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Isomerization enabling near-infrared electron acceptors†

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An isomerization method was utilized to yield a novel near-infrared nonfullerene acceptor DTA-IC-M. By simply changing the linking fashion between the anthracene and neighboring thiophenes, a remarkable redshift (~ 170 nm) of absorption was observed from DTA-IC-S to its isomer DTA-IC-M which shows a maximum absorption peak over 800 nm with a narrow bandgap of 1.35 eV. Due to the enhanced photo-to-current response in the near-infrared region, an improved short-circuit current of 12.96 mA cm^{-2} was achieved for the DTA-IC-M based OSCs.

Nonfullerene n-type organic semiconductors have attracted increasing attention¹ as electron acceptors in organic solar cells (OSCs) owing to their favorable features, such as widely tunable absorption windows and controllable energy levels.² It is well known that the majority of solar irradiation is located at near-infrared (NIR) region (>750 nm), hence exploring NIR nonfullerene acceptors (NNAs) is highly desirable in order to achieve efficient OSCs with high short-circuit current (J_{sc}). In addition, NNAs will be beneficial for realizing high-performance semitransparent and tandem OSCs. For example, a carbon–oxygen-bridged NNA reported by Ding *et al.* shows an ultralow bandgap (E_{g}) of 1.26 eV with the maximum absorption peak over 800 nm, which enables a remarkable J_{sc} of 27.39 mA cm^{-2} with a power conversion efficiency (PCE) of 14.62%³ in single-junction OSCs and a much higher PCE of 17.36%⁴ in tandem OSCs. Recently, Zou *et al.* invented a novel A-D-A-D-A type NNA with a low E_{g} of 1.33 eV as well as good electron-transporting ability which boosted the PCE to 15.7% and 16.5% in binary^{5a} and ternary⁵ single-junction OSCs, respectively. Hence, huge potential still exists for the development of NNAs.

Generally, there are three main strategies for designing NNAs as shown in Fig. 1. Taking the simplest fused-ring electron acceptor IDT-IC as the starting point, elongating the π -conjugation of molecular skeleton is the most widely used

method to redshift absorption (method A).⁶ As a result, a low E_{g} of 1.35 eV is obtained for NNA1 having a 8-ring fused backbone.⁷ Alternatively, introducing strong electron-donating functionalities such as alkoxy or alkyl amino groups can also reduce E_{g} (method B),^{4,8} e.g. a similar E_{g} of 1.34 eV is achieved for NNA2.^{8d} In addition, narrow E_{g} could be also realized by increasing quinoidal character⁹ as demonstrated in NNA3 (method C), which shows an E_{g} of 1.32 eV.^{9b,10} Despite of these achievements in red-shifting absorption and narrowing bandgap for electron acceptors, tedious synthesis and high-cost production are usually required due to the introduction of additional specific functional groups onto the initial molecular backbones. Therefore, the exploration of NNAs with simple structural modification and facile synthesis is preferred for future commercialization of OSCs.

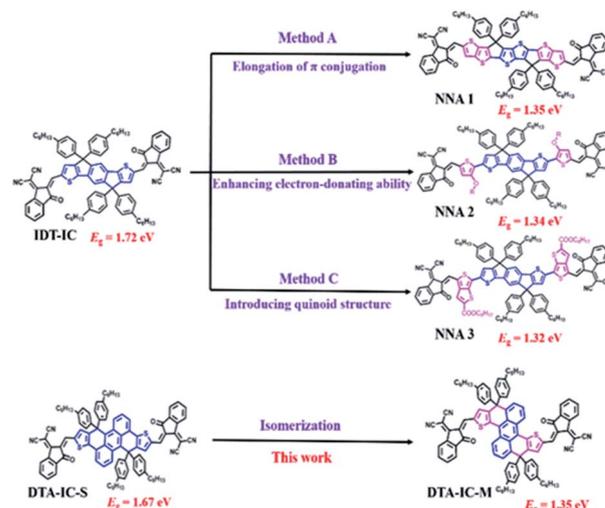


Fig. 1 Typical methods employed to reduce the bandgap (E_{g}) of nonfullerene acceptors.

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As reported in our previous work,¹¹ polycyclic aromatic hydrocarbon (PAH) based acceptors feature readily tailorable chemical structures and facile synthesis as well as fine-tuned optoelectronic properties. Based on such simple-to-use chemical backbones, here, an isomerization method was employed to produce a novel NNA without increasing the workload of synthesis. As depicted in Fig. 1, we have developed an anthracene based acceptor DTA-IC-S in our previous work,¹¹ which possesses a medium E_g of 1.67 eV similar to that of IDT-IC. Interestingly, after simply changing the linking fashion between the anthracene core and fused thiophene units, a significant redshift of nearly 170 nm could be realized for the newly-developed acceptor DTA-IC-M, rendering an absorption maximum (λ_{max}) at 822 nm as well as a low E_g of 1.35 eV. Owing to its wide photo-response in NIR region (750–950 nm), an enhanced J_{sc} of 12.96 mA cm⁻² could be achieved in the DTA-IC-M based solar cells compared to DTA-IC-S, enabling a PCE of 4.2%.

As shown in Scheme S1,[†] the synthetic route to DTA-IC-M is exactly the same with DTA-IC-S¹¹ except for different commercially available starting materials (*i.e.* 1,5-dibromoanthracene for DTA-IC-S and 9,10-dibromoanthracene for DTA-IC-M), ensuring low-cost and facile synthesis. The structure and purity of new acceptor DTA-IC-M was proved by ¹H NMR, ¹³C NMR, high resolution mass spectrometry, and elemental analysis as included in the ESI.[†]

The UV-vis absorption spectra of DTA-IC-S and DTA-IC-M are shown in Fig. 2a with the corresponding data summarized in Table 1. Surprisingly, a significant redshift of λ_{max} as large as 145 nm is observed for DTA-IC-M in solution *versus* its isomer DTA-IC-S, accompanied by an intensified molar extinction coefficient (ϵ) from $1.63 \times 10^5 \text{ M}^{-1} \text{ cm}^{-1}$ of DTA-IC-S to $2.23 \times$

$10^5 \text{ M}^{-1} \text{ cm}^{-1}$ of DTA-IC-M (Fig. S1[†] and Table 1). Ongoing from solution to thin film, the redshift of λ_{max} increases to 168 nm due to enhanced intermolecular interactions in solid state with the λ_{max} of DTA-IC-M reaching 822 nm in NIR region. As a result, a medium bandgap electron acceptor DTA-IC-S ($E_g = 1.67 \text{ eV}$) could be transformed to a NNA DTA-IC-M ($E_g = 1.35 \text{ eV}$) just by simple structural isomerization.

To unravel such an interesting phenomenon of absorption properties, density functional theory (DFT) and time-dependent DFT (TD-DFT) calculations were carried out at the B3LYP/6-31(d,p) level. As shown in Fig. S2a,[†] DTA-IC-S features a highly planar backbone with four 4-methylphenyl groups out of the plane, while DTA-IC-M exhibits a slightly twisted structure which might be due to the steric hindrance between S atoms in thiophene rings and hydrogens at 1 and 5 positions of anthracene. Fig. 2c and d reveal that the highest occupied molecular orbital (HOMO) is mainly localized on anthracene and fused thiophene rings for both DTA-IC-S and DTA-IC-M. However, a pronounced difference is observed for the lowest unoccupied molecular orbitals (LUMOs). The LUMO of DTA-IC-S is mainly concentrated on electron-withdrawing ending groups with marginal delocalization on anthracene, while the LUMO of DTA-IC-M is fully delocalized along the conjugated backbone covering the anthracene and fused thiophene rings as well as ending groups, indicative of a more conjugated backbone for DTA-IC-M than that for DTA-IC-S. This is reasonable by considering that unavoidable cross-conjugation mode exists in DTA-IC-S as displayed in Fig. S2c[†] whereas no conjugation node is found for DTA-IC-M despite of its slightly twisted structure.¹² Moreover, the upshifted HOMO level (-5.32 to -5.25 eV) and downshifted LUMO level (-3.21 to -3.50 eV) for DTA-IC-M suggest its narrower bandgap compared to DTA-IC-S. To further confirm the remarkable variation in absorption properties for these two isomers, the simulated absorption spectra were conducted by the TD-DFT calculation as shown in Fig. S2b.[†] Apparently, a large redshift of $\sim 140 \text{ nm}$ is also observed from DTA-IC-S to DTA-IC-M which is in good agreement with the experimental results. Our findings suggest that structural isomerization employed in this work could furnish a simple and effective means to construct NNA with facile synthesis.

The electrochemical properties of DTA-IC-S and DTA-IC-M were investigated by cyclic voltammetry¹³ (CV) as shown in Fig. 3a. Distinctive oxidative and reductive behaviors could be detected for both compounds, and the HOMO/LUMO levels of $-5.73/-3.83$ and $-5.53/-4.05 \text{ eV}$ were derived for DTA-IC-S and DTA-IC-M, respectively, based on the onsets of the redox peaks. The trend of energy level variation is in accordance with that from the DFT calculations. In addition to the frontier molecular orbital (FMO) energy levels, the electron mobility (μ_e) is also probed by the space-charge-limited-current (SCLC) method¹⁴ (Fig. S3[†] and Table 1). A relatively lower μ_e of $1.06 \times 10^{-6} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ is obtained for DTA-IC-M than that of DTA-IC-S ($3.90 \times 10^{-6} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$) which may be attributed to the twisted backbone of DTA-IC-M and the resultant weaker molecular packing in neat film.

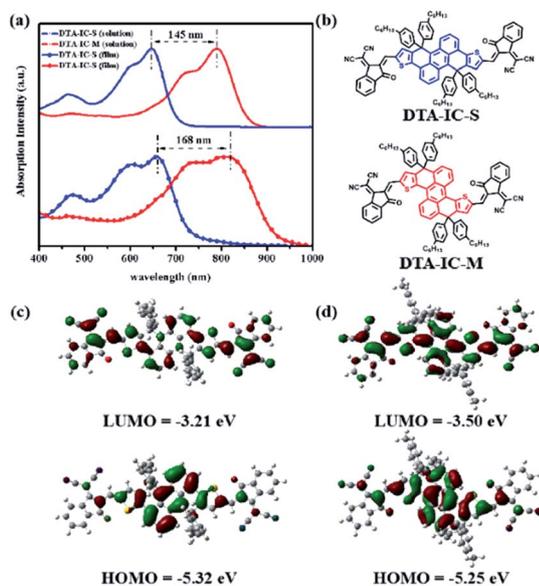


Fig. 2 (a) Normalized absorption spectra of DTA-IC-S and DTA-IC-M in CHCl₃ solutions and as thin films; (b) chemical structure of DTA-IC-S and DTA-IC-M; (c and d) frontier molecular orbitals of (c) DTA-IC-S and (d) DTA-IC-M.



Table 1 Optical properties, electronic energy levels and electron mobilities of DTA-IC-S and DTA-IC-M

Compound	$\lambda_{\max, \text{soln}}^a$ (nm)	ϵ ($\text{M}^{-1} \text{cm}^{-1}$)	$\lambda_{\max, \text{film}}^b$ (nm)	$\lambda_{\text{onset, film}}^b$ (nm)	E_g^c (eV)	HOMO ^d (eV)	LUMO ^d (eV)	E_g^{Cve} (eV)	μ_e ($\text{cm}^2 \text{V}^{-1} \text{s}^{-1}$)
DTA-IC-S	647	1.63×10^5	654	744	1.67	-5.73	-3.83	1.90	3.90×10^{-6}
DTA-IC-M	792	2.23×10^5	822	920	1.35	-5.53	-4.05	1.48	1.06×10^{-6}

^a Measured in chloroform solution with a concentration of $10^{-5} \text{ mol L}^{-1}$. ^b Measured in neat film. ^c Estimated from the absorption edge in neat film using the equation: $E_g = 1243/\lambda_{\text{onset}}$ (eV). ^d Measured in neat film using cyclic voltammetry (CV). ^e Electrochemical bandgap from CV measurement.

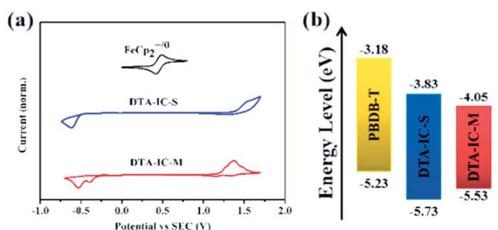


Fig. 3 (a) Cyclic voltammograms of DTA-IC-S and DTA-IC-M in acetonitrile with 0.1 M Bu_4NPF_6 as the supporting electrolyte; (b) energy level diagram of donor polymer PBDB-T and electron acceptors DTA-IC-S and DTA-IC-M.

To evaluate the potential of DTA-IC-M as a NNA, OSC devices are fabricated with the configuration of indium tin oxide (ITO)/poly(3,4-ethylenedioxythiophene): poly(styrene sulfonate) (PEDOT: PSS)/active layer/amino *N*-oxide perylene diimide (PDINO)¹⁵/Al, where PBDB-T¹⁶ is selected as the donor polymer. In addition, the acceptor DTA-IC-S is also used as a control. The optimized D: a weight ratio of PBDB-T: DTA-IC-S and PBDB-T: DTA-IC-M are 1 : 1.5 and 1 : 1, respectively, and no additives are needed during the device fabrication. The corresponding current density–voltage (J - V) characteristics of the devices measured under 100 mW cm^{-2} simulated sunlight illumination are displayed in Fig. 4a. The DTA-IC-S based device shows a PCE of 6.09%, with a V_{oc} of 0.94 V, a J_{sc} of 12.27 mA cm^{-2} and a FF of 0.529, which is in consistent with the

result reported in our previous work.¹¹ In contrast, the device based on DTA-IC-M shows a lower PCE of 4.20% which mainly resulted from the lower V_{oc} (0.69 V) and FF (0.468). Despite of this, the slightly enhanced J_{sc} (12.96 mA cm^{-2}) could be obtained from the DTA-IC-M based device which can be attributed to the additional NIR absorption of DTA-IC-M. As shown in Fig. 4b, the external quantum efficiency (EQE) spectra reveal that a much broader photo-to-current response from visible to NIR region (300 to 950 nm) is achieved for the DTA-IC-M based device which is in good agreement with its red-shifted absorption in film state. Although the EQE value is generally lower for the DTA-IC-M based device than that of the DTA-IC-S based on in the visible region, the absorption contribution in the NIR region enables a higher integrated J_{sc} of 12.56 mA cm^{-2} for the DTA-IC-M based cell compared to 11.93 mA cm^{-2} for the DTA-IC-S base one, suggesting the great significance of NIR harvesting for such low bandgap electron acceptor.

To understand the lower PCE of the NNA DTA-IC-M compared to the medium bandgap acceptor DTA-IC-S, energy levels are taken into consideration first. As shown in Fig. 3b, the lower-lying LUMO of DTA-IC-M is responsible for the decreased V_{oc} in the device compared to DTA-IC-S (0.69 V vs. 0.94 V).¹⁷ However, the energy loss (E_{loss}) is reduced to 0.66 eV for the DTA-IC-M based device as compared to that of the DTA-IC-S based cell (0.73 eV) according to the equation: $E_{\text{loss}} = E_g^{\text{opt}} - eV_{\text{oc}}$.¹⁸ This may be ascribed to suppressed nonradiative recombination losses for electron acceptors with a low bandgap.¹⁹ Subsequently, the charge transport properties in the blend films were investigated using the SCLC method as illustrated in Fig. 4c and d with the calculated data listed in Table S2.† The μ_e is comparable for both DTA-IC-S and DTA-IC-M based films ($2.96 \times 10^{-5} \text{ cm}^2 \text{V}^{-1} \text{s}^{-1}$ vs. $9.37 \times 10^{-5} \text{ cm}^2 \text{V}^{-1} \text{s}^{-1}$), whereas the hole mobility (μ_h) of DTA-IC-M blend film ($1.16 \times 10^{-5} \text{ cm}^2 \text{V}^{-1} \text{s}^{-1}$) is an order of magnitude lower than that of DTA-IC-S blend film ($1.96 \times 10^{-4} \text{ cm}^2 \text{V}^{-1} \text{s}^{-1}$) which leads to a much more unbalanced μ_e/μ_h (8.07) for DTA-IC-M based device and the resultant lower FF. This could be further confirmed by morphological analysis. Based on the atomic force microscopy (AFM) (Fig. S4†) and transmission electron microscopy (TEM) images (Fig. S5†), it can be seen that more obvious fiber-like aggregates exist in the PBDB-T:DTA-IC-S blend film with a larger root-mean-square (RMS) roughness of 3.12 nm, while the PBDB-T:DTA-IC-M blend shows a more homogenous film with a RMS roughness of 1.39 nm. This likely indicates that DTA-IC-M blend film suffers from an unsatisfactory phase separation which results in lower charge transporting abilities, increased charge recombination losses and

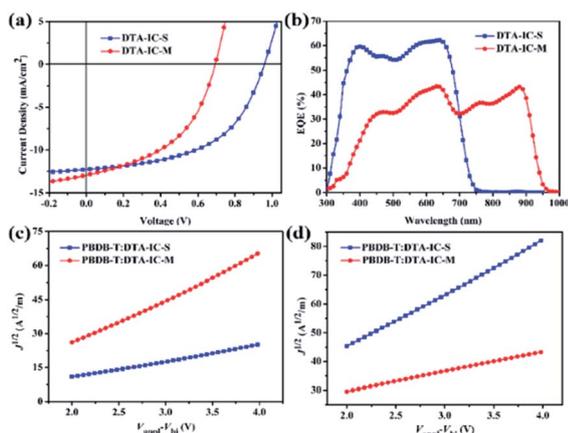


Fig. 4 (a) J - V characteristics of the optimized OSCs under the illumination of AM 1.5G, 100 mW cm^{-2} ; (b) EQE curves of the corresponding OSCs; (c) $J^{1/2}$ - V curves of the electron-only devices for the optimal active layer; (d) $J^{1/2}$ - V curves of the hole-only devices for the optimal active layer.



the lower FF, ultimately.^{8d} Similar results could also be obtained from two dimensional (2D) grazing-incidence wide-angle X-ray scattering (GIWAXS) (Fig. S6†). The weaker scattering along both in-plane and out-of-plane directions are observed for DTA-IC-M compared to DTA-IC-S in their pristine films due to the more twisted molecular backbone of DTA-IC-M and undermining intermolecular stacking. Similarly, the PBDB-T:DTA-IC-M blend film shows a slightly weaker π - π scattering along out-of-plane direction compared to the PBDB-T:DTA-IC-S blend film which may be owing to the more disruptive PBDB-T crystallization when blended with the DTA-IC-M with a lower ordering. Therefore, the nonplanar structure of DTA-IC-M is regarded as the main origin for the weaker packing, lower charge mobility, poorer phase separation and ultimately lower device performance in OSC devices. However, this issue can be tackled by further structural optimization, *e.g.* replacing fused thiophene units with furan units to reduce steric hindrance. Please note that the most significant finding in this work is the realization of a novel NIR electron acceptor featuring a simple structure and facile synthesis by a newly developed structural isomerization method.

Compared with several benchmark nonfullerene acceptors (Table S3†), the acceptor DTA-IC-M synthesized here displays a fairly narrow optical bandgap of 1.35 eV. It should be noted that the bandgap of the benchmark non-fullerene acceptors becomes narrower in principle as the field progresses (Table S3†). To date, the best-performance non-fullerene acceptor Y6 has a bandgap of 1.33 eV,¹⁴ which is highly comparable to that (1.35 eV) of our acceptor. The results demonstrate the greatly advantage of DTA-IC-M for non-fullerene solar cells in terms of optoelectronic properties. We must admit that the performance of our acceptor is greatly lower than that of the state-of-the-art acceptors, which reflects the fact that there is a great room for further performance improvement through donor polymer pairing and film morphology optimization. Hence the isomerization method presented in this work provides an effective way (even if not universal) to develop novel NIR acceptors, which are highly pursued for maximizing the device performance of non-fullerene organic solar cells. In addition, developing novel NIR nonfullerene acceptors is essential for efficient ternary solar cells or tandem solar cells. More efforts on structural optimization using this isomerization method are currently undergoing and further improvement of PCE could be anticipated in the future.

In summary, a new NIR electron acceptor DTA-IC-M is developed *via* a structural isomerization method with simple chemical structure and facile synthesis. Compared to its medium bandgap isomeric acceptor DTA-IC-S, a remarkable bathochromic shift (~ 170 nm) is attained for the new DTA-IC-M, which shows a maximum absorption peak over 800 nm and a narrowed bandgap of 1.35 eV. Due to the enhanced photo-current response in the NIR region, a higher J_{sc} of 12.96 mA cm^{-2} is achieved for the DTA-IC-M based solar cell with a PCE of 4.21%. Moreover, the universality of this isomerization approach and the isomerization effect on the stability are currently undergoing in our lab and further PCE enhancement could be anticipated in the near future.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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Notes and references

- (a) Y. Lin and X. Zhan, *Mater. Horiz.*, 2014, **1**, 470–488; (b) S. Li, L. Ye, W. Zhao, S. Zhang, S. Mukherjee, H. Ade and J. Hou, *Adv. Mater.*, 2016, **28**, 9423–9429; (c) W. Zhao, D. Qian, S. Zhang, S. Li, O. Inganäs, F. Gao and J. Hou, *Adv. Mater.*, 2016, **28**, 4734–4739; (d) J. Zhao, Y. Li, G. Yang, K. Jiang, H. Lin, H. Ade, W. Ma and H. Yan, *Nat. Energy*, 2016, **1**, 15027; (e) G. C. Bazan, *Sci. China: Chem.*, 2017, **60**, 1109–1110; (f) J. Hou, O. Inganäs, R. H. Friend and F. Gao, *Nat. Mater.*, 2018, **17**, 119; (g) B. Fan, D. Zhang, M. Li, W. Zhong, Z. Zeng, L. Ying, F. Huang and Y. Cao, *Sci. China: Chem.*, 2019, **62**, 746–752; (h) J. Yuan, Y. Zhang, L. Zhou, G. Zhang, H. L. Yip, T. K. Lau, X. Lu, C. Zhu, H. Peng, P. A. Johnson, M. Leclerc, Y. Cao, J. Ulanski, Y. Li and Y. Zou, *Joule*, 2019, **3**, 1140–1151.
- (a) H. Sun, B. Liu, C. W. Koh, Y. Zhang, J. Chen, Y. Wang, P. Chen, B. Tu, M. Su, H. Wang, Y. Tang, Y. Shi, H. Y. Woo and X. Guo, *Adv. Funct. Mater.*, 2019, 1903970; (b) Y. Wang, Z. Yan, M. A. Uddin, X. Zhou, K. Yang, Y. Tang, B. Liu, Y. Shi, H. Sun, A. Deng, J. Dai, H. Y. Woo and X. Guo, *Sol. RRL*, 2019, **3**, 1900107; (c) K. Yang, Q. Liao, C. W. Koh, J. Chen, M. Su, X. Zhou, Y. Tang, Y. Wang, Y. Zhang, H. Y. Woo and X. Guo, *J. Mater. Chem. A*, 2019, **7**, 9822–9830.
- H. Li, Z. Xiao, L. Ding and J. Wang, *Sci. Bull.*, 2018, **63**, 340–342.
- L. Meng, Y. Zhang, X. Wan, C. Li, X. Zhang, Y. Wang, X. Ke, Z. Xiao, L. Ding, R. Xia, H. L. Yip, Y. Cao and Y. Chen, *Science*, 2018, **361**, 1094.
- R. Yu, H. Yao, Y. Cui, L. Hong, C. He and J. Hou, *Adv. Mater.*, 2019, **31**, 1902302.
- (a) X. Li, T. Yan, H. Bin, G. Han, L. Xue, F. Liu, Y. Yi, Z. Zhang, T. P. Russelle and Y. Li, *J. Mater. Chem. A*, 2017, **5**, 22588–22597; (b) Y. Lin and X. Zhan, *Adv. Energy Mater.*, 2015, **5**, 1501063; (c) S. Dai, F. Zhao, Q. Zhang, T. K. Lau, T. Li, K. Liu, Q. Ling, C. Wang, X. Lu, W. You and X. Zhan, *J. Am. Chem. Soc.*, 2017, **139**, 1336–1343; (d) B. Jia, S. Dai, Z. Ke, C. Yan, W. Ma and X. Zhan, *Chem. Mater.*, 2018, **30**, 239–245.
- S. Dai, T. Li, W. Wang, Y. Xiao, T. K. Lau, Z. Li, K. Liu, X. Lu and X. Zhan, *Adv. Mater.*, 2018, **30**, 1706571.
- (a) X. Guo and M. D. Watson, *Org. Lett.*, 2008, **10**, 5333–5336; (b) K. H. Hendriks, W. Li, M. M. Wienk and R. A. Janssen, *J.*



- Am. Chem. Soc.*, 2014, **136**, 12130–12136; (c) X. Guo, Q. Liao, E. F. Manley, Z. Wu, Y. Wang, W. Wang, T. Yang, Y. E. Shin, X. Cheng, Y. Liang, L. X. Chen, K. J. Baeg, T. J. Marks and X. Guo, *Chem. Mater.*, 2016, **28**, 2449–2460; (d) H. Yao, Y. Chen, Y. Qin, R. Yu, Y. Cui, B. Yang, S. Li, K. Zhang and J. Hou, *Adv. Mater.*, 2016, **28**, 8283–8287; (e) J. Sun, X. Ma, Z. Zhang, J. Yu, J. Zhou, X. Yin, L. Yang, R. Geng, R. Zhu, F. Zhang and W. Tang, *Adv. Mater.*, 2018, **30**, 1707150; (f) Y. Cui, C. Yang, H. Yao, J. Zhu, Y. Wang, G. Jia, F. Gao and J. Hou, *Adv. Mater.*, 2017, **29**, 1703080; (g) J. Sun, X. Ma, Z. Zhang, J. Yu, J. Zhou, X. Yin, L. Yang, R. Geng, R. Zhu, F. Zhang and W. Tang, *Adv. Mater.*, 2018, **30**, 1707150.
- 9 (a) Z. X. Liu, T. K. Lau, G. Zhou, S. Li, J. Ren, S. K. Das, R. Xia, G. Wu, H. Zhu, X. Lu, H. L. Yip, H. Chen and C. Z. Li, *Nano Energy*, 2019, **63**, 103807; (b) F. X. Chen, J. Q. Xu, Z. X. Liu, M. Chen, R. Xia, Y. Yang, T. K. Lau, Y. Zhang, X. Lu, H. L. Yip, A. K. Y. Jen, H. Chen and C. Z. Li, *Adv. Mater.*, 2018, **30**, 1803769; (c) H. Xu, Y. Yang, C. Zhong, X. Zhan and X. Chen, *J. Mater. Chem. A*, 2018, **6**, 6393–6401.
- 10 F. Liu, Z. Zhou, C. Zhang, J. Zhang, Q. Hu, T. Vergote, F. Liu, T. P. Russell and X. Zhu, *Adv. Mater.*, 2017, **29**, 1606574.
- 11 Y. Wang, B. Liu, C. W. Koh, X. Zhou, H. Sun, J. Yu, K. Yang, H. Wang, Q. Liao, H. Y. Woo and X. Guo, *Adv. Energy Mater.*, 2019, **9**, 1803976.
- 12 (a) M. Gholami and R. R. Tykwinski, *Chem. Rev.*, 2006, **12**, 4997–5027; (b) G. W. P. van Pruissen, J. Brebels, K. H. Hendriks, M. M. Wienk and R. A. J. Janssen, *Macromolecules*, 2015, **8**, 2435–2443.
- 13 (a) J. Wang, J. Zhang, Y. Xiao, T. Xiao, R. Zhu, C. Yan, Y. Fu, G. Lu, X. Lu, S. R. Marder and X. Zhan, *J. Am. Chem. Soc.*, 2018, **140**, 9140–9147; (b) S. Dai, Y. Xiao, P. Xue, J. James Rech, K. Liu, Z. Li, X. Lu, W. You and X. Zhan, *Chem. Mater.*, 2018, **30**, 5390–5396.
- 14 P. N. Murgatroyd, *J. Phys. D: Appl. Phys.*, 1970, **3**, 151–156.
- 15 Z. G. Zhang, B. Qi, Z. Jin, D. Chi, Z. Qi, Y. Li and J. Wang, *Energy Environ. Sci.*, 2014, **7**, 1966–1973.
- 16 D. Qian, L. Ye, M. Zhang, Y. Liang, L. Li, Y. Huang, X. Guo, S. Zhang, Z. a. Tan and J. Hou, *Macromolecules*, 2012, **45**, 9611–9617.
- 17 Z. Ding, X. Long, C. Dou, J. Liu and L. Wang, *Chem. Sci.*, 2016, **7**, 6197–6202.
- 18 W. Li, K. H. Hendriks, A. Furlan, M. M. Wienk and R. A. Janssen, *J. Am. Chem. Soc.*, 2015, **137**, 2231–2234.
- 19 Y. Cui, H. Yao, J. Zhang, T. Zhang, Y. Wang, L. Hong, K. Xian, B. Xu, S. Zhang, J. Peng, Z. Wei, F. Gao and J. Hou, *Nat. Commun.*, 2019, **10**, 2515.

