



Cite this: RSC Adv., 2019, 9, 27176

 Received 22nd July 2019  
 Accepted 1st August 2019

 DOI: 10.1039/c9ra05670j  
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## Introduction

The development of selective methods for the direct conversion of carbon–hydrogen bonds into carbon–heteroatom and carbon–carbon bonds remains a critical challenge in organic chemistry. An interesting approach to address this issue involves the use of substrates that contain coordinating atoms (or directing groups) that bind to the metal center in a first step, and a further rearrangement allows the C–H bond activation. A variety of transition-metal compounds have been shown to be effective in this process. Palladium complexes are particularly attractive catalysts for such transformations because ligand-directed C–H functionalization at palladium centres can be used to obtain different types of C–Y bonds (Y being carbon, oxygen, nitrogen, sulphur, or halogen). Furthermore, palladium can activate C–H bonds both at  $sp^2$  and  $sp^3$  sites, and a wide range of catalytic processes have been described with different directing groups.<sup>1</sup>

The use of directing groups in C–H activation processes is a strategy that provides precise site selectivity. However, it often involves the introduction of functional groups in the structure of the substrates that remain in the skeleton of the products.<sup>2</sup> A

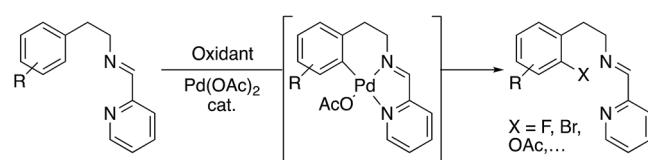
solution to this problem can be accomplished with the use of easily removable directing moieties.

In the last years, many efforts have been addressed to the use of transient directing groups in C–H activation.<sup>3</sup> Among them, we explored the use of bidentate imine-derived groups as removable directing groups. We were interested in the functionalization of phenyl groups using palladium and strong oxidants following Sanford's works with Pd(II)/Pd(IV) catalytic cycles.<sup>4</sup> Our previous experience with primary amines as directing groups in C–H functionalization reactions prompted us to investigate the use of imines derived from 2-pyridinecarboxaldehyde in acetoxylation processes (Scheme 1).<sup>5</sup>

## Results and discussion

We initially tested imine **1a** (Scheme 2) with several oxidants such as NBS for bromination or trimethylfluoropyridinium tetrafluoroborate for fluorination but complex mixtures were always obtained. Only when PhI(OAc)<sub>2</sub> was used as an oxidant, several compounds could be isolated but none of them was the anticipated acetoxylated product. Instead, all the isolated compounds contained an indoline framework with a new C–N bond formed where the C–OAc bond was expected (Scheme 2).

Indoline **4** structure was unequivocally determined by X-ray analysis (Fig. 1). The spectroscopic data of compounds **2a** and **3** suggested analogous structures to compound **4**.



Scheme 1 C–H functionalization directed by imines derived from 2-pyridinecarboxaldehyde.

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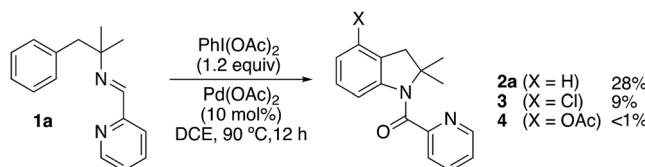
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† Electronic supplementary information (ESI) available. CCDC 1883868. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9ra05670j





Scheme 2 Indolines formation.

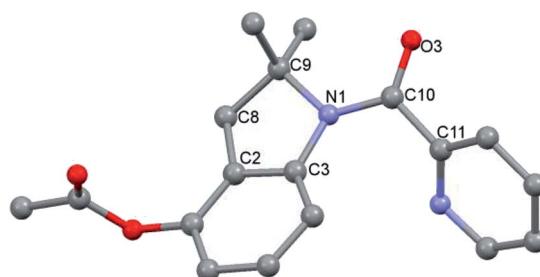
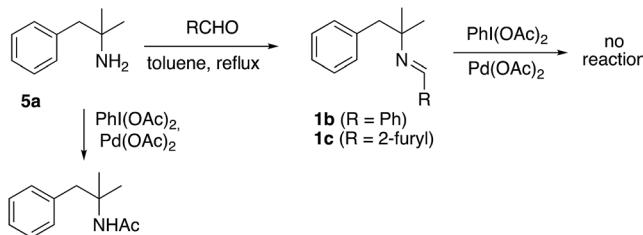


Fig. 1 Crystal structure of indoline 4. Hydrogen atoms have been omitted for clarity. Selected bond distances (Å) and angles (°): C2–C3 = 1.383(5), C2–C8 = 1.502(5), C8–C9 = 1.538(5), C9–N1 = 1.510(5), C3–N1 = 1.431(4), C10–O3 = 1.232(4), C10–N1 = 1.368(5), C3–C2–C8 = 110.5(3), C2–C3–N1 = 108.2(3), N1–C9–C8 = 101.3(3), C3–N1–C9 = 108.5(3), N1–C10–C11 = 118.9(3), C2–C8–C9 = 103.3(3), C3–C2–C8 = 110.5(3).

The unprecedented formation of the indoline framework involved the oxidation of an imine to an amide and an intramolecular C–H functionalization process but the detailed mechanism was unclear. Remarkably, several experiments showed that the imine derived from 2-pyridinecarboxaldehyde was necessary. Thus, similar imines derived from benzaldehyde or furfural, easily prepared from amine **5a** by the same procedure to prepare imine **1a**, did not afford the indoline structure under similar conditions (Scheme 3). Furthermore, unprotected amine **5a** was just acetylated when the same process was applied.

The C–N bond formation in aromatic compounds *via* Pd-catalyzed C–H functionalization has been a topic of interest for the last years.<sup>6</sup> The Pd-catalyzed synthesis of indolines from 2-phenylethylamines has been reported with picolinamides,<sup>7</sup> sulfonamides,<sup>8</sup> 1,2,3-triazole-4-carboxamides,<sup>9</sup> oxallyl amides,<sup>10</sup> 2-(methoxyimino)amides<sup>11</sup> and 2-(*tert*-butoxycarbonylamino)-amides<sup>12</sup> as directing groups in such C–H functionalization processes. Based on the above preliminary results and the



Scheme 3 Alternative nitrogen atoms as directing groups.

importance of the indoline scaffold in natural and pharmaceutical bioactive compounds,<sup>13</sup> we decided to explore the 2,2-disubstituted indoline formation using imines as directing groups. Due to the steric demand of the amino group, such indolines are usually formed from anilines and not by C–N bond formation as it is shown herein.<sup>13b</sup>

### Synthesis of indolines from imines

The initial results encouraged us to explore different reaction conditions to improve the yield of indoline **2a** (Table 1). An increase of equivalents of oxidant (entries 1 and 2) improved the yield and shortened the reaction time, being 1 hour the optimal conditions (entries 2–4). Once it was formed, the product **2a** slowly decomposed under the reaction conditions (entries 2 and 3). A reduction of catalyst amount resulted in a drop of the yield (entry 5), being necessary the presence of the Pd(II) catalyst to observe the reaction (entry 6). As expected, the removal of the oxidant was also adverse (entry 7). The amount of oxidant could be reduced up to 1.5 equiv. of PhI(OAc)<sub>2</sub> without a decrease of the yield (entry 8). Reduction of the temperature to 60 °C required longer reaction times to achieve similar yields (entries 9 and 10). As expected, the formation of the chlorinated indoline **3** was only observed in DCE. Reaction in other solvents than DCE such as AcOH, acetonitrile or toluene did not afford **3** and followed the same behavior but lower yields of **2a** were obtained: 20%, 40% and 50% respectively.

### Synthesis of indolines from imines formed *in situ*

The *in situ* formation of imines to promote the metal-catalyzed C–H activation has been recently used to functionalize *sp*<sup>2</sup> and *sp*<sup>3</sup> C–H bonds.<sup>14</sup>

Accordingly, we next explored the possibility of the *in situ* preparation of the imine followed by the indoline formation reaction. Our first attempt using DCE and molecular sieves for

Table 1 Indoline **2a** formation

Entry	Temp. (°C)	Time (h)	PhI(OAc) <sub>2</sub> (equiv.)	Pd(OAc) <sub>2</sub> (mol%)	<b>3</b> <sup>a</sup> (%)	<b>2a</b> <sup>a</sup> (%)
1	90	24	1	10	9	26
2	90	3	2	10	7	37
3	90	1	2	10	16	61
4	90	0.5	2	10	10	40
5	90	1	2	5	12	40
6	90	1	2	0	0	0
7	90	1	0	10	0	0
8	90	1	1.5	10	8	66
9	60	1	2	10	0	27
10	60	24	2	10	8	60

<sup>a</sup> Isolated yield.

the imine formation and subsequent addition of 2 equiv. of iodinane and 10 mol% of palladium acetate afforded only 25% yield of indoline **2a** (Scheme 4).

Since toluene gave also good yield in the indoline formation, we tried the two-step process in this solvent (Table 2, entry 1). Under these conditions, an overall yield of 54% was obtained which was comparable to the one obtained from pure imine **1a**.

The palladium catalyst can be also used to accelerate the imine formation. Thus, the time of imine formation (step 1) could be reduced to 1 hour when the  $\text{Pd}(\text{OAc})_2$  was added in this step without any loss of yield (55%). The amount of 2-pyridinecarboxaldehyde could also be reduced to 1 equiv. with similar overall yield (51%).

Once we confirmed that the imine could be prepared *in situ* using toluene as a solvent, we next studied the steric and electronic influence of substituents on the phenyl ring (Table 2).

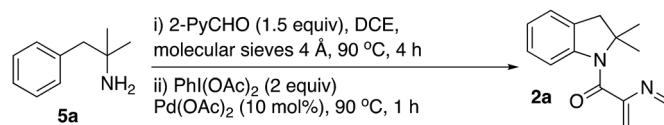
Electron donating groups in *para* and *ortho* positions afforded similar yields (entries 2–4) whereas an electron withdrawing group lowered clearly the yield (entry 5).

The  $\alpha$ -substitution of the amine is very critical. The absence of substituents  $\alpha$  to the amine, such in amine **5f** (Fig. 2), precluded the indoline formation.  $\alpha$ -Monosubstituted amines such 1-phenyl-2-propanamine (**5g**) or methyl phenylglycinate (**5h**) did not afford the expected indolines **2f–h**. Furthermore, the corresponding imines derived of 2-pyridinecarboxaldehyde of **5f–h** amines did neither afford the indolines. On the other hand, the use of either amine **5i** or the corresponding imine, with an additional methylene group in the aliphatic chain, did not afford the expected isoquinoline type structure (**2i**). Thus, we were facing to some structural limitations rather than to problems due to the *in situ* formation of the imines.

However, indolines could be formed when the primary amine was  $\alpha,\alpha$ -disubstituted as in compounds **5j–l**, (Scheme 5). Only, highly hindered  $\alpha,\alpha$ -diethyl-2-phenylethylamine afforded indoline **2l** in low yield.

To gain some mechanistic insight, the reaction between palladium acetate and imine **1a** was studied (Scheme 6). When this reaction was carried out in toluene at room temperature the coordination compound **6** was obtained. In this complex the imine acts as a bidentate ligand, coordinated by both nitrogen atoms. This coordination compound experiences easily a metatation process. Treatment of this compound in acetic acid or toluene for 1 h at 90 °C, produced the cyclometalated derivative **7** in good yield. Complex **8** was obtained by reaction between **7** and LiBr at room temperature in acetone.

It is also remarkable that the metatation process can even occur in some extension in solid state. Thus, the  $^1\text{H}$  NMR spectrum of a sample of **6**, containing initially less than 2% of



Scheme 4 *In situ* imine **1a** formation.

Table 2 *In situ* indolines **2a–e** formation

Entry	Amine	R	R'	Indoline	Yield <sup>a</sup> (%)
1	<b>5a</b>	H	H	<b>2a</b>	54
2	<b>5b</b>	Me	H	<b>2b</b>	49
3	<b>5c</b>	H	Me	<b>2c</b>	64
4	<b>5d</b>	OMe	H	<b>2d</b>	52
5	<b>5e</b>	NO <sub>2</sub>	H	<b>2e</b>	30

<sup>a</sup> Isolated yield.

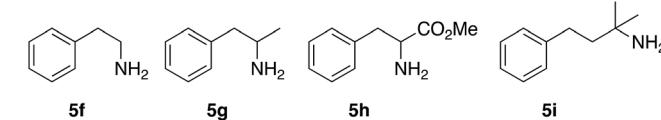
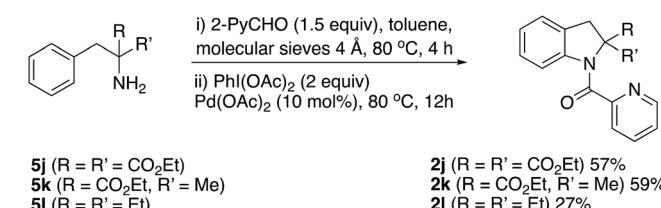
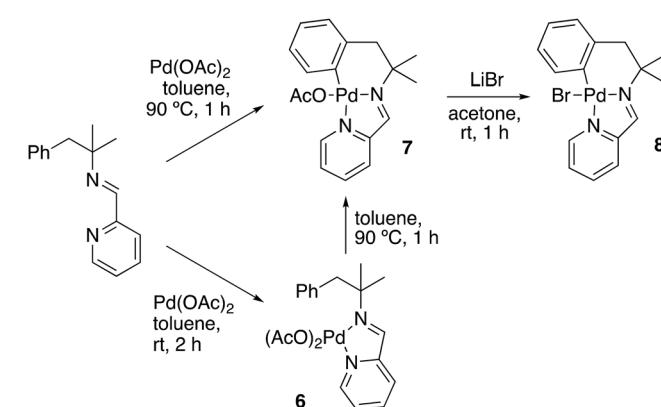


Fig. 2 Non- $\alpha,\alpha$ -disubstituted amines **5f–i**.



Scheme 5  $\alpha,\alpha$ -Disubstitution effect in indoline formation.

cyclometalated derivative **7**, shows an increase to an 8% of this complex after standing one week at room temperature in solid state. It has been previously reported that palladium(II) complexes containing two acetate ligands in the coordination



Scheme 6 Stoichiometric metatation of imine **1a**.

sphere are excellent starting materials to activate C–H bonds in a facile way, and also that these compounds experience the metalation in solid state in some extension.<sup>15</sup>

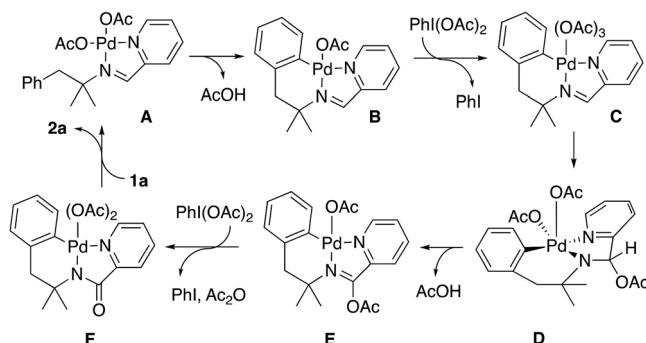
Three processes are mechanistically necessary for the indoline formation: (i) C–H functionalization of the phenyl ring, (ii) imine oxidation and (iii) C–N bond formation.

To answer whether the imine oxidation or C–H functionalization is the first step of the catalytic cycle, the reaction was performed with the imine derived of *tert*-butylamine (8). If the first step was imine oxidation, the reaction would afford the amide 9 since this imine cannot be C–H functionalized at the phenyl group. Under the standard reaction conditions, the structurally related imine 8 was not oxidized, and the imine was mainly recovered (Scheme 7).

On the other hand, when the reaction was initiated in absence of oxidant  $\text{PhI}(\text{OAc})_2$  and, once the palladacycle 7 was stoichiometrically formed, the oxidant was added, the indoline was formed.<sup>16</sup> According to the above results and based on related processes,<sup>17</sup> we propose a plausible mechanism based on a  $\text{Pd}(\text{n})/\text{Pd}(\text{iv})$  cycle (Scheme 8).

Initial coordination of the nitrogen atoms of the imine to  $\text{Pd}(\text{OAc})_2$  would afford complex A that can be converted to palladacycle B upon heating. In our hands, such a complex can be formed under stoichiometric conditions in absence of oxidant. In presence of the oxidant  $\text{PhI}(\text{OAc})_2$ , complex B would be oxidized to a  $\text{Pd}(\text{iv})$  complex (C). Acetate migration to the imine carbon would afford complex D.<sup>18</sup> The *cis* disposition between one acetato ligand and the  $\text{NCH}(\text{AcO})$  fragment in this compound permits the reduction of palladium in a process that involves the transference of the hydrogen atom in  $\beta$  position to the pendant carbonyl oxygen atom of the acetato ligand to form acetic acid. At this point, it is important the presence of the *gem*-dimethyl group, since only one hydrogen atom can be eliminated, and the process can be more selective. When this group was substituted by a methylene, a complex mixture was obtained under the reaction conditions, as it was the case with the imine derived of 2-pyridinecarboxaldehyde and the  $\alpha,\alpha$ -unsubstituted amine 5f. Pyridine-assisted deacetylation and oxidation of complex E by an additional equivalent of  $\text{PhI}(\text{OAc})_2$  would form a  $\text{Pd}(\text{iv})$  complex (F).

This complex could afford two different compounds by reductive elimination reactions. The formation of a C–N bond from this complex would afford the corresponding indoline 2a. In contrast the formation of a C–O (or a C–Cl bond when the reaction was performed in DCE)<sup>19</sup> would afford an imine derivative *ortho*-acetato (or *ortho*-chloro) substituted, which by a sequence of metalation, oxidative addition and reductive elimination with formation of C–N bond would afford the minor compounds 4 or 3 respectively.



Scheme 8 Indolines formation.

It should be noted that Sanford *et al.* have described that the distribution of different compounds obtained by competitive reductive elimination of C–Cl, C–C or C–N bonds form palladium(IV) complexes is highly sensitive to the reaction medium.<sup>20</sup>

The fact that similar imines derived from benzaldehyde or furfural did not afford the indoline structure under similar conditions can be explained because the coordination to the nitrogen atom of a pyridine fragment is a key point to afford the palladium(IV) derivatives. It should be noted that palladium catalyzed oxidative synthesis of pyrrolidine, indole and indoline derivatives needs the presence of the picolinamide fragment as a directing group.<sup>20,21</sup>

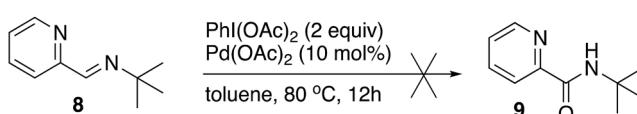
## Conclusions

In summary, a new convenient route for 2,2-disubstituted indolines was developed *via*  $\text{Pd}$ -catalyzed C–H activation of imines derived from 2-pyridinecarboxaldehyde that can be formed *in situ* in one-pot synthesis from 2-phenylethylamines. An oxidant such as  $\text{PhI}(\text{OAc})_2$  was necessary to promote a plausible  $\text{Pd}(\text{n})/\text{Pd}(\text{iv})$  cycle with concomitant oxidation of the imine to an amide. The ability of the pyridine moiety to coordinate  $\text{Pd}$  seems to be essential for this catalytic cycle. The absence of hydrogens in the  $\alpha$ -position of the amine is also a requirement for the formation of indolines.

## Experimental

### General considerations

Amines 5f–h are commercially available. Amines 5a–e and 5j–l were prepared according to known procedures.<sup>5a</sup> Amine 5i was obtained from 2-methyl-4-phenylbutan-2-ol by a Ritter–Jirgensons reaction.<sup>22</sup> Elemental analysis C, H and N analyses were performed with an Eager 1108 microanalyzer.  $^1\text{H}$  (400 MHz) and  $^{13}\text{C}$  (101 MHz) nuclear magnetic resonance (NMR) spectra were registered using a Varian Mercury 400 MHz NMR spectrometer.  $\text{CDCl}_3$  (99.9%) was used as solvent while  $\text{SiMe}_4$  (TMS) was used as reference. The coupling constants ( $J$ ) are expressed in Hz and the chemical shifts in part per million (ppm). The signals multiplicities are reported with the following abbreviations: s (singlet), d (doublet), t (triplet), q (quartet), m (multiplet), and br (broad). More complicated signals are described as



Scheme 7 Oxidation of imine 8.



combination of the indicated abbreviations: *e.g.* dd (double of doublets), dt (double of triplets), *etc.* Two-dimensional experiments (NOESY, HSQC) were performed to confirm the signal assignments. ESI (+) spectra were acquired either on an LC/MSD-TOF instrument or on a ZQ mass spectrometer, utilizing a mixture of  $\text{H}_2\text{O} : \text{CH}_3\text{CN}$  (1 : 1, v/v) as the eluent. Crystallographic data for **4** were collected on a Bruker APEX II QUAZAR diffractometer at the Group of Magnetism and Functional Molecules (GMMF) of the Universitat de Barcelona, equipped with a microfocus multilayer monochromator with Mo  $\text{K}\alpha$  radiation ( $\lambda = 0.71073 \text{ \AA}$ ). Data reduction and absorption corrections were performed by using SAINT<sup>23</sup> and SADABS,<sup>23</sup> respectively. The structure was solved using SHELXT<sup>24</sup> and refined with OLEX2 suite.<sup>25</sup>

### General procedure for the preparation of imines

Aldehyde (6.00 mmol) was added to a solution of amine **5a** (900 mg, 6.00 mmol) in toluene (60 mL). The solution was heated to reflux overnight with a Dean–Stark apparatus. The crude mixture was cooled to room temperature, dried with  $\text{MgSO}_4$  and filtered. The solvent was eliminated at reduced pressure to afford the pure imine.

#### *N*-(2-Methyl-1-phenylpropan-2-yl)-1-(pyridin-2-yl)methan-ime

**(1a).** Brown oil, 1.43 g (6.00 mmol, 100% yield).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.63–8.61 (m, 1H), 8.13 (s, 1H), 8.09–8.06 (m, 1H), 7.78–7.73 (m, 1H), 7.32–7.28 (m, 1H), 7.23–7.13 (m, 5H), 2.92 (s, 2H), 1.28 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  157.0, 155.4, 149.2, 138.3, 136.5, 130.7, 127.6, 126.1, 124.4, 120.8, 60.9, 49.6, 26.8. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 239.1543, found: 239.1542.

#### *N*-(2-Methyl-1-phenylpropan-2-yl)-1-phenylmethanimine (1b).

Brown oil, 1.42 g (6.00 mmol, 100% yield).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.99 (s, 1H), 7.75–7.70 (m, 2H), 7.42–7.38 (m, 3H), 7.23–7.13 (m, 5H), 2.91 (s, 2H), 1.26 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  157.1, 138.1, 136.4, 131.1, 129.8, 128.9, 128.5, 127.1, 125.0, 60.9, 49.6, 26.8. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 238.1590, found: 238.1590.

#### 1-(Furan-2-yl)-*N*-(2-methyl-1-phenylpropan-2-yl)methanimine (1c).

Brown oil, 1.36 g (6.00 mmol, 100% yield).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.77 (s, 1H), 7.53 (d,  $J = 1.8 \text{ Hz}$ , 1H), 7.22–7.10 (m, 5H), 6.66 (d,  $J = 3.4 \text{ Hz}$ , 1H), 6.47 (dd,  $J = 3.4, 1.8 \text{ Hz}$ , 1H), 2.91 (s, 2H), 1.24 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  152.2, 145.2, 144.4, 138.2, 130.8, 127.6, 126.0, 113.7, 111.4, 60.7, 49.3, 26.8. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 228.1383, found: 228.1383.

### General procedure for the preparation of indoline **2a** from imine **1a**

The imine (0.50 mmol) was dissolved in the suitable solvent (3 mL), then  $\text{Pd}(\text{OAc})_2$  (11 mg, 0.05 mmol) and  $\text{PhI}(\text{OAc})_2$  (322 mg, 1.00 mmol) were added. The reaction mixture was stirred at 90 °C in a sealed vial then it was filtered through a pad of Celite. The solvent was evaporated at reduced pressure to give a crude mixture that was purified by flash silica gel chromatography (hexane/ethyl acetate).

#### (2,2-Dimethylindolin-1-yl)(pyridin-2-yl)methanone (2a).

Pale oil, 77 mg (0.31 mmol, 61%) from imine **1a** in DCE for 1 h.  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.60 (d,  $J = 4.5 \text{ Hz}$ , 1H), 7.86–7.81 (m,

1H), 7.66 (d,  $J = 7.8 \text{ Hz}$ , 1H), 7.42–7.38 (m, 1H), 7.13 (d,  $J = 7.7 \text{ Hz}$ , 1H), 6.90–6.77 (m, 2H), 5.71 (br m, 1H), 3.05 (s, 2H), 1.70 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  167.2, 155.0, 149.2, 142.1, 137.2, 130.5, 126.4, 125.3, 125.2, 123.4, 123.1, 115.7, 67.5, 45.5, 25.9. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 253.1335, found: 253.1339.

#### (4-Chloro-2,2-dimethylindolin-1-yl)(pyridin-2-yl)methanone

**(3).** Pale oil, 23 mg (0.08 mmol, 16%) from imine **1a** in DCE for 1 h.  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.58 (d,  $J = 4.8 \text{ Hz}$ , 1H), 7.88–7.83 (m, 1H), 7.69–7.66 (m, 1H), 7.43–7.40 (m, 1H), 6.88 (d,  $J = 7.9 \text{ Hz}$ , 1H), 6.77–6.63 (m, 1H), 5.64 (br s, 1H), 3.08 (s, 2H), 1.71 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  167.3, 154.4, 149.1, 143.4, 137.3, 131.0, 129.0, 127.8, 125.5, 123.5, 122.9, 114.0, 67.6, 44.5, 26.0. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 287.0951, found: 287.0950.

### General procedure for the preparation of indolines from amines

2-Pyridinecarboxaldehyde (70  $\mu\text{L}$ , 0.75 mmol) was added to a solution of amine (0.50 mmol) in dry toluene (3 mL) in the presence of 4 Å molecular sieves. The solution was stirred for 4 hours at 80 °C in a sealed vial, then  $\text{Pd}(\text{OAc})_2$  (11 mg, 0.05 mmol) and  $\text{PhI}(\text{OAc})_2$  (322 mg, 1 mmol) were added. The suspension was stirred overnight at 80 °C in the sealed vial. The crude mixture was filtered through a pad of Celite, the solvent was eliminated at reduced pressure and the mixture was purified by flash silica gel chromatography (hexane/ethyl acetate) to afford the indoline.

**Pyridin-2-yl(2,2,6-trimethylindolin-1-yl)methanone (2b).** Pale oil, 66 mg (0.25 mmol, 49%) from amine **5b** (81 mg).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.62–8.60 (m, 1H), 7.86–7.82 (m, 1H), 7.64 (d,  $J = 7.8 \text{ Hz}$ , 1H), 7.43–7.39 (m, 1H), 7.01 (d,  $J = 7.5 \text{ Hz}$ , 1H), 6.71 (d,  $J = 7.5 \text{ Hz}$ , 1H), 5.51 (br s, 1H), 3.00 (s, 2H), 2.00 (s, 3H), 1.68 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  167.1, 155.1, 149.1, 142.2, 137.2, 136.1, 127.5, 125.0, 124.8, 123.9, 123.4, 116.7, 67.7, 45.2, 25.9, 21.6. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 267.1492, found: 267.1491.

**Pyridin-2-yl(2,2,4-trimethylindolin-1-yl)methanone (2c).** Pale oil, 85 mg (0.32 mmol, 64%) from amine **5c** (81 mg).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.58 (d,  $J = 4.6 \text{ Hz}$ , 1H), 7.82–7.81 (m, 1H), 7.64 (d,  $J = 7.8 \text{ Hz}$ , 1H), 7.40–7.38 (m, 1H), 6.73–6.68 (m, 2H), 5.51 (br m, 1H), 2.97 (s, 2H), 2.20 (s, 3H), 1.72 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  167.2, 155.0, 149.1, 141.7, 137.2, 134.6, 129.2, 126.5, 125.2, 124.1, 123.4, 113.2, 67.1, 44.3, 26.1, 18.6. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 267.1492, found: 267.1494.

**(6-Methoxy-2,2-dimethylindolin-1-yl)(pyridin-2-yl)methanone (2d).** Pale oil, 73 mg (0.26 mmol, 52%) from amine **5d** (90 mg).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.65 (d, 1H,  $J = 4.8 \text{ Hz}$ ), 7.86–7.82 (m, 1H), 7.65 (d,  $J = 7.8 \text{ Hz}$ , 1H), 7.42–7.38 (m, 1H), 7.01 (d,  $J = 8.2 \text{ Hz}$ , 1H), 6.44 (dd,  $J = 8.2, 2.2 \text{ Hz}$ , 1H), 5.30 (br s, 1H), 3.43 (s, 3H), 2.98 (s, 2H), 1.69 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  167.1, 158.5, 155.0, 149.2, 143.1, 137.3, 125.4, 125.2, 123.4, 122.6, 108.9, 102.6, 68.4, 55.1, 44.8, 25.9. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 283.1441, found: 283.1442.

#### (2,2-Dimethyl-6-nitroindolin-1-yl)(pyridin-2-yl)methanone

**(2e).** Pale oil, 35 mg (0.12 mmol, 30%) from amine **5e** (75 mg).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.60–8.56 (m, 1H), 7.98–7.94 (m,



1H), 7.80–7.77 (m, 2H), 7.52–7.49 (m, 1H), 7.27 (d,  $J$  = 7.7 Hz, 1H), 6.58 (br s, 1H), 3.13 (s, 2H), 1.72 (s, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  167.4, 153.5, 149.1, 147.2, 143.3, 137.9, 137.9, 126.2, 125.2, 123.9, 118.4, 110.8, 68.8, 45.1, 25.9. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 298.1186, found: 298.1189.

**Diethyl 1-picolinoylindoline-2,2-dicarboxylate (2j).** Pale oil, 39 mg (0.15 mmol, 57%) from amine **5j** (50 mg).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ , 50 °C):  $\delta$  8.49 (br m, 1H), 8.05 (br m, 1H), 7.87–7.82 (m, 1H), 7.40–7.35 (m, 1H), 7.15 (d,  $J$  = 7.7 Hz, 1H), 7.05 (br m, 1H), 7.04–6.98 (m, 1H), 4.22–4.08 (br m, 4H), 3.74 (s, 2H), 1.34–1.02 (br m, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ , 45 °C):  $\delta$  168.1, 165.5, 152.9, 147.2, 137.2, 127.8, 127.3, 125.4, 124.2, 124.1, 116.9, 75.2, 62.0, 40.7, 13.8. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 369.1445, found: 369.1450.

**Ethyl 2-methyl-1-picolinoylindoline-2-carboxylate (2k).** Pale oil, 91 mg (0.29 mmol, 59%) from amine **5k** (103 mg).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.62 (br m, 1H), 7.83 (td,  $J$  = 7.7, 1.7 Hz, 1H), 7.63 (br m, 1H), 7.42 (dd,  $J$  = 7.2, 5.3 Hz, 1H), 7.14 (d,  $J$  = 7.7 Hz, 1H), 6.95–6.75 (m, 2H), 5.63 (br m, 1H), 4.19 (br m, 2H), 3.53 (d,  $J$  = 15.7 Hz, 1H), 3.02 (d,  $J$  = 15.7 Hz, 1H), 1.79 (s, 3H), 1.21 (br s, 3H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  172.5, 166.0, 153.7, 149.5, 137.3, 129.4, 127.1, 125.4, 125.2, 123.8, 123.2, 114.7, 69.7, 61.5, 41.7, 21.6, 14.1. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 311.1390, found: 311.1385.

**(2,2-Diethylindolin-1-yl)(pyridin-2-yl)methanone (2l).** Pale oil, 38 mg (0.14 mmol, 27%) from amine **5l** (88 mg).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.62 (d, 1H,  $J$  = 4.6 Hz), 7.85–7.80 (m, 1H), 7.61 (d,  $J$  = 7.8 Hz, 1H), 7.41–7.38 (m, 1H), 7.12 (d,  $J$  = 7.3 Hz, 1H), 6.87–6.71 (m, 2H), 5.54 (br m, 1H), 3.10 (s, 2H), 2.54–2.40 (m, 2H), 1.88–1.78 (m, 2H), 0.93 (t,  $J$  = 7.4 Hz, 6H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ ):  $\delta$  167.1, 155.0, 149.4, 143.4, 137.2, 131.0, 126.4, 125.0, 124.9, 123.1, 122.9, 114.9, 73.6, 37.9, 30.4, 8.3. HRMS (ESI) ( $m/z$ ) calcd for  $[\text{M} + \text{H}]^+$ : 281.1648, found: 281.1649.

### Synthesis of coordination compound 6

Palladium acetate (112 mg, 0.50 mmol) was added to a solution of imine **1a** (120 mg, 0.50 mmol) in toluene (20 mL), the resulting solution was stirred 2 h at room temperature and then filtered to obtain **6** as yellow crystals (135 mg, 0.32 mmol, 60% yield).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.42 (d,  $J$  = 5.2, 1H), 8.05 (td,  $J$  = 7.8, 1.5 Hz, 1H), 7.65 (d,  $J$  = 7.5 Hz, 2H), 7.58–7.55 (m, 2H), 7.44 (s, 1H), 7.28 (t,  $J$  = 7.5 Hz, 2H), 7.19 (t,  $J$  = 7.5 Hz, 1H), 3.27 (s, 2H), 2.13 (s, 3H), 2.08 (s, 3H), 1.5 (s, 6H). MS (ESI) ( $m/z$ ) calcd for  $[\text{M} - 2\text{AcO}]^+$ : calcd for 344.04, found: 344.04; calcd for  $[\text{M} - 2\text{AcO} + \text{CH}_3\text{CN}]^+$ : 385.07, found: 385.07.

### Synthesis of metallated compound 7

200 mg (0.43 mmol) of **6** were dissolved in toluene and heated at 90 °C for 1 hour. The resulting solution was filtered, the solvent was evaporated, and the residue was treated with diethyl ether to obtain **7**, as yellow crystals, 120 mg, (0.30 mmol, 70% yield).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.60 (d,  $J$  = 5.1 Hz, 1H), 8.34 (s, 1H), 7.94 (td,  $J$  = 7.8, 1.6 Hz, 1H), 7.67 (d,  $J$  = 7.7 Hz, 2H), 7.57 (ddd,  $J$  = 7.8, 5.1, 1.3 Hz, 1H), 7.35 (m, 1H), 6.92 (m, 2H), 6.79 (m 1H), 2.95 (br s, 2H), 2.15 (br s, 3H), 1.40 (br s, 6H). EA (calc. for  $\text{C}_{18}\text{H}_{20}\text{N}_2\text{O}_2\text{Pd}$ ): C: 53.4% (53.68%); H: 5.1% (5.01%); N: 6.6%

(6.95%). MS (ESI) ( $m/z$ ) calcd for  $[\text{M} - \text{AcO}]^+$ : 343.04, found: 343.04.

### Synthesis of metallated compound 8

A mixture of 200 mg (0.50 mmol) of **7** and 130 mg of LiBr (1.50 mmol) was stirred in acetone at room temperature for 1 hour and then filtered to obtain **8** as yellow crystals (165 mg, 0.40 mmol, 80% yield).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ):  $\delta$  9.50 (d,  $J$  = 5.5 Hz, 1H), 8.51 (s, 11H), 8.06 (dd,  $J$  = 7.8, 1.2 Hz, 1H), 8.0 (td,  $J$  = 7.8, 2.0 Hz, 1H), 7.74 (d,  $J$  = 7.7, 1H), 7.66–7.62 (dd,  $J$  = 7.8, 5.5 Hz, 1H), 6.97 (td,  $J$  = 7.2, 1.3 Hz, 1H), 6.90–6.85 (m, 2H), 3.15 (br s, 2H), 1.45 (br s, 6H). EA (calc. for  $\text{C}_{16}\text{H}_{17}\text{BrN}_2\text{Pd}$ ): C: 45.4% (45.36%); H: 4.1% (4.04%); N: 6.4% (6.61%). MS (ESI) ( $m/z$ ) calcd for  $[\text{M} - \text{Br}]^+$ : 343.04, found: 343.04; calcd for  $[\text{M} - \text{Br} + \text{CH}_3\text{CN}]^+$ : 384.07, found: 384.07.

### Conflicts of interest

There are no conflicts to declare.

### Acknowledgements

This work was supported by the Generalitat de Catalunya (2017SGR474 and 2014SGR107) and the Spanish Ministerio de Economía y Competitividad (PGC2018-098630-B-I00, CTQ2015-65040-P, SAF2014-52223-C2-1-R and fellowship to A. M.).

### Notes and references

- (a) T. Gensch, M. N. Hopkinson, F. Glorius and J. Wencel-Delord, *Chem. Soc. Rev.*, 2016, **45**, 2900–2936; (b) J. F. Hartwig, *J. Am. Chem. Soc.*, 2016, **138**, 2–24; (c) G. Rouquet and N. Chatani, *Angew. Chem., Int. Ed.*, 2013, **52**, 11726–11743; (d) R.-Y. Zhu, M. E. Farmer, Y.-Q. Chen and J.-Q. Yu, *Angew. Chem., Int. Ed.*, 2016, **55**, 10578–10599; (e) Z. Chen, B. Wang, J. Zhang, W. Yu, Z. Liu and Y. Zhang, *Org. Chem. Front.*, 2015, **2**, 1107–1295; (f) T. W. Lyons and M. S. Sanford, *Chem. Rev.*, 2010, **110**, 1147–1169; (g) O. Baudoin, *Chem. Soc. Rev.*, 2011, **40**, 4902–4911; (h) A. J. Hickman and M. S. Sanford, *Nature*, 2012, **484**, 177–185; (i) J. Wencel-Delord, T. Droge, F. Liu and F. Glorius, *Chem. Soc. Rev.*, 2011, **40**, 4740–4761.
- (a) C. Sambiagio, D. Schönauer, R. Blieck, T. Dao-Huy, G. Pototschnig, P. Schaaf, T. Wiesinger, M. F. Zia, J. Wencel-Delord, T. Besset, B. U. W. Maes and M. Schnürch, *Chem. Soc. Rev.*, 2018, **47**, 6603–6743; (b) Y. Li and S. Xu, *Chem.-Eur. J.*, 2018, **24**, 16218–16245; (c) L. Ping, D. S. Chung, J. Bouffard and S. Lee, *Chem. Soc. Rev.*, 2017, **46**, 4299–4328.
- For reviews, see: (a) T. Bhattacharya, S. Pimparkar and D. Maiti, *RSC Adv.*, 2018, **8**, 19456–19464; (b) Q. Zhao, T. Poisson, X. Pannecoucke and T. Besset, *Synthesis*, 2017, **49**, 4808–4826; (c) M. R. Yadav, R. K. Rit, M. Shankar and A. K. Sahoo, *Asian J. Org. Chem.*, 2015, **4**, 846–864; (d) C. Wang and Y. Huang, *Synlett*, 2013, **24**, 145–149. For recent examples, see: (e) Y. Liu and H. Ge, *Nat. Chem.*,



2017, **9**, 26–32; (f) X.-Y. Chen and E. J. Sorensen, *J. Am. Chem. Soc.*, 2018, **140**, 2789–2792.

4 S. R. Neufeldt and M. S. Sanford, *Acc. Chem. Res.*, 2012, **45**, 936–946 and references therein.

5 (a) A. Mancinelli, C. Alamillo, J. Albert, X. Ariza, H. Etxabe, J. Farràs, J. Garcia, J. Granell and F. J. Quijada, *Organometallics*, 2017, **36**, 911–919; (b) A. Rodríguez, J. Albert, X. Ariza, J. Garcia, J. Granell, J. Farràs, A. La Mela and E. Nicolás, *J. Org. Chem.*, 2014, **79**, 9578–9585; (c) J. Albert, X. Ariza, T. Calvet, M. Font-Bardia, J. Garcia, J. Granell, A. Lamela, B. López, M. Martínez, L. Ortega, A. Rodriguez and D. Santos, *Organometallics*, 2013, **32**, 649–659; (d) B. López, A. Rodriguez, D. Santos, J. Albert, X. Ariza, J. Garcia and J. Granell, *Chem. Commun.*, 2011, **47**, 1054–1056.

6 (a) Y. N. Timsina, B. F. Gupton and K. C. Ellis, *ACS Catal.*, 2018, **8**, 5732–5776; (b) Y. Park, Y. Kim and S. Chang, *Chem. Rev.*, 2017, **117**, 9247–9301; (c) R. K. Rit, M. Shankar and A. K. Sahoo, *Org. Biomol. Chem.*, 2017, **15**, 1282–1293; (d) Y. Zhou, J. Yuan, Q. Yang, Q. Xiao and Y. Peng, *ChemCatChem*, 2016, **8**, 2178–2192; (e) W. A. Nack and G. Chen, *Synlett*, 2015, **26**, 2505–2511.

7 (a) G. He, Y. Zhao, S. Zhang, C. Lu and G. Chen, *J. Am. Chem. Soc.*, 2012, **134**, 3–6; (b) E. T. Nadres and O. Daugulis, *J. Am. Chem. Soc.*, 2012, **134**, 7–10; (c) G. He, C. Lu, Y. Zhao, W. A. Nack and G. Chen, *Org. Lett.*, 2012, **14**, 2944–2947.

8 (a) T.-S. Mei, D. Leow, H. Xiao, B. N. Laforteza and J.-Q. Yu, *Org. Lett.*, 2013, **15**, 3058–3061; (b) T.-S. Mei, X. Wang and J.-Q. Yu, *J. Am. Chem. Soc.*, 2009, **131**, 10806–10807; (c) J.-J. Li, T.-S. Mei and J.-Q. Yu, *Angew. Chem., Int. Ed.*, 2008, **47**, 6452–6455; (d) For analogous Cu-catalyzed functionalization, see: K. Takamatsu, K. Hirano, T. Satoh and M. Miura, *J. Org. Chem.*, 2015, **80**, 3242–3249.

9 X. Ye, Z. He, T. Ahmed, K. Weise, N. G. Akhmedov, J. L. Petersen and X. Shi, *Chem. Sci.*, 2013, **4**, 3712–3716.

10 C. Wang, C. Chen, J. Zhang, J. Han, Q. Wang, K. Guo, P. Liu, M. Guan, Y. Yao and Y. Zhao, *Angew. Chem., Int. Ed.*, 2014, **53**, 9884–9888.

11 Y.-P. He, C. Zhang, M. Fan, Z. Wu and D. Ma, *Org. Lett.*, 2015, **17**, 496–499.

12 Y. Zheng, W. Song, Y. Zhu, B. Wei and L. Xuan, *Org. Biomol. Chem.*, 2018, **16**, 2402–2405.

13 See for example: (a) D. Liu, G. Zhao and L. Xiang, *Eur. J. Org. Chem.*, 2010, 3975–3984 and references therein. For recent examples of 2,2-disubstituted indolines see: (b) Y. Ni, Q. Yu, Q. Liu, H. Zuo, H.-B. Yu, W.-J. Wei, R.-Z. Liao and F. Zhong, *Org. Lett.*, 2018, **20**, 1404–1408 and references therein.

14 S. St John-Campbell and J. A. Bull, *Org. Biomol. Chem.*, 2018, **16**, 4582–4595.

15 H. Font, M. Font-Bardia, K. Gómez, G. González, J. Granell, I. Macho and M. Martínez, *Dalton Trans.*, 2014, **43**, 13525–13536.

16 Compared with the Pd-catalyzed reaction, these stoichiometric conditions gave lower yield (20%).

17 (a) J. J. Topczewski and M. S. Sanford, *Chem. Sci.*, 2015, **6**, 70–76; (b) L. V. Desai, K. J. Stowers and M. S. Sanford, *J. Am. Chem. Soc.*, 2008, **130**, 13285–13293; (c) N. R. Deprez and M. S. Sanford, *Inorg. Chem.*, 2007, **46**, 1924–1935.

18 Intermolecular addition of the acetate group to complex **B** or **C** might also occur.

19 V. Botla, A. Akudari and C. Malapaka, *Tetrahedron Lett.*, 2019, **60**, 115–119.

20 S. R. Whitfield and M. S. Sanford, *J. Am. Chem. Soc.*, 2007, **129**, 15142–15143.

21 S. Karnakanti, Z.-L. Zang, S. Zhao, P.-L. Shao, P. Hu and Y. He, *Chem. Commun.*, 2017, **53**, 11205–11208.

22 A. Jirgensons, V. Kauss, I. Kalvinsh and M. R. Gold, *Synthesis*, 2000, 1709–1712.

23 G. M. Sheldrick, *SAINT and SADABS*, Bruker AXS Inc., Madison, Wisconsin, USA, 2012.

24 (a) G. M. Sheldrick, *Acta Crystallogr., Sect. A: Found. Crystallogr.*, 2008, **64**, 112–122; (b) G. M. Sheldrick, *SHLEXTL*.

25 O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. K. Howard and H. Puschmann, *J. Appl. Crystallogr.*, 2009, **42**, 339–341.

