a) Temperature dependent of fluorescence properties of L1 over the range 2–300 K. Thermal evolution of the (b) position (in eV), (c) FWHM and (d) integrated intensity of the emission peak I1, showing significant changes around 150 and 200 K.

bonds and the N–Fe–N bond angles, for the Fe^{II} environments (Fe1), reveals that the central metal ion adopts a regular FeN₆ octahedral geometry which is defined by six nitrogen atoms arising from six naph-trz molecules. The average values of the Fe–N distances at 296 K (2.189(4) Å) and at 100 K (1.995(4) Å) are in agreement with the presence of full HS and LS states, respectively. This observation is consistent with the presence of a HS ↔ LS SCO transition, as revealed by the magnetic data (Fig. 1). The octahedral geometry of the metal center (Fe1) is relatively regular in the HS state, and therefore does not regularize further during the spin transition as revealed by the distortion parameter Σ (Table 1) that is relatively small and remains almost unchanged between RT and 100 K. This peculiar behavior, typical of constrained rigid systems,¹⁵ is probably induced by the presence of six sterically hindered heterocyclic

bulky naph-trz ligands in the coordination sphere of the cationic complex. Examination of the crystal packing in C1 reveals that the shortest intermolecular distances correspond to the π – π (C...C = 3.50–3.52 Å) and to the O...C (3.20–3.30 Å) contacts between the planar naph-trz heterocyclic ligands of adjacent complexes (see Fig. S14†). To understand how the unit cell parameters are affected by the SCO behavior, we have studied the temperature dependence of the lattice parameters of a single crystal of C1 in the temperature range 90–296 K. As expected, the unit cell parameters (a , b , c) and volume (V) as well as the β angle show significant changes around the transition temperature (see Fig. S15†). Furthermore, the analysis of the thermal variation of the three metric parameters (a , b , c) reveals, anisotropic changes around the transition temperature since, upon cooling from HS to the LS, the ac plane shows significant

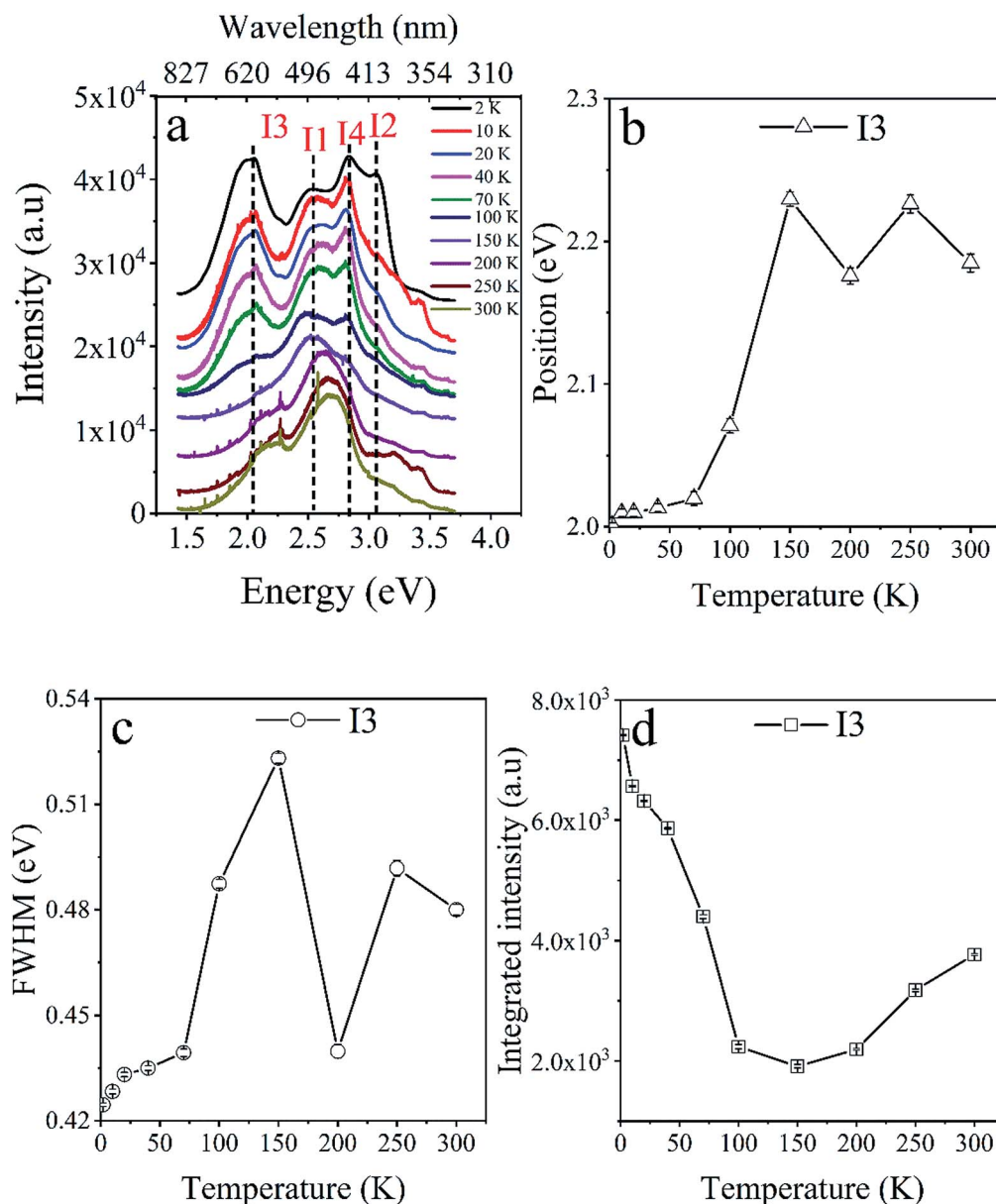


Fig. 6 (a) Fluorescence spectra of C1 over the temperature range 2–300 K showing the two extra components found (peaks I3 and I4) in comparison with the emission spectra of L1 shown in Fig. 5. Thermal dependence of (b) the position (in eV), (c) the FWHM and (d) the integrated intensity of I3 emission peak.

contraction along the two crystallographic directions, while the b parameter increases (see Fig. S15[†]), a behavior that is frequently encountered in SCO solids.¹⁶ The impact of this overall change is an increase of unit cell volume by 3.5% upon LS to HS transition, in excellent agreement with usual crystallographic data of SCO materials.¹⁶

The apparent discrepancy between the transition temperatures obtained by magnetic and crystallographic measurements originates from the fact that magnetic data result from the average thermal response of a large number of SCO particles having different transition temperatures while the structural data give the thermal transition of one single crystal. To compare the crystallographic and magnetic data, we derived the

corresponding HS fractions and plotted their variations as a function of the dimensionless parameter, $T/T_{1/2}$, where $T_{1/2}$ is the associated transition temperature (Fig. S16[†]). While the procedure is simple for the magnetic curve, for the lattice parameters, a , b , c , it requires subtracting their natural thermal expansion which interferes with the spin transition, before to plot their relative variations. A good agreement is obtained between all these data, although the structural transition appears sharper with respect to the magnetic one. Similar conclusions are also found when comparing magnetic data with optical microscopy studies, performed on various SCO single crystals, which demonstrated that the transition temperature crucially depends on the crystal size and shape.¹⁷

Optical investigations of **C1** were performed using optical absorption, fluorescence measurements (see ESI†). Fig. 3a shows the RT optical absorption spectrum of **C1** with a band centered at 380 nm. Irradiation of **C1** with a UV laser source at 325 nm at RT leads to the spectrum shown in Fig. 3, which has two bands peaking at 561 nm and 463 nm clearly confirming that **C1** is fluorescent at RT. Moreover, optical microscopy images, recorded at RT, allow observation of the bluish fluorescent color of **C1** as well as that of **L1** at RT (Fig. 4).

Fluorescence properties of **L1** were then investigated over the temperature range 2–300 K. The emission spectra of **L1** as a function of temperature (Fig. 5a and S17†) show a complex structure with two well separated components centered at 486 nm (I1) and 412 nm (I2). The behavior of the spectra of **L1** indicates that there are significant changes when **L1** is heated from helium to ambient temperature. These changes manifest through a monotonous increase of the position and the full width at half maximum (FWHM) of I1 peak while its integrated intensity shows a singularity around 150 K (Fig. 5b and c).

Moreover, important changes were also observed in the thermal behavior of the position, FWHM and the integrated intensity of I2 peak shown in Fig. S18.† Besides that, it is well known that the thermal dependence of the fluorescence line broadening may bring important information on the vibration/rotations of the ligand. Hence, the FWHM of peak I1, is derived using a Lorentzian line shape fitting procedure, based on the following law,

$$\Gamma(T) = \Gamma_0 + \Gamma_{AC} \times T + \Gamma_{LO} \exp\left(-\frac{E_{LO}}{k_B T}\right), \quad (1)$$

in which Γ_0 is the natural line width at 0 K, the second term represents the broadening induced by acoustic phonons, and the third term accounts for the inhomogeneous broadening of the emission line width caused by optical vibrations. There, Γ_{LO} is the exciton-phonon coupling constant and E_{LO} is the optical phonon energy. The best fit gives, $\Gamma_0 = 375 \pm 4$ meV, $\Gamma_{AC} = 1.2 \pm 0.1$ meV K⁻¹, $\Gamma_{LO} = 56 \pm 13$ meV and $E_{LO} = 28 \pm 3$ meV. At low temperature, a strong exciton-phonon coupling takes place, governed by the relatively high exciton-phonon interaction constant Γ_{LO} of 56 meV. In contrast, beyond 120 K, considered as the high temperature region, the broadening of the luminescence peak is mainly caused by acoustic phonons¹⁸ and shows a typical linear increase with a slope 1.2 meV K⁻¹, whose value is in good agreement with usual findings in hybrid perovskites, for example.¹⁸

The thermal response of the fluorescence spectra of **C1** was investigated on single crystals in the same temperature region as for **L1**. The obtained results are summarized in Fig. 6 and S19–23.† A large part of the emission spectra of **C1** is reminiscent of **L1**, where the main optical bands I1 and I2 are also observed, although they are slightly blue-shifted. Additionally, two extra contributions, not present in the fluorescence spectra of **L1**, emerge at 600 nm (~2 eV) and 445 nm (2.78 eV) noted as peaks I3 and I4, respectively. These emission bands are assigned to the fluorescence of the ligand when coordinated to the Fe(II) ion in the SCO metal complex. This behavior confirms

the contribution of the Fe(II) ions to the presence of the extra band I3 in the emission spectra of **C1**.

To further compare the behaviors of **L1** and **C1**, we have plotted in Fig. 7 the thermal dependence of the position of peak I1 that is present in both compounds. A significant difference is noted on the behavior of the two peaks. Indeed, while peak I1 displays a monotonous blue-shift for sample **L1**, it exhibits a more complex temperature dependence in material **C1**, with the presence of a regime change around 100 K, corresponding to the SCO transition region. Above 100 K **L1** and **C1** show the same behavior (Fig. 7).

The inspection of the thermal behaviour of the overall fluorescence of **C1** (see Fig. S23†) shows that its maximum intensity is almost two times smaller than that of the emission of the organic ligand **L1**, at RT, but it still remains strong enough to be seen with the naked eye, (as displayed in the optical microscopy images shown in Fig. 4).

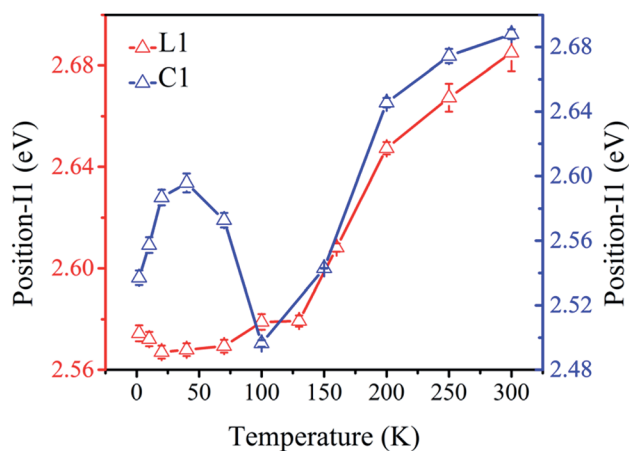


Fig. 7 Thermal-dependence of the positions (in eV) of peaks I1 in **L1** (red) and **C1** (blue) showing a significant difference in the thermal behavior of the emission at low-temperature region, i.e. in the LS region.

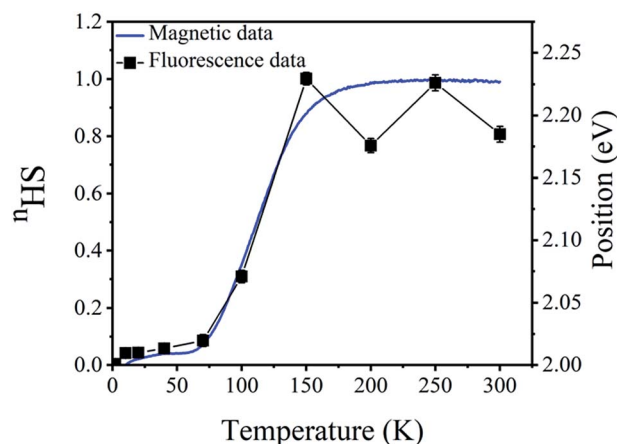


Fig. 8 Temperature dependence of peak I3 position in the emission spectra of the iron compound (black line with square symbol) in correlation with n_{HS} (blue line) derived from its magnetic properties.



In order to study the global consistency of the experimental data of thermal dependence of the fluorescence properties of **C1**, we have correlated the results of the thermal-dependence of the HS fraction derived from magnetic measurements and those of the thermal evolution of the position of the extra peak I3 around 2 eV (~ 600 nm) arising from the fluorescence response. The data, summarized in Fig. 8, show an excellent correlation, with the presence of a significant change in the temperature region 70–180 K where the thermal spin transition occurs in the magnetic response (Fig. 1a). These results clearly prove the existence of a direct synergy between the fluorescence properties of the complex and the spin transition of the central metal ion.

The spin transition causes an important change in the electronic density which is accompanied by a change of the Fe–ligand distances (structural reorganization), which undoubtedly affect the emission properties of the ligand. It is worth mentioning that at the present stage of the study it is not possible to bring more details about the coupling mechanism between the ligand and the SCO metal ion. We believe that DFT (Density Functional Theory) calculations, that we plan to study in a near future, will help to identify the intimate nature of this interaction. To further confirm the key role of the spin transition of the Fe(II) center, on the emission features of the complex **C1**, we have synthesized the homologous Cu-containing compound $[\text{Cu}(\text{naphtrz})_6](\text{tensme})_2 \cdot 4\text{CH}_3\text{CN}$ (**C2**). As expected, this sample does not show any spin transition. Fluorescence experiments performed on **C2** at various temperatures (Fig. 9, S24 and 25[†]) clearly show that the characteristic optical band I3 is absent in **C2** while the other bands look similar to I1 of **L1** and I2 (of **L1**) or I4 (of **C1**), although slightly shifted and broadened, due to the effect of the d-electrons configuration of Cu. This behavior confirms that the luminescence properties of

L1 are modified by the SCO behavior and therefore, constitutes a valuable and local probe of the spin transition phenomenon.

Conclusions

In summary, we have reported an iron(II) spin-crossover solid combining a spin transition and a strong fluorescence even at ambient temperature. Due to covalent character of the bonding between the six fluorescent ligands and the SCO metal center, the light emission spectra are significantly affected by the spin-transition of the metal. As a result, and at least in this compound, the SCO transition can be revealed through the fluorescence spectroscopy. This property is of high importance, in particular, in the study of SCO nanoparticles for which the magnetic signal is too low to be studied in single objects, while the emission spectra can be still detected for small objects. Moreover, the thermal dependence of the intensity, the width and even more, the position of the luminescence peaks constitute an excellent local probe of temperature and mechanical constraints, particularly when reducing the size of the SCO particles. On the other hand, the close correlation observed in the present **C1** complex between the gradual SCO and the luminescence behaviors, proves that the strong cooperativity is not needed for obtaining such synergetic multifunctional materials. Once again, this finding should be very useful for the design of fluorescent SCO thin films or nanoparticles, in which lower cooperativity is usually expected, for appropriate implementation in nanosized devices for enhanced sensing properties at the nanoscale. Finally, investigations of the photoinduced effects on the optical properties of the present SCO material are in progress as well as DFT calculations to get more insights about the quantum nature of the interrelationship of these two phenomena. Furthermore, in terms of innovative perspective to reach practical applications for the nano-switching devices, we are now designing systems based on new modified ligands involving similar triazole motif and fluorophore groups, separated by a flexible linker. Such ligands are able to favor the bridging coordination mode of the triazole motif at the origin of the formation 1D polymers,^{16c} exhibiting strong cooperative effects, and from which we wish to reach the bistability of magnetic and fluorescence properties, around ambient temperature.

Conflicts of interest

The authors declare no conflict of interest.

Acknowledgements

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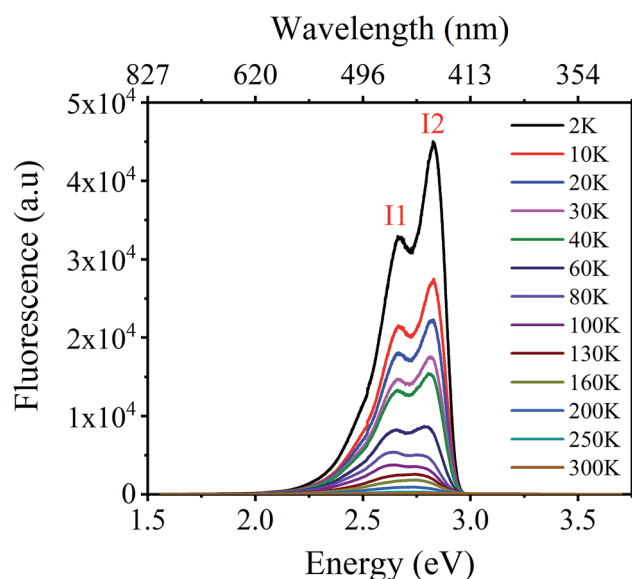


Fig. 9 Thermal-dependence of the fluorescence spectra of **C2** showing only two peak contributions as for the ligand **L1**.



Notes and references

- 1 (a) S. Basak, P. Hui and R. Chandrasekar, *Chem. Mater.*, 2013, **25**, 3408–3413; (b) E. Coronado, J. R. Galán-Mascarós, M. Monrabal-Capilla, J. García-Martínez and P. Pardo-Ibáñez, *Adv. Mater.*, 2007, **19**, 1359–1361; (c) M. Cavallini, *Phys. Chem. Chem. Phys.*, 2012, **14**, 11867–11876; (d) H. J. Shepherd, G. Molnár, W. Nicolazzi, L. Salmon and A. Bousseksou, *Eur. J. Inorg. Chem.*, 2013, **2013**, 653–661; (e) A. D. Naik, L. Stappers, J. Snauwaert, J. Fransaer and Y. Garcia, *Small*, 2010, **6**, 2842–2846; (f) M. Cavallini, I. Bergenti, S. Milita, J. C. Kengne, D. Gentili, G. Ruani, I. Salitros, V. Meded and M. Ruben, *Langmuir*, 2011, **27**, 4076–4081.
- 2 (a) A. Bousseksou, N. Negre, M. Goiran, L. Salmon, J. P. Tuchagues, M. L. Boillot, K. Boukheddaden and F. Varret, *Eur. Phys. J. B*, 2000, **13**, 451–456; (b) M. A. Halcrow, *Spin-Crossover Materials: Properties and Applications*, John Wiley & Sons Ltd, 2013; (c) J. Linares, E. Codjovi and Y. Garcia, *Sensors*, 2012, **12**, 4479–4492; (d) A. Bousseksou, G. Molnar, L. Salmon and W. Nicolazzi, *Chem. Soc. Rev.*, 2011, **40**, 3313–3335; (e) P. Gutlich, A. B. Gaspar and Y. Garcia, *Beilstein J. Org. Chem.*, 2013, **9**, 342–391.
- 3 (a) P. Lloveras, E. Stern-Taulats, M. Barrio, J. L. Tamarit, S. Crossley, W. Li, V. Pomjakushin, A. Planes, L. Mañosa, N. D. Mathur and X. Moya, *Nat. Commun.*, 2015, **6**, 8801; (b) P. J. v. Ranke, *Appl. Phys. Lett.*, 2017, **110**, 181909.
- 4 X. Zhang, T. Palamarciuc, J.-F. Létard, P. Rosa, E. V. Lozada, F. Torres, L. G. Rosa, B. Doudin and P. A. Dowben, *Chem. Commun.*, 2014, **50**, 2255–2257.
- 5 (a) M. Hiroyuki, K. Keita and K. Nobuo, *Chem. Lett.*, 2008, **37**, 446–447; (b) C. Piguet, E. Rivara-Minten, G. Hopfgartner and J.-C. G. Bünzli, *Helv. Chim. Acta*, 1995, **78**, 1541–1566; (c) C. Piguet, E. Rivara-Minten, G. Bernardinelli, J.-C. G. Bünzli and G. Hopfgartner, *J. Chem. Soc., Dalton Trans.*, 1997, 421–434; (d) C. Edler, C. Piguet, J. C. G. Bünzli and G. Hopfgartner, *Chem. - Eur. J.*, 2001, **7**, 3014–3024.
- 6 (a) Y. Garcia, F. Robert, A. D. Naik, G. Zhou, B. Tinant, K. Robeyns, S. Michotte and L. Piroux, *J. Am. Chem. Soc.*, 2011, **133**, 15850–15853; (b) C.-F. Wang, R.-F. Li, X.-Y. Chen, R.-J. Wei, L.-S. Zheng and J. Tao, *Angew. Chem., Int. Ed.*, 2015, **54**, 1574–1577; (c) C. F. Wang, M. J. Sun, Q. J. Guo, Z. X. Cao, L. S. Zheng and J. Tao, *Chem. Commun.*, 2016, **52**, 14322–14325; (d) M. Hasegawa, F. Renz, T. Hara, Y. Kichuki, Y. Fukada, J. Okubo, T. Hoshi and W. Linert, *Chem. Phys.*, 2002, **277**, 21–30; (e) R. González-Prieto, B. Fleury, F. Schramm, G. Zoppellaro, R. Chandrasekar, O. Fuhr, S. Lebedkin, M. Kappes and M. Ruben, *Dalton Trans.*, 2011, **40**, 7564–7570; (f) J.-L. Wang, Q. Liu, X.-J. Lv, R.-L. Wang, C.-Y. Duan and T. Liu, *Dalton Trans.*, 2016, **45**, 18552–18558; (g) B. Schäfer, T. Bauer, I. Faus, J. A. Wolny, F. Dahms, O. Fuhr, S. Lebedkin, H.-C. Wille, K. Schlage, K. Chevalier, F. Rupp, R. Diller, V. Schünemann, M. M. Kappes and M. Ruben, *Dalton Trans.*, 2017, **46**, 2289–2302; (h) H. J. Shepherd, C. M. Quintero, G. Molnár, L. Salmon and A. Bousseksou, in *Spin-Crossover Materials: Properties and Application*, pp. 347–373, ed. M. A. Halcrow, John Wiley & Sons Ltd, 2013.
- 7 (a) M. Matsuda, H. Isozaki and H. Tajima, *Thin Solid Films*, 2008, **517**, 1465–1467; (b) K. Senthil Kumar and M. Ruben, *Coord. Chem. Rev.*, 2017, **346**, 176–205; (c) M. Masaki, I. Hikaru and T. Hiroyuki, *Chem. Lett.*, 2008, **37**, 374–375.
- 8 (a) J.-L. Wang, Q. Liu, Y.-S. Meng, X. Liu, H. Zheng, Q. Shi, C.-Y. Duan and T. Liu, *Chem. Sci.*, 2018, **9**, 2892–2897; (b) C.-F. Wang, G.-Y. Yang, Z.-S. Yao and J. Tao, *Chem. - Eur. J.*, 2018, **24**, 3218–3224.
- 9 (a) L. Salmon, G. Molnár, D. Zitouni, C. Quintero, C. Bergaud, J.-C. Micheau and A. Bousseksou, *J. Mater. Chem.*, 2010, **20**, 5499–5503; (b) S. Titos-Padilla, J. M. Herrera, X. W. Chen, J. J. Delgado and E. Colacio, *Angew. Chem., Int. Ed.*, 2011, **50**, 3290–3293.
- 10 C. A. Tovee, C. A. Kilner, J. A. Thomas and M. A. Halcrow, *CrystEngComm*, 2009, **11**, 2069–2077.
- 11 (a) C. Lochenie, K. Schötz, F. Panzer, H. Kurz, B. Maier, F. Puchtler, S. Agarwal, A. Köhler and B. Weber, *J. Am. Chem. Soc.*, 2018, **140**, 700–709; (b) J. Yuan, S.-Q. Wu, M.-J. Liu, O. Sato and H.-Z. Kou, *J. Am. Chem. Soc.*, 2018, **140**, 9426–9433.
- 12 N. Zhang, A. B. Carter, A. J. Fitzpatrick and G. G. Morgan, *J. Coord. Chem.*, 2016, **69**, 2024–2037.
- 13 S. Lethu and J. Dubois, *Eur. J. Org. Chem.*, 2011, 3920–3931.
- 14 M. Marchivie, P. Guionneau, J. F. Létard and D. Chasseau, *Acta Crystallogr., Sect. B: Struct. Sci.*, 2005, **61**, 25–28.
- 15 E. Milin, V. Patinec, S. Triki, E.-E. Bendeif, S. Pillet, M. Marchivie, G. Chastanet and K. Boukheddaden, *Inorg. Chem.*, 2016, **55**, 11652–11661.
- 16 (a) S. pillet, J. Hubsch and C. Lecomte, *Eur. Phys. J. B*, 2004, **38**, 541–552; (b) M. Sy, F. Varret, K. Boukheddaden, G. Bouchez, J. Marrot and S. Kaizaki, *Angew. Chem., Int. Ed.*, 2014, **53**, 7539–7542; (c) N. Pittala, F. Thétiot, S. Triki, K. Boukheddaden, G. Chastanet and M. Marchivie, *Chem. Mater.*, 2017, **29**, 490–494.
- 17 (a) H. Fourati, G. Bouchez, M. Paez-Espejo, S. Triki and K. Boukheddaden, *Crystals*, 2019, **9**, 46; (b) H. Fourati, E. Milin, A. Slimani, G. Chastanet, Y. Abid, S. Triki and K. Boukheddaden, *Phys. Chem. Chem. Phys.*, 2018, **20**, 10142–10154.
- 18 (a) A. Yangui, S. Pillet, A. Mlayah, A. Lusson, G. Bouchez, S. Triki, Y. Abid and K. Boukheddaden, *J. Chem. Phys.*, 2015, **143**, 224201; (b) A. Yangui, S. Pillet, E.-E. Bendeif, A. Lusson, S. Triki, Y. Abid and K. Boukheddaden, *ACS Photonics*, 2018, **5**, 1599–1611; (c) A. Yangui, D. Garrot, J. S. Lauret, A. Lusson, G. Bouchez, E. Deleporte, S. Pillet, E. E. Bendeif, M. Castro, S. Triki, Y. Abid and K. Boukheddaden, *J. Phys. Chem. C*, 2015, **119**, 23638–23647; (d) M. Koubaa, T. Dammak, D. Garrot, M. Castro, E. Codjovi, A. Mlayah, Y. Abid and K. Boukheddaden, *J. Appl. Phys.*, 2012, **111**, 053521.

