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Inhibiting effect of CO₂ on the oxidative combustion thermodynamics of coal

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A thermal analysis experiment was conducted in O₂/N₂/CO₂ and O₂/N₂ atmospheres (O₂ concentrations were 21, 14, 8, and CO₂ concentrations were 0, 39, 46, 52) to investigate the thermal behavior of coal oxidation and combustion. Results demonstrated that an elevated CO₂ concentration or decreased O₂ concentration had a delaying effect on the thermogravimetric analysis and differential scanning calorimetry (DSC) curves; moreover, the characteristic temperatures were substantially augmented. When the O₂ concentration was 21 vol%, the total heat released by coals A (highly volatile bituminous coal) and B (anthracite coal) decreased by 5.8% and 4.1%, respectively, after CO₂ addition. The comprehensive combustion performance index was also lowered. The DSC curve can be divided into two exothermic peaks, and the ratio of the peak 1 to peak 2 areas decreased with the addition of CO₂, which indicated that CO₂ inhibited the oxidation of the active functional groups of coal structures. Apparent activation energy in O₂/CO₂/N₂ was less than that in O₂/N₂.

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1. Introduction

Coal is one of the most commonly used fossil energy resources in the world.¹ However, spontaneous combustion of coal is commonplace in many coal-producing countries, resulting in not only safety problems and resource waste^{2–4} but also ecological damage and environmental pollution.^{5,36} Therefore, understanding the characteristics of coal spontaneous combustion and developing preventative technology is crucial.⁶ Conventional techniques adopted to prevent and control coal spontaneous combustion are irrigation with water and loess mud.⁷ However, water and loess mud are unable to completely cover areas affected by the spontaneous combustion of coal. Consequently, a series of coal fire prevention and extinguishment materials, such as cellular grout, three-phase foam, gel foam, and suspension sand have been developed.^{8,37} CO₂ has not only a density greater than air but also an effective flame retardant effect. As a result, pouring CO₂ into a goaf will completely bury the lower area, which is beneficial for the prevention and control of coal fires.⁹ Therefore, CO₂ firefighting technology has been developed, and used in the prevention and control of coal fires. Therefore, the inhibiting effect of CO₂ on

spontaneous combustion of coal under different conditions is the key to improving and using CO₂ firefighting technology. But previous publications have focused on the characteristics of coal oxidation and combustion, and relevant experimental techniques such as thermogravimetric analysis (TGA) and differential scanning calorimetry (DSC) have been increasingly used.^{10–12} Rotaru *et al.*¹³ discovered that different types of coal demonstrated different apparent activation energies in the process of spontaneous combustion. Deng *et al.*¹⁴ determined that a large difference in activation energy existed between the oxidation and combustion stages. The study of oxygen-enriched combustion characteristics involves examining the impact of CO₂ on coal combustion.¹⁵ Liu *et al.*¹⁶ combusted two chars in O₂/CO₂ and O₂/N₂ atmospheres and found that substituting CO₂ for inert nitrogen in the oxidizer had a subtle effect on coal combustion. Várhegyi *et al.*¹⁷ demonstrated that during coal combustion in an O₂/CO₂ atmosphere, the net reaction rate of coal combustion was only proportional to the partial pressure of oxygen. Rathnam *et al.*¹⁸ revealed that combustion reactivity prominently increased at high temperature stages in a 2% O₂/CO₂ atmosphere compared with a 2% O₂/N₂ atmosphere. Zhou *et al.*¹⁹ discovered that CO₂ has a delayed effect on the maximum mass-loss rate and burnout temperature in the coal combustion process. Therefore, the influence of CO₂ on characteristics of coal oxidation and combustion should be further study.

To further examine the inhibiting effect of CO₂ on the characteristics of coal oxidation and combustion in different atmospheres, this study used the TGA and DSC methods to analyze the thermal behaviors and variation of apparent activation energies in O₂/N₂/CO₂ and O₂/N₂ atmospheres. The main

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purposes of this research were to: (1) investigate the thermal behavior in different atmospheres; (2) analyze the effect of CO₂ and oxygen concentration on apparent activation energies during oxidation and combustion of coal; (3) research the influence of CO₂ and oxygen concentration on coal oxidation and combustion. The results are useful for understanding the oxidation and combustion characteristics of coal in O₂/N₂/CO₂ and O₂/N₂ atmospheres and can aid development of CO₂ fire prevention technology.

2. Experiments and methods

2.1. Experimental samples

Coal A was highly volatile bituminous coal and was collected from the Juye mining area in Shandong Province, China, whereas coal B was anthracite coal and came from the Furong mining area in Sichuan Province, China. The proximate and ultimate analyses of the two samples are listed in Table 1. The samples had a particle size of 0.067–0.079 mm and were stored in an airtight container.

2.2. Experimental details

A synchronous thermal analyzer, STA449F3 (Netzsch, Germany), was employed to conduct the experimental tests. The heating rates were set as 5, 10, and 15 °C min⁻¹. The experimental samples were heated from 30 to 800 °C for each heating rate. The gases, mixtures of O₂, N₂, and CO₂, were constantly supplied at a 100 mL min⁻¹ flow rate. Table 2 presents the gas mixtures with different volume fractions of O₂, N₂, and CO₂.

2.3. Analysis methods

According to the heating rate used in the experiments, single heating rate and multi-rate heating rate methods were concluded to calculate the activation energy. In addition, the single heating rate method must determine the mechanism functions before calculating the activation energy,²⁰ which may cause errors in the process of selecting mechanism functions. However, the multi-heating rate method can avoid the selection of mechanism functions to obtain more accurate values of activation energy. For instance, Kissinger–Akahira–Sunose (KAS) method and Flynn–Wall–Ozawa (FWO) widely used to the calculation of activation energy, as eqn (1) and (2), respectively.²¹

$$\log(\beta_i) = \text{Const} - 1.052 \frac{E_\alpha}{RT_\alpha} \quad (1)$$

Table 2 Gases supplied when undergoing thermal analysis experiments

| Gas | Volume fraction (vol%) | | | | | |
|-----------------|------------------------|----|----|----|----|----|
| | 1 | 2 | 3 | 4 | 5 | 6 |
| O ₂ | 21 | 14 | 8 | 21 | 14 | 8 |
| CO ₂ | 0 | 0 | 0 | 39 | 46 | 52 |
| N ₂ | 79 | 86 | 92 | 40 | 40 | 40 |

$$\log\left(\frac{\beta_i}{T_{\alpha,i}^2}\right) = \text{Const} - \frac{E_\alpha}{RT_\alpha} \quad (2)$$

According to the analysis of Starink,²¹ KAS method developed by Starink offers more accurate estimates of E_α , as described in eqn (3):

$$\log\left(\frac{\beta_i}{(T + 273.15)_{\alpha,i}^{1.92}}\right) = \text{Const} - 1.0008 \frac{E_\alpha}{R(T + 273.15)_\alpha} \quad (3)$$

The comprehensive combustion performance index (K) was used to comprehensively evaluate the characteristics of coal spontaneous combustion, and is expressed in eqn (2):²²

$$K = \frac{(dw/dt)_{\max} \times (dw/dt)_{\text{av}}}{T_{\text{ig}}^2 \times T_b} \quad (4)$$

3. Results and discussion

3.1. Features on TG-DSC tests

According to the characteristic temperatures in Fig. 1, the mass variation of coal was divided into three stages: the low-temperature pyrolysis stage (stage 1), oxygen-adsorption stage (stage 2), and combustion stage (stage 3). The boundary temperatures for stage 2 were T_1 and T_{ig} , where T_1 is the initial temperature when the coal's mass continuously increases and T_{ig} is the ignition temperature.²⁰ Furthermore, T_2 was defined as the temperature at which the maximum exothermic peak is completed on the DSC curve. Stage 1 was from the initial temperature to T_1 . Stage 3 was from T_{ig} to T_2 .

3.2. Mass variation

Fig. 2 illustrates the TG and differential thermogravimetry (DTG) curves, and the characteristic temperatures are given in Table 3 with a heating rate of 5 °C min⁻¹.

Table 1 Proximate and ultimate analyses of the two coal samples

| Coal sample | Proximate analysis (air-dried basis, mass%) | | | | Ultimate analysis (% daf.) | | | | |
|-------------|---|-------|-----------------|--------------|----------------------------|------|------|------|------|
| | Moisture | Ash | Volatile matter | Fixed carbon | C | H | O | N | S |
| Coal A | 2.3 | 9.7 | 34.85 | 53.15 | 85.71 | 4.58 | 7.48 | 1.54 | 0.69 |
| Coal B | 1.4 | 26.02 | 8.85 | 63.73 | 89.69 | 3.61 | 1.71 | 2.85 | 2.14 |



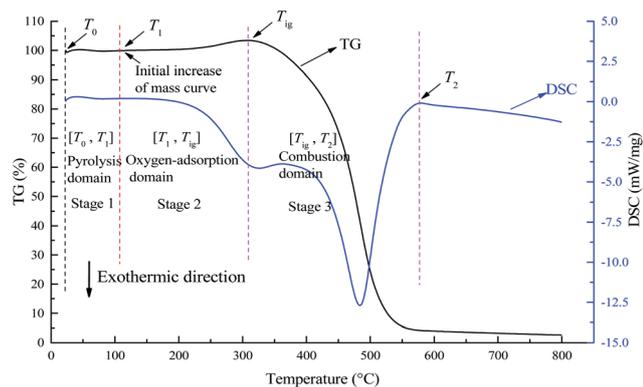


Fig. 1 TG and DSC curves of a coal sample showing the temperature boundaries for different stages. This individual diagram is obtained from the thermal analysis of coal A conversion under 21 vol% O₂/79 vol% N₂ and heating rate of 5 °C min⁻¹.

Both the increase of the CO₂ concentration and decrease of the O₂ concentration made the TG/DTG curve move to a high temperature range, as illustrated in Fig. 2. T_{ig} and T_{max} decreased with increasing O₂ concentration. For example, when the O₂ concentration was reduced from 21 to 8 vol%, the T_{ig} of coal A increased from 308.4 to 323.1 °C, and the T_{ig} of coal B increased from 309.2 to 325.5 °C (Table 3). The lower the oxygen concentration in the environment, the lower the number of activated oxygen molecules. Therefore, the opportunities of collision between activated oxygen molecules and active functional groups of coal is significantly reduced, and the reaction between coal and oxygen becomes difficult. This finding is in agreement with the reports by Deng *et al.*,¹⁴ Wang *et al.*,²³ and Liu *et al.*²⁴ When the O₂ concentration was constant, increasing the CO₂ concentration had a delay effect on the TG and DTG curves, whereas the characteristic temperatures increased along with CO₂ concentration. As Table 3 shows, when the O₂ concentration was 21 vol% and the CO₂ concentration was increased from 0 to 52 vol%, the ignition temperature of coal A increased from 323.1 to 325.5 °C and the ignition temperature of coal B increased from 382.3 to 387.4 °C. The diffusion rate of O₂ in CO₂ is lower than that in N₂, which has an adverse effect on the transportation of O₂ from the gas mixture to the coal particle surface.^{23,25} In addition, CO₂ is much easier for coal to

physically adsorb than O₂ and N₂.^{26,27} Therefore, adsorption of CO₂ occupies some of the active sites on the coal surface, inhibiting the ability of O₂ and coal to interact.

3.3. Exothermic properties

3.3.1. Differential scanning calorimetry. The DSC and first-order differential of DSC (DDSC) curves for the two coal samples from 30 to 800 °C were illustrated in Fig. 3. Total heat release is presented in Table 4.

Fig. 3 shows that exothermic peaks occur when the temperature exceeds 200 °C. The exothermic process finished when the temperature was more than 600 °C. The DSC curve of coal exhibited typical bimodal exothermic behavior. However, the first exothermic peak of the DSC curves for coal B was not clear. This phenomenon is caused by differences in the types and quantities of active structures contained in coal. Some active structures, such as alkyl chains, methyl groups, carboxyl groups, and hydroxyl groups, are easily oxidized at lower temperatures.^{28,29} However, the combustion or decomposition temperature of some stable structures, such as aromatic structures, is higher.³⁰ When the concentration of CO₂ was zero, the maximum heat intensity and total heat increased significantly with an increase in the O₂ concentration. When the O₂ concentration was increased from 8 to 21 vol%, the total heat released by coal rose from 17.296 to 19.247 kJ g⁻¹, as listed in Table 4. When the O₂ was sufficient, the coal-oxygen reaction was more intense. With the increased O₂ concentration, the coal and O₂ were able to have more contact; therefore, the oxidation/combustion reactions of the coal produced a high proportion of stable oxidation products, increasing heat intensity. When the O₂ concentration was kept constant, the DSC and DDSC curves moved to the high temperature region, whereas total heat release decreased substantially with increasing CO₂ concentration. As Lee³¹ noted, competitive adsorption occurs in coal when various gases exist. The adsorption of CO₂ causes a reduction in the amount of O₂ contacting coal. In addition, more unstable products are formed, leading to an apparent reduction in heat intensity.

3.3.2. Stage characteristics of heat release. The DSC curves for the two samples at a heating rate of 5 °C min⁻¹ and in O₂:N₂:CO₂ = 21:79:0 and O₂:N₂:CO₂ = 21:40:39 atmospheres are illustrated in Fig. 4. The sub-peaks of the DSC

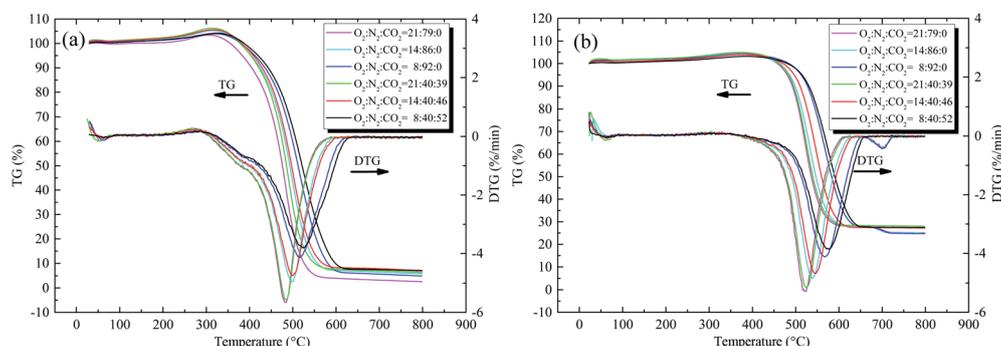


Fig. 2 TG and DTG curves of the two coal samples with the heating rate of 5 °C min⁻¹ under different atmospheres, (a) coal A and (b) coal B.



Table 3 Characteristic temperatures during combustion of coal with heating rate of 5 °C min⁻¹

| Sample Atmosphere (O ₂ : N ₂ : CO ₂) | Coal A | | | | Coal B | | | |
|---|---------------------|----------------------|-----------------------|---------------------|---------------------|----------------------|-----------------------|---------------------|
| | T ₁ (°C) | T _{ig} (°C) | T _{max} (°C) | T ₂ (°C) | T ₁ (°C) | T _{ig} (°C) | T _{max} (°C) | T ₂ (°C) |
| 21 : 79 : 0 | 81.3 | 308.4 | 482.9 | 579.5 | 83.35 | 367.6 | 523.2 | 617.9 |
| 14 : 86 : 0 | 79.1 | 313.6 | 500.1 | 591.1 | 81.45 | 370.1 | 539.7 | 627 |
| 8 : 92 : 0 | 83.1 | 323.1 | 516.2 | 616.4 | 70.25 | 382.3 | 567.1 | 649.2 |
| 21 : 40 : 39 | 71.9 | 309.2 | 485.4 | 583.4 | 84.55 | 371.7 | 524.9 | 618.8 |
| 14 : 40 : 46 | 80.6 | 317.9 | 499.7 | 598.7 | 81.45 | 378.6 | 546 | 634.7 |
| 8 : 40 : 52 | 75.7 | 325.5 | 523.9 | 627.5 | 80.55 | 387.4 | 576.5 | 660.2 |

curves were fitted using eqn (5) (Gaussian equation). The fitting parameters of the sub-peaks are given in Table 4. Fig. 5 illustrates the ratio of the peak 1 to peak 2 areas.

$$v_m = v_0 + \frac{S \left(\left(-4 \ln(2)(t - t_c)^2 \right) / w^2 \right)}{w \sqrt{\pi} / (4 \ln(2))} \quad (5)$$

The temperature corresponding to the maximum of peak 1 in coal A was in the range of 320–360 °C, whereas it was in the range of 410–440 °C for coal B, as listed in Table 4. Furthermore, the temperature corresponding to the maximum of peak 1 was close to the ignition temperature of coal. Some of the active structures in coal, such as alkyl chains, methyl groups, carboxyl groups, and hydroxyl groups, are oxidized at low temperatures, resulting in a clear oxidative-exothermic peak before coal combustion.³⁰ Peak 2 was caused by the breaking of numerous stable structures (such as aromatic rings) into small molecules that participated in the combustion reaction.³² As illustrated in Fig. 5, the ratio for coal A was clearly larger than that for coal B, which indicates that attribution of peak 1 to the total heat release was greater than coal B. The primary reason for this phenomenon is that the metamorphic grade of coal A was lower than that of coal B, and the molecules of coal A contained numerous alkyl chains, methyl groups, carboxyl groups, hydroxyl groups, and other active structures.³³ The ratio decreased with a decrease of the O₂ concentration or increase of the CO₂ concentration, which indicates that a low O₂ concentration or high CO₂ concentration inhibits oxidation of some reactive groups in coal.

3.4. Comprehensive combustion performance index

The dependence of *K* on the O₂ concentration for the two samples is illustrated in Fig. 6. The comprehensive combustion performance index of the two samples decreased with a decrease in the O₂ concentration; furthermore, the higher the O₂ concentration was, the more O₂ participated in coal combustion, leading to the combustibility being strengthened. For the same O₂ concentration, the addition of CO₂ resulted in a significant decrease in the comprehensive combustion performance index. The physical adsorption of CO₂ onto the coal surface occupied active sites, which hindering the physical adsorption and chemical adsorption of O₂ by coal, and further inhibiting coal oxidation and combustion.

3.5. Apparent activation energy

The calculation of the apparent activation energy uses the Starink method, which is a multiple heating rates calculation method. Therefore, a set of data points of multiple heating rates (5, 10 and 15 °C) calculated at intervals of 0.5 from 0.5 to 0.95 was obtained according to eqn (3), and $\ln(\beta/(T + 273.15)^{1.92})$ and $1/(T + 273.15)$ were calculated at different conversion rates as shown in Fig. 7. The apparent activation energies during stages 2 and 3 for the two coal samples are given in Fig. 8 and 9.

3.5.1. Oxygen-adsorption stage. The physical and chemical adsorption of O₂ and CO₂ is easier and the energy required for the chemical reaction between O₂ and coal is lower at the beginning of the oxygen-adsorption stage. The reaction between coal and oxygen is a step-by-step reactive process in which functional groups are gradually activated and oxidized. At low temperatures, only some active structures are activated and

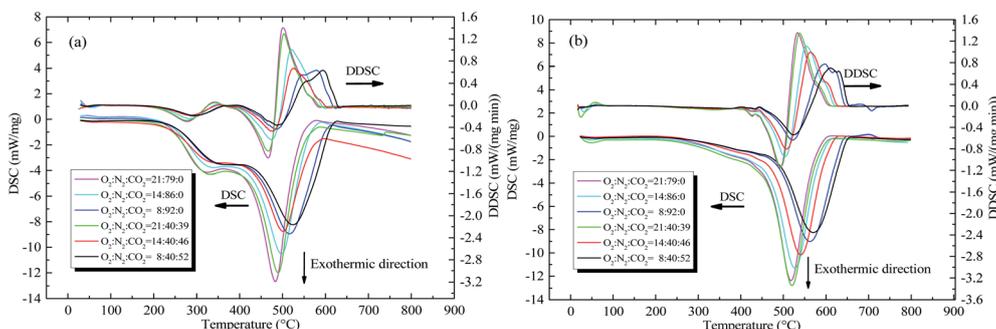


Fig. 3 DSC and DDSC curves of the two coal samples with the heating rate of 5 °C min⁻¹ under different atmospheres, (a) coal A and (b) coal B.



Table 4 Factors used to fit sub-peaks for DSC curves and total heat release and the maximum of exothermic peak with heating rate of 5 °C min⁻¹

| Coal sample | Atmosphere (O ₂ : N ₂ : CO ₂) | Peak | ν_0 | S | t_c | w | R^2 | Q (kJ g ⁻¹) |
|-------------|---|------|---------|----------|---------|---------|--------|---------------------------|
| Coal A | 21 : 79 : 0 | 1 | 0.306 | -172.495 | 66.473 | 35.291 | 0.9896 | 19.247 |
| | | 2 | | -162.325 | 90.641 | 13.913 | | |
| | 14 : 86 : 0 | 1 | 0.156 | -158.973 | 68.084 | 35.237 | 0.9939 | 17.891 |
| | | 2 | | -159.263 | 93.457 | 15.907 | | |
| | 8 : 92 : 0 | 1 | 0.042 | -139.887 | 69.772 | 34.700 | 0.9972 | 17.296 |
| | | 2 | | -175.342 | 97.945 | 19.815 | | |
| | 21 : 40 : 39 | 1 | -0.141 | -162.444 | 66.061 | 34.582 | 0.9889 | 18.177 |
| | | 2 | | -158.456 | 90.895 | 14.639 | | |
| | 14 : 40 : 46 | 1 | -0.274 | -118.875 | 67.107 | 33.344 | 0.9942 | 17.893 |
| | | 2 | | -150.177 | 94.420 | 18.585 | | |
| | 8 : 40 : 52 | 1 | -0.093 | -131.845 | 70.106 | 34.198 | 0.9938 | 17.794 |
| | | 2 | | -172.622 | 99.379 | 20.963 | | |
| Coal B | 21 : 79 : 0 | 1 | 0.038 | -103.340 | 82.755 | 38.798 | 0.9972 | 15.308 |
| | | 2 | | -138.154 | 97.773 | 12.381 | | |
| | 14 : 86 : 0 | 1 | -0.165 | -89.947 | 83.914 | 39.612 | 0.9984 | 13.812 |
| | | 2 | | -141.981 | 100.100 | 13.747 | | |
| | 8 : 92 : 0 | 1 | 0.004 | -68.126 | 84.428 | 39.436 | 0.9967 | 13.018 |
| | | 2 | | -152.792 | 106.658 | 16.9603 | | |
| | 21 : 40 : 39 | 1 | -0.200 | -106.997 | 83.199 | 39.943 | 0.9976 | 14.779 |
| | | 2 | | -144.890 | 98.404 | 12.736 | | |
| | 14 : 40 : 46 | 1 | 0.021 | -76.145 | 83.285 | 37.229 | 0.9983 | 13.197 |
| | | 2 | | -144.658 | 102.357 | 14.770 | | |
| | 8 : 40 : 52 | 1 | -0.081 | -68.306 | 87.135 | 42.810 | 0.9948 | 12.476 |
| | | 2 | | -143.390 | 108.245 | 17.740 | | |

react with O₂. Moreover, these oxidation reactions require less reaction energy, resulting in lower apparent activation energy of coal. As the temperature increases, more stable structures are

gradually activated and react with O₂. These oxidation reactions require more energy, leading to an increase in apparent activation energy. Therefore, we observed that with an increase of

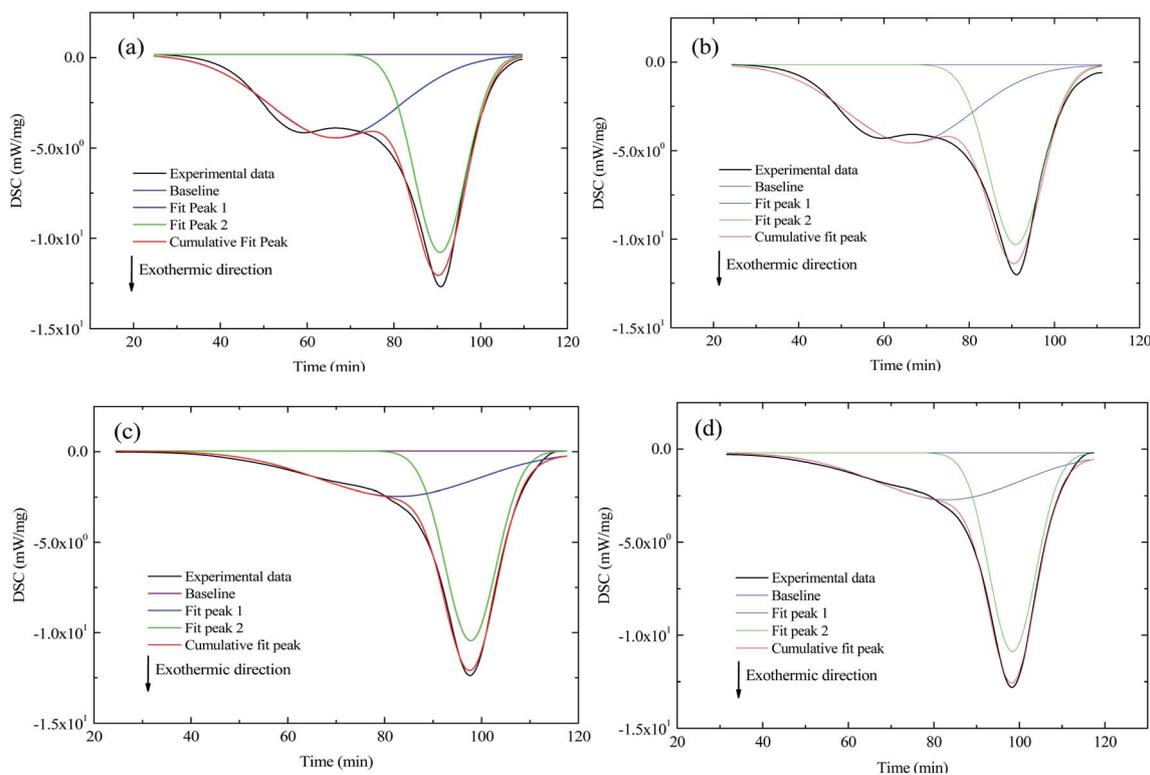


Fig. 4 DSC curves from the pyrolysis domain stage (stage 1) through the combustion domain stage (stage 3) for the two coal samples, (a) coal A in atmosphere O₂ : N₂ = 21 : 79, (b) coal A in atmosphere O₂ : N₂ : CO₂ = 21 : 40 : 39, (c) coal B in atmosphere O₂ : N₂ = 21 : 79, and (d) coal B in atmosphere O₂ : N₂ : CO₂ = 21 : 40 : 39.



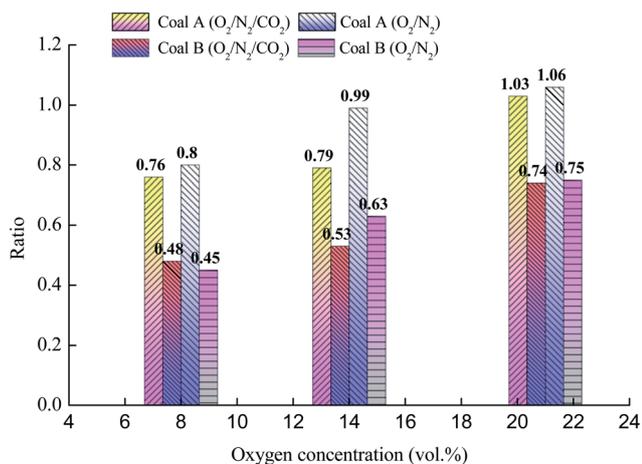
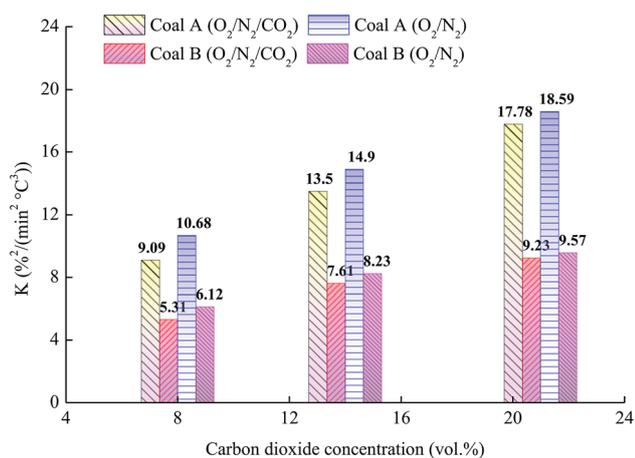


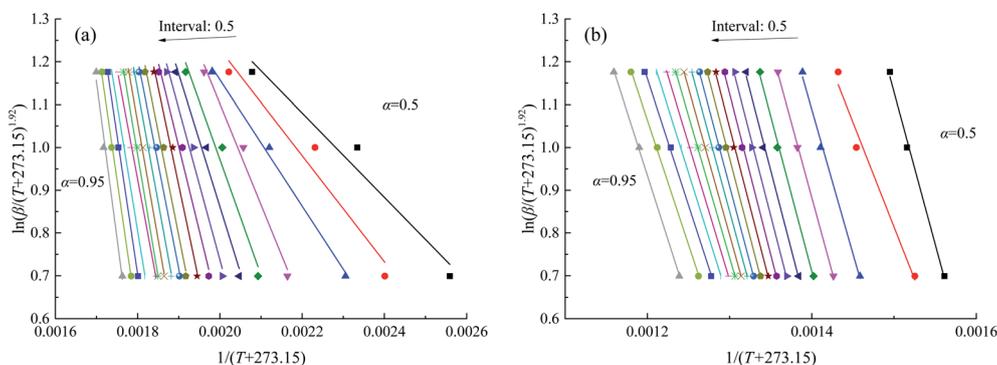
Fig. 5 Ratio of the peak 1/peak 2 areas.

Fig. 6 Relationship between the comprehensive combustion performance index and O₂ concentration.

the conversion rate, apparent activation energy increased. Furthermore, the tendency and value of apparent activation energy changed relative to the conversion rate at all percentages of O₂ in O₂/N₂ or O₂/CO₂/N₂, while tendency of apparent

activation energy was almost the same, as illustrated in Fig. 8. When the CO₂ concentration was 0, the apparent activation energy decreased with an increase in the O₂ concentration. In light of the criteria of competitive reactions, the primary reason for this behavior is that the active structures consume a considerable amount of O₂. The decrease of the O₂ concentration inhibits the reaction of the stable structures in coal with O₂, which requires more energy. In addition, coal B was anthracite, which provides more surface area for strongly competitive chemisorption,³⁴ resulting in complex changes of the apparent activation energies. When the O₂ concentration was the same, the apparent activation energy of coal oxidation in O₂/N₂ was greater than that in O₂/CO₂/N₂. Liu *et al.*²⁴ considered chemisorption in O₂/CO₂ conditions to be mainly controlled by O₂ chemisorption at the oxygen-adsorption stage. Moreover, CO₂ competes with O₂ when they are adsorbed by active sites on the coal surface, which hinders the chemical reaction of coal and O₂. In summary, O₂ competed with CO₂ for active sites; thus, more energy was required.

3.5.2. Combustion stage. The apparent activation energy of coal A first increased and then decreased with an increase in the conversion rate. Coal B exhibited a similar tendency when the O₂ concentration was more than 8 vol%. When the O₂ concentration was 8 vol%, the apparent activation energy of coal B decreased with an increase in the conversion rate, as illustrated in Fig. 9. Temperatures when the apparent activation energy reached the maximum were close to the devolatilization temperature of coal. When the temperature exceeded the devolatilization temperature, a large volume of volatiles evaporated out, leading to an increase in the diffusion resistance of O₂ to the coal pores.³⁵ Then, the apparent activation energy of coal combustion tended to decrease due to the large number of active structures resulting from pyrolysis. Values of the apparent activation energy in N₂/O₂ atmospheres were larger than those in O₂/N₂/CO₂ atmospheres. The cause of this phenomenon is that competition chemisorption of O₂ and CO₂ on the coal surface during combustion in the O₂/N₂/CO₂ atmosphere, which causes the chemical adsorption process of oxygen in the O₂/N₂/CO₂ atmosphere significantly longer than in the O₂/N₂ atmosphere.²⁴ Consequently, the probability of oxygen contacting with the active structure is reduced.

Fig. 7 Linear relationship between $\ln(\beta/(T + 273.15)^{1.92})$ and $1/(T + 273.15)$ of coal A in atmosphere O₂ : N₂ : CO₂ = 21 : 40 : 39, (a) in oxygen-adsorption stage and (b) combustion domain stage.

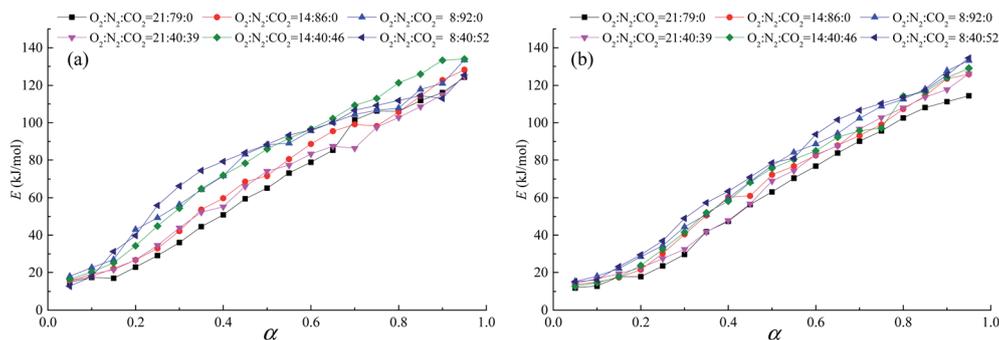


Fig. 8 Relationship between apparent activation energy and conversion during the oxygen-adsorption stage (stage 2), (a) coal A and (b) coal B.

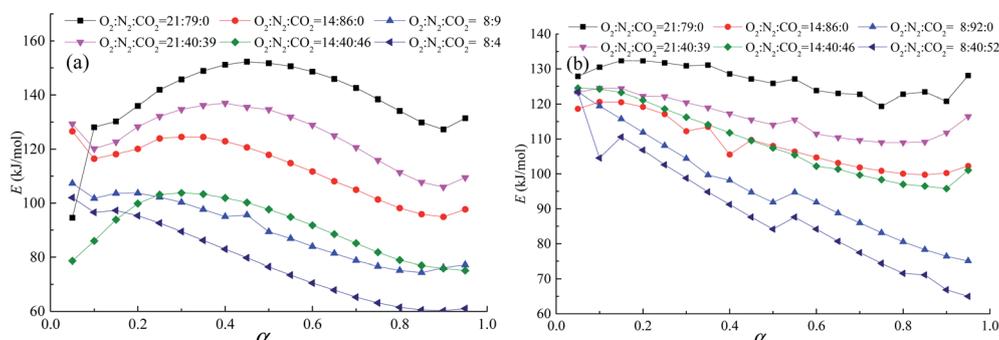


Fig. 9 Relationship between apparent activation energy and conversion during the combustion domain stage (stage 3), (a) coal A and (b) coal B.

Therefore, the combustion rate of coal in the $O_2/N_2/CO_2$ atmosphere is lower than that in the O_2/N_2 atmosphere (Fig. 2).

4. Conclusions

• With an increase in the CO_2 concentration or decrease in the O_2 concentration, the thermal analysis curves of coal were delayed and the characteristic temperatures increased.

• The DSC curve can be divided into two exothermic peaks, and the ratio of the peak 1 to peak 2 areas decreased with the addition of CO_2 , which indicated that CO_2 inhibited the oxidation of active functional groups of coal structures. Furthermore, the total heat release decreased. Higher concentrations of CO_2 or lower concentrations of O_2 resulted in a lower comprehensive combustion performance index.

• At the oxygen-adsorption stage, the apparent activation energy increased with an increase in the conversion rate, whereas it first increased and then decreased at the combustion stage. The apparent activation energy in the $O_2/CO_2/N_2$ atmosphere was less than that in the O_2/N_2 atmosphere.

Nomenclature

| | |
|------------|---|
| E_α | Apparent activation energy ($J mol^{-1}$) |
| I | Constant |
| K | Comprehensive combustion performance index ($\%^2 min^{-2} \text{ } ^\circ C^{-3}$) |

| | |
|-----------------|--|
| Q | Total heat release ($kJ mg^{-1}$) |
| R | Gas constant ($J mol^{-1} K^{-1}$) |
| S | Peak area ($mW min mg^{-1}$) |
| T | Temperature ($^\circ C$) |
| T_1 | Initial temperature when the coal's mass continuously increases ($^\circ C$) |
| T_2 | Temperature is exothermic completed on DSC curve ($^\circ C$) |
| T_b | Burnout temperature ($^\circ C$) |
| t_c | Center of the peak time (min) |
| T_{ig} | Ignition temperature ($^\circ C$) |
| v_o | Baseline of peak ($mW mg^{-1}$) |
| w | Full width at half maximum (min) |
| v_m | Change rate of exothermic (min) |
| $(dw/dt)_{av}$ | Average rate of mass loss ($\% min^{-1}$) |
| $(dw/dt)_{max}$ | Maximum rate of mass loss ($\% min^{-1}$) |
| α | Conversion rate (%) |
| β | Heating rate ($^\circ C min^{-1}$) |

Conflicts of interest

There are no conflicts to declare.

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