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A salt-based method to adapt stiffness and biodegradability of porous collagen scaffolds†

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Type I collagen scaffolds for tissue reconstruction often have impaired mechanical characteristics such as limited stiffness and lack of strength. In this study, a new technique is presented to fine-tune stiffness and biodegradability of collagen scaffolds by treatment with concentrated salt solutions. Collagen scaffolds were prepared by a casting, freezing and lyophilization process. Scaffolds were treated with 90% saturated salt solutions, the salts taken from the Hofmeister series, followed by chemical crosslinking. Treatment with salts consisting of a divalent cation in combination with a monovalent anion, e.g. CaCl₂, resulted in fast shrinkage of the scaffolds up to approximately 10% of the original surface area. Effective salts were mostly at the chaotropic end of the Hofmeister series. Shrunken scaffolds were more than 10 times stiffer than non-shrunken control scaffolds, and displayed reduced pore sizes and swollen, less organized collagen fibrils. The effect could be pinpointed to the level of individual collagen molecules and indicates the shrinking effect to be driven by disruption of stabilizing hydrogen bonds within the triple helix. No calcium deposits remained in CaCl₂ treated scaffolds. Subcutaneous implantation in rats showed similar biocompatibility compared to H₂O and NaCl treated scaffolds, but reduced cellular influx and increased structural integrity without signs of major degradation after 3 months. In conclusion, high concentrations of chaotropic salts can be used to adjust the mechanical characteristics of collagen scaffolds without affecting biocompatibility. This technique may be used in regenerative medicine to stiffen collagen scaffolds to better comply with the surrounding tissues, but may also be applied for e.g. slow release drug delivery systems.

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Introduction

Type I collagen has been extensively used in the field of tissue engineering and regenerative medicine, particularly as a scaffolding material and as a controlled release system.¹ Its excellent biocompatibility, widespread availability and straightforward processability make collagen a very useful biomaterial² that has been used in many (pre)clinical studies to

repair tissues like skin,³ cartilage⁴ and parts of the urogenital system.⁵ However, type I collagen scaffolds often show poor mechanical characteristics in applications where the scaffold is exposed to mechanical stress.⁶ When large defects are to be repaired, e.g. in diaphragmatic hernia, mechanical properties are generally insufficient.⁷

To strengthen collagen scaffolds, techniques such as chemical crosslinking using e.g. carbodiimide or glutaraldehyde are applied.^{8,9} Alternatively, components can be added to the collagen in order to enhance mechanical properties. Typically, synthetic polymers such as PLLA,¹⁰ PCL¹¹ or PGA¹² have been used in this respect. However, addition of other components generally complicates the preparation process, may impair biocompatibility, and imposes additional regulatory constraints.

Increasing the concentration of insoluble type I collagen in a porous scaffold is another way to enhance mechanical stiffness.¹³ However, this procedure is poorly controllable for insoluble fibrillar collagen and maximum concentrations correspond to about 2% (w/v) using current preparation techniques, which may still be insufficient for stiffer tissues.

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In this study, a method is presented that enables the construction of microporous collagen scaffolds with defined stiffness, that are biocompatible and display limited *in vivo* degradation for at least 3 months. The method is based on fast and extensive shrinkage of collagen scaffolds by treatment with highly concentrated salt solutions selected from the Hofmeister series of salts, a series based on the ability of anions and cations to alter protein stability, and to influence surface tension and solubility.¹⁴ The method may expand the applications of collagen-based constructs in the regeneration of tissues with defined stiffness.

Experimental section

Materials

N-Hydroxysuccinimide (NHS) was obtained from Fluka Chemie AG, Buchs, Switzerland. Acetic acid, ammonium chloride (NH₄Cl), ammonium sulfate ((NH₄)₂SO₄), bovine serum albumin, 1-ethyl-3-(3-dimethylaminopropyl)carbodiimide (EDC), ethanol, glutaraldehyde, hydrogen chloride, nitric acid, paraformaldehyde, sodium chloride (NaCl), sodium hydrogen phosphate (Na₂HPO₄) and sodium sulfate (Na₂SO₄) were from Merck, Darmstadt, Germany. Glycine was from Scharlau Chemicals, Barcelona, Spain. Barium perchlorate (Ba(ClO₄)₂), barium sulfate (BaSO₄), calcium chloride (CaCl₂), calcium sulfate (CaSO₄), magnesium chloride (MgCl₂), magnesium perchlorate (Mg(ClO₄)₂), 2-(*N*-morpholino)ethane sulfonic acid (MES) and trinitrobenzene sulfonic acid (TNBS) were from Sigma Aldrich, St. Louis, MO, USA. Barium chloride (BaCl₂) and magnesium sulfate (MgSO₄) were from VWR International, Radnor, PA, USA.

Preparation of shrunken collagen scaffolds

(1) **Porous scaffold preparation.** Type I collagen fibrils were purified from bovine achilles tendon using extractions steps with aqueous solutions of NaCl, urea, diluted acetic acid, acetone and demineralized water.¹⁵ Purified collagen was suspended in 0.25 M acetic acid (0.8% w/v), swollen overnight, and subsequently homogenized using a Silverson L5M-A laboratory mixer (Silverson, Chesham, UK) by mixing for 6 min at 2500 rpm (3 min with general purpose disintegrating work head and 3 min with a slotted work head). Next, the suspension was deaerated using centrifugation at 100g for 30 min. The suspension was poured into 12-well plates (1.5 mL) and 6-well plates (4 mL), frozen for 4 h at −20 °C on an aluminum surface for optimal thermal conduction and lyophilized (Zirbus sublimator 500II, Bad Grund, Germany).

(2) **Shrinkage of scaffolds by Hofmeister series of salt treatment.** To analyze the effect of different concentrated salt solutions, non-crosslinked porous scaffolds (∅ 22 mm) were exposed to 4 mL of 90% saturated salt solutions in demineralized water. Twelve salts were selected from the entire spectrum of the Hofmeister series (see ESI Table 1†). To enhance penetration of the salt solution into the scaffold, the surface tension of 4 mL 90% saturated salt solutions was reduced by the addition of 10 μL 100% (v/v) ethanol. As a control, scaffolds were

incubated in demineralized water with 10 μL ethanol. After 48 h, the surface of the scaffolds was scanned (Epson V750 PRO) and the surface area was measured using image analysis software (ImageJ version 1.47i, NIH, Bethesda, MD, USA). For the salts that showed high shrinkage (more than 90% compared to original surface), the shrinkage was determined at different concentrations. All experiments were performed three times independently.

(3) **Shrinkage process for CaCl₂.** We focused on CaCl₂, as it proved to be one of the most effective salts (see Results section). As controls, scaffolds were treated with NaCl (salt treatment, but no shrinkage) and demineralized water (H₂O). Scaffolds were shrunken in an ascending series of CaCl₂ to achieve gradual shrinkage. Briefly, they were first incubated for 15 min in 1 M CaCl₂ or 1 M NaCl, followed by 15 min 2 M, and 30 min 4 M salt solution. Control scaffolds were incubated for 30 min in demineralized water. Macroscopic images were taken and videos of the shrinking process with CaCl₂ were recorded with a Sony Cyber-shot DSC-H10 (Sony, Minato, Tokyo, Japan).

(4) **Stabilization of scaffolds by chemical crosslinking.** After incubation in 4 M salt solution or demineralized water, scaffolds were immediately crosslinked using a carbodiimide zero-length crosslinker. In brief, scaffolds were incubated in 4 mL 50 mM MES buffer, pH 5.0, containing 40% (v/v) ethanol, 33 mM EDC and 6 mM NHS for 3 h. Next, the scaffolds were extensively washed 6 times with demineralized water to remove residual crosslinking reagents and salt. Non-crosslinked scaffolds were also washed 6 times with demineralized water. Finally, scaffolds were placed in 70% (v/v) ethanol and stored at −20 °C until use.

(5) **Sterilization of the scaffolds.** Scaffolds used for *in vivo* evaluation were sterilized in PBS using gamma irradiation at 25 kGy from a ⁶⁰Co source (ISO 9001; Synergy Health BV, Ede, The Netherlands). Before sterilization, ∅ 6 mm discs were cut from the scaffolds after which they were washed 6 times with PBS to remove ethanol.

Scaffold characterization

(1) **Scanning electron microscopy.** For scanning electron microscopy, crosslinked scaffolds were washed in demineralized water to remove storage medium, and subsequently frozen at −80 °C. The frozen scaffolds were lyophilized, mounted on stubs with double-sided carbon tape, and coated with an ultrathin gold layer (Scancoat Six Sputter Coater, Edwards, Crawley, United Kingdom). The samples were evaluated with a Sigma 300 field emission scanning electron microscope (Carl Zeiss B. V., Sliedrecht, The Netherlands) with an accelerating voltage of 10 kV. The pore sizes of the different scaffolds were determined using ImageJ analysis software.

(2) **Transmission electron microscopy.** Scaffolds were fixed in 2% (w/v) glutaraldehyde in 0.1 M phosphate buffer (PB), pH 7.3 for 24 h at 4 °C. After washing steps in PB, scaffolds were treated with 1% osmium tetroxide in 0.1 M PB for 1 h. Next, scaffolds were washed in PB, dehydrated in graded series of alcohol and embedded in Epon 812. Sections (1 μm) were cut, stained with toluidine blue and examined with a light-



microscope (Dialux 20, Leitz). Ultrathin sections 60 nm were cut, followed by treatment with uranyl-acetate/lead-citrate for double contrast. Sections were imaged using a transmission electron microscope (JEOL 1010).

(3) Determination of stiffness. The stiffness of CaCl₂, NaCl and H₂O treated collagen scaffolds were measured with the PIUMA nano-indenter (Optics11, Amsterdam, The Netherlands; Fig. 3A).¹⁶ Before the measurement, scaffolds were incubated in PBS for 30 min followed by a passivation step of 60 min in 5% bovine serum albumin in PBS to prevent adhesion of the probe to the scaffolds. Per sample, 25 indentations were performed in a grid of 1 × 1 mm with a distance of 200 μm between individual indentations (Fig. 3B). For soft samples, an indenter probe with a stiffness of 0.05 N m⁻¹ was used and for stiffer samples a probe with a stiffness of 0.47 N m⁻¹. Both probes had a tip diameter of 180 μm. The indentation depth, reached with a speed of 5 μm s⁻¹, was 15 μm and the probe was kept in place for 2 s (indentation time). Each scaffold type was measured 3 times independently.

(4) Calcium analysis of CaCl₂ treated scaffolds. Two distinct methodologies were used to determine potential remnants of the CaCl₂ shrinking treatment. Inductively coupled plasma mass spectrometry (ICP-MS) was performed to detect elemental calcium, and Fourier transform infrared spectroscopy (FTIR) was performed to detect the presence of calcium-containing compounds which may have been formed after CaCl₂ treatment. For the ICP-MS measurement, scaffolds were first hydrolyzed with 50% (v/v) nitric acid and subsequently diluted to a final concentration of 0.05 mg mL⁻¹ in 1% nitric acid. Samples were analyzed using a X series I – ICP-MS (Thermo Fisher Scientific, Waltham, MA, USA). The total calcium content was calculated based on the measurement of the ⁴³Ca isotope as ⁴⁰Ca interfered with other elements. One-way ANOVA with Bonferroni *post-hoc* testing was performed to determine statistical differences. Fourier Transform Infrared Spectroscopy (FTIR, Spectrum Two, PerkinElmer, Waltham, MA, USA) was used to study the presence of calcium-containing compounds in the scaffolds. The FTIR spectra were measured in Attenuated Total Reflectance mode, with a spectral resolution of 4 cm⁻¹, a scan speed of 0.2 cm s⁻¹ and 10 scans per measurement.

(5) Differential Scanning Calorimetry. Differential scanning calorimetry (DSC Q1000, TA Instruments, New Castle, DE, USA) was used to assess potential denaturation of type I collagen caused by the salt treatment by measuring the denaturation temperature of non-crosslinked scaffolds treated with CaCl₂, NaCl and H₂O. Purified insoluble type I collagen (starting material for scaffold preparation) was used as a control. Samples (~1 mg) in aluminium pans with 25 μL PBS were heated from 1 to 80 °C with a rate of 5 °C min⁻¹. The endothermic peak was determined using TA Instruments Universal Analysis 2000 software (Version 4.5 A).

(6) Gel electrophoresis. Sodium dodecyl sulfate polyacrylamide gel electrophoresis (SDS-PAGE) was performed to assess potential denaturation of collagen in the scaffold after salt treatment. Briefly, 2–3 mg of non-crosslinked collagen scaffolds was digested overnight at ambient temperature in 75

μL 0.25 M acetic acid containing 3 mg mL⁻¹ pepsin after which an equal volume of a 2-fold concentrated non-reducing sample buffer was added. Samples and controls were loaded on an 8% (w/v) polyacrylamide gel without prior heating. After running at 150 V, the gel was stained with 0.1% (w/v) Coomassie Brilliant Blue solution.

(7) Evaluation of degree of crosslinking. The degree of crosslinking of the salt treated scaffolds was calculated by assaying the loss of primary amine groups after crosslinking using a TNBS assay.¹⁷ After salt treatment, crosslinking and rinsing steps with water, the scaffolds were lyophilized and subsequently incubated for 30 min in 4% (w/v) Na₂HPO₄ at 22 °C. Subsequently, samples were incubated for 2 h at 40 °C in 0.5% (w/v) TNBS in demineralized water, after which the samples were hydrolyzed with 6 M HCl for 1.5 h at 60 °C. Absorbance at 420 nm was measured with a spectrophotometer (Bio-Tek, Bad Friedrichshall, Germany). Glycine was used for the calibration curve. Non-crosslinked controls for each condition were used to calculate the percentage of crosslinked amine groups. Experiments were performed in triplicate in three different experiments.

In vivo evaluation of salt treated scaffolds in a rat model

The study was approved by the Ethical Committee on Animal Research of the Radboud University Nijmegen, The Netherlands (RU-DEC 2013-144) and executed according to Dutch legislation and guidelines for the use of laboratory animals. Nine male Wistar rats (Harlan Laboratories, Horst, The Netherlands) of 3 months old (~400 g) were housed in trios in a controlled environment (19–21 °C, 12–12 h day per night cycle) with *ad libitum* access to food and water. Using Microsoft Excel, the rats were randomly divided over three experimental groups. Rats were anesthetized with 2–3% (v/v) isoflurane, and received buprenorphine by intramuscular injection as analgesic. Three different scaffolds (a CaCl₂ treated scaffold, a NaCl treated scaffold, and a H₂O treated scaffold) were randomly implanted in subcutaneous pockets created on the back of each rat. At the time of implantation, researchers performing evaluations were blinded for the type of scaffold by naming them A, B and C. Reference sutures (Prolene 5.0, Ethicon, Somerville, NJ) were placed next to the scaffold. Wounds were sutured using Prolene 5.0 and for reinforcement of the sutures hog rings were placed over the wound. Immediately after surgery and 12, 24, 36 and 48 h later, rats received additional buprenorphine *via* intramuscular injection. During the whole experiments the general health condition of the rats was monitored. At 3 days (group 1), 14 days (group 2) and 90 days (group 3) after implantation the animals were sacrificed using carbon dioxide gas. Scaffolds were excised from the subcutaneous pockets, fixed using 4% (w/v) paraformaldehyde in phosphate buffer (pH 7.4) and embedded in paraffin. Sections (5 μm) were obtained using a HM 340E Electronic Rotary Microtome (Thermo Fisher Scientific, Waltham, MA, USA). Tissue sections were stained with haematoxylin/eosin and elastin Von Masson. To visualize calcifications, a von Kossa and an Alizarin red staining were performed. Blinding was applied for the animal experiments;



however, at time of evaluation the blinded feature of this study was lost due to the distinct characteristics of each scaffold type. Haematoxylin and eosin stained sections were scored semi quantitatively from 0 (not present) to 3 (abundantly present), where sporadic events were scored as 0.1. Statistical significance was determined using an Independent Samples Kruskal Wallis-test for comparison between scaffolds at the three time points. Differences were considered statistically significant when $p < 0.05$.

Results

To evaluate the effect of the Hofmeister salts on porous disc-shaped collagen scaffolds, scaffolds were incubated for 48 h in 90% saturated solutions. For a number of salts (e.g. CaCl_2), major shrinkage of the scaffolds was observed, whereas for other salts intermediate, minor or no shrinkage was noticed (Fig. 1). The surface area was shrunken for more than 90% in case of CaCl_2 ($93 \pm 1\%$, Fig. 1A), $\text{Ba}(\text{ClO}_4)_2$ ($91 \pm 1\%$), MgCl_2 ($93 \pm 1\%$), and $\text{Mg}(\text{ClO}_4)_2$ ($94 \pm 1\%$) (Fig. 1B). The process was

extremely fast: within 30 s the scaffold had almost completely shrunken (Movie 1†). Salts that caused only limited shrinkage were MgSO_4 ($8 \pm 11\%$), Na_2SO_4 ($13 \pm 7\%$), NaCl ($15 \pm 8\%$), NH_4Cl ($20 \pm 9\%$), CaSO_4 ($28 \pm 14\%$), and BaCl_2 ($39 \pm 8\%$). $(\text{NH}_4)_2\text{SO}_4$ and BaSO_4 displayed no shrinkage. In general, salts containing divalent cations in combination with monovalent anions, resulted in major shrinkage of the scaffold.

Focussing on salts that induced shrinkage of $>90\%$, we studied the minimal concentration required to achieve full shrinkage of the scaffold (Fig. 1C). For $\text{Mg}(\text{ClO}_4)_2$ and $\text{Ba}(\text{ClO}_4)_2$ 1 M was needed, whereas for CaCl_2 and MgCl_2 1.5 and 3 M were required. Being one of the most effective salts for shrinkage, CaCl_2 was selected for further studies, comparing this salt treatment to treatment using NaCl and H_2O .

Scaffolds treated with CaCl_2 , NaCl and H_2O were characterized by scanning and transmission electron microscopy. Cross-sectional images revealed that salt treated scaffolds remained porous (Fig. 2, left panels). CaCl_2 treated scaffolds displayed circular pores but, compared to the H_2O treated scaffold, the pore size decreased from 40–120 μm to 20–60 μm . In the NaCl treated scaffolds, the pore structure had somewhat flattened and the size of pores was slightly decreased. CaCl_2 treatment had a distinct swelling effect on the individual collagen fibrils in the scaffolds. The fibrils in CaCl_2 treated scaffolds appeared distended (Fig. 2B, right panels), resulting in the loss of the characteristic banding pattern, whereas clear banding patterns were visible in NaCl and H_2O treated scaffolds (Fig. 2B and C, right panels).

The shrinking process may have an effect on the mechanical properties of the scaffolds due to the increased collagen content per volume. Therefore, the stiffness was assessed using nano-indentation (Fig. 3). CaCl_2 , NaCl and H_2O treated scaffolds were measured three times independently with 25 indentations on different locations per scaffold (Fig. 3B). CaCl_2 scaffolds (7.4×10^3 Pa) were stiffer compared to the NaCl (1.2×10^3 Pa) and the H_2O treated scaffolds (4.4×10^2 Pa) as shown in (Fig. 3C). Peak values showed a stiffness of more than 3.0×10^4 Pa, most likely when the tip indented exactly on top of collagen structure in the porous scaffold indicating that the stiffness of CaCl_2 treated collagen material is higher than 7.4×10^3 Pa.

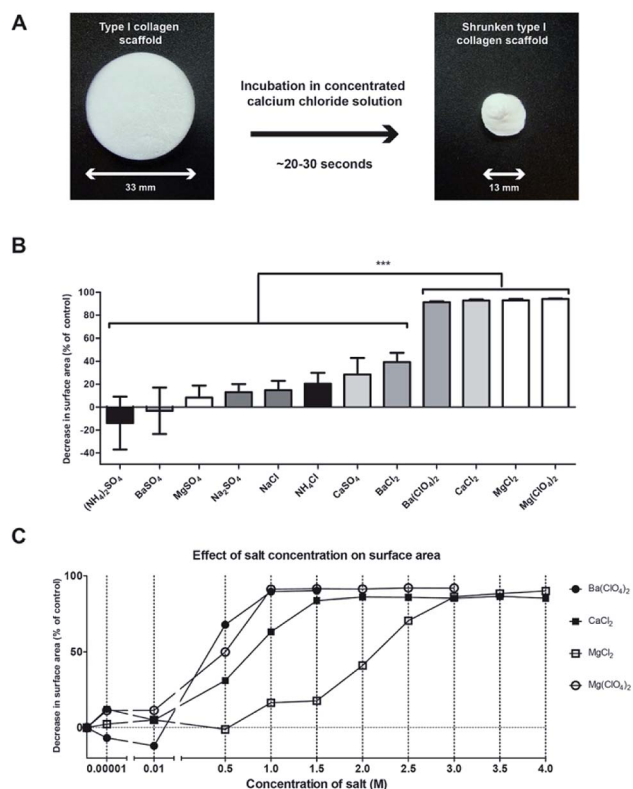


Fig. 1 The shrinkage effect of different salts from the Hofmeister series on 0.8% (w/v) collagen scaffolds. (A) Type I collagen scaffold before and after exposure to 4 M CaCl_2 . (B) The effect of 90% saturated salt solutions (x-axis) on the surface area of on collagen scaffolds (y-axis). Note that $\text{Ba}(\text{ClO}_4)_2$, CaCl_2 , MgCl_2 , and $\text{Mg}(\text{ClO}_4)_2$ induced shrinkage to over 90% compared to control scaffolds, $n = 3$. Statistical analysis: one-way ANOVA with Bonferroni *post-hoc* test, $p < 0.001$. Due to variabilities in solubility, the 90% saturated salt solutions vary in molarity (see ESI Table 1†). (C) The effect of salt concentration (x-axis) on the degree of shrinkage (y-axis) for a subset of salts that showed over 90% shrinkage in (B), $n = 3$.

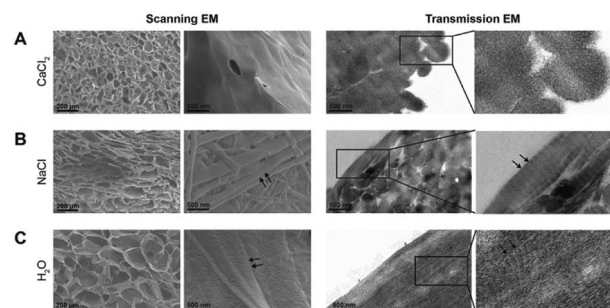


Fig. 2 Morphological characterization of collagen scaffolds. Scanning and transmission electron microscopy images of (A) a CaCl_2 treated scaffold showing the altered pore structure and the swollen fibrils without banding pattern; (B and C) NaCl (B) and H_2O (C) treated scaffolds showing the characteristic striation of collagen fibrils indicated with black arrows.



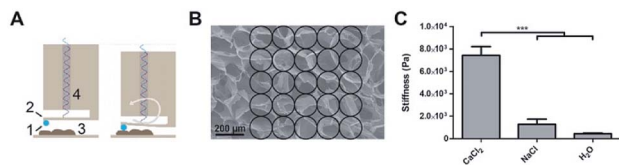


Fig. 3 Mechanical characterization using nanoindentation. (A) Working mechanism of PIUMA nanoindentation equipment: a probe (1) connected to a cantilever (2) indents the sample (3) and bends through. The degree of bending is measured via an optical cable (4) and converted to the degree of stiffness. (B) Locations of indentation on the scaffold visualized with a representative SEM image. (C) The stiffness of CaCl₂, NaCl and H₂O treated scaffolds measured with a PIUMA nanoindentation apparatus. Each condition was tested in triplicate with 25 indentations per scaffold. Bars represent mean \pm standard error of the mean. One-way ANOVA with Bonferroni *post-hoc* test, *** $p < 0.0001$.

Next, we assessed the presence of calcium or calcium-containing compounds in collagen scaffolds. Inductively coupled plasma mass spectrometry showed that the CaCl₂ treated scaffold contained only 17.8 ± 5.2 ppb calcium corresponding to 0.36 ± 0.03 ng per μg scaffold (Table 1). The NaCl and H₂O treated scaffolds contained 0.25 ± 0.02 and 0.29 ± 0.1 ng calcium per μg scaffold respectively, which were both not significantly different from the CaCl₂ treated scaffold (one-way ANOVA with Bonferroni *post-hoc* testing, $p > 0.05$). This suggests that no remnants of the calcium treatment were present in the CaCl₂ shrunken scaffolds. Using Fourier-transform infrared spectroscopy no absorption peaks were detected that could be attributed to calcium-containing compounds (Table 1).

The degree of crosslinking in the CaCl₂, NaCl, and H₂O treated scaffolds was estimated by the loss of primary amine groups (Fig. 4). Non-crosslinked CaCl₂, NaCl, and H₂O treated scaffolds contained 241 ± 41 , 244 ± 43 , 246 ± 37 nmol of primary amine groups per mg collagen, respectively, whereas crosslinked scaffolds had 81 ± 12 , 111 ± 8 and 139 ± 10 nmol. The reduction of primary amine groups was 66% for the CaCl₂ treated scaffold, 55% for the NaCl treated scaffold and 42% for the H₂O treated scaffold (Fig. 4A). The difference in crosslinking percentage between the three scaffolds types were statistically different as assessed by a one-way ANOVA with Bonferroni *post-hoc* testing, $p < 0.05$.

The extent of denaturation was assessed using electron microscopy and SDS-PAGE. Electron microscopical images of CaCl₂ treated scaffolds (non-EDC/NHS crosslinked) displayed swollen fibrils which may be caused by partial denaturation

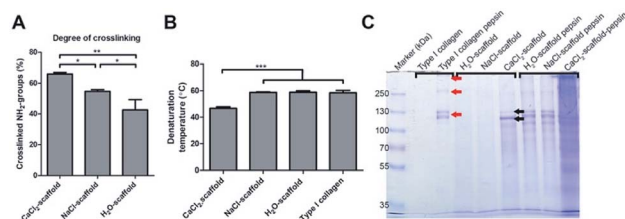


Fig. 4 Biochemical and biophysical characterization of CaCl₂, NaCl and H₂O treated scaffolds. (A) Degree of crosslinking (TNBS-assay) indicating differences between scaffolds ($n = 3$, one-way ANOVA with Bonferroni *post-hoc* test, * $p < 0.05$, ** $p < 0.01$). (B) Denaturation temperature (T_d) of purified insoluble type I collagen and various scaffold types indicated that CaCl₂ treatment lowered the T_d ($n = 6$, one-way ANOVA with Bonferroni *post-hoc* test, *** $P < 0.001$). (C) SDS-PAGE gel of purified insoluble type I collagen, CaCl₂, NaCl and H₂O treated scaffolds with and without pepsin treatment indicating that CaCl₂ treatment causes denaturation of the triple helix. The red arrows indicate the location of α , β , and γ bands after pepsin digestion and the black arrows indicates the location of α bands after CaCl₂ treatment.

induced by CaCl₂. Differential scanning calorimetry showed that the denaturation temperature was reduced from 58 ± 2 °C (type I collagen) to 47 ± 1 °C after calcium treatment indicating denaturation (Fig. 4B). The denaturation temperature of the H₂O (59 ± 1 °C) and NaCl (59 ± 1 °C) treated scaffolds did not change significantly from purified insoluble type I collagen. SDS-PAGE substantiated that CaCl₂ treatment induced denaturation (Fig. 4C). No bands will be visible on an SDS-PAGE gel when insoluble type I collagen is in its native state, as shown by the control samples (H₂O and NaCl treated scaffolds). The CaCl₂ treated collagen showed clear collagen α -chain bands (black arrows) indicating denaturation of the insoluble type I collagen fibrils. When scaffolds were subjected to pepsin treatment (liberating collagen molecules from fibrils), α , β , and γ bands were detected for H₂O and NaCl treated scaffolds (see red arrows in Fig. 4C), while for CaCl₂ treated scaffolds a smear was observed. Pepsin does not cleave intact type I collagen triple helices, but if the helix is partially denaturated it becomes susceptible for pepsin cleavage over the entire helix, resulting in a smear on gel. Consequently, the results of the SDS-PAGE indicate CaCl₂ induced denaturation of the triple helix.

Finally, we performed an *in vivo* biocompatibility study upon subcutaneous implantation in rats. All rats survived surgery without complications and did not exhibit any discomfort caused by the subcutaneously implanted scaffolds throughout the entire study. The surgical wounds healed normally. Two

Table 1 Presence of calcium in treated scaffolds

Scaffold type	ICP-MS calcium content (ng calcium per μg collagen scaffold)	FT-IR presence of calcium-containing compounds
CaCl ₂ scaffold	0.36 ± 0.03	None
NaCl scaffold	0.25 ± 0.02	None
H ₂ O scaffold	0.29 ± 0.1	None



scaffolds could not be retrieved upon explantation; *i.e.* 1 untreated scaffold at day 3 and 1 NaCl treated scaffold at day 14. Three days after implantation the original shape of all scaffolds was maintained, and they were covered with a macroscopically visible layer of fibrous tissue. Visual signs of inflammation were absent, but histologically a minor immune response was observed for all scaffolds based on the presence of granulocytes and macrophages/giant cells. For the CaCl₂ treated scaffolds, these cells were located at the rims of the scaffold (ESI Fig. 1†). At day 14 after implantation, all scaffolds were partially encapsulated, but the H₂O and NaCl treated scaffolds had clearly lost their original shape, while this was not macroscopically observed for CaCl₂ treated scaffolds. Microscopic analysis showed a somewhat increased cellular infiltration in H₂O and NaCl treated scaffolds, which may have contributed to the deformation and biodegradation of the scaffold (ESI Fig. 1†). At day 90 days after implantation, scaffolds treated with H₂O and NaCl showed clear bio-degradation, whereas the CaCl₂ treated scaffolds still kept their original shape. Statistical analysis revealed differences in cellular influx, degradation and presence of blood vessels between scaffolds at this time points (Table 2). All scaffolds were fully encapsulated as observed in the H&E staining in Fig. 5. H&E and Elastin von Masson (EvM) stainings indicated that cells (giant cells, monocytes, macrophages and fibroblasts) infiltrated the H₂O and NaCl treated scaffolds (Fig. 5B and C, left and middle panel). Cell migration was limited for the CaCl₂ treated scaffolds as cells were only observed at the outer edges of the scaffold (Fig. 5A, left and middle panel). For all scaffolds the immunological reaction was mild. Alizarin Red and von Kossa stainings did not reveal any

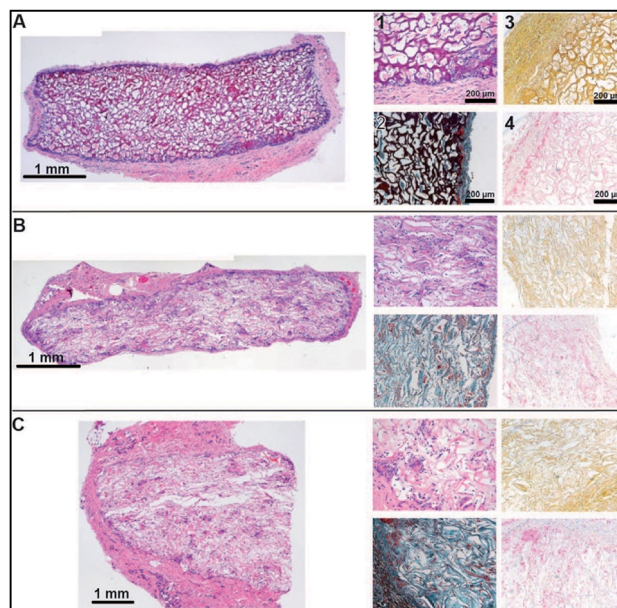


Fig. 5 Biocompatibility of CaCl₂, NaCl, and H₂O treated scaffolds in a subcutaneous rat model. In the left panel H&E overview images are depicted of a CaCl₂ treated scaffold (A), a NaCl treated scaffold (B), and a H₂O treated scaffold (C) 90 days after implantation. On the right, enlarged images of similar magnification are shown for every scaffold type. Image (1) H&E staining, (2) Elastin von Masson staining, (3) Alizarin Red staining, (4) von Kossa staining.

calcium/calcification in the implanted scaffolds (Fig. 5C, right panels). See ESI Fig. 2† for positive controls for calcium deposits (human bone tissue).

Table 2 Semi quantitative scores of haematoxylin and eosin stained sections of subcutaneously implanted collagen scaffolds treated with salt

Parameter	Scaffold	Mean ± SD		
		Day 3	Day 14	Day 90
Capsule	CaCl ₂	1.0 ± 0.5	2.8 ± 0.3	2.0 ± 0.5
	NaCl	0.3 ± 0.3	2.3 ± 0.4	1.7 ± 0.6
	Untreated	1.0 ± 0.7	2.5 ± 0.5	1.7 ± 1.3
Cellular influx	CaCl ₂	0.0 ± 0.0 ^a	0.5 ± 0.5	0.7 ± 0.3 ^a
	NaCl	0.5 ± 0.0 ^a	0.8 ± 0.4	2.8 ± 0.3 ^a
	Untreated	0.5 ± 0.0 ^a	1.0 ± 0.5	1.7 ± 0.6 ^a
Granulocytes	CaCl ₂	1.2 ± 0.6	0.1 ± 0.0	0.1 ± 0.0
	NaCl	0.5 ± 0.0	0.1 ± 0.0	0.1 ± 0.0
	Untreated	0.8 ± 0.4	0.4 ± 0.2	0.2 ± 0.2
Macrophages & giant cells	CaCl ₂	0.1 ± 0.0	1.2 ± 0.3	0.7 ± 0.3
	NaCl	0.1 ± 0.0	1.0 ± 0.0	1.7 ± 0.6
	Untreated	0.3 ± 0.3	1.2 ± 0.3	1.3 ± 0.3
Lymphocytes	CaCl ₂	0.2 ± 0.2	0.1 ± 0.0	0.4 ± 0.2
	NaCl	0.1 ± 0.0	0.1 ± 0.0	0.2 ± 0.2
	Untreated	0.1 ± 0.0	0.1 ± 0.0	0.4 ± 0.2
Blood vessels	CaCl ₂	0.0 ± 0.0	0.0 ± 0.1	0.2 ± 0.2 ^a
	NaCl	0.0 ± 0.0	0.1 ± 0.0	2.5 ± 0.5 ^a
	Untreated	0.0 ± 0.0	0.5 ± 0.5	1.0 ± 0.5 ^a
Degradation	CaCl ₂	0.0 ± 0.0	0.1 ± 0.0	0.4 ± 0.2 ^a
	NaCl	0.0 ± 0.0	0.5 ± 0.0	2.3 ± 1.2 ^a
	Untreated	0.0 ± 0.0	0.4 ± 0.2	1.2 ± 0.3 ^a

^a Significant differences by Kruskal Wallis test between scaffolds at one time point. Cellular influx day 3 $p = 0.030$; cellular influx day 90 $p = 0.029$; blood vessels day 90 $p = 0.031$; degradation day 90 $p = 0.044$.



Discussion

Over the years, methods to enhance stiffness of insoluble type I collagen scaffolds have been widely investigated. Generally, crosslinking techniques are applied to stabilize collagen scaffolds, which increase their overall strength, cohesion and stiffness.⁹ However, for a number of applications this is not sufficient and (synthetic) polymers have been added to further improve mechanical properties.¹⁸ Another option to strengthen collagen scaffolds is to modify the collagen by the addition of methacrylates that can be set by photopolymerization, which enhances the compressive modulus of *e.g.* collagen-based gels.¹⁹ In this study, a simple method is described to adapt the stiffness of collagen scaffolds and increase their strength without modifying the base material. The method is based on treatment with a concentrated salt solution resulting in major shrinkage of the scaffolds, thus increasing the collagen concentration in the walls of the porous constructs. The method resulted in rigid collagen scaffolds that were more than 10 times stiffer than untreated scaffolds, while *in vivo* biocompatibility was maintained. By adjusting the type and concentration of the salt used, the method may allow the construction of scaffolds with a specific stiffness. This is relevant since every tissue in the human body has its own characteristic stiffness caused by a distinct composition of the extracellular matrix. In recent years it has become clear that stiffness of the extracellular matrix plays an important role in differentiation of stem cells.²⁰ Consequently, to achieve optimal tissue regeneration it is important to take the stiffness of the target tissue into account when designing biomaterials. Current applications of collagen scaffolds are often limited to soft tissues, *e.g.* skin. Scaffolds treated with the shrinking method showed to have an average stiffness of 7.4×10^3 Pa with peak values of 3.0×10^4 Pa. Consequently, this may expand the application range of collagen scaffolds to harder tissues such as muscle (~ 7 – 18 kPa), cartilage (~ 25 kPa) and unmineralized bone (~ 35 kPa).^{21,22}

To investigate the molecular mechanism responsible for the shrinkage, the effect of salts from the Hofmeister series on collagen scaffolds was analyzed. The influence on collagen behavior of ions from the Hofmeister series has been reported (*e.g.* Lim, Komsa-Penkova *et al.* and Rabinovich).^{23–26} For instance, it has been observed for telopeptide-poor type I collagen molecules at low pH that kosmotropic ions induce fibrous aggregates and elastic 3D networks, whereas chaotropic ions result in finer precipitates.²⁶ In our study, only highly concentrated salts with divalent cations (Ca^{2+} , Mg^{2+} and Ba^{2+} , all at the chaotropic site of the series) in combination with an anion (not being sulfate) were able to induce shrinkage up to 90% of the initial surface area. Electron microscopical images of shrunken collagen scaffolds showed that fibrils were swollen and had lost their characteristic banding pattern, which is in line with earlier observations that CaCl_2 destabilizes both the fibrillar and triple helical structure of collagen.²⁷ A phenomenon described in leather research as ‘lyotropic swelling’, may explain the shrinkage.²⁸ During lyotropic swelling, cations such as Ca^{2+} disrupt the hydrogen bonds that stabilize the triple

helical structure. The cations interact with negatively charged carboxylic groups, thereby attracting water inside the helix, causing it to swell. For collagen, we propose the mechanism displayed in Fig. 6. Type I collagen molecules consist of 2 α_1 polypeptide chains and 1 α_2 chain, which are very close together (triple helix conformation) and are shifted one amino acid with respect to each other. This brings the carboxylic groups of the Glu and Asp amino acid residues in *e.g.* two α_1 chains (~ 100 Glu/Asp per chain) in close proximity to each other. Bivalent cations such as Ca^{2+} bind very well to two closely appositioned carboxylic groups, and this may be an initiating event in the denaturation of the triple helix and the subsequent swelling process. This mechanism would be different from reversible swelling by acid treatment (*e.g.* 0.25 M acetic acid at pH 2.5) where molecules/microfibrils start to repel each other due to similarity in charge caused by the pH. Lyotropic swelling takes place inside the triple helix and is irreversible.²⁹ When triple helices are exposed to high concentrations of cations such as Ca^{2+} , the triple helix gets disrupted (Fig. 6A, upper panel) causing the helix to swell. As a consequence, additional Ca^{2+} can enter the inner core of the helix and disrupt the hydrogen bonds that stabilize the helix structure resulting in (partial) denaturation of the triple helix. However, the (natural) covalent bonds between the individual triple helices are not affected (Fig. 6A, middle panel), preventing the fibril to fall apart. Consequently, as the helices swell, they become thicker and shorter (Fig. 6A, lower panel). On the level of fibrils, this leads to swollen but shortened fibrils (Fig. 6B). When this occurs throughout the scaffold it results in increased pore wall thickness and reduced pore size, ultimately leading to shrinkage and stiffening of the entire scaffold.

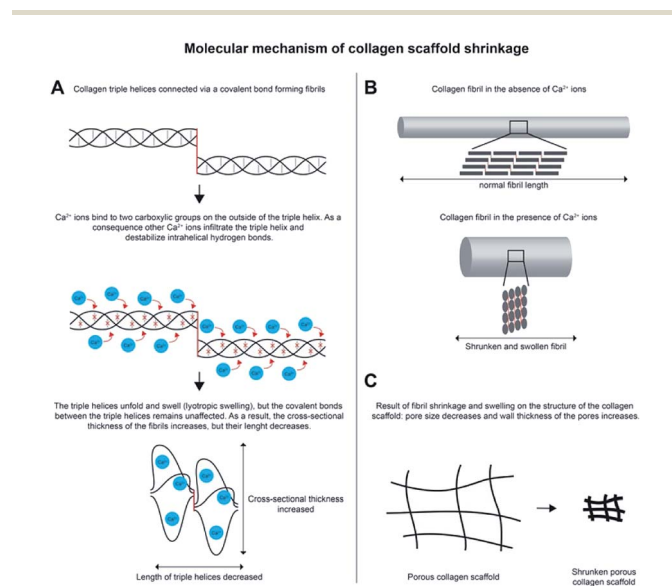


Fig. 6 Schematic representation of proposed molecular mechanism of shrinkage. (A) Ca^{2+} ions destabilize the triple helix by first binding to carboxylic groups and secondly by disrupting the hydrogen bonds inside the triple helix resulting in swelling of the triple helix, making it thicker but shorter. (B) The fibril structure changes dramatically after addition of Ca^{2+} ions: they become thicker and shorter. (C) The thicker and shorter fibrils result in shrinkage of the collagen scaffold.



Type I collagen has shown to be a biocompatible scaffolding material in many studies, also in combination with EDC/NHS crosslinking.^{30,31} After CaCl₂ treatment the collagen triple helix was partly denatured as indicated by a decrease in denaturation temperature and the appearance of pepsin-derived fragments on SDS-PAGE. However, subcutaneous implantations of the CaCl₂ treated scaffold revealed a similar cellular response as observed for H₂O and NaCl treated scaffolds indicating proper biocompatibility. The major difference was that the CaCl₂ treated scaffolds were barely infiltrated by cells explaining that the scaffolds were still looked intact 90 days after implantation maintaining their morphology. Biodegradability was clearly decreased and this may open ways to prepare collagen-based scaffold with long-term integrity.

The method developed to stiffen collagen scaffolds opens new opportunities in the field of regenerative medicine. By manipulating a number of parameters such as collagen concentration, CaCl₂ concentration or shrinkage time, scaffolds with varying stiffness may be created. The preparation of such a wide stiffness range may serve as an *in vitro* platform for stem cell research to further investigate the role of stiffness on stem cell differentiation in a natural 3D extracellular matrix. Currently, this role is generally evaluated using synthetic hydrogels prepared from *e.g.* polyacrylamide.^{32–34} A 3D environment created from a natural component of the human extracellular matrix would resemble the native stem cell niche more closely and may be advantageous to further elucidate the molecular mechanism involved in tissue regeneration and stem cell differentiation.

The *in vivo* results indicated that 90 days after implantation the CaCl₂ treated scaffolds were not deformed by surrounding tissue, and hardly populated by cells in contrast to untreated scaffolds. The CaCl₂ treated scaffolds seemed to act as a very slow degrading implant maintaining their morphological and mechanical properties over a long time period. In this respect, a function as a reservoir facilitating sustained release of drugs (*e.g.* anti-conception) or antigens (immunization) may be anticipated. The biocompatibility in combination with the slow degradation properties of CaCl₂ treated scaffolds may be interesting as a coating for medical implants such as glucose sensors as shown in a study by Ju *et al.*³⁵

Conclusions

A method was developed to enhance stiffness of porous scaffolds prepared from type I collagen applying concentrated salt solutions, as exemplified by CaCl₂. The observed shrinking effect, likely driven by the disruption of stabilizing hydrogen bonds within the triple helix, resulted in stiffer microporous scaffolds which were biocompatible, showing reduced biodegradability compared to control scaffolds. The method may be useful for the construction of collagen scaffolds with stiffness in line with the tissue to be regenerated, but may also be applied for other applications like a drug delivery system.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 S. Gomes, I. B. Leonor, J. F. Mano, R. L. Reis and D. L. Kaplan, *Prog. Polym. Sci.*, 2012, **37**, 1–17.
- 2 A. Sorushanova, L. M. Delgado, Z. Wu, N. Shologu, A. Kshirsagar, R. Raghunath, A. M. Mullen, Y. Bayon, A. Pandit, M. Raghunath and D. I. Zeugolis, *Adv. Mater.*, 2019, **31**, e1801651.
- 3 A. A. Chaudhari, K. Vig, D. R. Baganizi, R. Sahu, S. Dixit, V. Dennis, S. R. Singh and S. R. Pillai, *Int. J. Mol. Sci.*, 2016, **17**(12), E1974.
- 4 E. A. Makris, A. H. Gomoll, K. N. Malizos, J. C. Hu and K. A. Athanasiou, *Nat. Rev. Rheumatol.*, 2015, **11**, 21–34.
- 5 L. R. M. Versteegden, P. de Jonge, J. Int'Hout, T. H. van Kuppevelt, E. Oosterwijk, W. F. J. Feitz, R. B. M. de Vries and W. F. Daamen, *Eur. Urol.*, 2017, **72**, 594–606.
- 6 A. George and S. Ravindran, *Nano Today*, 2010, **5**, 254–266.
- 7 K. M. Brouwer, W. F. Daamen, H. R. Hoogenkamp, P. J. Geutjes, I. de Blaauw, W. Janssen-Kessels, W. de Boode, E. Versteeg, R. M. Wijnen, W. F. Feitz, M. Wijnen and T. H. van Kuppevelt, *J. Biomed. Mater. Res., Part B*, 2014, **102**, 756–763.
- 8 G. P. Huang, S. Shanmugasundaram, P. Masih, D. Pandya, S. Amara, G. Collins and T. L. Arinze, *J. Biomed. Mater. Res., Part A*, 2015, **103**, 762–771.
- 9 L. H. Olde Damink, P. J. Dijkstra, M. J. van Luyn, P. B. van Wachem, P. Nieuwenhuis and J. Feijen, *Biomaterials*, 1996, **17**, 765–773.
- 10 H. Lu, H. H. Oh, N. Kawazoe, K. Yamagishi and G. Chen, *Sci. Technol. Adv. Mater.*, 2012, **13**, 064210.
- 11 S. J. Lee, J. Liu, S. H. Oh, S. Soker, A. Atala and J. J. Yoo, *Biomaterials*, 2008, **29**, 2891–2898.
- 12 S. Toosi, H. Naderi-Meshkin, F. Kalalinia, M. T. Peivandi, H. HosseinKhani, A. R. Bahrami, A. Heirani-Tabasi, M. Mirahmadi and J. Behravan, *J. Biomed. Mater. Res., Part A*, 2016, **104**, 2020–2028.
- 13 C. M. Tierney, M. G. Haugh, J. Liedl, F. Mulcahy, B. Hayes and F. J. O'Brien, *J. Mech. Behav. Biomed. Mater.*, 2009, **2**, 202–209.
- 14 R. L. Baldwin, *Biophys. J.*, 1996, **71**, 2056–2063.
- 15 J. S. Pieper, A. Oosterhof, P. J. Dijkstra, J. H. Veerkamp and T. H. van Kuppevelt, *Biomaterials*, 1999, **20**, 847–858.



- 16 S. V. Beekmans and D. Iannuzzi, *Biomed. Microdevices*, 2016, **18**, 15.
- 17 L. Buttafoco, P. Engbers-Buijtenhuijs, A. A. Poot, P. J. Dijkstra, W. F. Daamen, T. H. van Kuppevelt, I. Vermes and J. Feijen, *J. Biomed. Mater. Res., Part B*, 2006, **77**, 357–368.
- 18 L. E. Freed, G. C. Engelmayr Jr, J. T. Borenstein, F. T. Moutos and F. Guilak, *Adv. Mater.*, 2009, **21**, 3410–3418.
- 19 J. Sun, W. Q. Xiao, Y. J. Tang, K. F. Li and H. S. Fan, *Soft Matter*, 2012, **8**, 2398–2404.
- 20 K. H. Vining and D. J. Mooney, *Nat. Rev. Mol. Cell Biol.*, 2017, **18**, 728–742.
- 21 A. Buxboim, I. L. Ivanovska and D. E. Discher, *J. Cell Sci.*, 2010, **123**, 297–308.
- 22 A. J. Engler, S. Sen, H. L. Sweeney and D. E. Discher, *Cell*, 2006, **126**, 677–689.
- 23 J. J. Lim, *Biopolymers*, 1976, **15**, 2371–2383.
- 24 R. Komsa-Penkova, R. Koynova, G. Kostov and B. G. Tenchov, *Biochim. Biophys. Acta*, 1996, **1297**, 171–181.
- 25 D. Rabinovich, The chemical reactivity control of collagen: the Hofmeister effect revisited, in *World Leather*, 2008, pp. 26–28.
- 26 A. M. Oechsle, M. Landenberger, M. Gibis, S. B. Irmscher, R. Kohlus and J. Weiss, *Int. J. Biol. Macromol.*, 2015, **79**, 518–526.
- 27 M. Meyer, *Biomed. Eng. Online*, 2019, **18**, 24.
- 28 A. D. Covington, *Tanning Chemistry: The Science of Leather*, Royal Society of Chemistry, Cambridge, 2011.
- 29 M. Redwood, W. Hummel and J. Cooke, *Tanning Chemistry: The Science of Leather*, Royal Society of Chemistry, 2015.
- 30 N. A. Hosper, A. J. Eggink, L. A. Roelofs, R. M. Wijnen, M. J. van Luyn, R. A. Bank, M. C. Harmsen, P. J. Geutjes, W. F. Daamen, T. H. van Kuppevelt, D. M. Tiemessen, E. Oosterwijk, J. J. Crevels, W. A. Blokk, F. K. Lotgering, P. P. van den Berg and W. F. Feitz, *Biomaterials*, 2010, **31**, 3910–3919.
- 31 R. Parenteau-Bareil, R. Gauvin and F. Berthod, *Materials*, 2010, **3**, 1863–1887.
- 32 J. H. Wen, L. G. Vincent, A. Fuhrmann, Y. S. Choi, K. C. Hribar, H. Taylor-Weiner, S. Chen and A. J. Engler, *Nat. Mater.*, 2014, **13**, 979–987.
- 33 J. S. Park, J. S. Chu, A. D. Tsou, R. Diop, Z. Tang, A. Wang and S. Li, *Biomaterials*, 2011, **32**, 3921–3930.
- 34 W. J. Hadden, J. L. Young, A. W. Holle, M. L. McFetridge, D. Y. Kim, P. Wijesinghe, H. Taylor-Weiner, J. H. Wen, A. R. Lee, K. Bieback, B. N. Vo, D. D. Sampson, B. F. Kennedy, J. P. Spatz, A. J. Engler and Y. S. Choi, *Proc. Natl. Acad. Sci. U. S. A.*, 2017, **114**, 5647–5652.
- 35 Y. M. Ju, B. Yu, T. J. Koob, Y. Moussy and F. Moussy, *J. Biomed. Mater. Res., Part A*, 2008, **87**, 136–146.

