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The opposite and amplifying effect of B ← N coordination on photophysical properties of regioisomers with an unsymmetrical backbone†

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1,3-Dipolar cycloaddition of pyrido[2,1-*a*]isoindole with internal alkynes functionalized by a BMes₂ph and an N-aromatic heterocycle leads to the formation of two types of regioisomers (major *a* and minor *b*) that have distinct physical and photophysical properties. Examination on 5 pairs of regioisomers unveils that the major isomers consistently have a smaller optical energy gap and emission energy than the corresponding minor isomers, which is greatly amplified by the formation of an internal B ← N bond. The regioisomers with a B ← N bond display contrasting temperature-dependent structural dynamics and response to fluoride ions, owing to an entropy-driven or fluoride initiated B ← N bond rupture/ring-opening process and the different B ← N bond strength. The opposite inductive effect and the Lewis pair properties of the dichotomic substituent units are responsible for the contrasting properties of the regioisomers in this system.

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Introduction

Isomerism plays various important roles in chemistry, especially in medicinal and pharmaceutical chemistry.¹ In the areas of optoelectronics and materials science,² isomers with varying substitution patterns also attract great research interest due to their distinct properties, which allows them to be used for fine-tuning molecular functions in both bulk materials and devices. Therefore, understanding and exploiting the influence of isomerism on the physical and photophysical properties of molecules is an important issue in materials science.

Owing to their diverse chemical properties, boron-containing molecules have important applications in materials chemistry, including optoelectronic materials³ and sensory devices.⁴ The utility of isomerism phenomena in organoboron systems has been demonstrated in several contexts previously, including structural change at a boron atom *via* photoisomerization,⁵ varying emission colours based on constitutional isomerism,⁶ and reversible B ← X (X = N or O) bond formation/cleavage where the open and closed forms display significantly different photophysical properties.⁷ The introduction of an internal B ← N bond is also known to be effective in

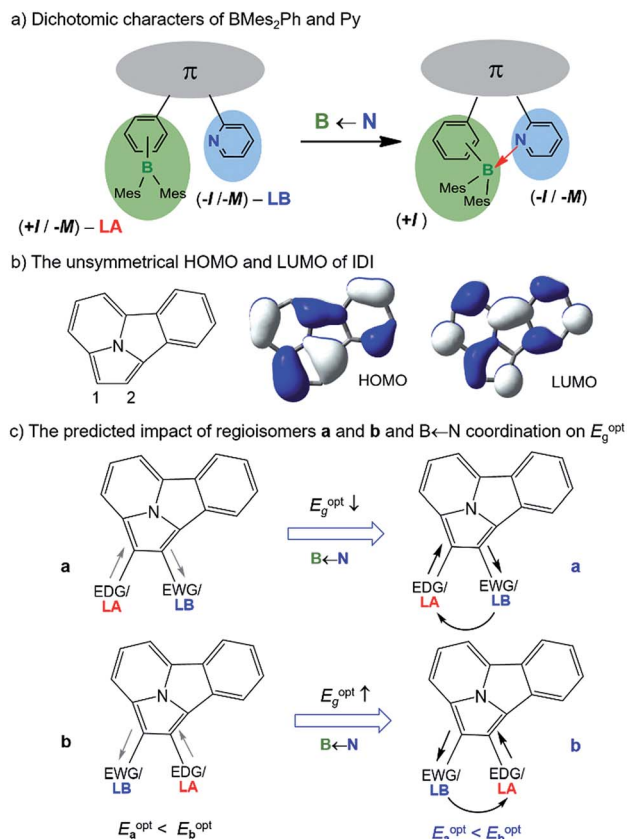
enhancing extended π -conjugation, tuning the electronic properties of molecules, and improving the performance of emissive/charge transport materials in high performance organic devices.^{3b,8} BMes₂ph and N-aromatic units (*e.g.* 2-pyridyl; py) are commonly used functional groups in conjugated materials. These two types of units can be inductively electron-donating (EDG/+I) and electron-withdrawing (EWG/−I), respectively, when bound to an appropriate arene substrate. BMes₂ph and py can also both exert a negative mesomeric effect (−M) on the substrate they are bound to *via* π -conjugation. On the other hand, as illustrated in Scheme 1a, they are also Lewis acid (LA) and Lewis base (LB), respectively, capable of forming a B ← N bond. If we consider the inductive effect only, the dual EDG(+I)-LA and EWG(−I)-LB characters possessed by the BMes₂ph and py, respectively, are seemingly opposing each other, hence such units may be described as electronically “dichotomic” groups. The formation of an internal B ← N bond would reinforce the −I/−M character of the py unit and greatly enhance the +I character of BMes₂Ph. The impact of an internal B ← N bond on the HOMO–LUMO gap/photophysical properties of molecules possessing such dichotomic units would depend on the electron density/orbital distributions of the π -backbone unit. If the substitution sites to the backbone are electron-rich and electron-poor, respectively, and contribute to HOMO and LUMO, respectively, the dichotomic units could have a distinct influence on the photophysical properties of regioisomers *via* both inductive/mesomeric and coordination effects. Despite the appealing features of such units, experimental confirmation on their influence on the properties of regioisomers has not been demonstrated owing to the lack of suitable systems for accessing the required regioisomers.

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Scheme 1 Illustration of impact of dichotomic units on E_g^{opt} . The straight arrows indicate inductive effect and the curved arrow shows the dative bond between the Lewis pair.

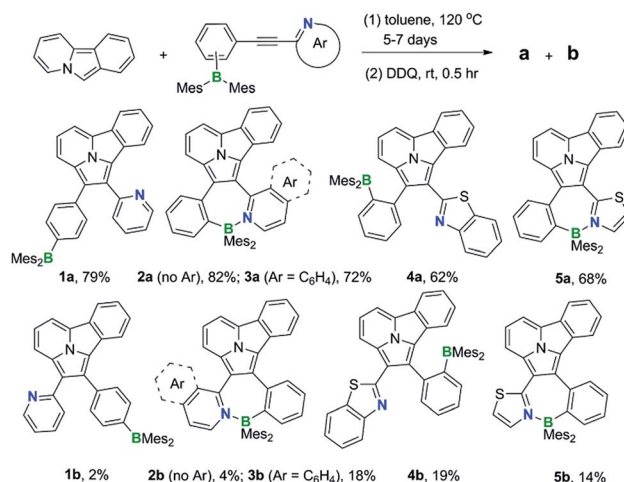
We have discovered that indolizino[3,4,5-*ab*]isoindole (IDI) can serve as the model system for the study of the influence of dichotomic substituents on regioisomers. The IDI molecule is an electron-rich substrate. BMes_2Ph and py attached to IDI are mainly inductively EDG and EWG, respectively.⁹ As shown in Scheme 1b, the C1 and C2 atoms in IDI have the highest and the lowest electron density, respectively, among the carbon atoms (electrostatic potential charge = -0.421 for C1, -0.110 for C2). The HOMO of IDI has a large contribution to C1 but no contribution to C2 while the LUMO contributes to both C1 and C2 (DFT, M06-2X/6-31g(d)). If we consider the inductive and coordination effects alone, two scenarios, **a** and **b** are possible as shown in Scheme 1c. In **a**, the substituents mismatch the electron density distribution of the IDI substrate, thus likely leading to the destabilization and stabilization of HOMO and LUMO, respectively, and a smaller optical energy gap (E_g^{opt}). In **b**, the substituents match the IDI unit, thereby stabilizing the HOMO and destabilizing the LUMO, likely leading to a larger E_g^{opt} . By successfully preparing and isolating 5 pairs of regioisomers of IDI based on BMes_2Ph and N-heterocycle units, we have established the impact of the dichotomic units and the amplification effect of a $\text{B} \leftarrow \text{N}$ bond on the electronic/photophysical properties of regioisomers for the first time. The details are presented herein.

Results and discussion

Synthesis and structures of **1a/1b–5a/5b**

1,3-Dipolar cycloaddition reactions of pyrido[2,1-*a*]isoindole⁹ with internal alkynes decorated by BMes_2ph and N-aromatic units were employed to synthesize the 5 pairs of regioisomers using the general one-pot, two-step procedure shown in Scheme 2. Previously, terminal alkynes containing a BMes_2Ph unit were investigated for this type of reaction, which led to the exclusive formation of species that have the boron unit at the C1 position.^{3b} Consistent with the previous findings and the well-established selection rules for 1,3-dipolar cycloadditions,¹⁰ regioisomers (**a**) were isolated as the dominating products in all reactions, along with the minor isomers (**b**). The overall isolated yields are excellent for all 5 pairs of compounds (81% to 90%). Each pair of regioisomers were separated by either column or preparative thin-layer chromatography (TLC). The low isolated yields of **1b** and **2b** were caused by their band overlap with **1a** and **2a**, respectively, on column and TLC plates, which made it necessary to run multiple column/TLC purifications to isolate the pure products, resulting in the loss of compounds. DFT computational analysis (M06-2X/6-31g(d) with CH_2Cl_2 as the solvent) indicated that isomer (**b**) is thermodynamically more stable than isomer (**a**) by ~ 0.5 to 3.1 kcal mol⁻¹ for all 5 pairs. Therefore, the preferential formation of isomer (**a**) can be attributed to kinetic factors. Attempts to increase the yield of isomer (**b**) by either adding a large excess of NBu_4F (TBAF) to the cycloaddition reaction or using alternative procedures were not successful. The 5 pairs of regioisomers were fully characterized by NMR, HRMS and single-crystal X-ray diffraction analyses.

The crystal structures of **3a**, **3b**, **5a**, and **5b** are shown in Fig. 1 while **4a** and **4b** are shown in Fig. 2. The structures of **1a/2a** and **1b/2b** are provided in the ESI.† For the tetracoordinate compounds, the boron atom is bound to the nitrogen atom of the heterocycle with a long $\text{B} \leftarrow \text{N}$ bond as shown in Table 1. DFT (M06-2X/6-31G(d)) optimized structures gave similar $\text{B} \leftarrow \text{N}$ bond lengths for **2a/2b** and **3a/3b**. For **5a/5b**, however, the DFT calculated $\text{B} \leftarrow \text{N}$ bond length is 1.705 and 1.691 Å,



Scheme 2 Synthetic procedures and structures of **1a/1b–5a/5b**.



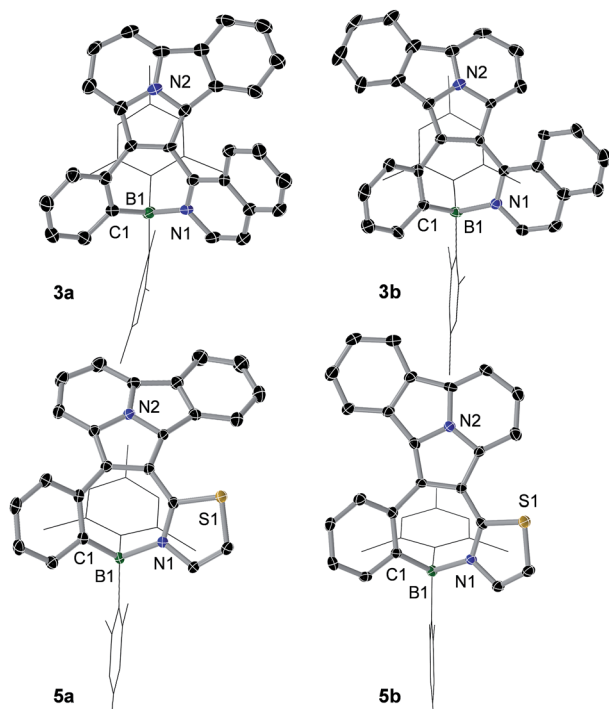


Fig. 1 Crystal structures of **3a**, **3b**, **5a** and **5b** with 35% thermal ellipsoids. For clarity, the Mes groups are shown in wireframe style and H atoms are omitted.

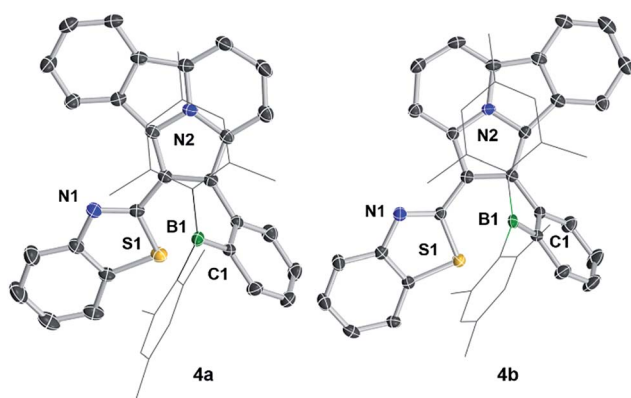


Fig. 2 Crystal structures of **4a** and **4b**. For clarity, the Mes groups are shown in wireframe style and H atoms are omitted.

respectively, which agrees with the trend observed for other pairs but disagrees with the crystal data of **5a/5b**. This discrepancy may be caused by the different intermolecular interactions in the crystal lattices of **5a** and **5b** (see ESI†). Natural bond orbital (NBO) analysis gave Wiberg bond index of

0.55/0.56, 0.56/0.57, 0.58/0.60 for the B ← N bond in **2a/2b**, **3a/3b**, **5a/5b**, respectively, supporting that the (a) isomer has a slightly weaker B ← N bond than that of (b) isomer. One of the mesityls in each of these molecules has π -stacking interactions with the pyrrole ring of the backbone with the shortest atomic separation distance being ~ 3.65 Å (see the side view of the structures of **3a** and **3b** in Fig. 3). The dihedral angles between the IDI ring and the N-heterocycle/the phenyl ring (bound to B) vary considerably, with **1a/1b** and **4a/4b** displaying the smallest φ_1 , **4a** and **4b** displaying the largest φ_2 , because of the highly congested *ortho*-BMe₂ unit (Fig. 2).

In solution, **2a/2b**, **3a/3b**, and **5a/5b** retain at least partially the tetracoordinate geometry, as evidenced by their ¹¹B NMR chemical shifts (22.5/10.3, 6.2/4.6, 7.7/3.3 ppm, respectively). The much down-field shifted ¹¹B NMR peak of the (a) isomers again agrees with their weaker B ← N bond, relative to that of the (b) isomers in solution. The boron unit in **1a/1b** and **4a/4b** has a typical trigonal planar geometry. For **4a/4b**, despite the *ortho*-BMe₂ group, there is no B ← N bond, which could be attributed to the unfavorable steric interactions between the benzothiazolyl and the Mes ring.

2a, **2b** and **5a** display temperature (*T*)-dependent NMR spectra while **3a** and **3b** do not show appreciable change up to 333 K, supporting the greater stability of the latter. The ¹¹B NMR signal (Fig. 4 and see ESI†) of **2a** shifts from 5.8 ppm at 253 K to 47.0 ppm at 333 K in CDCl₃ while that of **2b** from 5.6 ppm at 253 K to 24.1 ppm at 333 K. This shows that both isomers are in dynamic equilibrium with their open form, and that the closed form is more stable at low *T* while the open one becomes dominant at high *T* driven by entropy. The equilibrium constant/ $\Delta H/\Delta S$ values at 298 K for the closed to open form conversion were determined to be 0.44/6.88 kcal mol⁻¹/91 J mol⁻¹ K⁻¹ for **2a** and 0.10/8.02 kcal mol⁻¹/93 J mol⁻¹ K⁻¹ for **2b**, respectively, consistent with the DFT predicted relative stability of **2a** and **2b**. The relatively stronger B ← N bond, facilitated by the favourable inductive effect that makes the py a stronger donor and the B center a stronger acceptor, is likely a key

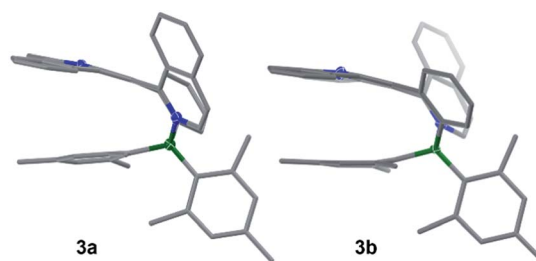


Fig. 3 Side view of the crystal structures of **3a** and **3b**.

Table 1 B ← N bond lengths (Å) and dihedral angles between IDI ring and the N-heterocycle (°, φ_1)/the phenyl ring bound to boron (°, φ_2)

	1a/1b	2a/2b	3a/3b	4a/4b	5a/5b
B ← N		1.721(5)/1.713(3)	1.709(6)/1.700(15)		1.688(2)/1.700(6)
φ_1	29.3/31.8	45.7/43.6	50.1/49.8	25.6/30.1	40.7/33.5
φ_2	38.8/50.1	42.1/41.5	43.8/42.9	71.9/63.4	43.8/33.1



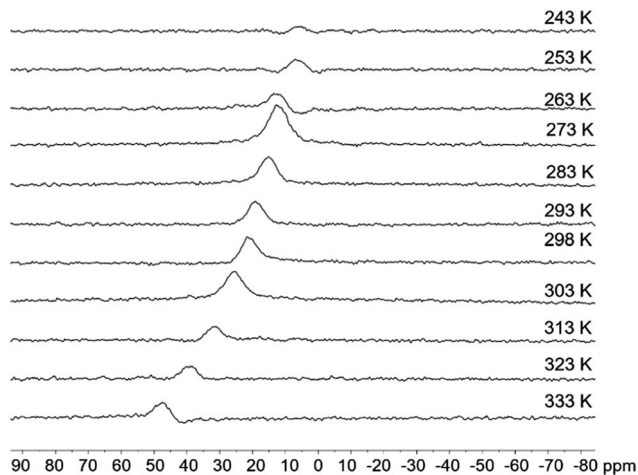


Fig. 4 VT-¹¹B NMR spectra of **2a** in CDCl₃.

contributor to the greater stability of **2b**, compared to **2a**. π -conjugation involving the BMes₂Ph unit in the corresponding open form appears to play a lesser role, given the fact that **1a** (a model for **2a**-open) is less stable than **1b**. The ¹¹B NMR spectral change with *T* for **5a** is similar to that of **2b** (1.6 ppm at 253 K to 16.1 ppm at 333 K), an indication that the B ← N bond in **5a** has a similar stability as that in **2b**. **5b** shows a change of 2 ppm (3.0 ppm at 253 K to 5.0 ppm at 333 K), indicating its greater stability than **5a**.

Photophysical and electrochemical properties of **1a/1b–5a/5b**

B ← N coordination leads to dramatically different m.p. for the regioisomers with the (b) isomers having a consistently higher m.p. than the (a) isomers (197 °C/309 °C for **2a/2b**; 197 °C/264 °C for **3a/3b**, 238 °C/262 °C for **5a/5b**, recorded for vacuum dried solids). According to DFT data, the tetracoordinate compounds have a much greater electrical dipole moment, compared to their open forms (see ESI†).

The most interesting impact of the dichotomic groups is on the electrochemical and photophysical properties of the 5 pairs of regioisomers. As shown by the absorption spectra in Fig. 5 and the data in Table 2, the regioisomers (a) consistently have a smaller E^{opt} and a longer λ_{max} (the first band) than those of the corresponding isomers (b). For example, the absorption λ_{max} of **1a** and **1b** is 478 nm and 472 nm, respectively ($\Delta E_{\text{g}}^{\text{opt}} = 0.06$ eV, $\Delta\lambda_{\text{max}} = -6$ nm) while that of **4a** and **4b** is 489 nm and 470 nm, respectively ($\Delta E_{\text{g}}^{\text{opt}} = 0.11$ eV, $\Delta\lambda_{\text{max}} = -19$ nm). Compounds **1a** and **1b** are constitutional isomers of **2a** and **2b**, and could serve as model compounds for the experimentally inaccessible **2a**-open/**2b**-open. This is in fact supported by TD-DFT calculated $S_0 \rightarrow S_1$ vertical excitation energies of **2a**-open and **2b**-open which are very close to those of **1a** and **1b** (see ESI†). From **1a** to **2a**, $E_{\text{g}}^{\text{opt}}$ decreases by 0.16 eV and λ_{max} increases by 29 nm, while, in contrast, from **1b** to **2b**, $E_{\text{g}}^{\text{opt}}$ increases by 0.04 V and λ_{max} decreases by 10 nm. As a consequence, the change of ΔE^{opt} (0.26 eV) and $\Delta\lambda_{\text{max}}$ (–45 nm) between **2a** and **2b** is much greater, compared to that between **1a** and **1b**, resulting the greater visual colour difference for **2a** and **2b** as shown by the photographs in

Fig. 5. Similarly, regioisomers **3a/3b**, and **5a/5b** also have a large E^{opt} difference ($\Delta E^{\text{opt}} = 0.32$ eV and 0.29 eV, respectively) with distinct colour difference. These data confirm that B ← N coordination indeed has an amplifying effect on E^{opt} between isomers (a) and (b).

The impact of B ← N coordination on fluorescence spectra of the regioisomers follows the same trend as observed in the absorption spectra/of E^{opt} . However, the emission colour difference displayed by each pair of the B ← N bound regioisomers is more dramatic and impressive, as shown in Fig. 5. For example, **2a** and **2b** emit a yellow and a blue colour, respectively while **3a** and **2b** emit a red and a yellow colour, respectively. All five pairs of compounds display impressive emission quantum yields in solution and the solid state and the emission colours/spectra cover nearly the entire visible region (Table 2). The sharp colour contrast displayed by these regioisomers indicates that B ← N coordination can be an effective approach in tuning optical bandgap and fluorescent colours of isomeric species.

Well-resolved oxidation and reduction peaks for **1a/1b–5a/5b** were observed in the differential pulse voltammetry (DPV) diagrams. The DPV diagrams for **1a/1b** and **2a/2b** are shown in Fig. 6, while those of other compounds are provided in the ESI.† In general, the electrochemical energy (HOMO–LUMO) gap of **1a/1b–5a/5b** follows approximately the same trend as that of E^{opt} (Table 2) and the HOMO–LUMO gap calculated by DFT (Fig. 6, see also ESI†). Although the calculated HOMO and LUMO values deviate significantly from those obtained from electrochemical data, the trend of these two sets of values for the 10 molecules agrees with each other. The impact of B ← N coordination on redox potentials is illustrated by comparing the $E_{\text{red/ox}}$ of **1a** and **2a**, **1b** and **2b**. From **1a** to **2a**, E_{ox} decreases by 0.20 V while E_{red} decreases by 0.05 V (less negative), indicating that B ← N coordination mainly destabilizes the HOMO level in (a), leading to a smaller energy gap. In contrast, from **1b** to **2b**, E_{ox} decreases by 0.08 V while E_{red} increases by 0.32 V (more negative), confirming that B ← N coordination mainly destabilizes the LUMO level in (b), leading to a larger energy gap. Within each pair of regioisomers, except **1a/1b**, the LUMO of the (b) isomer in each pair is consistently higher than that of the corresponding (a) isomer. For the HOMO level, the (b) isomer is consistently lower than that of the (a) isomer for all 5 pairs.

The distinct impact of B ← N coordination on $E^{\text{opt}}/\lambda_{\text{em}}$ of the (a) and (b) isomers is clearly caused by the different influence of the dichotomic units on the HOMO and LUMO levels. As shown in Scheme 1, in the (a) isomer, B ← N coordination makes BMes₂ph a better EDG (raising HOMO) and the N-heterocycle a better EWG (lowering LUMO). The N-heterocycle in (a) does not influence significantly HOMO level owing to the lack of HOMO at C2, but decreases LUMO *via* both –I and –M effect. The net result is the narrowing of E^{opt} in the (a) isomer. This agrees well with the DFT data calculated HOMO/LUMO diagrams of **2a** shown in Fig. 7, where the py at C2 contributes mainly to the LUMO level while the PhBMes₂ unit have a large contribution to the HOMO level. In the (b) isomer, B ← N coordination makes the N-heterocycle a better EWG (lowering HOMO/LUMO) and BMes₂ph a better EDG (increasing



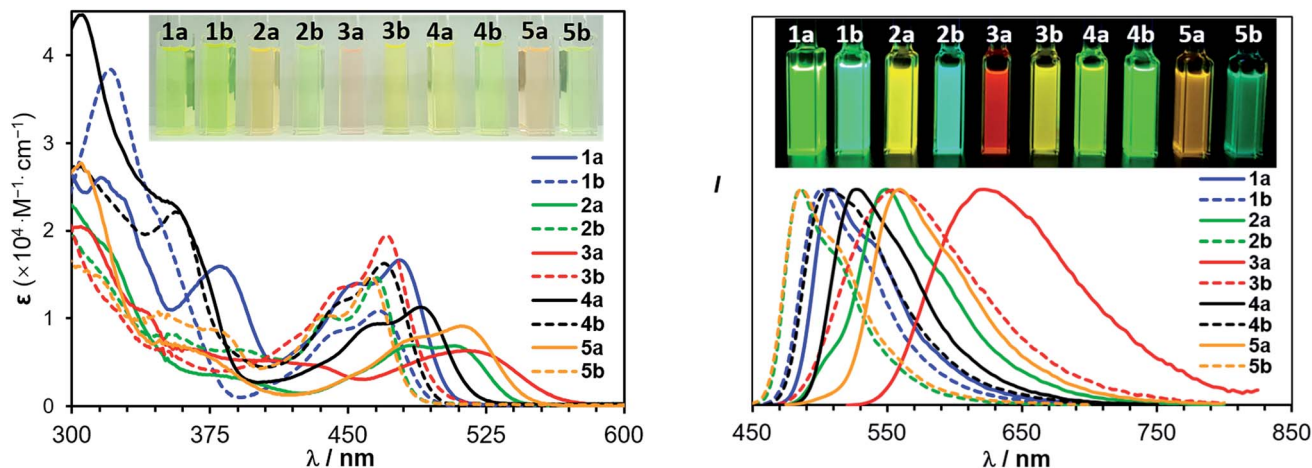


Fig. 5 Absorption (left) and normalized fluorescence (right) spectra of **1a/1b–5a/5b** in THF (1.0×10^{-5} M). Insets: photographs showing the absorption and emission colours of the 5 pairs of isomers.

LUMO), hence widening E^{opt} . Indeed, DFT data for **2b** show that the py at C1 contributes to both HOMO and LUMO levels while the Ph of the PhBMe₂ unit contributes mainly to LUMO. To further illustrate the impact of B ← N coordination, TD-DFT computations for the geometrically optimized open forms of **2a/2b**, **3a/3b** and **5a/5b** were performed. The computational data indeed show that from the open form to the closed form, the $S_0 \rightarrow S_1$ vertical transition energy decreases by ~ 0.2 eV for **2a**, **3a** and **5a** isomers but increases by about 0.1 eV for **2b** and **5b** isomers as illustrated by the **2a/2b** pair in Fig. 8 (see ESI† for data of other pairs). This supports the “amplification effect” exerted by B ← N coordination for the **2a/2b** and **5a/5b** pairs. For **3b**, from the open form to the closed form, the $S_0 \rightarrow S_1$ transition energy has a slight decrease of 0.02 eV, which is likely caused by the greater π -conjugation of the large isoquinoline ring with the backbone that partially offsets the coordination effect in **3b**.

The anomalous LUMO energy decrease for **1a** to **1b** is likely caused by the $-M$ effect of the PhBMe₂ unit that offsets its $+I$ effect to the LUMO, leading to the net stabilization of the LUMO level in **1b**. This is evidenced by the dominant contribution of

the Ph and B atom to the LUMO of **1b** shown in Fig. 7. This effect is only observed for the **1a/1b** pair because the BMe₂ unit is at the *para*-position of the phenyl ring and the entire molecule is much less congested, allowing effective π -conjugation, and maximizing the $-M$ effect, compared to the highly congested tricoordinate pair of **4a/4b**. These data unequivocally confirm that the B ← N bond and the dichotomic units indeed have a opposite effect on the HOMO/LUMO energies, leading to the contrasting electronic/photophysical properties of the regioisomers, which agrees with the postulated energy change shown in Scheme 1.

Temperature-dependent emission of **2a/2b**, **3a/3b**, and **5a/5b**

To examine the impact of B ← N bond dissociation on fluorescence colour change, the emission spectra of **2a/2b**, **3a/3b**, and **5a/5b** at various temperatures (T) were examined. Because of the relatively low m.p. (228 K) and high b.p. (403 K) of mesitylene, it was selected for the VT fluorescence experiments. The data of **2a/2b** and **3a/3b** are shown in Fig. 9 and that of **5a/5b** are provided in the ESI.† Consistent with the

Table 2 Photophysical data of **1a/1b** to **5a/5b**

	λ_{abs}^a (nm)	ϵ , ($\text{L M}^{-1} \text{cm}^{-1}$)	$E_{\text{g}}^{\text{opt}}$ (eV)	λ_{em}^b (nm)	$\Phi^{b,c}$	τ_{av} (ns) ^b	$k_{\text{r}}/k_{\text{nr}}$ (10^7s^{-1}) ^d	Exp. $E_{\text{LUMO}}/E_{\text{HOMO}}^e$ (eV)
1a	478	1.62×10^4	2.46	575/509	0.41/0.98	11.2/9.6	10.2/0.21	−2.07/−5.02
1b	467	1.09×10^4	2.52	549/502	0.22/0.89	6.8/13.4	6.64/0.82	−2.20/−5.03
2a	507	0.703×10^4	2.30	589/550	0.53/0.93	20.2/24.8	3.75/0.28	−2.12/−4.82
2b	465	1.52×10^4	2.56	550/486	0.27/0.65	13.1/11.2	5.80/3.13	−1.88/−4.95
3a	516	0.618×10^4	2.19	623/621	0.38/0.29	23.9/20.8	1.39/3.41	−2.38/−4.84
3b	471	2.04×10^4	2.51	575/556	0.59/0.54	23.2/7.0	7.71/6.57	−2.24/−5.03
4a	489	0.923×10^4	2.40	565/527	0.15/0.68	5.6/15.8	4.30/2.02	−2.07/−4.96
4b	470	1.60×10^4	2.51	559/507	0.30/0.86	14.2/10.2	8.43/1.37	−1.96/−4.99
5a	511	0.912×10^4	2.27	604/559	0.35/0.82	15.8/21.9	3.74/0.82	−2.20/−4.91
5b	462	1.43×10^4	2.56	571/487	0.31/0.49	14.2/6.0	8.17/8.50	−1.95/−5.00

^a 10^{-5} M in THF at 298 K. ^b Solid/THF (10^{-5} M). ^c The solution QY was determined in THF using an integration sphere under N₂. The solid-state QY was measured using an integration sphere. ^d Radiative decay rate constant $k_{\text{r}}/k_{\text{nr}}$ in THF is calculated by $k_{\text{r}} = \Phi/\tau_{\text{av}}$, $k_{\text{nr}} = (1 - \Phi)/\tau_{\text{av}}$. ^e E_{LUMO} or $E_{\text{HOMO}} = -(4.8 + E^{\text{red}}$ or E^{ox}).



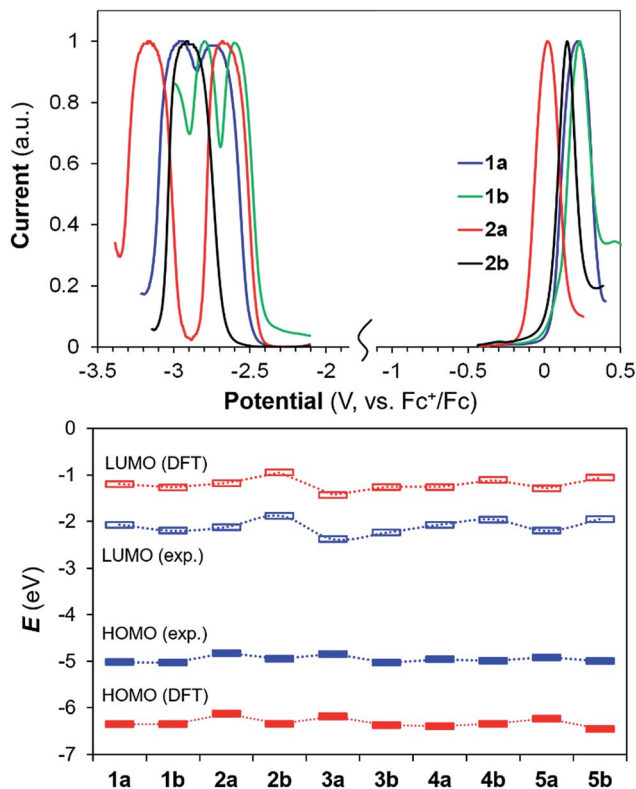


Fig. 6 Top: Differential Pulse Voltammetry (DPV) diagrams of **1a/1b** and **2a/2b** showing the reduction (recorded in THF) and oxidation peak (recorded in CH₂Cl₂) using NBu₄PF₆ (0.1 M) as the electrolyte and a scan rate of 100 mV s⁻¹. Bottom: HOMO and LUMO energies obtained from DFT (red) and electrochemical data (blue).

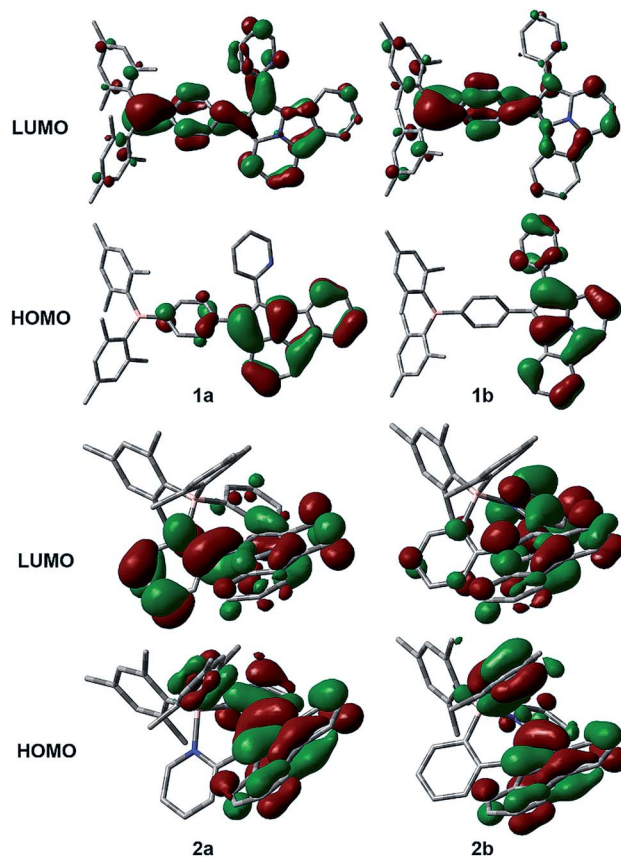


Fig. 7 HOMO and LUMO diagrams of **1a/1b** and **2a/2b** (DFT, M06-2X/6-31g(d)).

dynamic equilibrium of **2a/2b** with their corresponding open forms in solution, the fluorescence spectra of **2a** and **2b** show contrasting dependence on T with **2a** being more sensitive (Fig. 9). For **2a**, upon increasing T from 233 K to 403 K, the emission peak gradually blue shifts, with the appearance of a well resolved new peak at 517 nm. In contrast, compound **2b** undergoes a small red shift from 489 nm to 500 nm with increasing T . The emission spectral change of **2a** and **2b** with T reverses the $B \leftarrow N$ amplification effect shown in Fig. 8, thus consistent with the $B \leftarrow N$ bond dissociation and the formation of the open isomer. The trend of the emission spectral change of **5a** and **5b** with increasing T is similar to that of **2a** and **2b**, respectively (see ESI[†]), and in agreement with the closed to open form $S_0 \rightarrow S_1$ transition energy change predicted by TD-DFT data. For **3a**, with increasing T , the emission spectrum only undergoes a slight blue shift with a shoulder peak rising at ~ 515 nm, which can be attributed to **3a**-open, based on TD-DFT prediction. For **3b** the emission spectrum blue shifts by about 10 nm with increasing T , which also agrees with the DFT predicted $S_0 \rightarrow S_1$ transition energy change between **3b** and **3b**-open. The T -dependent emission spectral change of **2a/2b**, **3a/3b** and **5a/5b** is fully reversible, supporting the potential of this new class of $B \leftarrow N$ coordination bond-based system as temperature probes.¹¹

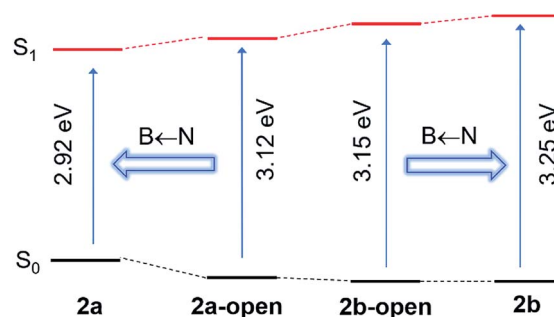


Fig. 8 TD-DFT calculated $S_0 \rightarrow S_1$ vertical transition energies for **2a/2b** and their corresponding open forms (DFT, M06-2X/6-31g(d)), illustrating the "amplification effect" of $B \leftarrow N$ coordination.

Switching $B \leftarrow N$ bond with fluoride ions

Fluoride titration experiments are commonly used for probing the Lewis acidity and the electronic properties of triarylboron compounds because of their relatively high affinity/selectivity toward fluoride ions.^{36,4} This approach is used by us to gauge the relative $B \leftarrow N$ bond strength in **2a/2b**, **3a/3b**, and **5a/5b**. Both absorption and emission spectra with the addition of NBu₄F (TBAF) were recorded and those of **2a/2b** and **2a/2b** are shown in Fig. 10 (for **5a/5b**, see the ESI[†]). The addition of ~ 10 eq. TBAF to the THF solution of **2a** fully converts **2a** to its F⁻-



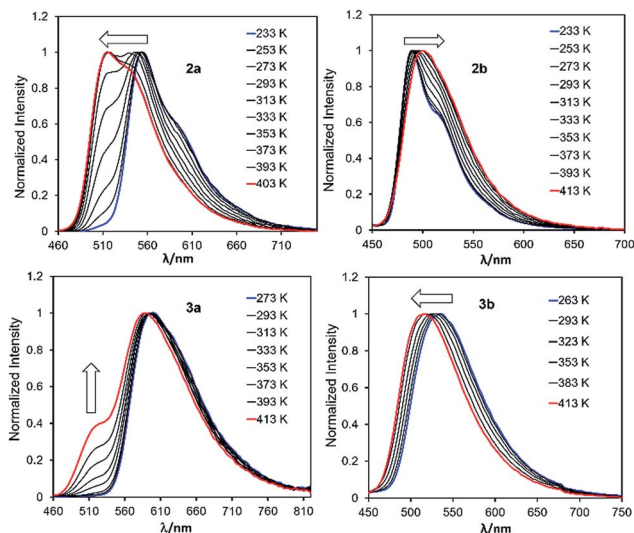


Fig. 9 Temperature-dependent fluorescence spectra of **2a/2b** and **3a/3b** in mesitylene under N_2 .

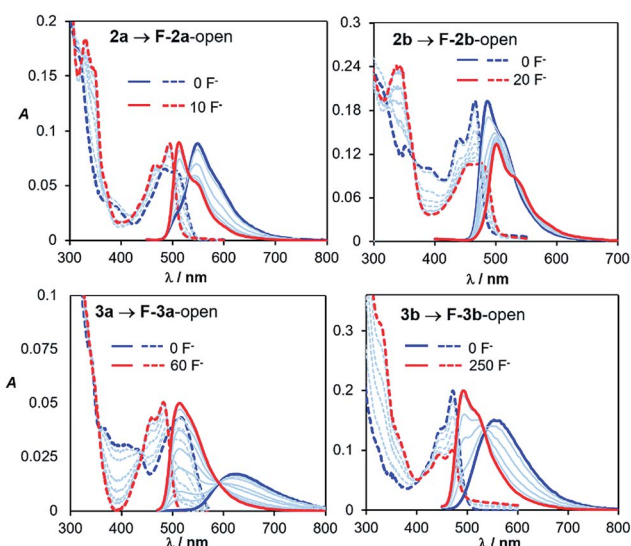


Fig. 10 Absorption (dashed lines) and emission (solid lines) spectral change of **2a**, **2b**, **3a** and **3b** with TBAF addition in THF (1.0×10^{-5} M).

bound open form, accompanied by the emission colour change from yellow to green (Fig. 10). In contrast, for **2b**, about 20 eq. TBAF are needed to reach the saturation point with the colour change from blue-green to green. In fact, the change of the emission spectra and the CIE coordinates of **2a** and **2b** with TBAF is similar to that observed with increasing T (Fig. 9 and 10; and ESI^+). For **3a/3b**, and **5a/5b**, the trend of the emission spectral/CIE coordinates change with TBAF is also parallel to that observed with increasing T (Fig. 11). Because of the tetra-coordinate geometry in the F^- -bound open form, the boron unit has the +I effect only. Thus, the similar emission spectral change with T and TBAF displayed by the three pairs of regioisomers indicates that the inductive effect exerted by the boron unit likely has a greater role in the distinct photophysical

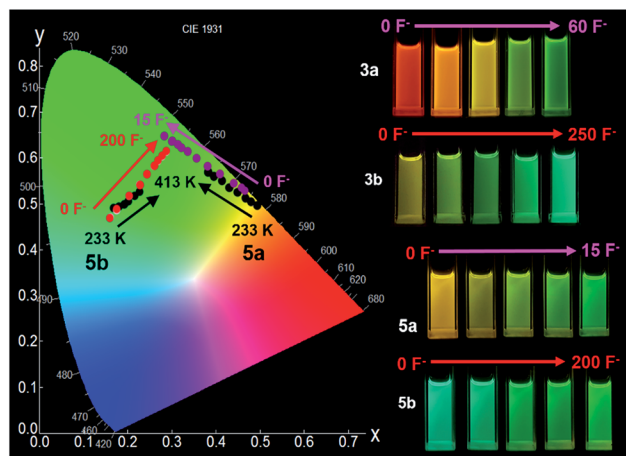


Fig. 11 Left: a CIE diagram showing the emission colour change of **5a** and **5b** with T (black dots) in mesitylene and the addition of TBAF (purple and red, respectively) in THF. Right: photographs showing the contrasting emission colour change of **3a/3b** and **5a/5b** with the addition of TBAF in THF (1.0×10^{-5} M).

properties of the (a) and (b) isomers. Furthermore, the fluoride titration experiments show that the isomer (a) consistently requires much less TBAF to reach the saturation point than the isomer (b) for all three pairs (10 eq./20 eq. for **2a/2b**; ~60 eq./250 eq. for **3a/3b**; ~15 eq./100 eq. for **5a/5b**), which supports the greater stability of the $B \leftarrow N$ bond in the (b) isomer, compared to that in the (a) isomer. Notably, compounds **3a**, **3b** and **5b** require much more TBAF, compared to the other isomers, to reach saturation, an indication that the $B \leftarrow N$ bond in these three compounds are likely much stronger than those in **2a/2b** and **5a**. This is in good agreement with the VT ^{11}B NMR data. Based on the VT NMR data and the TBAF titration data, the $B \leftarrow N$ bond strength follows approximately the order of $2a < 2b \approx 5a < 3a < 5b < 3b$. Clearly both Lewis basicity of the N-heterocycles and the cooperative or uncooperative inductive effect of dichotomic units in the (a) and the (b) isomers are the main contributors to the $B \leftarrow N$ strength. Steric/entropy factors must also play a role.

Conclusions

5 pairs of regioisomers based on an IDI backbone and dichotomic substituents have been achieved. The distinct electronic/photophysical properties for each pair of the regioisomers have been established. The contrasting impact of the dichotomic substituents to the regioisomers is ascribed to both inductive and mesomeric effects. For compounds that contain an *ortho*-BMes₂, the +I effect appears to dominate. The formation of an internal $B \leftarrow N$ bond amplifies the impact of the dichotomic units, leading to a smaller optical energy gap for the (a) isomer and a greater optical gap for the (b) isomer, respectively. The (a) isomers show consistently a weaker $B \leftarrow N$ bond than the (b) isomers. This work demonstrates that for π -conjugated systems that have unsymmetrical electron density/orbital distributions on the backbone, utilizing internal B



← N coordination and regioisomerism could be highly effective in tuning the optoelectronic properties of organic materials.

Experimental

General

Solvents were freshly distilled over Na or CaH₂ and stored under nitrogen prior to use. Starting materials were purchased from Sun Chemical Technology Co., Ltd. and used without further purification. 2-(2'-(Dimesitylboryl)phenyl-ethynyl)pyridine (**2**),¹² (4-bromophenyl)dimesitylborylborane,¹³ 2-ethynylpyridine,¹⁴ pyrido[1,2-*a*]isoindole,¹⁵ 1-dimesitylboryl-2-ethynylbenzene¹⁶ were synthesized according to procedures reported in literature. All the reactions were performed under nitrogen and in dry solvent. ¹H, ¹³C and ¹¹B NMR spectra were recorded on a Bruker Ascend 400 or 700 MHz spectrometer. Melting point (m.p.) determination was performed using a SGW X-4A microscopic melting point instrument. UV-vis spectra were obtained on an Agilent Cary 300 UV-vis spectrophotometer. Fluorescence spectra were recorded on an Edinburgh Instruments FLS980 spectrophotometer. High-resolution mass spectra (HRMS) were obtained from an Agilent Q-TOF 6520 LC-MS spectrometer. Fluorescent quantum efficiencies were determined using a Hamamatsu Quantaurus-QY spectrometer (C11347). The electrochemical cell was a standard three-compartment cell composed of a Pt working electrode, a Pt auxiliary electrode, and a Pt wire reference electrode. All measurements were performed using 0.10 M of NBu₄PF₆ in THF or CH₂Cl₂ as the supporting electrolyte. The potentials are reported relative to the ferrocene/ferrocenium couple. Analytical thin-layer chromatography was performed on glass plates of Silica Gel GF-254. The synthetic procedures for PhBMes₂/M-heterocycle functionalized alkyne starting materials (**1**, **3–5**) are provided in the ESI.†

Synthesis of compound **1a** and **1b**

Pyrido[1,2-*a*]isoindole (738 mg, 4.42 mmol) and compound **1** (947 mg, 2.21 mmol) were reacted in toluene (15.0 mL) at 120 °C in a 50 ml sealed tube for 6 days, after the mixture was cooled to room temperature, DDQ (1.10 g, 3.98 mmol) was added. After reacting for 30 min, the solvent was removed under reduced pressure and the product was isolated by column chromatography on neutral alumina using petroleum ether/ethyl acetate (50 : 1 to 10 : 1) as the eluent. The product **1a** was isolated as a yellow solid (1.038 g, 79% yield), the product **1b** isolated as a yellow solid (30 mg, 2% yield).

1a. ¹H NMR (400 MHz, CDCl₃): δ 8.93 (d, *J* = 4.7 Hz, 1H), 8.70 (d, *J* = 8.1 Hz, 1H), 8.44 (d, *J* = 7.9 Hz, 1H), 8.17–8.10 (m, 2H), 7.73 (m, 2H), 7.62–7.52 (m, 6H), 7.43 (d, *J* = 7.9 Hz, 1H), 7.27–7.24 (m, 1H), 6.86 (s, 4H), 2.33 (s, 6H), 2.11 (s, 12H). ¹³C NMR (101 MHz, CDCl₃): δ 154.74, 149.92, 143.46, 141.84, 140.80, 139.35, 138.50, 137.06, 135.66, 130.30, 129.92, 129.22, 128.74, 128.31, 128.17, 127.28, 126.98, 125.15, 123.52, 122.14, 122.03, 121.68, 121.60, 120.77, 119.13, 115.30, 108.79, 23.51, 21.22. HR-ESIMS (*m/z*): [M + H]⁺ calcd. for C₄₃H₃₈BN₂, 593.3123; found: 593.3119.

1b. ¹H NMR (400 MHz, CDCl₃): δ 8.75 (d, *J* = 4.4 Hz, 1H), 8.65 (d, *J* = 8.3 Hz, 1H), 8.44 (d, *J* = 7.9 Hz, 1H), 8.15–8.08 (m, 2H), 7.82 (t, *J* = 7.8 Hz, 1H), 7.76 (d, *J* = 7.7 Hz, 2H), 7.67 (d, *J* = 7.7 Hz, 2H), 7.64 (t, *J* = 7.7 Hz, 1H), 7.53 (t, *J* = 7.5 Hz, 1H), 7.45 (t, *J* = 7.6 Hz, 1H), 7.22 (d, *J* = 8.0 Hz, 1H), 7.13–7.07 (m, 1H), 6.87 (s, 4H), 2.34 (s, 6H), 2.13 (s, 12H). ¹³C NMR (176 MHz, CDCl₃): δ 155.38, 149.71, 144.95, 141.77, 140.85, 139.14, 138.71, 136.86, 135.66, 135.46, 130.07, 129.80, 129.33, 128.40, 128.37, 128.25, 128.15, 127.82, 123.90, 123.33, 122.53, 121.84, 120.88, 120.05, 119.47, 118.04, 117.25, 108.92, 23.54, 21.24. HR-ESIMS (*m/z*): [M + H]⁺ calcd. for C₄₃H₃₈BN₂, 593.3123; found: 593.3122.

Synthesis of compounds **2a** and **2b**

Pyrido[1,2-*a*]isoindole (558 mg, 3.34 mmol) and compound **2** (476 mg, 1.11 mmol) were reacted in toluene (20.0 mL) at 120 °C in a 50 ml seal tube for 7 days, after the mixture was cooled to room temperature, DDQ (759 mg, 3.34 mmol) was added. After reacting 30 min, the solvent was removed *in vacuo* and the product was isolated by column chromatography on neutral alumina using petroleum ether/ethyl acetate (50 : 1 to 10 : 1) as the eluent. Product **2a** was isolated as an orange solid (541 mg, 82% yield), product **2b** was isolated as a yellow solid (27 mg, 4% yield).

2a. ¹H NMR (400 MHz, CDCl₃): δ 9.21 (d, *J* = 5.6 Hz, 1H), 8.42 (d, *J* = 7.9 Hz, 1H), 8.16 (d, *J* = 7.8 Hz, 1H), 8.03 (m, 3H), 7.91 (t, *J* = 7.5 Hz, 1H), 7.86 (d, *J* = 7.4 Hz, 1H), 7.71–7.63 (m, 2H), 7.60–7.50 (m, 2H), 7.47 (td, *J* = 7.4, 1.0 Hz, 1H), 7.38–7.30 (m, 1H), 7.28–7.24 (m, 1H), 7.00 (s, 1H), 6.81 (s, 1H), 5.35 (s, 1H), 5.20 (s, 1H), 2.35 (s, 3H), 1.89 (s, 6H), 1.31 (s, 6H), 0.99 (s, 3H). ¹³C NMR (176 MHz, CDCl₃): δ 155.16, 154.02, 150.42, 147.46, 145.34, 144.73, 141.86, 140.80, 138.71, 138.53, 138.24, 138.04, 136.12, 132.26, 130.15, 129.66, 129.18, 128.84, 128.57, 128.53, 127.77, 127.66, 127.14, 126.56, 126.33, 125.76, 125.34, 123.97, 122.71, 122.15, 121.02, 120.64, 119.91, 119.08, 116.46, 109.07, 27.15, 25.40, 24.41, 22.88, 20.89, 19.38. ¹¹B NMR (225 MHz, CDCl₃): δ 22.5. HR-ESIMS (*m/z*): [M + H]⁺ calcd. for C₄₃H₃₈BN₂, 593.3123; found: 593.3148.

2b. ¹H NMR (400 MHz, CDCl₃): δ 9.16 (d, *J* = 6.0 Hz, 1H), 8.41 (d, *J* = 7.9 Hz, 1H), 8.27 (d, *J* = 8.0 Hz, 1H), 8.06–7.87 (m, 4H), 7.83 (d, *J* = 8.2 Hz, 1H), 7.77–7.65 (m, 2H), 7.62–7.52 (m, 2H), 7.48 (td, *J* = 7.4, 1.0 Hz, 1H), 7.32–7.27 (m, 1H), 7.25–7.20 (m, 1H), 7.10 (s, 1H), 6.86 (s, 1H), 5.38 (s, 1H), 5.03 (s, 1H), 2.38 (s, 3H), 2.03 (s, 3H), 1.88 (s, 3H), 1.24 (s, 3H), 1.15 (s, 3H), 1.00 (s, 3H). ¹³C NMR (176 MHz, CDCl₃): δ 156.97, 153.45, 150.71, 148.74, 146.06, 145.47, 141.93, 141.58, 138.89, 138.40, 138.35, 137.26, 135.56, 134.55, 131.53, 130.55, 130.12, 130.00, 129.50, 128.97, 128.67, 128.19, 128.18, 127.98, 126.69, 126.58, 126.26, 126.11, 123.23, 122.14, 121.71, 120.81, 119.47, 119.39, 115.32, 113.10, 108.25, 27.75, 25.98, 24.44, 22.92, 20.87, 19.41. ¹¹B NMR (225 MHz, CDCl₃): δ 10.3. HR-ESIMS (*m/z*): [M + H]⁺ calcd. for C₄₃H₃₈BN₂, 593.3123; found: 593.3112.

Synthesis of compound **3a** and **3b**

Pyrido[1,2-*a*]isoindole (140 mg, 0.84 mmol) and compound **3** (134 mg, 0.28 mmol) were reacted in toluene (10.0 mL) at 120 °C in a 50 ml seal tube for 6 days, after the mixture was cooled to



room temperature, DDQ (191 mg, 0.84 mmol) was added. After reacting 30 min, the volatiles were removed *in vacuo* and the product was isolated by column chromatography using petroleum ether/toluene (20 : 1 to 5 : 1) as the eluent. Product **3a** was isolated as a red solid (130 mg, 72% yield), product **3b** was isolated as a yellow solid (33 mg, 18% yield).

3a. $^1\text{H NMR}$ (700 MHz, CDCl_3): δ 9.17 (d, $J = 7.0$ Hz, 1H), 8.53 (d, $J = 8.7$ Hz, 1H), 8.41 (d, $J = 7.9$ Hz, 1H), 8.10 (d, $J = 8.2$ Hz, 1H), 8.07 (d, $J = 7.2$ Hz, 1H), 7.99 (d, $J = 8.2$ Hz, 1H), 7.88 (t, $J = 7.5$ Hz, 1H), 7.82 (d, $J = 7.3$ Hz, 1H), 7.77–7.69 (m, 2H), 7.59 (d, $J = 7.8$ Hz, 1H), 7.51–7.43 (m, 2H), 7.38 (t, $J = 7.3$ Hz, 1H), 7.35 (t, $J = 7.6$ Hz, 1H), 7.17 (t, $J = 7.5$ Hz, 1H), 7.14 (s, 1H), 6.88 (s, 1H), 6.50 (d, $J = 8.0$ Hz, 1H), 5.23 (s, 1H), 4.82 (s, 1H), 2.41 (s, 3H), 2.03 (s, 3H), 1.81 (s, 3H), 1.13 (s, 6H), 1.02 (s, 3H). $^{13}\text{C NMR}$ (176 MHz, CDCl_3): δ 158.71, 155.89, 148.79, 146.34, 146.21, 142.59, 141.99, 140.68, 139.38, 138.13, 137.46, 136.39, 135.31, 133.01, 131.45, 130.69, 130.58, 130.17, 128.92, 128.86, 128.84, 128.81, 128.65, 127.82, 127.60, 127.36, 126.88, 126.29, 126.10, 126.01, 125.88, 125.12, 123.22, 122.58, 122.01, 120.19, 119.06, 118.95, 116.98, 109.86, 28.22, 26.15, 24.64, 24.59, 20.87, 19.49. $^{11}\text{B NMR}$ (225 MHz, CDCl_3): δ 6.2. HR-ESIMS (m/z): $[\text{M} + \text{H}]^+$ calcd. for $\text{C}_{47}\text{H}_{40}\text{BN}_2$, 643.3285; found: 643.3260.

3b. $^1\text{H NMR}$ (400 MHz, CDCl_3): δ 9.04 (d, $J = 7.1$ Hz, 1H), 8.40 (m, 1H), 8.38 (m, 1H), 8.24 (d, $J = 8.0$ Hz, 1H), 8.00 (d, $J = 7.4$ Hz, 1H), 7.94 (d, $J = 7.4$ Hz, 1H), 7.90 (d, $J = 8.1$ Hz, 1H), 7.84 (t, $J = 7.4$ Hz, 1H), 7.69 (t, $J = 7.6$ Hz, 1H), 7.66–7.54 (m, 4H), 7.54–7.49 (m, 1H), 7.45 (t, $J = 7.3$ Hz, 1H), 7.24–7.22 (m, 1H), 7.13 (s, 1H), 7.07 (d, $J = 8.3$ Hz, 1H), 6.87 (s, 1H), 5.09 (s, 1H), 5.02 (s, 1H), 2.40 (s, 3H), 2.01 (s, 3H), 1.83 (s, 3H), 1.18 (s, 3H), 1.10 (s, 3H), 1.00 (s, 3H). $^{13}\text{C NMR}$ (176 MHz, CDCl_3): δ 158.75, 155.79, 148.97, 146.27, 146.17, 142.58, 142.11, 141.00, 139.30, 137.82, 137.26, 136.85, 136.81, 135.25, 132.88, 132.26, 130.62, 130.55, 130.41, 130.14, 130.11, 129.15, 129.11, 128.89, 128.63, 128.33, 127.48, 127.44, 126.27, 125.78, 125.68, 125.66, 123.47, 122.08, 121.91, 121.49, 119.36, 117.65, 114.16, 112.44, 108.16, 27.99, 26.14, 24.59, 24.57, 20.87, 19.47. $^{11}\text{B NMR}$ (128 MHz, CDCl_3): δ 4.6. HR-ESIMS (m/z): $[\text{M} + \text{H}]^+$ calcd. for $\text{C}_{47}\text{H}_{40}\text{BN}_2$, 643.3285; found: 643.3267.

Synthesis of compound 4a and 4b

Pyrido[1,2-*a*]isoindole (86 mg, 0.51 mmol) and compound **4** (100 mg, 0.21 mmol) were reacted in toluene (10.0 mL) at 120 °C in a 50 ml seal tube for 5 days, after the mixture was cooled to room temperature, DDQ (117 mg, 0.51 mmol) was added. After reacting 30 min, the solvent was removed *in vacuo* and the product was isolated by column chromatography using petroleum ether/ CH_2Cl_2 (5 : 1) as the eluent. Product **4a** was isolated as a yellow solid (83 mg, 62% yield), product **4b** was isolated as a yellow solid (25 mg, 19% yield).

4a. $^1\text{H NMR}$ (400 MHz, CDCl_3): δ 9.12 (d, $J = 6.5$ Hz, 1H), 8.40 (d, $J = 7.9$ Hz, 1H), 8.24 (d, $J = 8.1$ Hz, 1H), 7.98 (d, $J = 7.2$ Hz, 1H), 7.89 (d, $J = 7.4$ Hz, 1H), 7.78 (m, 2H), 7.74–7.66 (m, 2H), 7.65–7.45 (m, 5H), 7.34 (t, $J = 7.5$ Hz, 1H), 6.14 (s, 4H), 2.38–1.18 (m, 18H). $^{13}\text{C NMR}$ (176 MHz, CDCl_3): δ 163.03, 153.51, 150.37, 141.98, 139.61, 136.77, 136.17, 133.37, 130.49, 129.90, 128.92, 128.10, 127.79, 127.49, 125.79, 124.45, 123.40, 123.00, 122.26,

121.69, 121.18, 119.57, 116.06, 108.81, 22.86, 20.74. HR-ESIMS (m/z): $[\text{M} + \text{H}]^+$ calcd. for $\text{C}_{45}\text{H}_{38}\text{BN}_2\text{S}$, 649.2849; found: 649.2840.

4b. $^1\text{H NMR}$ (400 MHz, CDCl_3): δ 8.67 (s, 1H), 8.34 (d, $J = 7.8$ Hz, 1H), 8.06 (s, 1H), 8.02–7.93 (m, 2H), 7.91–7.79 (m, 2H), 7.77 (t, $J = 7.4$ Hz, 1H), 7.73–7.50 (m, 5H), 7.43 (t, $J = 7.5$ Hz, 1H), 7.31–7.21 (m, 1H), 6.58 (m, 2H), 5.55 (s, 1H), 5.10 (s, 1H), 2.17 (s, 3H), 1.98–1.35 (m, 15H). $^{13}\text{C NMR}$ (176 MHz, CDCl_3): δ 163.22, 153.77, 148.63, 142.49, 142.26, 141.80, 141.34, 140.37, 139.05, 137.71, 137.49, 137.10, 135.76, 135.32, 131.34, 130.81, 130.35, 130.22, 129.37, 128.25, 128.15, 127.51, 127.37, 126.11, 125.58, 123.60, 123.36, 122.96, 122.19, 122.05, 120.86, 120.22, 116.30, 115.33, 109.11, 23.79, 22.45, 21.01, 20.59. HR-ESIMS (m/z): $[\text{M} + \text{H}]^+$ calcd. for $\text{C}_{45}\text{H}_{38}\text{BN}_2\text{S}$, 649.2849; found: 649.2829.

Synthesis of compound 5a and 5b

Pyrido[1,2-*a*]isoindole (694 mg, 4.16 mmol) and compound **5** (720 mg, 1.66 mmol) were reacted in toluene (25.0 mL) at 120 °C in a 50 ml seal tube for 5 days, after the mixture was cooled to room temperature, DDQ (943 mg, 4.16 mmol) was added. After reacting 30 min, the solvent was removed *in vacuo* and the product was isolated by column chromatography using petroleum ether/ethyl acetate (20 : 1) or petroleum ether/ CH_2Cl_2 (5 : 1) as the eluent. Product **5a** was isolated as a red solid (558 mg, 68% yield), product **5b** was isolated as a yellow solid (262 mg, 14% yield).

5a. $^1\text{H NMR}$ (400 MHz, CDCl_3): δ 8.40 (m, 2H), 8.20 (s, 1H), 8.12 (d, $J = 8.4$ Hz, 1H), 8.05 (d, $J = 7.2$ Hz, 1H), 7.81 (d, $J = 7.5$ Hz, 1H), 7.69 (m, 2H), 7.55 (m, 2H), 7.39 (t, $J = 7.4$ Hz, 1H), 7.31 (d, $J = 3.9$ Hz, 1H), 7.26–7.18 (m, 1H), 7.05 (s, 1H), 6.82 (s, 1H), 5.39 (s, 1H), 5.19 (s, 1H), 2.35 (s, 3H), 2.08 (s, 3H), 1.93 (s, 3H), 1.38 (s, 3H), 1.22 (s, 3H), 0.89 (s, 3H). $^{13}\text{C NMR}$ (176 MHz, CDCl_3): δ 164.58, 157.25, 149.54, 145.86, 144.45, 142.71, 142.00, 139.34, 138.85, 136.79, 135.55, 131.47, 130.11, 129.17, 128.70, 128.60, 128.33, 128.11, 127.12, 126.55, 126.38, 126.20, 125.97, 125.87, 123.39, 122.36, 120.71, 120.58, 118.80, 117.48, 116.13, 114.81, 109.97, 27.36, 25.93, 24.44, 22.74, 20.92, 19.23. $^{11}\text{B NMR}$ (128 MHz, CDCl_3): δ 7.7. HR-ESIMS (m/z): $[\text{M} + \text{H}]^+$ calcd. for $\text{C}_{41}\text{H}_{36}\text{BN}_2\text{S}$, 599.2692; found: 599.2686.

5b. $^1\text{H NMR}$ (400 MHz, CDCl_3): δ 8.40 (d, $J = 8.0$ Hz, 1H), 8.33 (d, $J = 8.1$ Hz, 1H), 8.20–7.98 (m, 4H), 7.80 (t, $J = 8.0$ Hz, 1H), 7.72 (t, $J = 7.7$ Hz, 1H), 7.64 (d, $J = 7.8$ Hz, 1H), 7.59 (t, $J = 7.6$ Hz, 1H), 7.45 (t, $J = 7.4$ Hz, 1H), 7.31 (t, $J = 7.5$ Hz, 1H), 7.15 (s, 1H), 7.11 (s, 1H), 6.86 (s, 1H), 5.47 (s, 1H), 5.23 (s, 1H), 2.39 (s, 3H), 2.17 (s, 3H), 1.97 (s, 3H), 1.37 (s, 3H), 1.23 (s, 3H), 0.97 (s, 3H). $^{13}\text{C NMR}$ (176 MHz, CDCl_3): δ 163.87, 158.63, 150.49, 146.06, 145.47, 144.85, 143.05, 142.90, 139.29, 139.00, 136.43, 135.49, 135.38, 131.40, 130.38, 130.36, 130.32, 130.29, 129.43, 129.28, 128.88, 128.21, 127.07, 126.88, 125.99, 125.80, 124.10, 123.13, 122.48, 121.39, 119.89, 113.67, 112.65, 108.75, 27.47, 26.17, 24.57, 22.74, 21.02, 19.40. $^{11}\text{B NMR}$ (128 MHz, CDCl_3): δ 3.3. HR-ESIMS (m/z): $[\text{M} + \text{H}]^+$ calcd. for $\text{C}_{41}\text{H}_{36}\text{BN}_2\text{S}$, 599.2692; found: 599.2683.

Computational Study

DFT and TD-DFT calculations were performed using the Gaussian 09 suite of programs.¹⁷ Geometry optimizations of all



compounds were obtained at the M06-2X¹⁸/6-31g(d)¹⁹ level of theory using Polarizable Continuum Models (PCMs) and the resulting structures were confirmed to be stationary points through vibrational frequency analysis. The Natural Bond Orbital (NBO)²⁰ analysis were performed at the same level of theory, ¹¹B-NMR of **2a/2a**-open was obtained at the B972²¹/def2TZVP²² level of theory in CHCl₃ using SMD.²³ All the calculations were performed at 298 K. Orbital composition analysis with Ros-Schuit (C-squared Population Analysis, SCPA) method using Multiwfn software version 3.4²⁴ Electrostatic potential (ESP) charge calculated according to Merz-Kollman (MK) Methods.²⁵

Single-crystal X-ray Diffraction Analysis

The crystal data were collected on a Bruker D8-Venture diffractometer with Mo-target ($\lambda = 0.71073 \text{ \AA}$) at 180 K for all compounds. Data were processed on a PC with the aid of the Bruker SHELXTL software package and corrected for absorption effects. All non-hydrogen atoms were refined anisotropically. The positions of hydrogen atoms were calculated and refined isotropically. The crystal data were deposited to the Cambridge Crystallographic Data Centre with deposition numbers of CCDC 1839688 (**B1a**), 1839689 (**B1b**), 1839690 (**B2a**), 1839691 (**B2b**), 1839692 (**B3b**), 1839693 (**B3a**), 1839694 (**B4a**), 1839695 (**B4b**), 1854311 (**B5a**), 1853312 (**B5b**). This data can be obtained free of charge from the Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.

Conflicts of interest

There are no conflicts to declare.

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