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# A theoretical study on the formation and oxidation mechanism of hydroxyalkylsulfonate in the atmospheric aqueous phase

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Hydroxymethanesulfonate (HMS) is an important organosulfur compound in the atmosphere. In this work, we studied the formation mechanism of HMS *via* the reaction of formaldehyde with dissolved SO<sub>2</sub> using the quantum chemistry calculations. The results show that the barrier (9.7 kcal mol<sup>-1</sup>) of the HCHO + HSO<sub>3</sub><sup>-</sup> reaction is higher than that (1.6 kcal mol<sup>-1</sup>) of the HCHO + SO<sub>3</sub><sup>2-</sup> reaction, indicating that the HCHO + SO<sub>3</sub><sup>2-</sup> reaction is easier to occur. For comparison, the reaction of acetaldehyde with dissolved SO<sub>2</sub> also was discussed. The barriers for the CH<sub>3</sub>CHO + HSO<sub>3</sub><sup>-</sup> reaction and CH<sub>3</sub>CHO + SO<sub>3</sub><sup>2-</sup> reaction are 16.6 kcal mol<sup>-1</sup>, 2.5 kcal mol<sup>-1</sup>, respectively. This result suggests that the reactivity of HCHO with dissolved SO<sub>2</sub> is higher than that of CH<sub>3</sub>CHO. The further oxidation of CH<sub>2</sub>(OH)SO<sub>3</sub><sup>-</sup> and CH<sub>3</sub>CH(OH)SO<sub>3</sub><sup>-</sup> by an OH radical and O<sub>2</sub> shows that the SO<sub>5</sub><sup>-</sup> radical can be produced.

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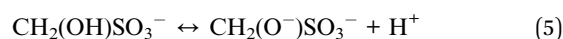
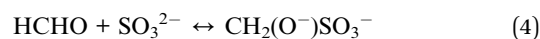
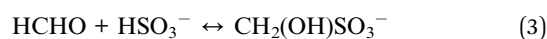
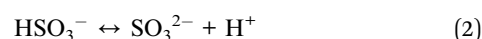
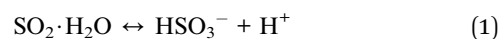
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## Introduction

Organosulfur compounds (OS), including organosulfates (ROSO<sub>3</sub><sup>-</sup>), sulfones (RSO<sub>2</sub>R'), and sulfonates (RSO<sub>3</sub><sup>-</sup>),<sup>1,2</sup> have been identified to widely exist in fog, rainwater and in ambient atmosphere aerosols.<sup>3-6</sup> These OS can be produced from marine sources including dimethylsulfide (DMS) emissions and oxidation of primary marine biomass,<sup>7,8</sup> and their formation can be affected by aerosol acidity, relative humidity and concentration of nitrogen oxides (NO<sub>x</sub>).<sup>3,9,10</sup> Hydroxymethanesulfonate (HMS) is one of the important organosulfur compounds and is a significant contributor to secondary aerosol formation.<sup>6</sup> The other hydroxyalkylsulfonate species are considered to be less important than HMS.<sup>11</sup> HMS has been misidentified as inorganic sulfate (SO<sub>4</sub><sup>2-</sup>) for a long time, which results in discrepancies between sulfate observation and model results. As a matter of fact, HMS is an important OS compound, and it may account for about 1/3 of the missing sulfate in Beijing winter haze aerosols.<sup>12</sup>

The formation of HMS has been studied by many researchers. The reaction between SO<sub>2</sub> and HCHO contributes to its formation. Wagner *et al.* proposed that the reaction rate was determined by both HSO<sub>3</sub><sup>-</sup> and SO<sub>3</sub><sup>2-</sup> (HCHO + HSO<sub>3</sub><sup>-</sup> → CH<sub>2</sub>(OH)SO<sub>3</sub><sup>-</sup>; H<sub>2</sub>O + HCHO + SO<sub>3</sub><sup>2-</sup> → CH<sub>2</sub>(OH)SO<sub>3</sub><sup>-</sup> + OH<sup>-</sup>).<sup>13</sup> But in subsequent study, Peter *et al.* disagreed with Wagner's conclusion and thought that termolecular process

(H<sub>2</sub>O + HCHO + SO<sub>3</sub><sup>2-</sup> → CH<sub>2</sub>(OH)SO<sub>3</sub><sup>-</sup> + OH<sup>-</sup>) was impossible, and the kinetics should be explained by such reactions: HCHO + HSO<sub>3</sub><sup>-</sup> → CH<sub>2</sub>(OH)SO<sub>3</sub><sup>-</sup>; HSO<sub>3</sub><sup>-</sup> ↔ SO<sub>3</sub><sup>2-</sup> + H<sup>+</sup>. They believed that only HSO<sub>3</sub><sup>-</sup> could react with HCHO and determine the reaction rate.<sup>14</sup> However, the experiment using spectrophotometer showed that the SO<sub>3</sub><sup>2-</sup> reacts rapidly with HCHO.<sup>15</sup> In the spectrophotometric study on the reaction between dissolved SO<sub>2</sub> and HCHO, it can be concluded that HSO<sub>3</sub><sup>-</sup> and SO<sub>3</sub><sup>2-</sup> can react with HCHO, and the rate for SO<sub>3</sub><sup>2-</sup> is higher than HSO<sub>3</sub><sup>-</sup> obviously.<sup>16</sup> In a recent research, the important organosulfur compound (HMS) was been investigated for its important role in haze aerosols.<sup>12</sup> And this paper proposed a potentially HMS chemical mechanism. According to the analysis above, the chemical formation of HMS can be illustrated as:<sup>16,17</sup>



Although the HMS formation mechanism has been proposed in many researches, there is no theoretical calculation on it. Thus, in order to confirm whether the mechanism is reliable, theoretical calculation is necessary to be done.

In this paper, we investigated the reaction HCHO with HSO<sub>3</sub><sup>-</sup> and SO<sub>3</sub><sup>2-</sup> using quantum chemical calculations. In

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order to understand the influence of different aldehydes for the reactions, we also discussed the reaction between  $\text{CH}_3\text{CHO}$  and  $\text{HSO}_3^-$  or  $\text{SO}_3^{2-}$ . Besides, the total rate constant ( $k_{\text{total}}$ ,  $\text{M}^{-1} \text{s}^{-1}$ ) for the individual reaction pathway within the range of 200 to 298 K were calculated. The further oxidation of the reaction products ( $\text{CH}_2(\text{OH})\text{SO}_3^-$  and  $\text{CH}_3\text{CH}_2(\text{OH})\text{SO}_3^-$ ) by OH radical and  $\text{O}_2$  also be talked about.

## Computational methods

The Gaussian-09 suite of programs was used to perform all the quantum chemistry calculations described in this paper. The density functional theory (DFT) was used for calculations.<sup>18</sup> All the geometrical structures (including reactants, pre-reactive complexes, transition states and products) calculated in this paper were optimized using the M06-2X functional at the 6-311++G(d,p) basis set.<sup>19,20</sup> Vibrational frequencies were calculated at the same level of theory to ascertain the local minimum points and the transition states, which supposed to have zero and one imaginary frequency. We also performed intrinsic reaction coordinate (IRC)<sup>21</sup> calculations in order to prove whether the transition state we found were the correct. Single-point energies were refined using the CCSD(T) method<sup>22</sup> with the aug-cc-pVTZ basis set.<sup>23–25</sup> Very recently, many high level quantum chemical methods have been used for the atmosphere reactions in order to obtain more reliable and excellent intrinsic accuracy.<sup>26–28</sup> In this article, under the consider of computational speed and accuracy, we think the dual-level strategy (CCSD(T)/aug-cc-pVTZ//M06-2X/6-311++G(d,p) level of theory) is appropriate. M06-2X functional has been widely used in theoretical calculation and can be better for ionic hydrogen-bonding interactions and identifying the global minimum conformer.<sup>29–36</sup> Been considered as the “gold standard” of quantum chemistry, CCSD(T) method was widely been used for reactions of organic matters and has high level accuracy. Thus we choose M06-2X functional at the 6-311++G(d,p) basis set for geometrical structures optimization and CCSD(T) method with the aug-cc-pVTZ basis set for single point energy calculation in this paper.<sup>22,37–40</sup> The Gibbs free energies were calculated by the following equation:  $G = E$  (single-point energy) +  $G_{\text{corr}}$  (thermal correction to Gibbs free energy). The geometries were drawn using the CYLview software package.<sup>41</sup>

For the kinetics analysis, the conventional transition-state theory (TST)<sup>42</sup> with Wigner tunneling correction was used to calculate the rate constants. All rate constants were calculated by using the KiSThELP program.<sup>43</sup>

## Results and discussion

### Reaction of HCHO with dissolved $\text{SO}_2$

It aqueous phase, dissolved  $\text{SO}_2$  can be dissociated to form  $\text{HSO}_3^-$  and  $\text{SO}_3^{2-}$ . Thus, we focus on the reactions of  $\text{HSO}_3^-$  + HCHO and  $\text{SO}_3^{2-}$  + HCHO.

As shown in Fig. 1(a), when HCHO approaches  $\text{HSO}_3^-$ , the reaction occurs *via* a five-membered cyclic prereactive complex  $\text{HCHO} \cdots \text{HSO}_3^-$  (C1). C1 is regarded as the initial step of the HCHO and  $\text{HSO}_3^-$  reaction because C1 is connected with the

transition state TS1. The C1 is held together through a hydrogen bond and an electron donor–acceptor type of interaction between the two molecules. The change of the distances of  $\text{C}_b \cdots \text{O}_d$  and  $\text{S}_c \cdots \text{C}_e$  are all shorten from 2.11 Å (C1) to 1.70 Å (TS1) and 3.24 Å (C1) to 2.11 Å (TS1), respectively. The reaction proceeds *via* a transition state TS1 with a barrier of 9.7 kcal mol<sup>-1</sup> to produce  $\text{CH}_2(\text{OH})\text{SO}_3^-$  (HMS).

When  $\text{SO}_3^{2-}$  reacts with HCHO, the process is similar with the HCHO +  $\text{HSO}_3^-$  reaction. From Fig. 1(b), it can be found that the free energy of the complex C2 is 22.5 kcal mol<sup>-1</sup> lower than reactants. Similar with C1, C2 is regarded as the initial step of the HCHO and  $\text{SO}_3^{2-}$  reaction. The complex C2 is held together by one van der Waals interaction. The change of distance between  $\text{S}_a \cdots \text{C}_b$  is from 2.91 Å to 2.67 Å. The result is contrary to that of the HCHO +  $\text{HSO}_3^-$ , in which the higher free energy of the C1 can be observed. Once the C2 is formed, it can easily transform to product  $\text{CH}_2(\text{O}^-)\text{SO}_3^-$  because the free energy of transition state TS2 is only 1.6 kcal mol<sup>-1</sup> higher than the complex C2.

According to the analysis above, it is clear that the  $\text{SO}_3^{2-}$  is more likely to react with HCHO. However, the dissolved  $\text{SO}_2$  exist in the form of  $\text{HSO}_3^-$  in the pH range of 2–7,<sup>44</sup> and the acidic condition of aerosol particles can be found in Beijing winter haze.<sup>45</sup> Thus, the  $\text{HSO}_3^-$  predominates in aerosol particles of Beijing winter haze so as to that the main reaction is HCHO +  $\text{HSO}_3^-$  in these aerosol particles. And only the further oxidation of HMS will be discussed in the latter part.

### Reaction of $\text{CH}_3\text{CHO}$ with dissolved $\text{SO}_2$

In order to study the effect of different aldehydes on the reaction, we also discussed the reaction between  $\text{CH}_3\text{CHO}$  with dissolved  $\text{SO}_2$ . As shown in Fig. 2(a),  $\text{CH}_3\text{CHO}$  +  $\text{HSO}_3^-$  reaction is firstly considered. The reaction is initiated with the formation of a five-ringlike structure complex  $\text{CH}_3\text{CHO} \cdots \text{HSO}_3^-$  (C3), followed by the formation of a transition state TS3 to produce  $\text{CH}_3\text{CH}(\text{OH})\text{SO}_3^-$  (HES). The complex C3 is held together by one hydrogen bond and one van der Waals interaction. As shown in Fig. 2(a), the hydrogen atom ( $\text{H}_b$ ) of  $\text{HSO}_3^-$  interacts with the oxygen atom ( $\text{O}_d$ ) of  $\text{CH}_3\text{CHO}$  to form one hydrogen bond, and the sulfur atom ( $\text{S}_c$ ) of  $\text{HSO}_3^-$  reacts with the carbon atom ( $\text{C}_e$ ) of  $\text{CH}_3\text{CHO}$ . It is obviously that the distances between  $\text{S}_c \cdots \text{C}_e$  and  $\text{H}_b \cdots \text{O}_d$  are all shorten. The barrier in this reaction is 16.6 kcal mol<sup>-1</sup>, which is larger than that of HCHO +  $\text{HSO}_3^-$  reaction.

For  $\text{CH}_3\text{CHO}$  +  $\text{SO}_3^{2-}$  reaction in Fig. 2(b), the formed complex C4 is also lower in energy than the reactants, which is similar with the HCHO +  $\text{SO}_3^{2-}$  reaction. The product  $\text{CH}_3\text{CH}(\text{O}^-)\text{SO}_3^-$  can be produced from the C4. The complex C4 is held together by one van der Waals interaction. As shown in Fig. 2(b), and sulfur atom ( $\text{S}_a$ ) is involved in the formation of van der Waals interaction with a carbon atom ( $\text{C}_b$ ) of  $\text{HSO}_3^-$ . The distance between  $\text{S}_a$  and  $\text{C}_b$  changes from 3.30 Å (C4) to 2.49 Å (TS4). The process needs to cross a transition state TS4 and to overcome the barrier of 2.5 kcal mol<sup>-1</sup>, which is higher than that in HCHO +  $\text{SO}_3^{2-}$  reaction.

On the basis of the results, it can be concluded that the reactivity of HCHO with dissolved  $\text{SO}_2$  is higher than that of another aldehydes with dissolved  $\text{SO}_2$ .



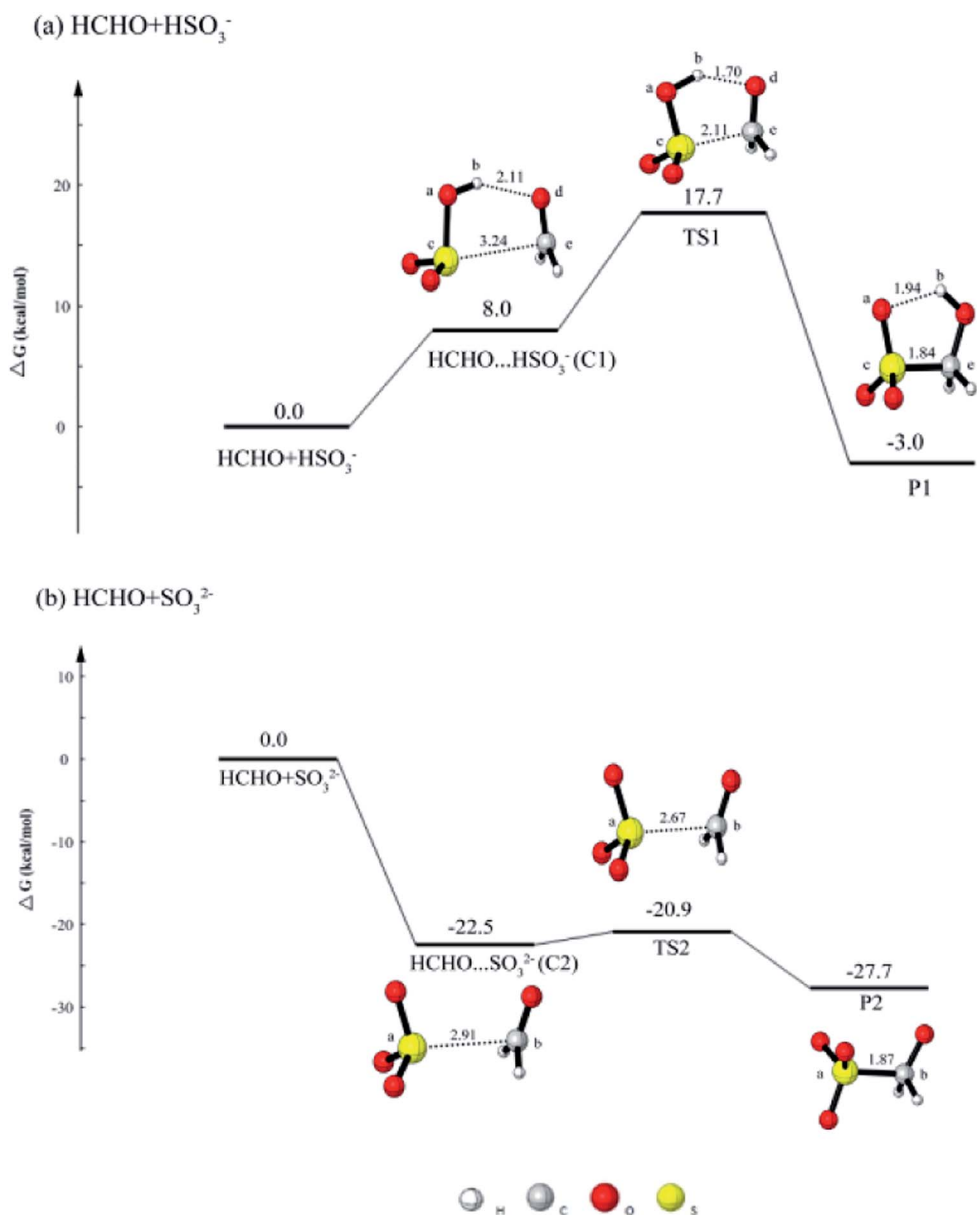
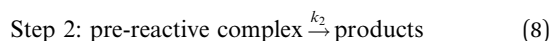
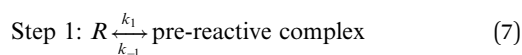


Fig. 1 The calculated free energy diagram for (a) the HCHO + HSO<sub>3</sub><sup>-</sup> reaction, and for (b) the HCHO + SO<sub>3</sub><sup>2-</sup> reaction calculated at the CCSD(T)/aug-cc-pVTZ//M06-2X/6-311++G(d, p) level of theory.

### Rate calculation

In terms of the theoretical results discussed above, the reactions occur through a two-step mechanism, involving firstly a fast pre-equilibrium between the reactants and a pre-reactive complex, and the irreversible formation of the products, which can be characterized by eqn (7) and (8).



In the above reactions,  $k_1$  is the kinetic rate constant characterizing the forward bimolecular reaction step (in cm<sup>3</sup> per molecule per s), whereas  $k_{-1}$  and  $k_2$  represent the backward and

forward unimolecular reaction rate constants (in s<sup>-1</sup>). A steady-state analysis of the total reaction pathway's rate constant is formulated as:

$$k_{\text{total}} = \frac{k_1 k_2}{k_{-1} + k_2} = k_{\text{eq}} k_2 \quad (9)$$

And  $K_{\text{eq}}$  and  $k_2$  are the equilibrium constant of the first step and the rate constant of the second step in the reactions, respectively. The computed data are shown in Table 1.

The rate constants of HCHO + SO<sub>3</sub><sup>2-</sup> are ~10<sup>9</sup> to 10<sup>7</sup> times than that of HCHO + HSO<sub>3</sub><sup>-</sup>, whereas rate constants of CH<sub>3</sub>CHO + SO<sub>3</sub><sup>2-</sup> is 8–6 orders of magnitude larger than that of CH<sub>3</sub>CHO + HSO<sub>3</sub><sup>-</sup> within range of 200–298 K. These results show that the reaction between aldehydes and SO<sub>3</sub><sup>2-</sup> is faster, which is consistent with the analysis above. For HCHO and CH<sub>3</sub>CHO, it



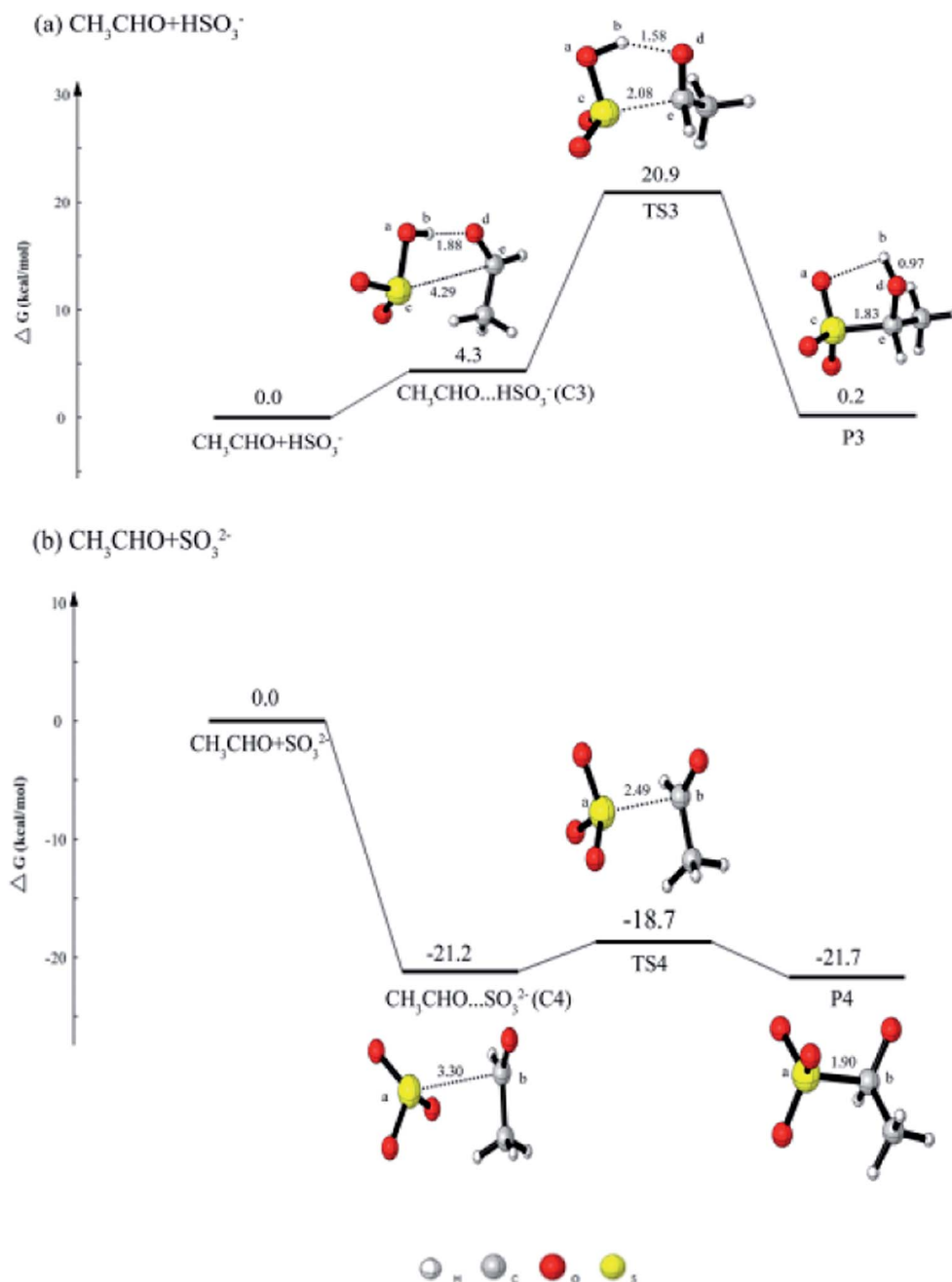


Fig. 2 The calculated free energy diagram for (a) the  $\text{CH}_3\text{CHO} + \text{HSO}_3^-$  reaction, and for (b) the  $\text{CH}_3\text{CHO} + \text{SO}_3^{2-}$  reaction calculated at the CCSD(T)/aug-cc-pVTZ//M06-2X/6-311++G(d, p) level of theory.

is clear that the rate constants between  $\text{HCHO}$  and  $\text{HSO}_3^-$  or  $\text{SO}_3^{2-}$  are larger than that between  $\text{CH}_3\text{CHO}$  and  $\text{HSO}_3^-$  or  $\text{SO}_3^{2-}$ , which is coincident with the discussion above.

Although the reaction between aldehydes and  $\text{SO}_3^{2-}$  is faster,  $\text{HSO}_3^-$  is the main form of dissolved  $\text{SO}_2$  in the aerosol particles, leading to that the reaction of aldehydes with  $\text{HSO}_3^-$

Table 1 The total rate constant ( $k_{\text{total}}$ ,  $\text{M}^{-1} \text{s}^{-1}$ ) for the individual reaction pathway within the temperature range from 200 to 298 K

	200 K	220 K	240 K	260 K	280 K	298 K
$\text{HCHO} + \text{HSO}_3^- \rightarrow \text{CH}_2(\text{OH})\text{SO}_3^-$	$4.00 \times 10^{-3}$	$1.70 \times 10^{-2}$	$6.06 \times 10^{-2}$	$1.73 \times 10^{-1}$	$4.27 \times 10^{-1}$	$8.60 \times 10^{-1}$
$\text{HCHO} + \text{SO}_3^{2-} \rightarrow \text{CH}_2(\text{O}^-)\text{SO}_3^-$	$3.21 \times 10^6$	$2.74 \times 10^6$	$2.42 \times 10^6$	$2.20 \times 10^6$	$2.04 \times 10^6$	$1.93 \times 10^6$
$\text{CH}_3\text{CHO} + \text{HSO}_3^- \rightarrow \text{CH}_3\text{CH}(\text{OH})\text{SO}_3^-$	$3.52 \times 10^{-6}$	$2.41 \times 10^{-5}$	$1.20 \times 10^{-4}$	$4.69 \times 10^{-4}$	$1.52 \times 10^{-3}$	$3.80 \times 10^{-3}$
$\text{CH}_3\text{CHO} + \text{SO}_3^{2-} \rightarrow \text{CH}_3\text{CH}(\text{O}^-)\text{SO}_3^-$	$3.96 \times 10^2$	$5.85 \times 10^2$	$8.23 \times 10^2$	$1.10 \times 10^3$	$1.43 \times 10^3$	$1.76 \times 10^3$



predominates in the aqueous phase of aerosol particles. Thus, in the next part, only  $\text{CH}_2(\text{OH})\text{SO}_3^-$  and  $\text{CH}_3\text{CH}(\text{OH})\text{SO}_3^-$  will be talked about.

### Oxidation of HMS and HES

As the oxidation of HMS is of great importance, it is meaningful and necessary to investigate the oxidation mechanism. Previous research<sup>46</sup> has investigated the oxidation of HMS with  $\text{H}_2\text{O}_2$  and  $\text{O}_3$ , there was no calculation on the oxidation by OH radical. Thus, we will calculate the HMS + OH reaction so as to confirm whether peroxysulfate radicals ( $\text{SO}_5^{\cdot-}$ ) can be produced.

Fig. 3(a) shows the potential energy profile for  $\text{CH}_2(\text{OH})\text{SO}_3^- + \text{OH}$  reaction. In the reaction process, HMS firstly reacts with

OH radical to form the complex  $\text{CH}_2(\text{OH})\text{SO}_3^-\cdots\text{OH}$  (C5) with the free energy release of  $35.2 \text{ kcal mol}^{-1}$ . The C5 can evolve *via* TS5 with a barrier of  $29.9 \text{ kcal mol}^{-1}$  into  $\text{CH}_2(\text{OH})_2$  and  $\text{SO}_3^{\cdot-}$ . The  $\text{SO}_3^{\cdot-}$  can continue to react with  $\text{O}_2$  to produce  $\text{SO}_5^{\cdot-}$ . The addition reaction between  $\text{SO}_3^{\cdot-}$  and  $\text{O}_2$  can occur *via* a transition state TS6 with the barrier of  $12.2 \text{ kcal mol}^{-1}$ .

The further oxidation mechanism of  $\text{CH}_3\text{CH}(\text{OH})\text{SO}_3^-$  (HES) by OH radical is similar with that of HMS by OH radical. As can be seen from Fig. 3(b), OH radical is added to HES to produce the complex  $\text{CH}_3\text{CH}(\text{OH})\text{SO}_3^-\cdots\text{OH}$  (C6). The reaction proceeds *via* a transition state TS7 with a barrier of  $31.0 \text{ kcal mol}^{-1}$ . The formed  $\text{SO}_3^{\cdot-}$  in this reaction also react with  $\text{O}_2$  to form  $\text{SO}_5^{\cdot-}$ , and the process has the same barrier of  $12.2 \text{ kcal mol}^{-1}$ .

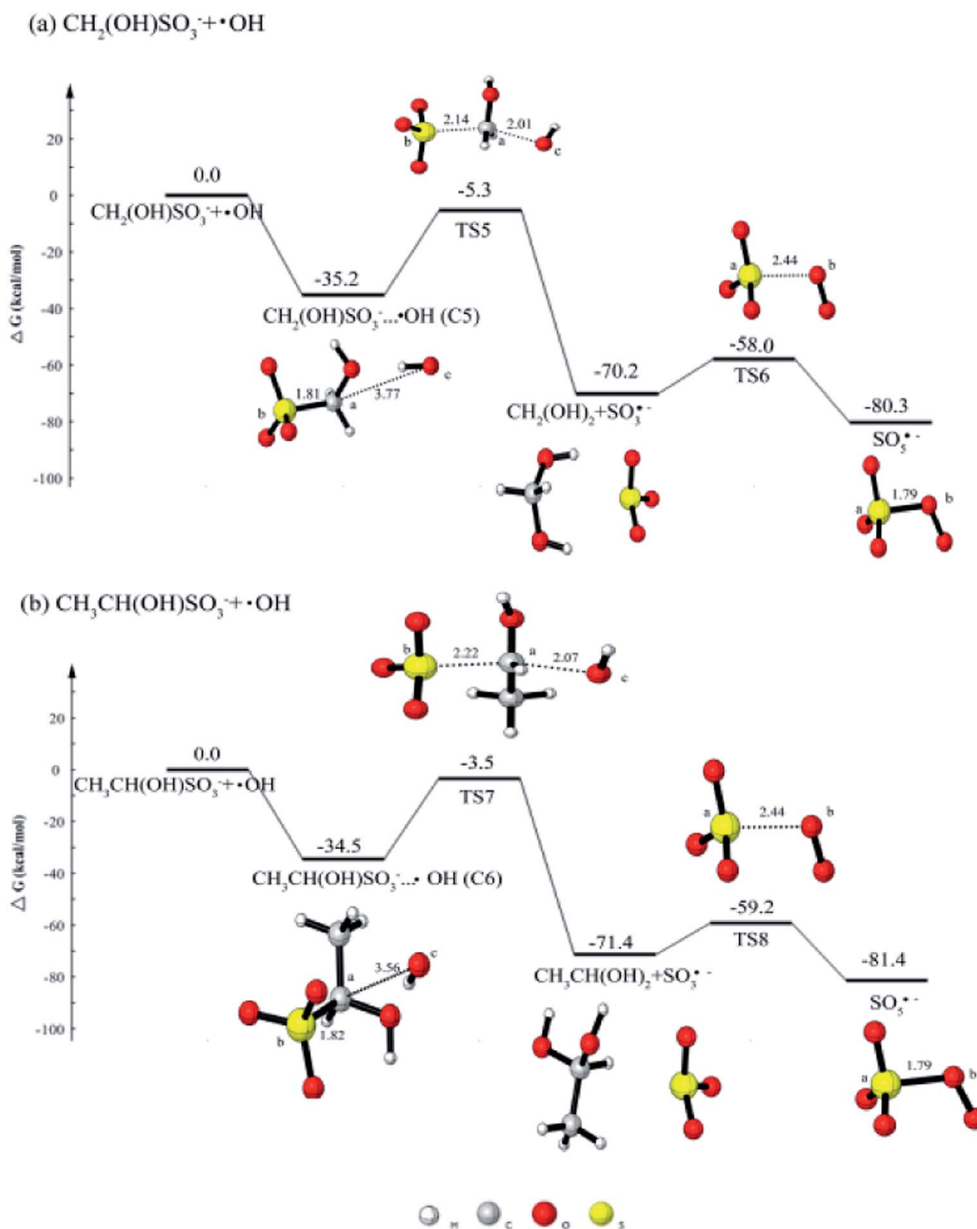


Fig. 3 The calculated free energy diagram for (a) the  $\text{CH}_2(\text{OH})\text{SO}_3^- + \cdot\text{OH}$  reaction, and for (b) the  $\text{CH}_3\text{CH}(\text{OH})\text{SO}_3^- + \cdot\text{OH}$  reaction calculated at the CCSD(T)/aug-cc-pVTZ//M06-2X/6-311++G(d, p) level of theory.



## Conclusions

HMS is the major OS species. In this paper, we investigated the formation of HMS using the quantum chemical calculations. Besides, other aldehydes like acetaldehyde also exist in the atmosphere. The similar structure between formaldehyde and acetaldehyde makes us to think the effect of different aldehydes on their reaction with dissolved  $\text{SO}_2$ . Thus, the reaction between acetaldehyde and dissolved  $\text{SO}_2$  also was discussed.

The result shows that the energy barrier for  $\text{CH}_3\text{CHO} + \text{HSO}_3^-$  reaction is  $16.6 \text{ kcal mol}^{-1}$ , which is a little higher than  $9.7 \text{ kcal mol}^{-1}$  for reaction of  $\text{HCHO}$  with  $\text{HSO}_3^-$ . The barrier ( $2.5 \text{ kcal mol}^{-1}$ ) of  $\text{CH}_3\text{CHO} + \text{SO}_3^{2-}$  reaction is larger than that ( $1.6 \text{ kcal mol}^{-1}$ ) of  $\text{HCHO}$  and  $\text{SO}_3^{2-}$  reaction. These results indicate that the reaction of aldehydes with  $\text{SO}_3^{2-}$  is easier than that with  $\text{HSO}_3^-$ . However, the  $\text{HSO}_3^-$  is the main form of dissolved  $\text{SO}_2$  in the aerosol particles, leading to that the aldehydes +  $\text{HSO}_3^-$  reaction dominates. Thus, the main products are  $\text{CH}_2(\text{OH})\text{SO}_3^-$  and  $\text{CH}_3\text{CH}(\text{OH})\text{SO}_3^-$ . Their further oxidation by OH radical and  $\text{O}_2$  shows that the  $\text{SO}_3^{\cdot-}$  and  $\text{SO}_5^{\cdot-}$  radical can be formed.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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