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Oxidative cyanation of *N*-aryltetrahydroisoquinoline induced by visible light for the synthesis of α -aminonitrile using potassium thiocyanate as a "CN" agent[†]

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A novel method for the synthesis of α -aminonitrile through visible-light-induced oxidative cyanation of *N*-aryltetrahydroisoquinoline with potassium thiocyanate has been developed. The process does not require the use of a photocatalyst, transition metal reagent, strong oxidizing agent, or toxic cyano-containing compound, which makes the reaction simple and green.

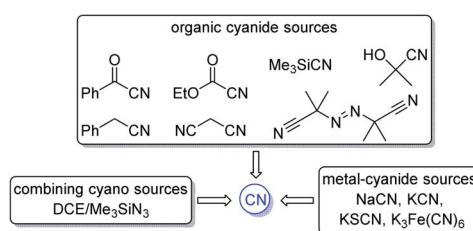
α -Aminonitriles are versatile intermediates which can be readily converted into a range of multifunctional organics,¹ such as α -aminocarbonyl compounds, α -amino acids, 1,2-diamines and many others. Moreover, many nitrile-containing compounds have been widely used as drugs for treating various diseases.² For example, amphetaminil³ is a stimulant that can be employed to treat obesity and narcolepsy; others like Saxagliptin,⁴ NVP-DPP728 (ref. 5) and Vildagliptin⁶ are potent reversible inhibitors of dipeptidyl peptidase-4 (DPP-4), and can be applied as novel antidiabetic drugs. Therefore, the development of novel, efficient and clean methods for the synthesis of α -aminonitriles has attracted wide attention.

Traditionally, α -aminonitriles are synthesized through the Strecker reaction.⁷ Another momentous strategy for synthesis of α -aminonitriles is oxidative cyanation of tertiary amines with the involvement of transition metal reagents (such as Ru,⁸ Au,⁹ Cu,¹⁰ Fe,¹¹ Co,¹² V,¹³ Mo,¹⁴ etc.) and non-metallic reagents (such as PhI(OAc)₂,¹⁵ tropylium salt,¹⁶ AIBN,¹⁷ TBAI,¹⁸ AcOH,¹⁹ PIFA,²⁰ TBHP,²¹ thiourea,²² DDQ,²³ TBPP,^{24a} etc.). Among the reported cyanation of tertiary amines, the cyano sources mainly includes (i) metal-cyanides such as NaCN,^{8a-8c,10c} KCN,¹⁶ K₃[Fe(CN)₆],²⁵ KSCN,²¹ etc., (ii) organic cyanides such as trimethylcyanosilane,^{9a,10e,11a-11c,11f} malononitrile,^{10a,15,23} ethyl cyanoformate,^{11d,8e,11d,12} benzoyl cyanide,^{11e} phenylacetonitrile,^{18a} cyanoacetic acid,^{18b} tetrabutylammonium cyanide,¹⁹ cyano-benziodoxolones,²⁴ AIBN,²⁶ etc., and (iii) combined cyano source such as the combination of 1,2-dichloroethane and trimethylsilyl azide.^{10b} (Scheme 1). However, most of these methods

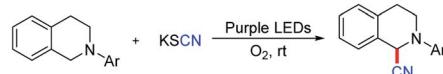
require transition metal reagents, strong oxidizing agents, and cyano sources that are highly toxic and difficult to obtain. Therefore, it is necessary to find a cyanide source that is easily available and of low toxicity for the cyanidation of tertiary amines.

In organic synthesis, thiocyanate has received extensive attention due to its low toxicity, low cost and easy availability.²⁷ Using the thiocyanato group of thiocyanate, thiocyanation of organic molecules can be achieved by nucleophilic substitution or through a free radical pathway.²⁸ More importantly, thiocyanate can be utilized as a clean and safe source of cyano group.^{21,29} However, these methods have the following disadvantages: using a transition metal catalyst, or a strong oxidizing agent, or heating. The utilization of light energy in organic synthesis is highly commendable because it is sustainable, clean, environmentally benign, as well as widely and easily available. In recent years, great progresses have been made in

Previous cyano source



This work



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Scheme 1 Cyano sources for cyanidation reactions of tertiary amine.



the formation of carbon–carbon bond that is induced by visible light,³⁰ among which the oxidative cyanation of tertiary amine is a good example.³¹ However, to the best of our awareness, the cyanation of tertiary amine with thiocyanate induced by visible light has not been documented yet. Herein, we report the visible-light-induced oxidative cyanation reaction of tertiary amines with potassium thiocyanate.

We began our investigation with the use of *N*-phenyl-tetrahydroisoquinoline (**1a**) and potassium thiocyanate (**2a**) as the model reaction (Table 1). First, the reaction was carried out in absolute ethanol illuminated by the compact fluorescent lamps (CFL) of 30 W at room temperature under air atmosphere. Fortunately, the desired product 2-phenyl-1,2,3,4-tetrahydroisoquinoline-1-carbonitrile (**3a**) was obtained in 50% yield after 24 h (entry 1). Then, we explored the effects of oxidants such as O₂, TBHP, H₂O₂, K₂S₂O₈, DDQ, and PhI(OAc)₂ on the reaction (entries 2–7). Among them, O₂ showed the best result (entry 2). The light source also had an important influence on this reaction. When the compact fluorescent lamps was replaced by a red LEDs, the yield of the reaction significantly reduced from 55% to 6% (entry 8). When green LEDs was chosen as the light source, the yield of the reaction was increased to 68% (entry 9). Besides, we investigated the effect of other visible light sources on the reaction, and purple LEDs was proved to be the best choice, up to 83% yield (entries 10 and 11). Based on the above experimental results, it is deduced that the

shorter the visible-light wavelength, the higher is the reaction efficiency. In addition, we investigated the effect of solvents such as CH₃OH, DMSO, PhMe, and CH₃CN on the reaction (entries 12–15). Among them, CH₃CN was the most suitable, giving **3a** in 92% yield. When KSCN was replaced with NH₄SCN, the yield of the corresponding product was 71% (entry 16). It is worth noting that the reaction did not proceed in the absence of either light or oxidant (entries 17 and 18).

With the optimized conditions, we explored the substrate scope of this visible-light-induced cyanation reaction between *N*-aryltetrahydroisoquinoline (**1**) and potassium thiocyanate (**2a**) (Scheme 2). The *N*-aryltetrahydroisoquinolines with either an electron-withdrawing or electron-donating group reacted well with potassium thiocyanate to afford the desired products in moderate to good yields. Specifically, when the benzene ring attached to the nitrogen atom had an electron donating group (Me, Et, and Ph) at the *para* position, the desired products were in 79%, 86% and 62% yields, respectively (**3b–d**). In the cases of having an electron-withdrawing group (F, Cl, Br, CN, CF₃, and OCF₃) attached to the benzene ring at *para* position, the desired products were obtained in moderate to good yields, and the highest one was up to 88% (**3e–j**). Besides, *ortho*- and *meta*-substituted *N*-aryltetrahydroisoquinolines also exhibited high reactivity, and the target products were obtained in moderate yields (**3k–n**).

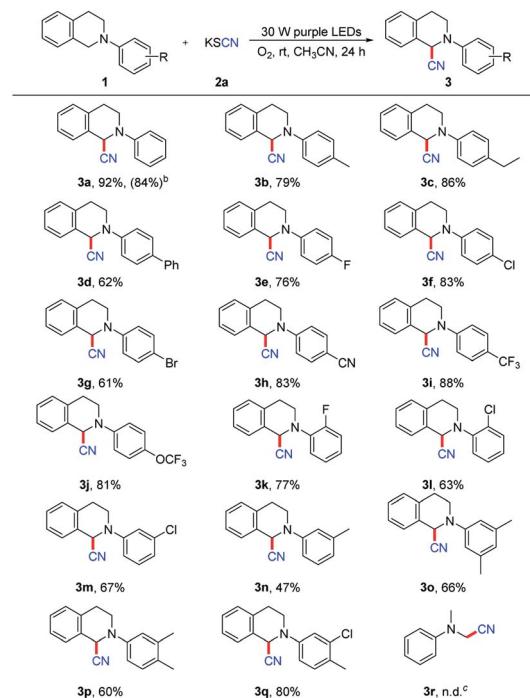
What's more, the substrates with two identical or different substituents on the phenyl ring were also suitable for this

Table 1 Optimization of cyanation of *N*-phenyl-tetrahydroisoquinoline (**1a**) with potassium thiocyanate (**2a**)^a

Entry	Oxidant	Light	Solvent	Yield ^b (%)
1	Air	30 W CFL	C ₂ H ₅ OH	50
2	O ₂	30 W CFL	C ₂ H ₅ OH	55
3	TBHP	30 W CFL	C ₂ H ₅ OH	25
4	H ₂ O ₂	30 W CFL	C ₂ H ₅ OH	17
5	K ₂ S ₂ O ₈	30 W CFL	C ₂ H ₅ OH	19
6	DDQ	30 W CFL	C ₂ H ₅ OH	3
7	PhI(OAc) ₂	30 W CFL	C ₂ H ₅ OH	5
8	O ₂	30 W red LEDs	C ₂ H ₅ OH	6
9	O ₂	30 W green LEDs	C ₂ H ₅ OH	68
10	O ₂	30 W blue LEDs	C ₂ H ₅ OH	74
11	O ₂	30 W purple LEDs	C ₂ H ₅ OH	83
12	O ₂	30 W purple LEDs	CH ₃ OH	11
13	O ₂	30 W purple LEDs	DMSO	8
14	O ₂	30 W purple LEDs	PhMe	11
15	O ₂	30 W purple LEDs	CH ₃ CN	92
16 ^c	O ₂	30 W purple LEDs	CH ₃ CN	71
17	—	30 W purple LEDs	C ₂ H ₅ OH	0
18	O ₂	—	C ₂ H ₅ OH	0

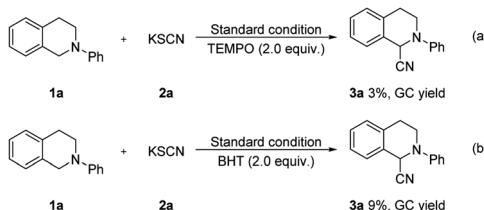
^a Reaction conditions: **1a** (0.2 mmol), **2a** (0.1 mmol), oxidant (1.5 equiv.), solvent (1.5 mL), rt, 24 h. ^b Isolated yields based on **2a**.

^c NH₄SCN was used instead of KSCN.

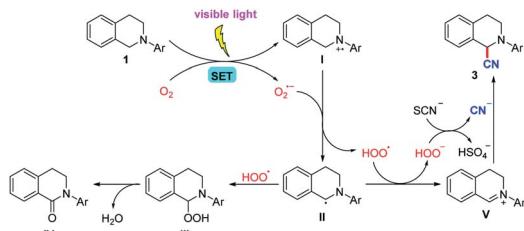


Scheme 2 Substrate scope for cyanation of *N*-aryltetrahydroisoquinolines (**1**) with potassium thiocyanate (**2a**)^a. ^b Conditions: **1** (0.2 mmol), **2a** (0.1 mmol), CH₃CN (1.5 mL), 24 h, isolated yields of column chromatography based on **2a**. ^c The yield was based on 1 mmol of KSCN. ^d No product was detected by GC-MS.





Scheme 3 Control experiments.



Scheme 4 Possible mechanism.

reaction, and the yields of the products vary from 60% to 80% (3o–q). Regrettably, dialkyl substituted aromatic amine (such as *N,N*-dimethylaniline) was not suitable for this reaction (3r).

To gain reasonable insight into the reaction mechanism, we conducted the following control experiments (Scheme 3). When 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) and 2,6-di-*tert*-butyl-4-methylphenol (BHT), both are radical quenchers, was added to the reaction under the standard conditions, the product 3a was obtained in 3% and 9% yield, respectively, detected by GC. The results reveal that the reaction may proceed with the participation of free radicals.

Based on our experimental results and those of previous reports,^{19,21,32,33} a mechanism for this visible-light-induced oxidative cyanation reaction is proposed as illustrated in Scheme 4. Initially, substrate 1 is transformed into cationic radical I in the presence of visible light and oxygen, and anionic superoxide radical O₂^{·-} is generated simultaneously. Then, cationic radical I reacts with O₂^{·-} to generate radical II which loses an electron to form intermediate V. Under the action of HOO[·], thiocyanate is converted to cyano anion which then reacts with intermediate V to generate the desired product 3. At the same time, the intermediate II is combined with HOO[·] to form III, and the latter III is dehydrated to form a by-product IV.

Conclusion

In summary, we have developed a novel method for the synthesis of α -aminonitrile using *N*-aryltetrahydroisoquinoline and potassium thiocyanate as raw materials. The product α -aminonitriles can be obtained in moderate to good yields under the illumination of 30 W purple LEDs at room temperature for 24 h in the oxygen atmosphere. The process does not involve any photocatalysts, transition metal reagents, strong oxidizing agents, and cyano-containing compounds that are highly toxic and difficult to obtain. This reaction provides a simple, green,

and efficient method for the synthesis of α -aminonitriles. Further studies on reaction mechanism and synthetic applications are currently underway.

Conflicts of interest

There are no conflicts to declare.

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