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Silkworm cocoon derived N, O-codoped hierarchical porous carbon with ultrahigh specific surface area for efficient capture of methylene blue with exceptionally high uptake: kinetics, isotherm, and thermodynamics†

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Dyes are typical water contaminants that seriously affect water quality. In this study, silkworm cocoon derived N, O-codoped hierarchical porous carbon was successively developed via a facile pre-carbonization and chemical activation method, and characterized thoroughly by SEM, TEM, HRTEM, XRD, Raman, N₂ adsorption and XPS. The as-prepared N, O-HPC showed a well-developed porous structure with an ultra-high specific surface area of 2270.19 m² g⁻¹, which proved to be a high-efficiency adsorbent. Batch adsorption experiments demonstrated that MB adsorption was highly dependent on contact time, initial MB concentration, temperature and initial solution pH. However, no remarkable effects of humic acid and ionic strength were observed. In the kinetic studies, the good applicability of a pseudo-second-order kinetic model was demonstrated. The adsorption isotherm study showed that a Langmuir isotherm model can describe the experimental data much more suitably with a maximum monolayer adsorption capacity value of 2104.29 mg g⁻¹, which is among the highest in previously reported adsorbents and ascribed to multiple adsorption mechanisms including pore filling, π–π stacking interaction and electrostatic interaction between MB and N, O-HPC. Thermodynamic analyses suggested that MB adsorption onto N, O-HPC was spontaneous and endothermic. Furthermore, the as prepared adsorbent showed highly efficient adsorption for MB in tap water and synergistic adsorption performance toward MB and MO. Therefore, N, O-HPC derived from silkworm cocoon could be considered as an efficient, novel and advantageous material for wastewater remediation.

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1 Introduction

Nowadays, dyes are widely used in the textile, leather, paper-making, dye synthesis, printing, food, cosmetics and plastics industries.¹ Water contamination caused by these industries has drawn more and more attention, since almost all dyes are poorly biodegradable or resistant to environmental conditions and, therefore, create major problems in the treatment of wastewater stemming from the dyeing industry.² Currently, various treatment methods including physical, chemical and biological schemes have been developed to remove dyes from wastewater, such as photocatalytic degradation,^{3–6} electrochemical degradation,^{7,8} photochemical degradation,^{9,10}

sonochemical degradation,^{11–13} sonocatalytic degradation,^{14,15} Fenton-like degradation,^{16–18} photo-Fenton degradation,^{19–21} advanced oxidation process,^{22–25} electro-Fenton degradation^{26,27} and adsorption.^{28–35} Among those, adsorption is proven to be a potentially powerful technique for removing pollutants from contaminated aqueous media. It is preferred over other methods because of its relative simplicity of design, ease of operation, cost effectiveness, insensitivity to noxious pollutants, energy efficiency and environmental benignity.^{36–39} Numerous adsorbents, such as MOF-based material,³³ geopolymers,³¹ graphene-based material,^{30,32} palygorskite and its composites,^{40,41} clay-based minerals,^{42,43} agricultural peels,⁴⁴ hydrogels,^{45–47} carbon nanotubes,⁴⁸ *etc* have been studied. Among the adsorbents mentioned, the carbonaceous materials usually have the highest dye removal efficiency due to their large surface area and the functional groups on their surface that facilitate interactions with various dyes.¹

Methylene blue (MB) is a cationic dye that is most commonly used for colouring. It is generally used for dyeing cotton, wool, and silk. MB can cause eye burns in humans and animals, methemoglobinemia, cyanosis, convulsions, tachycardia,

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dyspnea, irritation to the skin, and if ingested, irritation to the gastrointestinal tract, nausea, vomiting, and diarrhea.⁴⁹ Therefore, MB often serves as a model compound for removing organic contaminants and coloured bodies from aqueous solutions.

Biomass materials are promising precursors that can be used to prepare hierarchically porous carbons with high surface area and rich heteroatom doping. As a natural bio-polymer, silkworm cocoon is mainly composed of protein with fibroin being enveloped by the glue-like sericin, and the rich heteroatoms including N atoms is beneficial for the preparation of nitrogen-doped carbons during pyrolysis.⁵⁰ Moreover, the trace inorganic salts dispersing in the silk is a natural template that is favourable for the homogeneously creation and expansion of pores through adjusting the activation condition during pyrolysis, facilitating the formation of hierarchically porous structure.⁵¹

Herein, we reported a successful synthesis of N, O-codoped hierarchical porous carbon (N, O-HPC) with ultrahigh specific surface area by a pre-carbonization and a subsequent KOH activation of silkworm cocoon. The as-prepared N, O-HPC with an ultrahigh specific surface area of 2270.19 m² g⁻¹ proved to be novel and highly efficient adsorbent for MB. The effects of contact time, initial MB concentration, temperature, initial solution pH, adsorbent dosage, ionic strength and co-existing humic acid were systematically investigated. The adsorption process fitted well with the Langmuir isotherm model and pseudo-second-order kinetic model, and the maximum adsorption capacity was 2104.29 mg g⁻¹. To the best of our knowledge, this is among the highest in the previously reported adsorbents. Moreover, MB adsorption onto N, O-HPC was spontaneous and endothermic, and the as prepared adsorbent showed highly efficient adsorption for MB in tap water and synergistic adsorption performance toward MB and MO. Therefore, N, O-HPC derived from silkworm cocoon is expected to be a promising adsorbent for the efficient removal of dyes from wastewater.

2 Experimental

2.1 Materials

The raw silkworm cocoon was directly obtained from sericulturists of Sichuan, China. Potassium hydroxide (KOH) were purchased from Aladdin Chemistry Co. Ethanol absolute and hydrochloric acid (37 wt%, HCl) were purchased from Sino-pharm Chemical Reagent Co., Ltd. Methylene blue (MB) was purchased from Shanghai Macklin Biochemical Co., Ltd. All the chemicals used in the sample preparation were used as received without further purification. Furthermore, deionized water was used to prepare different concentrations of MB solutions.

2.2 Preparation of N, O-codoped hierarchical porous carbon (N, O-HPC)

In a typical process, after cut into pieces, silkworm cocoon was washed and dried at 80 °C for 48 h. Subsequently, the above dry and clean precursor was loaded into a corundum crucible and

pre-carbonized in a horizontal tubular furnace at 450 °C for 0.5 h with a heating rate of 5 °C min⁻¹ under N₂ flow. After cooling down to room temperature, the resulting carbon material was ground into fine powder, which was thoroughly impregnated in KOH aqueous solution at a mass ratio of C : KOH = 1 : 1 and dehydrated at 80 °C to form a slurry. Afterward, the mixture was transferred into a corundum crucible and activated in a tubular furnace at 900 °C for 2 h with a heating rate of 2 °C min⁻¹ under N₂ flow. After being cooled to ambient temperature, the activated product was adjusted to pH ≈ 7 by 1 M HCl solution and magnetically stirred for several hours. After that, the sample was centrifuged and repeatedly washed abundantly with deionized water and finally dried at 80 °C overnight to obtain the final product (denoted as N, O-HPC).

2.3 Characterization of N, O-HPC

Scanning electron microscopy (SEM) measurements were conducted on a Hitachi S-4800 field emission scanning electron microscope. Transmission electron microscopy (TEM) and high-resolution transmission electron microscopy (HRTEM) images were taken from a JEM-2100F (field emission) scanning transmission electron microscope (JEOL, Japan) with an acceleration voltage of 200 kV. X-ray powder diffraction (XRD) pattern was recorded on a Rigaku Dmax 2200 diffractometer with Cu K α 1 radiation ($\lambda = 1.5416 \text{ \AA}$, operating at 40 kV and 40 mA). Raman spectrum was recorded on a LABRAM-HR confocal laser micro-Raman spectrometer (Jobin-Yvon, France) in the backscattering configuration excited with an Ar⁺ ion laser ($\lambda = 514.5 \text{ nm}$). X-ray photoelectron spectroscopy (XPS) analyses were measured using an ESCALAB 250Xi X-ray photoelectron spectrometer with Al K α monochromatic radiation at a constant pass energy of 1486.6 eV (ThermoFisher Scientific, USA). The nitrogen adsorption-desorption isotherms were carried out by N₂ physisorption at 77 K on an automatic Micromeritics ASAP 2020M nitrogen adsorption analyzer, and with the adoption of the Brunauer-Emmett-Teller (BET) model and nonlocal density functional theory (NLDFT) methods to calculate the specific surface area (SSA) and the pore size distribution (PSD). Fourier transform infrared (FTIR) spectra were obtained using a FTIR spectrometer (IRAffinity-1, Shimadzu, Japan).

2.4 Adsorption experiments

Batch adsorption experiments was conducted in a set of 200 mL glass beakers containing 100 mL MB solutions at a consistent stirring rate. The initial solution pH was adjusted by using 0.1 M NaOH or 0.1 M HCl solution. Typically, the adsorption kinetics was investigated at 293 K by adding 0.05 g N, O-HPC into 700 mg L⁻¹ MB solution without adjusting pH value. The adsorption isotherm experiments were carried out at 293 K by adding 0.05 g N, O-HPC into different MB solutions with various initial concentrations (400–1000 mg L⁻¹) and without adjusting pH. The adsorption thermodynamics was studied by conducting the adsorption experiment at 293 K, 303 K and 313 K. The effects of experimental parameters such as N, O-HPC dosage, initial solution pH, ionic strength and humic acid on the



adsorption process were also investigated. Moreover, the adsorption of MB in tap water and synergistic adsorption performance toward MB and MO were also evaluated. At pre-determined time intervals, 1 mL MB solution was withdrawn and filtrated using a 0.22 μm syringe filter, and then measured by UV-vis spectrophotometer at its maximum wavelength of 664 nm. The amount of MB adsorbed q_t and the dye removal efficiency (R) at any time t , was calculated using the following equations:

$$q_t = \frac{(C_0 - C_t)V}{m}$$

$$R = \frac{(C_0 - C_t) \times 100}{C_0}$$

where C_0 and C_t (mg L^{-1}) are the liquid-phase concentrations of MB at initial and any time t , respectively. V (L) is the volume of MB solution, and m (g) is the mass of dry adsorbent used.

2.5 Determination of point of zero charge (pHpzc)

For the determination of pHpzc of the adsorbent, 10 mL of 0.01 M NaCl solution was added into a series of 20 mL capped glass bottles. The initial pH (pH_i) of solutions was adjusted between 2 to 10 by using 0.1 M HCl or 0.1 M NaOH. Then, 10 mg of N, O-HPC was added to each glass bottle, and the obtained suspensions were shaken in controlled temperature (20 ± 2 °C) with a speed of 150 rpm. After 48 h, the suspensions were filtered and final pH values (pH_f) of the filtrates were measured again. The difference between pH_i and pH_f values ($\Delta\text{pH} = \text{pH}_i - \text{pH}_f$) was plotted *versus* the pH_i . The point of intersection of the resulting curve with abscissa, at which $\Delta\text{pH} = 0$, provided the point of zero charge.^{52,53}

2.6 Adsorption kinetics

The Lagergren pseudo-first order,⁵⁴ pseudo-second order⁵⁵ and intra-particle diffusion⁵⁶ models were used to describe the experimental adsorption kinetic data. The Lagergren pseudo-first-order equation is expressed as:

$$q_t = q_e[1 - \exp(-k_1t)]$$

where q_e and q_t are the amount of adsorbate adsorbed (mg g^{-1}) at equilibrium and at time t , respectively. t is the contact time (min) and k_1 , the rate constant (min^{-1}).

The pseudo-second order model, proposed by Ho and McKay, can be represented in the following linear form:

$$\frac{t}{q_t} = \frac{t}{q_e} + \frac{1}{k_2q_e^2}$$

where k_2 is the pseudo-second-order rate constant ($\text{g mg}^{-1} \text{min}^{-1}$); $k_2q_e^2$ is the initial adsorption rate (mg g^{-1}). The values of q_e and k_2 were obtained by a linear plot t/q_t *versus* t .

The intraparticle diffusion model equation can be written as:

$$q_t = k_{\text{intra}}\sqrt{t} + C$$

where k_{intra} is the intraparticle diffusion rate constant ($\text{mg g}^{-1} \text{min}^{-0.5}$) and C represents the intercept (mg g^{-1}). A plot of q_t *versus* $t^{0.5}$ should be in straight line if the adsorption process follows the intraparticle diffusion; meanwhile if the line passes through the origin then intra particle diffusion will be the sole controlling step.

2.7 Adsorption isotherms

In this study, the equilibrium adsorption data were analysed by Langmuir,^{57,58} Freundlich⁵⁹ and Temkin⁶⁰ models. Langmuir isotherm model has been successfully applied to many adsorption processes. It assumes that adsorption occurs at specific homogenous sites within the adsorbent. Once a molecule occupies a site, no further adsorption will take place at the site. The linear form of the Langmuir isotherm equation can be represented by the following equation:

$$\frac{C_e}{q_e} = \frac{C_e}{q_m} + \frac{1}{k_Lq_m}$$

where C_e (mg mL^{-1}) is the equilibrium concentration of adsorbate in the solution; q_m (mg g^{-1}), the maximum adsorption capacity that corresponds to complete monolayer coverage, and k_L , the equilibrium constant (mL g^{-1}).

The essential characteristics of the Langmuir isotherm can be expressed by the dimensionless constant separation factor (R_L) which is given by:

$$R_L = \frac{1}{1 + k_L C_0}$$

The adsorption process can be defined as irreversible ($R_L = 0$), favorable (R_L between 0 and 1), linear ($R_L = 1$) or unfavorable ($R_L > 1$).

The Freundlich isotherm assumes a heterogeneous system with different energies of active sites and reversible adsorption, which is not restricted to monolayer sorption. It is an empirical equation which is used for the description of multilayer adsorption with interaction between adsorbed molecules. The Freundlich isotherm is shown as below:

$$\log q_e = \log k_F + \frac{1}{n} \log C_e$$

where k_F (mg g^{-1}) is Freundlich constant representing the adsorption capacity; n (dimensionless), an empirical parameter related to the adsorption intensity. The value of n varies with the heterogeneity of the adsorbent and for favourable adsorption process the value of n should lie in the range of 1–10.

Temkin isotherm assumes that the heat of adsorption of all the molecules in the layer would decrease linearly with coverage and the adsorption is characterized by a uniform distribution of binding energies, up to some maximum binding energy. The Temkin isotherm equation is as follow:

$$q_e = \frac{RT}{b} \ln A_T + \frac{RT}{b} \ln C_e$$

where R is the gas constant; T , the absolute temperature; A_T (L mg^{-1}), the Temkin equilibrium binding constant and b (dimensionless), the Temkin isotherm constant.



2.8 Adsorption thermodynamics

Thermodynamics parameters including enthalpy change (ΔH^0), Gibbs free energy change (ΔG^0) and entropy change (ΔS^0) can be determined as follows:⁶¹

$$\Delta G^0 = \Delta H^0 - T\Delta S^0$$

$$\Delta G^0 = -RT \ln K_d$$

$$K_d = \frac{q_e}{C_e}$$

$$\ln K_d = \frac{\Delta S^0}{R} - \frac{\Delta H^0}{RT}$$

where K_d is the equilibrium constant; q_e (mg g⁻¹) is the amount of adsorbate that is adsorbed at equilibrium; C_e is equilibrium concentration (mg L⁻¹) of adsorbate in the solution; T is the absolute temperature (K) and R is the universal gas constant.

3 Results and discussion

3.1 Characterization of N, O-HPC

N, O-HPC was successively developed *via* a facile pre-carbonization and chemical activation method (Scheme S1†). The SEM and TEM images of N, O-HPC exhibit their morphologies and porous structures. As shown in Fig. 1A and B, N, O-HPC displays a loose and well-developed porous structure with many open cavities. Specifically, under high-resolution, N, O-HPC exhibits a typical honeycomb-like porous structure with abundant interconnected mesopores and macropores (Fig. 1C). TEM and HRETEM images also reveal a porous structure for N, O-HPC. As shown in Fig. 1D–F, the synthesized sample has an interconnected 3D framework and porous structure with high-density macropores, mesopores and micropores. The high-density pores and open cavities are apparently favourable for the quick transportation of adsorbate molecules. Promisingly, some bent lattice fringes with a spacing of about 0.34 nm can be observed in the HRTEM image (Fig. 1F), corresponding to (002) plane of the graphitic stacking, which indicates the partial graphitization of the prepared carbon sample.⁶²

As shown in Fig. 2A, XRD pattern of N, O-HPC shows two broad diffraction peaks at 15–26° and 40–47° ranges, consistent with the graphitic carbon (JCPDS card file no. 75–1621), suggesting a turbostratic structure of carbon with low crystallinity. The broad diffraction peak centered at ~23° can be assigned to the (002) plane spacing corresponding to the interlayer spacing of the graphitic carbon, and the peak centered at ~43° can be attributed to a superposition of the (100) and (101) reflections of the interlayer condensation with a limited degree of graphitization.^{63,64} Moreover, it is noticed that the pattern presented a significantly enhanced intensity in its small-angle region ($2\theta < 15^\circ$), indicating the presence of high-density micropores,⁶⁵ thus confirming the results of TEM studies.

The Raman spectrum (Fig. 2B) exhibits two prominent first-order band peaks and two weak second-order band peaks. The peak centered at 1350 cm⁻¹ (D band) is commonly assigned to the disordered or defective graphitic structure with the breathing mode of *k*-point phonons of A_{1g} symmetry;⁶⁶ while the peak located at 1590 cm⁻¹ (G band) is associated with the ordered graphitic structure with an E_{2g} mode of graphite.⁶⁷ As known, the D band should be originated from the vibration of carbon atoms with dangling bonds in planar terminations of disordered graphite; while the G band should stem from the tangential vibration of sp²-bonded carbon atoms in a two-dimensional hexagonal lattice. The D band implies the appearance of structural disorder in the carbon specimen owing to the porous structure and heteroatom doping. It is generally recognized that the intensity ratio of D to G band (I_D/I_G) can be used to evaluate the degree of structural disorder in the carbon materials. Herein, the I_D/I_G ratio of the prepared carbon can be calculated to be about 1.01, indicating the existence of abundant disordered carbon structures. Moreover, the full width at half maximum of the D band is distinctly wider than that of G band, revealing a relatively high proportion of defects and disorder in the resultant carbon matrix, which is in good accordance with the XRD result. Additionally, a large graphitic domain size of 4.31 nm can be calculated by using the formula $L_a = 4.35 \times (I_G/I_D)$.⁶⁸ The larger domain size is usually advantageous for the higher surface reactivity of the carbon materials, and will promote the adsorption process by providing more unsaturated carbon atoms on the surface. Furthermore, the widening and weakening of the 2D band at around 2690 cm⁻¹ and the combination band of G + D at about 2940 cm⁻¹ also can be observed. The 2D band and (G + D) band are ascribed to the two-phonon process activated by double resonance and the combination of phonons with different momentum around the Γ and *K* points, respectively.^{67,69} These observations further demonstrate that the as-prepared carbon possesses a partial graphitization structure and a mass of defects (heteroatom co-doping and porosity).

To better understand the porous structure of N, O-HPC, N₂ adsorption–desorption and pore size distribution were performed as shown in Fig. 2C and D. As can be seen, the isotherm shows a combined Type I and IV with H4 hysteresis loop according to the International Union of Pure and Applied Chemistry (IUPAC) classification, indicating the coexistence of micropores and mesopores/macropores. The dramatic increase in the isotherm at a relatively low pressure ($P/P_0 < 0.2$) can be attributed to the presence of high-density micropores. At medium and high relative pressure regions ($P/P_0 = 0.4–1.0$), a small hysteresis loop associated with a capillary condensation step indicates the presence of mesopores and macropores. The specific surface area (SSA) is calculated to be 2270.19 m² g⁻¹, demonstrating a good activation effect caused by KOH. As well-demonstrated in the literature,^{70,71} the activation mechanism for the resulting N, O-HPC with KOH might be explained as follows: (1) the etching of carbon by KOH and the resulting K₂CO₃/K₂O through redox reactions ($6\text{KOH} + 2\text{C} \rightarrow 2\text{K} + 3\text{H}_2 +$



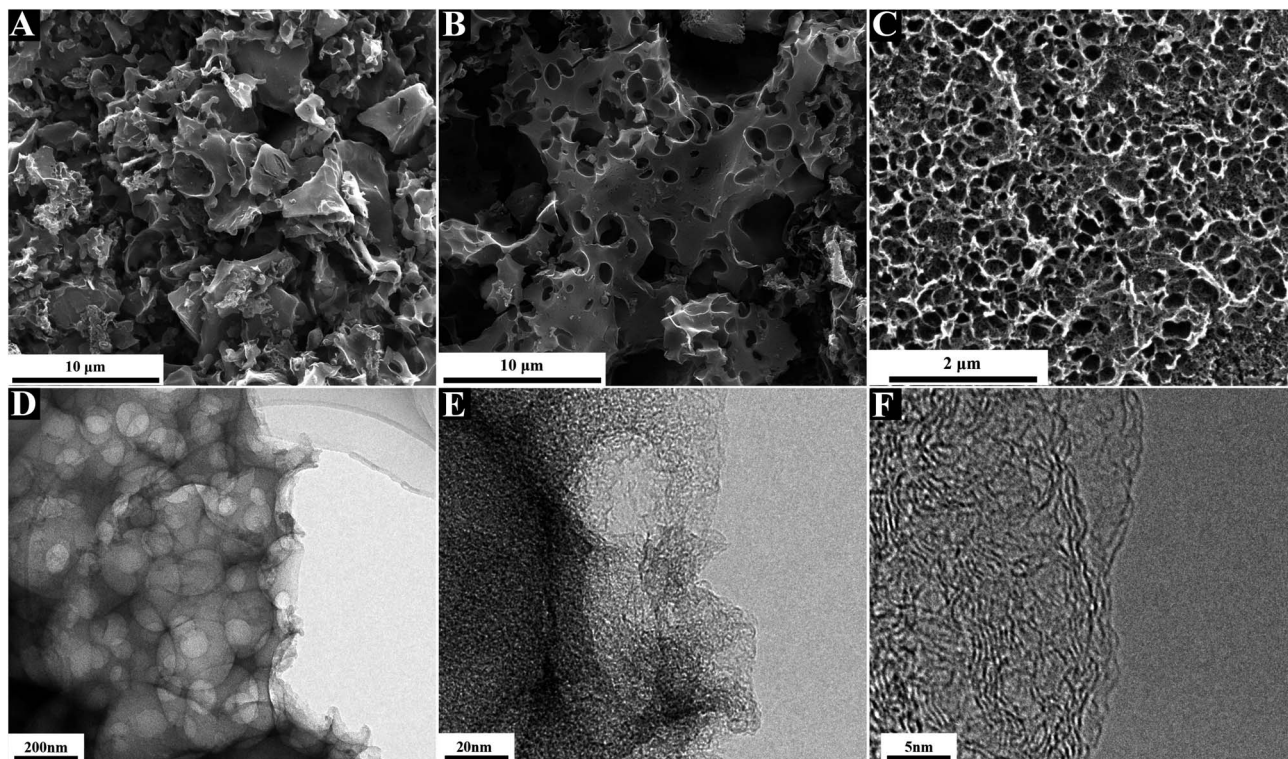


Fig. 1 (A–C) SEM, (D) TEM and (E and F) HRTEM images of N, O-HPC.

$2K_2CO_3$; $K_2CO_3 \rightarrow K_2O + CO_2$; $K_2CO_3 + 2C \rightarrow 2K + 3CO$; $K_2O + C \rightarrow 2K + CO$); (2) the gasification of carbon by the formed water vapor and CO_2 ($H_2O + C \rightarrow CO + H_2$; $CO_2 + C \rightarrow 2CO$); and (3) the intercalation of the intermediate K into the carbon framework. Furthermore, the plot of pore size distribution shows that the sample primarily composes of mesopores with the pore size ranging from 3.5 to 5 nm. The high SSA and well-developed hierarchically porous structure of N, O-HPC make it promising as ideal absorbent material for dye removal from coloured wastewater.

The chemical composition related to the surface functional groups of the prepared carbon materials were conducted by XPS measurement. As shown in Fig. 3A, the XPS survey scan of the sample showing obvious peaks of C 1s, N 1s, and O 1s, indicating the successful synthesis of the *in situ* N, O-codoped hierarchically porous carbon. 86.93 wt% C, 1.71 wt% N and 11.36 wt% O were detected in the N, O-HPC sample. The high-

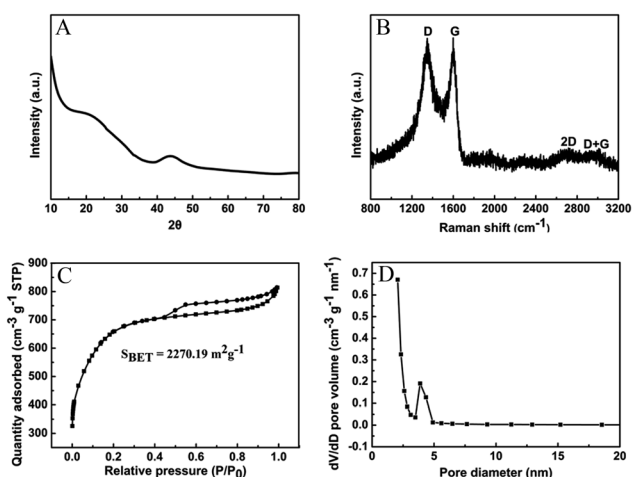


Fig. 2 (A) XRD pattern, (B) Raman spectrum, (C) N_2 adsorption-desorption isotherm and (D) pore size distribution of N, O-HPC.

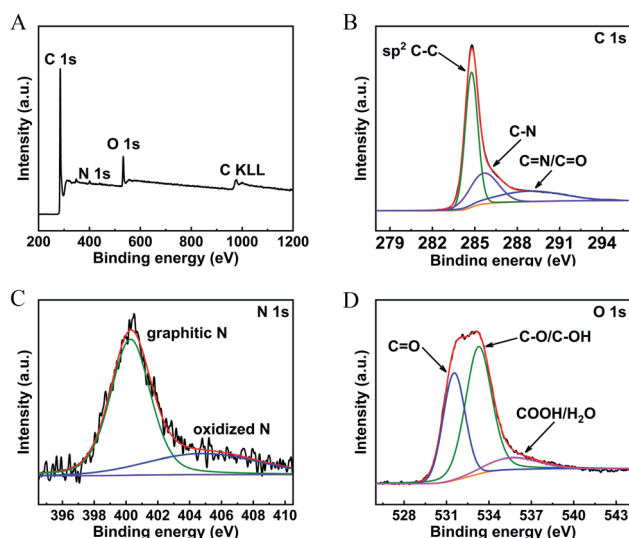


Fig. 3 (A) XPS survey spectrum and the high-resolution spectra for (B) C 1s, (C) N 1s and (D) O 1s together with the fitting peaks of N, O-HPC.



resolution C 1s spectrum (Fig. 3B) can be divided into three individual peaks, centered at about 284.8, 285.7, and 288.9 eV, corresponding to C–C, C–N and C=O/C=N binding.^{72,73} The high-resolution N 1s spectra shown in Fig. 3C can be deconvoluted into two peaks, centered at about 400.3 and 404.7 eV, which can be attributed to characteristic nitrogen functionalities of graphitic nitrogen and oxidized nitrogen.^{74,75} The high-resolution O 1s spectrum (Fig. 3D) revealed the presence of three oxygen-based components, which can be resolved into three peaks at 531.5 eV (keto oxygen or quinone group, O-I), 533.3 eV (esters or phenolic group, O-II) and 535.5 eV (carboxylic group or water, O-III), respectively.⁷⁶

3.2 Effect of contact time and adsorption kinetics

The effect of contact time on the adsorption capacities of N, O-HPC for MB was investigated. As shown in Fig. S1† there was a rapid increase in MB adsorbed amount in the first 15 min, which could be ascribed to the presence of numerous adsorption active sites on the N, O-HPC surface. During the following period, the adsorption rate slowed down, which was arisen from: (1) a reduction in adsorption active sites present on the adsorbent surface; (2) occupancy of the pores in the adsorbent by MB molecules, leading to slower internal diffusion.⁷⁷ After 45 min, adsorption equilibrium was reached gradually, indicating that the adsorbents owned high adsorption efficiency towards MB.

The Lagergren pseudo-first order, pseudo-second order and intra-particle diffusion models were employed to describe the kinetics of MB adsorption. As shown in Fig. 4, it can be found that the regression coefficient R^2 obtained from the pseudo-second order model (0.9998) is much higher than those from other models, suggesting that the pseudo-second order model fits better with the experimental data. Furthermore, Fig. 4C shows the multilinear plots of intraparticle

diffusion process of MB adsorption onto N, O-HPC. It suggests that two steps have taken place during the adsorption process. The first region represents the film diffusion stage which is the diffusion of MB molecules from solution to the external surface of adsorbent. And the second region can be ascribed to the intraparticle diffusion stage due to the rough surface and abundant pores of adsorbent.⁷⁸ The slope of the film diffusion stage is obviously larger than intraparticle diffusion stage, indicating the intraparticle diffusion stage is a gradual process. Moreover, the plot does not pass through the origin indicating the intraparticle diffusion is not the rate-limiting step and boundary layer phenomenon could also direct the adsorption process.

3.3 Effect of initial MB concentration and adsorption isotherms

Fig. S2A† is the time dependence of MB adsorption capacities on N, O-HPC at various initial concentrations. As shown, despite the different initial concentrations, the MB uptake was very rapid and efficient, and adsorption capacity of MB onto the adsorbent drastically increase during the initial stage and then at a slow speed until reach a state of equilibrium. The results can be ascribed to the fact that most vacant adsorption active sites are available for adsorption during the initial stage and the remaining vacant surface sites are hard to be utilized due to repulsive forces between the MB molecules on adsorbent and the bulk phase. The equilibrium adsorption capacities of MB at various initial concentrations demonstrate an increasing trend, since initial MB concentrations can provide a driving force to overcome the mass transfer resistance of the dye, which indicates that the adsorption process is highly dependent on initial MB concentration. Furthermore, despite various initial concentrations, all the adsorption kinetics can be well fitted by the pseudo-second order model with regression coefficient R^2 all higher than 0.9996 (Fig. S2B†), which is consistent with the former adsorption kinetic study.

The equilibrium adsorption isotherm was studied in detail, since it could provide information about the surface properties of adsorbent and the adsorption behaviour. In this study, the equilibrium adsorption data is analysed by Langmuir, Freundlich and Temkin models. As shown in Fig. 5, the regression coefficient R^2 obtained from the Langmuir isotherm model is much higher than those from other models, suggesting that Langmuir isotherm model can describe the experimental data much more suitably, which also indicating the presence of homogeneously adsorption active sites on the surface of N, O-HPC. Moreover, the maximum monolayer adsorbent capacity value of 2104.29 mg g⁻¹ predicted by the Langmuir model is among the highest in the previously reported adsorbents (Table S1†), which proves N, O-HPC as an effective adsorbent for MB uptake from aqueous solution. Furthermore, the feasibility of the adsorption process was evaluated using the dimensionless separation factor, the value of RL was found to be in the range of 0.005–0.013, which showed that the adsorption of MB onto N, O-HPC was favourable.

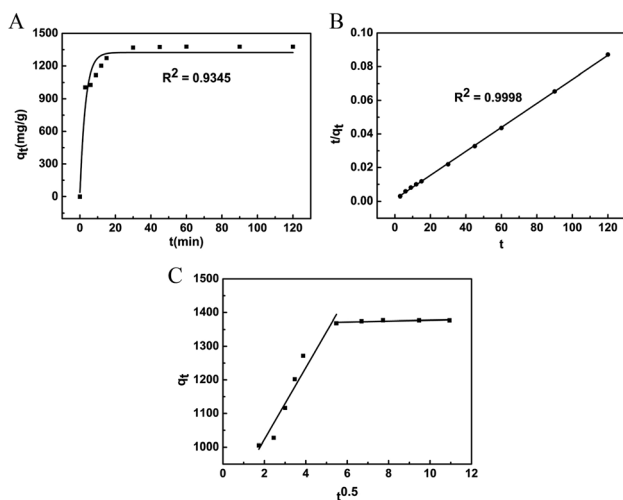


Fig. 4 The fitting of different kinetic models for MB adsorption onto N, O-HPC. (A) Lagergren pseudo-first order, (B) pseudo-second order, (C) intra-particle diffusion.



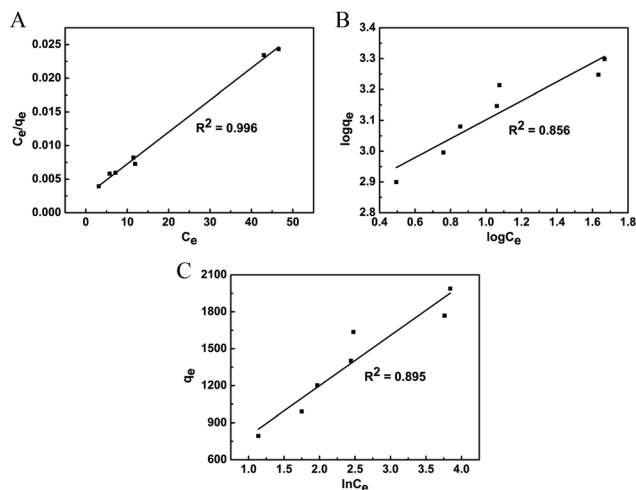


Fig. 5 The fitting of different isotherm models for MB adsorption onto N, O-HPC. (A) Langmuir, (B) Freundlich, (C) Temkin.

3.4 Effect of temperature and adsorption thermodynamics

Fig. S3A† shows the time dependence of MB adsorption capacities on N, O-HPC at different temperatures. During the first 30 min, the adsorption capacity presents a significant increasing trend with the rise in temperature. The equilibrium adsorption capacity increases from 1368 to 1381 mg g⁻¹ as the temperature increases from 293.15 K to 313.15 K. The results suggest that MB adsorption on the adsorbent is favoured at higher temperatures within the appropriate temperature range, which may be ascribed to that the mobility of MB molecules increases with the rise in temperature and more dye molecules can interact with the active sites of N, O-HPC.

Thermodynamics parameters including enthalpy change (ΔH^0), Gibbs free energy change (ΔG^0) and entropy change (ΔS^0), enable proper understanding of the inherent energetic changes occurring during adsorption and show whether the reaction can happen spontaneously or not. ΔG^0 was determined by the Gibbs isotherm equation ($\Delta G^0 = -RT \ln K_d$), while ΔH^0 and ΔS^0 were calculated from the slope and intercept of the Van't Hoff plot (Fig. S3B†). The calculated values were presented in Table S2.† The negative values of ΔG^0 confirmed the spontaneous nature and the feasibility of the adsorption process. Besides, the decrease in the negative value of ΔG^0 with an increase in temperature suggested the adsorption is more favourable at higher temperature. The positive value of ΔH^0 indicated this adsorption process is endothermic, which is consistent with the effect of temperature. The positive value of ΔS^0 implied good affinity of MB molecules towards the adsorbent and an increasing randomness at the solid–solution interface during the adsorption process increases. All the thermodynamic parameters mentioned above indicate that N, O-HPC can be used as a high-efficiency adsorbent to remove MB from aqueous solution.

3.5 Effect of adsorbent dosage

Fig. S4A† shows the time dependence of MB adsorption capacities on N, O-HPC at different adsorbent dosage. As

shown, there was a rapid adsorption capacity increase with increasing adsorbent mass dosage from 10 mg to 50 mg and afterwards almost reached a saturation level at higher dosage. This resulted from the creation of additional surface area and adsorption active sites as the adsorbent dosage increases, which enhances MB uptake. Furthermore, all the adsorption kinetics can also be well fitted by the pseudo-second order model with regression coefficient R^2 all higher than 0.9992 (Fig. S4B†), which is also consistent with the former adsorption kinetic study.

3.6 Effect of humic acid (HA) and ionic strength

HA is a common model compound of natural organic matter and widely exists in the aquatic environment.⁷⁹ Effect of HA on MB adsorption was studied (Fig. 6A and B). Negligible impact can be observed when the HA concentrations below 20 mg L⁻¹, while higher HA concentrations gave a negative effect on the adsorption process. The adsorption rate decreased and a much longer time was needed to reach the adsorption equilibrium stage as the HA concentration increased from 20 to 160 mg L⁻¹. Moreover, the remove efficiency decreased from 97.69% to 91.48% as the HA concentration increased from 20 to 160 mg L⁻¹. The reduction of adsorption rate and remove efficiency could be resulted from the competitive adsorption between HA and MB molecules toward the adsorbent.⁸⁰ Actually, the HA concentration in realistic water environment was much lower than that of in this study, implying that N, O-HPC exhibited excellent anti-interference ability and had potential to be used in real water matrix on MB removal.

The impact of ionic strength was investigated using NaCl with different concentrations. Generally, ionic strength has two potential impacts in organic contaminants adsorption by carbonaceous adsorbents. On the one hand, increasing ionic strength can enhance the activity coefficient of hydrophobic

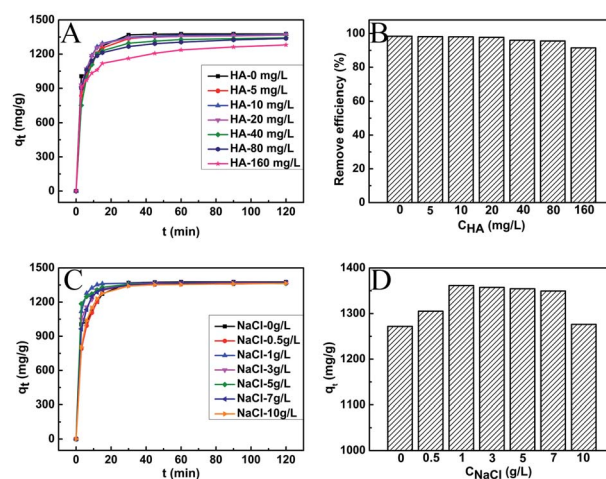


Fig. 6 Effect of HA concentration and ionic strength on (A and C) the time-dependent adsorption capacity, (B) the equilibrium remove efficiency and (D) the remove efficiency at 15 min of MB onto N, O-HPC ([MB] = 700 mg L⁻¹, adsorbent dosage = 0.05 g, 293 K).



organic compounds, leading to a decrease in their solubility (*i.e.* salting out effect), which is favourable for organic contaminants adsorption.⁸¹ Besides, the increase of ionic strength give rise to the compression of the diffuse double layer on the adsorbent, which facilitated the electrostatic attraction and contributed to the adsorption consequently.⁸² On the other hand, the increase in salt ions can exert a “squeeze-out” effect on carbonaceous adsorbents, which is unfavourable for organic contaminants adsorption.^{83,84} As shown in Fig. 6C, the equilibrium adsorption capacities basically kept unchanged even the salt concentrations were up to 10 g L⁻¹. However, the variation of NaCl concentration exhibited a significant effect on the adsorption rate and high salt concentration was contributive to dye removal, especially in the first 15 min (Fig. 6D). This facilitation for MB adsorption could be resulted from both the salting-out effect and the enhanced electrostatic attraction.

3.7 Effect of initial solution pH

The pH is one of the most significant environmental factors influencing the adsorption process. Solution pH would influence both aqueous chemistry and surface binding sites of the adsorbents. The pHPzc is very crucial to judge the surface charge of adsorbents. The surface of the adsorbent is negatively charged at pH > pHPzc, or it is positively charged at pH < pHPzc. The pHPzc of N, O-HPC was determined by the classical immersion technique and found to be 7.3 (Fig. 7A). To study the influence of initial solution pH on the adsorption capacities of N, O-HPC for MB, experiments were carried out using various pHs varying from 1 to 9. From Fig. 7B, it is observed that the adsorption was highly dependent on solution pH, and the dye adsorption capacities increased with the rise of solution pH. At low pH (≤ 7), the dye adsorption capacities decreased in the order: 1 > 3 > 7, which was ascribed to the surface of the adsorbent was positively charged, which generated repulsive forces between MB⁺ cations and positively charged adsorption sites triggering decline in dye removal. Moreover, the competition between excess amount of H⁺ ions and MB⁺ cations for negatively charged adsorption sites also reduced the adsorption capacities. At pH = 9, the negatively charged adsorption sites on adsorbents were enriched causing high electrostatic attraction between MB⁺ and adsorbent, leading to enhanced adsorption capacities. Similar results can be found in previous literatures.

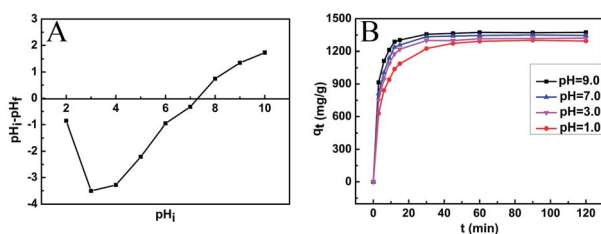


Fig. 7 (A) pHpzc of N, O-HPC, (B) Effect of initial solution pH on the time-dependent adsorption capacity of MB onto N, O-HPC ([MB] = 700 mg L⁻¹, adsorbent dosage = 0.05 g, 293 K).

3.8 Adsorption mechanism

Fourier transform infrared (FTIR) spectra of N, O-HPC, MB adsorbed N, O-HPC and pure MB were performed to gain more insights into the adsorption mechanism. FTIR spectrum of N, O-HPC (Fig. 8A) confirmed the presence of –OH (3435 cm⁻¹), –C=C– (1635 cm⁻¹), C=O/C=N (1589 cm⁻¹) and C–N/C–O (1120 cm⁻¹), which led to the hydrophilic nature of N, O-HPC and acted as anchoring sites for MB molecules. In the case of MB (Fig. 8C), the peak at 1597 cm⁻¹ was assigned to the stretching vibrations of C=N (and C=C) bond in the heterocycle of MB, while the peaks at 1350 and 1329 cm⁻¹ were attributed to the stretching vibration of C–N bond connected with benzene ring and N–CH₃ bond. The peak at about 1487 cm⁻¹ was ascribed to the CH₂ deformation vibration, while the peak around 1388 cm⁻¹ was owing to the CH₃ deformation vibration. Furthermore, the peaks at about 1246 and 1217 cm⁻¹ were due to Ar–N deformation vibration, the peaks at 1168 and 1136 cm⁻¹ to the stretching vibrations of C=S and C–S, and the band at about 884 cm⁻¹ to the wagging vibration of C–H in aromatic ring of MB. After adsorption of MB (Fig. 8B), the peak at 3435 cm⁻¹ of N, O-HPC remained unchanged, which indicated that hydrogen bonding may not be an important process for MB adsorption onto N, O-HPC. However, the new absorption peak at 1383 and 880 cm⁻¹ could be clearly observed, indicating that MB has been anchored on the surface of N, O-HPC during the adsorption process. Moreover, after adsorbed by N, O-HPC, obvious shifts were observed to 1593 cm⁻¹ for C=N (and C=C) bond in the heterocycle of MB, to 1455 cm⁻¹ for CH₂ deformation vibration in benzene ring and to 880 cm⁻¹ for the wagging vibration of C–H in aromatic ring of MB, which corresponded to the attachment of MB on the surface of N, O-HPC by π - π stacking interaction between the aromatic backbone of MB and hexagonal skeleton of N, O-HPC since MB was an ideally planar molecule.^{85,86} Besides, the electronegativity of N is

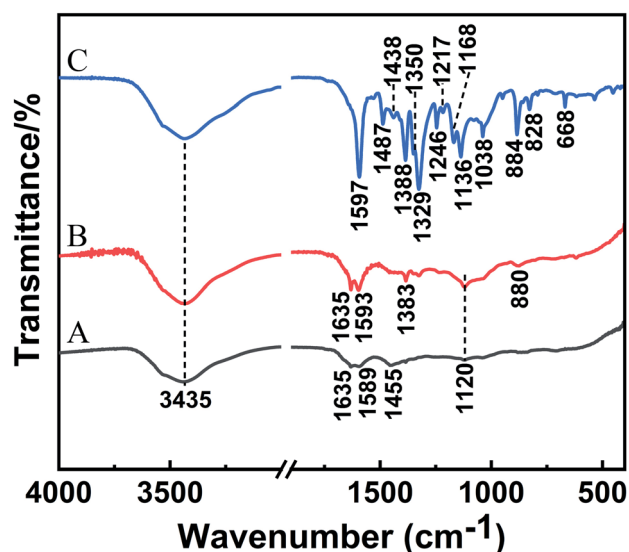


Fig. 8 FTIR spectra of N, O-HPC before (A) and after (B) adsorption of MB and pure MB (C).



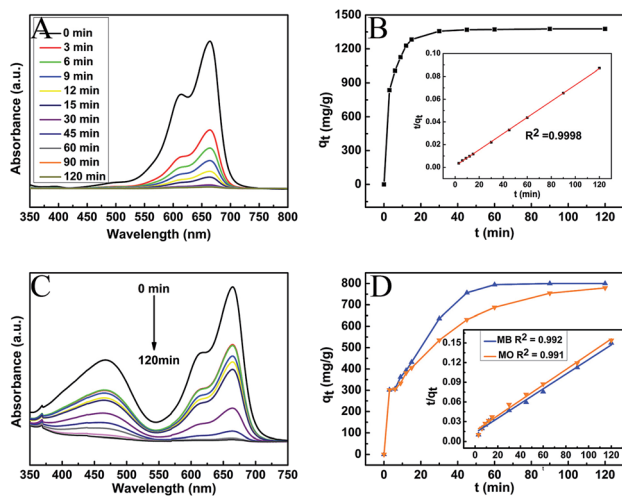


Fig. 9 (A) Time-dependent UV-vis absorption spectra of MB solutions, (B) effect of contact time on MB adsorption capacities. Inset: corresponding pseudo-second order model fit ($[MB] = 700 \text{ mg L}^{-1}$, adsorbent dosage = 0.05 g, 293 K, solvent: tap water), (C) time-dependent UV-vis absorption spectra of MB and MO solutions, (D) effect of contact time on MB and MO adsorption capacities and their corresponding pseudo-second order kinetic model fittings ($[MB] = [MO] = 400 \text{ mg L}^{-1}$, adsorbent dosage = 0.05 g, 293 K).

larger than that of C ($\chi_N(3.04) > \chi_C(2.55)$), the electronegative graphitic N atoms could positively polarize their neighbouring carbon atoms, thus the corresponding negatively charged graphitic N atoms were able to act as adsorption sites to absorb cationic MB *via* electrostatic attraction. Based on the above, the exceptionally high uptake of MB onto N, O-HPC resulted from not only high specific surface area and abundant mesopores/micropores of N, O-HPC, but also multiple adsorption mechanisms including pore filling, π - π stacking interaction and electrostatic interaction between MB and N, O-HPC.

3.9 Adsorption of MB in tap water and synergistic adsorption of MB and MO

As indicated in Fig. 9A and B, despite in tap water solution, the maximum adsorption peak at 664 nm of MB gradually decreased with the increase of time and the adsorption process reached equilibrium state at about 30 min. Furthermore, the as prepared adsorbent showed synergistic adsorption performance for MB and methyl orange (MO). As shown in Fig. 9C, the maximum adsorption peak at 464 nm of MO and at 664 nm of MB both gradually decreased as the rise of time. Furthermore, all the adsorption kinetics can also be well fitted by the pseudo-second order model with regression coefficient R^2 all higher than 0.99 (Fig. 9D). These results indicated N, O-HPC could be considered as efficient, novel and advantageous adsorbent for the removal of both cationic and anionic dyes.

4 Conclusions

In summary, N, O-codoped hierarchical porous carbon (N, O-HPC) derived from silkworm cocoon with an ultrahigh

specific surface area of $2270.19 \text{ m}^2 \text{ g}^{-1}$ was successfully synthesized *via* a facile pre-carbonization and chemical activation method, and characterized thoroughly. The as-prepared N, O-HPC proved to be a high-efficiency adsorbent. Batch adsorption experimental results demonstrated that the MB adsorption was highly dependent on contact time, initial MB concentration, temperature and initial solution pH, while no remarkable effects of humic acid and ionic strength are observed. Kinetics experiments clearly indicated the adsorption process followed pseudo-second-order kinetic model. Besides, Weber's intraparticle diffusion model indicated two steps have taken place during the adsorption process and the intraparticle diffusion was not the rate-limiting step. The adsorption isotherm study showed that Langmuir isotherm model can describe the experimental data much more suitably with a maximum monolayer adsorption capacity value of $2104.29 \text{ mg g}^{-1}$, which is among the highest in the previously reported adsorbents and ascribed to multiple adsorption mechanisms including pore filling, π - π stacking interaction and electrostatic interaction between MB and N, O-HPC. By thermodynamic analysis, it was shown that MB adsorption onto N, O-HPC was spontaneous and endothermic. Furthermore, the as prepared adsorbent showed highly efficient adsorption for MB in tap water and synergistic adsorption performance toward MB and MO. Therefore, N, O-HPC derived from silkworm cocoon can be developed as an efficient, novel and advantageous material for wastewater remediation.

Conflicts of interest

We have no competing interests.

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