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Reactive Coordinatively Saturated Mo(III) Complex: Exploiting the Adaptive Hapticity of Tris(*tert*-butoxy)silanolate Ligands.

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Coordinatively unsaturated Mo(III) complexes have been identified as highly reactive species able to activate dinitrogen without the need of a sacrificial reducing agent. Here, we report a coordinatively saturated octahedral Mo(III) complex stabilized by κ^2 -tris(*tert*-butoxy)silanolate ligands, which is yet highly reactive towards dinitrogen and small molecules. The combined high stability and activity is ascribed to the dual binding mode of the tris(*tert*-butoxy)silanolate ligands that allows unlocking a coordination site in presence of reactive small molecules to promote their activation at low temperatures.

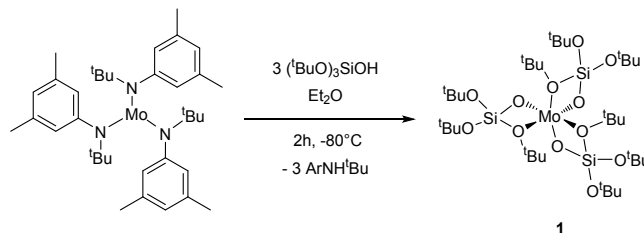
Introduction

The presence of a Mo center in the active site of the FeMo cofactor of nitrogenase enzymes has stimulated numerous studies on the synthesis of low valent Mo complexes and their reactivity towards small molecules.^[1] Tri-coordinated complexes with bulky amide ligands like Mo[N(R)Ar]₃ (R = *t*Bu, Ar = 3,5-Me₂C₆H₃) have shown to be particularly reactive and able to cleave dinitrogen.^[2] Such reactivity contrasts this observed for many mononuclear coordinatively saturated octahedral Mo(III) complexes, which are quite inert towards small molecules. To date, there are still a small number of examples^[3] of tri-coordinated unsaturated Mo(III) compounds due to their propensity to form rather unreactive triply-bonded Mo≡Mo dimers.^[4] The use of bulky chelating ligands has thus been key to obtain isolable, yet reactive coordinatively unsaturated Mo(III) centers.^[3c, 5] This approach has enabled the synthesis of the first mononuclear Mo catalytic systems able to catalytically reduce N₂ into ammonia.^[6] The tris(*tert*-butoxy)silanolate (TBOS) ligand is not very sterically demanding but presents a variety of binding modes that allows for the formation of mono-nuclear complexes with a broad range of metals through coordination as a terminal κ^1 , terminal κ^2 , bridging $\kappa^1\mu^2$ and bridging $\kappa^2\mu^2$ ligand.^[7] Terminal κ^1 binding modes are observed for small metal ions such as trivalent Cr or Ga, while terminal κ^2 modes are found for larger metal ions such as trivalent Sc, Yb, Lu^[8] or Dy.^[9] Mo(III) possesses an intermediate ionic radius between these two groups, we

thus reasoned that the TBOS ligands had the potential to stabilize Mo(III), while maintaining its reactivity towards small molecules, as such ligand can adopt different coordination modes and interconvert between them. Here, we present the synthesis and the reactivity of Mo(OSi(*t*Bu)₃)₃, **1**, which displays a distorted octahedral geometry with three κ^2 -TBOS ligands. Coordination of the TBOS ligand in a bidentate κ^2 -fashion prevents dimerization and allows isolating a mononuclear Mo(III) complex, while the flexibility of its binding mode provides high reactivity as testified by the activation at low temperature of small molecules like N₂, N₂O and CO₂, among others.

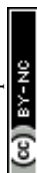
Results and discussion

Mo(OSi(*t*Bu)₃)₃, **1**, is prepared by reaction of [Mo(N(R)Ar)₃] with (tBuO)₃SiO-H (TBOS-H) in a slurry of diethyl ether at -80°C; it is isolated as dark brown crystals in 36% yield after evacuation of the solvent and aniline byproduct under high vacuum at low temperature, followed by crystallization from pentane. NMR studies indicate a quantitative reaction, but the tedious isolation of this compound, which is required to separate the highly soluble compound **1** from the residual amine, diminishes significantly the overall isolated yield for the pure crystalline form. The compound **1** is however stable for months when kept in the solid state under argon atmosphere at -35 °C and is rather stable in apolar solvents such as benzene, allowing investigating its reactivity with small molecules in solution.



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Scheme 1. Synthesis of [Mo(TBOS)₃] (1).

Single-crystal X-ray diffraction^[10] shows that the complex **1** adopts a distorted octahedral geometry with three κ^2 -TBOS ligands, where the Mo-O distance of the silanolate is ca. 0.3 Å shorter than the tert-butoxy one (2.020(4) Å vs. 2.302(4) Å). The ¹H NMR spectrum of **1** shows a single very broad resonance at 3.82 ppm ($\nu_{1/2}$ = 164 Hz, RT), suggesting a fluxional coordination of the tert-butoxy moieties in solution. At -80°C, the ¹H NMR spectrum displays three peaks at 22.6, 4.9 and -1.36 ppm, in agreement with the three inequivalent tert-butoxy groups (See ESI for details). However, the complex **1** still displays broad peaks at this temperature, suggesting fast dynamics of the methyl groups. The magnetic susceptibility (μ_{eff} = 3.74 μ_{B}) measured by Evans method at 298K is close to the expected value for a d³ complex with three unpaired electrons (3.75 μ_{B}).

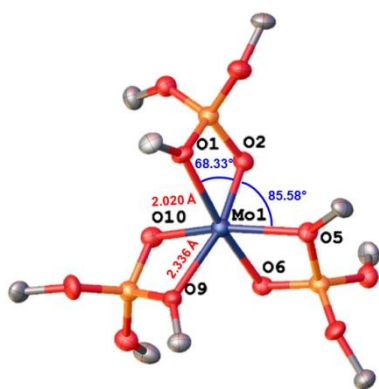
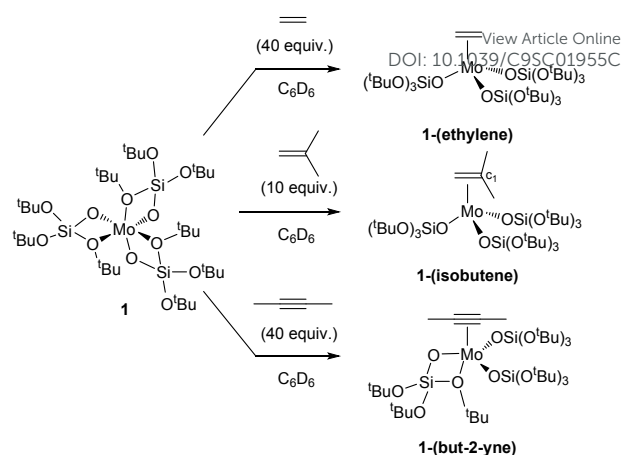


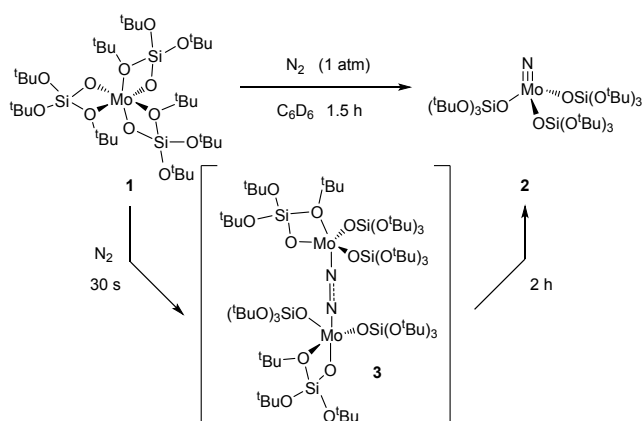
Figure 1. Thermal ellipsoid plot at the 50% probability of [Mo(TBOS)₃] (**1**). Hydrogen atoms and siloxide methyl groups have been omitted for clarity.

We first explored the reaction of complex **1** with alkenes (ethylene, isobutene) as well as alkynes (2-butyne). Condensing the gas at low temperature (-196°C) in the solution and then warming to room temperature, the corresponding π -complexes were obtained in quantitative yields (Scheme 2, see ESI for details). The alkene and alkyne adducts [Mo(C₂H₄)(OSi(OtBu)₃)₃], **1-(ethylene)**, [Mo(C₄H₈)(OSi(OtBu)₃)₃], **1-(isobutene)**, and [Mo(C₄H₆)(OSi(OtBu)₃)₃], crystallize in a tetrahedral geometry, while **1-(but-2-yne)**,^[11] presents a square-based pyramidal one. In the latter, the π -accepting ligand occupies the apical site and three monodentate silanolate ligands are in the basal plane (XRD structures are presented in Fig. S11-13). The significant lengthening of the C-C bond upon coordination (1.406(7), 1.391(9) and 1.283(5) Å for **1-(ethylene)**, **1-(isobutene)** and **1-(but-2-yne)**, respectively vs. 1.337(5)^[12], 1.331(3)^[13] and 1.202(5)^[11] Å for the corresponding organic molecules) indicates a strong back-donation of the Mo(III) center to these π -acid ligands.^{[14][15]}



Scheme 2. Reaction of **1** with unsaturated hydrocarbons. Selected bond distances: Mo-C1 = 2.135(4); 2.145(12) and 2.007(4) Å for **1-(ethylene)**, **1-(isobutene)** and **1-(but-2-yne)**.

The fluxionality of the ligand prompted us to further investigate the reactivity of **1** towards small molecules, namely N₂, N₂O and CO₂. Exposing **1** to a N₂ atmosphere at room temperature led to an immediate color change from amber to dark blue within seconds. This blue color fades away after 2h at room temperature, leaving a colorless solution.



Scheme 3. Reaction of **1** with N₂ yields [Mo(N)(TBOS)₃] (**2**) via (μ -N₂)[Mo(TBOS)₃]₂ intermediate (**3**).

Removal of all volatiles followed by extraction in pentane and crystallization at -35°C affords the Mo(VI) nitride complex [Mo(N)(TBOS)₃], **2**, in 40% yield. The relatively low isolated yield reflects the high solubility of the material, despite the quantitative reaction yield according to NMR. Single crystal X-ray analysis of **2** (Figure 2a) shows a tetrahedral terminal nitride complex of Mo(VI) with three κ_1 -TBOS ligands (Scheme 3); the limited quality of the collected data however precludes an accurate discussion of bond distances. Presence of a nitride ligand was confirmed by FTIR, as evidenced by a band at 1011 cm⁻¹ associated to $\nu(\text{Mo}\equiv\text{N})$.^[16] To verify unequivocally the signal, we carried out the reaction with ¹⁵N₂. The Infrared spectrum of ¹⁵N≡Mo(TBOS)₃ shows a band at 997 cm⁻¹ that is red shifted compared to its ¹⁴N analogue, consistent



with their attribution to Mo≡N vibration. ^1H NMR of **2** shows a single and sharp peak at 1.49 ppm, in agreement with a Mo(VI) diamagnetic species. The blue intermediate observed upon N_2 addition can be isolated by rapid cooling of the reaction mixture (-196°C) followed by a fast solvent-exchange at 5°C (from benzene- d_6 to pentane) and crystallization at -35°C . A dimeric Mo dinitrogen complex, $(\mu\text{-N}_2)[\text{Mo}(\text{TBOS})_2]_2$, **3**, is isolated as dark blue crystals. Single crystal X-ray analysis of **3** (Figure 2b) displays two Mo ions in a distorted TBP (Trigonal BiPyramidal) geometry with two κ^1 and one κ^2 TBOS ligands along with the end-on bound dinitrogen molecule sandwiched between the two Mo centers (Mo1-N1-N2 and N1-N2-Mo2 angle of $172.5(3)^\circ$ and $169.0(3)^\circ$, respectively). The N1-N2 distance is notably elongated with respect to the free N_2 molecule ($1.209(5) \text{ \AA}$ vs $1.0976(2)^{[11]} \text{ \AA}$) indicating a significant degree of activation. When **3** is warmed to room temperature the subsequent step of the overall $6 e^-$ reduction of N_2 occurs, leading to the formation of two equivalents of **2** (Scheme 3). While this mechanism is similar to the one observed with coordinatively unsaturated tris-amido complexes, $^{[17]}$ the complete process is much faster with **1** (2 h for the quantitative reaction at room temperature vs. less than 5% conversion after 12h at 28°C for $\text{Mo}(\text{NtBuAr})_3$). $^{[18]}$ Worth of note is that the corresponding complex bearing three very large silox ligands – $\text{Mo}(\text{OSi}t\text{Bu}_3)_3$ – does not activate dinitrogen. $^{[5c]}$ This high reactivity towards N_2 reduction is unusual for an octahedral Mo(III) complex, and highlights the particularity of the TBOS ligand, which can adopt mono or bidentate binding modes while being relatively small compared to larger siloxy (e.g. $t\text{Bu}_3\text{SiO}$) $^{[19]}$ or $\text{N}(\text{tBu})(3,5\text{-Me}_2\text{C}_6\text{H}_3)$ ligands. $^{[7g]}$

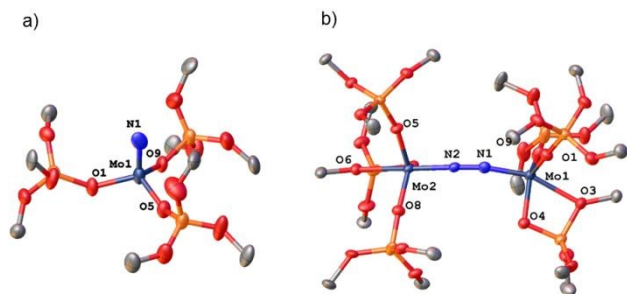
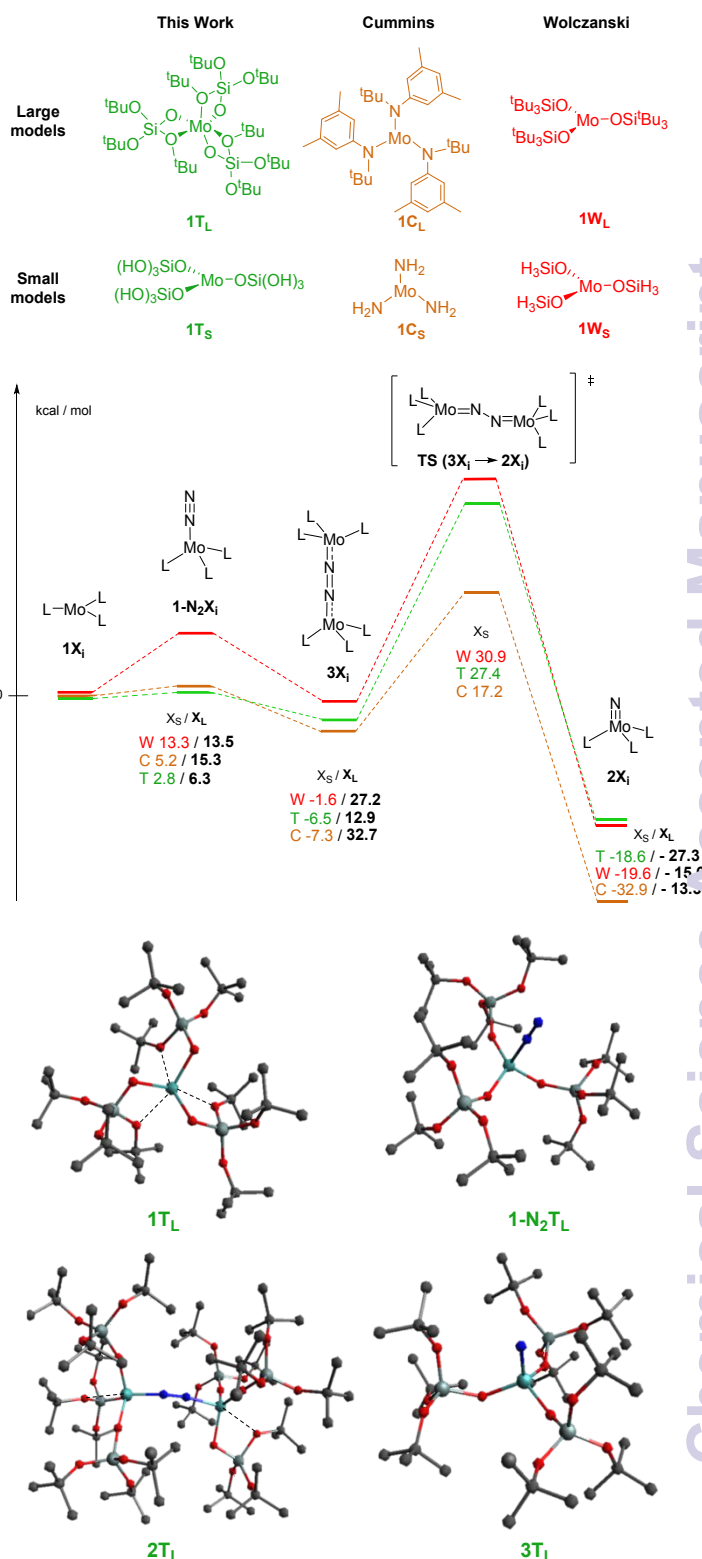


Figure 2. Thermal ellipsoid plot at the 50% probability of a) **2**, and b) **3**. Hydrogen atoms, siloxide methyl groups and disorder have been omitted for clarity.

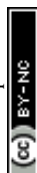
To corroborate our data, we performed DFT calculations in order to better understand the reactivity difference between these three systems. Due to the very large size of the investigated systems, we calculated the entire potential free energy surface with simplified models $^{[20]}$ and calculated all reaction intermediates with the full ligand set (Scheme 4). The DFT calculations show that if electronic effects were the dominant factor (small models), one would predict the tris-amido system to be the most effective for N_2 cleavage (TS of 17.2 kcal/mol above separated reactants), while the siloxy derivatives should have similar significantly lower reactivity (with TS energies of 27.4 and 30.9 kcal/mol , for **1** and the model silox derivatives) in contrast with the observed trend, i.e. **1** reacts much faster with N_2 than

the Cummins Mo(III), while the Wolczanski analog is unreactive towards N_2 (Scheme 4).

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Scheme 4 Top: Calculated ΔG in kcal/mol for the three compared systems 1T_L , 2T_L and 3T_L . Colored numbers refer to the small models and black numbers to the large ones. DFT models of complex **1** are in agreement with the XRD data, confronting the bond lengths and bond angles (calculated vs experimental Mo-O1 $2.410/2.302(4)$ Mo1-O6 $2.043/2.046(4)$, Si-O1 $1.707/1.677(5)$, O2-Mo-O1 $67.2/68.33(16)$, O2-Mo1-O5 $87.6/85.58(15)$), so are the other values for intermediate **c** and for the final nitride **e** (see ESI for details). Bottom:

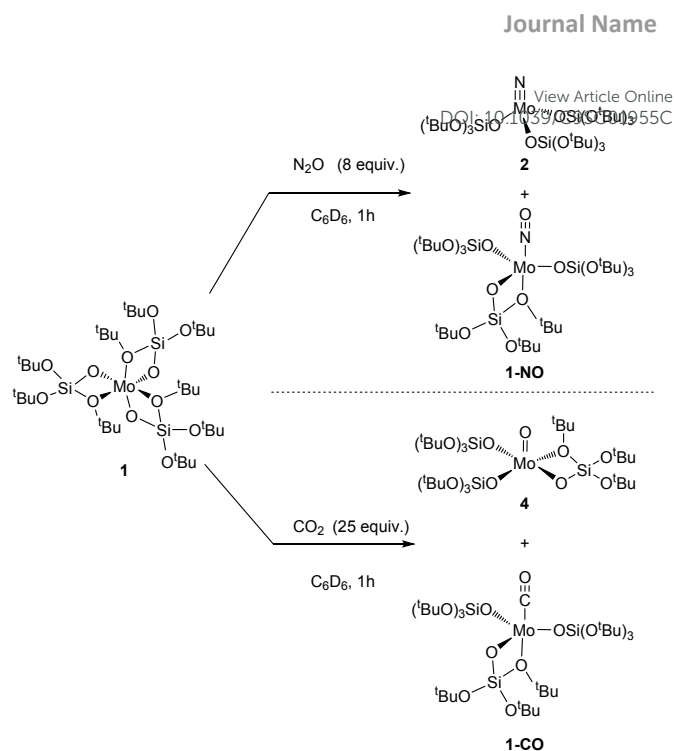


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Calculated structure for the large model of complex **1** for the reaction intermediates.

Once steric contributions are taken into consideration, it is notable how **1** yields more energetically accessible reactions intermediates and a significantly more stable nitrido complex: dinitrogen coordination costs only 6 kcal/mol for **1** vs 13-15 kcal/mol for the two others and the formation of the dimeric species prior to N₂ activation costs only 13 kcal/mol for **1** vs. 27 and 33 kcal/mol for the Cummins and Wolczanski systems, respectively, in line with their relative difference of reactivity. Overall, the DFT calculations show that the TBOS ligand favors the N₂ activation because of its hemilability while protecting Mo(III) thanks to the possible κ^2 -configuration.

We then further explored the reactivity of **1** towards other small molecules. Upon exposure to N₂O, a solution of **1** undergoes a fast color change from amber to bright yellow upon warming up from liquid nitrogen temperature to room temperature (Scheme 4). Analysis of the reaction mixture by ¹H NMR shows the presence of two peaks at 1.51 and 1.50 ppm. FTIR analysis of this reaction mixture after removal of the solvent *in vacuo* reveals the presence of an intense band at 1678 cm⁻¹ which is attributed to Mo nitrosyl.^[16, 21] Two species are obtained as crystals, and X-ray crystallography analyses revealed the species **2** and [Mo(NO)(TBOS)₃], **1-NO** (Scheme 5). The nitrosyl compound **1-NO** crystallizes in a distorted TBP geometry, with a linear NO ligand (Mo-N-O = 178.6(2)°) in apical position and a κ^2 -TBOS ligand. The Mo-NO bond distance of 1.715(3) Å is consistent with literature value for linear nitrosyl complexes of Cummins and Zubieta, two molybdenum thiolate species with a metal-nitrogen distance of 1.769(13) and 1.766(6) Å respectively.^[21a, 22] It is slightly shorter consistent with the weaker trans influence of the siloxy ligand. Despite a high kinetic barrier to reduction, the activation of N₂O has been observed with numerous low valent metal complexes.^[23] However, the majority of metal-mediated activation of N₂O generally involves the scission of N-O bond, as the N-N bond is significantly stronger (~481 kJ mol⁻¹ and ~167 kJ mol⁻¹, for N-N and N-O respectively) with few examples of N-N bond activation,^[6a, 21e, 24] highlighting the high azophilicity of the Mo complex **1**. The extremely challenging purification^[21a] of the reaction mixture did not allow any further characterization of **1-NO** and of the impurities which appear to be paramagnetic.



Scheme 5. Reaction of **1** with N₂O and CO₂.

Contacting **1** with CO₂ also results in a rapid color change from amber to red upon warming from liquid nitrogen temperature to room temperature, to yield a 1:1 ratio of two compounds according to NMR spectroscopy. Crystallization of the reaction mixture from pentane at -35°C affords two sets of red crystals, assigned to [Mo(O)(TBOS)₃], **4**, and [Mo(CO)(TBOS)₃], **1-CO** (Figure 3) according to X-ray single-crystal diffraction studies. The X-ray diffraction analysis of the mono-carbonyl complex **1-CO** shows that the CO-adduct adopts a distorted TBP geometry with a κ^2 -TBOS ligand. The CO ligand occupies an axial position and is almost linear with a Mo-C-O angle of 175.2(3)°. The Mo-C and C-O distances of 1.931(3) Å and 1.156(3) Å respectively are in between those of Mo(CO)₆ (Mo-C 2.059(3) Å and C-O 1.125(5) Å)^[25] and [Mo(CO)(Ph₂PCH₂CH₂PPh₂)₂] (Mo-C 1.90(1) Å and C-O 1.19(1) Å)^[26] indicating a π -back-donation on the carbonyl, comparable to that of *cis*-(diethylenetriamine)molybdenum tricarbonyl.^[27] This is consistent with the red shift of the ν (C=O) band at 1869 cm⁻¹ observed in the FTIR spectrum, which can be attributed to the weak σ -donor ability of TBOS ligand.^[28] Complex **1-CO** can as be synthesized upon exposing **1** to an excess of carbon monoxide, see ESI for details. The structure of **4** presents a Mo(V) complex in a distorted TBP geometry with an equatorial oxo moiety with no trans ligand. Mo=O bond length is 1.665(2) Å, which lies in the usual range for terminal Mo(V) oxo complexes.^[29]



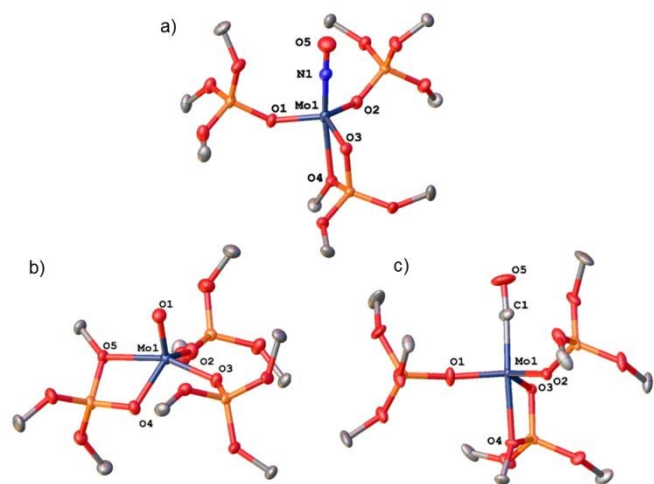


Figure 4. Thermal ellipsoid plot at the 50% probability of a) **1-NO**, b) **4** and c) **1-CO**. Hydrogen atoms, siloxide methyl groups and disorder have been omitted for clarity.

Conclusions

In conclusion, we have shown that the TBOS ligand can bind Mo(III) in a bidentate κ^2 -fashion, providing an isolable – yet highly reactive – mononuclear octahedral Mo(III) complex. This complex thus activates a large variety of small – often rather inert – molecules, cleaving N-N, C-O and N-O bonds at low temperatures. An example is the 6-electron reduction of N_2 , which occurs faster than with other tri-coordinated Mo(III) complexes. The high reactivity of this coordinatively saturated complex illustrates the high adaptability of the TBOS ligand, able to protect a highly reactive Mo(III) sites towards dimerization, without quenching its reactivity. The use of TBOS ligand thus constitutes an alternative strategy to employ highly coordinatively unsaturated complexes towards small molecule activation at Mo centers.

Conflicts of interest

There are no conflicts to declare.

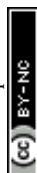
Acknowledgements

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Notes and references

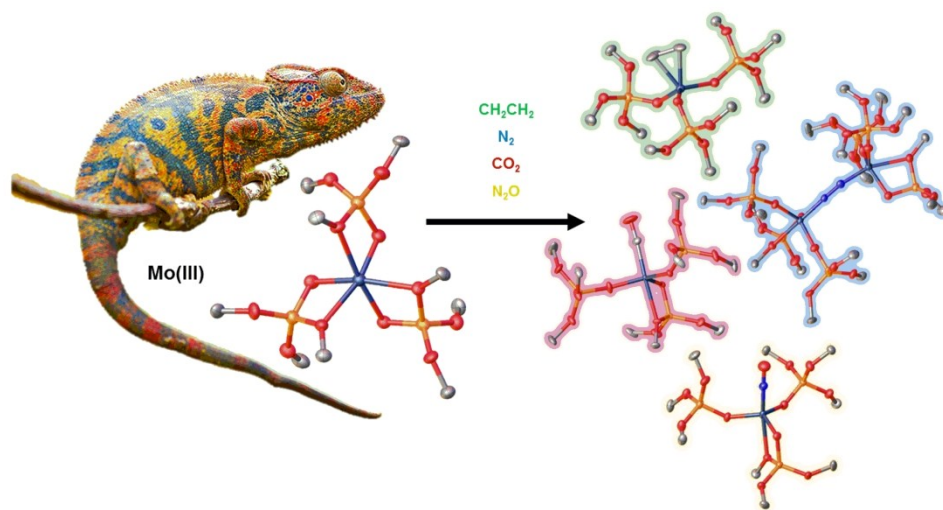
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- 1826963 (1-(but-2-yne)), CCDC 1826968 (1-NO), CCDC 1826969 (4), CCDC 1826965 (1-CO) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Crystallographic Data Centre.
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