



Cite this: *RSC Adv.*, 2019, 9, 36027

Nanoceria-modified platinum supported on hierarchical zeolites for selective alcohol oxidation†

Marisa Ketkaew,^a Duangkamon Suttipat,^a Pinit Kidkhunthod,^b Thongthai Witoon^c and Chularat Wattanakit^{d,*a}

The highly selective oxidation of alcohols to aldehydes has been achieved due to the synergic effect of Pt and CeO₂ supported on hierarchical zeolites. The combination of Pt and CeO₂ strongly enhances the catalytic performance of the oxidation of benzyl alcohol to benzaldehyde with respect to the isolated materials. In addition, the hierarchical zeolite not only increases the fraction of exposed active sites because of its high surface area that can prevent the aggregation of Pt and CeO₂ nanoparticles, but also affects the oxidation state of cerium. The presence of a high content of trivalent Ce species (Ce³⁺) on the hierarchical zeolite benefits the oxidation reaction, eventually leading to almost 100% yield of an aldehyde product. Moreover, the catalytic performance can be further improved by the easily tunable Si to Al ratio of zeolite catalysts.

Received 25th September 2019

Accepted 31st October 2019

DOI: 10.1039/c9ra07793f

rsc.li/rsc-advances

Introduction

The selective oxidation of alcohols to aldehydes is of crucial importance to produce various intermediate compounds in fine chemicals.¹ In particular, the oxidation of benzyl alcohol to benzaldehyde is one of the most fascinating processes because aromatic aldehydes can be used in various potential applications, for example as an industrial solvent, a commercial food flavouring, and an interesting key intermediate for perfumes and dyes.² Typically, the synthesis of benzaldehyde has been achieved by the hydrolysis of benzylidene chloride or the partial oxidation of toluene.³ However, it often suffers from a low yield of product and a wastewater disposal problem.⁴ An alternative way to efficiently and cleanly produce benzaldehyde is the direct selective oxidation of liquid benzyl alcohol, which has been utilized under mild reaction conditions at a moderate reaction temperature.⁵

Indeed, various catalysts have been applied for the selective oxidation, in particular, Cu, Pd, Au and Pt,^{6–8} for example. Although noble metals such as Pt, Pd and Au^{9–11} play a crucial role as beneficial metal catalysts, there are still problems

according to their cost effectiveness. Recently, several researchers have developed the catalysts to circumvent this limitation by decreasing the amount of noble metal *via* the combination with other materials. For instance, Evangelisti *et al.*¹² developed the hybrid bimetallic Au/CuO core-shell nanoparticles to enhance the catalytic performance of alcohol oxidation due to the synergistic effect of Au and CuO nanostructures. Chen *et al.*¹³ also studied the synergistic effect of the bimetallic catalyst, which is Au–Pd supported on ceria–zirconia. They revealed that the influence of the different Au–Pd interactions obtained by the different synthesis method directly related to the catalytic activity. In addition, the promising strategy to reduce the cost of catalysts and to sustainably improve the catalytic activity would be achieved by adding inexpensive metal oxide materials such as CeO₂, ZrO₂ and TiO₂.^{14,15} Among them, the CeO₂ modified noble metal is one of the most interesting catalysts for the alcohol oxidation because of their outstanding characteristics such as the redox properties of Ce³⁺ and Ce⁴⁺ species, high oxygen vacancy capacity and poisoning tolerance of hydrocarbon on metal surfaces.^{16,17} The combination of them can greatly enhance the catalytic performance compared to isolated metals. In addition, it has been described that doping of ceria with silica or alumina can also transform ceria crystallites, eventually resulting in changing its oxidation states.¹⁸

As above-mentioned development of effective catalysts, other solid materials such as mesoporous silicas,¹⁹ molecular sieves,²⁰ montmorillonite²¹ and medicinal stone²² have been used to enhance the catalytic performance in many crucial oxidation processes. To date, many reports of the developed catalysts for selective oxidation of benzyl alcohol have been launched. It has

^aDepartment of Chemical and Biomolecular Engineering, School of Energy Science and Engineering, Nanocatalysts and Nanomaterials for Sustainable Energy and Environment Research Network NANOTEC, Vidyasirimedhi Institute of Science and Technology, Rayong 21210, Thailand. E-mail: Chularat.w@vistec.ac.th

^bSynchrotron Light Research Institute (Public Organization), 111 University Avenue, Muang District, Nakhon Ratchasima, 30000 Thailand

^cFaculty of Engineering, Department of Chemical Engineering, Kasetsart University, Bangkok 10900, Thailand

† Electronic supplementary information (ESI) available: Details of experimental procedures and material characterization. See DOI: 10.1039/c9ra07793f



been emphasized to modify the surface properties of the support in order to not only enhance the active sites dispersion but also facilitate in activating alcohol molecules.^{23,24} However, the factor that has an influence on the catalytic performance for an oxidation is not only an easy accessibility of reactant molecules into the active sites but also the surface properties such as acid/base property of the support surface.^{25–27} Therefore, the tunability of solid support properties would make more benefit for the catalyst development.

It is well-known that a zeolite is also one type of aluminosilicate compounds and has been widely used as a solid support for metal oxides because of its unique properties, such as high-surface area, high metal ion exchange capacity, and high thermal/hydrothermal stability.²⁸ However, the presence of sole micropores imposes the intracrystalline diffusion limitation in the zeolite framework, eventually leading to low utilization of active sites. Recently, there have been several reports on the development of hierarchical zeolites possessing at least two levels of porous structures, such as a microporous feature together with meso- or/and microporous networks,^{29,30} which have been used to improve the accessibility and molecular transportation of guest molecules into the zeolite framework.³¹

Although the hierarchical zeolite might be one of the most promising candidates as a solid support for nanometal oxides due to an increase in surface area, resulting in improving the dispersion of metal, the development of combined catalysts between the nanoceria-modified platinum and the hierarchical zeolite as a solid support has not yet been demonstrated even though there would be many advantages, in particular the beneficial effect of hierarchical structures on the improvement of redox active species.

In this context, we present the development of nanoceria-modified platinum supported on hierarchical zeolites for the selective oxidation of alcohols to aldehydes. The effect of alumina content in the hierarchical zeolite framework on the catalytic performance was also systematically investigated. Furthermore, the structural analysis of developed catalysts was completed by means of various techniques such as X-ray diffraction (XRD), scanning electron microscopy (SEM), transmission electron microscopy (TEM), H₂ temperature programmed reduction (H₂-TPR) and X-ray absorption near edge structure (XANES) to gain insights into the mechanistic perspectives of redox species on different supports, which affect the catalytic activity of the selective oxidation of alcohols.

Experimental

Materials

All chemicals were of analytical grade and used as received without further purification. Aluminium isopropoxide (Sigma-Aldrich, ≥98.0%), tetraethyl orthosilicate (TEOS, Sigma-Aldrich, ≥99.0%), sodium hydroxide (NaOH, Carlo Erba, ≥98.0%) were used as aluminium source, silica source and mineralizing agent, respectively. Tetrapropylammonium hydroxide (TPAOH, 1.0 M in H₂O) and tetrabutylphosphonium hydroxide (TBPOH, Sigma-Aldrich, 40% in H₂O) were used as structure directing agents (SDAs) for conventional zeolites and

hierarchical zeolites, respectively. Cerium acetate (TCI, >98.0%) and tetraamineplatinum(II) nitrate (Sigma-Aldrich, ≥50.0% Pt basis) were used as precursor materials for cerium oxide and platinum active sites, respectively. To study catalytic activity, benzyl alcohol (Sigma-Aldrich, ≥99.0%), toluene (Merck, ≥99.9%) and decane (Sigma-Aldrich, ≥99.0%) were used as reactant, solvent and internal standard, respectively.

Catalyst preparation

Synthesis of conventional silicalite-1 zeolite (Silicalite1-CON).³² The conventional silicalite-1 zeolite was synthesized with the molar composition of 10SiO₂ : 1TPAOH : 1.03NaOH : 400H₂O. After mixing TEOS (7 g) and TPAOH (3.42 g) as a silica-SDA solution, the sodium hydroxide (0.14 g) mixed with DI-water was slowly added into the silica-SDA solution, and it was then stirred at room temperature for 2 h. After the aging process, the obtained gel was transferred to the Teflon-lined hydrothermal synthesis reactor and heated to 180 °C for 3 days. After that, the prepared sample was washed with DI-water, filtered and dried at 110 °C. Finally, the Silicalite1-CON zeolite was calcined at 650 °C for 8 h to remove SDA.

Synthesis of hierarchical silicalite-1 zeolite (Silicalite1-HIE).³² The hierarchical silicalite-1 zeolite was synthesized with the molar composition of 60SiO₂ : 18TBPOH : 0.75NaOH : 600H₂O. After mixing TEOS (8.67 g) and TBPOH (8.62 g) as a silica-SDA solution, the sodium hydroxide (0.02 g) mixed with DI-water was slowly added into silica-SDA solution, and it was then stirred at room temperature for 12 h. After the aging process, the synthesis gel was transferred to the Teflon-lined hydrothermal synthesis reactor and heated to 130 °C for 2 days. After that, the prepared sample was washed with DI-water, filtered and dried at 110 °C. Finally, the Silicalite1-HIE zeolite was calcined at 650 °C for 8 h to remove SDA.

Synthesis of hierarchical ZSM-5 zeolite (ZSM5-HIE).³² The hierarchical ZSM-5 zeolite with different Si/Al ratios (100 and 400) were synthesized with the molar composition of 60SiO₂ : xAl₂O₃ : 18TBPOH : 0.75NaOH : 600H₂O, where *x* was 0.075 and 0.3. Firstly, TEOS, aluminium isopropoxide were mixed as a silica-alumina solution. The second solution containing TBPOH (8.62 g), sodium hydroxide (0.02 g) and DI-water was slowly added into a silica-alumina solution, and then it was stirred at room temperature for 12 h. After the aging process, the obtained gel was transferred to the Teflon-lined hydrothermal synthesis reactor and heated to 130 °C for 2 days. After that, the prepared sample was washed with DI-water, filtered and dried at 110 °C. Finally, the ZSM5-HIE zeolite was calcined at 650 °C for 8 h to remove SDA.

Preparation of Pt–CeO₂ supported on zeolites. The as-synthesized zeolites were mixed with the desired amount of cerium acetate and DI-water at room temperature for 24 h, and then the solvent was evaporated by rotary evaporator. Subsequently, the obtained sample was dried at 100 °C overnight. The prepared catalysts were calcined at 550 °C for 6 h and denoted as *y*CeO₂-Silicalite1-HIE, *y*CeO₂-ZSM5-HIE, and *y*CeO₂-



Silicalite1-CON for the hierarchical silicalite-1, the hierarchical ZSM-5, and the conventional silicalite-1, respectively ($y = 0, 5, 10, 20$ and 30 wt% of CeO_2). After that, the platinum loading process was proceeded in a similar procedure. The $y\text{CeO}_2$ -Silicalite1-HIE, $y\text{CeO}_2$ -ZSM5-HIE, and $y\text{CeO}_2$ -Silicalite1-CON were stirred with aqueous platinum solution at room temperature for 24 h followed by the solvent removal and drying at 100°C overnight. Finally, the prepared catalysts were calcined at 550°C for 6 h and denoted as $z\text{Pt}/y\text{CeO}_2$ -Silicalite1-HIE, $z\text{Pt}/y\text{CeO}_2$ -ZSM5-HIE, $z\text{Pt}/y\text{CeO}_2$ -Silicalite1-CON, for the hierarchical silicalite-1, the hierarchical ZSM5, and the conventional silicalite-1, respectively ($z = 0, 0.5, 1$ and 2 wt% of Pt).

Characterization

The XRD patterns of the prepared catalysts were investigated using Bruker D8 ADVANCE instrument with CuK_α radiation (40 kV, 40 mA) in the 2θ range of 5° to 60° with the step size of 0.02° and the scan rate of 1°min^{-1} . To study the morphology of prepared catalysts, SEM and TEM images were obtained from JEOL JSM-7610F microscope and JEOL JEM-ARM200F microscope at 200 kV, respectively. The textural properties were determined by a N_2 adsorption-desorption technique at -196°C operated on a MicrotracBEL, BELSORP-max model and all the prepared samples were degassed at 350°C for 24 h before the measurement. The specific surface area (SBET), micropore surface area and pore volume were calculated by the Brunauer-Emmett-Teller (BET) method and t -plot method, respectively. The elemental analysis was studied by using wavelength-dispersive X-ray fluorescence spectrometer (WDXRF) performed on Bruker S8 TIGER ECO instrument. The H_2 -TPR measurements were used to investigate the reducibility properties of the obtained catalysts. The measurement was conducted using a BELCAT II instrument equipped with thermal conductivity detectors (TCD). Typically, the samples (0.05 g) were pre-treated at 300°C for 1 hour with the heating rate of 10°Cmin^{-1} under the flow of Ar and then cooled down to 50°C . After that, the temperature was increased from 50°C to 900°C with the heating rate of 5°Cmin^{-1} in a flow of 2 v/v% of H_2 in Ar (50 ml min^{-1}). To observe the metal dispersion, it was performed on a BELCAT II instrument equipped with thermal conductivity detector (TCD) using H_2 -pulse method. The catalysts were first treated in the air flow at 350°C for 30 min followed by purging with Ar for 15 min, and then pretreated in the flow of H_2 at 350°C for 20 min followed by purging with Ar for 20 min to remove weakly adsorbed H_2 . Finally, the samples were cooled down to room temperature in the flow of Ar, and a series of H_2 pulses was injected until the amount of H_2 pulses reached a steady state value. The amount of adsorbed H_2 was measured and the metal dispersion (MD) and the average diameter of particles (assuming with spherical metal particles) were calculated by using the following equations:

$$\text{MD}(\%) = \frac{V_m \times \text{MW} \times 10^4}{W\% \times \text{SF}}$$

$$D_a(\text{nm}) = \frac{6 \times 10^3}{\text{MS} \times \rho}$$

where V_m is monolayer adsorbed volume (mol g^{-1}), MW is metal atomic weight (g mol^{-1}), $W\%$ is metal percentage in sample, SF is stoichiometric factor (molecule of gas per metal atom), MS is metal surface area ($\text{m}^2 \text{g}^{-1}$) and ρ is metal density (g cm^{-3}), respectively.³³ Ce $L_{3\text{-edge}}$ XANES spectra were carefully characterized to study the oxidation state of Ce by using X-ray absorption near-edge structure in transmission mode at the BL5.2 station at Siam Photon Laboratory (Synchrotron Light Research Institute (Public Organization), SLRI), Thailand.

Catalytic activity test

The benzyl alcohol oxidation was carried out using a batch reactor. Firstly, the prepared catalysts were activated at 350°C under the flow of H_2 (5 ml min^{-1}) for 2 h. After that, a 0.08 g of catalyst was transferred to two-neck round-bottom flask connected with a condenser column as a batch reactor. Then, a 6 ml of toluene and decane were mixed together with the catalyst as a solvent and an internal standard, respectively. The reactor was sealed and heated to 80°C under vigorous stirring. The reactant, which is a 50 μl of benzyl alcohol was added into the reactor. The flow of air was introduced into the reactor under an atmospheric pressure using the balloon. The quantification of products was done by analysis using the internal calibration method. The reactant and products were sampled and analyzed by a gas chromatograph (GC, Agilent 7890B) equipped with a flame ionization detector (FID) and PoraBOND Q capillary column (25 m \times 0.32 mm). Subsequently, benzyl alcohol conversion ($X_{\text{benzyl alcohol}}$) and benzaldehyde selectivity ($S_{\text{benzaldehyde}}$) were calculated by the following equations:

$$X_{\text{benzyl alcohol}}(\%) = \frac{n_0 - n_i}{n_0} \times 100\%$$

$$S_{\text{benzaldehyde}}(\%) = \frac{n_{\text{benzaldehyde}}}{n_0 - n_i} \times 100\%$$

where n_0 and n_i are mole of benzyl alcohol at initial and subsequent time, respectively. The mass balance was calculated for all experiments in the range of $97.59 \pm 10.04\%$.

Results and discussion

To verify the effect of different contents of metals on the physicochemical properties of catalysts, the crystalline structures of nanoceria-modified platinum supported on hierarchical silicalite-1 (Silicalite1-HIE) with various metal loadings were examined by X-ray powder diffraction (XRD) as shown in Fig. S1 and S2.† As expected, when increasing the CeO_2 and Pt loading the characteristic peaks at 2θ of $28, 33, 48,$ and 56° and 40° corresponding to CeO_2 and Pt, respectively,³⁴ were visibly observed. In strong contrast to this, the crystallinity of the silicalite-1 framework corresponding to the MFI topology significantly decreases as a function of CeO_2 loading. Compared to other support materials, including the conventional silicalite-



1 (Silicalite1-CON), hierarchical ZSM-5 (ZSM5-HIE) and commercial CeO₂ nanoparticles, XRD patterns reveal that the nanoceria-modified platinum (1 wt% of Pt and 20 wt% of CeO₂) supported on the conventional silicalite-1 (1Pt/20CeO₂-Silicalite1-CON) consists of the characteristic peaks of CeO₂, which are similar to the pattern of 1Pt/CeO₂, whereas the corresponding peaks disappear in cases of both hierarchical supports (1Pt/20CeO₂-Silicalite1-HIE and 1Pt/20CeO₂-ZSM5-HIE) (Fig. 1A). It is therefore reasonable to assume that small CeO₂ nanoparticles can be somewhat formed on hierarchical zeolite surfaces. Therefore, it clearly demonstrates that the 1Pt/20CeO₂-Silicalite1-HIE and 1Pt/20CeO₂-ZSM5-HIE compose of well-dispersed CeO₂ and Pt nanoparticles, whereas the large aggregated particles are preferably deposited on the outermost surfaces of Silicalite1-CON.

The morphologies of all the prepared catalysts are observed by scanning electron microscope (SEM) images as shown in Fig. S3† confirming that the structures are not destroyed after the metal-loading process. The large particle size of Silicalite1-CON can be clearly observed, whereas hierarchical supports show a smaller particle formed from the self-assembled nanolayers with the size in the range of 150 to 250 nm.

To reveal the morphologies of metal nanoparticles on supports, transmission electron microscopy (TEM) images of 1 wt% of Pt and 20 wt% of CeO₂ loaded on different supports (conventional and hierarchical ones, which are 1Pt/20CeO₂-Silicalite1-CON and 1Pt/20CeO₂-Silicalite1-HIE, respectively) are shown in Fig. 1B and Fig. S4.† Obviously, the aggregation of

metals on the conventional samples is a main problem, whereas the metal distribution over Silicalite1-HIE is greatly improved. These behaviors are also confirmed by TEM-EDX (Fig. S5†) demonstrating a well-dispersed Pt and CeO₂ over the hierarchical sample.

Furthermore, N₂ physisorption isotherms are displayed in Fig. S6† and the textural properties are summarized in Table S1.† The conventional zeolite typically shows the type I isotherm, which can be observed for 1Pt/20CeO₂-Silicalite1-CON, implying that it composes of a pure microporous structure. In contrast, the hierarchical zeolites exhibit mixed isotherms between types I and IV due to the micropore filling at the low relative pressure with a high uptake and a hysteresis loop at the high relative pressure attributed to a capillary condensation within mesopores and macropores, respectively.³² However, the isotherm of 1Pt/CeO₂ is totally different demonstrating that there is almost no adsorption ability at low relative pressure with the small hysteresis loop at high relative pressure, indicating non-porous structures with the interparticle void.³⁵ These observations clearly show that hierarchical zeolites exhibit the higher specific surface area, which could provide the improved metal dispersion, eventually leading to enhance the catalytic performance. Although the 1Pt/CeO₂ composes of the smallest particle size obtained from SEM images (Fig. S3D†), it still suffers from the agglomeration of CeO₂ nanoparticles, resulting in a very low surface area of metal oxides.

To clarify the metal dispersion ability, the H₂ pulse chemisorption technique was also applied over nanoceria-modified platinum supported on different supports. The evaluation of the number of exposed active surfaces of Pt by chemisorption is based on the measurement of the amount of H₂ gas that is adsorbed during the analytical cycle. In order to inhibit the hydrogen spillover, it has been recommended to chemisorb hydrogen at lower temperature.³⁶ The adsorbed quantities are used to indicate the metal dispersion in accordance with the calculation.³³ Interestingly, hierarchical zeolite supports can enhance the metal dispersion compared with the conventional zeolite (Table S1†). These observations again confirm the improved dispersion of Pt and CeO₂ supported on hierarchical zeolites.

To investigate the redox properties of active sites, hydrogen temperature-programmed reduction (H₂-TPR) profiles reveal that both 1Pt/20CeO₂-Silicalite1-CON and 1Pt/20CeO₂-Silicalite1-HIE compose of three predominant peaks at 200–350, 340–550 and 540–800 °C corresponding to the reduction of PtO_x, the surface-active oxygen close to the Pt–ceria interface and the bulk ceria,³⁷ respectively (Fig. 1C). Notably, the reducibility of PtO_x at the Pt–CeO₂ interface plays an important role in an oxidation reaction. Interestingly, for 1Pt/20CeO₂-Silicalite1-HIE, H₂ consumption was observed at lower temperature compared with that of 1Pt/20CeO₂-Silicalite1-CON, implying that Pt atoms were highly dispersed on the hierarchical support. These observations also agree with XRD and TEM results. It is therefore reasonable to assume that the hierarchical zeolite would promote highly efficient active sites to activate the interfacial lattice oxygen. These behaviors would increase the catalytic activity for the oxidation reaction. In addition, for 1Pt/

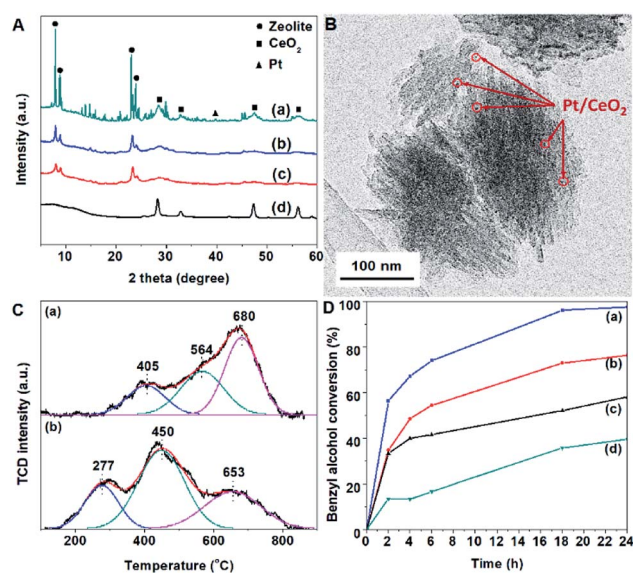


Fig. 1 Structural characterization of catalysts: (A) XRD patterns of nanoceria-modified platinum supported on different supports including (a) 1Pt/20CeO₂-Silicalite1-CON, (b) 1Pt/20CeO₂-Silicalite1-HIE, (c) 1Pt/20CeO₂-ZSM5-HIE, and (d) Pt–CeO₂; (B) TEM image of 1Pt/20CeO₂-Silicalite1-HIE; (C) H₂-TPR profiles of (a) 1Pt/20CeO₂-Silicalite1-CON and (b) 1Pt/20CeO₂-Silicalite1-HIE; (D) Catalytic activity of benzylalcohol oxidation over various catalysts: (a) 1Pt/20CeO₂-ZSM5-HIE, (b) 1Pt/20CeO₂-Silicalite1-HIE, (c) 1Pt/CeO₂ and (d) 1Pt/20CeO₂-Silicalite1-CON (100% selectivity towards benzaldehyde).



20CeO₂-ZSM5-HIE, the reducibility at PtO_x site is similar to what has been described in the case of 1Pt/20CeO₂-Silicalite1-HIE. Meanwhile, more distinct peaks can be observed in the sample of 1Pt/CeO₂ (Fig. S7†) due to the reduction of ceria surface and bulk ceria appeared at 400 and 800 °C, respectively.³⁸

To illustrate the beneficial effect of nanoceria-modified platinum and hierarchical zeolite on the selective oxidation of alcohols to aldehydes, the catalytic activities were tested using different Pt and CeO₂ contents over the metal supported hierarchical silicalite-1 (Silicalite1-HIE) (Fig. 2). Notably, both isolated Pt and CeO₂ exhibit a low catalytic activity for the oxidation of benzyl alcohol to benzaldehyde. However, the catalytic performance can greatly enhance when using a combined Pt and CeO₂ catalyst, in particular, with 1 wt% of Pt and 20 wt% of CeO₂ on hierarchical silicalite-1 as a solid support.

To study the effect of catalyst supports on the catalytic performance, as can be seen in Fig. 1D and 3, the catalytic activity of 1Pt/20CeO₂-Silicalite1-HIE is obviously improved compared with the 1Pt/20CeO₂-Silicalite1-CON. Interestingly, the highest yield of benzaldehyde can be achieved over 1Pt/20CeO₂-ZSM5-HIE. In this case, ZSM5-HIE has the same framework as Silicalite1-HIE but it also consists of the Al sites in the framework, whereas the Silicalite1-HIE contains a pure silica (Table S1†). The reason for the higher catalytic activity of 1Pt/20CeO₂-ZSM5-HIE with respect to 1Pt/20CeO₂-Silicalite1-HIE relates to the fact that the presence of alumina in the solid support can enhance the metal-support interaction.³⁹ In addition, the catalytic performances also relate to the Si/Al of hierarchical zeolites (Fig. 3). The low Si/Al provides the higher number of Al atoms incorporated into the framework, eventually resulting in the improvement of the catalytic activity. However, in the case of 1Pt/CeO₂, the catalytic performance is much lower than that of the hierarchical ones.

To understand the effect of active site structures on the catalytic activity, X-ray absorption near-edge spectroscopy (XANES) technique was used to investigate the oxidation state of

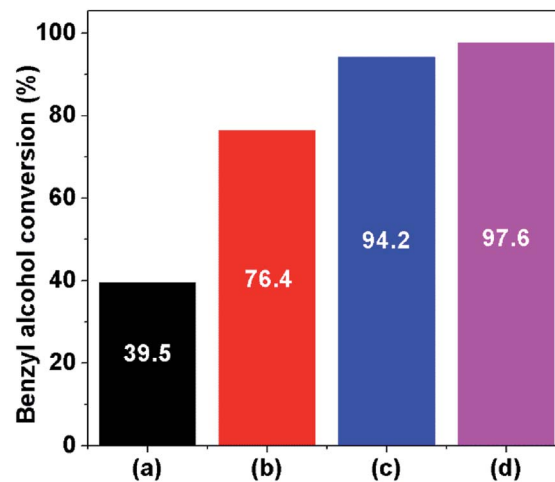


Fig. 3 Catalytic activity of the selective oxidation of benzyl alcohol to benzaldehyde (selectivity of benzaldehyde = 100%) using a batch reactor at 80 °C and 24 h over 1 wt% Pt and 20 wt% CeO₂ supported on different support materials: (a) 1Pt/20CeO₂-Silicalite1-CON; (b) 1Pt/20CeO₂-Silicalite1-HIE; (c) 1Pt/20CeO₂-ZSM5(Si/Al = 400)-HIE; (d) 1Pt/20CeO₂-ZSM5(Si/Al = 100)-HIE.

Ce incorporated into the catalysts. Fig. 4 illustrates the measured Ce L₃-edge XANES spectra of Pt and CeO₂ supported on different materials and two reference standards, Ce(NO₃)₃·6H₂O and CeO₂, which were used as Ce³⁺ and Ce⁴⁺ references, respectively. The peak I is associated to the absorption into the 5d level with the 4f occupancy in the initial state, and therefore it indicates the presence of Ce³⁺ in the sample. While the double peaks (peaks II and III) relating to fully oxidized CeO₂, are associated with the occupancy of the 4f level in the final state and the absorption into the 5d level with no occupancy in the 4f level in either the initial or final state, respectively. Therefore, this peak is only expected when Ce⁴⁺ is present.⁴⁰ The fitting of XANES spectra of the Ce L₃ edge is demonstrated in Fig. S8†. The experimental curve can be reproduced by the combination of an arctan function in order to simulate the edge jump and fit

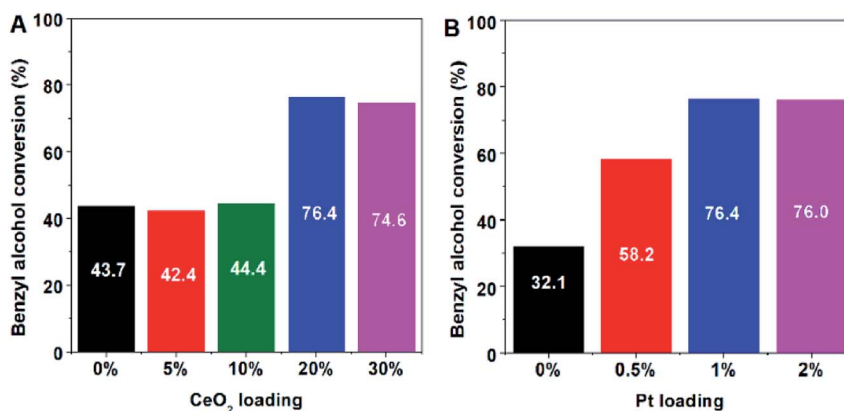


Fig. 2 Catalytic activity of the selective oxidation of benzyl alcohol to benzaldehyde (selectivity of benzaldehyde = 100%) using a batch reactor at 80 °C and 24 h of reaction time over: (A) 1 wt% Pt supported on Silicalite1-HIE with different CeO₂ loading; (B) 20 wt% CeO₂ supported on Silicalite1-HIE with different Pt loading.



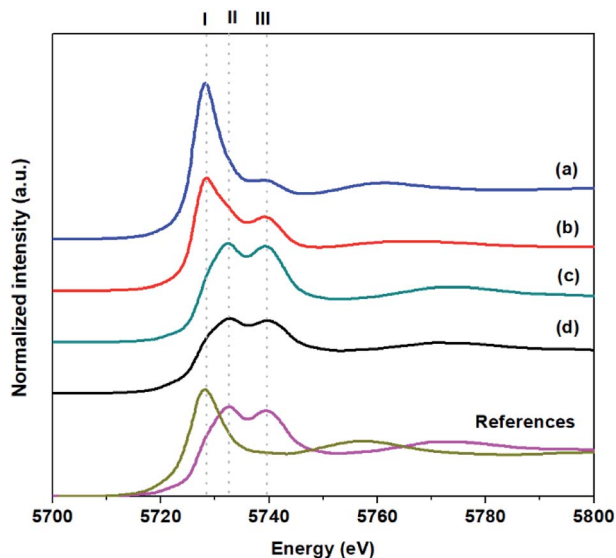


Fig. 4 Normalized Ce L_3 edge XANES spectra of different catalysts: (a) 1Pt/20CeO₂-ZSM5-HIE, (b) 1Pt/20CeO₂-Silicalite1-HIE, (c) 1Pt/20CeO₂-Silicalite1-CON, (d) 1Pt/CeO₂ and standard references of Ce³⁺ (yellow) and Ce⁴⁺ (pink).

three Gaussian functions (I, II, III).⁴¹ The quantitative analysis of Ce oxidation from XANES for prepared catalysts is summarized as shown in Table S2.† In cases of 1Pt/20CeO₂-Silicalite1-CON and 1Pt/CeO₂, their peaks can be absolutely fitted with Ce⁴⁺, whereas 1Pt/20CeO₂-Silicalite1-HIE and 1Pt/20CeO₂-ZSM5-HIE exhibit a mixed phase between two valences with Ce³⁺ as a majority. The results of XANES indicate that support materials affect the oxidation state of Ce, resulting in the difference of catalytic performances. To clearly explain, Ce³⁺

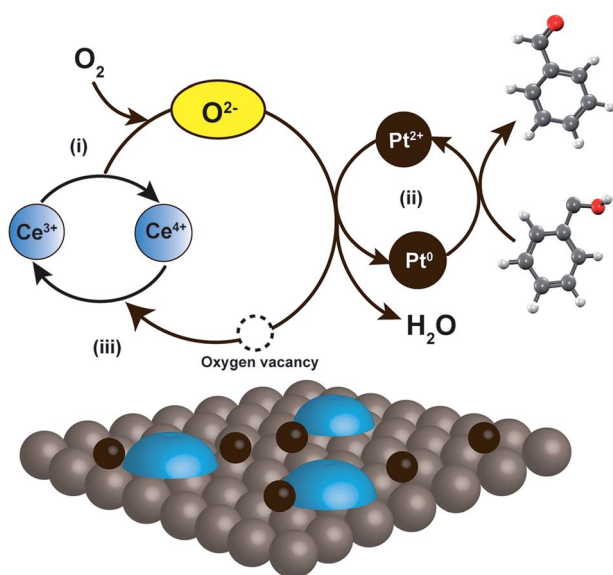


Fig. 5 Schematic illustration of proposed reaction mechanism obtained over nanoceria-modified platinum supported on hierarchical zeolites.

usually comes from small particles of ceria nanoparticles, which is highly well-dispersed on solid supports.⁴² In cases of hierarchical catalysts, they show the higher catalytic activity compared to the conventional catalyst due to the oxygen vacancy of ceria, which can accelerate the mobility of lattice oxygen. However, in the case of 1Pt/CeO₂, the high Ce⁴⁺ content affects a lower catalytic performance compared with hierarchical supports. Interestingly, the oxidation state of Ce did not change after the reaction (Table S2†). This makes it clear that the active sites in this system were recyclable.

To propose the reaction mechanism over hierarchical catalysts, Fig. 5 shows the schematic illustration of the reaction mechanism including the following three main steps: (i) the reduced Ce³⁺ transferring an electron to O₂, and producing Ce⁴⁺ and surface-active oxygen species (O²⁻);⁴³ (ii) the formation of Pt-hydride by H atom adsorption from benzyl alcohol, which reacts with O²⁻ to produce H₂O and an O vacancy; (iii) an O vacancy transferring to Ce⁴⁺, recovering reactive Pt⁰ and Ce³⁺ to produce fresh active sites for next catalytic cycles.⁴⁴ Therefore, these features can be described as an electron bridge that facilitates the electron transfer from reductants to oxidants in the catalytic oxidation process.

Conclusions

The development of hierarchical zeolite catalysts was studied for the selective oxidation of benzyl alcohol to benzaldehyde as a model reaction. The synergic effect of Pt and CeO₂ can greatly enhance the catalytic performance. Compared with a conventional zeolite, hierarchical zeolites strongly affect the catalytic activity because of not only improving the metal dispersion but also increasing Ce³⁺ fraction. Moreover, almost 100% yield of benzaldehyde can be achieved over 1 wt% of Pt combined with 20 wt% of CeO₂ supported on hierarchical ZSM-5 having a low Si/Al ratio. This example demonstrates the development of highly efficient catalyst with reducing the amount of noble-metal, which relates to environmental concerns and cost effectiveness, eventually leading to a potentially improved method for sustainable production of fine-chemicals.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the Vidyasirimedhi Institute of Science and Technology, TTSF research project supported by Thailand Toray Science Foundation, Thailand Research Fund (MRG6180099), and the Office of Higher Education Commission. In addition, this work has been partially supported by the National Nanotechnology Center (NANOTEC), NSTDA, Ministry of Science and Technology, Thailand, through its program of Research Network NANOTEC.



Notes and references

- 1 S. E. Davis, M. S. Ide and R. J. Davis, *Green Chem.*, 2013, **15**, 17–45.
- 2 R. Marotta, I. Di Somma, D. Spasiano, R. Andreozzi and V. Caprio, *Chem. Eng. J.*, 2011, **172**, 243–249.
- 3 J. A. Satrio and L. Doraiswamy, *Chem. Eng. J.*, 2001, **82**, 43–56.
- 4 C. Parmeggiani and F. Cardona, *Green Chem.*, 2012, **14**, 547–564.
- 5 K. R. Seddon and A. Stark, *Green Chem.*, 2002, **4**, 119–123.
- 6 P. Gamez, I. W. Arends, J. Reedijk and R. A. Sheldon, *Chem. Commun.*, 2003, 2414–2415.
- 7 B. Karimi, S. Abedi, J. H. Clark and V. Budarin, *Angew. Chem., Int. Ed.*, 2006, **45**, 4776–4779.
- 8 V. R. Choudhary, A. Dhar, P. Jana, R. Jha and B. S. Uphade, *Green Chem.*, 2005, **7**, 768–770.
- 9 H. Tsunoyama, H. Sakurai, Y. Negishi and T. Tsukuda, *J. Am. Chem. Soc.*, 2005, **127**, 9374–9375.
- 10 P. Maity, C. S. Gopinath, S. Bhaduri and G. K. Lahiri, *Green Chem.*, 2009, **11**, 554–561.
- 11 Q. Wang, X. Cai, Y. Liu, J. Xie, Y. Zhou and J. Wang, *Appl. Catal., B*, 2016, **189**, 242–251.
- 12 M. Marelli, A. Jouve, A. Villa, R. Psaro, A. Balerna, L. Prati and C. Evangelisti, *J. Phys. Chem. C*, 2019, **123**, 2864–2871.
- 13 C. M. Olmos, L. E. Chinchilla, A. Villa, J. J. Delgado, A. B. Hungria, G. Blanco, L. Prati, J. J. Calvino and X. Chen, *J. Catal.*, 2019, **375**, 44–55.
- 14 D. I. Enache, J. K. Edwards, P. Landon, B. Solsona-Espriu, A. F. Carley, A. A. Herzing, M. Watanabe, C. J. Kiely, D. W. Knight and G. J. Hutchings, *Science*, 2006, **311**, 362–365.
- 15 P. S. Murphin Kumar, S. Thiripuranthagan, T. Imai, G. Kumar, A. Pugazhendhi, S. R. Vijayan, R. Esparza, H. Abe and S. K. Krishnan, *ACS Sustainable Chem. Eng.*, 2017, **5**, 11290–11299.
- 16 A. Tanaka, K. Hashimoto and H. Kominami, *J. Am. Chem. Soc.*, 2012, **134**, 14526–14533.
- 17 C. Xu and P. K. Shen, *Chem. Commun.*, 2004, 2238–2239.
- 18 E. Rocchini, A. Trovarelli, J. Llorca, G. W. Graham, W. H. Weber, M. Maciejewski and A. Baiker, *J. Catal.*, 2000, **194**, 461–478.
- 19 S. Wei, H. He, Y. Cheng, C. Yang, G. Zeng, L. Kang, H. Qian and C. Zhu, *Fuel*, 2017, **200**, 11–21.
- 20 C. Yang, K. Zhao, Y. Cheng, G. Zeng, M. Zhang, J. Shao and L. Lu, *Sep. Purif. Technol.*, 2016, **163**, 153–161.
- 21 L. Kang, H. Liu, H. He and C. Yang, *Fuel*, 2018, **234**, 1229–1237.
- 22 L. Qiu, Y. Cheng, C. Yang, G. Zeng, Z. Long, S. Wei, K. Zhao and L. Luo, *RSC Adv.*, 2016, **6**, 17036–17045.
- 23 P. Tamizhdurai, S. Sakthinathan, P. S. Krishnan, A. Ramesh, A. Abilarasu, V. L. Mangesh, S. Narayanan, K. Shanthi and T.-W. Chiu, *Polyhedron*, 2018, **155**, 390–397.
- 24 J. Liu, Q. Yuan, H. Zhao and S. Zou, *Catal. Lett.*, 2018, **148**, 1093–1099.
- 25 S. S. Moeini, C. Battocchio, S. Casciardi, I. Luisetto, P. Lupattelli, D. Tofani and S. Tuti, *Catalysts*, 2019, **9**, 847.
- 26 J. Xu, J.-K. Shang, Y. Chen, Y. Wang and Y.-X. Li, *Appl. Catal., A*, 2017, **542**, 380–388.
- 27 C. S. Ramirez-Barria, M. Isaacs, C. Parlett, K. Wilson, A. Guerrero-Ruiz and I. Rodríguez-Ramos, *Catal. Today*, 2019, DOI: 10.1016/j.cattod.2019.05.057.
- 28 A. Corma, *J. Catal.*, 2003, **216**, 298–312.
- 29 L. Meng, B. Mezari, M. G. Goesten and E. J. M. Hensen, *Chem. Mater.*, 2017, **29**, 4091–4096.
- 30 L. Meng, X. Zhu, W. Wannapakdee, R. Pestman, M. G. Goesten, L. Gao, A. J. F. van Hoof and E. J. M. Hensen, *J. Catal.*, 2018, **361**, 135–142.
- 31 J. Pérez-Ramírez, C. H. Christensen, K. Egeblad, C. H. Christensen and J. C. Groen, *Chem. Soc. Rev.*, 2008, **37**, 2530–2542.
- 32 W. Wannapakdee, C. Wattanakit, V. Paluka, T. Yutthalekha and J. Limtrakul, *RSC Adv.*, 2016, **6**, 2875–2881.
- 33 M. Fadoni, L. Lucarelli and A. Dąbrowski, *Stud. Surf. Sci. Catal.*, 1999, **120**, 177–225.
- 34 X. Wang, D. Liu, S. Song and H. Zhang, *J. Am. Chem. Soc.*, 2013, **135**, 15864–15872.
- 35 F. J. Sotomayor, K. A. Cychosz and M. Thommes, *Acc. Mater. Surf. Res.*, 2018, **3**, 34–50.
- 36 V. Perrichon, L. Retailleau, P. Bazin, M. Daturi and J. C. Lavalley, *Appl. Catal., A*, 2004, **260**, 1–8.
- 37 J. Ke, W. Zhu, Y. Jiang, R. Si, Y.-J. Wang, S.-C. Li, C. Jin, H. Liu, W.-G. Song and C.-H. Yan, *ACS Catal.*, 2015, **5**, 5164–5173.
- 38 R. Singha, A. Shukla, A. Yadav, T. Sasaki, A. Sandupatla, G. Deo and R. Bal, *Catal. Sci. Technol.*, 2017, **7**, 4720–4735.
- 39 J. Chen, Q. Zhang, Y. Wang and H. Wan, *Adv. Synth. Catal.*, 2008, **350**, 453–464.
- 40 S. Overbury, D. Huntley, D. Mullins and G. Glavee, *Catal. Lett.*, 1998, **51**, 133–138.
- 41 S. Phokha, S. Pinitsoontorn and S. Maensiri, *Nano-Micro Lett.*, 2013, **5**, 223.
- 42 C. M. Sims, R. A. Maier, A. C. Johnston-Peck, J. M. Gorham, V. A. Hackley and B. C. Nelson, *Nanotechnology*, 2018, **30**, 085703.
- 43 Y. Chen, J. Chen, W. Qu, C. George, M. Aouine, P. Vernoux and X. Tang, *Chem. Commun.*, 2018, **54**, 10140–10143.
- 44 Z. Guo, B. Liu, Q. Zhang, W. Deng, Y. Wang and Y. Yang, *Chem. Soc. Rev.*, 2014, **43**, 3480–3524.

