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Theoretical investigation of various aspects of two dimensional holey boroxine, B_3O_3 †

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By means of first-principles calculations, we study the structural, electronic and mechanical properties of the newly synthesized boron–oxygen holey framework (*Chem. Comm.* 2018, **54**, 3971). It has a planar structure formed by B_3O_3 hexagons, which are joined via strong covalent boron–boron bonds. The six B_3O_3 units are connected with six-fold symmetry exhibiting a large hole with a surface area of 23 \AA^2 , which is ideal for the adsorption of alkalis. For neutral alkalis, we found that the adsorption energy of potassium is 14 and 12 kcal mol⁻¹ larger than those determined for sodium and lithium, respectively. In contrast, for alkali cations, there is a clear preference for lithium over sodium and potassium. With regard to its electronic properties, it is an insulator with an electronic band gap of 5.3 eV, at the HSE level of theory. We further investigate the effect of strain on the band gap and find it a less efficient technique to tune the electronic properties. The wide optical gap of B_3O_3 can be utilized in ultraviolet (UV) applications, such as UV photodetectors, etc. Additionally, the 2D elastic modulus of B_3O_3 (53.9 N m⁻¹) is larger than that of Be_3N_2 , silicene, and germanene. Besides, we also report bilayer and graphite-like bulk B_3O_3 and furthermore, find that the optoelectronic properties of the bilayer can be tuned with an external electric field. The great tunability of optical properties from UV to the visible range offers a vast range of applications in optoelectronics.

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1. Introduction

The preparation of graphene in 2004 opened new research directions in all branches of science.¹ It is difficult to make a full list of the two dimensional (2D) materials available. Without trying to be exhaustive, we can mention the transition metal chalcogenides^{2,3} MoS_2 , WS_2 , $MoSe_2$, WSe_2 , $MoTe_2$, the MXenes,⁴ which are layered compounds formed by transition metal nitrides or carbides, the p-block based 2D materials like silicene,^{5,6} germanene,⁷ antimonene,⁸ bismuthene,⁹ black¹⁰ and blue phosphorene¹¹ as well as graphyne¹² and the isoelectronic analog of graphene, boron nitride^{13,14} among many others. Upon reading the latter list it is possible to postulate that 2D materials can only exist when their bulk materials are layered. However, it has been shown that it is possible to synthesize an atomic single layer material when the bulk is not layered, being silicene a clear example. Recently, Stredansky *et al.*¹⁵ reported the synthesis of a novel 2D material based on the boronic acid condensation. Using three tetrahydroxyboron ($H_2B_2O_4$) molecules, a six-membered boroxine ring molecule was formed

along with three molecules of water. Although the boroxine layers could readily be identified in the STM studies, the authors suggested that further improvements in the synthetic procedure should be made in order to reduce the occurrence of non-six-fold symmetry motifs. Concomitantly, Lin *et al.*¹⁶ proposed a porous boron oxide monolayer with a robust band gap, which is similar to the one synthesized by Stredansky *et al.*¹⁵ They showed that boroxine is perfectly planar with uniform pores with a 6.27 \AA diameter and exhibit an indirect wide band gap with the possible applications in optoelectronics. This work is the extension to the previous predicted (theoretical) and synthesized (experimentally) boroxine by exploring the microscopic science and including new interesting capabilities such as the high affinity for alkalis and the study of bilayer boroxine as well as the bulk layered material. As by-product, we have found a minor discrepancy in the electronic structure of boroxine. In effect, our HSEH1PBE calculations indicated that the gap is direct, in contrast with the results reported by Lin *et al.*¹⁶ In the text, the only “ B_3O_3 ” will refer to the monolayer B_3O_3 unless it is followed by “ring” or “hexagon”.

2. Methods

We study boroxine (B_3O_3) by means of spin-polarized vdW-DF,¹⁷ M06-L,¹⁸ and HSEH1PBE¹⁹ density functional theory (DFT) calculation. We use SIESTA^{20,21} code to perform vdW-DF calculations. We select the double-zeta basis set with polarization

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functions and fixed the orbital confining cut-off to 0.01 Ry with the default value (0.15) of split-norm. The interaction between ionic cores and valence electrons is described by the Troullier-Martins norm-conserving (NC) pseudopotentials.²² Geometry optimizations were carried out using the conjugate gradient algorithm until all the residual forces were smaller than 0.01 eV Å⁻¹. The unit cells were optimized and they were sampled using a 50 × 50 × 1 (about 900 *k*-points gamma centered) Monkhorst-Pack sampling, the periodic M06-L and HSEH1PBE calculations were carried out with Gaussian 09.²³ The basis set selected were Pople's 6-31G* and 6-311G*.²⁴ In general, we did not observe significant differences in the results obtained using both basis sets. 1000 *k*-points were used to sample the unit cell and the ultrafine grid was employed. For comparative calculations, we also investigate the adsorption of alkalis employing a boroxine (B₃O₃) cluster formed by six B₃O₃ units and terminated with hydrogen atoms. In general, the results indicate the same trend as that obtained employing infinite models.

We also perform PBE-D2 (ref. 25) and PBE-D3 (ref. 26) calculations with the help of VASP code.²⁷ We use generalized gradient approximation (GGA) with the flavor proposed by Perdew, Burke, and Ernzerhof (PBE)^{28,29} assisted with the projector augmented wave (PAW).³⁰

Finally, we also use a hybrid functional (not in a true sense) which is a mixture of LDA-GGA^{28,29,31} as implemented in SIESTA, especially for the calculation of in-plane elastic modulus to compare our results with our previous work on graphene and Be₃N₂.³² The details are in ref. 31. Throughout the manuscript, the hybrid is referring to the aforementioned method. Finally, we also report *ab initio* molecular dynamics (AIMD) simulations at 300–600 K with the aid of SIESTA code. Further details regarding the computational methods can be seen in the ESI.†

3. Results and discussion

3.1 Structural properties of 2D boroxine

As suggested by the experimental study¹⁵ and the theoretical study by Lin *et al.*,¹⁶ we find that 2D boroxine can adopt six-fold symmetry. The unit cell employed in our calculations is shown in Fig. 1.

The six-fold symmetry structure is obtained by joining the B₃O₃ hexagons obtained after the condensation of three tetrahydrodiboron molecules. The B–B bonds are 1.71–1.72 Å long, a value which is bracketed by the B1–B1 and B1–B2 bond distances found in *Pmmn* borophene, namely, 1.613 and 1.879 Å, respectively, which are in excellent agreement with the PBE results previously reported.¹⁶ The B–O bond length 1.38–1.39 Å is among the highest values of which B–O can adopt in molecular forms. As suggested by Coulson,³² this may indicate a moderate to weak participation of the π orbitals of boron and oxygen. A fact that seems to be corroborated by the partial density of states presented in Fig. 2. When the six B₃O₃ molecules are joined, they leave a hole in which the distance between the opposite O atoms is 6.21–6.28 Å (depending on the level of theory) and the surface is close to 23 Å², ideal for the adsorption of large atoms or molecules, as observed for crown ethers. For example, the 18-crown-6 ether has a high affinity for the

hydronium ion H₃O⁺. In addition to this, it allows the dissolution of KMnO₄ in benzene thanks to the complexation of the potassium ion. We show in Section 3.8 that 2D boroxine has also a strong affinity towards alkalis, especially to potassium.

To report the structural stability, we calculate the formation energy and since we use the atomic form of B and O as the chemical potential, the formation energy here is equal to the cohesive energy. We make use of the following expression:

$$E_{\text{coh}} = \frac{E_{\text{B}_3\text{O}_3} - 6E_{\text{B}} - 6E_{\text{O}}}{12}, \text{ where } E_{\text{B}_3\text{O}_3} \text{ is the total energy of the unit cell containing } 6\text{B and } 6\text{O atoms. } E_{\text{B}} \text{ and } E_{\text{O}} \text{ are taken as the gas-phase energies of B and O, respectively. The cohesive energy of } \text{B}_3\text{O}_3 \text{ PBE-D3 level is } -6.72 \text{ eV per atom, just } 0.04 \text{ lower than the value reported by Lin } \text{et al.}^{16} \text{ For the comparison, the cohesive energy of graphene at the same level of theory is } -8.07 \text{ eV per atom. This means that } \text{B}_3\text{O}_3 \text{ has a comparable strength to that of graphene. At the M06-L level, the cohesive energy of } \text{B}_3\text{O}_3 \text{ } -6.3 \text{ eV per atom which is twice the value calculated for phosphorene } (-3.1 \text{ eV per atom}) \text{ at the same level. Further information is given in ESI.† The data regarding structural properties and stability at various levels of theory is shown in Table S1.†}$$

3.2 Dynamical stability

In the experiment, 2D boroxine (B₃O₃) was grown on Au (111) surface and, therefore, it is necessary to check the dynamical stability of free-standing boroxine (B₃O₃). To that end, we calculate phonon spectrum by supercell approach as implemented in SIESTA code along the G–M–K–G high symmetry points. We use a 2 × 2 × 1 supercell to calculate the phonon band spectrum, whereas, a 1 × 1 × 1 cell also give converged result, keeping in mind that the unit cell size is already large enough. The phonon spectrum graph is shown in Fig. 2 where it is clear that 2D boroxine is certainly a local minimum in the potential energy surface due to the absence of any negative frequency. The highest frequency of B₃O₃ is \sim 1450 cm⁻¹.

3.3 Electronic properties

The electronic properties of 2D boroxine are predicted by calculating the electronic band structure and for additional comprehension, the orbital projected density of states (PDOS) is also computed. With SIESTA (both the levels employed), 2D boroxine (B₃O₃) is a wide and indirect band gap material with conduction band minima (CBM) lie at M point, while valence band maxima (VBM) at gamma point. However, the energy difference between gamma and M points is quite low (55 meV). This indirect gap nature is also confirmed by VASP code with a slight disagreement as the CBM appears at the K point which is in line with the previous findings by Lin *et al.*¹⁶ The calculated value of band gap, at vdW-DF level is 3.80 eV. However, the M06-L level of theory revealed a direct gap with a value of 3.9 eV. As expected, the band gap was significantly underestimated by the vdW-DF method and also by the M06-L functional. In effect, HSEH1PBE calculations indicated that the band gap of boroxine is direct, namely 5.30 eV which is a bit higher than the previously reported value of 5.23 eV. The latter value is smaller than



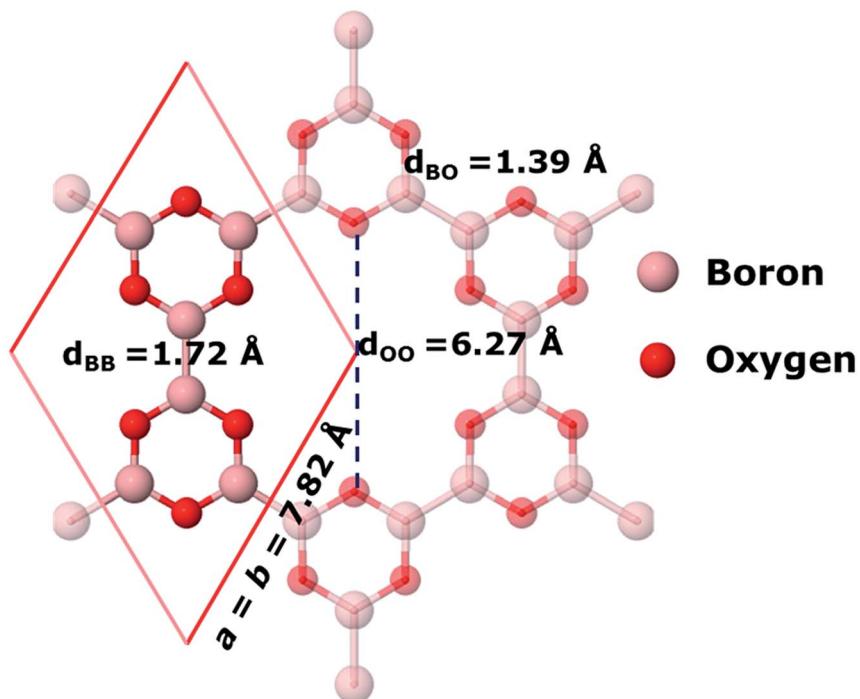


Fig. 1 The unit cell and bond lengths of 2D boroxine (B_3O_3) at PBE-D3 level of theory. Periodically repeated atoms are made translucent.

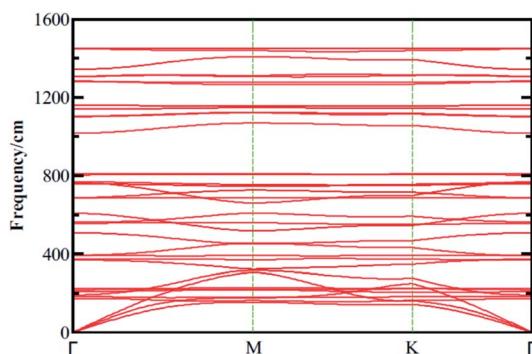


Fig. 2 Phonon dispersion curves calculated for boroxine (B_3O_3) calculated at vdW-DF/DZP level of theory with SIESTA code.

the gap computed for boron nitride at the HSEH1PBE level, namely, 6.1 eV. Moreover, B_3O_3 bears the same direct-indirect band gap controversy as that of boron nitride.

From the PDOS, the CBM has a π character and mostly composed of B-p_z with a slight involvement from O-p_z. However, the O-p_(x+y) strongly hybridized with B-p_(x+y) to form the VBM and, therefore, has a σ -like character. Furthermore, the valence and conduction bands are mostly governed by O and B, respectively. Additionally, this outcome also reflects in LDOS plots shown in Fig. 3 where the density around O atoms can be seen in VB case, whereas, for CB, the density can be spotted around B and O with a larger LDOS amplitude on B in comparison with O. Besides, we also check the effect of strain on the electronic structure of B_3O_3 and find that it is a less efficient technique to alter the electronic properties. The band structures are given in Fig. S1 in ESI.†

3.4 Chemical analysis and bonding character

We calculate the charge transfer with the help of various charge transfer mechanisms namely Hirshfeld and Voronoi as in SIESTA, Mulliken as in Gaussian, and Bader as in VASP. Each of these methods reveals qualitatively the same results of charge transfer from B to O as in the spirit of their electronegativity values. The total charge density calculated at PBE-D3 level is plotted in Fig. 4(a), where charge accumulation on O atoms can be seen quite vividly. The quantitative results are given in Table S1.† Furthermore, to picture the electron localization in B_3O_3 , we calculate electron localization function (ELF) which is quite useful in predicting the bonding character, as well as the stability mechanism. The details regarding the interpretation of ELF plots can be seen in ref. 31 and 33. In B_3O_3 rings, we see electron localization on O atoms broadens towards B atoms and a complete charge delocalization in the vicinity of B atoms. This is in perfect agreement with the previous experimental + DFT study on B_3O_3 rings held together by phenyl molecules where the authors reported the ultra-fast charge delocalization with the charge localization on O atoms.³⁴ Such an analysis, however, is absent for phenyl-free B_3O_3 rings which are bonded *via* B atoms. Consequently, the ELF becomes crucial and much more meaningful. Interestingly, we find a strong charge localization between B-B atoms, the region between two B_3O_3 rings. This indicates that the B_3O_3 rings are held together by strong covalent bonds with the ionic bonds (by virtue of charge transfer) present between B-O. This strong covalent bond might be responsible for the stability of B_3O_3 monolayer consisting of B_3O_3 rings. Moreover, we also calculate simulated scanning tunneling microscopy (STM) image (based on Tersoff-Hamann



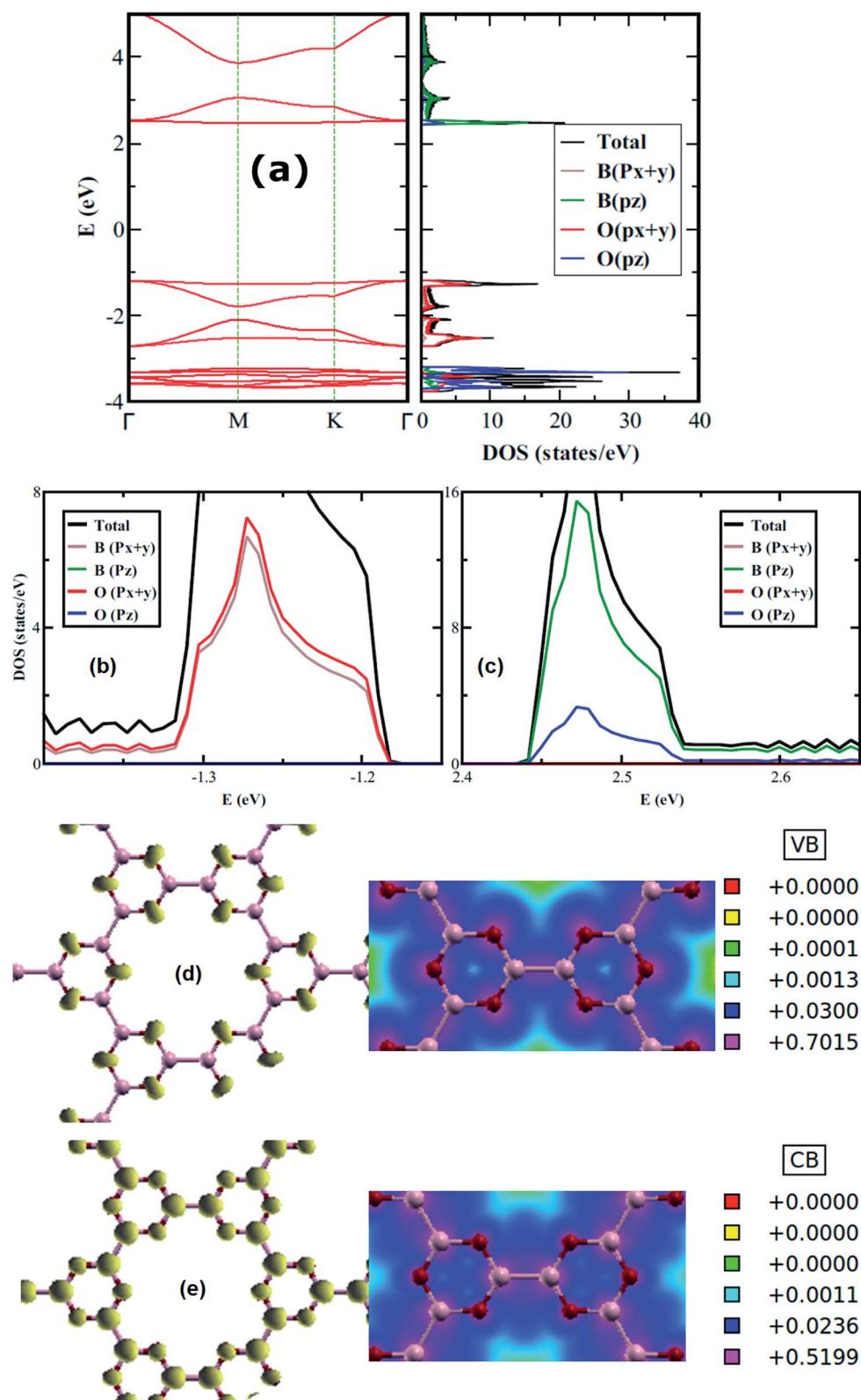


Fig. 3 Electronic band structure along with (total and projected) density of states (a); projected density of states near the (b) valence band maximum and (c) the conduction band minimum; 3D plots along with their 2D cuts of LDOS of filled (d) and empty states (e) of boroxine calculated at the hybrid level of theory.

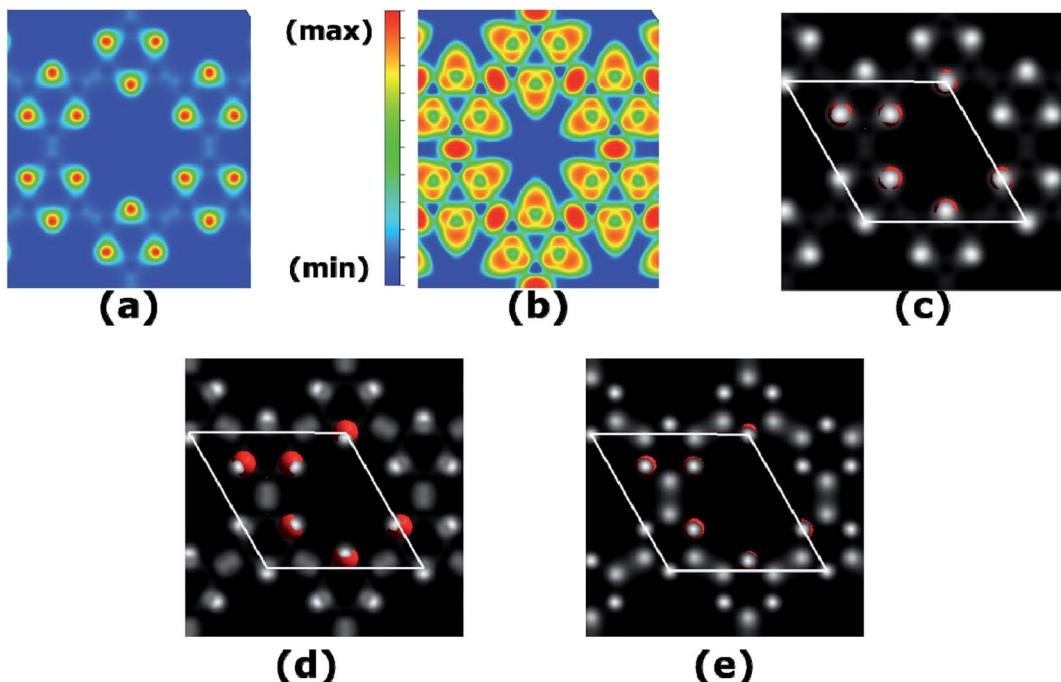


Fig. 4 (a) Total charge density, (b) ELF, where red means strong electron localization and blue means the opposite, whereas, the green represents a region with a homogenous electron density, (c) STM image of B_3O_3 , (d) STM image of filled and (e) empty states.

approximations)³⁵ of B_3O_3 monolayer which is given in (c) and is in agreement with the experimental result. O atoms in the unit cell are shown for the sake of brevity. In addition, we also calculate STM images of the filled and empty states which are in agreement with the PDOS (and LDOS) calculations.

3.5 Kinetic stability

We also test boroxine with high temperatures (300 and 600 K) to see the response. The initial and final structures treated at 600 K are shown in Fig. 5. The negligible structural modifications suggest that B_3O_3 is quite stable at high temperatures. This point has already been proved experimentally. In addition, we also calculate the average change appeared in the B–O and B–B

bond lengths at a temperature of 600 K. We find a slight variation of 0.986% and 0.602% in B–O and B–B bonds, respectively. That being said, boroxine can be used in high temperatures environment, such as space, aircraft, oil, and gas research applications.³⁶ Nevertheless, it will be interesting to see the highest temperature range that B_3O_3 can withstand and, therefore, propose a separate investigation to find the melting point of B_3O_3 .

3.6 Mechanical properties

In Fig. 6, we present the stress curve employed to calculate the in-plane elastic modulus (Y_{2D}) of 2D boroxine. We vary the lattice constant in small steps and observe the energy change as

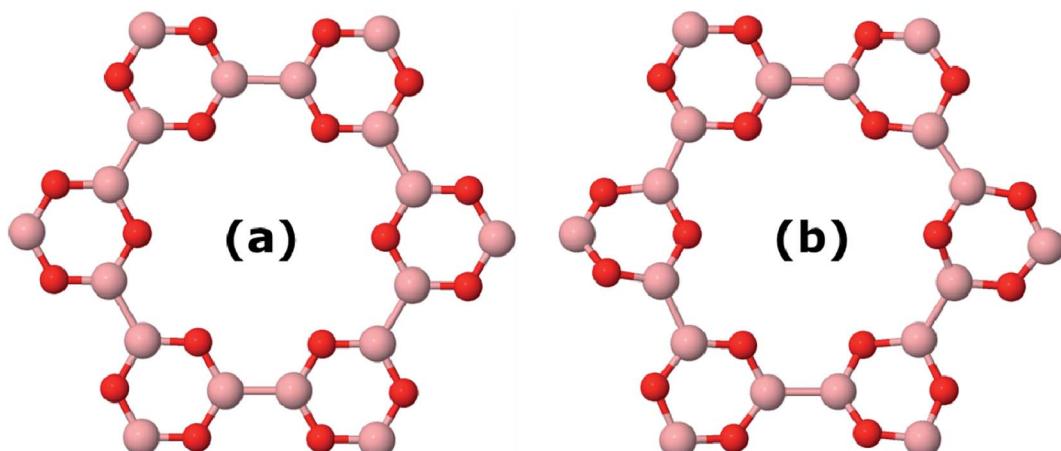


Fig. 5 Snapshots of the results of AIMD of boroxine at (a) 0 K, and (b) 600 K. The negligible distortion in the structure reflects its robustness.

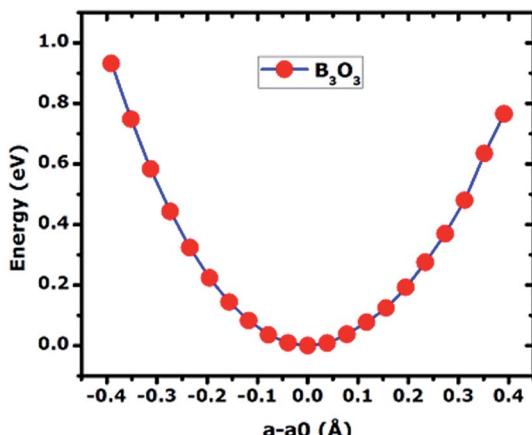


Fig. 6 The curve of total energy (in eV) vs. change in lattice constant (in Å) calculated of 2D B_3O_3 at the hybrid level of theory. The in-plane stiffness is calculated by taking the second derivative at the minimum.

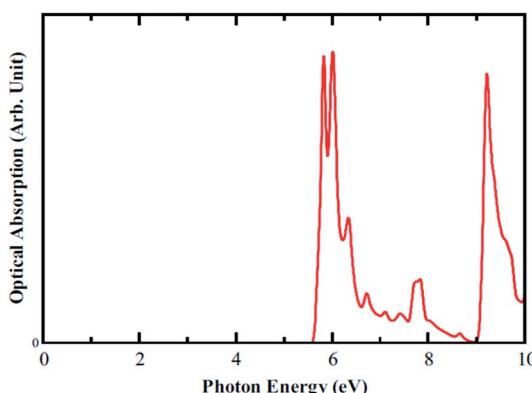


Fig. 7 Optical Absorption spectra calculated for B_3O_3 , at the hybrid level of theory.

Table 1 Adsorption energies for lithium, sodium, and potassium at the M06-L/6-31G* level of theory

	Infinite boroxine	$(B_3O_3)_6H_6$ flake
Li	-52.0	-55.2
Na	-50.6	-54.2
K	-64.1	-67.8
Li^+		-60.7
Na^+		-51.63
K^+		-47.3

a function of lattice parameter. The second derivative from the plot gives the Y_{2D} and is given

by: $Y_{2D} = A_0 \frac{\partial^2 E}{\partial A^2} \Big|_{A_0} = \frac{\sqrt{3}}{6} \frac{\partial^2 E}{\partial a^2} \Big|_{a_0}$, $A = \frac{\sqrt{3}a^2}{6}$, represents the 2D area of B_3O_3 and a is the lattice constant. At the hybrid level, the 2D bulk modulus of B_3O_3 is 53.9 N m^{-1} , about 17.93 N m^{-1} smaller than the values reported by Lin *et al.*¹⁶ The latter values are significantly smaller than the one determined for graphene

(206 N m^{-1}) in our earlier study at the same level of theory.³¹ The earlier reported value of Y_{2D} of graphene (206.6 N m^{-1}), hBN (177 N m^{-1}), SiC (116.5 N m^{-1}), and GeC (101 N m^{-1}) are higher than that of B_3O_3 . However, this value is slightly higher than that of Be_3N_2 (46.5 N m^{-1}),³¹ silicene (44.5 N m^{-1}), and germanene (29.6 N m^{-1}).³⁷

3.7 Optical properties

Optical absorption spectra are also calculated for parallel polarization without taking into account the excitonic effect. The onset of optical absorption, which is the optical gap, appears around 5.7 eV as can be seen in Fig. 7. This value is sufficiently larger than the electronic band gap of 3.66 eV, meaning that the optical transition from first valence to first conduction band is forbidden as per dipole selection rules. This can be understood by looking at the orbital character shown in PDOS plot. The first transition occurs from the blue peak to the green peak (p_z – p_z). Moreover, the separation between these two pi bands is around 5.7 eV, thus agreeing well the optical gap. Additionally, the optical gap is in the deep UV range.

3.8 Adsorption of alkalis

As expressed in the first section, based on the chemistry displayed by the 18-crown-6 ether, we expect a strong affinity of 2D boroxine towards potassium. In Table 1, we list the adsorption energies computed for lithium, sodium, and potassium using finite and infinite models. The calculations were performed at the M06-L/6-31G* level of theory and we used a larger unit cell composed by six B_3O_3 rings so that the unit cell contains a complete hole (as shown in Fig. 1, but including the translucent atoms). The adsorption energies of lithium, sodium, and potassium on the hexagonal hole of boroxine are -52.0 , -50.6 and $-64.1 \text{ kcal mol}^{-1}$, respectively. Large differences, which may indicate that it is possible to separate selectively adsorb the latter alkali. However, for the adsorption of cations, the situation is reversed and lithium becomes clearly favored. In addition to this, it will also be interesting to integrate a metal atom in the hole of B_3O_3 and utilize it as a B_3O_3 -supported metal single-atom catalyst, just like C_2N .^{38,39} With regard to the electronic properties, in Fig. 8 we present the band structure and total and projected density of states for boroxine with a potassium atom adsorbed in the hole. The band gap determined for boroxine with a Li, Na and K atoms adsorbed in the hole are: $0.16/3.86$, $0.03/3.79$ and $0.08/3.93 \text{ eV}$, respectively, at the M06-L/6-31G* level of theory (spin-up/spin-down notation). Thus, alkali atom adsorption dramatically changes the gap for one spin channel while the other is minimally affected. Therefore, it may be possible to attain selective pins polarized transport using these elements.

3.9 Bilayer boroxine

We also calculate B_3O_3 bilayers in AA, AB, and AB2 stacking as shown in Fig. 9. In AA stacking, the interlayer distance is about 3.57 \AA . In AB stacking, the second layer is translated about 1.39 \AA (the B–O bond length) along the x -axis, whereas, in AB2 stacking, this translation is about 2.78 \AA (the next neighboring



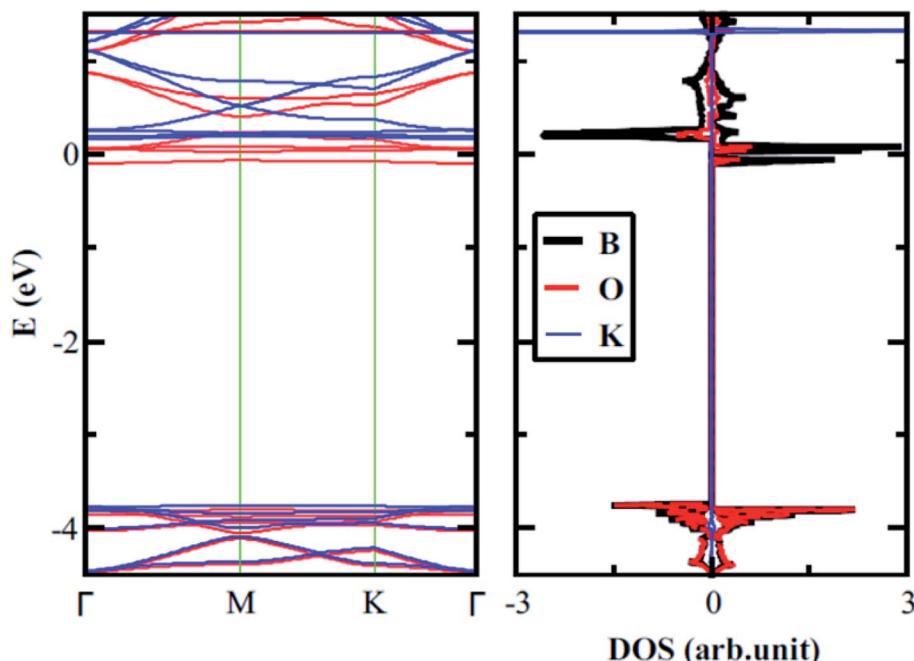


Fig. 8 Band structure and density of states determined for boroxine with a potassium atom adsorbed in the central hole, at the VDW-DF/DZP level of theory.

B–O bond length) at PBE-D3 level. The interlayer separation in AB ($h_1 = 3.18$, $h_2 = 3.3$ Å) is larger than that of AB2 ($h_1 = 3.02$, $h_2 = 3.28$ Å). Here, h_1 is the interlayer separation between B–O (or O–B), whereas, h_2 refers to the separation between B–B (O–O) in AB (AB2) bilayer. We show the periodic structures of these bilayers in Fig. S2 bilayer in ESI.† For comparison, the interlayer

distance in bilayer graphene at PBE-D3 level is 3.33 Å, which is larger than that of AB2 stacking. Furthermore, the interlayer separation has a direct impact on the stability of the bilayers, AB2 being the most favorable, whereas, AA is the least. We also calculate the interlayer interaction energy (IIE) as:

$$\text{IIE} = \frac{2E_1 - E_{\text{tot}}}{N}, \text{ where } E_1 \text{ and } E_{\text{tot}} \text{ are the energies of}$$

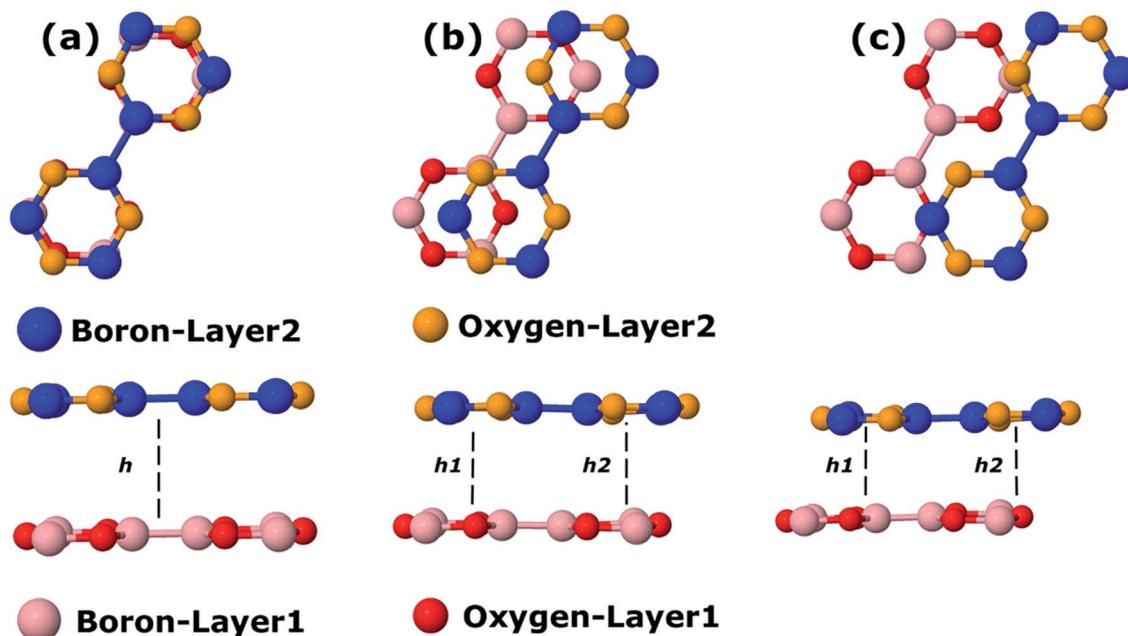


Fig. 9 Top and side view bilayer B₃O₃ with (a) AA, (b) AB, and (c) AB2 stacking. In the AA pattern, h represents the interlayer distance. In the other two stackings, h_1 and h_2 correspond to the separation between B–O (O–B) and B–B (O–B), respectively.

Table 2 Relative energy (in meV), interlayer interaction energy (IIE) (in meV per atom), exfoliation energy (E_{exf}) (in meV A^{-2}), and interlayer separation at PBE-D3 level in B_3O_3 patterns. The h represents interlayer distance in the AA pattern, whereas, h_1/h_2 represents B–O/B–B (or O–O) separation in different layers of other patterns

	Relative energy (meV)	IIE (meV per atom)	E_{exf} (meV A^{-2})	Interlayer separation, $h, h_1/h_2$ (Å)
B_3O_3	743.71			
AA-bilayer	589	6.45		3.57
AB-bilayer	443	12.5		3.18/3.3
AB2-bilayer	358	16.05		3.02/3.28
AA-bulk	447	12.4	2.8	3.48
AB-bulk	150	24.7	5.6	3.15/3.15
AB2-bulk	0	31	7	3.09/3.11
Graphene-bilayer		24		3.33
Graphite		52.7	20	3.34

monolayer and bilayer, respectively. In addition, we compare the outcome with bilayer graphene. Interestingly, we find that IIE in AB2 stacking is comparable to that of bilayer graphene. Table 2 is decorated with relevant data at the PBE-D3 level. The results obtained using other methods are gathered in Table S2.† The electronic band structures are given in Fig. 10. Furthermore, the bilayer gaps are found to be lower than that of the single layer.

We also consider the effect of an external electric field on the electronic structure of AB2 stacked bilayer B_3O_3 . We keep on increasing the field until we reach a limit where the gap becomes zero. The band structure plots are shown in Fig. S3 in ESI† and the band gap values as a function of the increasing field are shown in Fig. 11. We divide the plot into three parts where green, orange, and red colors represent reasonably-high, very-high, and extremely-high external electric field range. It is not known how much external electric field can be applied to

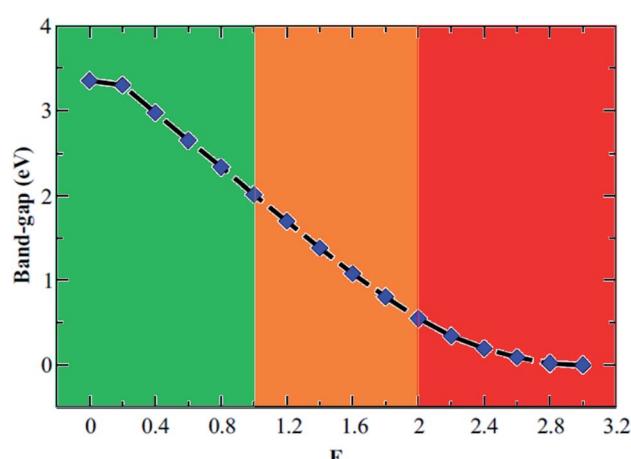


Fig. 11 Band gap (in eV) variation as a function of external electric field (in V A^{-1}) for AB2 stacked bilayer boroxine.

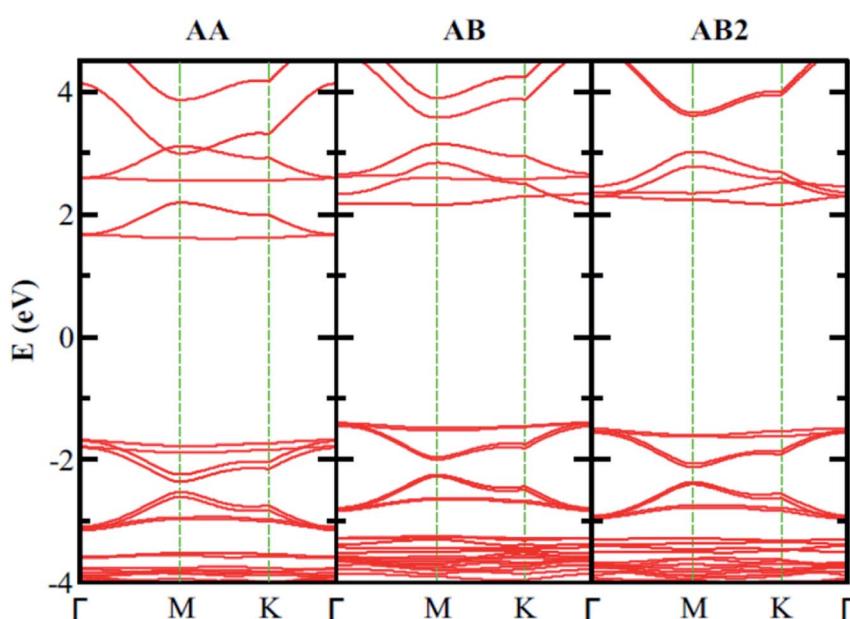


Fig. 10 Electronic band structure of bilayer B_3O_3 with different stacking at vdW-DF level. The band gaps in AA, AB, and AB2 stacking are 3.28 eV, 3.56, and 3.65 eV, respectively, at the vdW-DF level.



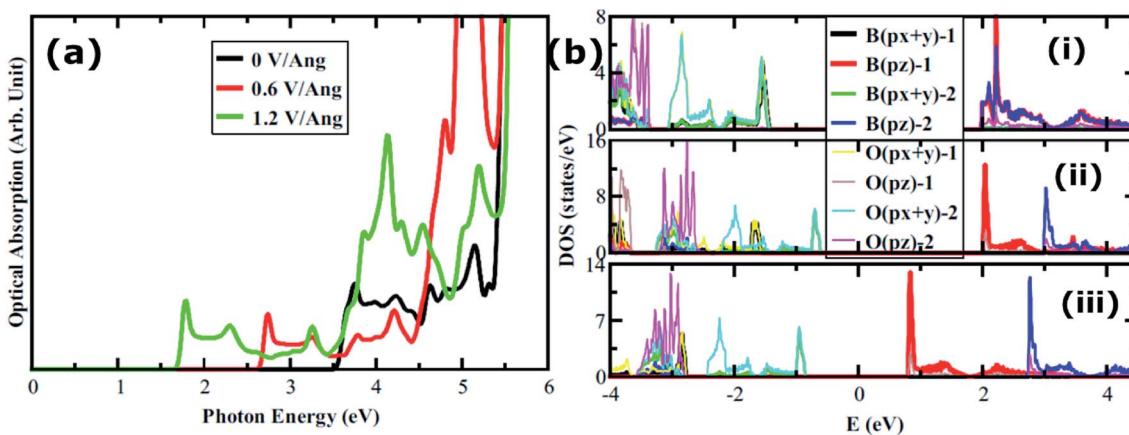


Fig. 12 (a) Optical absorption spectra of AB2 bilayer at various values of the external electric field. (b) Layer-resolved PDOS of AB2 bilayer (i) at 0, (ii) 0.6, and (iii) 1.2 V \AA^{-1} .

B_3O_3 until it reaches its breakdown region. However, this study can motivate both theoreticians and experimentalists to initiate such an investigation. Moreover, it can be noted that the external electric field can significantly alter the electronic properties. The gap reaches a value of 1.7 eV at a field of 1.2 V \AA^{-1} from 3.47 eV (at 0 V \AA^{-1}). In addition to this, we also calculate the optical properties of AB2 bilayer with and without the external electric field. Here, we consider the structure at 0, 0.6, and 1.2 V \AA^{-1} external electric field. The layer-resolved PDOS along with the optical absorption spectra is given in Fig. 12. Unlike the monolayer, it seems that the σ to π transition becomes allowed by the dipole selection rule in bilayer. Furthermore, there induce some extra peaks which seem to be coming from the transition between the broad B–O– σ band and the narrow B–O– π band. Most interestingly, these peaks correspond to interlayer transition and, furthermore, can be tuned

right in the visible range. Nevertheless, those intense peaks in the deep UV range also look like the result of an interlayer transition. In photovoltaics, such transitions have an edge over the intralayer ones.^{40–44} However, we propose that more accurate calculations, such as GW/BSE are required to confirm these outcomes.

3.10 Graphite-like bulk B_3O_3

Finally, we consider graphite-like B_3O_3 in bulk form which is a new phase of B_3O_3 at PBE-D3 level. We consider the same patterns as that of the bilayers. From the total energy calculations, it is evident that these patterns are local minima in the total energy surface, again AB2 pattern being the most favorable. The same outcome is confirmed by all the employed levels of theory. In bulk form, they exhibit the same band gaps like

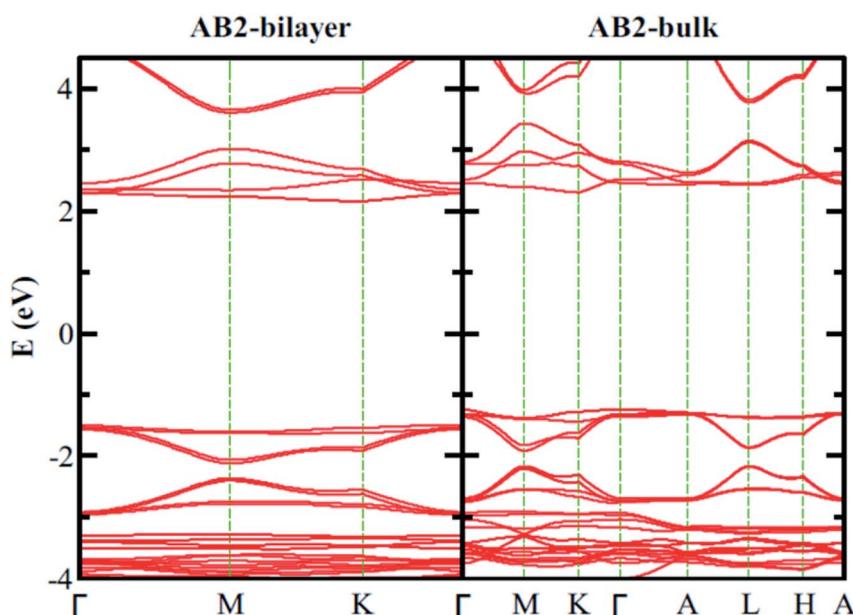


Fig. 13 Comparison of the electronic band structure of AB2 stacking of B_3O_3 in bilayer and bulk form at vdW-DF level.



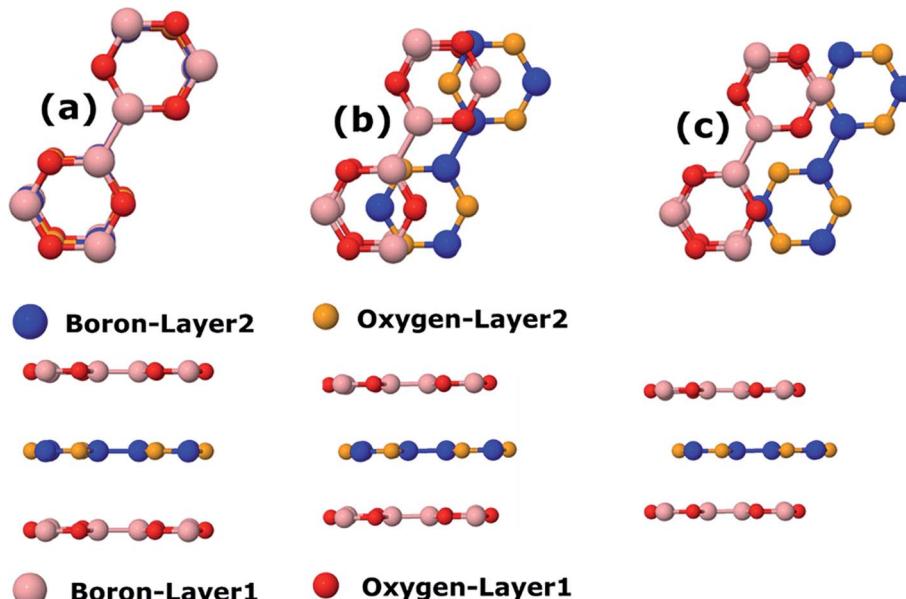


Fig. 14 Top and side view of (a) AA, (b) AB, and (c) AB2 patterns of B_3O_3 in bulk form at PBE-D3 level.

that in the bilayer-form and as a comparison, the bilayer and bulk band structures of AB2 stacking are given in Fig. 13. We also calculate the exfoliation energy as reported elsewhere using the following equation:³¹ $E_{\text{exf}} = \frac{E_{\text{B}_3\text{O}_3} - E_{\text{g-B}_3\text{O}_3}}{A_{\text{B}_3\text{O}_3}}$, where the energy difference between monolayer B_3O_3 and graphite-like bulk B_3O_3 is divided by the monolayer area. From Table 2, it is evident that B_3O_3 in bulk form is not only comparably robust to that of graphite but the monolayer can be exfoliated with a lot ease as well. The geometries are given in Fig. 14.

4. Conclusions

We employed first-principles calculations to perform a thorough theoretical characterization of the various aspects of the recently synthesized boroxine B_3O_3 monolayer. In agreement with previous theoretical investigations by Lin *et al.*,¹⁶ we found that the structure is perfectly planar formed by B_3O_3 hexagons with hexagonal symmetry, joined by a strong covalent bond between BB as reveal by ELF. The six B_3O_3 units joined with six-fold symmetry exhibiting a large hole with a surface of 23 \AA^2 , which is ideal for the adsorption of potassium and the hydroxium ion. In effect, we found that boroxine preferentially binds potassium over sodium. In contrast with previous works, we found that the band gap is direct. In effect, HSE calculations indicate that it is an insulator with a direct electronic band gap equal to 5.3 eV. Based on the cohesive energy calculations, B_3O_3 monolayer is comparably strong as that of graphene and remains negligibly distorted even at 600 K as from AIMD simulations. The elastic modulus is higher than that of Be_3N_2 , silicene, and germanene. Besides, we also explore the bilayer and bulk form of B_3O_3 and find that those phases are equally robust in comparison to bilayer graphene and graphite. Furthermore, the exfoliation energy calculations suggest that

the exfoliation of monolayer B_3O_3 from the bulk phase is quite easier than that of graphene-graphite. In addition, we find a great tenability of the optoelectronic properties in the presence of external electric field. That being said, the optical properties can be tuned from deep UV to the visible range thus offering a vast range of applications in optoelectronics.

Conflicts of interest

The authors declare no conflicts of interest.

References

- 1 K. S. Novoselov, A. K. Geim, S. V. Morozov, D. Jiang, Y. Zhang, S. V. Dubonos, I. V. Grigorieva and A. A. Firsov, *Science*, 2004, **306**, 666–669.
- 2 Z. Zeng, Z. Yin, X. Huang, H. Li, Q. He, G. Lu, F. Boey and H. Zhang, *Angew. Chem., Int. Ed.*, 2011, **50**, 11093–11097.
- 3 G. Ye, Y. Gong, J. Lin, B. Li, Y. He, S. T. Pantelides, W. Zhou, R. Vajtai and P. M. Ajayan, *Nano Lett.*, 2016, **16**, 1097–1103.
- 4 Y. Gogosti and B. Anasori, *ACS Nano*, 2019, **13**, 8491–8494.
- 5 P. Vogt, P. De Padova, C. Quaresima, J. Avila, E. Frantzeskakis, M. C. Asensio, A. Resta, B. Ealet and G. Le Lay, *Phys. Rev. Lett.*, 2012, **108**, 155501.
- 6 P. A. Denis, *Phys. Chem. Chem. Phys.*, 2015, **17**, 5393–5402.
- 7 M. E. Davila, L. Xian, S. Cahangirov, A. Rubio and G. Le Lay, *New J. Phys.*, 2014, **16**, 095002.
- 8 J. Ji, X. Song, J. Liu, Z. Yan, C. Huo, S. Zhang, M. Su, L. Liao, W. Wang, Z. Ni, Y. Hao and H. Zeng, *Nat. Commun.*, 2016, **7**, 13352.
- 9 F. Reis, G. Li, L. Dudy, M. Bauernfeind, S. Glass, W. Hanke, R. Thomale, J. Schäfer and R. Claessen, *Science*, 2017, **357**, 287–290.



10 H. Liu, A. T. Neal, Z. Zhu, D. Tomanek and P. D. Ye, *ACS Nano*, 2014, **8**, 4033.

11 Z. Zhu and D. Tomanek, *Phys. Rev. Lett.*, 2012, **112**, 176802.

12 M. Zhang, H. Sun, X. Wang, H. Du, J. He, Y. Long, Y. Zhang and C. Huang, *J. Phys. Chem. C*, 2019, **123**, 5010–5016.

13 J. Wang, F. Ma and M. Sun, *RSC Adv.*, 2017, **7**, 16801–16822.

14 P. A. Denis and F. Iribarne, *J. Phys. Chem. C*, 2018, **122**, 18583–18587.

15 M. Stredansky, A. Sala, T. Fontanot, R. Costantini, C. Africh, G. Comelli, L. Floreano, A. Morgante and A. Cossaro, *Chem. Commun.*, 2018, **54**, 3971.

16 S. Lin, J. Gu, H. Zhang, Y. Wang and Z. Chen, *FlatChem*, 2018, **9**, 27–32.

17 M. Dion, H. Rydberg, E. Schroder, D. C. Langreth and B. I. Lundqvist, *Phys. Rev. Lett.*, 2004, **92**, 246401.

18 Y. Zhao and D. G. Truhlar, *J. Chem. Phys.*, 2006, **125**, 194101.

19 J. Heyd and G. E. Scuseria, *J. Chem. Phys.*, 2004, **120**, 7274–7280.

20 J. M. Soler, E. Artacho, J. D. Gale, A. Garcia, J. Junquera, P. Ordejon and S.-D. Portal, The SIESTA Method for *Ab Initio* Order-N Materials Simulation, *J. Phys.: Condens. Matter*, 2002, **14**, 2745–2779.

21 P. Ordejon, E. Artacho and J. M. Soler, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **53**, R10441–R10444.

22 N. Troullier and J. L. Martins, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1991, **43**, 1993–2006.

23 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision D.01*, Gaussian, Inc., Wallingford CT, 2009.

24 W. Hehre, L. Radom, P. v. R. Schleyer, and J. A. Pople, *Ab initio Molecular Orbital Theory*, Wiley, New York, 1986.

25 S. Grimme, *J. Comput. Chem.*, 2006, **27**, 1787–1799.

26 S. Grimme, S. Ehrlich and L. Goerigk, *J. Comput. Chem.*, 2011, **32**, 1456–1465.

27 G. Kresse and J. Hafner, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **49**, 14251–14269.

28 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.

29 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1997, **78**, 1396.

30 P. E. Blöchl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953–17979.

31 G. Kresse and J. Furthmüller, *Comput. Mater. Sci.*, 1996, **6**, 15.

32 S. Ullah, P. A. Denis, R. B. Capaz and F. Sato, *New J. Chem.*, 2019, **43**, 2933–2941.

33 C. A. Coulson and T. W. Dingle, *Acta Crystallogr., Sect. B: Struct. Crystallogr. Cryst. Chem.*, 1968, **24**, 153.

34 A. Savin, R. Nesper, S. Wengert and T. F. Fässler, *Angew. Chem., Int. Ed.*, 1997, **36**, 1808–1832.

35 D. Toffoli, M. Stredansky, Z. Feng, G. Balducci, S. Furlan, M. Stener, H. Ustunel, D. Cvetko, G. Kladnik, A. Morgante, A. Verdini, C. Dri, G. Comelli, G. Fronzoni and A. Cossaro, *Chem. Sci.*, 2017, **8**, 3789–3798.

36 J. Tersoff and D. R. Hamann, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1985, **31**, 805–813.

37 C. Buttay, D. Planson, B. Allard, D. Bergogne, P. Bevilacqua, C. Joubert, M. Lazar, C. Martin, H. Morel, D. Tournier and C. Raynaud, *J. Mater. Sci. Eng. B*, 2011, **176**, 283–288.

38 R. C. Andrew, R. E. Mapasha, A. M. Ukpong and N. Chetty, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2012, **85**, 125428.

39 L. Liu and A. Corma, *Chem. Rev.*, 2018, **118**, 4981–5079.

40 J. Mahmood, F. Li, C. Kim, H.-J. Choi, O. Gwon, S.-M. Jung, J.-M. Seo, S.-J. Cho, Y.-W. Ju, H. Y. Jeong, G. Kim and J.-B. Baek, *Nano Energy*, 2018, **44**, 304–310.

41 S. Ullah, P. A. Denis, M. G. Menezes and F. Sato, *Appl. Surf. Sci.*, 2019, **493**, 308–319.

42 S. Gélinas, O. Paré-Labrosse, C.-N. Brosseau, S. Albert-Seifried, C. R. McNeill, K. R. Kirov, I. A. Howard, R. Leonelli, R. H. Friend and C. Silva, *J. Phys. Chem. C*, 2011, **115**, 7114–7119.

43 S. Gélinas, A. Rao, A. Kumar, S. L. Smith, A. W. Chin, J. Clark, T. S. van der Poll, G. C. Bazan and R. H. Friend, *Science*, 2014, **343**, 512.

44 M. M. Furchi, A. Pospischil, F. Libisch, J. Burgdörfer and T. Mueller, *Nano Lett.*, 2014, **14**, 4785–4791.

