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Structure of formylglycine-generating enzyme in complex with copper and a substrate reveals an acidic pocket for binding and activation of molecular oxygen†

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The formylglycine generating enzyme (FGE) catalyzes oxidative conversion of specific peptidyl-cysteine residues to formylglycine. FGE mediates O₂-activation and hydrogen-atom abstraction in an active site that contains Cu(I) coordinated to two cysteine residues. Similar coordination geometries are common among copper-sensing transcription factors and copper-chaperone but are unprecedented among copper-dependent oxidases. To examine the mechanism of this unusual catalyst we determined the 1.04 Å structure of FGE from *Thermomonospora curvata* in complex with copper and a cysteine-containing peptide substrate. This structure unveils a network of four crystallographic waters and two active site residues that form a highly acidic O₂-binding pocket juxtaposed to the trigonal planar tris-cysteine coordinated Cu(I) center. Comparison with structures of FGE in complex with Ag(I) and Cd(II) combined with evidence from NMR spectroscopy and kinetic observations highlight several structural changes that are induced by substrate binding and prime the enzyme for O₂-binding and subsequent activation.

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Introduction

Copper is a versatile catalyst for oxygen-dependent reactions. Combined with appropriate ligands, copper can cycle between the oxidation states I, II and III to activate molecular oxygen (O₂) and to form reactive oxygen species that can initiate very difficult reactions.^{1–3} Mechanistic appreciation of this reactivity is an important prerequisite to understand the fundamental role of copper in biology and to exploit its potential in chemical synthesis.⁴ There are two groups of copper-proteins that are most relevant for the subject of the present report: mononuclear copper enzymes that utilize O₂ as an oxidant, and copper-binding proteins that suppress unspecific and cytotoxic activities of copper in the cell.⁵

The group of known mononuclear copper enzymes includes lytic polysaccharide monooxygenases (LPMO),^{6,7} particulate methane monooxygenase (pMMO),⁸ peptidylglycine- α -hydroxylating monooxygenase (PHM),^{9,10} copper amine oxidase (CAO),¹¹ and dopamine- β -hydroxylase (DBH).¹² These enzymes bind copper in histidine-dominated tetrahedral or square planar coordination spheres. The structure of these enzymes have inspired the development of abiotic copper-ligands in

order to examine specific aspects of copper-catalysis, but also with an eye on application in organic synthesis.^{2,13} Not surprisingly, most of these compounds also contain nitrogen-rich ligand sets.

The mononuclear copper proteins of the second group do not necessarily catalyze reactions but rather contribute to cellular copper-transport and sensing. Representative examples are the Cu-responsive transcriptional activator CueR,¹⁴ and the copper chaperone Atox1.¹⁵ These proteins bind Cu(I) in a linear bis-cysteine coordination sphere with extremely high affinity.¹⁶ This coordination-type provides an ideal solution to form redox inert protein:copper complexes with maximal thermodynamic stability, while enabling reversible transfer of Cu(I) from one protein to another.^{15,17}

The formylglycine generating enzyme (FGE) is unique, since it cannot be attributed to either the first or the second group of copper-binding proteins. FGE binds a single Cu(I) by linear bis-cysteine coordination and catalyzes copper-dependent O₂ activation.^{18,19} FGE catalyzes the oxidative conversion of specific peptidyl-cysteines to formylglycine *via* abstraction of the pro-(*R*)- β -hydrogen atom (fGly, Fig. 1). This posttranslational modification is important for sulfatases which use the hydrated form of fGly as catalytic nucleophile.²⁰ FGE has attracted scientific interest for several reasons. Initially, this enzyme was discovered in the pursuit of finding the molecular cause for the storage disease *multiple sulfatase deficiency*.^{20,21} The subsequent realization that FGE can be used as a tool to introduce

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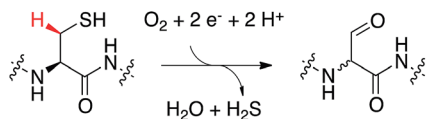


Fig. 1 FGE-catalyzed oxidation of peptidyl-cysteine to formylglycine (fGly) is initiated by abstraction of the pro-(*R*)- β -hydrogen atom from the substrate (red).

electrophilic aldehyde functions into recombinant proteins highlighted the considerable biotechnological potential of this enzyme.^{22–24} Finally, the recent discovery that FGE is a copper-dependent enzyme opened the door for mechanistic studies on this unique enzyme.^{18,19,25,26} The crystal structure of FGE from *Thermomonospora curvata* (*tcFGE*) in complex with Ag(I) (*tcFGE*_Ag, PDB: 5NXL), and most recently, the crystal structure of FGE from *Streptomyces coelicolor* in complex with Cu(I) (*scFGE*_Cu, PDB: 6MUJ), combined with biochemical characterization showed that two active site cysteines are the only metal ligands and that the Cu(I) bound state is the catalytic resting state.^{25–27} In this way, FGE represents an entirely new type of copper-enzyme. Structural and mechanistic studies on this enzyme may inspire new directions in the design of abiotic copper-catalysts.² Conversely, the similarities of the primary metal coordination sphere in FGE and non-catalytic Cu-binding proteins raise the questions as to what secondary features may enable O₂-activation by FGE. Identification of these features may help to identify redox enzymes, masquerading as copper-trafficking proteins, or could provide guidelines to engineer novel redox-activity into existing copper-proteins.

In this report we describe the crystal structure of *tcFGE* in complex with Cu(I) and a 17-residue substrate analog. Based on this structure together with NMR spectroscopy and kinetic characterization, we identify an acidic O₂-binding pocket juxtaposed to the copper center as key determinant for efficient O₂-activation.

Results and discussion

Crystallization of *tcFGE* in complex with Cu(I) and a substrate-analog

Recombinant *tcFGE* was produced in *Escherichia coli* (see ESI†).¹⁹ The purified enzyme was crystallized in the presence of 1.2 equivalents of Cu(I) and 2.8 equivalents of a synthetic 17-residue substrate-analog. To stabilize the enzyme:substrate complex, crystallization and room-temperature handling of the crystals were executed under anoxic conditions. The sequence of the substrate-analog (FGE-27: Abz-ATTPLCGPSRASILSGR, Abz = *o*-aminobenzoic acid) was derived from the active site segment of a putative arylsulfatase (WP_012850446) from *T. curvata*. Using isothermal titration calorimetry (ITC), we established that this peptide binds to an inactive variant of *tcFGE* (*tcFGE*_{W228F}, see below) with a $K_d = 40 \pm 10$ nM (Fig. S2†), indicating that the wild type enzyme would also form a stable complex with this peptide. Single crystals were isolated from solutions containing purified *tcFGE* (21.5 mg mL⁻¹), CuCl, peptide, PEG 8000 (10%), Tris-HCl (0.1 M, pH 7.0) and MgCl₂

(0.2 M). Diffraction data could be processed up to a resolution of 1.04 Å (Fig. 2A). At this resolution conservative estimates of standard deviations (r.m.s.) are ± 0.05 Å for bond lengths and $\pm 2^\circ$ for bond angles.²⁸ The unbiased electron density revealed a continuous polypeptide chain from -1 to 302 for *tcFGE* and from Abz (-6) to Gly (11) for the substrate FGE-27. Two Ca(II) ions were identified in two conserved calcium binding sites (Fig. S3†).^{26,29,30} The high resolution of this data set allowed placement of 494 crystallographic waters, in particular four waters in the active site (H₂O_{1–4}) with high confidence (Fig. S4†). The overall structure of *tcFGE* and its similarity to human FGE (*hsFGE*) and *scFGE* has been discussed previously.^{26,29,30} Therefore, the main focus here will be on the structure of the substrate binding pocket and the metal binding site. For this discussion we compare *tcFGE* in complex with Cu(I) and substrate (*tcFGE*_Cu_S) with the structure of *tcFGE* in complex Ag(I) (*tcFGE*_Ag) as a model for *tcFGE* in complex with Cu(I) (*tcFGE*_Cu). The structure of *scFGE* in complex with Cu (*scFGE*_Cu) has also become available very recently.²⁷ The metal binding sites in *scFGE*_Cu and *tcFGE*_Ag were found to be essentially isostructural.

Position of the substrate

The unbiased electron density for the peptide reveals ordered positions for all residues except for the C-terminal Arg10 (Fig. 2B). The residues in the N-terminal part (Ala-5–Ser4), except for Cys1, adopt extended backbone conformations ($\phi = -50^\circ$ to -136° ; $\psi = 120^\circ$ – 170°). The conformation of Cys1 ($\phi = -108^\circ$; $\psi = -20^\circ$) induces a sharp turn allowing this residue to access the copper localized at the deepest point of the substrate binding pocket (Fig. 2). Pro-2 and Pro3 make extensive hydrophobic interactions with Phe38/Tyr273 and Phe86/Trp109 respectively. In contrast, Ala-5, Thr-4, Thr-3 and Leu-1 make little or no contact with the protein surface. The C-terminal residues Arg3–Ile6 adopt dihedral angles consistent with a helical conformation ($\phi = -90$ to 62 ; $\psi = -30$ to -5). Indeed, the network of intramolecular backbone-backbone hydrogen bonds is reminiscent of a short ₃₁₀-helix (Fig. 2C). The side chain of Ser4 hydrogen bonds to the backbone amide of Ser7 (3.0 Å), serving as a N-terminal helix cap.³¹ In return, the side chain of Ser7 hydrogen bonds with the backbone carbonyl of Ser4 (2.7 Å). The guanidinium side chain of Arg5 makes a π -cation interaction with the indole ring of Trp84, in addition to hydrogen-bonds with Ser290 (2.8 Å), and the backbone carbonyl of Trp84 (2.9 Å). The side chain of Ile8 binds to a hydrophobic dell lined by the side chains of Trp84, Phe86, Met99 and Ala101. The side chains of Ala6, Leu9 and the following residues (Ser10, Gly11 and Arg12) make no interaction with the protein surface. The helical structure of the substrate may be further stabilized by weak interactions with neighboring *tcFGE* chains (Fig. 2D and E).

hsFGE has been crystallized with a seven-residue substrate (LCTPSRA, *hsFGE*_S, PDB: 2AIK).³² This structure contains no metal in the active site and Cys366 (Cys274 in *tcFGE*) was mutated to Ser. Instead, the substrate cysteine is linked *via* disulfide bond to Cys341 (Cys269 in *tcFGE*). Despite these damages, the equivalent substrate residues in *tcFGE*_Cu_S and



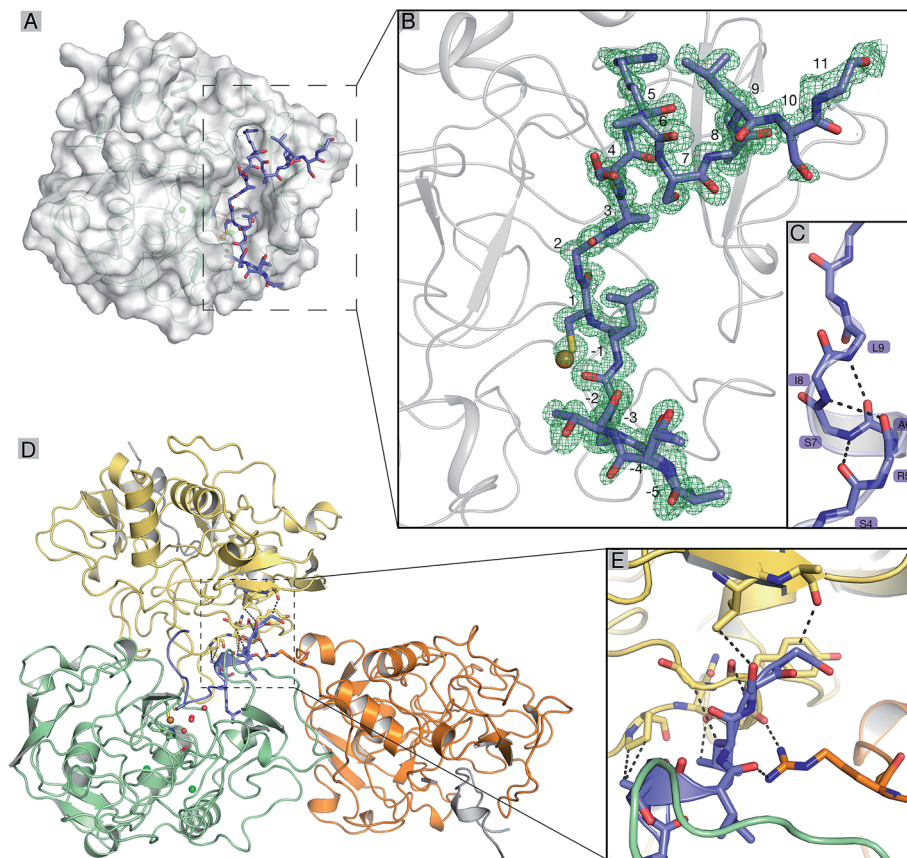


Fig. 2 (A) Structure of *tcFGE* in complex with Cu(I) and substrate (*tcFGE_Cu_S*). The 17-residue peptide substrate (violet). (B) The $m|F_o| - D|F_c|$ omit map for the substrate model is contoured at σ -level = 3.0. (C) The C-terminus of the substrate folds into a short 3_{10} -helix and interacts with neighboring *tcFGE* chains in the crystal (D and E). The dashed lines in (E) indicate nearest contacts, not necessarily attractive interactions.

hsFGE_S adopt almost identical positions in the substrate binding pocket (Fig. S5†). Apparently, residues 0–5, the conserved residues Pro3 and Arg5 in particular (numbering according to *tcFGE_Cu_S*) make sufficient interactions with the substrate binding pocket to steer the targeted cysteine residue into the active site – even in the absence of Cu(I).

Indeed, the biotechnological potential of FGE depends on its ability to target cysteine residues embedded in the minimal LCXPXR motif. The limited length of this motif can be inserted into N- or C-termini or surface exposed loops of unrelated proteins without significant disturbance of the structure and stability of the parent protein.^{22,23,33,34} On the other hand, comparison of kinetic parameters from *in vitro* studies with 9-, 14- and 23-residue substrates showed that longer substrates are recognized with significantly higher affinity.^{18,19,35} The structure of *tcFGE_Cu_S* provides a convincing explanation for this observation. The C-terminal residues Arg5–Ile8 adopt a defined secondary structure that positions Ile8 for additional hydrophobic contact with FGE. Likewise, Pro-2 also makes a specific contact with the enzyme that has not been observed in the structure of *hsFGE_S*. The more comprehensive picture of the enzyme:substrate interaction seen in the structure of *tcFGE_Cu_S* could provide a valuable starting point for the design of FGE variants that accept alternative motifs with applications in orthogonal bifunctionalization of recombinant proteins.³⁶

Primary and secondary coordination sphere

The structure of *FGE_Cu_S* demonstrates that the peptidyl cysteine (Cys1) from the substrate is a direct ligand of the catalytic Cu(I). This finding corroborates XAS observations suggesting trigonal coordination of Cu(I) by *scFGE* and its substrate.²⁷ The three thiol side chains of Cys269 (Cu–S_{Cys269}: 2.2 Å), Cys274 (Cu–S_{Cys274}: 2.3 Å) and the Cys1 (Cu–S_{Cys1}: 2.2 Å) coordinate Cu(I) in an irregular trigonal planar complex with significant deviations from the expected 120° S–Cu–S bond angles (S_{Cys269}–Cu–S_{Cys274}: 108°; S_{Cys274}–Cu–S_{Cys1}: 104°; S_{Cys269}–Cu–S_{Cys1}: 147°). The angle between the two FGE-derived cysteines and Cu(I) is significantly smaller than the value (134°) derived from DFT-based modelling based on the structure of *scFGE_Cu*.²⁷ Instead, the much larger S_{Cys269}–Cu–S_{Cys1} angle opens the triangle enabling two crystallographic waters (H₂O_1, 3.3 Å; H₂O_2, 3.5 Å) to form near van der Waals contact with Cu(I). Notably, these two waters are also immobilized by planar trigonal solvation spheres. H₂O_2 is in contact with H₂O_1, H₂O_3 and H₂O_4 (all 2.8 Å, O₁–O₂–O₄: 111°; O₁–O₂–O₃: 124°; O₃–O₂–O₄: 119°). H₂O_1 also binds to the indole side chain of Trp228 (2.9 Å) and Ser266 (2.8 Å, O_{Ser266}–O₁–N_{Trp228}: 89°; N_{Trp228}–O₁–O₂: 129°; O₂–O₁–O_{Ser266}: 140°). In addition, H₂O_1 and H₂O_2 make weak contacts with the thiolates of Cys269 (3.1 Å) and Cys-1 (3.3 Å). This solvation geometry deviates significantly from the more stable tetrahedral solvation spheres



observed in bulk water,³⁷ suggesting that the two crystallographic waters are held in unstable positions. Release of H₂O₁ and H₂O₂ from this pocket may provide important enthalpic driving force to enhance O₂-binding (see below).^{38–41}

Inspection of the next neighbors to Ser266, H₂O₃ and H₂O₄ provides information about the position of the protons in this hydrogen-bonding network. In addition to contacting H₂O₁, the side chain of Ser266 also hydrogen bonds with the backbone amide of Leu268 (3.0 Å). This interaction suggests that Ser266 is a hydrogen-bond donor with respect to H₂O₁. H₂O₃ is in hydrogen-bonding contact with Trp108 (2.9 Å), Trp109 (2.9 Å) and the backbone carbonyl of Leu-1 of the substrate (2.7 Å). This tetrahedral system suggests that H₂O₃ is a hydrogen bond donor with respect to H₂O₂. H₂O₄ is in contact with the backbone carbonyl of Leu268 (2.8 Å) and is otherwise engulfed by hydrophobic moieties of Phe194, Tyr273 and Cys269. This environment too suggests that H₂O₄ is a hydrogen bond donor with respect to H₂O₂. Hence, the pocket that hosts H₂O₁ and H₂O₂ in the *tcFGE_Cu_S* structure appears as a highly acidic microenvironment.

Hydrogen bonding to metal-coordinated thiolates are known to affect the nucleophilicity and redox potential of metal-thiolate complexes.^{42–45} Such interactions are indeed observable in *tcFGE_Cu_S*. Cys1 makes two weak hydrogen bonds to H₂O₂ and Arg276, 3.3 Å, Cys269 hydrogen bonds to H₂O₁ and the backbone amide of Ser266 (3.3 Å), whereas Cys274 only points to the backbone amide of His270 (3.6 Å). This network may play an important role in stabilizing the high electron density of the tris-thiolate Cu(i) complex. Incidentally, Cys274 appears as the least stabilized thiolate in the complex suggesting that this residue may be most vulnerable to oxidation during catalysis or as a result of abortive side reactions (see below).

Comparison of Ag, Cu and Cd-bound FGE reveals a sequence of coordination changes

The structure of *tcFGE_Cu_S* (resolution: 1.04 Å) complements the structures of the same enzyme in complex with Ag(i) (resolution: 1.66 Å) or with Cd(ii) (*tcFGE_Cd*, PDB: 5NYY, resolution: 1.28 Å). Comparison of the three structures provide a high-resolution picture of the structural changes that take place as the coordination sphere expands from linear to trigonal to tetrahedral in order to allow inner-sphere coordination of substrate and O₂.

Superposition of the *tcFGE_Ag* and *tcFGE_Cu_S* structures highlights three significant changes that are induced by substrate binding. First, the metal center moves by 1.1 Å as it changes from linear to trigonal coordination (Fig. 3). This movement reduces the distance from the metal to H₂O₁ and H₂O₂ by about 0.3 Å and 0.8 Å respectively. As a second change, H₂O₁ and H₂O₂ are displaced by 0.6 Å from predominantly tetrahedral to purely trigonal planar hydrogen bonding environments. In addition, the interaction partners of H₂O₂ change. In *tcFGE_Ag* H₂O₂ hydrogen bonds with four partners: H₂O₁ (2.6 Å) H₂O₅₀ (equivalent to H₂O₃, 2.8 Å), H₂O₁₀₅ (2.7 Å) and Tyr273 (2.8 Å, Fig. S7†). Substrate binding displaces H₂O₁₀₅, leaving H₂O₂ with only three nearest

neighbors in *tcFGE_Cu_S*. In addition, the side chain of Tyr273 is rotated out of the active site to establish a hydrogen-bond to the backbone amide of Leu-1 of the substrate. Instead, H₂O₄ is recruited as third neighbor of H₂O₂ in *tcFGE_Cu_S*. This exchange is likely accompanied with a change in hydrogen bond polarity. In *tcFGE_Ag* Tyr273 hydrogen bonds with the backbone carbonyl of Leu268 (2.7 Å) (Fig. S7†), indicating that Tyr273 is a hydrogen bond acceptor with respect to H₂O₂. In contrast, H₂O₄ in *tcFGE_Cu_S* likely acts as a hydrogen bond donor (see above). These changes suggest that substrate binding induces destabilization of H₂O₁ and H₂O₂. A third substrate-induced change occurs at Cys274. The dihedral angle along Cα–Cβ of this residue changed by 58° (Nα–Cα–Cβ–S) and the amide function between Cys274 and Asn275 rotated by 135° presumably as a result of the change from linear to trigonal Cu(i)-coordination (Fig. S6†). A similar rotation was observed in the *tcFGE_Cd* structure (135°). Hence, this reorganization also seems to be driven by substrate binding.

The structure of *tcFGE_Cd* provides a model for the enzyme with a tetrahedral coordination site which may be required simultaneous coordination of the substrate and a reactive oxygen species.²⁷ Superposition of the structures of *tcFGE_Cu_S* and *tcFGE_Cd* shows that the transition from trigonal to tetrahedral coordination moves the metal again by 0.8 Å and reduces the distance between the metal and the position of H₂O₁ and H₂O₂ down to 2.8 Å and 2.7 Å (Fig. 3). Tyr273 adopts the same position as in *tcFGE_Ag*, confirming that the orientation of this residue is dependent on the presence of substrate, rather than on the geometry of the metal complex (Fig. 3). Otherwise, the *tcFGE_Cu_S* and *tcFGE_Cd* structures are remarkably similar, suggesting that most active site reorganization that occurs en route from the Cu(i) resting state to the reactive complex may be induced by substrate-binding.

Trp228 and Ser266 form an anion binding site

The structure of *tcFGE_Cu_S* identifies the position of H₂O₁ and H₂O₂ as the only available pocket for O₂ to initiate an interaction with the copper center. The opposite face of the trigonal complex is blocked by the side chains of Trp228 and Arg276. Hence, further scrutiny of the subtle changes that occur at this site in response to substrate binding may provide additional insight as to how this unusual metal center can activate O₂. The apo structures of *hsFGE* (PDB: 1Y1E)²⁹ and *scFGE* (PDB: 2Q17),³⁰ the metalated forms *tcFGE_Ag*,²⁶ *scFGE_Cu*,²⁷ and *tcFGE_Cu_S* all contain one crystallographic water (H₂O₁) in hydrogen-bonding distance to Ser266 (2.6–2.8 Å) and Trp228 (2.9–3.2 Å, Fig. S7†). In the apo structures the hydrogen bonds with Trp228 are likely weak due to a small bond angle (N–H–O < 130°, Table S4†). Nevertheless, several crystal structures of *hsFGE* with halides bound to the approximate position of H₂O₁ indicate that this pocket is endowed with significant affinity for anionic species.^{32,46} Most interestingly, addition of metal and substrate to *tcFGE* increased the bond angle between H₂O₁ and the indole of Trp228 to 148° (*tcFGE_Ag*) and 157° (*tcFGE_Cu_S*) (Table S4†), suggesting that the presence of metal and substrate also increases bonding interactions at the O₂-



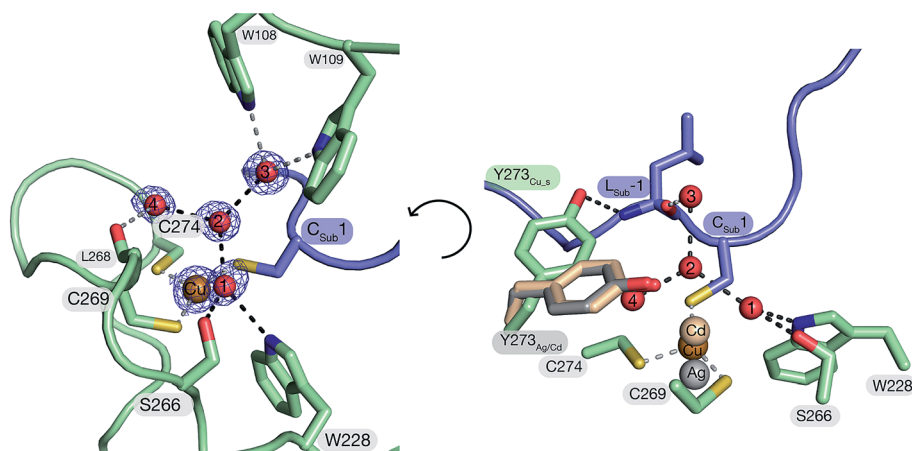


Fig. 3 Left: The primary and secondary coordination sphere of Cu(I) in the crystal structure of *tcFGE_Cu_S* (protein: green, substrate: violet). The $2m|F_o| - D|F_c|$ omit map of Cu(I) and the four crystallographic water molecules (H_2O_{1-4}) is contoured at σ -level = 1.0. Right: Active site of *tcFGE_Cu_S* including the metals and the side chain of Tyr273 from the structures of *tcFGE_Ag* (gray) and *tcFGE_Cd* (beige).

binding site. Also, based on the observed anion-affinity of this pocket we have proposed previously that H_2O_1 in *tcFGE_Ag* might be a hydroxide that serves as the base that deprotonates the thiol of the incoming peptidyl cysteine.²⁶

Observation of Trp228 by NMR

To examine the hydrogen bonding interaction between H_2O_1 and Trp228 in more detail we used NMR spectroscopy. In addition to directly interrogate the chemical environment of specific protons NMR spectroscopy has the key advantage over X-ray measurements (at low temperatures) that the enzyme is examined in solution and at room temperature. 1H - ^{15}N TROSY HSQC spectra (transverse relaxation-optimized spectroscopy heteronuclear single quantum correlation) were measured using uniformly ^{15}N -labelled *tcFGE* and *tcFGE_{W228F}* in apo form and in complex with Cu(I) on a 600 MHz Bruker NMR instrument.⁴⁷ Superposition of these spectra, in combination with a triple resonance HNCO experiment that distinguishes amide-specific resonances from indole-specific resonances (Fig. S8 and S9[†]),⁴⁸ allowed unambiguous identification of the $H^{\epsilon 1}$ and $N^{\epsilon 1}$ signals relating to the indole side chain of Trp228 (Fig. S10–S12[†]). The chemical shifts of indole $H^{\epsilon 1}$ and $N^{\epsilon 1}$ signals are highly sensitive to their environment, reporting on the solvation of the N–H function and the aromatic ring system.^{49,50} Therefore, we reasoned that the Trp228 signal could be used to monitor the protonation state and movements of H_2O_1 . The recorded spectra show that addition of Cu(I) to apo *tcFGE* induced a significant downfield shift of the $H^{\epsilon 1}$ signal (Table 1 and Fig. 4). Addition of the substrate to *tcFGE_Cu* induced a further $H^{\epsilon 1}$ shift of similar magnitude. Addition of metal also caused slight deshielding of $N^{\epsilon 1}$, consistent with increased strength of the hydrogen bond between H_2O_1 and Trp228. In contrast, addition of substrate caused a marked shielding effect. Inspection of the crystal structures shows, that the indole ring of Trp228 is solvent exposed in *tcFGE_Ag* and is buried in *tcFGE_Cu_S* (Fig. S13[†]). Hence, addition of the substrate reduces the polarity of the medium above the indole ring

consistent with the observed shielding effect. The observed deshielding of the indole proton is fully consistent with the proposition that metal and substrate binding increase bonding interactions at the O_2 -binding site.

Complexation of apo *tcFGE* with Ag(I) induced a significantly smaller downfield shift of the $H^{\epsilon 1}/N^{\epsilon 1}$ signal than observed for Cu(I) (Table 1 and Fig. 4). A likely cause for this small but significant difference is that the bond angle between the sulfur ligands and the metal is larger for Ag(I) than for Cu(I).²⁷ This slight movement of the metal would push H_2O_1 closer to Trp228, and therefore strengthen the interaction between the two. The structures of *tcFGE_Ag* (S–Ag–S: 178°) and *scFGE_Cu* (S–Cu–S: 171°) indeed document a small structural difference.²⁷ Addition of the substrate to *tcFGE_Ag* shifted the $H^{\epsilon 1}/N^{\epsilon 1}$ signal of *tcFGE_Ag_S* to exactly the same position as measured with *tcFGE_Cu_S* (Table 1), suggesting that the two complexes adopt very similar structures. The crystallographic and spectroscopic findings validate Ag(I) as a reliable structural mimic of Cu(I) in FGE.

Using this sensitive NMR methodology, we examined the proposition that H_2O_1 in the *tcFGE_Cu* is stabilized as a hydroxide. Even though H_2O_1 is not a metal ligand, it is possible that the proximity of an additional cationic charge, compounded by the interactions with Ser266 and Trp228 could stabilize a hydroxide at physiological pH. Since the pK_a of water closely matches that of indole rings, a hydrogen bond between indole and hydroxide is likely stronger than between indole and water. Superposition of three 1H - ^{15}N TROSY HSQC spectra for *tcFGE_Cu* measured at pH 6, 7 and 8 showed no change of the Trp228 signal relative to other indole-specific resonances (Fig. 4). Based on this observation we concluded that either the pK_a of H_2O_1 is far below 6, or that H_2O_1 is neutral at physiological pH. The second option is far more likely.

Role of Ser266 and Trp266 in the resting state and in catalysis

Finally, we examined the effect of the Ser266 to Ala mutation on the interaction of H_2O_1 with Trp228. 1H - ^{15}N TROSY HSQC



Table 1 The observed H^{e1} and N^{e1} chemical shifts of Trp228 sidechain upon metalation and substrate binding^a

Enzyme	Metal	$\delta H^{e1}/N^{e1}$ Trp228, ppm			$\Delta \delta H^{e1}/N^{e1}$ (Trp228), ppm		
		E	E : M	E : M : S	E \rightarrow E : M	E : M \rightarrow E : M : S	E \rightarrow E : M : S
<i>tcFGE</i>	Ag(I)	9.69/128.93	9.88/129.95	10.26/126.81	0.19/1.02	0.38/−3.14	0.57/−2.12
<i>tcFGE</i>	Cu(I)	9.69/128.93	9.98/130.75	10.22/126.66	0.29/1.82	0.24/−4.09	0.53/−2.27
<i>tcFGE</i> _{S266A}	Cu(I)	9.75/129.29	10.04/130.78	10.22/126.77	0.29/1.49	0.18/−4.01	0.47/−2.52

^a The errors of the given values are estimated to be <0.02 ppm.

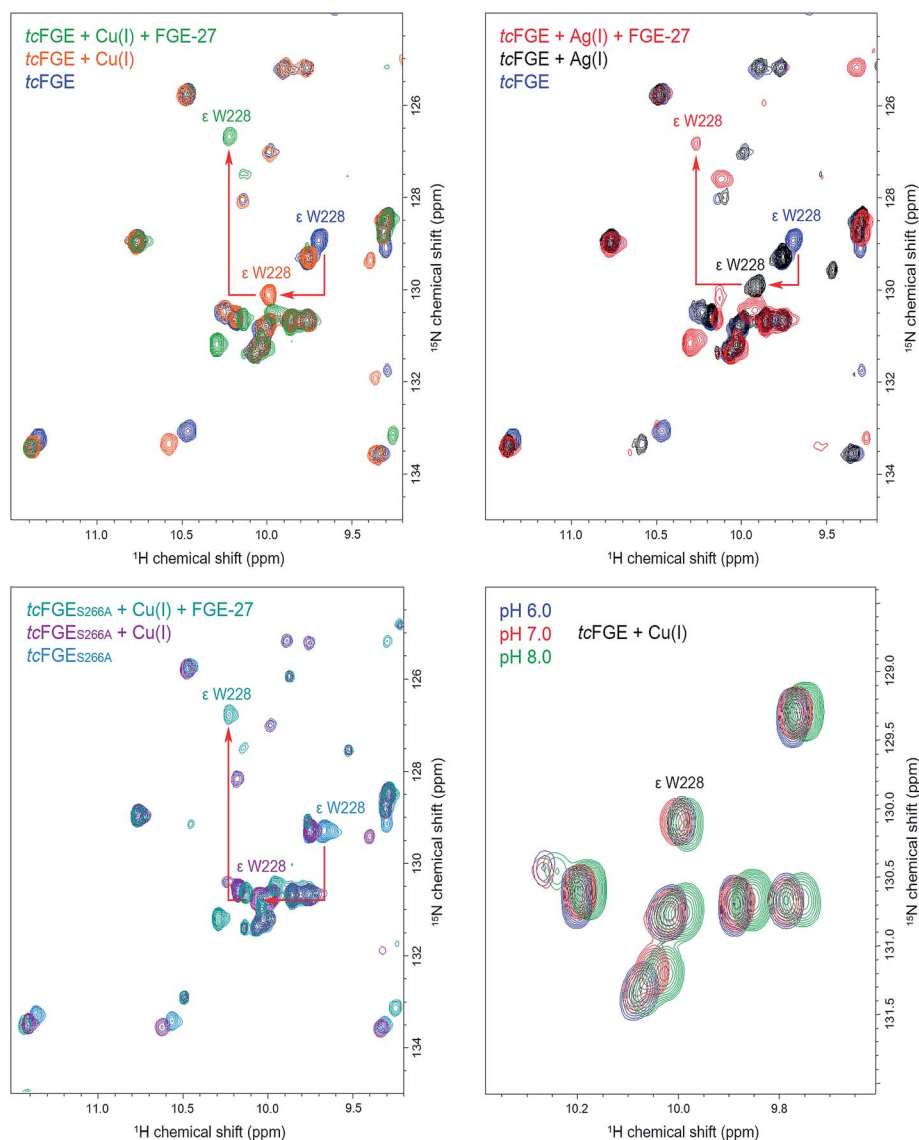


Fig. 4 1H - ^{15}N TROSY HSQC spectra of uniformly ^{15}N -labelled 300 μM *tcFGE* in 20 mM phosphate buffer, 50 mM NaCl, 2 mM DTT, 25 C. The downfield area of the amide region is presented. Top left: The overlay of the spectra of apo-form *tcFGE*, pH 8.0 (dark blue), *tcFGE* in the presence of 450 μM $CuSO_4$, pH 8.0 (orange), *tcFGE* in the presence of 450 μM $CuSO_4$ and 1 mM FGE-27, 2 mM glucose, 2 units (U) glucose oxidase (GO), 200 U catalase (CAT), pH 8.1 (green). Top right: The overlay of the spectra of apo-form *tcFGE*, pH 8.0 (dark blue), *tcFGE* in the presence of 450 μM $AgNO_3$, pH 8.0 (black), *tcFGE* in the presence of 450 μM $AgNO_3$ and 1 mM FGE-27, 2 mM glucose, 2 U GO, 200 U CAT, pH 8.1 (red). Bottom left: The overlay of the spectra of apo-form *tcFGE*_{S266A}, pH 8.0 (blue), *tcFGE*_{S266A} in the presence of 450 μM $CuSO_4$, pH 8.0 (magenta), *tcFGE*_{S266A} in the presence of 450 μM $CuSO_4$ and 1 mM FGE-27, 2 mM glucose, 2 U GO, 200 U CAT, pH 8.1 (dark cyan). Bottom right: The overlay of the spectra of *tcFGE* in the presence of 450 μM $CuSO_4$ measured at different pH.



spectra recorded for $tcFGE_{S266}$, $tcFGE_{S266A-Cu}$ and $tcFGE_{S266A-Cu_S}$ revealed only small or no deshielding effects on the H^{e1}/N^{e1} signal, suggesting that the absence of the Ser266 side chain does not affect the interaction between H_2O_1 and Trp228. Hence, the role of Ser266 in immobilizing H_2O_1 in the $tcFGE_Cu_S$ complex is limited. In stark contrast, the same mutation causes a 50-fold reduction in catalytic turnover (k_{cat}) without affecting substrate binding (K_m , Table 2).²⁵ The kinetic and spectroscopic properties of $tcFGE_{S266A}$ suggest that Ser266 provides a hydrogen bond that become important after substrate-binding. Similarly, mutation of Trp228 to Phe ($tcFGE_{W288F}$) also specifically affects k_{cat} without hurting substrate binding. The lack of this hydrogen bond donor in the O_2 -binding site is likely to contribute to the demise of catalysis by this variant.

The unusual resting state of FGE

The active site of FGE is unusual for several reasons. First, catalysts that use a bis-cysteine coordinated Cu(I) as a starting point for O_2 -dependent C-H cleavage is unexplored territory. A large number of studies have shown quite conclusively that imidazole, amine, amide and thioether ligands are particularly well suited to facilitate copper-mediated O_2 -activation, not the least because these ligands stabilize high-valent copper species.^{1-3,51,52} From this perspective, bis- or tris-thiolate coordination spheres seem odd choices for designing copper-dependent oxidation catalysts. Thiolate ligands stabilize Cu(I) which should render electron transfer to O_2 more difficult. On the other hand, thiolates are intrinsically vulnerable to oxidation which may open efficient paths for catalyst destruction. A specific complication of the FGE-catalyzed reaction is that the two active site cysteines and the substrate contain altogether six chemically equivalent cysteinyl C_{β} -H bonds. Hence, the geometry of the active site must ensure that the reactive oxygen species exclusively attacks the pro-(R)- β -hydrogen on the substrate.¹⁹ The structure of $tcFGE_Cu_S$ visualizes how the enzyme achieves regioselectivity. All three cysteines approach Cu(I) in a similar angle (C_{β} -S-Cu: 103–120°) and place their C_{β} at a similar distance to the copper center (3.2–3.5 Å). However, only the pro-(R)- β -hydrogen of the substrate points towards the presumed O_2 -binding site (Fig. 5).

The second unusual feature of FGE is that the substrate and the reactive oxygen species both coordinate to the metal.^{1,27} Metal-dependent oxidases often bind O_2 via displacement of a weakly bound metal ligand, such as water, at a pre-formed coordination

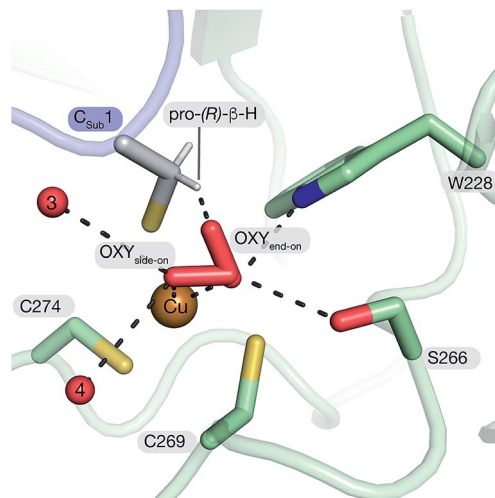


Fig. 5 Structure-based model of $tcFGE_Cu_S$ with superoxide coordinated to Cu(II) in side-on mode (η^1) or end-on mode (η^2).

site. Formation of the coordination bond is accompanied by reduction of O_2 to superoxide via inner-sphere electron transfer. For example, PHM coordinates Cu(I) via the side chains of two histidines and one methionine in a near tetrahedral geometry. The fourth ligand is a water molecule that is displaced by O_2 .^{9,10} In LPMO Cu(I) is coordinated by a histidine brace ($N\alpha/N\pi$) and the side chain of a second histidine.⁷ The square planar coordination sphere is completed by a chloride or a water molecule filling the dedicated O_2 binding site. CAO coordinates copper with three histidines and a loosely bound water. Displacement by O_2 occurs without change in the coordination geometry.¹¹ We note that the path leading to the cupric superoxide in CAO is controversial. In the conventional mechanism O_2 forms a coordination bond with Cu(I).⁵³ An alternative proposal suggests that prebound O_2 is first reduced by a protein-derived redox cofactor, and then combines as superoxide with Cu(II).⁵⁴

In contrast to these enzymes, the substrate complex of $tcFGE$ contains no preformed coordination site for a fourth ligand. The tris-thiolate Cu(I) complex is completely planar and the closest crystallographic waters (H_2O_1 and H_2O_2) are too distant for attractive interactions. Hence, direct binding of O_2 would require a change in coordination geometry. The structure of $tcFGE$ in complex with Cd(II) showed a tetrahedral coordination sphere around this bivalent metal. Therefore it is plausible that oxidation of Cu(I) to Cu(II) may trigger a change from trigonal planar to tetrahedral coordination.²⁶ Importantly, if O_2 -coordination requires structural change, and this change requires oxidation of Cu(I), it is unlikely that this redox reaction occurs via inner-sphere electron transfer as in other mononuclear copper enzymes. A more plausible scenario is that O_2 first binds to the active site without direct contact to the metal (C, Fig. 6). Outer sphere electron transfer from Cu(I) to pre-bound O_2 followed by coordination would produce the predicted cupric superoxide complex (D). The efficiency of outer sphere electron transfer would critically rely on a binding site that increases the redox potential of pre-bound O_2 .

Table 2 The Characterization of the catalytic activity of $tcFGE$ variants using HPLC-based kinetic assay^a

Enzyme	k_{cat} , [min ⁻¹]	K_m , [μ M]	k_{cat}/K_m , [min ⁻¹ M ⁻¹]
$tcFGE$	1.4 ± 0.07	530 ± 40	2700 ± 200
$tcFGE_{S266A}$ ^b	0.025 ± 0.002	520 ± 240	49 ± 8
$tcFGE_{W228F}$	0.0073 ± 0.0002	390 ± 20	19 ± 1

^a Kinetic parameters were determined in the presence of 2 μ M $CuSO_4$, 5 mM DTT, 50 mM NaCl and 50 mM Tris buffer pH 8.0, at 25°C. Rate determined by monitoring the formation of fGly. ^b From ref. 25.



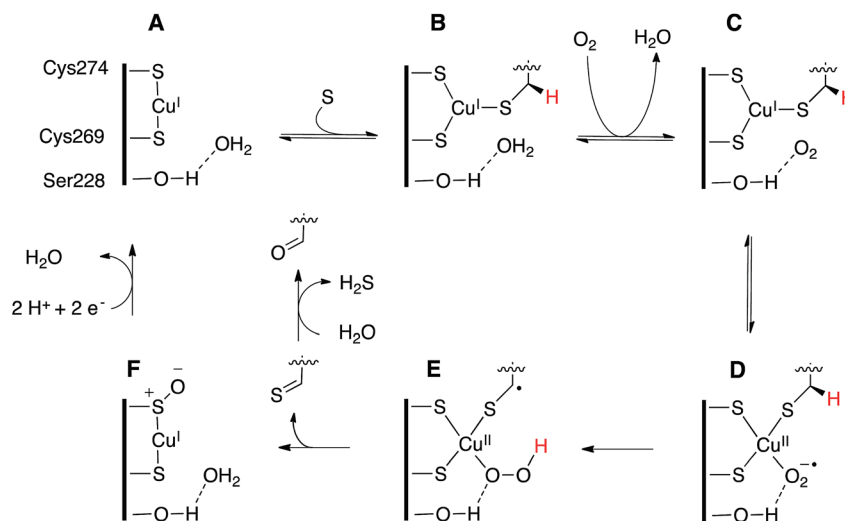


Fig. 6 Proposed catalytic mechanism of FGE. The reaction starts with sequential binding of substrate and O_2 forming first complex **B** (crystallized) and then **C**. Outer sphere electron transfer from $Cu(I)$ to O_2 could produce the $Cu(II)$ superoxide species **D** (modelled in Fig. 5). Homolytic abstraction of the pro-(*R*)- β -hydrogen atom from the substrate (**E**), followed by release and hydrolysis of the oxidized peptide product and reduction of the oxidized enzyme form (**F**) return the enzyme to the catalytically active resting state **A**.

This requirement highlights the third unusual feature of FGE. The presumed O_2 -binding site in *tcFGE* is lined with an extensive array of hydrogen bond donors. Hydrogen-bonding has been recognized as important modulator of the stability and reactivity of copper-coordinated oxygen species in enzymes and in abiotic catalysts.^{1,2,55–57} For example, hydrogen-bonding to the proximal oxygen atom of a synthetic cupric superoxide complex has been shown to increase its stability and its propensity for hydrogen atom abstraction in direct correlation with hydrogen-bonding strength.⁵⁵ Furthermore, specific hydrogen-bonding to the distal or to the proximal oxygen has been shown to affect the activity of copper-oxygen adducts in very different ways.⁵⁶ On the other hand, oxygen-binding pockets lined by four well-positioned hydrogen bond donors are rare among metal-dependent catalysts. Studies on synthetic cages showed that solvation of O_2 with multiple amide- or amine-donated hydrogen bonds can significantly stabilize superoxide and peroxide anions.^{58,59} The acidic O_2 -binding pocket in FGE may play a similar stabilizing role and thereby facilitate electron transfer from the electron-rich tris-thiolate $Cu(I)$ to O_2 . The observation that mutation of Ser266 or Trp228 dramatically reduce the activity of *tcFGE* is consistent with this interpretation.

Conclusions

In this report we describe the crystal structure of *tcFGE* in complex with $Cu(I)$ and a 17-residue peptide. Comparisons of this complex with the structures of *tcFGE* bound to $Ag(I)$ or $Cd(II)$ highlight substrate-induced changes that may prime the active site for O_2 -binding. On the other hand, we also found that substrate binding does not induce a preformed coordination site for O_2 at the metal center. These observations raise the possibility that FGE activates O_2 in a different way than other

metallo-enzymes. Specifically, we propose that O_2 first extracts an electron from $Cu(I)$ via outer sphere electron transfer before combining to a cupric superoxide species. This and other testable hypotheses emanating from our structural investigating set the stage for deciphering the detailed mechanism of this unusual oxidation catalyst. We also hope to inspire the synthesis and characterization of abiotic catalysts that exploit this reactivity of bis- and tris-thiolate coordinated copper species.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 E. I. Solomon, D. E. Heppner, E. M. Johnston, J. W. Ginsbach, J. Cirera, M. Qayyum, M. T. Kieber-Emmons, C. H. Kjaergaard, R. G. Hadt and L. Tian, *Chem. Rev.*, 2014, **114**, 3659–3853.
- 2 C. E. Elwell, N. L. Gagnon, B. D. Neisen, D. Dhar, A. D. Spaeth, G. M. Yee and W. B. Tolman, *Chem. Rev.*, 2017, **117**, 2059–2107.



- 3 W. Keown, J. B. Gary and T. D. Stack, *J. Biol. Chem.*, 2017, **22**, 289–305.
- 4 S. E. Allen, R. R. Walvoord, R. Padilla-Salinas and M. C. Kozlowski, *Chem. Rev.*, 2013, **113**, 6234–6458.
- 5 J. T. Rubino and K. J. Franz, *J. Inorg. Biochem.*, 2012, **107**, 129–143.
- 6 R. J. Quinlan, M. D. Sweeney, L. L. Leggio, H. Otten, J. C. N. Poulsen, K. S. Johansen, K. B. R. M. Krogh, C. I. Jorgensen, M. Tovborg, A. Anthonsen, T. Tryfona, C. P. Walter, P. Dupree, F. Xu, G. J. Davies and P. H. Walton, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, **108**, 15079–15084.
- 7 K. E. Frandsen, T. J. Simmons, P. Dupree, J. C. Poulsen, G. R. Hemsworth, L. Ciano, E. M. Johnston, M. Tovborg, K. S. Johansen, P. von Freiesleben, L. Marmuse, S. Fort, S. Cottaz, H. Driguez, B. Henrissat, N. Lenfant, F. Tuna, A. Baldansuren, G. J. Davies, L. Lo Leggio and P. H. Walton, *Nat. Chem. Biol.*, 2016, **12**, 298–303.
- 8 L. Cao, O. Caldararu, A. C. Rosenzweig and U. Ryde, *Angew. Chem., Int. Ed.*, 2018, **57**, 162–166.
- 9 S. T. Prigge, B. A. Eipper, R. E. Mains and L. M. MAMzel, *Science*, 2004, **304**, 864–867.
- 10 S. T. Prigge, A. S. Kolhekar, B. A. Eipper, R. E. Mains and M. Amzel, *Nat. Struct. Biol.*, 1999, **6**, 976–983.
- 11 C. M. Wilmot, J. Hajdu, M. J. McPherson, P. F. Knowles and S. E. Phillips, *Science*, 1999, **286**, 1724–1728.
- 12 T. V. Vendelboe, P. Harris, Y. Zhao, T. S. Walter, K. Harlos, K. El Omari and H. E. M. Christensen, *Sci. Adv.*, 2016, **2**, e1500980.
- 13 J. Y. Lee and K. D. Karlin, *Curr. Opin. Chem. Biol.*, 2015, **25**, 184–193.
- 14 A. Changela, K. Chen, Y. Xue, J. Holschen, C. E. Outten, T. V. OHalloran and A. Mondragon, *Science*, 2003, **301**, 1383–1387.
- 15 A. K. Wernimont, D. L. Huffman, A. L. Lamb, T. V. OHalloran and A. C. Rosenzweig, *Nat. Struct. Biol.*, 2000, **7**, 766–771.
- 16 A. K. Boal and A. C. Rosenzweig, *Chem. Rev.*, 2009, **109**, 4760–4779.
- 17 L. Banci, I. Bertini, F. Cantini, I. C. Felli, L. Gonnelli, N. Hadjiladis, R. Pierattelli, A. Rosato and P. Voulgaris, *Nat. Chem. Biol.*, 2006, **2**, 367–368.
- 18 P. G. Holder, L. C. Jones, P. M. Drake, R. M. Barfield, S. Banas, G. W. de Hart, J. Baker and D. Rabuka, *J. Biol. Chem.*, 2015, **290**, 15730–15745.
- 19 M. Knop, P. Engi, R. Lemnar and F. P. Seebeck, *ChemBioChem*, 2015, **16**, 2147–2150.
- 20 T. Dierks, B. Schmidt, L. V. Borissenko, J. H. Peng, A. Preusser, M. Mariappan and K. von Figura, *Cell*, 2003, **113**, 435–444.
- 21 M. P. Cosma, S. Pepe, I. Annunziata, R. F. Newbold, M. Grompe, G. Parenti and A. Ballabio, *Cell*, 2003, **113**, 445–456.
- 22 I. S. Carrico, B. L. Carlson and C. R. Bertozzi, *Nat. Chem. Biol.*, 2007, **3**, 321–322.
- 23 M. J. Appel and C. R. Bertozzi, *ACS Chem. Biol.*, 2015, **10**, 72–84.
- 24 T. Krüger, T. Dierks and N. Sewald, *Biol. Chem.*, 2019, **4000**, 289–297.
- 25 M. Knop, T. Q. Dang, G. Jeschke and F. P. Seebeck, *ChemBioChem*, 2017, **18**, 161–165.
- 26 M. Meury, M. Knop and F. P. Seebeck, *Angew. Chem., Int. Ed.*, 2017, **56**, 8115–8119.
- 27 M. J. Appel, K. K. Meier, J. Lafrance-Vanasse, H. Lim, C. L. Tsai, B. Hedman, K. O. Hodgson, J. A. Tainer, E. I. Solomon and C. R. Bertozzi, *Proc. Natl. Acad. Sci. U. S. A.*, 2019, **116**(12), 5370–5375.
- 28 M. Jaskolski, M. Gilski, Z. Dauter and A. Wlodawer, *Acta Crystallogr., Sect. D: Biol. Crystallogr.*, 2007, **63**, 611–620.
- 29 T. Dierks, A. Dickmanns, A. Preusser-Kunze, B. Schmidt, M. Mariappan, K. von Figura, R. Ficner and M. G. Rudolph, *Cell*, 2005, **121**, 541–552.
- 30 B. L. Carlson, E. R. Ballister, E. Skordalakes, D. S. King, M. A. Breidenbach, S. A. Gilmore, J. M. Berger and C. R. Bertozzi, *J. Biol. Chem.*, 2008, **283**, 20117–20125.
- 31 A. J. Doig, M. W. MacArthur, B. J. Stapley and J. M. Thornton, *Protein Sci.*, 1997, **6**, 147–155.
- 32 D. Roeser, A. Preusser-Kunze, B. Schmidt, K. Gasow, J. G. Wittmann, T. Dierks, K. von Figura and M. G. Rudolph, *Proc. Natl. Acad. Sci. U. S. A.*, 2006, **103**, 81–86.
- 33 M. Knop, R. Lemnar and F. P. Seebeck, *ChemBioChem*, 2017, **18**, 1755–1761.
- 34 E. L. Smith, J. P. Giddens, A. T. Iavarone, K. Godula, L. X. Wang and C. R. Bertozzi, *Bioconjugate Chem.*, 2014, **25**, 788–795.
- 35 J. H. Peng, S. Alam, K. Radhakrishnan, M. Mariappan, M. G. Rudolph, C. May, T. Dierks, K. von Figura and B. Schmidt, *FEBS J.*, 2015, **282**, 3262–3274.
- 36 T. Krüger, S. Weiland, G. Falck, M. Gerlach, M. Boschanski, S. Alam, K. M. Müller, T. Dierks and N. Sewald, *Angew. Chem., Int. Ed.*, 2018, **57**, 7245–7249.
- 37 T. Urbic and K. A. Dill, *J. Am. Chem. Soc.*, 2018, **140**, 17106–17113.
- 38 A. Biela, N. N. Nasief, M. Betz, A. Heine, D. Hangauer and G. Klebe, *Angew. Chem., Int. Ed.*, 2013, **52**, 1822–1828.
- 39 M. C. Chervenak and E. J. Toone, *J. Am. Chem. Soc.*, 1994, **116**, 10533–10539.
- 40 P. W. Snyder, J. Mecinovic, D. T. Moustakas, S. W. r. Thomas, M. Harder, E. T. Mack, M. R. Lockett, A. Héroux, W. Sherman and G. M. Whitesides, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, **108**, 17889–17894.
- 41 F. Biedermann, W. M. Nau and H. J. Schneider, *Angew. Chem., Int. Ed.*, 2014, **53**, 11158–11171.
- 42 P. C. Roehm and J. M. Berg, *J. Am. Chem. Soc.*, 1998, **120**, 13083–13087.
- 43 S. J. Chiou, C. G. Riordan and A. L. Rheingold, *Proc. Natl. Acad. Sci. U. S. A.*, 2003, **1000**, 3695–3700.
- 44 A. Dey, T. A. Okamura, N. Ueyama, B. Hedman, K. O. Hodgson and E. I. Solomon, *J. Am. Chem. Soc.*, 2005, **127**, 12046–12053.
- 45 X. Yang, S. Niu, T. Ichiye and L. S. Wang, *J. Am. Chem. Soc.*, 2004, **126**, 15790–15794.
- 46 D. Roeser, B. Schmidt, A. Preusser-Kunze and M. G. Rudolph, *Acta Crystallogr., Sect. D: Biol. Crystallogr.*, 2007, **63**, 621–627.



- 47 K. Pervushin, R. Riek, G. Wider and K. Wüthrich, *Proc. Natl. Acad. Sci. U. S. A.*, 1997, **94**, 12366–12371.
- 48 S. Grzesiek and A. Bax, *J. Magn. Reson.*, 1992, **96**, 432–440.
- 49 W. Sicinska, W. M. Westler and H. F. DeLuca, *Proteins*, 2005, **61**, 461–467.
- 50 E. T. Mollova, D. E. Metzler, A. Kintanar, H. Kagamiyama, H. Hayashi, K. Hirotsu and I. Miyahara, *Biochemistry*, 1997, **36**, 615–625.
- 51 S. M. Adam, G. B. Wijeratne, P. J. Rogler, D. E. Diaz, D. A. Quist, J. J. Liu and K. D. Karlin, *Chem. Rev.*, 2018, **118**, 10840–11022.
- 52 L. Ciano, G. J. Davies, W. B. Tolman and P. H. Walton, *Nat. Catal.*, 2018, **1**, 571–577.
- 53 E. M. Shepard, K. M. Okonski and D. M. Dooley, *Biochemistry*, 2008, **47**, 13907–13920.
- 54 S. A. Mills, Y. Goto, Q. Su, J. Plastino and J. P. Klinman, *Biochemistry*, 2002, **41**, 10577–10584.
- 55 M. Bhadra, J. Y. C. Lee, R. E. Cowley, S. Kim, M. A. Siegler, E. I. Solomon and K. D. Karlin, *J. Am. Chem. Soc.*, 2018, **140**, 9042–9045.
- 56 S. I. Mann, T. Heinisch, T. R. Ward and A. S. Borovik, *J. Am. Chem. Soc.*, 2017, **139**, 17289–17292.
- 57 S. Hong, Y.-M. Lee, K. Ray and W. Nam, *Coord. Chem. Rev.*, 2017, **334**, 25–42.
- 58 N. Lopez, D. J. Graham, R. J. McGuire, G. E. Alliger, Y. Shao-Horn, C. C. Cummins and D. G. Nocera, *Science*, 2012, **335**, 450–453.
- 59 E. W. Dahl, J. J. Kiernicki, M. Zeller and N. K. Szymczak, *J. Am. Chem. Soc.*, 2018, **140**, 10075–10079.

