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## Stable radical *versus* reversible $\sigma$ -bond formation of (porphyrinyl)dicyanomethyl radicals†

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(Porphyrinyl)dicyanomethyl radicals were produced by oxidation of dicyanomethyl-substituted porphyrins with  $\text{PbO}_2$ . These radicals constitute a rare example displaying stable radical *versus* dynamic covalent chemistry (DCC) depending upon the substitution position of the dicyanomethyl radical. *meso*-Dicyanomethyl-substituted radicals exist as stable monomeric species and do not undergo any dimerization processes either in the solid state or in solution. In contrast, *β*-dicyanomethyl-substituted radicals are isolated as  $\sigma$ -dimers that are stable in the solid-state but display reversible  $\sigma$ -dimerization behavior in solution; monomeric radical species exist predominantly at high temperatures, while  $\sigma$ -dimerization is favoured at low temperatures. This dynamic behaviour has been confirmed by variable-temperature  $^1\text{H}$  NMR, UV-vis and EPR measurements. The structures of the stable radical and  $\sigma$ -dimer have been revealed by single-crystal X-ray diffraction analysis. The observed different reactivities of the two (porphyrinyl)dicyanomethyl radicals have been rationalized in terms of their spin delocalization behaviours.

### Introduction

Radical-based dynamic covalent chemistry (DCC) has continued to attract considerable attention in recent years due to its intriguing nature to form stable molecular self-assemblies *via* a clean radical–radical coupling process.<sup>1–3</sup> In particular, the stimuli-responsive nature of radicals to form  $\sigma$ -bonds in a reversible manner has been used as an attractive tool for various applications.<sup>2</sup> In recent studies, a number of dicyanomethyl-substituted aromatic compounds were reported to show such reversible self-dimerization-dissociation processes.<sup>3</sup> Among them, the study by Seki and co-workers demonstrated that *p*-diphenylamino-substituted dicyanomethyl radical **1** underwent clean DCC to form its  $\sigma$ -dimer (**1**)<sub>2</sub> in a reversible manner.<sup>3a</sup> In addition to this, they have also reported reversible formation of the  $\pi$ -dimer of the more stable aryl-dicyanomethyl radical **2** possessing a fused nitrogen atom (Chart 1).<sup>3d</sup> Recently, Chi and co-workers have reported three-dimensional DCC of aryl-dicyanomethyl radicals, which resulted in the construction of a large cyclic hexamer, whose structure was elucidated by single-crystal X-ray analysis.<sup>3e</sup> Further, Winter and co-workers described the effect of the structure–activity relationship on aryl-dicyanomethyl radicals.<sup>3g</sup>

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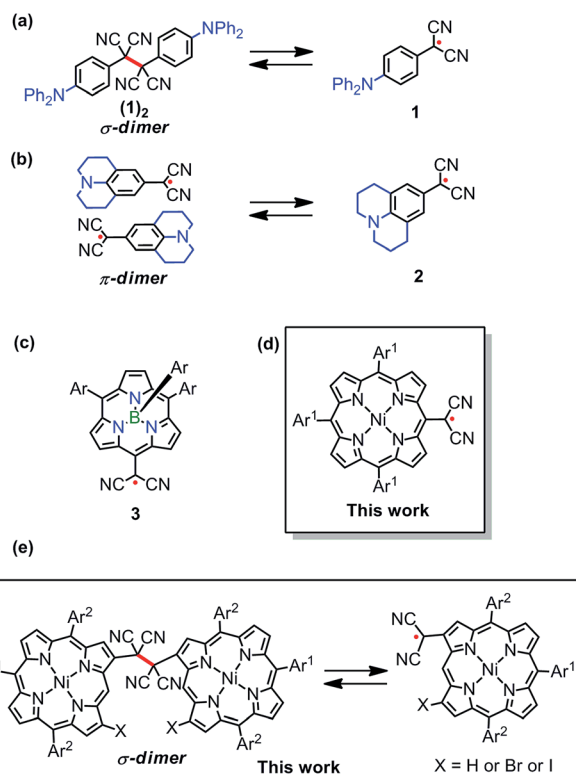


Chart 1 (a–c) Reported dicyanomethyl based radicals. (d and e) Porphyrin based stable radical vs. reversible  $\sigma$ -dimer (DCC). Ar = 4-methylphenyl, Ar<sup>1</sup> = 3,5-di-*tert*-butylphenyl and Ar<sup>2</sup> = mesityl.

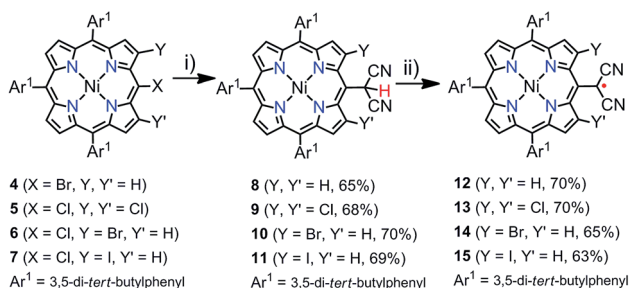


Following these results, we have reported that (*meso*-subporphyrinyl)dicyanomethyl radical **3** is quite stable and does not undergo any dimerization.<sup>4</sup> As an extension, we examined the chemistry of (porphyrinyl)dicyanomethyl radicals, since porphyrins have been demonstrated to be a nice platform to stabilize neighboring radicals owing to effective spin-delocalization.<sup>5,6</sup> While various types of porphyrin-based stable radicals have been explored, those undergoing reversible  $\sigma$ -dimerization are still quite rare,<sup>7</sup> except the important examples.<sup>8</sup> In the present work, (*meso*- and  $\beta$ -porphyrinyl)dicyanomethyl radicals were examined and displayed contrasting behaviors in solutions. (Porphyrinyl)dicyanomethyl radicals constitute a rare example displaying stable radical *versus* dynamic covalent chemistry (DCC) depending upon the substitution position of the dicyanomethyl radical. In particular, (*meso*-porphyrinyl)dicyanomethyl radicals do not form any  $\sigma$ - or  $\pi$ -dimers and exist predominantly as a monomer radical in solution, while ( $\beta$ -porphyrinyl)dicyanomethyl radicals show DCC by forming  $\sigma$ -dimers in a reversible manner. These results contrast sharply with the previous reports that oxidation of 5,10- and 5,15-dicyanomethyl-substituted porphyrins gave closed shell quinonoid compounds.<sup>9</sup>

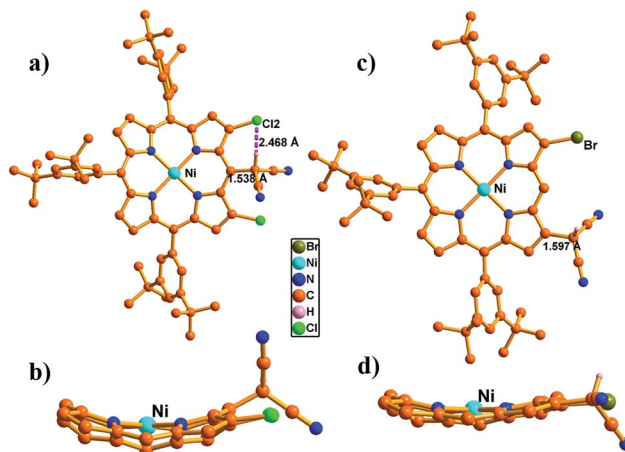
## Results and discussion

### (*meso*-Porphyrinyl)dicyanomethyl radicals

In the synthesis of *meso*-dicyanomethylporphyrins, we employed a simple nucleophilic aromatic substitution ( $S_NAr$ ) of *meso*-halogenated porphyrins<sup>4</sup> **4–7** with malononitrile in the presence of NaH in 1,3-dimethyl-2-imidazolidinone (DMI) at 60 °C, which produced the desired products **8–11** in 65–70% yields as stable compounds (Scheme 1). Notably, these  $S_NAr$  reactions proceeded nicely even for sterically hindered substrates **5**, **6** and **7**. The <sup>1</sup>H NMR spectrum of **8** shows a singlet at 7.01 ppm due to the dicyanomethyl group (Fig. S17<sup>†</sup>), while peaks of **9–11** are observed in the range of 7.35–8.00 ppm (Fig. S19, S21 and S23<sup>†</sup>). These trends are comparable with the previously reported subporphyrin results,<sup>4</sup> and have been ascribed to effective hydrogen bonding interaction between the methine proton in the dicyanomethyl group and the  $\beta$ -halogens. The structure of **9** was confirmed by X-ray crystallographic analysis (Fig. 1a), which revealed the distance between the methine hydrogen and  $\beta$ -chlorine atom to be 2.468(1) Å,



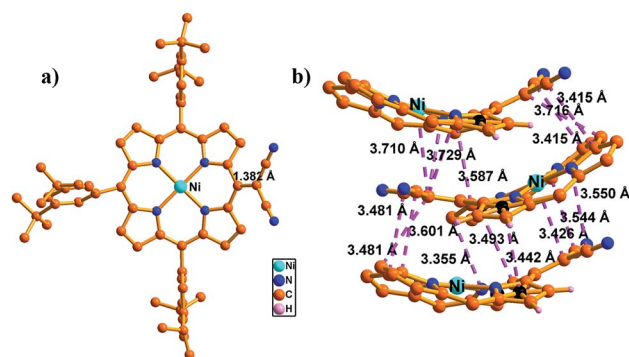
**Scheme 1** Synthesis of **8–15**. (i) 10 equiv. malononitrile, 10 equiv. NaH, DMI, 60 °C, 1 h. (ii) 200 equiv. PbO<sub>2</sub>, CH<sub>2</sub>Cl<sub>2</sub>, rt, 2 h.



**Fig. 1** Single crystal X-ray structure of **9** and **21**. (a and c) Top view and (b and d) side view. The *meso*-aryl groups in (b and d) and hydrogen atoms in (a–d) are omitted for clarity.

suggesting intramolecular hydrogen bonding interaction. The *meso*-dicyanomethyl group is largely tilted by 52.17(9)° from the mean plane of the Ni(II) porphyrin.

Oxidation of *meso*-dicyanomethyl porphyrins **8–11** with an excess amount of PbO<sub>2</sub> in CH<sub>2</sub>Cl<sub>2</sub> gave the corresponding dicyanomethyl radicals **12–15** in 63–70% yields (Scheme 1). All these radicals were stable and tolerant towards moisture and air and separable by silica-gel column chromatography. These radicals were NMR silent even at low temperatures and EPR active (Fig. S45<sup>†</sup>). The structure of radical **12** was revealed by single crystal X-ray analysis (Fig. 2). In the unit cell, three radicals of slightly different structures formed an offset stacking aggregate with interporphyrin separations of *ca.* 3.5 and 3.7 Å. Therefore, the structure of **12** was studied based on averaged values. The dicyanomethyl radical group was tilted towards the porphyrin plane with a small dihedral angle of 38.08° and a C–C bond distance of 1.395 Å, which is shorter as compared to the regular C–C single bond (1.45 Å), indicating a double bond character and effective spin delocalization over the porphyrin ring. These structural features are comparable with those of the



**Fig. 2** Single crystal X-ray structure of **12**. (a) Top view and (b) close interaction between radical pairs. The *meso*-aryl groups in (b) and the peripheral hydrogen atoms in (a and b) have been omitted for clarity.



previously reported subporphyrins and other dicyanomethyl derivatives.<sup>3d,g,4</sup>

The absorption spectra of **8–11** in CH<sub>2</sub>Cl<sub>2</sub> showed Soret-like bands around 420 nm and Q-like bands between 550 and 600 nm as typical absorption features of porphyrin (Fig. S47–S50†). The absorption spectra of radicals **12–15** in CH<sub>2</sub>Cl<sub>2</sub> showed red shifted Soret-like bands at around 470 nm and Q-like bands between 600 and 900 nm along with a radical-characteristic broad absorption band extending to 1600 nm (Fig. S47–S50†). Importantly, there were no significant changes of the absorption spectra of **12–15** upon temperature change in line with the discussion above (Fig. S51–S55†).

Further, the magnetic properties of radicals **12–15** in the solid state were examined by SQUID magnetometry (Fig. S46†). The observed  $\chi T$ - $T$  curves clearly indicated intermolecular weak anti-ferromagnetic interactions, which were reproduced with a modified Bleaney–Bowers model. The best simulation plots of **12–15** were obtained with parameters of  $(f_1, f_2, J_1/k_B) = (0.713, 0.105, -4.85 \text{ K})$ ,  $(0.225, 0.313, -33.7 \text{ K})$ ,  $(0.170, 0.374, -29.1 \text{ K})$  and  $(0.141, 0.397, -24.7 \text{ K})$ , respectively.<sup>10</sup> Collectively, the dicyanomethyl radicals **12–15** are concluded to exist as a monomeric stable species in solution and in the solid-state.

### ( $\beta$ -Porphyrinyl)dicyanomethyl radicals

Our initial attempt to produce simple  $\beta$ -dicyanomethyl-substituted porphyrin **20** by nucleophilic aromatic substitution of  $\beta$ -bromoporphyrin **16** did not proceed well, while the same reaction proceeded nicely for  $\beta,\beta$ -dihalogenated porphyrins **17–19** with malononitrile in the presence of NaH in DMI at 100 °C to afford **21–23** in moderate yields (Scheme 2). The simple  $\beta$ -dicyanomethyl porphyrin **20** was prepared in 74% yield by a Pd-catalyzed coupling reaction of **16** with malononitrile in the presence of NaH in 1,2-dimethoxyethane (DME) at 80 °C (Scheme 2). The structure of **21** was confirmed by single crystal X-ray analysis (Fig. 1c), which revealed the bond length between the dicyanomethyl unit and porphyrin to be 1.597(5) Å and the tilt angle of the dicyanomethyl unit from the mean plane of the porphyrin to be 49.83(2)°.

Oxidation of **20–23** proceeded smoothly with an excess amount of PbO<sub>2</sub> in CH<sub>2</sub>Cl<sub>2</sub> to give dimers (**24**)<sub>2</sub>–(**27**)<sub>2</sub> as stable compounds. These dimers were separated by silica gel column chromatography in 50–71% yields. The molecular ion peaks of (**24**)<sub>2</sub>–(**27**)<sub>2</sub> were not detected but the half values were observed at  $m/z = 993.5096, 1073.4174, 1119.4040$  and  $979.2468 \text{ (M}^+)$ ,

respectively, matching with the expected radical compositions (Fig. S13–S16†).

Dimers (**24**)<sub>2</sub>–(**27**)<sub>2</sub> showed common temperature-dependent <sup>1</sup>H NMR spectral features. As a typical example, the <sup>1</sup>H NMR spectra of (**27**)<sub>2</sub> at various temperatures are shown in Fig. 3a. The spectrum was rather broad at 298 K but became increasingly sharper upon decreasing the temperature, which, judging from the previous related studies,<sup>3</sup> has been interpreted in terms of DCC between the monomeric radical **27** and its  $\sigma$ -dimer (**27**)<sub>2</sub>. At high temperatures a certain amount of monomeric radical exists but the  $\sigma$ -dimer dominates at 193 K. We succeeded in the X-ray structural determination of (**27**)<sub>2</sub>. Fig. 3b and c show the structure of (**27**)<sub>2</sub> which is a covalently linked *syn*-dimer, in which the formed C–C bond length is very long,

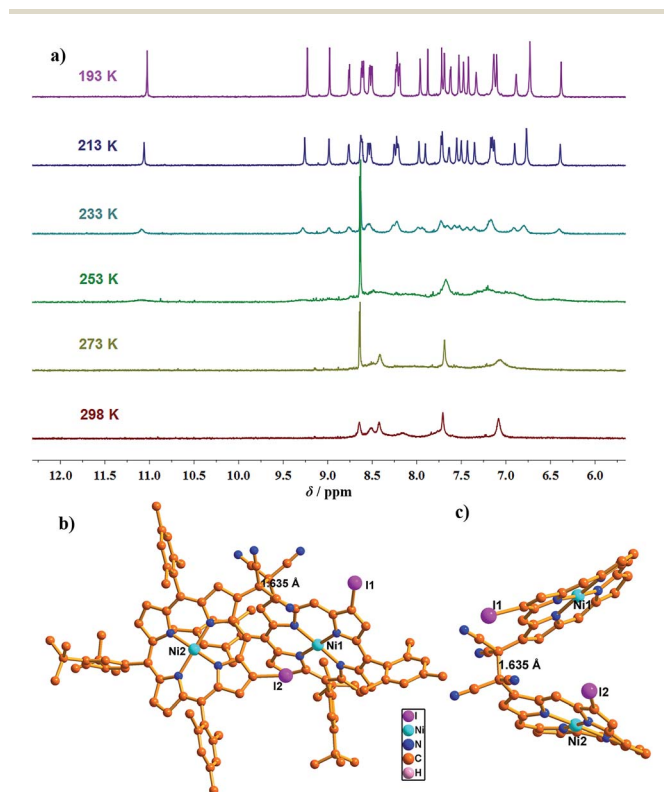
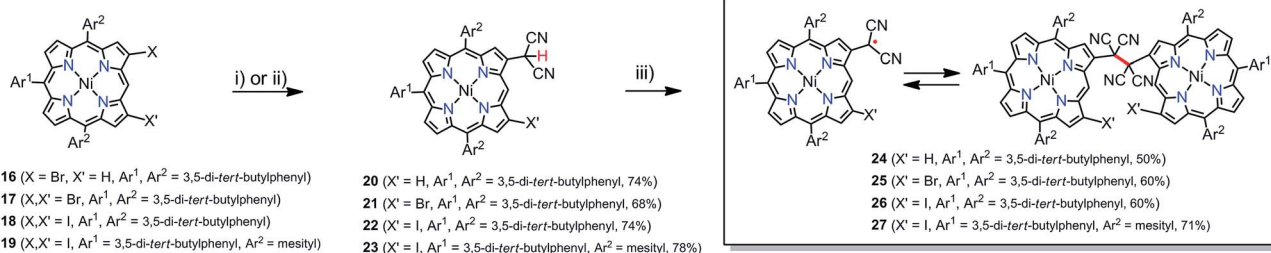


Fig. 3 (a) <sup>1</sup>H-NMR spectrum of (**27**)<sub>2</sub> in CD<sub>2</sub>Cl<sub>2</sub> at various temperatures; (b and c) X-ray crystal structure of (**27**)<sub>2</sub>. *meso*-Aryl groups in (c) and hydrogen atoms in (b and c) have been omitted for clarity.



Scheme 2 Synthesis of **20–27**. For synthesis of **20**, (i) 5 equiv. malononitrile, 8 equiv. NaH, DME, Pd(PPh<sub>3</sub>)<sub>4</sub> (20 mol%), 80 °C, 2 h. For synthesis of **21–23** (ii) 10 equiv. malononitrile, 10 equiv. NaH, DMI, 100 °C, 6 h. (iii) 200 equiv. PbO<sub>2</sub>, CH<sub>2</sub>Cl<sub>2</sub>, rt, 2 h.





1.635(1) Å, which is marginally longer than that of the reported examples undergoing DCC.<sup>3a,e</sup>

In line with DCC, (24)<sub>2</sub>–(27)<sub>2</sub> showed reversible temperature dependent absorption spectra. As the representative example, the variable temperature absorption spectra of (27)<sub>2</sub> in toluene are shown in Fig. 4. Upon increasing the temperature, the typical absorption bands of normal porphyrins, *i.e.* the Soret band at 414 nm and Q-bands at 535 and 571 nm, decreased and broad bands around 490 and 672 nm as well as a very broad tail in the NIR range of 1000–1400 nm emerged. These spectral changes have been interpreted to indicate that (27)<sub>2</sub> exists predominantly at room temperature and undergoes homolysis to give persistent dicyanomethyl radical 27 at high temperatures. This interpretation was strongly reinforced by variable temperature EPR measurement, which actually showed increased EPR signal intensity at higher temperatures (Fig. 4, inset). Similar variable temperature absorption spectra and EPR spectra were observed for (24)<sub>2</sub> (Fig. S60†), (25)<sub>2</sub> (Fig. S61†) and (26)<sub>2</sub> (Fig. S62†). On the basis of these experiments, the thermodynamic parameters of DCC were determined (Table 1). Overall, the VT-NMR, VT-UV, and VT-EPR studies clearly supported the conclusion that the products exist in an equilibrium state between the  $\sigma$ -dimers and radicals.

### Cyclic voltammograms and DFT calculations

To gain further insight into the electronic properties, electrochemical analysis was accomplished by cyclic voltammetry in CH<sub>2</sub>Cl<sub>2</sub> solution, and the redox potentials for 12–15 are

summarized in Table 2. The stable neutral radicals 12–15 exhibit two reversible oxidation waves and two reversible reduction waves. The energy gap between the first oxidation and reduction potentials of compounds 12–15 was very small, characteristic of radical species. We also examined the electrochemical properties of the  $\sigma$ -dimers (24)<sub>2</sub>–(27)<sub>2</sub>. Cyclic voltammetry showed two similar reversible oxidation waves and two reversible reduction waves and the electrochemical redox gaps were determined to be 0.83, 0.86, 0.85, and 0.85 eV for (24)<sub>2</sub>–(27)<sub>2</sub>, respectively. Because the redox gaps of (24)<sub>2</sub>–(27)<sub>2</sub> (<1 eV) were too small considering their optical HOMO–LUMO gaps (~2 eV), the observed electrochemically active species were determined to be monomeric radicals 24–27 that were in equilibrium with the corresponding dimers.

To investigate the spin density distribution, spin-unrestricted density functional theory (DFT) calculations were conducted at the UB3LYP/6-311g(d) level (Fig. S75–S78†).<sup>11</sup> As representative examples, the spin density maps of 12 and 24 are shown in Fig. 5. For *meso*-dicyanomethyl radicals 12–15, the spin density maps clearly show that the spin is fully delocalized over the whole  $\pi$ -network of porphyrin backbones (Fig. S73†). On the other hand, the spin density in  $\beta$ -dicyanomethyl radicals 24–27 is less delocalized and large spin densities are found at the dicyanomethyl unit and neighbouring  $\beta$ -carbon. Spin

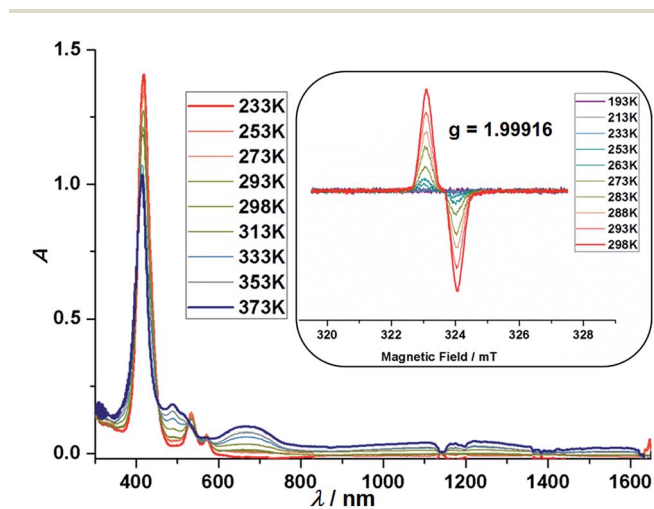


Fig. 4 Absorption spectra of (27)<sub>2</sub> in toluene at various temperatures. Inset: EPR spectra of (27)<sub>2</sub> in toluene at various temperatures.

Table 1 Thermodynamic parameters of (24)<sub>2</sub>–(27)<sub>2</sub>

	$\Delta H_{\text{dim}}/\text{kJ mol}^{-1}$	$\Delta S_{\text{dim}}/\text{J K}^{-1} \text{mol}^{-1}$	$\Delta G_{\text{dim},298 \text{ K}}/\text{kJ mol}^{-1}$
24	–55.7	–77.8	–32.5
25	–60.6	–89.5	–33.9
26	–66.3	–112.2	–32.9
27	–60.6	–91.7	–33.3

Table 2 Electrochemical potentials of 12–15 and (24)<sub>2</sub>–(27)<sub>2</sub>

	$E_{\text{OX1}}^{1/2}/\text{V}$	$E_{\text{OX2}}^{1/2}/\text{V}$	$E_{\text{RED1}}^{1/2}/\text{V}$	$E_{\text{RED2}}^{1/2}/\text{V}$	$\Delta E_{\text{HL}}/\text{eV}$
12	0.26	1.21	–0.33	–2.05	0.59
13	0.31	1.21	–0.29	–1.95	0.60
14	0.29	1.22	–0.31	–2.02	0.60
15	0.28	1.20	–0.30	–1.95	0.58
(24) <sub>2</sub>	0.50	1.29	–0.33	–1.99	0.83
(25) <sub>2</sub>	0.55	1.33	–0.31	–1.88	0.86
(26) <sub>2</sub>	0.55	1.37	–0.30	–1.88	0.85
(27) <sub>2</sub>	0.54	1.39	–0.31	–1.96	0.85

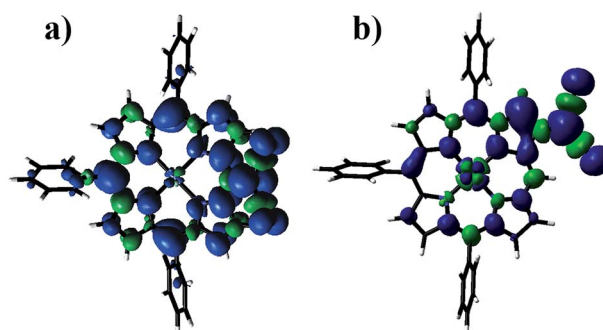


Fig. 5 Spin density distribution plots of (a) 12 and (b) 24 calculated at the UB3LYP/6-311G\* level (isovalue: 0.001). Consistent with the electrochemical data, the energy gaps between the SOMO( $\alpha$ ) and SOMO( $\beta$ ) of 12–15 (1.14–1.17 eV) are distinctly smaller than those of 24–27 (1.35–1.48 eV). Further TD-DFT calculations were carried out to interpret the absorption features of 12–15 (Fig. S86–S89†). The calculated absorption transitions were consistent with the observed data, which were similar to the previously reported subporphyrin radicals.



densities at the dicyanomethyl center clearly indicate a difference depending on the substitution position: 0.19–0.29 for **12–15** and 0.47–0.49 for **24–27**. For **12–15**, the spin density at the dicyanomethyl center largely depends on the neighboring  $\beta$ -substituent(s): 0.289 (**12**) > 0.242 (**14**)  $\sim$  0.234 (**15**) > 0.191 (**13**). This trend is related to the average tilt angles of the dicyanomethyl group toward the porphyrin plane; **12** ( $32.3^\circ$ ) < **14** ( $39.1^\circ$ )  $\sim$  **15** ( $39.7^\circ$ ) < **13** ( $44.7^\circ$ ) (Fig. S85†). We consider that the less delocalized spin nature makes **24–27** suitable for DCC.

## Conclusions

In summary, we synthesized (*meso*-porphyrinyl)dicyanomethyl radicals and ( $\beta$ -porphyrinyl)dicyanomethyl radicals by a  $S_NAr$  reaction and Pd-catalyzed coupling reaction of the corresponding porphyrin halides with malononitrile and subsequent oxidation with  $PbO_2$ . (*meso*-Porphyrinyl)dicyanomethyl radicals are very stable, exist as monomeric species and do not undergo any recombination in solution or in the solid state. On the other hand, ( $\beta$ -porphyrinyl)dicyanomethyl radicals are isolated as covalent  $\sigma$ -dimers, which are stable in the solid state but display DCC behavior in solution; namely, they undergo homolysis to form a persistent dicyanomethyl radical at high temperatures. This marked difference can be ascribed to better spin delocalization in (*meso*-porphyrinyl)dicyanomethyl radicals as compared with ( $\beta$ -porphyrinyl)dicyanomethyl radicals. Further studies on (porphyrinyl)dicyanomethyl radicals are actively ongoing in our laboratories.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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## Notes and references

- (a) H. Yokoi, S. Hiroto and H. Shinokubo, *J. Am. Chem. Soc.*, 2018, **140**, 4649; (b) S. N. Spisak, A. V. Zabula, M. Alkan, A. S. Filatov, A. Y. Rogachev and M. A. Petrukhina, *Angew. Chem., Int. Ed.*, 2018, **57**, 6171; (c) R. J. Wojtecki, M. A. Meador and S. J. Rowan, *Nat. Mater.*, 2010, **10**, 14; (d) Y. Wang, X. Tan, Y.-M. Zhang, S. Zhu, I. Zhang, B. Yu, K. Wang, B. Yang, M. Li, B. Zou and S. X.-A. Zhang, *J. Am. Chem. Soc.*, 2015, **137**, 931.
- (a) Y. Kishimoto and J. Abe, *J. Am. Chem. Soc.*, 2009, **131**, 4227; (b) K. Imato, J. C. Natterodt, J. Sapkota, R. Goseki, C. Weder, A. Takahara and H. Otsuka, *Polym. Chem.*, 2017, **8**, 2115; (c) K. Imato, A. Irie, T. Kosuge, T. Ohishi, M. Nishihara, A. Takahara and H. Otsuka, *Angew. Chem., Int. Ed.*, 2015, **54**, 6168; (d) K. Imato, M. Nishihara, T. Kanehara, Y. Amamoto, A. Takahara and H. Otsuka, *Angew. Chem., Int. Ed.*, 2012, **51**, 1138.
- (a) H. D. Hartzler, *J. Org. Chem.*, 1966, **31**, 2654; (b) T. Kobashi, D. Sakamaki and S. Seki, *Angew. Chem., Int. Ed.*, 2016, **55**, 8634; (c) J. P. Peterson, M. R. Geraskina, R. Zhang and A. H. Winter, *J. Org. Chem.*, 2017, **82**, 6497; (d) K. Okino, S. Hira, Y. Inoue, D. Sakamaki and S. Seki, *Angew. Chem., Int. Ed.*, 2017, **56**, 16597; (e) L. Yuan, Y. Han, T. Tao, H. Phan and C. Chi, *Angew. Chem., Int. Ed.*, 2018, **57**, 9023; (f) D. Wang, C. C. Ferron, J. Li, S. Gamez-Valenzuela, R. P. Ortiz, J. T. L. Navarrete, V. H. Jolin, X. Yang, M. P. Alvarez, V. G. Baonza, F. Hartl, M. C. R. Delgado and H. Li, *Chem.-Eur. J.*, 2017, **23**, 13776; (g) R. Zhang, J. P. Peterson, L. J. Fischer, A. Ellern and A. H. Winter, *J. Am. Chem. Soc.*, 2018, **140**, 14308.
- B. Adinarayana, D. Shimizu and A. Osuka, *Chem.-Eur. J.*, 2019, **25**, 1706.
- D. Shimizu and A. Osuka, *Chem. Sci.*, 2018, **9**, 1408.
- (a) T. Koide, G. Kashiwazaki, M. Suzuki, K. Furukawa, M.-C. Yoon, S. Cho, D. Kim and A. Osuka, *Angew. Chem., Int. Ed.*, 2008, **47**, 9661; (b) D. Shimizu, J. Oh, K. Furukawa, D. Kim and A. Osuka, *Angew. Chem., Int. Ed.*, 2015, **54**, 6613; (c) D. Shimizu, J. Oh, K. Furukawa, D. Kim and A. Osuka, *J. Am. Chem. Soc.*, 2015, **137**, 15584; (d) K. Kato, W. Cha, J. Oh, K. Furukawa, H. Yorimitsu, D. Kim and A. Osuka, *Angew. Chem., Int. Ed.*, 2016, **55**, 8711; (e) D. Shimizu, K. Furukawa and A. Osuka, *Angew. Chem., Int. Ed.*, 2017, **56**, 7435; (f) K. Kato, K. Furukawa, T. Mori and A. Osuka, *Chem.-Eur. J.*, 2018, **24**, 572; (g) D. Shimizu and A. Osuka, *Angew. Chem., Int. Ed.*, 2018, **57**, 3733; (h) K. Kato, K. Furukawa and A. Osuka, *Angew. Chem., Int. Ed.*, 2018, **57**, 9491.
- Only one report which related to norcorrole system. B. Liu, T. Yoshida, X. Li, M. Stępień, H. Shinokubo and P. J. Chmielewski, *Angew. Chem., Int. Ed.*, 2016, **55**, 13142.
- (a) J. H. Fuhrhop, S. Besecke, J. Subramanian, C. Mengersen and D. Riesner, *J. Am. Chem. Soc.*, 1975, **97**, 7141; (b) A. L. Balch, B. C. Noll, S. M. Reid and E. P. Zovinka, *J. Am. Chem. Soc.*, 1993, **115**, 2531.
- (a) I. M. Blake, H. L. Anderson, D. Beljonne, J.-L. Brédas and W. Clegg, *J. Am. Chem. Soc.*, 1998, **120**, 10764; (b) I. M. Blake, L. H. Rees, T. D. W. Claridge and H. L. Anderson, *Angew. Chem., Int. Ed.*, 2000, **39**, 1818; (c) M. J. Smith, W. Clegg, K. A. Nguyen, J. E. Rogers, R. Pachter, P. A. Fleitz and H. L. Anderson, *Chem. Commun.*, 2005, 2433; (d) Q. Chen, Y.-Z. Zhu, Q.-J. Fan, S.-C. Zhang and J.-Y. Zheng, *Org. Lett.*, 2014, **16**, 1590; (e) K.-i. Yamashita, S. Sakamoto, A. Suzuki and K.-i. Sugiura, *Chem.-Asian J.*, 2016, **11**, 1004.
- The dimeric structure was not fixed by strong interaction such as covalent bond, we assumed that some of radicals were not involved in the dimeric packing. Thus, we used  $f_1$  and  $f_2$  to as weights for the dimeric diradical and isolated radicals, respectively.



- 11 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. V. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov, J. L. Sonnenberg, D. Williams-Young, F. Ding, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery Jr, J. E. Peralta, F. Ogliaro, M. J. Bearpark, J. J. Heyd, E. N. Brothers, K. N. Kudin, V. N. Staroverov, T. A. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. P. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman, and D. J. Fox, *Gaussian 16, Revision A.03*, Gaussian, Inc., Wallingford CT, 2016.

