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An efficient and recyclable nanocatalyst for the green and rapid synthesis of biologically active polysubstituted pyrroles and 1,2,4,5-tetrasubstituted imidazole derivatives†

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An effective process for the green and rapid synthesis of biologically active polysubstituted pyrroles and 1,2,4,5-tetrasubstituted imidazoles derivatives using Cu@imine/Fe₃O₄ MNPs catalyst under solvent-free conditions is explained. This catalyst showed high reactivity for the synthesis of a set of different derivatives of polysubstituted pyrroles and 1,2,4,5-tetrasubstituted imidazole derivatives under appropriate reaction conditions and short times. Moreover, the catalyst was also recycled and reused for six runs with no considerable reduction in reactivity and yields. Compared to the reported procedures, this method consistently demonstrates the advantages of low catalyst loading, short reaction times, easy separation and purification of the products, high yields, and high recoverability and recoverability of the catalyst.

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Introduction

Recently, studies for finding separable and reusable nanocatalysts are seeing rapid progress in chemistry, especially in modern catalysis research. The immobilization of organic and inorganic functional groups on heterogeneous supports is in fact an effective process to achieve this goal.¹ Nowadays, compared to the typical separation, magnetic separation is receiving more attention due to activity as a highly efficient and rapid separation procedure for products and catalysts.² Magnetic nanoparticles (MNPs) have unique physical and chemical features such as large surface area per weight that causes these materials to act as supports of homogeneous catalysts or heterogeneous promoters for catalytic reactions. The functionalized magnetic nanoparticles can be simply isolated from the reaction solution by an appropriate external magnetic field, without any tedious workup procedures, which reduces the catalyst loss during separation.^{3,4} In addition, magnetic-supported catalysts can be recovered several times with no considerable reduction in their initial activity. High coercivity, greater selectivity, environmental friendliness, noncorrosive nature, operational simplicity, moisture

insensitivity, and low toxicity are the other factors that have attracted more attention of scientists to these nanoscale materials.^{5–9}

Poly substituted imidazoles are one of the core structural skeleton in many key molecules that have attracted considerable interest because of their biological activities such as histidine, histamine, and biotin.^{10–12} These were firstly prepared by *via* cyclocondensation reaction between 1,2-dicarbonyl compounds, different aldehydes, and a nitrogen resource.^{13,14} A variety of catalysts in the papers following this method have been reported.^{15–22} Although most of these synthetic approaches provide an improvement in the production of the above mentioned heterocyclic compounds, many of these procedures suffer from one or more drawbacks like applying dangerous organic solvents and costly reagents, harsh reaction conditions, difficult and complex work-up and purification, considerable values of waste ingredients, use of toxic reagents and non-recoverability of the catalyst as well as long reaction time and low efficiencies.^{23,24} Consequently, the improvement of easy, effective, and mild procedures with easily separable and reusable novel catalysts to overcome these shortcomings.

Pyrroles, as an important class of N-heterocyclic derivatives and significant building blocks, are one of the most important groups of heterocyclic compounds that have attracted considerable interest in organic production as well as in the drug exploration. They have been extensively utilized due to their promising pharmacological and therapeutic activities such as antianginal, antitumor antifungal, anti-oxidant, antibacterial, and anti-inflammatory.²⁵ Moreover, polysubstituted pyrroles owing to their enormous importance exhibited other medicinal

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applications. Because of the aforementioned merits of the polysubstituted pyrroles, to date various approaches have been improved to synthesize these kinds of heterocyclic compounds. These were synthesized through Paal–Knorr condensation reaction, Knorr reaction and Hantzsch reaction.²⁶ Herein we decided to study the preparation of polysubstituted pyrrole derivatives with various substituents from the reaction of aromatic aldehydes, ethyl acetoacetate, nitromethane and aniline under solvent-free conditions using Cu@imine/Fe₃O₄ MNPs as a new, eco-friendly, reusable and promising heterogeneous nanocatalyst. Although numerous methods have been utilized for the preparation of polysubstituted pyrrole derivatives, some of these strategies suffer from disadvantages such as use of high temperature, the need for excess amounts of the catalyst, long reaction times, and need for microwave or ultrasound irradiation. Therefore, it is necessary to develop an improved strategy for the preparation of polysubstituted pyrrole derivatives under mild reaction conditions.

Experimental

General

All the pure chemical substances were prepared from Fluka, Aldrich, and Merck chemical companies and applied with no further purification. Melting points of the substrate were measured with Electrothermal-9100 apparatus and uncorrected. Fourier transform infrared (FT-IR) spectra were recorded with a PerkinElmer PXI spectrometer in KBr wafers. X-ray diffraction (XRD) patterns of samples were taken on a Siemens D-5000 X-ray diffractometer (Germany) with Cu K α radiation. Magnetic susceptibility measurements were measured by a vibrating sample magnetometry (VSM; Lake Shore 7200 at 300 kVsm). Thermogravimetric analysis was examined by a PerkinElmer instrument under nitrogen atmosphere at a heating rate of 10 °C min⁻¹. Scanning electron microscope (SEM) images were recorded with a SEM-LEO 1430VP instrument. The chemical composition of synthesized nanoparticles was obtained with an energy dispersive X-ray spectroscopy (EDX) (ESEM, Philips, and XL30).

Catalyst synthesis

Preparation of Fe₃O₄. The magnetite nanoparticles were synthesized based on the reported chemical co-precipitation method, and the detailed process has been described herein-after. Briefly, 5.4 g of FeCl₃·6H₂O and 2 g of FeCl₂·4H₂O were dissolved in 40 mL of distilled water followed by adding 0.9 mL of concentrated HCl in a three-necked bottom (250 mL) under vigorous agitating and nitrogen protection. Next, 250 mL of 1.5 M NaOH solution was dropped into the resultant mixture in the absence of temperature under intense mechanical stirring. In the last step, the attained magnetic nanoparticles were isolated using a powerful magnet and rinsed several times with 200 mL deionized water. Eventually, the Fe₃O₄ nanoparticles dried under vacuum at 40 °C for 24 h.

Preparation of the Fe₃O₄@APTES. 1 g of Fe₃O₄ nanoparticles was dispersed poured in 50 mL dry toluene by ultrasonic

vibration; then, 2 mL of aminopropyltriethoxysilane (APTES) was added, and the obtained solution was mechanically agitated for 24 h under reflux conditions. At the end of the reaction, the Fe₃O₄@APTES nanoparticles were recovered from the solution using an external magnet, washed with deionized water and anhydrous ethanol to eliminate any unreacted chemicals, and dried under vacuum oven.

Preparation of the Fe₃O₄@APTES/AMTBA. To connect of 4-(4-amino-5-mercapto-4H-[1,2,4]H-1,2,4 triazol-3-yl)-benzoic acid (AMTBA) to the NH₂ functional group on the surface Fe₃O₄@APTES nanoparticles, dried Fe₃O₄@APTES (2 g) was poured in ethanol (100 mL) in a 250 mL round-bottom flask and then AMTBA (4 mmol) was poured in to this solution under dry nitrogen atmosphere. The reaction mixture was refluxed for 24 h and after that Fe₃O₄@APTES/AMTBA was separated by using an external magnet, washed thoroughly several times with ethanol and deionized water to remove the unreacted reagents, and finally dried at 80 °C in a vacuum oven.

Preparation of the imine@Fe₃O₄ MNPs. To prepare imine@Fe₃O₄ MNPs, 2 g of Fe₃O₄@APTES/AMTBA was poured in 20 mL of dry CH₂Cl₂ for 30 min using ultrasonication. After that, 4 mmol of 2-formylpyridine was added to the reaction solution and the obtained mixture was refluxed for 8 h. Finally, these precipitates were isolated using an external magnetic field, washed with 20 mL of water to eliminate the unreacted residue of the 2-formylpyridine, and next dried in a vacuum oven.

Preparation of Cu@imine/Fe₃O₄ MNPs. 2 g of imine@Fe₃O₄ MNPs was suspended in 50 mL of CH₂Cl₂ under a nitrogen atmosphere. Thereafter, CuCl₂ (0.5 mmol) was added to the reaction solution, and agitated for 24 h. After 24 h, the sediment formed was isolated by a strong magnet, rinsed twice with ethanol (15 mL) to remove unreacted metal precursors and dried under the reduced pressure to obtain a pure product. All stages of the Cu@imine/Fe₃O₄ MNPs synthesis are demonstrated in Scheme 1.

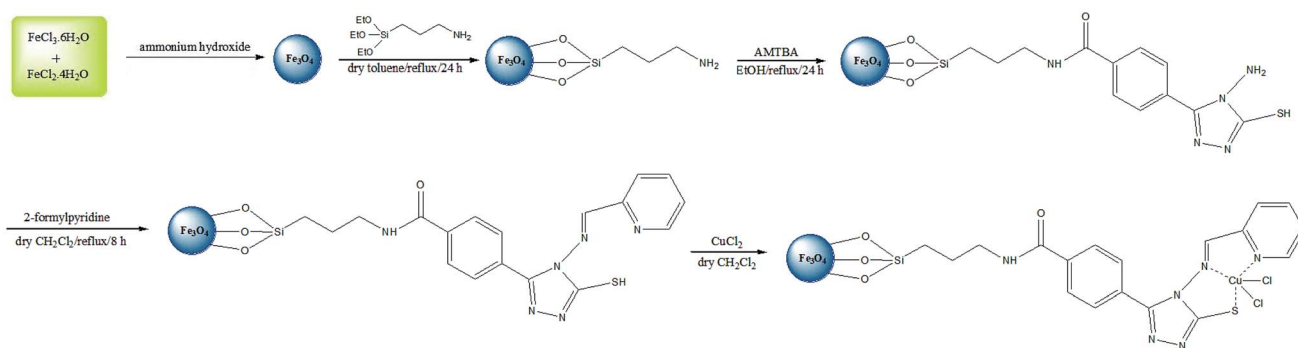
Results and discussion

Catalyst characterization

FTIR analysis of Cu@imine/Fe₃O₄ MNPs. The infrared spectra corresponding to the Fe₃O₄ and Cu@imine/Fe₃O₄ MNPs samples in the area of 400–4000 cm⁻¹ are demonstrated in Fig. 1. The bands at 586 cm⁻¹ in the Fe₃O₄ are associated to the stretching vibration for the Fe–O bond while the absorption peak at the 3405 cm⁻¹ region can be dedicated to the water molecule OH stretching modes. Also, the absorption peaks at 1617 cm⁻¹ is related to the bending vibrations of the water molecule HO–H bonds. In the FTIR spectrum of Cu@imine/Fe₃O₄ MNPs catalyst, the bands between 1400–1600 cm⁻¹ and the absorption peaks at 1045 cm⁻¹ region can be respectively attributed to the C=C and Si–O groups. The bands at 1635, 2584 and 2987 cm⁻¹ have been attributed to C=N, S–H and C–H stretching vibrations, respectively.

Thermal analysis of Cu@imine/Fe₃O₄ MNPs. TGA curves of Fe₃O₄, Fe₃O₄@APTES, Cu@imine/Fe₃O₄ MNPs was also investigated by thermogravimetric analyzer (Fig. 2). The weight loss





Scheme 1 All steps of the Cu@imine/Fe₃O₄ MNPs synthesis.

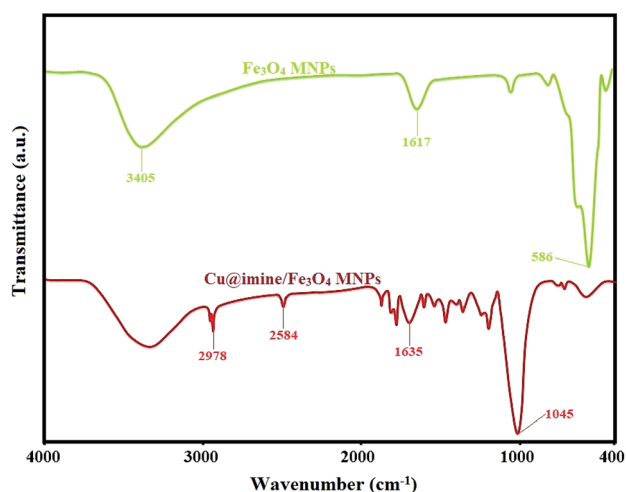


Fig. 1 FT-IR spectra of Fe₃O₄ and Cu@imine/Fe₃O₄ MNPs.

due to evaporation of physically adsorbed H₂O molecules in the temperature range below 220 °C was observed for all samples. Two weight loss stages were observed in thermogram for

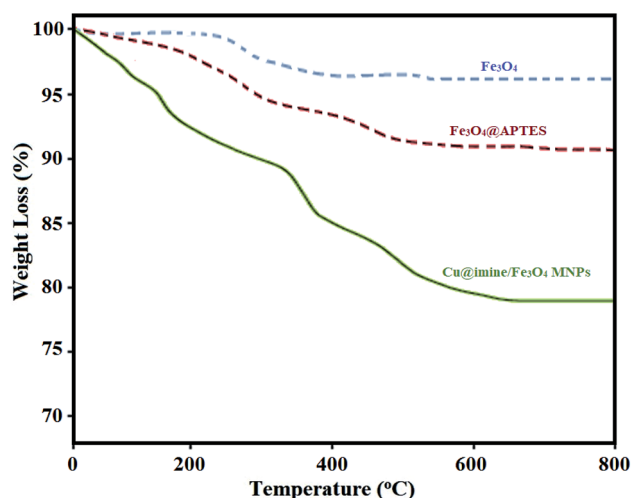


Fig. 2 TGA thermogram of Fe₃O₄, Fe₃O₄@APTES, Cu@imine/Fe₃O₄ MNPs.

Fe₃O₄@APTES, one below 200 °C which can be related to evaporating water from the MNPs surface; however the other one up to 430 °C was ascribed to the decomposition of APTES. Cu@imine/Fe₃O₄ MNPs have three weight loss stages; the first weight loss up to 180 °C, which is related to evaporation of H₂O molecules connected to the MNPs surface. The second and third weight loss up to 370 and 520 °C, respectively, associated to the decomposition of functionalized groups on MNPs surface. These weight losses exhibits the existence of functional groups grafted to the MNPs surface.

VSM analysis of Cu@imine/Fe₃O₄ MNPs. Fig. 3 shows the magnetization measurements of Fe₃O₄ MNPs and Cu@imine/Fe₃O₄ MNPs, determined by VSM technique. The saturation magnetization (M_s) is observed to be 35.94 emu g⁻¹ for Cu@imine/Fe₃O₄ MNPs, less than the Fe₃O₄ MNPs (62.59 emu g⁻¹). This difference suggests that functional groups are attached on the surface of Fe₃O₄ nanoparticles.

XRD analysis of Cu@imine/Fe₃O₄ MNPs. The XRD patterns corresponding to the synthesized Fe₃O₄, imine@Fe₃O₄ MNPs, Cu@imine/Fe₃O₄ MNPs are shown in Fig. 4. The characteristic broad diffraction peaks shown in high-angle XRD pattern can be corresponded to cubic spinel magnetite Fe₃O₄ structure, which shows diffraction peaks, corresponding to (220), (311), (400), (422), (511) and (440). The relative intensities and position of imine@Fe₃O₄ MNPs and Cu@imine/Fe₃O₄ MNPs peaks agree well with the Fe₃O₄ MNPs, representing retention of the crystalline cubic spinel structure over immobilization of Fe₃O₄ MNPs.

EDX analysis of Cu@imine/Fe₃O₄ MNPs. Based on the EDX spectrum (Fig. 5), the existence of copper in the Cu@imine/Fe₃O₄ MNPs catalyst structure represents that copper coordinates to the functional groups of imine@Fe₃O₄ MNPs. The EDX spectrum also illustrates other elements including Fe, S, Si, C and N which are existent in the catalyst substrate.

This survey is aimed to report the efficient and rapid production of polysubstituted pyrrole and 1,2,4,5-tetrasubstituted imidazoles derivatives using lower loading of the very recently introduced nanocatalyst, *i.e.*, Cu@imine/Fe₃O₄ MNPs, under solvent-free conditions. The catalytic activity of Cu@imine/Fe₃O₄ MNPs was investigating to report the efficient and rapid production of polysubstituted pyrrole derivatives under solvent-free conditions. To achieve the optimum conditions



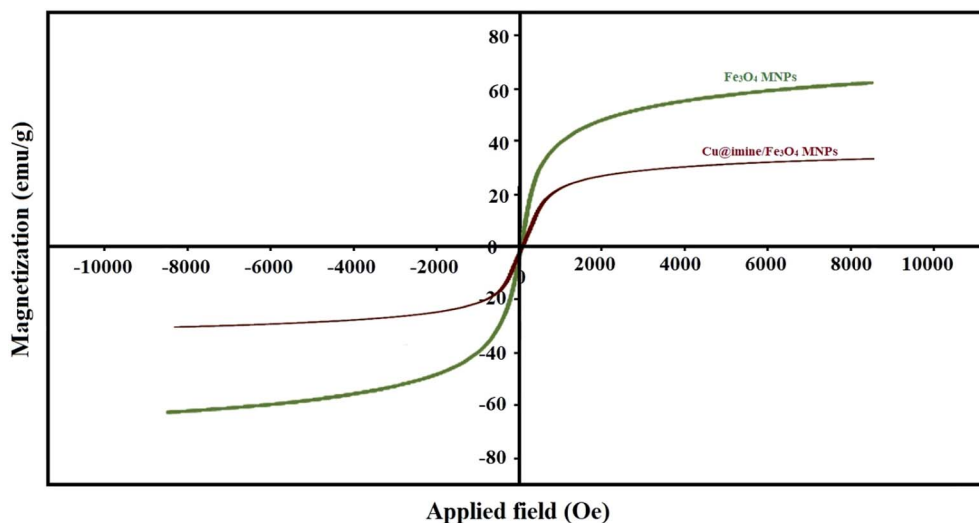


Fig. 3 Magnetization curves of Fe_3O_4 and $\text{Cu@imine/Fe}_3\text{O}_4$ MNPs.

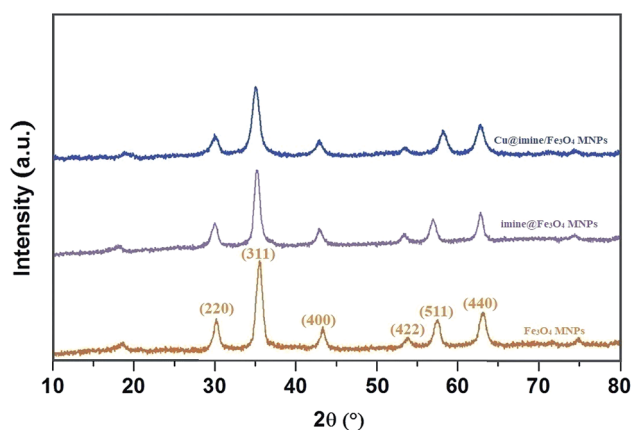


Fig. 4 The XRD patterns of Fe_3O_4 , imine@ Fe_3O_4 MNPs, $\text{Cu@imine/Fe}_3\text{O}_4$ MNPs.

especially in agreement with green chemistry, the reaction between ethyl acetoacetate, nitromethane, benzaldehyde, and aniline with 1 : 1 : 1 : 1 molar ratios were selected as the model

reactions for synthesizing polysubstituted pyrrole to examine the impacts of the amount of used catalyst, solvent, and temperature. The results of the optimization reactions were presented in Table 1. The outcomes were evaluated qualitatively using thin layer chromatography (TLC). As observed, in polar organic solvents such as ethanol and acetonitrile, under reflux conditions, the reaction yield was good in long reaction times (Table 1, Entries 1–2). In order to improve the eco-friendly of the reaction conditions, the reaction was occurred in water, but this solvent wasn't observed to be effective and only 37% of product was achieved after 70 min (Table 1, Entry 3). Finally, the reaction was performed in non-polar organic solvents such as dichloromethane, chloroform and tetrahydrofuran under reflux conditions, but none of these solvents were observed to be impressive (Table 1, Entries 4–6). The outcomes presented that when the reaction was performed under solvent-free conditions at 100 °C, the target product afforded in best optimal conditions in terms of rate and yield (Table 1, Entry 7). Due to the key role of the catalyst amount as a significant factor for a reaction, we studied the catalytic performance of the $\text{Cu@imine/Fe}_3\text{O}_4$

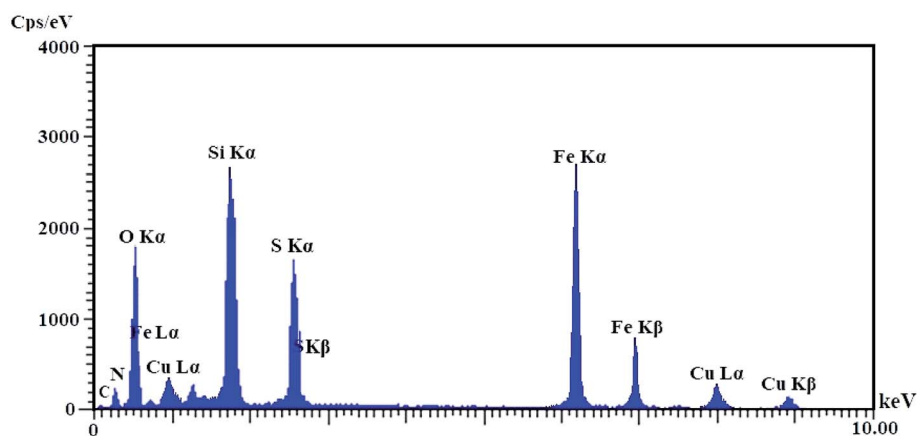
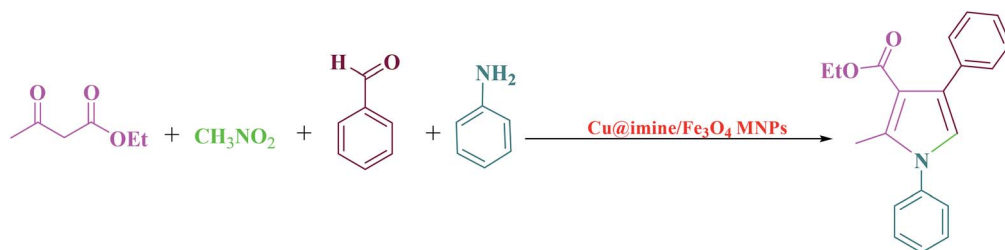


Fig. 5 EDX spectra of $\text{Cu@imine/Fe}_3\text{O}_4$ MNPs.



Table 1 Optimization corresponding to the four-component reaction for ethyl acetoacetate, nitromethane, benzaldehyde, and aniline under various conditions^a

Entry	Solvent	Catalyst (mol%)	Temp.	Time (min)	Yield ^b (%)
1	Ethanol	0.36	Reflux	50	79
2	Acetonitrile	0.36	Reflux	50	71
3	Water	0.36	Reflux	70	37
4	Dichloromethane	0.36	Reflux	70	52
5	Chloroform	0.36	Reflux	70	56
6	Tetrahydrofuran	0.36	Reflux	70	49
7	Solvent-free	0.36	100 °C	25	95
8	Solvent-free	0.24	100 °C	45	76
9	Solvent-free	0.48	100 °C	30	94
10	Solvent-free	—	100 °C	70	Trace
11	Solvent-free	0.36	25 °C	70	31
12	Solvent-free	0.36	60 °C	40	82
13	Solvent-free	0.36	70 °C	35	85
14	Solvent-free	0.36	80 °C	35	88
15	Solvent-free	0.36	90 °C	30	92
16	Solvent-free	0.36	110 °C	25	93
17	Solvent-free	0.36	120 °C	25	91
18	Solvent-free	BF ₃ –SiO ₂ /0.36	100 °C	120	71
19	Solvent-free	MgCl ₂ /0.36	100 °C	90	59
20	Solvent-free	SbCl ₅ –SiO ₂ /0.36	100 °C	100	62

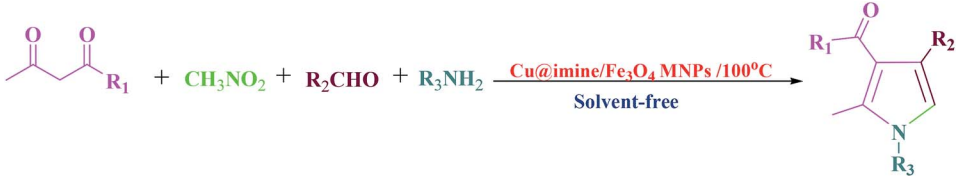
^a Reaction conditions: ethyl acetoacetate (1 mmol), nitromethane (1 mmol), benzaldehyde (1 mmol), aniline (1 mmol), and the required amount of catalyst. ^b The yields represent the separated product.

MNPs using 0.24–0.48 mol% of the catalyst under solvent-free conditions (Table 2, Entries 7–9), and the yield has increased remarkably in the existence of 0.36 mol% of catalyst, whereas the reaction time has decreased (Table 2, Entry 7). By reducing the value of catalyst to 0.24 mol%, the performance of the product was dropped (Table 1, Entry 8). It was observed that the reaction yield was remained unchanged even in quantities more than the optimum catalyst amount, and 0.48 mol% of the catalyst could not cause the obvious increase in yield of the product (Table 1, Entry 9). When the reaction was surveyed without using the catalyst, no formation of product was found in this case (Table 1, Entry 10). To complete the reaction, effect of temperature on the rate of model reaction was investigated by diverse temperatures from room temperature to 120 °C under solvent-free conditions (Table 1, Entries 11–17). It was observed that the yield of the product increases significantly by increasing the temperature from room temperature to 100 °C (Table 1, Entries 7 and 11–15), but a further enhance in temperature did not demonstrate any significant impact on the product yield (Table 1, Entries 16–17). It should be noted that when the reaction was carried out in the presence of metal

halides (Table 3, Entries 18–20), low to moderate yields of the desired product was obtained.

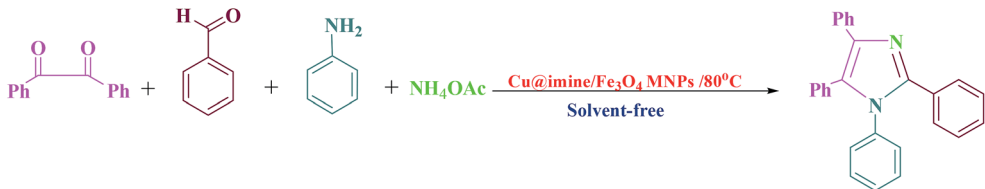
After optimization of the reaction conditions and to establish the effectiveness and the acceptability of the technique, we explored the scope and generalization of the reaction with a wide range of aldehydes and amines including benzylamine, aniline and substituted anilines under the optimal conditions. The outcomes were listed in Table 2. It was observed that under similar conditions, a variety of benzaldehydes containing electron-withdrawing and electron-donating groups, such as H, CH₃, CH₃O, NO₂ and Cl in the *para* positions of the benzene ring as well as heteroaromatic aldehydes in the existence of methyl acetoacetate, nitromethane, benzylamine, and the required amount of the Cu@imine/Fe₃O₄ MNPs, simply converted to the resulted products in short reaction times with good to excellent isolated yields (Table 2, Entries 1–7). Also, using above aromatic aldehydes, ethyl acetoacetate, nitromethane and aniline the target products were achieved in good to high separated yields (Table 2, Entries 8–11). Various aromatic aldehydes containing electron-withdrawing and electron-donating groups (H, CH₃, CH₃O, and Cl) and anilines



Table 2 Preparation of polysubstituted pyrrole derivatives using Cu@imine/Fe₃O₄ MNPs as a catalyst^a


Entry	R ₂ CHO (3)	R ₁	R ₃	Product	Time (min)	Yield (%)	TON ^b	TOF ^c (h ⁻¹)	Mp (obsd) (°C)	Mp (lit) (°C)
1	C ₆ H ₅ –	OMe	Ph–CH ₂ –	5a	20	93	258	774	Oil	Oil ²⁷
2	4-Cl–C ₆ H ₄ –	OMe	Ph–CH ₂ –	5b	15	95	264	1056	Oil	Oil ²⁸
3	4-NO ₂ –C ₆ H ₄ –	OMe	Ph–CH ₂ –	5c	15	97	269	1076	Oil	Oil ²⁸
4	4-CH ₃ –C ₆ H ₄ –	OMe	Ph–CH ₂ –	5d	20	91	253	759	Oil	Oil ²⁸
5	4-OCH ₃ –C ₆ H ₄ –	OMe	Ph–CH ₂ –	5e	15	92	256	1024	Oil	Oil ²⁷
6	2-Furyl–	OMe	Ph–CH ₂ –	5f	20	94	261	783	128–131	130–132 (ref. 30)
7	2-Thienyl–	OMe	Ph–CH ₂ –	5g	15	96	266	1064	120–123	122–124 (ref. 30)
8	4-Cl–C ₆ H ₄ –	OEt	Ph–CH ₂ –	5h	15	93	258	1032	Oil	Oil ²⁸
9	4-NO ₂ –C ₆ H ₄ –	OEt	Ph–CH ₂ –	5i	20	95	264	792	Oil	Oil ²⁹
10	4-CH ₃ –C ₆ H ₄ –	OEt	Ph–CH ₂ –	5j	20	91	253	759	Oil	Oil ²⁷
11	4-OCH ₃ –C ₆ H ₄ –	OEt	Ph–CH ₂ –	5k	20	89	247	741	Oil	Oil ²⁹
12	C ₆ H ₅ –	OEt	C ₆ H ₅ –	5l	25	95	264	634	Oil	Oil ³¹
13	4-Cl–C ₆ H ₄ –	OEt	C ₆ H ₅ –	5m	20	90	250	750	126–128	127–129 (ref. 32)
14	4-CH ₃ –C ₆ H ₄ –	OEt	4-Cl–C ₆ H ₄ –	5n	20	91	253	759	119–122	120–122 (ref. 32)
15	4-Cl–C ₆ H ₄ –	OEt	4-OCH ₃ –C ₆ H ₄ –	5o	20	93	258	774	128–131	130–132 (ref. 28)
16	4-OCH ₃ –C ₆ H ₄ –	OEt	C ₆ H ₅ –	5p	20	91	253	759	106–109	108–110 (ref. 32)

^a Reaction conditions: ethyl acetoacetate (1 mmol), nitromethane (1 mmol), benzaldehyde (1 mmol), and aniline (1 mmol) and Cu@imine/Fe₃O₄ MNPs (0.36 mol%) at 100 °C. ^b Number of moles of product produced from 1 mole of catalyst. ^c TON per unit of time.

Table 3 Optimization of the four-component reaction of benzil, benzaldehyde, benzylamine and ammonium acetate under various conditions^a


Entry	Solvent	Catalyst (mol%)	Temp.	Time (min)	Yield ^b (%)
1	Ethanol	0.36	Reflux	55	69
2	Acetonitrile	0.36	Reflux	55	61
3	Water	0.36	Reflux	80	78
4	Dichloromethane	0.36	Reflux	70	45
5	Chloroform	0.36	Reflux	70	49
6	Tetrahydrofuran	0.36	Reflux	70	43
7	Solvent-free	0.36	80 °C	35	95
8	Solvent-free	0.24	80 °C	55	79
9	Solvent-free	0.48	80 °C	35	92
10	Solvent-free	0.36	25 °C	120	35
11	Solvent-free	0.36	60 °C	65	69
12	Solvent-free	0.36	70 °C	45	75
13	Solvent-free	0.36	90 °C	35	93
14	Solvent-free	0.36	100 °C	35	91
15w	Solvent-free	BF ₃ –SiO ₂ /0.36	80 °C	150	68
16	Solvent-free	MgCl ₂ /0.36	80 °C	80	56
17	Solvent-free	SbCl ₅ –SiO ₂ /0.36	80 °C	150	65

^a Reaction conditions: benzil (1 mmol), benzaldehyde (1 mmol), benzylamine (1 mmol), ammonium acetate (4 mmol), and the required amount of catalyst. ^b The yields refer to the separated product.



with the above-mentioned substituted electron-withdrawing and electron-donating groups in the *para* positions of the benzene ring were efficiently reacted with ethyl acetoacetate and nitromethane in excellent yields and times (Table 2, Entries 12–16).

To obtain the optimized reaction conditions of the production of 1,2,4,5-tetrasubstituted imidazoles derivatives, the reaction between benzil, benzylamine, benzaldehyde and ammonium acetate with 1 : 1 : 1 : 4 molar ratios using 0.36 mol% of catalyst was examined in the existence of diverse solvents, the concentration of used catalyst, and temperatures. To find the solvent effect, the reaction was carried out in the existence of various polar and non-polar solvents such as ethanol, acetonitrile, water, dichloromethane, chloroform and tetrahydrofuran. When water was applied as a solvent, the reaction proceeded smoothly for producing the desired product (**5a**) in good yield (Table 3, Entry 3). The model reaction in the rest of the solvents mentioned above demanded long reaction times, and provided the product **5a** in poor to moderate yields (Table 1, Entries 1–2 and 4–6). Interestingly, none of these solvents were impressive and the model reaction proceeded rapidly to obtain the desired product in excellent yield 95% under solvent-free conditions (Table 3, Entry 7). Over optimizing the reaction conditions, the impact of catalytic activity of the Cu@imine/Fe₃O₄ MNPs on the model reaction was also studied at 0.24, 0.36, and 0.48 mol% to test their efficiency under solvent-free conditions (Table 3, Entries 7–9). It was observed that the product yield has increased in the existence of 0.36 mol% of catalyst (Table 3, Entry 7). Hence, 0.36 mol% was chosen as optimum catalyst amount for the reaction. Following choice of the optimal concentration of catalyst, the effect of the temperature on yield of the model reaction was evaluated by diverse temperatures

from room temperature to 100 °C in the existence of required amount of catalysts under solvent-free conditions (Table 3, Entries 7 and 10–14). It was found that by raising the reaction temperature, yield of the product increases. Maximum yield (95%) was obtained when the temperature of 80 °C was applied to the reaction system (Table 3, Entry 7). Further rise in the temperature had no considerable impact on the product yield (Table 3, Entries 13–14). It should be mentioned that the product yield has decreased remarkably at temperatures below 80 °C, whereas the reaction time has increased (Table 3, Entries 10–12). Therefore, as shown in Table 3, using 0.36 mol% of Cu@imine/Fe₃O₄ MNPs under solvent-free and 80 °C gave the maximum yield of the target product in the shortest time. It is noteworthy that when the reaction was carried out in the presence of metal halides (Table 3, Entries 15–17), low to moderate yields of the desired product was obtained.

With the optimized condition in hand, we explored the scope and generalization of the catalyst over a variety of aldehydes and amines as substrates for synthesizing 1,2,4,5-tetrasubstituted imidazoles. The outcomes were presented in Table 4. These results illustrate that the reactions are similarly facile with both electron-withdrawing and electron-donating groups, such as H, CH₃, CH₃O, NO₂ and Cl in the meta and *para* positions of the benzene ring and amines, causing good to high yields of the relating imidazoles.

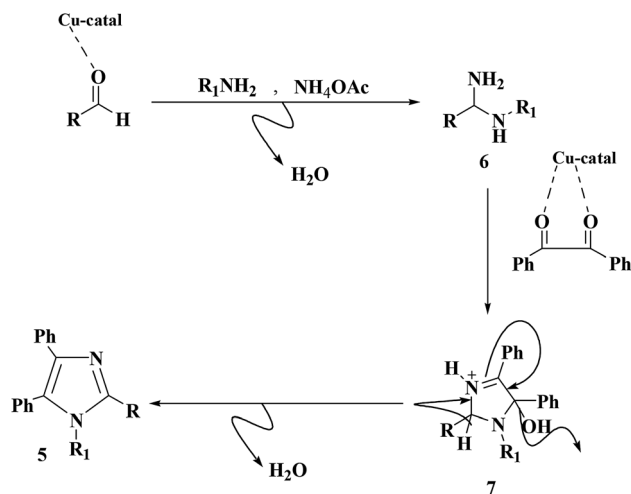
A proposed mechanism of the synthesis of replaced imidazoles catalyzed by Cu@imine/Fe₃O₄ MNPs is depicted in Scheme 2. In the beginning, carbonyl oxygen of aldehyde is coordinated to copper center on the nanocomposites surface to generate a better electrophilic species which more readily reacts with amine to give diamine intermediate **6**. The next step involves formation of imino intermediate **7** by condensing

Table 4 Preparation of 1,2,4,5-tetrasubstituted imidazoles derivatives using Cu@imine/Fe₃O₄ MNPs as a catalyst^a

Entry	RCHO (2)	R ₁	Product	Time (min)	Yield (%)	TON ^b	TOF ^c (h ⁻¹)	Mp (obsd) (°C)	Mp (lit) (°C)
1	C ₆ H ₅ –	C ₆ H ₅ –	5a	35	95	264	453	215–217	216–218 (ref. 33)
2	4-Cl–C ₆ H ₄ –	C ₆ H ₅ –	5b	30	95	264	528	150–152	149–151 (ref. 33)
3	3-NO ₂ –C ₆ H ₄ –	C ₆ H ₅ –	5c	30	97	269	538	254–257	255 (ref. 34)
4	4-CH ₃ –C ₆ H ₄ –	C ₆ H ₅ –	5d	35	93	258	442	186–188	185–188 (ref. 33)
5	4-OCH ₃ –C ₆ H ₄ –	C ₆ H ₅ –	5e	35	92	256	439	170–172	171–173 (ref. 35)
6	3,4-(OCH ₃) ₂ –C ₆ H ₄ –	C ₆ H ₅ –	5f	35	94	261	447	173–176	175–178 (ref. 35)
7	C ₆ H ₅ –	<i>n</i> -Pr–NH ₂	5g	30	96	266	532	88–91	87–89 (ref. 36)
8	4-Cl–C ₆ H ₄ –	<i>n</i> -Pr–NH ₂	5h	20	97	269	807	84–86	87–85 (ref. 36)
9	4-NO ₂ –C ₆ H ₄ –	<i>n</i> -Pr–NH ₂	5i	20	98	272	816	159–162	161–163 (ref. 37)
10	3-NO ₂ –C ₆ H ₄ –	<i>n</i> -Pr–NH ₂	5j	20	98	272	816	139–141	141–142 (ref. 37)
11	4-CH ₃ –C ₆ H ₄ –	<i>n</i> -Pr–NH ₂	5k	25	93	258	619	80–83	78–83 (ref. 36)
12	C ₆ H ₅ –	4-EtC ₆ H ₄	5l	35	96	266	456	141–143	142–144 (ref. 37)
13	3-NO ₂ –C ₆ H ₄ –	4-EtC ₆ H ₄	5m	25	98	272	653	130–133	131–132 (ref. 37)
14	4-CH ₃ –C ₆ H ₄ –	4-EtC ₆ H ₄	5n	35	94	261	447	157–160	159–161 (ref. 37)

^a Reaction conditions: benzil (1 mmol), benzaldehyde (1 mmol), benzylamine (1 mmol), ammonium acetate (4 mmol) and Cu@imine/Fe₃O₄ MNPs (0.36 mol%) at 80 °C. ^b Number of moles of product produced from 1 mole of catalyst. ^c TON per unit of time.



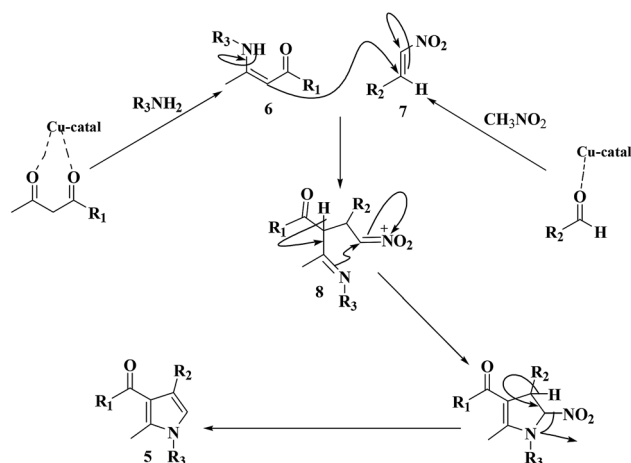


Scheme 2 A plausible mechanism of the creation of 1,2,4,5-tetra-substituted imidazoles derivatives in the existence of Cu@imine/Fe₃O₄ MNPs under solvent-free conditions.

diamine with 1,2-diketone, followed by dehydrating the relating products 5.

The mechanism corresponding to production of greatly substituted pyrrole mediated with the Cu@imine/Fe₃O₄ MNPs catalyst was presented in Scheme 3. The reaction occurs *via* initial production of the enamine 6 by nucleophilic addition of amine to α -ketoester followed by dehydration. Then, nitromethane reacts with aldehyde to generate intermediate 7. Subsequently, the enamine 6 reacts with intermediate 7 to afford Michael adduct 8. Following intramolecular nucleophilic cyclization by dehydration delivers the desired pyrrole.

Recycling and recoverability of the Cu@imine/Fe₃O₄ MNPs were evaluated in the reaction between ethylacetoacetate, nitromethane, benzaldehyde, and aniline. At the end of the reaction, 10 mL hot mixture of ethyl acetate and ethanol (4 : 10 ratio) was poured in to the reaction solution to dissolve the raw product. Then, the catalyst was isolated from the reaction



Scheme 3 A plausible mechanisms for the creation of polysubstituted pyrrole derivatives in the existence of Cu@imine/Fe₃O₄ MNPs under solvent-free conditions.

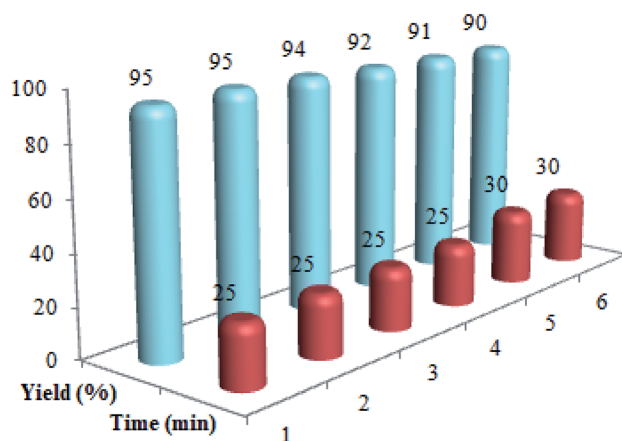


Fig. 6 The recycling of Cu@imine/Fe₃O₄ MNPs in the preparation of polysubstituted pyrrole derivatives.

solution by a permanent magnet, washed with chloroform, dried and recovered for a minimum of six runs (Fig. 6).

To evaluate recycling and recoverability of the Cu@imine/Fe₃O₄ MNPs, after six periods of recycling, this catalyst was reused in the reaction of benzil, benzaldehyde, benzylamine and ammonium acetate under the optimal reaction conditions (Fig. 7). To achieve this purpose, at the end of the reaction, a hot mixture of ethyl acetate and ethanol (4 : 10 ratio) was added to the reaction solution and then the catalyst was removed simply by means of a permanent external magnet. The recovered Cu@imine/Fe₃O₄ MNPs was rinsed with chloroform, dried, and recovered for alternative reaction.

Table 5 shows the efficiency of Cu@imine/Fe₃O₄ MNPs as the novel heterogeneous magnetic nanocatalyst in preparing 1-benzyl-4-(4-chloro-phenyl)-2-methyl-1H-pyrrole-3-carboxylic acid methyl ester and 1,2,4,5-tetraphenyl-1H-imidazole compared with those obtained by some of the other reported homogeneous and heterogeneous catalysts. As clearly shown in Table 5, although all the reported catalysts can catalyze the reactions, most of them suffer from one or more disadvantages, such as long reaction times, low yields of the products, use of toxic catalyst, tedious work-up procedures, and high catalyst

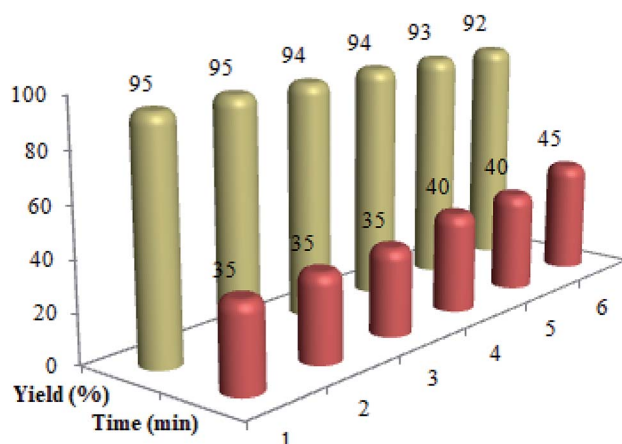
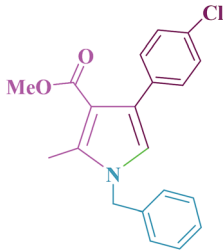
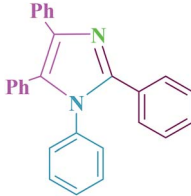


Fig. 7 The recycling of Cu@imine/Fe₃O₄ MNPs in the synthesis of 1,2,4,5-tetrasubstituted imidazoles derivatives.



Table 5 Comparison of the current methods with other reported strategies for synthesizing 1-benzyl-4-(4-chloro-phenyl)-2-methyl-1H-pyrrole-3-carboxylic acid methyl ester and 1,2,4,5-tetraphenyl-1H-imidazole derivatives

Entry	Catalyst	Temp (°C)	Time (min)	Yield (%)	Ref.
					
1	Nickel(II) chloride hexahydrate/solvent-free	rt	480	78	30
2	Nickel ferrite nanoparticles/solvent-free	100	240	96	28
3	Nano copper oxide/solvent-free	100	720	78	29
4	Heterogenized tungsten complex/solvent-free	100	270	85	27
5	Amberlyst-15/solvent-free	110	30	91	25
6	SbCu@silica-Fe ₃ O ₄ /solvent-free	rt	25	88	32
7	Cu@imine/Fe ₃ O ₄ MNPs/solvent-free	100	25	95	This work
					
1	BF ₃ /SiO ₂ /solvent-free	140	120	92	39
2	TiCl ₄ /SiO ₂ /solvent-free	110	190	75	40
3	Glycerol as solvent	90	180	96	41
4	PEG-400 as a solvent	110	360	86	38
5	NaH ₂ PO ₄ /solvent-free	120	35	90	42
6	Fe ₃ O ₄ -PEG-Cu/solvent-free	110	55	96	43
7	γ-Fe ₂ O ₃ @TiO ₂ -EG-Cu(II)/solvent-free	100	20	98	44
8	Cu@imine/Fe ₃ O ₄ MNPs/solvent-free	80	35	95	This work

loading. For these reasons, the catalytic behaviour of our catalytic system is significant in terms of easy work-up procedures, low catalyst loading, low reaction times, and simple recovery of the catalyst.

Conclusion

We have described a simple, facile, and effective process for the eco-friendly and fast production of biologically active poly-substituted pyrroles and 1,2,4,5-tetrasubstituted imidazoles derivatives. The reaction system was significantly affected by catalyst loading and reaction temperature. The catalyst could be recovered and reused at least six times without any significant decrease in its activity. The attractive properties of this procedure are short reaction times, simple workup, high yields, reusability and re-activity of the catalyst as well as simple purification of the products.

Conflicts of interest

There are no conflicts to declare.

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