

Cite this: *RSC Adv.*, 2019, 9, 34102

# *Tinospora cordifolia* derived biomass functionalized ZnO particles for effective removal of lead(II), iron(III), phosphate and arsenic(III) from water†

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Owing to the vast diversity in functional groups and cost effectiveness, biomass can be used for various applications. In the present study, biomass from *Tinospora cordifolia* (*TnC*) was prepared and grafted onto the surface of ZnO particles following a simple method. The *TnC* functionalized ZnO particles (*ZnO@TnC*) were characterized and exhibited excellent adsorption properties towards  $\text{Pb}^{2+}$  (506 mg  $\text{g}^{-1}$ ),  $\text{Fe}^{3+}$  (358 mg  $\text{g}^{-1}$ ) and  $\text{PO}_4^{3-}$  (1606 mg  $\text{g}^{-1}$ ) and the  $\text{Fe}^{3+}$  adsorbed *ZnO@TnC* adsorbs  $\text{AsO}_2^{1-}$  (189 mg  $\text{g}^{-1}$ ); the metal ions and anions were analyzed by ICP and IC. For reuse of *ZnO@TnC*, a desorption study was successfully carried out using NaOH and EDTA for  $\text{PO}_4^{3-}$  and  $\text{Pb}^{2+}$ , respectively;  $\text{Fe}^{3+}$  was further used for adsorption of  $\text{As(III)}$ . The adsorption fits well with the Langmuir adsorption isotherm model and the adsorption kinetic data are best fitted with a pseudo-second-order equation. The system developed may be useful for treatment of waste water and industrial effluents.

Received 3rd September 2019

Accepted 15th October 2019

DOI: 10.1039/c9ra07042g

rsc.li/rsc-advances

## 1. Introduction

The unavailability of safe drinking water is one of the major threats to all living species due to the globally polluted environment.<sup>1,2</sup> Toxic heavy metal ion and anion disposal through industrial, mining and agriculture processes is the major reason for the increasing pollution in air, soil and water.<sup>3,4</sup> Heavy metal ions such as  $\text{Pb}^{2+}$ ,  $\text{Hg}^{2+}$  and  $\text{AsO}_2^{1-}$  are highly toxic to humans and many other living species even at a very low concentration.<sup>5-7</sup> Lead-acid batteries are highly used for various purposes including in motor vehicles and a recent study suggested that 82% of the global lead consumption is due to lead batteries, which is a major source of lead pollution in the environment.<sup>8,9</sup> Consumption of  $\text{Pb}^{2+}$  can damage kidneys and severely affect the neurological system.<sup>10,11</sup> Arsenic is another highly toxic ion, the anthropogenic source of its pollution is the metal smelting and coal combustion.<sup>12</sup> Other reasons for arsenic pollution are wood preservation, waste incineration and volcanic emission.<sup>13</sup> Its consumption by human being causes heart disease and cancer of lung, liver, skin and kidney.<sup>14</sup> Iron is an essential metal ion for human body, however above a certain concentration it becomes harmful. It catalyzes formation of reactive oxygen species (ROS), which used to damage nucleic

acid and lipid causing Huntington's and Parkinson's diseases.<sup>15</sup> Iron contamination in water is mainly due to its presence in earth crust and other than this natural source, industries of mining and steel are significantly contributing to iron contamination in environment.<sup>4,16</sup> In human body, many intracellular pathways use phosphate ion for important cellular reactions, therefore, control of phosphate is one of the most critical biological regulations. Improper phosphate balance can affect the functionality of almost every human system and can cause kidney disorder, bone problems, weakness and fatigue.<sup>17</sup> Phosphate is key ingredient in fertilizers and above significant concentration can cause algal growth which have significant health and economic hazards. One main example is *Microcystis* algae, which produces microcystin that is lethal to both human as well as aquatic life.<sup>18</sup>

For removal of the above mentioned metal ions and anions, different techniques used are reverse osmosis, ion exchange, solvent extraction, electrochemical methods *etc.*<sup>19-22</sup> ZnO is a biocompatible material and because of its unique physical and chemical properties, it has been used for multifunctional purposes. In the area of separation science, it has been used as an adsorbent for removal of hazardous compound such as  $\text{H}_2\text{S}$ ,<sup>23</sup> heavy metal ions *etc.*<sup>24</sup> ZnO particles are available with various size and shape, its surface can also be modified with different coating agent to make it functional material and it has been used for removal of metal ions such as  $\text{Cu}^{2+}$ ,  $\text{Pb}^{2+}$  *etc.*<sup>25,26</sup> By proper choice of surface modifying agents, the performance of ZnO can be enhanced in terms of adsorption capacity and selectivity towards specific ions/molecules.

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9ra07042g



With the aim to develop cost effective materials for removal of toxic metal ions and anions from waste/polluted water for its possible recycle, ZnO particles were prepared and the surface of the particles were decorated with the biomass of *Tinospora cordifolia* (ZnO@TnC). *Tinospora cordifolia* (TnC) is a shrub from the Menispermaceae family, mainly available in the tropical areas, it contains various alkaloids, terpenoids and phenolic compounds having functional groups such as hydroxyl (-OH), carbonyl (-C=O) etc., which can effectively interact with metal ions.<sup>27</sup> This functional material (ZnO@TnC) exhibited high adsorption capacity towards the metal ions such as Pb<sup>2+</sup>, Fe<sup>3+</sup>, AsO<sub>2</sub><sup>1-</sup> and anion like PO<sub>4</sub><sup>3-</sup> in aqueous media. This material can be reused for several time in the case of lead and phosphate, while in the case of iron, it was reused for adsorption of arsenic from waste water. Herein we report preparation of ZnO based biomass decorated functional material, its characterization, metal ion adsorption property and its reusability.

## 2. Material and methods

### 2.1. Materials

Zn(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O was purchased from Alfa Aesar Chemical Company. NH<sub>2</sub>-NH<sub>2</sub>·H<sub>2</sub>O was purchased from TCI chemical company. Lead chloride (PbCl<sub>2</sub>), potassium phosphate monobasic (KH<sub>2</sub>PO<sub>4</sub>), sodium(*meta*)arsenite (AsNaO<sub>2</sub>), iron chloride (FeCl<sub>3</sub>) were purchased from Sigma Aldrich. *Tinospora cordifolia* (TnC) was collected from Bhavnagar (GPS location: lat. - 21.746746 and long. - 72.14018), Gujarat, India. Synthesis and other study was carried out in double distilled water. All chemicals were used without any further purification.

### 2.2. Instrumentation

Powder XRD was recorded on a PANalytical Empyrean (PIXcel 3D detector) instrument using Cu K $\alpha$  ( $\lambda = 0.15406$  nm) radiation, operated at 40 kV in the  $2\theta$  range of 5° to 80°. FT-IR spectra were recorded on a PerkinElmer instrument, model Spectrum GX, as KBr pellet. Scanning Electron Microscopy (SEM) images were recorded on a JEOL, model JSM 7100F FE-SEM instrument. EDX data was recorded on the Oxford Instrument X-Max<sup>np</sup> model. Metal ion analysis was carried out on an ICP-OES instrument from PerkinElmer, model Optima 2000. Ion chromatographic analysis was carried out on a Dionex ICS-5000 + DC model of Thermo Fischer for anions.

### 2.3. Synthesis of Zn(OH)<sub>2</sub> particles

Zn(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O (2.3 g, 20 mM) was dissolved in 400 mL of double distilled water. In another 400 mL of water, hydrazine hydrate (1.5 mL, 80 mM) was dissolved. Solution of hydrazine hydrate was then added into the solution of Zn(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O and the reaction mixture was stirred for 4 h at room temperature. The solution was clear at the beginning; however, formation of white precipitates was observed after 10–15 minutes. Solution was allowed to settle for 1 h after completion of stirring for 4 h and then the clear water from the top was decanted. The remaining solution with solid was then centrifuged at 8000 rpm for 15 minutes. The white precipitate thus separated was

isolated, washed with double distilled water (100 mL) for 3 times to remove any excess or unreacted reactants. It was then allowed to dry under vacuum for overnight. The white compound (Zn(OH)<sub>2</sub>) was characterized by Powder-XRD analysis.

### 2.4. Preparation of *Tinospora cordifolia* (TnC) powder

Leaves and skin of TnC were removed manually. Then it was chopped into small pieces and dried at 65 °C for 6 hours. The dried material was crushed by mortar and pestle. The resultant powder was then washed with 0.1 N HCL (100 mL) to remove any existing metal ion and impurities. After acid wash, powder was washed with plenty of water. The powder was characterized by IR and Powder-XRD analysis and was further used for surface modification of ZnO particles.

### 2.5. Preparation of ZnO@TnC particles

TnC powder (4 g) and Zn(OH)<sub>2</sub> (1 g), prepared as described above, are mixed well in a crucible and the reaction mixture was heated at 400 °C for 4 h in a muffle furnace. Zn(OH)<sub>2</sub> particles at high temperature was converted to ZnO particles due to calcination. After completion of heating, the dark gray colored crystalline mass was obtained, which was then allowed to cool to room temperature, which yielded microcrystals of ZnO@TnC. The compound was then characterized on the basis of IR, Powder-XRD and SEM analysis and was used for removal of metal ions.

### 2.6. Adsorption study with metal ions and anions

Adsorption of various metal ions such as Li<sup>+</sup>, Na<sup>+</sup>, K<sup>+</sup>, Ca<sup>2+</sup>, Mg<sup>2+</sup>, Mn<sup>2+</sup>, Fe<sup>3+</sup>, Co<sup>2+</sup>, Ni<sup>2+</sup>, Cu<sup>2+</sup>, Hg<sup>2+</sup>, Pb<sup>2+</sup>, Cd<sup>2+</sup> and anions such as AsO<sub>2</sub><sup>1-</sup>, Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup>, CN<sup>-</sup>, F<sup>-</sup>, PO<sub>4</sub><sup>3-</sup> were investigated using ZnO@TnC material. For this purpose, stock solutions of 10 ppm concentration of each metal ion from its respective chloride salt and of each anion from its respective sodium salt were prepared in 100 mL volumetric flask. In a typical experimental procedure, 10 mg of ZnO@TnC was stirred with 100 mL of stock solution of each metal ion and anion at 750 rpm for 6 h. Concentration of metal ions and anions were investigated in samples collected before and after adsorption by inductively coupled plasma (ICP) and ion chromatography (IC).

### 2.7. Metal ion removal study

The metal ion removal study was carried out at room temperature and at normal pH of Mili-Q water (6.8 ± 0.2) unless otherwise stated. Stock solution of 10, 50 and 100 ppm concentration of Pb<sup>2+</sup>, PO<sub>4</sub><sup>3-</sup> and Fe<sup>3+</sup> was prepared by dissolving required amount of PbCl<sub>2</sub>, KH<sub>2</sub>PO<sub>4</sub> and FeCl<sub>3</sub> in double distilled water, respectively. In a typical experiment, 10 mg of ZnO@TnC was dispersed into 50 mL of individual stock solution of metal ions and anions of 10, 50 and 100 ppm concentration with stirring at 750 rpm for 6 h. Aqueous samples of 500  $\mu$ L were collected at regular time interval of 60 minutes by micro pipette and metal ion and anion concentration in the samples were analyzed by ICP and IC, respectively. The ion adsorption



capacity of ZnO@TnC was then calculated using the following eqn (1).<sup>28</sup>

$$q_e = (C_0 - C_e) \frac{V}{m} \quad (1)$$

where  $q_e$  = quantity of metal ion/anion adsorbed on the adsorbent at the time of equilibrium ( $\text{mg g}^{-1}$ ),  $C_0$  = initial concentration of metal ion/anions in aqueous solution ( $\text{mg L}^{-1}$ ),  $C_e$  = final concentration of metal ion/anion in aqueous solution at the time  $t$ , ( $\text{mg L}^{-1}$ ),  $V$  = volume of the solution (L) and  $m$  = mass of adsorbents (g).

## 2.8. pH effect study

Adsorption of metal ions and anions by ZnO@TnC was studied at different pH in the range of 2 to 12. The solutions of that pH range were prepared by adjusting the pH using HCl and NaOH in double distilled water. Stock solution of  $\text{Pb}^{2+}$ ,  $\text{PO}_4^{3-}$  and  $\text{Fe}^{3+}$  were prepared of 100 ppm concentration at different pH. In a typical procedure, ZnO@TnC (10 mg) was treated with 50 mL of the stock solution of metal ions and anion with stirring at 750 rpm for 6 h. Aqueous samples (10 mL) were collected from the solutions after 6 h of stirring and analyzed by ICP and IC, as described above.

## 2.9. Competitive adsorption study

Adsorption capacity of ZnO@TnC towards  $\text{Pb}^{2+}$ ,  $\text{PO}_4^{3-}$  and  $\text{Fe}^{3+}$  in the same solution was also studied. A single stock solution containing all of these metal ions and anion was prepared with 100 ppm concentration of each ion in double distilled water. This stock solution (50 mL) containing mixture of ions was treated with 50 mg of ZnO@TnC following the procedure described in the Section 2.8. From the reaction mixture, 10 mL of sample was collected after 6 h of stirring and analyzed by ICP and IC, following the method described above.

## 2.10. Reusability of ZnO@TnC

Possible recycling of ZnO@TnC was examined by desorption experiment for ions adsorbed into ZnO@TnC. Stock solutions of  $\text{Pb}^{2+}$ ,  $\text{PO}_4^{3-}$  and  $\text{Fe}^{3+}$  were prepared of 500 ppm concentration in double distilled water. Then 20 mg of ZnO@TnC for each metal ion and was added into the stock solution (20 mL) of  $\text{Pb}^{2+}$ ,  $\text{PO}_4^{3-}$  and  $\text{Fe}^{3+}$  separately, stirred for 6 h (750 rpm) in a conical flask, and then centrifuged at 8000 rpm for 15 minutes. Supernatant liquid was analyzed by ICP and IC to determine the amount of metal ions and anion not adsorbed and ZnO@TnC particles with adsorbed ions were collected and dried in vacuum for overnight. Desorption of  $\text{PO}_4^{3-}$  from ZnO@TnC material was carried out with the aid of NaOH. The dried ZnO@TnC containing  $\text{PO}_4^{3-}$  was stirred with 50 mL of  $1.0 \text{ mol L}^{-1}$  NaOH solution for 1 h at 750 rpm. The resulting solution after stirring was centrifuged for 15 minutes at 8000 rpm and supernatant was collected and analyzed with ion chromatography (IC). The solid ZnO@TnC mass settled after centrifuge was then treated with 0.1 N HCl, washed with water (50 mL) and then reused for the adsorption of phosphate ( $\text{PO}_4^{3-}$ ). In case of  $\text{Pb}^{2+}$ , desorption experiment was carried out with the aid of EDTA.  $\text{Pb}^{2+}$  adsorbed

ZnO@TnC particles were dispersed in 50 mL of  $1.0 \text{ mol L}^{-1}$  EDTA solution and stirred for 1 h at 750 rpm. The resultant solution was then centrifuged for 10 minutes at 8000 rpm and supernatant was analyzed by ICP for  $\text{Pb}^{2+}$ . The settled ZnO@TnC particles were then washed with water (50 mL) and reused for the adsorption of  $\text{Pb}^{2+}$  in aqueous media. Desorption of iron ( $\text{Fe}^{3+}$ ) from the iron adsorbed ZnO@TnC particles was not successful, however the same was further used for the adsorption of  $\text{AsO}_2^{1-}$  in aqueous media.

## 2.11. Synthesis of Fe-ZnO@TnC and its use for removal of $\text{AsO}_2^{1-}$

$\text{Fe}^{3+}$  adsorbed ZnO@TnC particles were synthesized by stirring 100 mg of ZnO@TnC with 1000 ppm solution of  $\text{Fe}^{3+}$  at 750 rpm for 6 h followed by centrifugation at 7000 rpm for 10 minutes, the separated particles were then dried for overnight in vacuum. The  $\text{Fe}^{3+}$  adsorbed ZnO@TnC (Fe-ZnO@TnC) particles (10 mg) were then dispersed into solutions (50 mL) containing  $\text{AsO}_2^{1-}$  of 10, 50 and 100 ppm and stirred at 750 rpm for 6 h. An aliquot of 500  $\mu\text{L}$  from the solution was collected by micro pipette at a regular interval of 1 h and was analyzed by ion chromatography (IC) to determine the concentration of  $\text{AsO}_2^{1-}$  in the solution and thereby amount of  $\text{AsO}_2^{1-}$  adsorbed till the maximum adsorption was achieved. The eqn (1), as written above, was then used to calculate  $\text{AsO}_2^{1-}$  adsorption capacity of Fe-ZnO@TnC material.

## 2.12. $\text{AsO}_2^{1-}$ adsorption at different pH

Adsorption of  $\text{AsO}_2^{1-}$  by Fe-ZnO@TnC was studied at different pH in the range 2 to 12. Solutions of  $\text{AsO}_2^{1-}$  of 100 ppm was prepared with different pH in the range 2 to 12. The solution of  $\text{AsO}_2^{1-}$  (50 mL) of a particular pH was stirred with Fe-ZnO@TnC (10 mg) at 750 rpm for 6 h, then 10 mL of the solution was taken and analyzed for  $\text{AsO}_2^{1-}$  by ion chromatography (IC).

## 2.13. Adsorption isotherm study

Langmuir and Freundlich adsorption isotherm of ZnO@TnC and Fe-ZnO@TnC were studied. For this purpose, solutions of  $\text{Pb}^{2+}$ ,  $\text{Fe}^{3+}$ ,  $\text{AsO}_2^{1-}$  and  $\text{PO}_4^{3-}$  of concentration ranging from 10 to 100 ppm were prepared from the respective salts in double distilled water. ZnO@TnC and Fe-ZnO@TnC of 10 mg each were added into the solution of metal ion (50 mL) separately and was stirred for 6 h at 750 rpm. A sample of 10 mL was then collected from the mixture and analyzed for the metal ion/anion by ICP/IC to calculate adsorbed amount of metal ion/anion. The results obtained were then fitted in the equation of Langmuir and Freundlich adsorption isotherm.

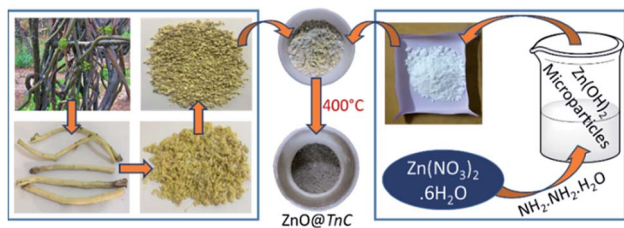
# 3. Results and discussion

## 3.1. Preparation and characterization of ZnO@TnC particles

The preparation of ZnO@TnC is described in the Experimental section and a schematic presentation is shown in Scheme 1.

In the calcination process at  $400 \text{ }^\circ\text{C}$ ,  $\text{Zn}(\text{OH})_2$  was converted to ZnO particle, surface of which was modified by TnC forming





Scheme 1 Schematic presentation of the stepwise synthesis of ZnO@TnC particles.

ZnO@TnC as crystalline material. This new material was characterized on the basis of powder-XRD, IR, SEM and EDX analysis. The Powder-X-ray diffractogram of the ZnO@TnC along with TnC and Zn(OH)<sub>2</sub> are shown in Fig. 1, which clearly shows that the diffractogram of ZnO@TnC is distinctly different from that of Zn(OH)<sub>2</sub> and TnC and there is no unreacted Zn(OH)<sub>2</sub> in the TnC containing new material.

The characteristic  $2\theta$  peaks for Zn(OH)<sub>2</sub> observed at 20.26°, 20.99, 25.13°, 27.29°, 27.87°, 32.94°, 39.60°, 40.86° and 42.20°, which are correspond to the orthorhombic zinc hydroxide (Fig. 1, JCPDS No: 01-089-0138)<sup>29</sup> are disappeared in the diffractogram of ZnO@TnC. The new  $2\theta$  peaks noted at 31.83°, 34.70°, 36.46°, 47.73°, 56.75°, 63.01°, 66.54°, 68.12° and 69.27° are correspond to ZnO (Fig. 1, JCPDS file no: 01-089-1397),<sup>29</sup> that confirmed the conversion of Zn(OH)<sub>2</sub> to ZnO. The  $2\theta$  peaks observed for TnC at 23.18°, 29.50, 36.10°, 39.53°, 43.28°, 47.58°, 48.61°, 57.51° and 60.79° (Fig. 1) are almost disappeared or shifted in the diffractogram observed for ZnO@TnC, indicating that the added TnC is consumed for functionalization of the surface of the ZnO particles and almost nothing is left free contaminating the functionalized composite material (ZnO@TnC). The IR spectra of TnC and that of ZnO@TnC are displayed in Fig. S1 (ESI),<sup>†</sup> the TnC exhibits IR bands at 3450, 1633, 1435 and 1076 cm<sup>-1</sup>, which are due to -OH, carbonyl group, carboxyl stretching and asymmetric stretching of carboxyl group.<sup>30</sup> The IR spectrum of ZnO@TnC exhibits similar bands with slight shift in position at 3425, 1627, 1457 and 1093 cm<sup>-1</sup>; with decrease in the intensity of bands at 1457 and

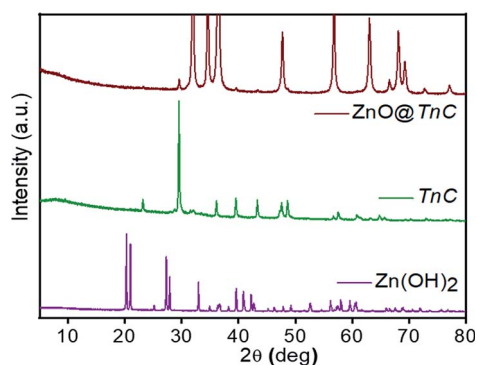


Fig. 1 Powder X-ray diffractograms of Zn(OH)<sub>2</sub> (purple color), biomass of *Tinospora cordifolia* (green color) and ZnO@TnC particles (brown color).

1093 cm<sup>-1</sup> confirming functionalization of TnC over ZnO particles. The bands were shifted due to interaction of the functional groups with the ZnO particles. The IR data therefore, suggest the presence of TnC in the crystals of ZnO@TnC and probably anchoring through the surface of the particles took place through the carboxylic groups, as its band intensity and position affected significantly due to surface modification by TnC. The crystallite size of ZnO and ZnO@TnC was calculated following the literature procedure<sup>31</sup> and the results showed that the crystallite size of ZnO before and after functionalization are 236.022 and 295.021 nm, respectively, it is increased significantly confirming the functionalization. The modification of the surface of ZnO is also evident from SEM images and EDX analysis. The FE-SEM images of ZnO, before and after coating with TnC, and EDX analysis of the ZnO@TnC are shown in Fig. 2. ZnO for SEM image was prepared from Zn(OH)<sub>2</sub> by calcination in a muffle furnace by heating at 400 °C for 4 h, the XRD data of which was matched with literature values (JCPDS No: 01-089-1397).

The changes on the surface of the particles after anchoring of TnC is clearly seen in Fig. 2C comparing it with the surface without modification (Fig. 2A). The EDX analysis of ZnO and that of ZnO@TnC, shown in Fig. 2B and D, show the presence of carbon, nitrogen and high percentage of oxygen in ZnO@TnC compared to ZnO, confirming the presence of TnC in ZnO@TnC.

### 3.2. Adsorption property of ZnO@TnC towards metal ions and anions

Adsorption property of ZnO@TnC was investigated towards the metal ions Li<sup>+</sup>, Na<sup>+</sup>, K<sup>+</sup>, Mg<sup>2+</sup>, Mn<sup>2+</sup>, Fe<sup>3+</sup>, Co<sup>2+</sup>, Ni<sup>2+</sup>, Cu<sup>2+</sup>, Hg<sup>2+</sup>, Pb<sup>2+</sup>, Cd<sup>2+</sup> and anions AsO<sub>2</sub><sup>1-</sup>, Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup>, CN<sup>-</sup>, F<sup>-</sup>, PO<sub>4</sub><sup>3-</sup> following the methods described in the Experimental section.

Bar diagram plotting the concentration of various metal ions and anions before and after adsorption is depicted in Fig. 3, which exhibited that the metal ions Fe<sup>3+</sup> and Pb<sup>2+</sup> and the anions PO<sub>4</sub><sup>3-</sup> are absorbed almost the entire amount used in the experiment. Further, the Fe<sup>3+</sup> adsorbed ZnO@TnC material

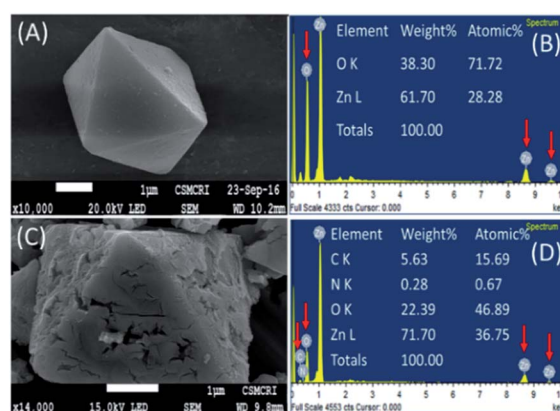


Fig. 2 SEM images of ZnO particles (A), ZnO@TnC particles (C) and their EDX spectra respectively (B and D). Enlarged SEM images with recording parameters is submitted as ESI (S9).<sup>†</sup>



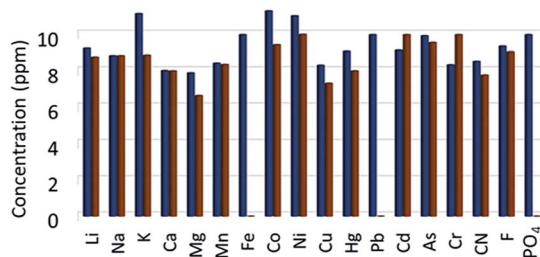


Fig. 3 Bar diagram showing concentration of different metal ions and anions before adsorption (blue bars) and after adsorption by ZnO@TnC particles (brown bars). (Oxidation state of metal ions are omitted for clarity; all metal ions are in "2+" oxidation state except iron (Fe) and Arsenic (As), which are in "3+" and chromium is in "6+" oxidation state).

(Fe-ZnO@TnC) was used for adsorption of  $\text{AsO}_2^{1-}$  from aqueous solution following the method described in the experimental section, and the concentration of  $\text{AsO}_2^{1-}$  measured before and after adsorption by IC is shown in Fig. S2 (ESI),† which clearly suggested substantial adsorption of  $\text{AsO}_2^{1-}$ . The presence of  $\text{Fe}^{3+}$ ,  $\text{Pb}^{2+}$  and  $\text{PO}_4^{3-}$  in ZnO@TnC and  $\text{AsO}_2^{1-}$  in Fe-ZnO@TnC were confirmed with the aid of PXRD, IR spectroscopy and SEM microscopy, as described below.

### 3.3. Characterization of adsorbed ions

The adsorption of metal ions and anions on to the surface of the ZnO@TnC was confirmed on the basis of PXRD, IR and SEM analysis of the isolated ion adsorbed species of ZnO@TnC. Each of these ion adsorbed species were obtained following the procedure described in the Experimental section. The powder X-ray diffractograms of the ion adsorbed ZnO@TnC materials are shown in Fig. 4.

It may be noted that there are substantial changes in the diffractograms compared to that of ZnO@TnC after adsorption of metal ions/anions. However, the peaks observed in pure ZnO@TnC are noted as it is in all the diffractograms, the additional peaks are due to  $\text{Pb}^{2+}/\text{Fe}^{3+}/\text{PO}_4^{3-}$  adsorbed species of ZnO@TnC. In the case of  $\text{AsO}_2^{1-}$ , probably after its interaction

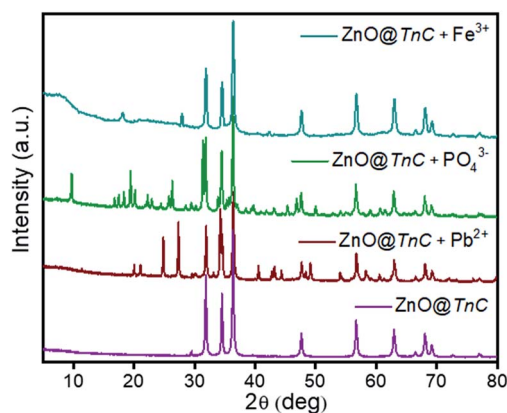


Fig. 4 XRD spectral analysis of ZnO@TnC (purple color) and that after adsorption of  $\text{Pb}^{2+}$  (brown color),  $\text{PO}_4^{3-}$  (green color) and  $\text{Fe}^{3+}$  (dark cyan color).

with Fe-ZnO@TnC, its crystallinity is lost and amorphous product is formed, as evident from the X-ray diffractogram, Fig. S3 (ESI).† The interaction of metal ion/anion with ZnO@TnC is also investigated by IR spectroscopy, shown in Fig. 5.

In Fig. 5, it may be noted that there is significant change in band position and band intensity in the region  $1700\text{--}500\text{ cm}^{-1}$  and around  $3500\text{ cm}^{-1}$  for all of the ion bound ZnO@TnC compared to the mother compound. Prominent changes are noted in the regions  $1700\text{--}1400\text{ cm}^{-1}$  and  $3400\text{--}3500\text{ cm}^{-1}$ , which are due to carbonyl/carboxylic acid groups and OH/NH groups, respectively, which suggest that the ions interacted with ZnO@TnC through the C=O/COOH and OH/NH functional groups of TnC embedded onto the surface of the ZnO particles. The IR spectrum of the  $\text{AsO}_2^{1-}$  bound Fe-ZnO@TnC, shown in Fig. S4 (ESI),† also showed significant change compared to its mother compound in the finger print region confirming the interaction of  $\text{AsO}_2^{1-}$  with Fe-ZnO@TnC. The morphology of the ion adsorbed ZnO@TnC is investigated by FE-SEM and the presence of ions ( $\text{Pb}^{2+}/\text{Fe}^{3+}/\text{PO}_4^{3-}$ ) in it is confirmed by EDX analysis. The FE-SEM images and the EDX analysis of all the ion adsorbed ZnO@TnC are shown in Fig. 6. It may be noted that on adsorption of ion, the morphology of ZnO@TnC has changed substantially and the EDX data show the presence of respective metal ion/element (P of phosphate ion). Therefore, PXRD, IR, FE-SEM and EDX analysis have confirmed the adsorption of ions by ZnO@TnC.

### 3.4. Removal of metal ions and anion and their quantification

The ZnO@TnC particles were used for adsorption of  $\text{Pb}^{2+}$ ,  $\text{Fe}^{3+}$  and  $\text{PO}_4^{3-}$  and Fe-ZnO@TnC was used for adsorption of  $\text{AsO}_2^{1-}$  in aqueous media at different concentration of metal ions/anions.

Details of the procedure is given in the Experimental section and the plot of amount of metal ion/anion adsorbed as a function of time at different concentrations are shown in Fig. 7. It is observed that in the case of  $\text{Pb}^{2+}$  and  $\text{Fe}^{3+}$ , saturation of

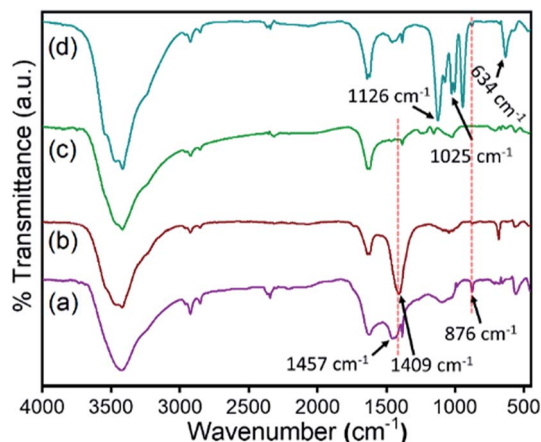


Fig. 5 IR spectra of the ZnO@TnC (a) and that after adsorption of  $\text{Pb}^{2+}$  (b),  $\text{Fe}^{3+}$  (c) and  $\text{PO}_4^{3-}$  (d).



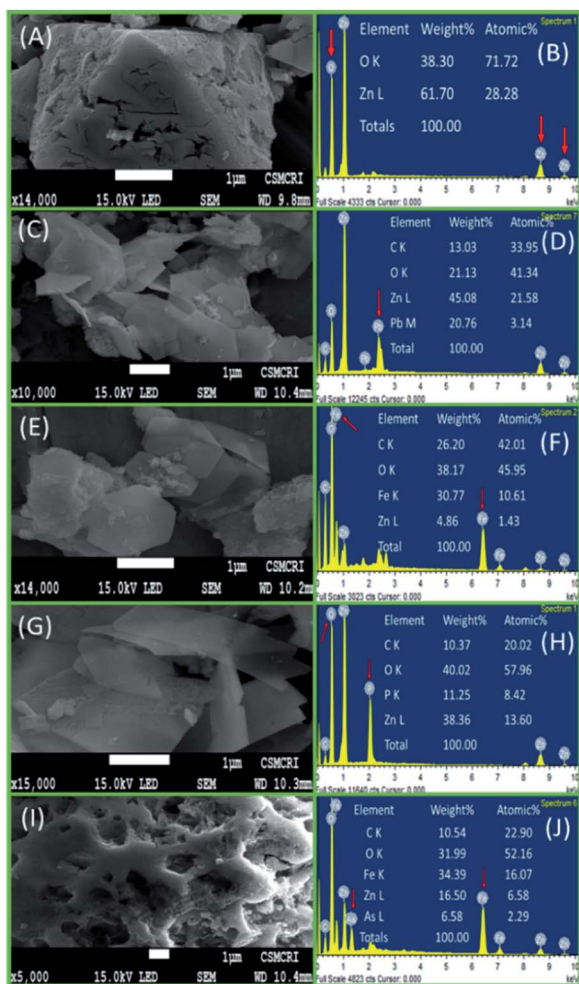


Fig. 6 SEM images along with EDX spectra of (A and B) ZnO@TnC particles after adsorption of (C and D) Pb<sup>2+</sup>, (E and F) Fe<sup>3+</sup>, (G and H) PO<sub>4</sub><sup>3-</sup> and (I and J) AsO<sub>2</sub><sup>1-</sup> adsorbed Fe-ZnO@TnC. Enlarged SEM images with recording parameters is submitted as ESI (S9).†

adsorption is reached by 180 minutes, while in the case of PO<sub>4</sub><sup>3-</sup> and AsO<sub>2</sub><sup>1-</sup>, saturation was observed after 240 and 360 minutes, respectively. The adsorption capacity of Pb<sup>2+</sup>, Fe<sup>3+</sup>, PO<sub>4</sub><sup>3-</sup>, calculated by using eqn (1), was found to be 506.1 mg g<sup>-1</sup>, 358 mg g<sup>-1</sup> and 1606 mg g<sup>-1</sup>, respectively and the same for AsO<sub>2</sub><sup>1-</sup> was to be 189 mg g<sup>-1</sup>. The adsorption capacities thus obtained were compared with that of a number of recently published systems, functionalized with various materials (Table 1) and it is noted that the data obtained by the present study are comparable and in many cases better than the reported values. Apparently, the removal capacity obtained for PO<sub>4</sub><sup>3-</sup> by present study is one of the highest among the reported values. The experimental data thus obtained was further used to determine kinetics and adsorption isotherm model.

### 3.5. Adsorption isotherm of ZnO@TnC

To evaluate the adsorption efficiency of ZnO@TnC towards Pb<sup>2+</sup>, Fe<sup>3+</sup>, PO<sub>4</sub><sup>3-</sup> and of Fe-ZnO@TnC towards AsO<sub>2</sub><sup>1-</sup>, the experimental data was fitted with classical Langmuir isotherm

and Freundlich isotherm models. According to the Langmuir model, most of the adsorption occurs on the surface of the adsorbent and migration of adsorbate molecules do not occur. Mathematically, Langmuir isotherm can be expressed by the following equation.<sup>53</sup>

Langmuir isotherm

$$q_e = \frac{Q_m K_L C_e}{1 + K_L C_e} \quad (2)$$

where “ $q_e$ ” is the quantity of the adsorbed adsorbate at the time of equilibrium, “ $C_e$ ” is the concentration of adsorbate remained at the time of equilibrium, “ $Q_m$ ” is the quantity of adsorbate adsorbed on the adsorbent at saturation point and “ $K_L$ ” is Langmuir constant. The above Langmuir isotherm can be rewritten in the following way and the plot  $1/q_e$  Vs  $1/C_e$  is expected to be linear.

$$\frac{1}{q_e} = \frac{1}{Q_m} + \frac{1}{Q_m K_L} \frac{1}{C_e} \quad (3)$$

An essential characteristic of the Langmuir adsorption isotherm can be expressed by dimensionless factor, also called equilibrium parameter ( $R_L$ ), which can be calculated using the following eqn (4).<sup>54</sup>

$$R_L = \frac{1}{1 + K_L C_0} \quad (4)$$

where  $K_L$  is the Langmuir constant and  $C_0$  is the initial concentration of the analyte. If the  $R_L$  values is  $0 < 1$  then the Langmuir adsorption is favorable and if  $R_L > 1$ , then the adsorption is unfavorable.

Freundlich adsorption isotherm is mainly describe multi-layer adsorption in liquid to solid phase. Freundlich isotherm can be expressed in the following way.<sup>53</sup>

Freundlich isotherm

$$q_e = k_f C_e^{1/n} \quad (5)$$

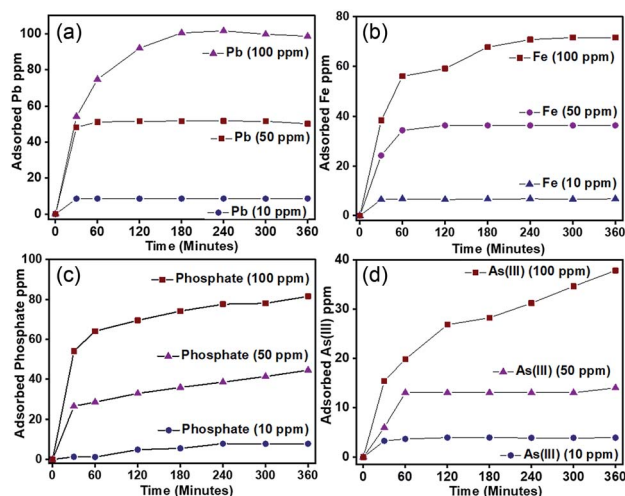


Fig. 7 Kinetic study of ions (a) Pb<sup>2+</sup>, (b) Fe<sup>3+</sup>, (c) PO<sub>4</sub><sup>3-</sup> and (d) AsO<sub>2</sub><sup>1-</sup> at their initial concentration of 10, 50 and 100 ppm with time.



Table 1 Comparison of different previously reported tool and their adsorption capacity towards Pb<sup>2+</sup>, Fe<sup>3+</sup>, PO<sub>4</sub><sup>3-</sup> and AsO<sub>2</sub><sup>1-</sup> in aqueous media

Material used	Adsorption capacity (mg g <sup>-1</sup> )				Ref.
	Pb <sup>2+</sup>	Fe <sup>3+</sup>	AsO <sub>2</sub> <sup>1-</sup>	PO <sub>4</sub> <sup>3-</sup>	
Saraca indica leaf powder	1.19	—	—	—	32
Titanate/Fe <sub>3</sub> O <sub>4</sub> nano	382.3	—	—	—	33
Poly(ethyleneimine) nano rod	240	—	—	—	34
Nano vaterite poly(ethyleneimine)	2762	—	—	—	35
MFe <sub>2</sub> O <sub>4</sub> (M = Mn, Co)–MoS <sub>2</sub> –carbon dot	588.2	—	—	—	36
ZnO/MMT nanocomposite	88.50	—	—	—	37
Fly ash	—	—	—	42.5	38
Magnetic cellulose	—	—	—	22.2	39
Biomass based Zr(OH) <sub>4</sub>	—	—	—	31.9	40
Biochar (Mg-enriched tomato tissues)	—	—	—	>100	41
Mesoporous silica	—	—	—	43.3	42
Ferric sludge	—	—	—	30	43
Natural zeolite	—	1.15	—	—	44
Natural smectite clay	—	12.8	—	—	45
Vegetable biomass	—	3.5	—	—	46
Activated carbon	—	3.60	—	—	47
Nano zero-valent iron	—	—	3.5	—	48
Fe/Cu nanoparticles	—	—	19.6	—	49
Magnetic Fe <sub>3</sub> O <sub>4</sub> nano	—	—	19.6	—	50
(Sn <sub>0.95</sub> Fe <sub>0.05</sub> O <sub>2-δ</sub> )@GO	—	—	105	—	51
Nano cellulose	—	—	51	—	52
ZnO@TnC particles	506.1	358	189	1606	This study

where “ $q_e$ ” and “ $C_e$ ” terms are same as described in Langmuir equation, “ $K_f$ ” is the Freundlich constant and “ $n$ ” is the Freundlich exponent which is related to the heterogeneity of the surface of adsorbent. The simplified form of the above equation can be rewritten as below.

$$\log q_e = \log K_f + \frac{1}{n} \log C_e \quad (6)$$

The experimental adsorption data were fitted in the eqn (3) and (5) and the plot of Langmuir and Freundlich adsorption isotherms for metal ions and anion are shown in Fig. 8 and S5

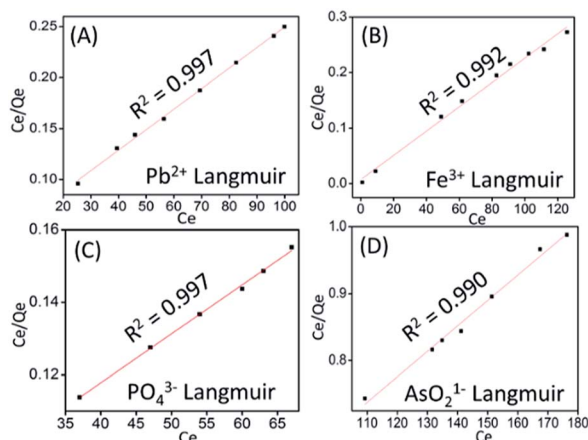


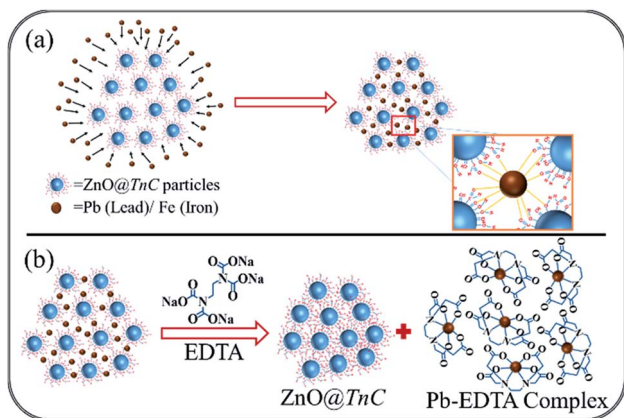
Fig. 8 Plot of  $C_e/Q_e$  Vs  $C_e$  representing linearized form of Langmuir adsorption isotherm of (A) Pb<sup>2+</sup> (B), Fe<sup>3+</sup> (C) PO<sub>4</sub><sup>3-</sup> and (D) AsO<sub>2</sub><sup>1-</sup>.

(ESI),<sup>†</sup> respectively. The dimensionless factor ( $R_L$ ) is calculated using the eqn (4) and all the calculated parameters and the correlation coefficient ( $R^2$ ) values obtained from the plot are summarized in Table 2. The  $R_L$  values obtained is  $0 < 1$ , which suggests that the Langmuir adsorption is favorable. Schematic presentation of monolayer adsorption of Pb<sup>2+</sup> and Fe<sup>3+</sup> is shown in the Scheme 2a. Higher value of Freundlich adsorption exponent “ $n$ ”, which is equal to or greater than 1, suggests that adsorption is favorable on surface of the particles.<sup>55</sup> It may be noted that the adsorption isotherm study with different ions were carried out with the solutions of concentrations 10, 50 and 100 ppm and the adsorption capacity thus obtained has given in Table 1. However, it was noted that in case of Pb<sup>2+</sup> and PO<sub>4</sub><sup>3-</sup>, the saturation point was not achieved with 100 ppm solution, therefore to examine the maximum adsorption capacity, experiment was repeated with 1000 ppm solution for these two ions and the data obtained is presented in Table 2. To maintain

Table 2 Langmuir and Freundlich adsorption isotherm model parameters for Pb<sup>2+</sup>, Fe<sup>3+</sup>, PO<sub>4</sub><sup>3-</sup> and AsO<sub>2</sub><sup>1-</sup> adsorption by ZnO@TnC particles

Metal ions	Langmuir adsorption isotherm				Freundlich adsorption isotherm		
	$K_L$	$Q_m$	$R_L$	$R^2$	$K_F$	$n$	$R^2$
Pb <sup>2+</sup>	0.042	497.51	0.15	0.997	3.68	2.72	0.924
Fe <sup>3+</sup>	0.302	456.62	0.02	0.992	4.99	6.10	0.906
PO <sub>4</sub> <sup>3-</sup>	0.021	740.74	0.29	0.997	1.60	0.90	0.988
AsO <sub>2</sub> <sup>1-</sup>	0.011	262.46	0.31	0.990	0.06	0.33	0.977





Scheme 2 Schematic presentation of (a) monolayer adsorption of  $\text{Pb}^{2+}$  and  $\text{Fe}^{3+}$  by  $\text{ZnO@TnC}$  particles and (b) regeneration of  $\text{ZnO@TnC}$  particles by desorption of  $\text{Pb}^{2+}$  after treatment with EDTA.

uniformity, the adsorption capacity data for all the ions with 100 ppm solution has given in Table 1.

### 3.6. Adsorption kinetics

For insight of mechanistic aspects, kinetics of the adsorption process was also carried out, experimental procedure for which is described in the Experimental section. For this purpose, two semi empirical kinetic models (i) pseudo-first-order and (ii) pseudo-second-order, which are based on adsorption equilibrium capacity, is used.<sup>56</sup> The equation for pseudo first order reaction is shown below.

$$\frac{dQ_t}{dt} = k_1(Q_e - Q_t) \quad (7)$$

where  $Q_e$  and  $Q_t$  are the adsorbed amounts of metal ions/anions at equilibrium and time ' $t$ ', respectively;  $k_1$  is the pseudo-first-order kinetic constant, expressed in inverse minutes. Equation integration and rearrangement yields the linear form as shown below.

$$\ln(Q_e - Q_t) = \ln Q_e - k_1 t \quad (8)$$

The equation for pseudo second order reaction can be expressed mathematically as shown below.

$$\frac{dQ_t}{dt} = k_2(Q_e - Q_t)^2 \quad (9)$$

where  $Q_e$  and  $Q_t$  are same as in eqn (6);  $k_2$  ( $\text{g}(\text{mol min})^{-1}$ ) is the pseudo-second-order rate constant. The differential equation is usually integrated and transformed in its linear form as shown below.

$$\frac{t}{Q_t} = \frac{1}{k_2 Q_e^2} + \frac{t}{Q_e} \quad (10)$$

Experimental data obtained from the kinetic study following the procedure described in the Experimental section is fitted in eqn (7) and (9) and the plot obtained for pseudo-second-order

and pseudo-first-order are shown in Fig. 9 and S6 (ESI),<sup>†</sup> respectively.

The values of correlation co-efficient ( $R^2$ ) obtained for all metal ions and anions at different concentrations using both the equations are summarized in Table S1.<sup>†</sup> The  $R^2$  values suggests that the experimental data are better fitted with the pseudo-second-order equation. The rate constant and other parameters are therefore, calculated using the pseudo-second-order equation and the data are summarized in Table 3. It can be observed from Table 3 that, in case of  $\text{Pb}^{2+}$  and  $\text{Fe}^{3+}$ , initial adsorption ( $h$ ) increases with increase in metal ion concentration from 10 to 50 ppm and then decreases it at 100 ppm, which indicated the saturation of  $\text{ZnO@TnC}$  with metal ions. However, in case of  $\text{PO}_4^{3-}$  and  $\text{AsO}_2^{1-}$ , " $h$ " increases with increase in the concentration of ions from 10 to 100 ppm, which indicated availability of more binding site on the particles.<sup>57</sup> Rate constant of pseudo-second-order kinetic " $k_2$ " decreases with increase in initial concentration of ions. In all cases, the amount of metal ions adsorbed increases with increase in their initial concentration. Adsorption capacity of  $\text{Pb}^{2+}$ ,  $\text{Fe}^{3+}$  and  $\text{PO}_4^{3-}$  was found to be  $506.1 \text{ mg g}^{-1}$ ,  $358 \text{ mg g}^{-1}$  and  $1606 \text{ mg g}^{-1}$ , respectively by  $\text{ZnO@TnC}$  and of  $\text{AsO}_2^{1-}$  to be  $189 \text{ mg g}^{-1}$  by  $\text{Fe-ZnO@TnC}$ .

### 3.7. Adsorption with variation in pH

Adsorption of all ions were conducted in the pH range 2 to 12 following the procedure described in the Experimental section. For  $\text{PbCl}_2$ , however the study was conducted up to pH 9 due to insufficient solubility of  $\text{PbCl}_2$  at higher pH. The bar diagram showing the amount of ions adsorbed at different pH is displayed in Fig. S7 (ESI),<sup>†</sup> which shows that except  $\text{PO}_4^{3-}$ , adsorption of all other ions are independent of pH. In case of  $\text{PO}_4^{3-}$ , its adsorption was observed to be higher at acidic pH and lower at higher (basic) pH. This is probably due to the fact that  $\text{PO}_4^{3-}$  binds with  $\text{ZnO@TnC}$  through H-bonding and at higher pH the binding sites get deprotonated.

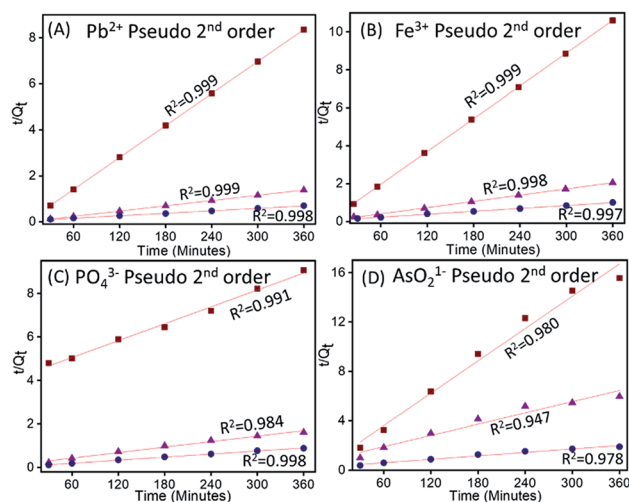


Fig. 9 Plot of pseudo-second-order kinetic study of (A)  $\text{Pb}^{2+}$  (B)  $\text{Fe}^{3+}$  (C)  $\text{PO}_4^{3-}$  and (D)  $\text{AsO}_2^{1-}$  adsorption by  $\text{ZnO@TnC}$ .





Table 3 Pseudo first and second order kinetic model parameters for  $\text{Pb}^{2+}$ ,  $\text{Fe}^{3+}$ ,  $\text{PO}_4^{3-}$  and  $\text{AsO}_2^{1-}$  adsorption by  $\text{ZnO@TnC}$ 

Initial conc. (PPM)	$k_2$ ( $\text{g mg}^{-1} \text{min}^{-1}$ )	$h$ ( $\text{mg g}^{-1} \text{min}^{-1}$ )	$Q_e$ ( $\text{mg g}^{-1}$ ) practical	$Q_e$ ( $\text{mg g}^{-1}$ ) calculated
<b><math>\text{Pb}^{2+}</math></b>				
10	$3.68 \times 10^{-2}$	68.57	43.11	43.15
50	$2.35 \times 10^{-3}$	160.72	260.10	261.09
100	$6.69 \times 10^{-5}$	20.89	512.25	558.65
<b><math>\text{Fe}^{3+}</math></b>				
10	$1.17 \times 10^{-2}$	13.75	33.97	34.19
50	$9.09 \times 10^{-4}$	29.32	175.25	179.53
100	$8.95 \times 10^{-5}$	13.65	358.35	390.62
<b><math>\text{PO}_4^{3-}</math></b>				
10	$1.23 \times 10^{-4}$	0.23	39.68	38.49
50	$1.12 \times 10^{-4}$	6.33	222.90	237.52
100	$1.10 \times 10^{-4}$	20.03	408.00	425.53
<b><math>\text{AsO}_2^{1-}</math></b>				
10	$3.67 \times 10^{-4}$	0.19	23.15	22.90
50	$8.73 \times 10^{-5}$	0.38	60.50	66.26
100	$2.45 \times 10^{-5}$	1.13	189.00	214.59

### 3.8. Recovery of $\text{ZnO@TnC}$ for reuse

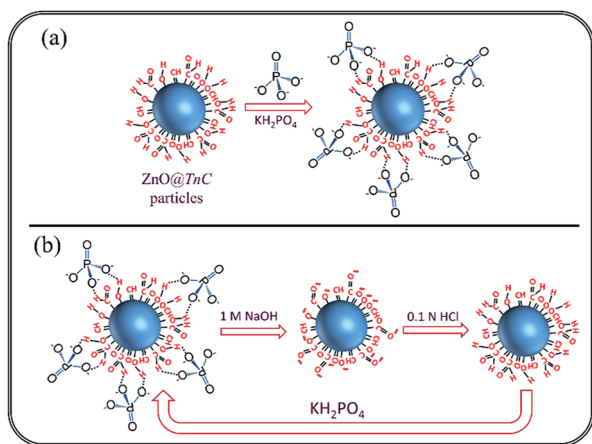
The possibility to recover the  $\text{ZnO@TnC}$  by desorption of adsorbed metal ion to reuse it for further adsorption study is investigated. Detail experimental procedure has given in the Experimental section. For desorption of  $\text{Pb}^{2+}$ , a strong chelating agent called ethylenediaminetetraacetate (EDTA) was added into  $\text{Pb}^{2+}$ - $\text{ZnO@TnC}$ , which resulted in formation of water soluble  $\text{Pb}^{2+}$ -EDTA complex, the suspended  $\text{ZnO@TnC}$  was then separated by centrifugation and after washing the solid by water it is reused for adsorption of  $\text{Pb}^{2+}$  up to 3 consecutive cycles as shown in the Fig. S8A (ESI).† The schematic presentation of desorption of  $\text{Pb}^{2+}$  and formation of  $\text{Pb}^{2+}$ -EDTA complex is depicted in Scheme 2b. As mentioned earlier,  $\text{PO}_4^{3-}$  at higher pH does not interact with  $\text{ZnO@TnC}$  due to

deprotonation of its binding sites, therefore  $\text{ZnO@TnC}$  after adsorption of  $\text{PO}_4^{3-}$  was treated with NaOH solution, which induced deprotonation of the functional group (Scheme 3), resulting in the release of  $\text{PO}_4^{3-}$  in the solution. After NaOH treatment, it was neutralized with HCl solution and then reused for further adsorption of  $\text{PO}_4^{3-}$ .

$\text{ZnO@TnC}$  was reused for  $\text{PO}_4^{3-}$  adsorption up to 3 consecutive cycle as shown in Fig. S8B (ESI).† The chelating agent EDTA, however was not effective for releasing  $\text{Fe}^{3+}$  from  $\text{Fe-ZnO@TnC}$ . However,  $\text{Fe-ZnO@TnC}$  was used further for the adsorption of  $\text{AsO}_2^{1-}$  in aqueous media. Attempt has been made for regeneration of  $\text{ZnO@TnC}$  from  $\text{AsO}_2^{1-}$  adsorbed  $\text{Fe-ZnO@TnC}$  materials by desorption of  $\text{AsO}_2^{1-}$  and  $\text{Fe}^{3+}$  by variation of pH and treatment with EDTA but it was not successful.

## 4. Conclusions

The present study demonstrated an easy and simple method for the preparation of a biomass grafted ZnO microcrystalline material as an adsorbent for selected metal ions and anions. The biomass powder was prepared from *Tinospora cordifolia* and it was anchored on to the surface of the ZnO particles by calcination of  $\text{Zn(OH)}_2$  in presence of the biomass. The surface functionalized ZnO particles ( $\text{ZnO@TnC}$ ) adsorbed  $\text{Pb}^{2+}$ ,  $\text{Fe}^{3+}$  and  $\text{PO}_4^{3-}$  effectively out of a large number of metal ions and anions examined and the  $\text{Fe}^{3+}$  adsorbed  $\text{ZnO@TnC}$  material further adsorbed  $\text{AsO}_2^{1-}$ . The amount of  $\text{PO}_4^{3-}$  absorbed,  $1606 \text{ mg g}^{-1}$ , is the highest so far reported to the best of our knowledge. The amount of  $\text{Pb}^{2+}$  ( $506.1 \text{ mg g}^{-1}$ ) and  $\text{Fe}^{3+}$  ( $358 \text{ mg}$ ) adsorbed is also very high and are comparable to some of the best systems reported recent time. For  $\text{Pb}^{2+}$  and  $\text{PO}_4^{3-}$ , the reusability of  $\text{ZnO@TnC}$  is demonstrated by desorption of the metal ion and anion for three cycle without deactivation. For  $\text{Fe}^{3+}$ , the metal ion adsorbed material was further used to



Scheme 3 Schematic presentation of (a) adsorption of  $\text{PO}_4^{3-}$  by  $\text{ZnO@TnC}$  and (b) regeneration of  $\text{ZnO@TnC}$  by desorption of  $\text{PO}_4^{3-}$  after treatment with  $1 \text{ mol L}^{-1}$  NaOH solution.



adsorb  $\text{AsO}_2^{1-}$ . The experimental adsorption data fitted excellently with the Langmuir adsorption isotherm model and the adsorption kinetic data follow the pseudo-second-order equation. All of the calculated parameters are reported and the possible mechanism of adsorption is presented. The material developed following a simple method using low cost starting materials and is very useful for removal of  $\text{Pb}^{2+}$ ,  $\text{Fe}^{3+}$ ,  $\text{PO}_4^{3-}$  and  $\text{AsO}_2^{1-}$  from the waste water and industrial effluents.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

Manuscript number: CSIR-CSMCRI – 102/2019. P.P. gratefully acknowledge financial assistance in the form of CSIR-Emeritus Scientist scheme. GV and SB gratefully acknowledge Senior Research Fellowship award by CSIR. The authors acknowledge CSIR-CSMCRI for generous support towards infrastructures and core competency development. Authors also thank V. Agrawal, A. Bhatt, J. Chaudhary, R. Patidar, B. Rabary for recording IR spectra, XRD analysis, SEM images, ICP-OES and Ion chromatography analysis, respectively.

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