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A novel *in situ* N₂ generation system assisted by authigenic acid for formation energy enhancement in an oilfield†

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The *in situ* N₂ generation technique is a promising method for enhancing oil recovery (EOR) and blockage removal. Conventionally, N₂ is generated through the redox reactions of sodium nitrite (NaNO₂) and ammonium chloride (NH₄Cl). The effect of several parameters on N₂ production was studied in this work. The reaction kinetics between NaNO₂ and NH₄Cl were developed. The results revealed that the optimal molar ratio was 1 : 1 for the reactants and a higher concentration of reactants, H⁺ and higher temperature would lead to an increase in the gas production volume and a better gas-generating efficiency. The kinetic equation of the N₂ generation reaction was defined as $dc/dt = -7.103 \times 10^7 C_{H^+}^{1.3291} C_O^{2.0949} e^{(-51.28 \text{ kJ mol}^{-1})/RT}$. Methyl formate, ethyl formate, ammonium persulphate as well as paraformaldehyde were selected as authigenic acids by measuring their pH values and their effect on the rate of gas-generation, and the gas production volume was studied. An energy enhancement experiment was conducted to examine the influence of reactant concentration, initial pressure and different authigenic acids on energy enhancement and gas-generating efficiency. The results showed that the increase in reactant concentration can improve the energy enhancement while neither the reactant concentration nor the initial pressure could hardly influence the gas-generating efficiency. The presence of authigenic acids can considerably enhance both the gas production volume and the gas-generating efficiency of the system. The ammonium persulphate assisted N₂ generation system is considered to be suitable for field applications.

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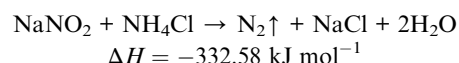
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1. Introduction

In situ gas generation technology has been an increasingly effective method to enhance oil recovery¹ and to reform the formation.^{2,3} The commonly used *in situ* gas generation systems can generally be divided into two types: *in situ* CO₂ generation system⁴ and *in situ* N₂ generation system.⁵ The most widely used *in situ* N₂ generation system consists of sodium nitrite (NaNO₂) and ammonium chloride (NH₄Cl),⁶ which has been applied in reducing the viscosity of paraffin-rich crude oil,⁷ improving the matrix acidizing,⁸ the removal of hydrate plug under subsea⁹ and enhancing the formation energy¹⁰ due to the remarkable generation of gas and heat. This highly exothermic chemical reaction is represented as:



An *in situ* N₂ generation system was first proposed by the Shell company^{11,12} to replace the vaporized liquified nitrogen, which required a convenient source of liquid nitrogen and suitable high-pressure equipment. Considering the remarkable thermal performance, this system has been an effective alternative method towards steam and hot oil treatments to stimulate the production from paraffin plugged wells in California and Wyoming.¹³ Dom Joao field in Brazil adopted a new method called SGN/EMULSAO to stimulate the wells and the water/oil emulsified system could extend the reaction time and increase the radius of influence.¹⁴ Another application of the *in situ* N₂ generation system was to promote the effect of the acidizing operation by heating the acid fluid.⁸ With the development of marine oil and gas, the *in situ* N₂ generation system was applied in the removal of hydrate and paraffin blockages by heating subsea production pipes because low temperature was considered as the major element in the organic solid deposition.¹⁵ Recently, *in situ* steam generation assisted with the N₂ release reaction between NaNO₂ and NH₄Cl was introduced for

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heavy oil production by mobilizing the low API crude oil.¹⁶ For an oilfield with low permeability as well as low formation energy, $\text{NaNO}_2/\text{NH}_4\text{Cl}$ system was also used to enhance the formation energy to successfully boost the oil production, *e.g.* in Jiangnan oilfield.¹⁷

In the N_2 generation system, H^+ was found to have a key effect on the reaction rate, reaction degree and the thermal effect of the system.¹⁸ In order to acquire a better performance of the *in situ* N_2 generation system, some additives such as hydrochloric acid and citric acid were used to boost the reaction in order to get a much higher temperature and more gas volume. Singh *et al.*¹⁹ and Nguyen *et al.*²⁰ treated the $\text{NaNO}_2/\text{NH}_4\text{Cl}$ system as a fused chemical reaction by adding acid-filled capsules, which could release citric acid slowly. Apart from that, some chemicals, like methyl formate and tetrachloroethane,^{21,22} which can generate H^+ at certain conditions were also used widely to strengthen the system. Nevertheless, some drawbacks accompanied with uncontrolled H^+ led to an intensive reaction rate and side reactions, which resulted in by-products (*e.g.* NO and NO_2).²³ In order to make the best use of the catalytic H^+ , the concentration boundary of H^+ for promoting the main reaction has to be determined accurately. Dusenbury *et al.*²⁴ reported that the $\text{NaNO}_2/\text{NH}_4\text{Cl}$ system would not generate NO_x at $\text{pH} = 1\text{--}6$ according to their study on the kinetics and the mechanism. Wu Wengang *et al.*²⁵ suggested that the *in situ* N_2 generation system can generate N_2 at room temperature without NO_x at $\text{pH} = 3$. Besides, Nguyen *et al.*²⁶ focused on the functions of H^+ , temperature and reactants in the reaction of the $\text{NH}_4^+/\text{NO}_2^-$ system and presented the predicted kinetics equation at $\text{pH} = 3\text{--}5$. The reported ranges of pH were quite inconsistent, which would confuse readers. Thus, a worker cannot choose the appropriate acid or the latent acid in the application of the *in situ* N_2 generation system. This study was carried out to figure out the effect of each parameter on the gas production volume and to form an efficient authigenic acid assisted N_2 generation system.

In this work, the effect of several reaction parameters (concentration of reactants, H^+ and the temperature) on the N_2 -producing ability of the $\text{NaNO}_2/\text{NH}_4\text{Cl}$ mixture was studied. Meanwhile, the kinetics study was also conducted. Four chemicals that served as authigenic acids were investigated through pH measurements and their capacity of assisting the *in situ* N_2 generation system was examined by an energy enhancement test. A novel ammonium persulphate assisted N_2 generation system that can be used for energy enhancement in an oilfield was proposed.

2. Materials and methods

2.1. Chemicals

Chemicals used in this study include, sodium nitrite (NaNO_2), ammonium chloride (NH_4Cl), hydrogen chloride (HCl), methyl acetate (MA), ethyl acetate (EA), paraformaldehyde (PFA) and ammonium persulphate (APS); the above chemicals were analytical reagents and purchased from Sinopharm Chemical Reagent Co., Ltd. Methyl formate (MF) and ethyl formate (EF) were chemically pure and purchased from Sinopharm Chemical

Reagent Co., Ltd, and sodium chloroacetate (SC) of analytical reagent was purchased from Aladdin, Shanghai, China. Distilled water was used for dissolving the above compounds.

2.2. Experiments

2.2.1. Gas-producing reaction test. Fig. 1 is the schematic diagram of the gas-producing reaction experiment. Different concentrations of NaNO_2 and NH_4Cl solutions (50 mL each) were prepared separately in beakers and then they were placed in a thermostat to maintain the same initial temperature. The concentration of H^+ in the reaction was adjusted with HCl solution. The flasks were placed in the water bath and heated beforehand to minimize the influence of the gas expansion. NaNO_2 and NH_4Cl solutions were mixed in the flask and stirred magnetically for 4 h. A measuring cylinder was used to quantify the volume of the gas.

2.2.2. The pH measurement of various authigenic acids. The test was carried out to evaluate the H^+ generating characteristic of authigenic acids. Aqueous solutions of seven authigenic acids (MF , EF , MA , EA , SC , PFA , and AP) in different concentrations (0.001, 0.005, 0.01, 0.05, 0.1, and 0.5 M) were prepared in Erlenmeyer flasks at 20°C and the pH values of the aqueous solutions were measured by a pH meter. The prepared solutions were then put in a water bath and the pH was monitored at different temperatures within 2 h.

Several authigenic acids were selected through the pH measurement to form new N_2 generation systems. Then, the capability of the systems with authigenic acids to generate gas was investigated through gas-producing experiments.

2.2.3. Energy enhancement experiments. Fig. 2 illustrates the layout of the energy enhancement experiment. In order to simulate the real formation conditions, the experiment was performed at various pressures in a high-temperature reactor. NaNO_2 and NH_4Cl solutions were injected into the reactor one after the other at 20°C . The reactor was then heated rapidly and the reactants began to release gas. The pressure and the temperature of the reactor were recorded, and all the pressures were measured by a pressure gauge.

The detailed procedure is as follow: (1) reactant solution (volume: 150 mL) was injected into the reactor (volume: 300 mL) at 20°C ; (2) the pressure of the reactor was adjusted to the

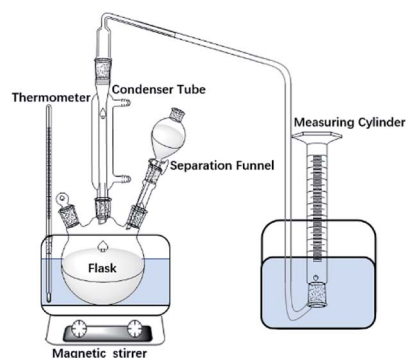


Fig. 1 The schematic diagram of gas production experiments.



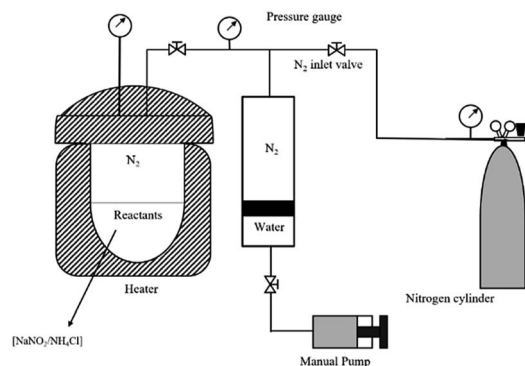


Fig. 2 The schematic diagram of energy enhancement experiments.

targeted pressure by injecting N_2 ; (3) the reactor's temperature was increased to 100 °C and heated for 4 h while magnetically stirring at 300 rpm. Afterwards, the system was cooled naturally to 25 °C. Temperature and pressure of the experiment were recorded continuously.

3. Results and discussion

3.1. Evaluation of the gas-producing reaction

3.1.1. Molar ratio of $NaNO_2/NH_4Cl$. The molar ratio of reactants is one of the most essential factors that influence the gas production volume (GPV). As mentioned in Fig. 1, the GPV was measured at atmospheric pressure (1 atm). The effect of the molar ratio ($NaNO_2/NH_4Cl = 1.5 : 1.5, 2.0 : 1.0, 1.0 : 2.0, 1.0 : 1.5, 1.5 : 1.0$) on the GPV in the presence of 0.0178 M HCl at 30 °C is shown in Fig. 3. As the total concentration of the reactants increased from 2.5 M to 3 M, the GPV rose from 1550 mL to 1870 mL. Under the circumstances of the same total concentration of 3.0 M ($NaNO_2/NH_4Cl = 1.5 : 1.5, 2.0 : 1.0, 1.0 : 2.0$), the tested group whose molar ratio of reactants was 1.5 : 1.5 had a maximum GPV compared to other groups ($NaNO_2/NH_4Cl = 2.0 : 1.0, 1.0 : 2.0$). Also, the maximum GPV per unit mass was found at a molar ratio of 1 : 1. This means the reaction had better performance to generate gas when the molar ratio of $NaNO_2$ and NH_4Cl was 1 : 1. Therefore, the optimal molar ratio of reactants was 1 : 1 and this molar ratio was applied in the following experiments.

3.1.2. Concentration of the reactant. In order to determine the optimal concentration of the reactant, GPV and Gas-Generating Efficiency (GGE) were introduced as two evaluation criteria. The GGE was calculated as follows:

$$GGE = \frac{\text{Real GPV (mL)}}{\text{Theoretical GPV (mL)}} \times 100 \quad (1)$$

The effect of the reactant concentration (0.5–3.0 M for each reactant) on the total GPV at 30 °C is shown in Fig. 4. As the concentration of each reactant was increased from 0.5 to 3.0 M, the GPV soared from 505 to 3910 mL. This is because the concentration of NO_2^- and NH_4^+ increased with the increase in the reactant concentration, leading to a rapid increase in the GPV. Meanwhile, GGE generally went up from 40.99 to 56.82%

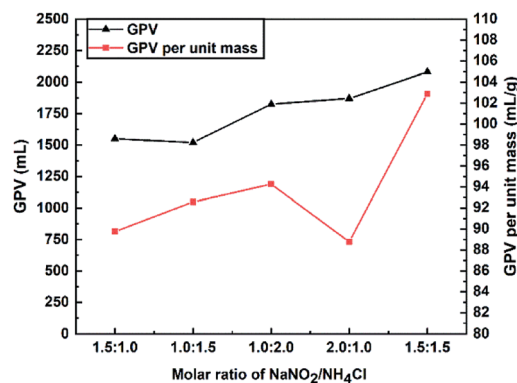


Fig. 3 GPV and GPV per unit mass vs. molar ratio with 0.0178 M HCl at 30 °C.

as the concentration of the $NaNO_2/NH_4Cl$ system was increased from 0.5 to 2.5 M. However, a slight decrease of GGE at 3.0 M (from 56.82% to 52.90%) was observed. The GGE seems to remain at around 55% once the reactant concentration is over 1.5 M. As a result, 1.5 M was recognized as the appropriate concentration for each reactant in the following experiments.

3.1.3. Concentration of H^+ . The effect of different $c(H^+)$ from 0.0001 M to 0.0356 M on the GPV at 30 °C in the presence of 1.5 M each reactant is illustrated in Fig. 5. As the $c(H^+)$ was increased from 0 to 0.0356 M, the GPV rose from 0 to 2435 mL and the GGE experienced a considerable growth. When $c(H^+)$ is higher than 0.001 M, it would promote the reaction more strongly.

Nguyen *et al.*²⁶ found that the activation energy of the reaction does not change when the $c(H^+)$ is varied by two orders of magnitude and they thought that H^+ would affect the concentration of the reactants by influencing the rate-limiting step. Due to the presence of NO_2^- , which could combine with H^+ to form an unstable substance (HNO_2), the reaction is prone to produce by-products, such as NO and NO_2 , which are harmful for the workers and the environment. According to the results of the experiment, when the $c(H^+)$ reaches 0.0356 M, a reddish brown gas (NO_x) could be observed. Therefore, the concentration of H^+ was set at 0.0178 M in the following experiments.

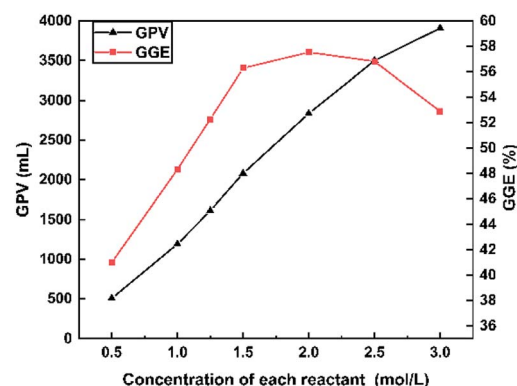


Fig. 4 GPV and GGE vs. concentrations of each reactant in the presence of 0.0178 M HCl at 30 °C.



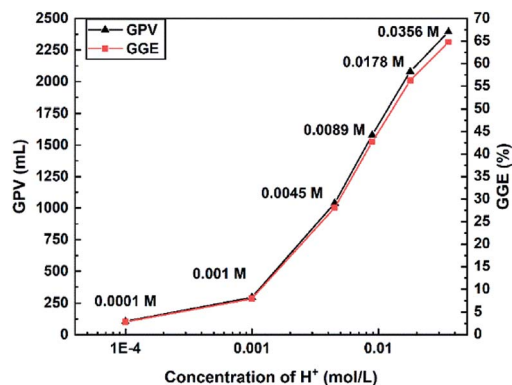


Fig. 5 GPV and GGE vs. $c(\text{H}^+)$ in the presence of 1.5 M of each reactant at 30 °C.

3.1.4. Temperature. GPV and GGE at different temperatures (from 30 to 70 °C) in the presence of 1.5 M of each reactant and 0.01780 M HCl are illustrated in Fig. 6. As expected, GPV and GGE increased regularly with the rise of temperature. This is because the activity of reactants will rise at higher temperatures and the activated reactants would have more opportunities to attack each other.

3.1.5. Reaction kinetics study. The kinetic equation of the reaction reflects the relationship between the reaction rate and the parameters. In the practical field application, the gas-generating rate of the *in situ* N_2 generation system can be predicted by the kinetic equation of the $\text{NaNO}_2/\text{NH}_4\text{Cl}$ mixture according to certain formation conditions.

The kinetic equation of the reaction between NaNO_2 and NH_4Cl can be expressed as eqn (2).

$$\frac{dc}{dt} = -Ae^{\left(-\frac{\Delta E}{RT}\right)} c_{\text{NaNO}_2}^{n_1} c_{\text{NH}_4\text{Cl}}^{n_2} \quad (2)$$

where A and ΔE (kJ mol^{-1}) are the frequency factor and the apparent Arrhenius activation energy; n_1 and n_2 are reaction orders of NaNO_2 and NH_4Cl . c_{NaNO_2} and $c_{\text{NH}_4\text{Cl}}$ (M) are the concentrations of NaNO_2 and NH_4Cl ; R is the ideal gas constant ($8.314 \text{ J mol}^{-1} \text{ K}^{-1}$) and T (K) is the absolute temperature.

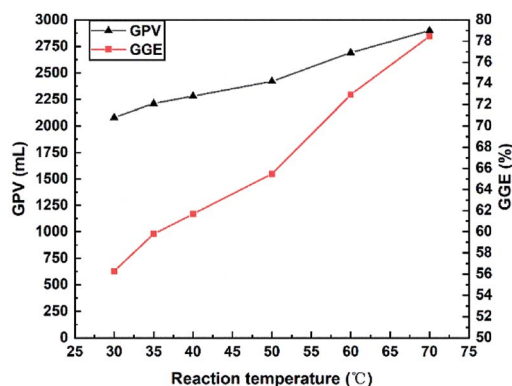


Fig. 6 GPV and GGE vs. reaction temperature with 1.5 M of each reactant and 0.01780 M HCl.

According to the previous experiments, the concentration of H^+ had a significant impact on the reaction rate and GPV, but the effect of H^+ is not reflected in eqn (2). Nguyen *et al.*²⁶ have reported the effect of H^+ based on eqn (2) and the kinetic equation was shown as follow:

$$\frac{dc}{dt} = Ae^{\left(-\frac{\Delta E}{RT}\right)} c_{\text{H}^+}^\alpha c_{\text{HNO}_2}^\beta c_{\text{NH}_4\text{Cl}}^\gamma \quad (3)$$

where α , β , and γ are the reaction orders of H^+ , NaNO_2 and NH_4Cl ; c_{H^+} (M) is the concentration of H^+ .

Wu Anming *et al.*²⁷ considered that the reaction rate constant (K_T) of the same reaction increased with the rise of the concentration of H^+ and established the reaction kinetic equation as:

$$\frac{dc}{dt} = -1.26 \times 10^7 c_{\text{H}^+}^{1.17} c_0^{1.88} e^{\left(-\frac{46.81 \text{ kJ mol}^{-1}}{RT}\right)} \quad (4)$$

where c_0 (M) is the concentration of NaNO_2 and NH_4Cl .

According to the reported results, the kinetic equation could be expressed as follows:

$$\frac{dc}{dt} = -Ac_{\text{H}^+}^m c_{\text{NaNO}_2}^{n_1} c_{\text{NH}_4\text{Cl}}^{n_2} e^{\left(-\frac{\Delta E}{RT}\right)} \quad (5)$$

where m , n_1 and n_2 are the reaction orders of H^+ , NaNO_2 and NH_4Cl .

The concentration variation of the two reactants was transformed into the volume of produced gas in this study. c_0 (the initial concentration of reactants) was used to indicate c_{NaNO_2} and $c_{\text{NH}_4\text{Cl}}$ due to the same concentration of the two reactants in the experiments. Eqn (5) can be transformed as eqn (6).

$$\frac{dc}{dt} = -Ac_{\text{H}^+}^m c_0^n e^{\left(-\frac{\Delta E}{RT}\right)} \quad (6)$$

where $n = n_1 + n_2$.

Transposing and integrating eqn (6) gives:

$$t = \frac{k^{n-1} - 1}{n - 1} A^{-1} c_{\text{H}^+}^{-m} c_0^{1-n} e^{\left(\frac{\Delta E}{RT}\right)} \quad (7)$$

Taking logarithm of eqn (6) and the result is shown as eqn (8):

$$\ln t = \ln \left(\frac{k^{n-1} - 1}{n - 1} \right) - \ln A - m \ln c_{\text{H}^+} + (1 - n) \ln c_0 + \frac{\Delta E}{RT} \quad (8a)$$

$$\lg t = \lg \left(\frac{k^{n-1} - 1}{n - 1} \right) - \lg A - m \lg c_{\text{H}^+} + (1 - n) \lg c_0 + \frac{\Delta E}{2.303RT} \quad (8b)$$

where k is the ratio of the initial concentration and the current concentration of each reactant (theoretical GPV/(theoretical GPV – real-time GPV)), c_0 (M) is the initial concentration of NaNO_2 and NH_4Cl , t (s) is the time consumed for collecting a certain gas volume.

The unknown parameters in eqn (6) are A , m , n , and ΔE and these parameters can be obtained through eqn 8(a) and (b) and the results of the gas-producing reaction test. A linear relationship could be found between $\ln t$ and $1/T$ when k , c_0 , c_{H^+} were



the same and the slope was $\Delta E/R$. The same linear relationship could also be found between $\ln t$ and $-\ln c_{H^+}$ or $\ln c_0$.

The time that the reaction generated a certain gas volume (500, 750, 1000, 1250 mL) was recorded at various temperatures (303–333 K) with the presence of 1.5 M of each reactant and 0.0178 M H^+ . As shown in eqn (8a), when the C_0 , C_{H^+} and the ratio k were all same, a linear relationship could be observed between $\ln t$ and $1/T$ and the line slope was $\Delta E/R$. The same k meant same GPV because of the same C_0 . The linear relationship between $\ln t$ and $1/T$ is shown in Fig. S1.† According to the average slope of four lines and the equation slope = $\Delta E/R$, the calculated apparent Arrhenius activation energy is 51.28 kJ mol⁻¹.

The time that the reaction generated a certain gas volume (500, 750, 1000 mL) was recorded under various C_{H^+} (0.00445, 0.0089, 0.0178, 0.0356 M) with the presence of 1.5 M of each reactant at 303 K. As shown in eqn (8b), when the C_0 , T and the ratio k were all same, a linear relationship could be observed between $\lg t$ and $(-\lg C_{H^+})$ and the line slope was the order of H^+ (m). The linear relationship between $\lg t$ and $(-\lg C_{H^+})$ is shown in Fig. S2† and the order of H^+ (m) is 1.3291.

The time that the reaction generated a certain gas volume ($k = 1.1386, 1.2546, 1.3219$) was recorded under various reactant concentrations (1.0, 1.25, 1.5, 2.0 M) in the presence of 0.0178 M H^+ at 303 K. As shown in eqn (8b), when the C_{H^+} , K and the ratio k were all same, a linear relationship could be observed between $\lg t$ and $\lg C_0$ and the line slope was the order of reactants (n). According to the average slope of the three lines in Fig. S3,† the order of reactants (n) is 2.0949.

On the basis of the intercepts of linear relationship in Fig. S1–S3,† the frequency factor A can be calculated as the average value of eqn (9a)–(9c) and the frequency factor A is 7.103×10^7 .

$$\text{Intercept} = \ln \left(\frac{k^{n-1} - 1}{n - 1} \right) - \ln A - m \ln C_{H^+} + (1 - n) \ln C_0 \quad (9a)$$

$$\text{Intercept} = \lg \left(\frac{k^{n-1} - 1}{n - 1} \right) - \lg A + (1 - n) \lg C_0 + \frac{\Delta E}{2.303RT} \quad (9b)$$

$$\text{Intercept} = \lg \left(\frac{k^{n-1} - 1}{n - 1} \right) - \lg A - m \lg C_{H^+} + \frac{\Delta E}{2.303RT} \quad (9c)$$

The resulting values of the undetermined kinetic parameters are listed in Table S1† and the kinetic equation of the reaction between $NaNO_2$ and NH_4Cl under experimental conditions is as follows:

$$\frac{dc}{dt} = -7.103 \times 10^7 c_{H^+}^{1.3291} c_0^{2.0949} e^{\left(\frac{-51.28 \text{ kJ mol}^{-1}}{RT} \right)} \quad (10)$$

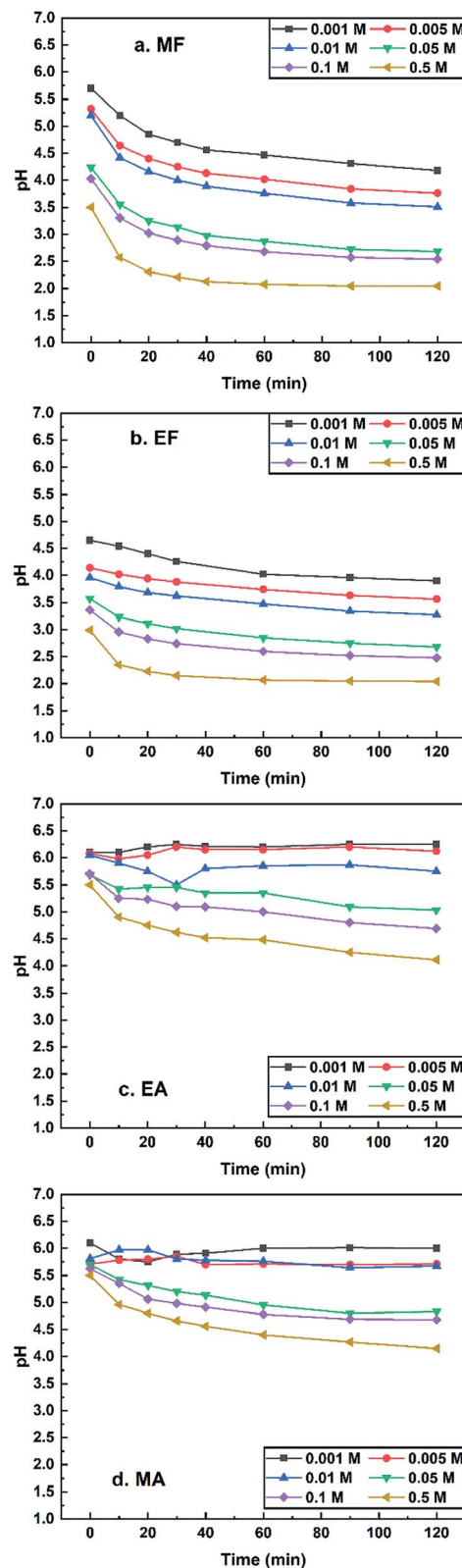


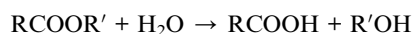
Fig. 7 pH values of methyl formate (a.MF), ethyl formate (b.EF), methyl acetate (d.MA) and ethyl acetate (c.EA) vs. time at 70 °C.



3.2. Acidic characteristic of authigenic acids

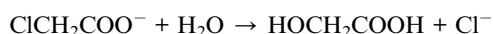
According to the results of Section 3.1, the reaction between NaNO_2 and NH_4Cl could benefit from the presence of H^+ . However, it is impractical to use HCl as a source of H^+ in field applications because the strong electrolyte could generate H^+ immediately and thus lead to a rapid generation of N_2 . The chemicals known as 'authigenic acids' would hardly generate H^+ at room temperature but release H^+ at high temperature, and thus authigenic acids can boost the system only at high temperatures. Meanwhile, the pH range for selecting the potential acids should be determined. Wu Wengang *et al.*²⁵ suggested that NO_x would not be formed by the $\text{NaNO}_2/\text{NH}_4\text{Cl}$ system when the pH is controlled at around 3 ($c(\text{H}^+) = 0.001 \text{ M}$). On the basis of previous experiments in Section 3.1, a much better gas production performance could be achieved without the generation of NO_x at a $c(\text{H}^+)$ range of 0.0045–0.0178 M. Therefore, the pH boundary for selecting authigenic acids is about 2–4 in this paper. The esters and salts are the two main kinds of authigenic acids used in the acidic characteristic experiments.

3.2.1. The esters. The four esters, methyl formate (MF), ethyl formate (EF), methyl acetate (MA), and ethyl acetate (EA) can hydrolyze spontaneously to generate acids and alcohols in aqueous solution.²⁸ These four esters are commonly used as authigenic acids in acidification. The hydrolysis temperature of the four esters rose with the increase in their molecular weight. Methyl formate was suitable for medium-low temperature formation. Ethyl formate, methyl acetate and ethyl acetate were suitable for high temperature formation. The hydrolysis equation of the ester is shown as follows.



A concentration range of 0.001–0.5 M was used to study the esters' potential acidic characteristic at 70 °C and the results of the experiment are displayed in Fig. 7. The pH values of the four ester solutions tended to decline with the increase of concentration and duration time. MA and EA whose pH values did not reach 4 under experimental conditions were not suitable for being authigenic acids in the N_2 generation system. For MF and EF, a concentration range of 0.01–0.1 M could cover the requirements mentioned above.

3.2.2. The salts. The three salts used in this study include sodium chloroacetate (SC), paraformaldehyde/ammonium chloride (PFA/ NH_4Cl) and ammonium persulphate (APS). SC is one of the chlorocarboxylates that features high-solubility and can hydrolyze moderately at a temperature range of 40–150 °C. The hydrolysis equation of chlorocarboxylates is as follows:²⁹

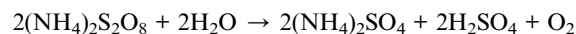


PFA is a linear polymer of formaldehyde that is soluble in hot water and gets decomposed by releasing formaldehyde (CH_2O). CH_2O and NH_4Cl would react to form HCl and

hexamethylenetetramine under certain conditions as the following reaction equation:^{30,31}



APS is an unstable substance whose aqueous solution can get decomposed to generate sulfuric acid when heated. The reaction equation is as follows:³²



The acidic characteristics of the three salts are shown in Fig. 8. The pH values of the three salt solutions (0.001–0.5 M) decreased with the rise of the concentration of salts and the pH value of SC remained essentially unchanged during the experiment. The PFA/ NH_4Cl and APS aqueous solutions had a downward trend with the increase of the measuring time. It is worth noting that the pH of the APS aqueous solution had a sharp drop at all tested concentrations but the pH of the PFA/ NH_4Cl mixture only declined considerably at higher concentrations, such as 0.05 M. It can be summarized that PFA/ NH_4Cl (0.1–0.5 M) and AP (0.001–0.01 M) can meet the requirements mentioned above.

Another pH test of APS at 30 °C was performed to estimate its stability at room temperature due to the sharp pH drop at 70 °C. If the pH value of APS solution at 30 °C declines as fast as that of 70 °C, this salt will not be the authigenic acid searched. As shown in Fig. 9, the pH value decreased slightly with the increasing measuring time, which suggested the good thermal stability of APS.

In conclusion, methyl formate (MF), ethyl formate (EF), paraformaldehyde (PFA) and ammonium persulphate (APS) can cover the requirements of authigenic acids. Therefore, MF, EF, PFA and APS were chosen as authigenic acids in the following gas-producing tests.

3.3. Gas-producing tests with authigenic acids

3.3.1. *In situ* N_2 generation system with MF. The gas-producing experimental results of the *in situ* N_2 generation system in the presence of MF (0.01–0.10 M) and 1.5 M of each reactant at various temperatures (30–70 °C) are illustrated in Fig. S4.† GPV and the reaction rate increased gradually as the concentration of MF grew at the same temperature. The temperature had a significant impact on gas production and reaction rate under the identical concentration of MF. It was remarkable that with the same concentration of MF, the *in situ* N_2 generation system could generate a large amount of gas at 70 °C, but only a small amount of gas was collected at 30 °C. The previous experimental results (Fig. 5) had shown that the addition of H^+ could raise the GPV and GGE immensely. However, commonly used acids, such as hydrochloric acid and acetic acid, can accelerate the process immediately, as soon as the reactants are in contact with acids. The *in situ* N_2 generation system assisted by authigenic acids is designed to solve this predicament. Authigenic acids can promote the reaction to



increase the GPV at high temperature and will not lead to an intensive reaction rate like HCl at room temperature. With the addition of MF (0.05, 0.10 M), the GPV of the system at 70 °C even surpassed that of the experimental group with 0.0178 M HCl at 70 °C (2900 mL, shown in Fig. 6). Meanwhile, the GPV of the system in the presence of MF (0.01–0.10 M) at 30 °C was merely 150 mL, which is much lower than the GPV (2080 mL) when 0.0178 M HCl was added at 30 °C.

3.3.2. *In situ* N₂ generation system with EF. The results of the gas-producing experiment with different concentrations of EF in the presence of 1.5 M of each reactant at various temperatures (30–70 °C) are shown in Fig. S5.† The trend of GPV of the *in situ* N₂ system assisted by EF was similar to the MF system, and the GPV had an upward trend with the rise in the concentration of EF and temperature. However, the total GPV of

the EF assisted system was always less than that of the MF assisted system when the temperature and addition were identical, which suggested that MF is easier to hydrolyze to generate H⁺ than EF. Under the same conditions, MF is more suitable for authigenic acid to boost the system. Also, the fact that EF is poorly soluble in water puts it in an inferior position again. When the total GPV of the EF assisted system is compared with the group with 0.0178 M HCl at 70 °C (2900 mL), only the 0.10 M EF assisted one had a slightly higher GPV than that of 0.0178 M HCl at 70 °C. The MF system could also reach the same GPV with a lower concentration (0.05 M). Therefore, MF would be a better choice as an authigenic acid than EF.

3.3.3. *In situ* N₂ generation system with APS. The gas-producing experimental results of the system with APS in the presence of 1.5 M of each reactant at various temperatures (30–70 °C) are displayed in Fig. S6.† With the increase in the concentration of APS and temperature, the GPV and reaction rate increased. One thing worth noticing is that the experimental group with 0.005 and 0.01 M APS at 30 °C had a larger GPV than the group with 0.001 M APS at 50 °C, which means that the concentration of APS had a greater impact on GPV than the temperature. This phenomenon was caused by the thermal instability of the APS aqueous solution, which led to the slow decomposition of APS, generating H⁺ at room temperature. The time that the experimental group began to generate gas quickly moved up as the concentration of APS was increased at 30 °C. The GPV of the group with 0.01 M APS at 30 °C reached 1880 mL, which was just 200 mL fewer than the experimental group with 0.0178 M HCl at 30 °C (2080 mL). It was obvious that 0.01 M was not the desired concentration of APS due to the large GPV at 30 °C. As for the other groups without the generation of NO₂ and NO, the group with 0.005 M APS could enhance the reaction better than the group with 0.001 M APS. Therefore, the optimal concentration of APS was 0.005 M.

3.3.4. *In situ* N₂ generation system with PFA. Actually, PFA could not originally serve as an authigenic acid at high temperatures. Due to the presence of NH₄Cl in the N₂ generation mixture, PFA could react with NH₄Cl to generate H⁺. The GPV of *in situ* N₂ system assisted by various concentrations of PFA in the presence of 1.5 M of each reactant at targeted temperatures (30, 50, 70 °C) are shown in Fig. S7.† Similar to the previous three

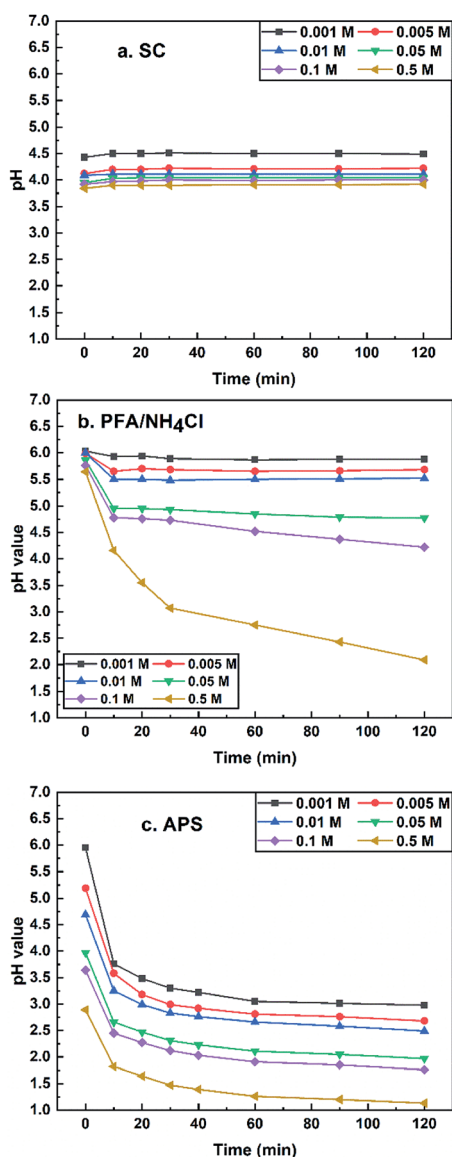


Fig. 8 pH values of sodium chloroacetate (a: SC), paraformaldehyde/ammonium chloride (b: PFA/NH₄Cl) and ammonium persulphate (c: APS) vs. time at 70 °C.

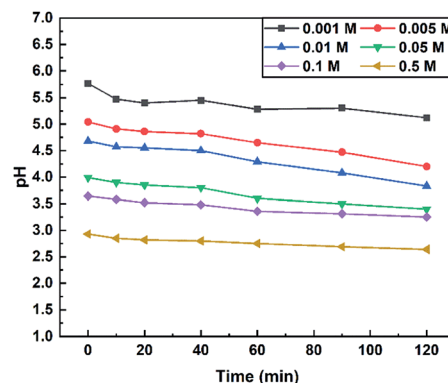


Fig. 9 pH values of ammonium persulphate (APS) vs. time at 30 °C.



authigenic acids, the GPV went up with the increase in the concentration of PFA and reaction temperature. The GPV of the system with the addition of PFA (0.05, 0.10 M) at 70 °C was much more than that of the group with 0.0178 M HCl at 70 °C (2900 mL). The experimental group with 0.10 M PFA had the maximum GPV in all the experiments, but a reddish brown gas (NO_x) could be seen during the experiment. Therefore, the *in situ* N_2 generation system assisted by 0.05 M PFA was selected to conduct the energy enhancement experiments. It is worth noting that PFA is a solid and is substantially insoluble in water at room temperature, which may hinder the field application of PFA.

3.4. Energy enhancement examination

According to the results of the gas-producing tests, an *in situ* N_2 generation system with a higher concentration of reactants, H^+ and a higher temperature could promote the GPV. However, the ability of this system to increase the pressure and temperature is still needed to be estimated by energy enhancement experiments in a constant volume reactor. Energy-enhancing means the increase of pressure and temperature in oilfield formation and the energy enhancement experiments were conducted to investigate the effects of reactant concentration and initial pressure on increasing the pressure and temperature. The performance of the N_2 generation system assisted by four different authigenic acids was compared through the energy enhancement experiments.

Before the experiments, several parameters had to be defined: P_c was the container pressure at 25 °C; P_0 was the initial container pressure at room temperature (20 °C); ΔP was the difference between P_c and P_0 ($\Delta P = |P_c - P_0|$); T_{\max} and P_{\max} were the maximum temperature and pressure during the experiments. The gas production volume V was calculated according to the ideal-gas equation ($PV = nRT$).

3.4.1. Effect of reactant concentration. The effect of $c(\text{reactant})$ (1.5–3.0 M for each reactant) on energy enhancement at 100 °C is illustrated in Table 1. The system with a higher concentration of reactants showed a better ability to increase the pressure and temperature. The T_{\max} and P_c increased from

125 to 218 °C and from 2.88 to 6.15 MPa, respectively, with the increase in the concentration of each reactant from 1.5 to 3.0 M, and the GGE was maintained around 82% except for 1.5 M (77.48%), which suggested that the increase in the reactant concentration had a little influence on GGE.

3.4.2. Effect of P_0 . The effect of the initial pressure (2–8 MPa) on energy enhancement with 1.5 M of each reactant at 100 °C is displayed in Table 2. With the rise of P_0 from 2 to 8 MPa, the P_{\max} and P_c increased gradually but the ΔP , GPV and GGE remained relatively stable. The T_{\max} of the four groups was around 133 °C, which suggested a similar thermal effect of the four experimental groups. Considering the influence of the pressure change on gas compression and dissolution, it could be recognized that the P_0 had a little effect on the *in situ* N_2 generation system.

3.4.3. Effect of four authigenic acids. Four kinds of temperature sensitive authigenic acids were used to assist the *in situ* N_2 generation system. The temperature of experiments was 100 °C and the concentration of each reactant was 1.5 M. The results of the system with four authigenic acids on energy enhancement are shown in Table 3. As expected, the T_{\max} increased from 125 °C to around 135 °C and the P_c rose from 2.88 MPa to around 3.30 MPa, respectively; the much higher T_{\max} and P_c showed that the four *in situ* N_2 generation systems assisted by authigenic acids are more effective in increasing the pressure and temperature than the blank group. The experimental groups with MF (0.05 M) and APS (0.005 M) could increase the pressure from 0 to 3.31 MPa, which was the best performance among all groups. Compared with the results in Tables 1 and 2, the addition of MF (0.05 M) and APS (0.005 M) could increase GGE from approximately 80.00% to 89.07%, which was higher than the other groups. Although MF and AP could both improve the energy-enhancing effect effectively, there are still some questions to be settled. Compared with APS, the minimum concentration of MF (0.05 M) is much more than that of APS (0.005 M) to achieve the same effect. Therefore, APS might be a better choice as the authigenic acid that can be used to assist the *in situ* N_2 generation system.

Table 1 The results of energy enhancement measurement at four reactant concentrations

Temperature/°C	P_0 /MPa	$c(\text{reactant})/\text{M}$	$T_{\max}/^\circ\text{C}$	P_{\max}/MPa	P_c/MPa	$\Delta P/\text{MPa}$	GPV/L	GGE/%
100	0	1.5	125	3.62	2.88	2.88	4.184	77.48
		2.0	163	5.26	4.15	4.15	6.030	83.75
		2.5	201	8.00	5.10	5.10	7.410	82.33
		3.0	218	10.90	6.15	6.15	8.936	82.74

Table 2 The results of energy enhancement measurements at four initial pressures

Temperature/°C	P_0 /MPa	$c(\text{reactant})/\text{M}$	$T_{\max}/^\circ\text{C}$	P_{\max}/MPa	P_c/MPa	$\Delta P/\text{MPa}$	GPV/L	GGE/%
100	2	1.5	135	6.20	5.10	3.10	4.455	82.50
	4		136	8.70	7.20	3.20	4.550	84.26
	6		130	10.75	8.90	2.90	4.065	75.28
	8		133	11.00	11.00	3.00	4.161	77.06



Table 3 The results of energy enhancement measurements of the system with authigenic acids

Authigenic acids	Temperature/°C	P_0 /MPa	T_{\max} /°C	P_{\max} /MPa	P_c /MPa	ΔP /MPa	GPV/L	GGE/%
Blank group	100	0	125	3.62	2.88	2.88	4.184	77.48
0.005 M APS			135	4.08	3.31	3.31	4.810	89.07
0.050 M MF			137	4.10	3.31	3.31	4.810	89.07
0.050 M MF			135	4.05	3.30	3.30	4.795	88.80
0.050 M PFA			128	3.85	3.19	3.19	4.621	85.57

4. Conclusion

The main purpose of this study was to find out the effect of several factors on the gas production volume (GPV) and develop the kinetics equation of the $\text{NaNO}_2/\text{NH}_4\text{Cl}$ system. In order to form an authigenic acid assisted *in situ* N_2 generation system, several potential chemicals were selected by monitoring the pH value. The gas generating and energy enhancing capacities of these systems were examined and the main conclusions are summarized below.

(1) With the increasing reactant concentration, $c(\text{H}^+)$ and the temperature, the volume of nitrogen production increased. The optimal molar ratio of $\text{NaNO}_2/\text{NH}_4\text{Cl}$ is 1 : 1. The experimental kinetic equation of the N_2 generating process by NaNO_2 and NH_4Cl is $\text{dc}/\text{dt} = -7.103 \times 10^7 c_{\text{H}^+}^{1.3291} c_0^{2.0949} e^{(-51.28 \text{ kJ mol}^{-1})/RT}$.

(2) The optimized concentrations of authigenic acids are 0.05 M for methyl formate (MF) and 0.005 M for ammonium persulphate (APS). Both these authigenic acids could clearly increase the gas production volume of the system.

(3) The energy-enhancing capacity increased with the increasing reactant concentration while the initial pressure had a little impact on energy enhancement. The presence of authigenic acids in the $\text{NaNO}_2/\text{NH}_4\text{Cl}$ mixture would provide a higher pressure and temperature. The gas-generating efficiency of the system with 0.05 M methyl formate and 0.005 M ammonium persulphate could reach 89.07%. The N_2 generation system assisted by ammonium persulphate would be a better choice for oil field application than methyl formate.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 S. S. Wang, C. L. Chen, B. Shiao and J. H. Harwell, *Fuel*, 2018, **217**, 499–507.
- 2 D. Y. Zhu, J. R. Hou, J. F. Wang, X. Wu, P. Wang and B. J. Bai, *Fuel*, 2018, **215**, 619–630.
- 3 F. Wang, H. L. Chen, S. Alzobaidi and Z. M. Li, *Energy Fuels*, 2018, **32**, 9093–9105.
- 4 K. K. Gumersky, I. S. Dzhabarov, A. K. Shakhverdiev and Y. G. Mamedov, *presented at SPE European Petroleum Conference*, Paris, France, October, 2000, SPE 65170.
- 5 D. A. Nguyen, H. S. Fogler and S. Chavadej, *Ind. Eng. Chem. Res.*, 2001, **40**, 5058–5065.
- 6 D. R. Davies, E. A. Richardson and D. Anthunis, *presented at SPE Middle East Oil Technical Conference and Exhibition*, Manama, Bahrain, March, 1981, SPE 9653.
- 7 H. W. Mcspadden, M. L. Tyler and T. T. Velasco, *presented at SPE California Regional Meeting*, Oakland, California, April, 1986, SPE 15098.
- 8 C. N. Khalil and Z. A. Franco, *presented at SPE Latin America Petroleum Engineering Conference*, Rio de Janeiro, Brazil, October, 1990, SPE 21113.
- 9 E. A. Evangelista, C. M. Chagas, J. A. F. Melo, J. D. H. Rocha, N. B. Filho and L. C. C. Marques, *presented at SPE Offshore Technology Conference*, Houston, Texas, USA, May, 2009, SPE 19730.
- 10 T. Moussa, S. Patil and M. Mahmoud, *presented at SPE Kingdom of Saudi Arabia Annual Technical Symposium and Exhibition*, Dammam, Saudi Arabia, April, 2018, SPE 192432-MS.
- 11 E. A. Richardson, R. F. Scheuerman and D. C. Berkshire, *US Pat.*, 4219083, Aug. 26, 1980.
- 12 E. A. Richardson and R. F. Scheuerman, *US Pat.*, 4178993, Dec. 18, 1979.
- 13 T. I. Mitchell, S. C. Donovan, J. B. Collesi and H. W. Mcspadden, *presented at SPE California Regional Meeting*, Long Beach, California, USA, April, 1984, SPE 12776.
- 14 R. K. Romeu, C. N. Khalil and A. Rabinovitz, *presented at SPE Latin America Petroleum Engineering Conference*, Rio de Janeiro, Brazil, October, 1990, SPE 21108.
- 15 K. Singh, A. S. Mohamed, S. S. Alian, M. Ismail, M. Anwar, W. W. Mohamed and S. A. Ghani, *presented at SPE Offshore Technology Conference*, Houston, Texas, USA, May, 2013, SPE 23933.
- 16 A. R. Alnakhli, L. A. Sukkar, J. Arukhe, A. Mulhem, A. Mohannad, M. Ayub and M. Arifin, *presented at SPE Heavy Oil Conference & Exhibition*, Kuwait City, Kuwait, December, 2016, SPE 184118MS.
- 17 Y. X. Liu, Q. Xie and X. Y. Chen, *Pet. Geol. Recovery Effic.*, 2002, **9**, 85–88.
- 18 J. P. Ashton, L. J. Kirspel, H. T. Nguyen and D. J. Credeur, *SPE Prod. Eng.*, 1989, **4**, 157–160.



- 19 P. Singh and H. S. Fogler, *Ind. Eng. Chem. Res.*, 1998, **37**, 2203–2207.
- 20 D. A. Nguyen, F. F. Moraes and H. S. Fogler, *Ind. Eng. Chem. Res.*, 2004, **43**, 5862–5873.
- 21 N. C. Lu, K. C. Zhang, P. H. Tan, T. Z. Yang and L. J. Wu, *J. Chem. Ind. Eng.*, 1998, **49**, 261–267.
- 22 F. L. Zhao, *Oilfield Chemistry*, China University of Petroleum Press, Dongying, China, 2nd edn, 2012, pp. 175–182.
- 23 E. D. Hughes, C. K. Ingold and J. H. Ridd, *J. Chem. Soc.*, 1958, 12–21.
- 24 J. H. Dusenbury and R. E. Powell, *J. Am. Chem. Soc.*, 1951, **73**, 3266–3268.
- 25 W. G. Wu, Y. Xiong, J. Wang and D. J. Chen, *Oilfield Chem.*, 2007, **24**, 106–108.
- 26 D. A. Nguyen, M. A. Iwaniw and H. S. Fogler, *Chem. Eng. Sci.*, 2003, **58**, 4351–4362.
- 27 A. M. Wu, M. T. Chen, S. R. Gu and W. J. Wang, *Oil Drill. Prod. Technol.*, 1995, **17**, 60–64.
- 28 Z. S. Gao, *Research of in situ generation gas technology cleanup and viscosity reduction adaptability in low permeability reservoirs*, MS thesis, Xi'an Shiyou University, China, May 2011.
- 29 Q. Y. Xing, W. W. Pei, R. Q. Xu and J. Pei, *Fundamentals of Organic Chemistry*, Higher Education Press, Beijing, China, 3rd edn, 2005, pp. 602–605.
- 30 W. F. Gao, *Study and preparation on new sandstone type latent acid*, MS thesis, China University of Petroleum (East China), China, May 2011.
- 31 W. M. Salathiel and C. M. Shaughnessy, *US Pat.*, 4136739, Jan. 30, 1979.
- 32 Y. N. Liu, J. H. Lei and L. S. Wang, *Inorganic Chemistry*, Science Press, Beijing, China, 2nd edn, 2013, pp. 328–330.

