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Acetone fractionation: a simple and efficient method to improve the performance of lignin for dye pollutant removal

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In this work, it was found that the adsorption capacity of lignin to cationic dye (methylene blue, MB) from aqueous solution could be significantly improved by simple acetone fractionation. The removal efficiency of MB by acetone insoluble kraft lignin (AIKL) was 10 times that of unfractionated kraft lignin (KL). And the maximum capacity of AIKL could reach up to 623.4 mg g^{-1} . And the high removal rate could be achieved even at low concentrations. The effects of ionic strength, temperature, adsorbent dosage were systematically investigated. Adsorption kinetics showed the adsorption behavior obeyed the pseudo-second-order kinetic model. The equilibrium data was more consistent with the Langmuir isotherm model. Thermodynamic analyses proved that the adsorption was a spontaneous and endothermic physisorption process. In addition, the reasons for the enhanced adsorption effect by fractionation were clarified based on characterization by FT-IR. The enhancement of $\pi-\pi$ interaction between AIKL and MB caused by fractionation plays an important role in the adsorption process.

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1. Introduction

The discharge of wastewater containing dyes is a key environmental challenge.¹ Not only do they cause damage to the ecosystem but they are also harmful to human beings.² Methylene blue (MB) is one of the most commonly used cationic dyes.³ It could cause many harmful effects on human beings, such as nausea, mental confusion, jaundice, quadriplegia and tissue necrosis.⁴ Consequently, MB must be removed from wastewater effluents before discharging.

Various techniques have been used to remove dyes that include biodegradation, flocculation-coagulation, chemical oxidation, ion-exchange, reverse osmosis, ultrafiltration and adsorption.⁵⁻⁹ Among these techniques, adsorption is considered to be efficient and economical.¹⁰⁻¹² Activated carbon was a commonly adsorbent for dye removal,¹³ while the high cost limited its application on a larger scale.¹⁴

Recently, low cost adsorbents have been developed such as natural materials, agricultural by-products and industrial solid wastes.¹⁵⁻¹⁹ Especially, they were more suitable for treating solutions containing dilute MB concentrations.²⁰ Due to the

advantages of low cost, biodegradability, usability, renewability and high adsorption efficiency, they are considered promising potential alternatives to conventional methods of removing MB.

Lignin is considered as an excellent candidate among the low-cost class.²¹⁻²³ It is one of the most abundant natural polymers, second only to cellulose.²⁴ Many researchers have reported the feasibility of using various lignin as the precursors of adsorbents for MB removal. A formaldehyde modified formic lignin and a lignin-chitosan extruded blends²⁵ shown the adsorption capacity of about 35 mg g^{-1} for MB. Alkali extracted lignin, chitosan-kraft lignin composites,²⁶ and the activated carbon²⁷ prepared from lignin exhibited higher adsorption capacity among $121-147 \text{ mg g}^{-1}$. After deacetylated and fractionated, the adsorption capacity of acetic acid lignin for MB increased from 18 to 63 mg g^{-1} .²⁸ The sulfonate lignin-based hydrogels were prepared by grafting sulfonic groups on KL and the adsorption capacity could be improved to 495 mg g^{-1} .

However, kraft technology was used by mostly pulp mills for pulp production, and thus KL is easier to use for value-added production.³⁰ Although KL can be directly used as a promising adsorbent owing to their insolubility in water, its application for MB adsorption did not arouse much attention. To our delight, the authors recently found that the performance of KL for MB removal could be significantly enhanced just by easy fractionation with acetone.

Acetone was extensively used to dissolve lignin as a facile solvent. As we know, the acetone soluble kraft lignin (ASKL) has been widely reported in many applications, due to the soluble

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lignin is easier to process and react chemically in solution. For example, ASKL could improve the thermal stability of its blends with polyethylene, due to the better antioxidant activity.³¹ Also, ASKL is more easily to form lignin colloidal spheres with high sun protection factor.³² ASKL shown its tremendous potential for thermoplastic materials relations.³³ Comparing to the reports for the application of ASKL, the application of AIKL is rarely reported. Obviously, the usage of AIKL could improve the utilization efficiency for the all-components of lignin. As a bio-sorbent coming from the abandoned biomass, AIKL was obtained just by a simple physical chemistry process without any chemical modification. Thus, it has the advantage of cost effective, biodegradability, high usability, minimization of waste, in addition to good adsorption efficiency.

In this study, KL was fractionated by acetone to ASKL and AIKL. Then, the first application of the AIKL, as an adsorption material for MB was investigated. The influence of adsorption parameters such as initial dye concentration, contact time, temperature, ionic strength was investigated. The adsorption process was studied by equilibrium data on batch adsorption, and the adsorption kinetics were calculated to establish the efficiency of AIKL. The adsorption behavior of KL and ASKL were also reported as reference. The interactions between the lignin and MB was investigated by FTIR, and then the reason for the performance improvement by solvent fractionation was deduced.

2. Materials and methods

2.1. Materials

KL (softwood), purchased from the Quanlin Paper Mill in Shandong Province of China, was obtained from the straw. Cationic dye MB was supplied by Tianjin Fenghua Chemical Reagent Technology Co., Ltd. All the reagents were analytical reagent grade and working solutions were prepared by distilled water.

2.2. Methods

2.2.1. Fractionation of kraft lignin. The raw KL was suspended in acetone solutions (50 g/500 mL) with a concentration of 99.5%, and the suspension was stirred 24 h at 30 °C. The solid residue (AIKL) was then removed by filtration. ASKL was recovered by evaporating the supernatant solvent in a rotary evaporator at 100 °C. Both of the AIKL and ASKL were then followed by drying in a vacuum oven at room temperature for 12 h. The yields of ASKL and AIKL were 46.6% and 53.4%, respectively.

2.2.2. Batch adsorption experiments. The working solutions were obtained by diluting stock solution of MB (1 g L⁻¹) to the desired concentrations for adsorption experiments. All adsorption experiments were performed at on a constant temperature shaker with 0.04 g of the lignin in a 100 mL conical flask containing 50 mL of dye solutions. The characteristic absorbance was measured by a UV-vis spectrometer (Purkinje Genera, China) to determine the concentration of the dyes after adsorption, and the characteristic absorbance of MB was 664 nm. The adsorption capacity of MB on

lignin, Q_e (mg g⁻¹) was calculated according to the following equation:

$$Q_e = \frac{(C_0 - C_e)V}{m} \quad (1)$$

where, C_e and C_0 are the equilibrium and initial concentrations (mg L⁻¹) of the adsorbed dye solutions, respectively; V is the solution volumes (L) of the tested solution, respectively, m is the dry weight (g) of the lignin adsorbent.

In the solution with methylene blue concentration of 120 mg L⁻¹ at 30 °C, the lignin concentration was changed from 0.1 g L⁻¹ to 12.0 g L⁻¹ to investigate the effect of the amount of adsorbent to adsorption capacity of MB. Under the same experimental conditions, the influences of adding one electrolytes (NaCl) were investigated. The MB adsorption effected by temperature were measured in three different temperature (30, 40 and 50 °C). The MB adsorption kinetics were measured in 30 °C, samples collection time is from 5 minutes to 24 hours. The adsorption isotherms were measured at 30 °C, the contact time was 24 h, and the initial MB concentration was in the range from 10 to 300 mg L⁻¹.

2.2.3. Isotherm models

2.2.3.1. Langmuir isotherm model. The Langmuir and Freundlich models were used to interpret the equilibrium absorption data, respectively.

The Langmuir expression and its linearized form are calculated according to the following equations:

$$q_e = \frac{q_m K_L C_e}{1 + K_L C_e} \quad (2)$$

$$\frac{C_e}{q_e} = \frac{1}{q_m K_L} + \frac{C_e}{q_m} \quad (3)$$

where K_L is the Langmuir constant (L mg⁻¹) which was connected to the affinity of binding sites, q_e (mg g⁻¹) represents the equilibrium adsorption capacity per gram adsorbent to MB, C_e (mg L⁻¹) represents the equilibrium concentration of MB, and q_m (mg g⁻¹) represents the theoretical maximum of adsorption capacity. Thus, the validity of the Langmuir model was confirmed by the linear plot of C_e/q_e versus C_e . And R_L is defined as:

$$R_L = \frac{1}{1 + K_L C_0} \quad (4)$$

where C_0 (mg L⁻¹) represents the initial concentration in the liquid phase; K_L (L mg⁻¹) represents the Langmuir constant; R_L is the dimensionless parameter of the equilibrium or adsorption intensity.

2.2.3.2. Freundlich isotherm model. The Freundlich expression and its linearized form are calculated according to the following equations:

$$q_e = K_F C_e^{\frac{1}{n}} \quad (5)$$

$$\ln q_e = \ln K_F + \ln \frac{C_e}{n} \quad (6)$$

where n is the heterogeneity factor and the $1/n$ value related to the adsorption intensity. K_F is the Freundlich constant (mg g⁻¹)



and K_F value is related to the adsorption capacity. A plot of $\ln q_e$ versus $\ln C_e$ shows a straight line, where K_F and $1/n$ are determined by intercept and slope respectively.

2.2.4. Adsorption kinetics. In order to explore the adsorption kinetics, pseudo-first-order (7) and pseudo-second-order (8) kinetic equations were used to analyze adsorption kinetic data:

$$\ln (q_e - q_t) = \ln q_e - k_1 t \quad (7)$$

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \quad (8)$$

where q_e and q_t are the adsorption capacity of methylene blue adsorbed (mg g^{-1}) and time t (min) at equilibrium, respectively. k_1 (min^{-1}) and k_2 ($\text{mg}^{-1} \text{min}$) are the rate constants of pseudo-first-order and pseudo-second-order kinetics equations, respectively.

2.2.5. Thermodynamic study. The thermodynamic parameters were confirmed by van't Hoff analysis to elucidate the thermodynamic origins of the adsorption process. Through van't Hoff equation, the relationship between Langmuir adsorption constant K_L and temperature T were described shown as follows:

$$\ln K_L = \frac{\Delta S}{R} - \frac{\Delta H}{RT} \quad (9)$$

where T (K) is the absolute temperature, R ($8.314 \text{ J mol}^{-1} \text{ K}^{-1}$) is the universal gas constant, ΔH ($\text{J mol}^{-1} \text{ K}^{-1}$) is the enthalpy change; and ΔS ($\text{J mol}^{-1} \text{ K}^{-1}$) is the entropy change. ΔH and ΔS were calculated from the slope and intercept of linear plots of $\ln K_L$ vs. T^{-1} , respectively.

The Gibbs energy change, ΔG (J mol^{-1}), was estimated from the following relationship:

$$\Delta G = -RT \ln K_L \quad (10)$$

3. Results and discussion

3.1. Effects of adsorbent dosage and treating effect

Fig. 1 shows the effects of the sample dosage of KL, ASKL and AIKL on MB adsorption. The removal rate of MB increased with

the increasing adsorbent dosage because of the increase of the binding sites in adsorption. And the adsorption capacity of MB decreases with increasing adsorbent. This is owing to the higher adsorbent dosage could provide a large excess of the active sites which causing a decrease in the utility of active sites. Obviously, AIKL is superior to KL and ASKL in MB removal under the lower dosage. The removal rate of MB by AIKL could approached 100% when the dosage was 1.8 g L^{-1} . However, for the same removal rate, no less than 11 g L^{-1} of KL was needed. And the removal rate of MB was only 44% while the dosage of ASKL was up to 10 g L^{-1} . It means that, compared to KL and ASKL, only a little AIKL was needed for MB removal.

From UV-vis spectra (Fig. 2), the characteristic peaks of the MB show obvious decrease or completely disappear after removal by the three samples at the same dosage. Meantime, the colour of solution (inset of Fig. 2) were also getting lighter. Especially, the solutions treated by AIKL were almost completely decolorized, indicating the most efficient removal of the MB from water.

3.2. Adsorption kinetics

The time-dependent absorption of MB on three kinds of lignin samples were measured, and results are presented in Fig. 3. The adsorption of MB on the three samples start out fast, then slow down and finally reach to equilibrium. Steep slope of MB adsorption on the three samples at the initial period indicated the instantaneous adsorption which might be due to the abundant adsorption sites for MB.

The values of k_1 was calculated by the plots of $\ln(q_e - q_t)$ versus t (Fig. 5). The linear plots of t/q_t versus t (Fig. 4), and values of k_2 and q_e were calculated from the intercept and slope of the plots. The experimental adsorption capacities and characteristic parameters calculated were shown in Table 1. The correlation coefficients (R^2) for pseudo-second-order kinetics were better than that for pseudo-first-order kinetics. Furthermore, the q_e calculated from pseudo-second-order kinetic model was more consistent with the experimental data (q_{ex}) than the pseudo-first-order kinetic model. Thus, the pseudo-second-order kinetics could describe the adsorption process more reasonably in comparison to the pseudo-first-order models.

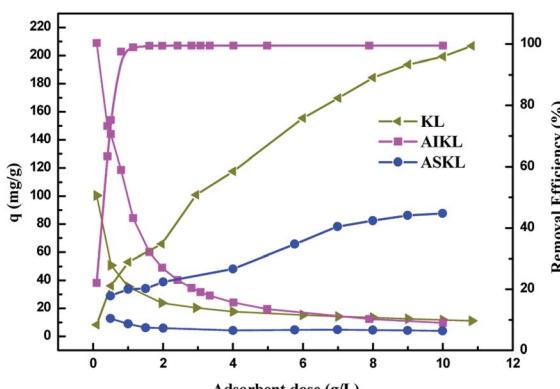


Fig. 1 Effect of the adsorbent dosage on the adsorption of MB (initial C_{MB} : 0.12 g L^{-1} , temp.: 30°C).

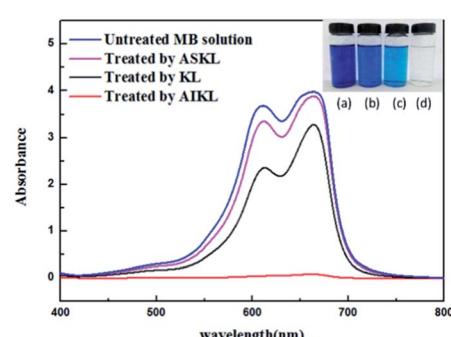


Fig. 2 UV-vis absorption spectra of the MB solution before and after treating. Inset: color of MB solution adsorbed by: (a) the control, (b) ASKL, (c) KL, (d) AIKL.



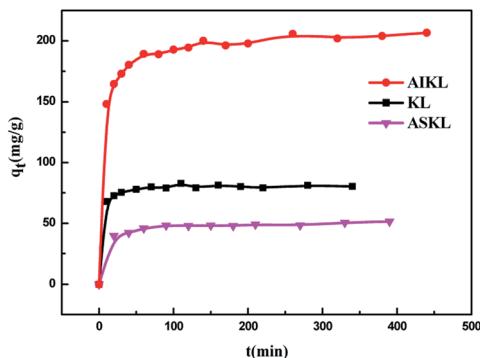


Fig. 3 Effect of contact time on the adsorption of MB onto KL, ASKL and AIKL at 30 °C.

From Fig. 3, the time required for ASKL and KL to reach the adsorption equilibrium was less than AIKL. The adsorption process of MB on AIKL in the following stage need a relative long time. Also, the values of k_2 for KL and ASKL was higher than AIKL, indicating the adsorption process of the latter system was slow.

As a matter of fact, AIKL is easier to form larger aggregates due to it is more hydrophobic. Thus, there are more inner cavities or porous structures formed by aggregation. The adsorption sites in the inner space of AIKL aggregates was more tough to contact MB, resulting in the longer adsorption equilibrium time and a lower adsorption rate.

3.3. Adsorption isotherms

Fig. 6 shows the adsorption isotherms of MB on KL, AIKL and ASKL and the uniformity and non-uniformity of adsorbent surface can be obtained by its shape. The adsorption capacity of MB on the three samples all increased with the increase of the equilibrium concentration and finally reached the saturation value. At the same equilibrium concentration, the adsorption capacity of MB on AIKL is much higher than that of KL, while the adsorption capacity of MB on ASKL is a little lower than KL. It is noteworthy that the absorption amount of MB on AIKL increased almost vertically with initial lower dye concentration,

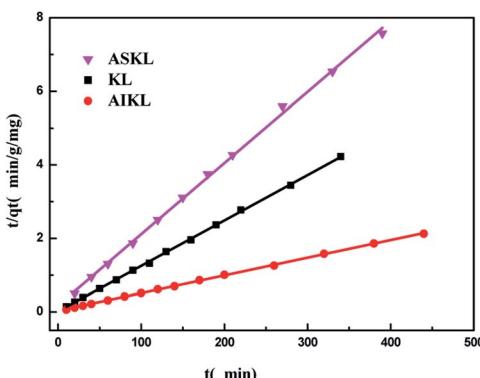


Fig. 4 Second-order kinetic equation for adsorption of methylene blue on to KL, ASKL and AIKL.

making the AIKL have the advantage of removing lower concentration wastewater.

Fig. 7 shows the Langmuir plot for the adsorption of MB onto AIKL, KL and ASKL, from the linearized Langmuir equation. The Langmuir constants K_L and q_m are shown in Table 2, the equilibrium data has a higher correlation coefficient and well adapted of Langmuir model. The adsorption isotherm data was also analyzed by Freundlich model, but its poor fit (Table 2) indicated that it was not suitable for MB/lignin system.

Due to the better fits with Langmuir isotherm, the adsorption of MB onto the three adsorbents was monolayer physical adsorption on homogeneous surfaces.

The separation factor R_L , calculated by eqn (4), was the dimensionless parameter of the adsorption intensity. The type of the isotherm was either unfavorable ($R_L > 1$), linear ($R_L = 1$), favorable ($0 < R_L < 1$) or irreversible ($R_L = 0$) could be inferred by the value of R_L (eqn (4)). For adsorption of MB onto three adsorbents in this study, all the R_L values (Table 2) are in the range of 0.0188 to 0.4615, indicating the adsorption is a favorable process.

MB adsorption had been studied by a variety of different adsorbents. MB adsorption had been studied by a variety of different adsorbents. Table 3 compared the maximum adsorption capacity of MB for various biomass adsorbents. Compared with the other modified natural materials such as banana pseudostem, *Platanus orientalis* leaf and *Luffa cylindrica* fibers,³⁷⁻³⁹ the advantage of AIKL is that the aromatic ring of lignin could have $\pi-\pi$ interaction with MB which is beneficial to the adsorption. As for lignin-based hydrogels prepared by crosslinking poly used methyl vinyl ether co-maleic acid,³⁵ the preparation of AIKL only requires one step of dissolution, the entire process is simple to operate and can be easily scalable.

Although its adsorption capacity is not the highest, it possesses the merits of renewable and easy to mass produce. Thus, AIKL could be a promising low-cost adsorbent for MB removal from solution.

According to the Langmuir equation, the maximum uptake capacity of MB on AIKL were 200.0 mg g⁻¹. While, the saturated adsorption capacity of KL and ASKL were 79.9 mg g⁻¹ and 49.9 mg g⁻¹, respectively. The saturated adsorption capacity of

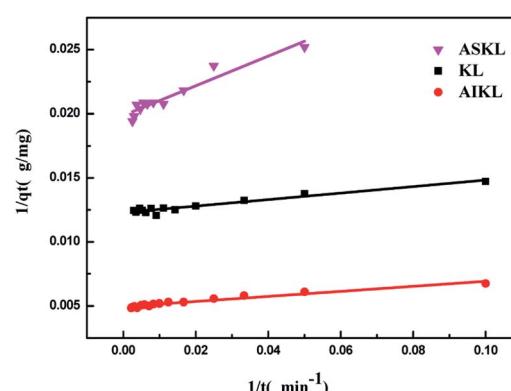


Fig. 5 Pseudo-first-order kinetic fit curve of MB adsorption on AIKL, KL and ASKL.



Table 1 Adsorption kinetic parameters for the adsorption of MB onto AIKL, ASKL and KL

Sorbent	Pseudo-first-order			Pseudo-second-order			
	k_1 (min ⁻¹)	q_e (mg g ⁻¹)	R^2	q_{ex}	k_2 (g mg ⁻¹ min ⁻¹)	q_e (mg g ⁻¹)	R^2
AIKL	3.9939	202.4291	0.9499	206.6405	0.0007	207.9002	0.9997
KL	2.0717	81.4332	0.9293	80.3529	0.0084	80.9717	0.9996
ASKL	5.8524	50.3778	0.9113	51.4837	0.0021	51.6796	0.9985

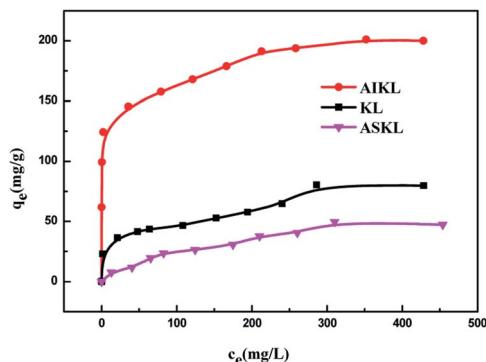


Fig. 6 Isotherms of MB adsorption onto KL, AIKL and ASKL at 30 °C.

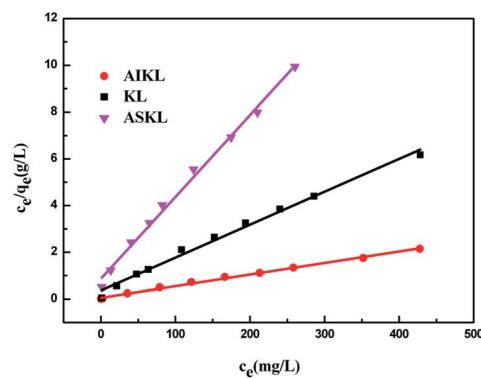


Fig. 7 Langmuir plot of adsorption of MB onto KL, ASKL and AIKL.

AIKL for MB increased about 2.5 times than KL. It is worth nothing that, the data of adsorption capacities for the three samples could not be explained by the conservation of mass. When the KL was fractionated by acetone, the yields of ASKL and AIKL were 46.7% and 53.3%, respectively. If we sum up the adsorption capacities ASKL and AIKL in proportion, that is:

$$q_{e1} \times \omega_1 + q_{e2} \times \omega_2 = Q'_{KL} > Q_{KL}$$

where q_{e1} and q_{e2} are the adsorption capacity (mg g⁻¹) of MB by AIKL and ASKL, respectively. The ω_1 and ω_2 are the yields of AIKL and ASKL, respectively. The Q'_{KL} represents the adsorption capacity calculated by mass conservation. Q_{KL} is the adsorption capacity of KL. It could be found Q'_{KL} (118.6 mg g⁻¹) is much larger than Q_{KL} (79.9 mg g⁻¹). Obviously, this did not obey the mass conservation. Thus, there must be some underlying reasons for explain this unpredictable result, which was described in the following paragraphs.

3.4. Effects of operating conditions on MB adsorption

3.4.1. Effect of temperature. Dye effluents are usually produced at relatively high temperatures.³⁷ To figure out the influence of temperature, adsorption experiments were carried out at 30, 40 and 50 °C, respectively. As shown in Fig. 8, the MB adsorption capacity increased significantly from 200.0 to 345.0 mg g⁻¹ with the increase in temperature from 30 to 50 °C. Which means the AIKL could be a high-efficiency adsorbent for MB removal at higher temperatures. The thermodynamic parameters calculated from eqn (9) and (10) are listed in Table 4. As we know, the change in free energy for physical absorption is between -20 and 0 kJ mol⁻¹.⁴⁰ The values of ΔG obtained in this study are within this ranges, indicating that the adsorption is physisorption. Moreover, with the increase of temperature the value of ΔG became more negative, indicating that a higher temperature was more favorable to the adsorption process (Fig. 9).

The positive enthalpy changes suggest that the adsorption process of MB onto AIKL is endothermic. Generally, adsorption was exothermic processes due to the binding of solute and adsorbent, and the adsorption capacity decreases with the increase of temperature. A reasonable explanation for the positive enthalpy changes of this system was the high energy water molecules were released from the desolvation of MB and

Table 2 The adsorption parameters of Langmuir and Freundlich model for the adsorption of MB onto AIKL, KL and ASKL

Sorbent	Langmuir isotherm parameters			Freundlich isotherm parameters			
	q_m (mg g ⁻¹)	K_L (L mg ⁻¹)	R^2	R_L	K_F (mg ^{1-1/n} L ^{1/n} g ⁻¹)	n	R^2
AIKL	200.0000	0.1042	0.9933	0.0188–0.1610	91.9741	7.800	0.9318
KL	79.9361	0.0398	0.9910	0.0537–0.3620	21.2886	4.587	0.9453
ASKL	49.8753	0.0233	0.9940	0.0789–0.4615	1.3265	1.713	0.9753

Table 3 Comparison of the adsorption capacity for MB by several adsorbents

Adsorbent	Adsorption capacity (mg g ⁻¹)	Reference
AIKL	200.00	This work
Lignin–chitosan composite	36.25	25
Lignin–silica composite	60.00	34
Lignin-based hydrogels	69 to 629	35
Sawdust	27.78–59.17	36
Kapok fiber	110.13	14
<i>Platanus orientalis</i> leaf	114.94	37
<i>Luffa cylindrica</i> fibers	122.00	38
Activated carbon prepared from lignin	147.00	27
Banana pseudostem biochar	146.23	39

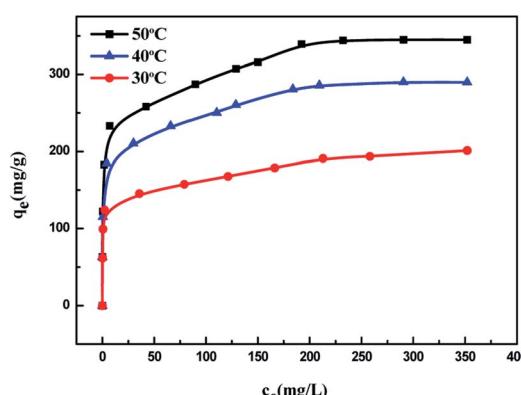


Fig. 8 Effect of the temperature on the adsorption of MB.

AIKL. As we know, the high-energy water molecules were released from the cavity upon polycyclic aromatic compounds, leads to the endothermic reaction. Increasing the temperature could reduce of hydration layer remarkably, resulting the enhance of hydrophobic interaction which could increase the adsorption capacity.

This opinion could be further confirmed by the positive value of the entropy change (ΔS). It indicates that the adsorption of MB on AIKL was completely driven by the entropic increase (Table 4) which is a typical characteristic of hydrophobic interaction in adsorption.⁴¹ Thus, the hydrophobic interaction could be considered as one of the main binding contributions for the adsorption of MB onto AIKL.

3.4.2. Effect of ionic strength. The sodium chloride as a stimulator is commonly used in dying processes, and the concentration of salt in industrial MB wastewater is high.⁴⁰

Table 4 Thermodynamic parameters for the adsorption of MB onto AIKL

ΔH (kJ mol ⁻¹)	ΔS (J mol ⁻¹ K ⁻¹)	ΔG (kJ mol ⁻¹)		
		303 K	313 K	323 K
74.3996	10.8422	-11.7051	-12.4357	-13.1937

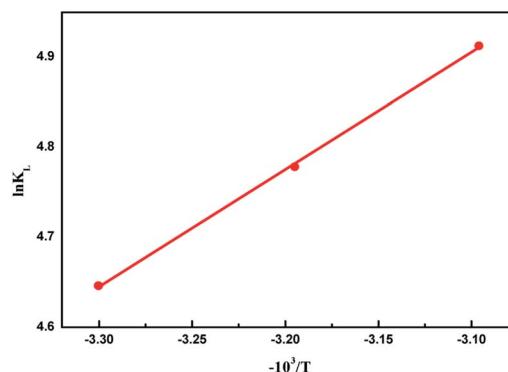


Fig. 9 Thermodynamic plots for the adsorption of MB onto AIKL.

Therefore, the effect of ionic strength on the removal of MB by AIKL was invested at different concentrations of NaCl. As shown in Fig. 10, at first, the adsorption capacity decreased slightly with the increase of ionic strength, but then increased sharply when the concentration exceeded 0.316 mol L⁻¹.

When the NaCl was added to the solution, it would screen the charges of MB and AIKL, and thus inhibits the cation–anion interactions between them. So, NaCl plays an adverse effect on the electrostatic attraction for MB adsorption. When enough

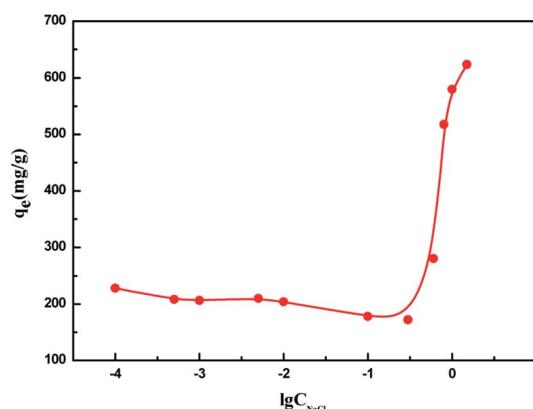


Fig. 10 Effects of ionic strength (NaCl concentration) on the adsorption of MB.



salt was added, the electrostatic attraction between MB and AIKL becomes negligible. However, the adsorption amount increased obviously. It indicated that there must exist some other important driving forces besides the electrostatic interaction. As we know, the increase of ionic strength could reduce the barrier of the hydration layer,⁴³ and the hydration layer barriers the combination of MB and AIKL molecules by hydrophobic interaction. Thus, the experimental results further proved that the hydrophobic interaction was beneficial to the adsorption.

3.5. Strengthening mechanism of acetone fractionation

As discussed in the effect of solvent conditions above on, the electrostatic attraction and hydrophobic interaction were the two main forces of adsorption. As we know, after acetone fractionation, due to the reduction of phenolic hydroxyl ($-\text{OH}$) and carboxyl groups ($-\text{COOH}$), AIKL has fewer negative charges than ASKL and KL.⁴⁴ Interestingly, AIKL with fewer negative charge has much higher adsorption capacity. Thus, the electrostatic attraction may not play a central role for the adsorption between lignin and MB. While, the hydrophobic interaction may play the dominant role.

The band around 1460 cm^{-1} of these lignin samples, ascribed to vibration of aromatic ring, all shifted to low wavenumber when MB was adsorbed on them. And the main aromatic peaks of these samples, before and after adsorption, were listed in Table 5. The shifts in the main aromatic peak position to lower wavenumber in vibrational spectroscopy provides direct evidence for $\pi-\pi$ interactions. Moreover, the peaks around 1460 cm^{-1} of AIKL caused by MB adsorption shift about 6.04 cm^{-1} , while, it is only 1.52 cm^{-1} for KL and 0.46 for ASKL. This information indicated that the $\pi-\pi$ interactions of AIKL-MB is stronger than that KL-MB.

Indeed, the $\pi-\pi$ interactions between AIKL and MB not only can be confirmed by the results of FT-IR, but also could be explained reasonably (Fig. 11). The strengthening mechanisms of acetone fractionation were schematically described in Fig. 12. The MB is an ideal planar molecule which has C-C double bonds and contains π -electrons. The aromatic rings of lignin are oriented preferentially parallel to the surface. Also, the associated KL complexes is a flexible lamellar configuration. Obviously, this spatial structure feature of lignin could contribute to the formation of $\pi-\pi$ stacking between MB and lignin. The similar $\pi-\pi$ electron donor acceptor interaction also

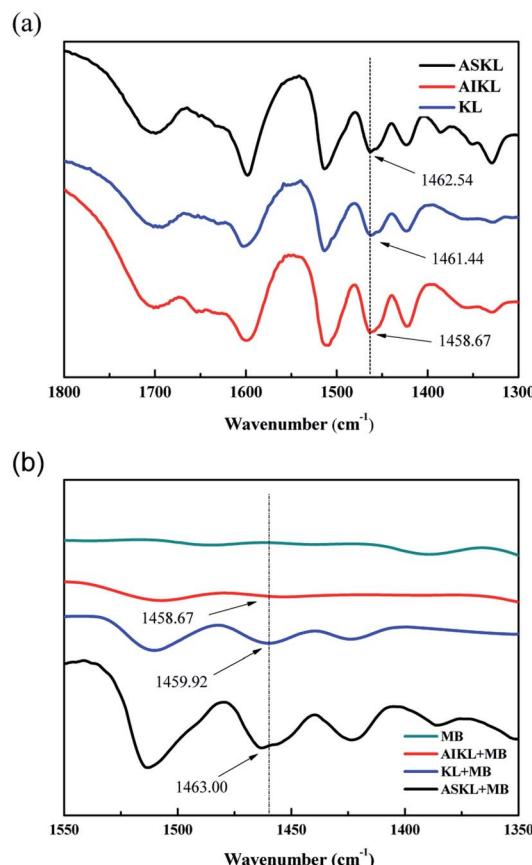


Fig. 11 FTIR spectra of KL, ASKL and AIKL (a); and KL, ASKL and AIKL after MB loading (b).

play an important role when MB were adsorbed onto graphene oxide⁴⁵ and carbon nanotubes.⁴²

It is reported that, the lignin after fractionation by acetone-aqueous solution of different concentrations, the molecular weight of lignin decreases with the decrease of acetone concentration, and the content of phenolic hydroxyl group and carboxyl group increases.^{44,46,47} And the AIKL with higher molecular weight after fractionation tend to have more aromatic groups.⁴⁸ On the one side, more aromatic groups could provide more effective adsorption sites for the reception of MB by $\pi-\pi$ interactions. On the other side, less phenolic hydroxyl and carboxyl groups could reduction the hydration layer which was beneficial to the $\pi-\pi$ interaction between AIKL and MB

Table 5 The aromatic skeletal vibration (around 1460 cm^{-1}) of different lignin samples before and after adsorption of MB

Before and after adsorbed by AIKL			
Adsorbent	Unloaded by MB (cm^{-1})	Loaded by MB (cm^{-1})	Δ (cm^{-1})
KL	1461.44	1459.92	-1.52
ASKL	1462.54	1463.00	0.46
AIKL	1458.67	1452.63	-6.04

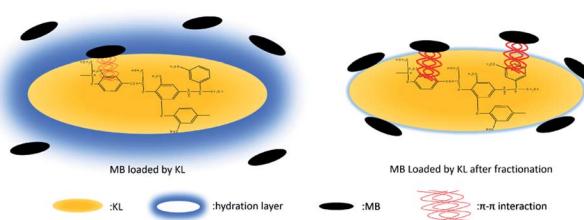


Fig. 12 Schematic illustration of enhanced adsorption after acetone grading.



aromatic rings. By the promotion of these two reasons, the performance of AIKL has been significantly improved.

4. Conclusions

In summary, the surface properties of KL were obviously improved by solvent fractionation treatment. The reduction of hydration layer on AIKL surface was beneficial to the π -electron coupling between the aromatic molecules of MB and AIKL. The increase of ionic strength and temperature would cause the reduction of hydration layer which lead to the increase of adsorption capacity. Compared to other adsorbents, AIKL has advantage of inherent adsorption ability for aromatic compounds through stronger π - π interaction. Therefore, the fractionated AIKL adsorbents derived from abundant, low-cost and non-toxic KL, would be a promising adsorbent for aromatic pollutants removal in wastewater.

Conflicts of interest

There are no conflicts to declare.

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