



Cite this: RSC Adv., 2019, 9, 9937

Received 7th March 2019
Accepted 25th March 2019

DOI: 10.1039/c9ra01742a

rsc.li/rsc-advances

Formation mechanism of nanoporous silver during dealloying with ultrasonic irradiation

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Nanoporous silver (NPS) with an extreme coarsened 3-dimensional bi-continuous ligament and nanopore structure could be prepared by chemical dealloying with high-intensity ultrasonic irradiation (UI). The formation mechanism of NPS dealloying with UI was different from NPS obtained through free corrosion. It evolved into NPS with a new lump forming-disintegrating mechanism. Ultrasonic irradiation had strong effects on the dealloying process of NPS. The stirring effect produced by ultrasonic vibration could promote the corrosion of Cu and facilitate the diffusion of Ag atoms. Therefore, the coarsening rate of the ligament was increased significantly. Dealloying assisted by UI could generate an extremely coarsened microstructure of which ligament and pore sizes were much larger than those obtained from free corrosion dealloying.

Introduction

Due to their unique structure-related properties, nanoporous metals that can exhibit a bi-continuous interpenetrating ligament/channel structure have attracted intense research interest in recent years in many fields such as catalysis, sensing, optical, and actuation.^{1–3} There are many methods to prepare nanoporous metals up to now. The template method which is the process of depositing the metallic material into the hole or surface of the template by physical or chemical methods and removing the template to obtain the nanometer material with the standard morphology and size of the template was a common method used for the fabrication of nanoporous materials.^{4–6} It offers convenient control of the microstructure (mainly mean pore size) by precisely changing the original template. However, this method shows significant disadvantages because it is technically difficult and time-consuming. Dealloying which means the selective dissolution of one or more components out of an alloy or a compound has been reported to be a simple and effective strategy for the fabrication of nanoporous metals with homogeneous porous structure.^{7,8}

To date, dealloying has been applied to fabricate nanoporous metals from different alloy systems, such as Ag–Al, Al–Cu, Au–Cu, Au–Ag, Cu–Zr alloys.^{9–13} In this research, a comparative commercial Ag–Cu alloy system was chosen to be investigated. According to existing research, the microstructure of the as-dealloyed sample could be affected by many factors, such as

nature of the precursor alloy, composition of the precursor alloy, dealloying time, dealloying temperature, and dealloying solution.^{14–17} A slight change in modulation of these factors may lead to a remarkable difference in the microstructure of nanoporous metals. Based on this effect, an efficient dealloying process and a tunable nanoporous structure were presented.^{18–20}

However, the dealloying process can be more efficient and the microstructure will be more desirable when the dealloying process is carried out with the assistant treatments.^{16,21} Nanoporous Ag micro-particles have been prepared by ultrasonic assisted dealloying Mg–Ag alloy.²² The length scale of the ligament is not only sensitive to the concentration of the dealloying solution but also to the intensity of ultrasonic irradiation. The influence of a magnetic field on the dealloying of Al–25Ag has also been elucidated.²³ According to the experimental results, an external magnetic field accelerates the dealloying process and the microstructure obtained from dealloying with magnetic treatment is more homogeneous than those from free corrosion dealloying. Moreover, the nanoporous Ag foam was fabricated by dealloying Ag–Al thin film through supercritical fluid corrosion.⁴ Recently, a three-dimensional bicontinuous nanoporous materials could be obtained by vapor phase dealloying.²⁴

As we all know, ultrasonic vibration has been widely applied for purification, degassing, refining metallic melts and fabricating particulate-reinforced metal matrix composites.^{25–28} In the present work, the nanoporous Ag with an extreme coarsened microstructure is obtained by chemical dealloying with ultrasonic irradiation at different frequencies. Therefore, a time saving method for preparing nanoporous Ag was presented and a lower acid concentration was acquired to obtain NPS. It could avoid time consuming and reduce environment pollution simultaneously. Not only the surface diffusion during

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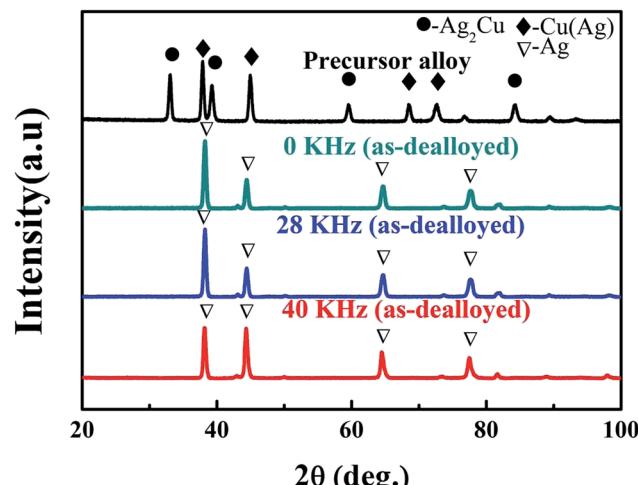


Fig. 1 XRD results of $\text{Ag}_{25}\text{Cu}_{75}$ precursor alloy and ribbons dealloyed without or with UI at different frequencies (28, 40 KHz) for 20 min.

dealloying process with ultrasonic irradiation (UI) is studied, but also the effect of UI on the resulting nanoporous structure is discussed. Extremely large variations between the microstructure of the nanoporous Ag dealloyed under free corrosion and nanoporous Ag dealloyed with UI are observed. In addition, different UI frequencies would have different effects on the dealloying process and diffusion of Ag atoms. Therefore, the relationship between the nanoporous Ag ligament and other

dealloying parameters is elucidated. A new formation mechanism of nanoporous Ag is also presented.

Experiment

$\text{Ag}-\text{Cu}$ alloy with the nominal composition of $\text{Ag}_{25}\text{Cu}_{75}$ (at%) was prepared by melting pure Ag (99.95 wt%) and pure Cu (99.95 wt%) in a graphite crucible using a high-frequency induction furnace under an argon atmosphere. The pre-alloyed $\text{Ag}-\text{Cu}$ alloy ingots were re-melted in a highly pure argon atmosphere and were subsequently injected onto a copper roller (cooling rate: $50\text{--}300\text{ }^{\circ}\text{C s}^{-1}$; rotational speed: 2500 rpm). The ingots were then rapidly solidified into ribbons, which are normally $40\text{--}60\text{ }\mu\text{m}$ thick, 10–15 mm wide and several centimeters long.

Dealloying of the samples was then performed at $45\text{ }^{\circ}\text{C}$ under different conditions: (1) under free corrosion in a 10 wt% HNO_3 aqueous solution for 5, 10 and 20 min; (2) with UI at the constant frequency of 28 kHz in a 10 wt% HNO_3 aqueous solution for 5, 10 and 20 min; and (3) with UI at the frequency of 40 kHz in a 10 wt% HNO_3 aqueous solution for 5, 10 and 20 min. The dealloyed samples were then rinsed three or four times sequentially using alcohol and distilled water. Under the effect of UI, the temperature of dealloying system would increase $1\text{--}2\text{ }^{\circ}\text{C}$.

Phase patterns of the $\text{Ag}-\text{Cu}$ precursor alloys and dealloyed samples were analyzed by X-ray diffraction (XRD, MAXIMA-7000) using $\text{Cu K}\alpha$ radiation with a scanning speed of

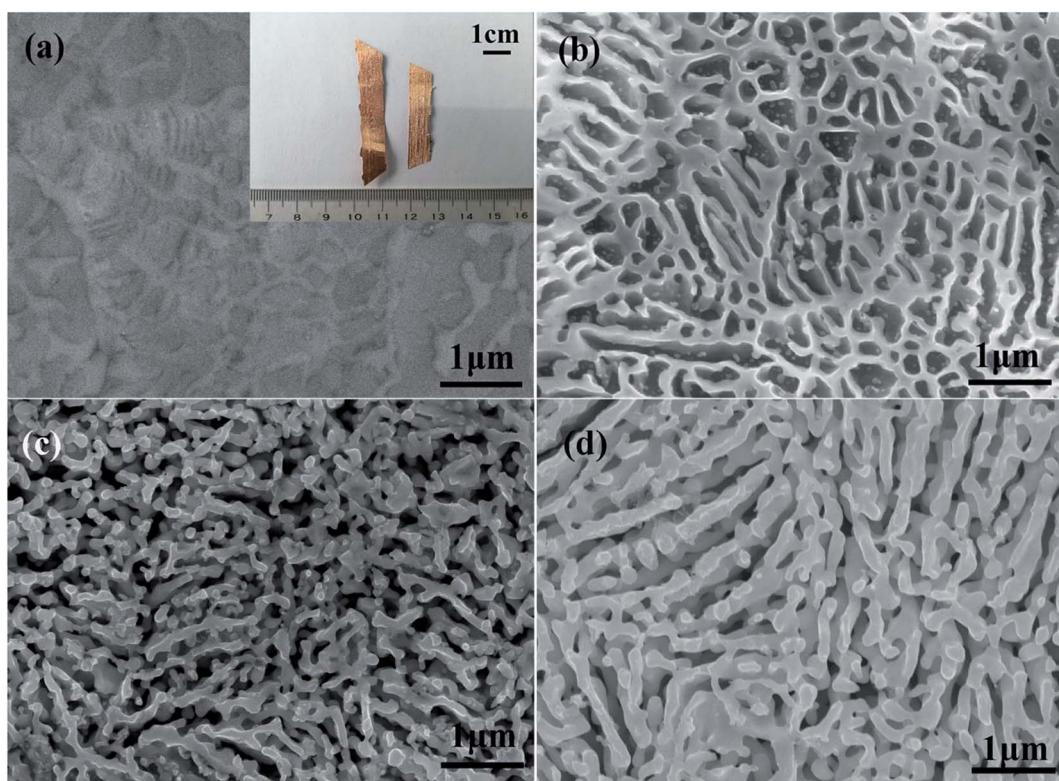


Fig. 2 SEM images of the $\text{Ag}_{25}\text{Cu}_{75}$ samples (plan view) dealloyed without UI for different periods of time: (a) 0 min, (b) 5 min, (c) 10 min, and (d) 20 min. The inset of (a) is the photograph of $\text{Ag}_{25}\text{Cu}_{75}$ initial alloy.



2° min $^{-1}$. The 2θ scans were performed between 20° and 90° . Microstructures and elemental content of the Ag–Cu precursor alloys and the dealloyed samples were characterized by transmission electron microscopy (TEM, JEOL-2100F) and scanning electron microscopy (SEM, Hitachi SU8000) equipped with an energy dispersive X-ray spectrometer (EDS, BRUKER). The ligament or pore sizes of nanoporous Ag in SEM images were measured by an image processing software ImageJ.

Results and discussion

Microstructure characterization and phase composition

The XRD results of the $\text{Ag}_{25}\text{Cu}_{75}$ alloy dealloyed in a 10 wt% HNO_3 solution at 45°C without UI and with ultrasonic irradiation at different frequencies are presented. The XRD patterns of $\text{Ag}_{25}\text{Cu}_{75}$ initial alloy sample was shown in Fig. 1 and the photograph of $\text{Ag}_{25}\text{Cu}_{75}$ initial alloy was inset in Fig. 2a. It can

be seen that two phases were present in the initial alloy: α - $\text{Cu}(\text{Ag})$ solid solution and an Ag_2Cu intermetallic compound. This result was close to those presented previously.⁷ As shown in Fig. 1, only a single Ag phase was left in the samples after dealloying for 20 min without UI and with UI at the frequencies of 28 and 40 kHz respectively. It is obvious that Cu from the $\text{Ag}_{25}\text{Cu}_{75}$ initial alloy was etched away during the dealloying process and all samples were dealloyed thoroughly after 20 min. In fact, the dealloying speed of $\text{Ag}_{25}\text{Cu}_{75}$ alloy was faster when dealloyed with UI than free corrosion. Technically, there was always a small amount of residual Cu existing in the final NPS. So it was not taken into account since the content of residual Cu was too little.¹⁹

The evolution of alloy microstructure during the dealloying process without UI was characterized. Fig. 2 shows the SEM images of the NPS ribbons (plan view) dealloyed from the $\text{Ag}_{25}\text{Cu}_{75}$ alloys for different durations in a 10 wt% HNO_3

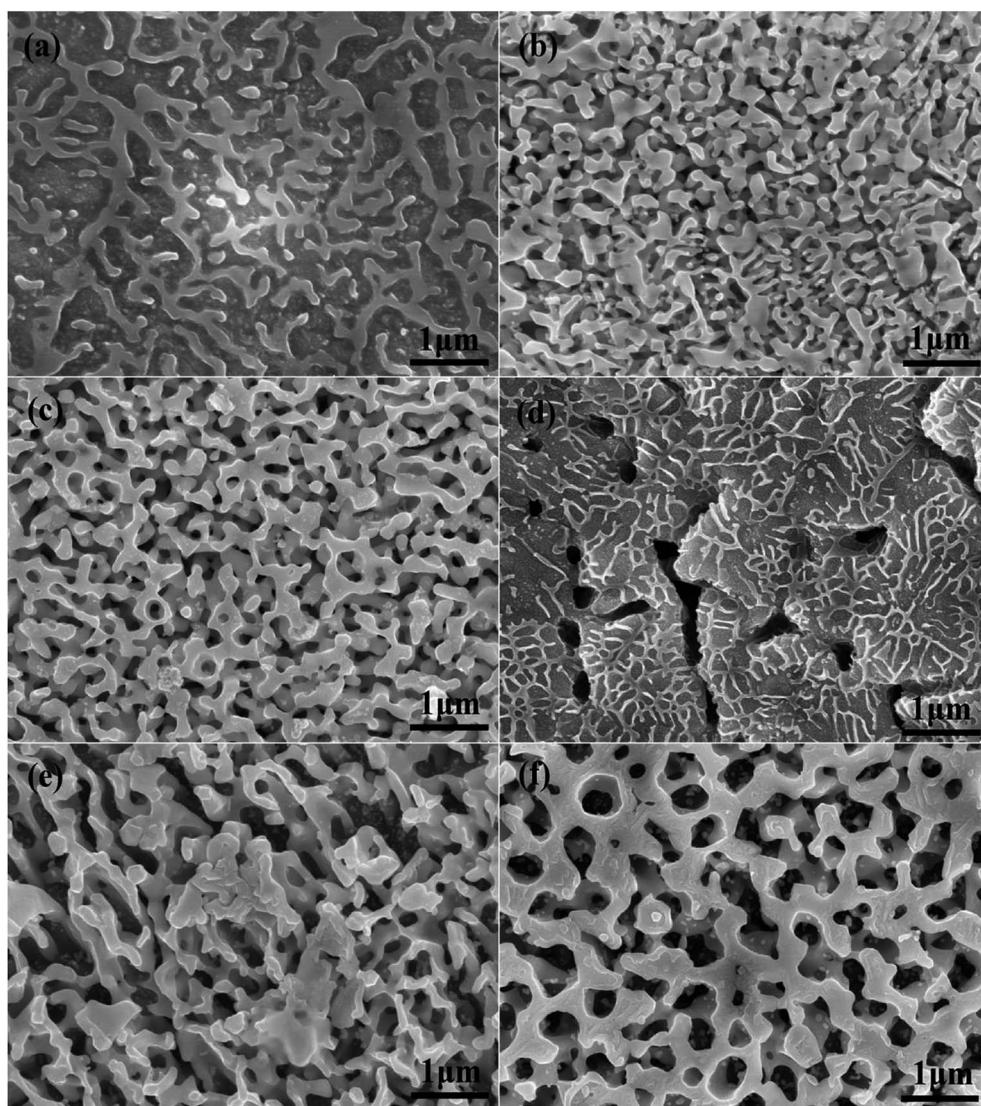


Fig. 3 (a–c) SEM images of the $\text{Ag}_{25}\text{Cu}_{75}$ samples (plan view) dealloyed with UI (frequency: 28 kHz) for different periods of time: (a) 5 min, (b) 10 min, and (c) 20 min. (d–f) SEM images of the $\text{Ag}_{25}\text{Cu}_{75}$ samples (plan view) dealloyed with UI (frequency: 40 kHz) for different periods of time: (a) 5 min, (b) 10 min, and (c) 20 min.



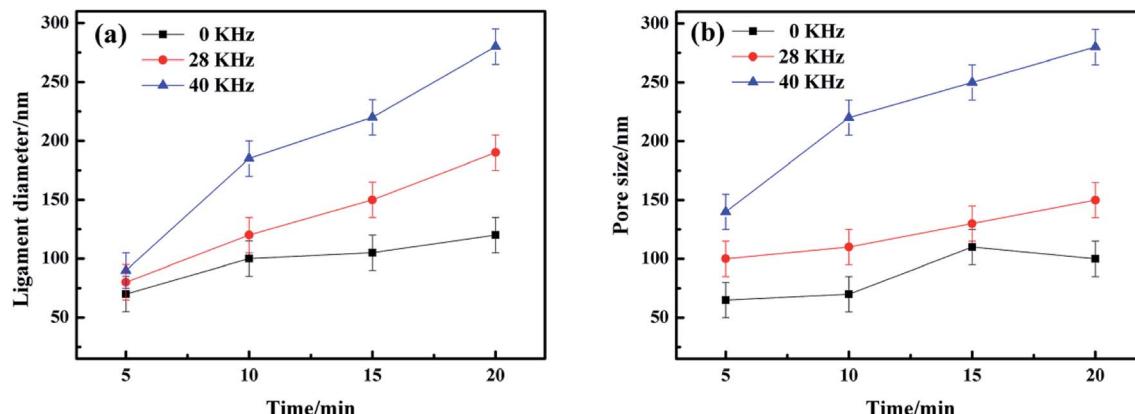


Fig. 4 (a) Ligament diameters and (b) pore sizes of $\text{Ag}_{25}\text{Cu}_{75}$ at different dealloying time under UI at frequencies of 0, 28 and 40 kHz.

solution at 45 °C. The morphology of the resulting NPS ribbon exhibited a homogeneous, ligament-channel structure. After dealloying for 20 min, the diameter of the ligament increased, and the size changed from 70 ± 15 to 120 ± 15 nm.

The SEM images of the NPS ribbons (plan view) dealloyed under UI (frequency: 28 kHz) for different durations are shown in Fig. 3a-c. The final microstructure was much coarsened than that of the $\text{Ag}_{25}\text{Al}_{75}$ alloy ribbon which was dealloyed without UI treatment. After dealloying for 20 min, the diameter of the ligament increased from 80 ± 15 to 190 ± 15 nm. Fig. 3d-f show

the SEM images of the NPS ribbons (plan view) dealloyed with UI (frequency: 40 kHz) for different durations. Many large cracks appeared on the alloy surface (Fig. 3d) which suggested that the high intensity of ultra-sonication gives rise to the cracks even fracture during the dealloying process. The as-dealloyed sample exhibited an microstructure with coarsened ligaments and enlarged pores. Moreover, the diameter of the ligament prepared from the $\text{Ag}_{25}\text{Al}_{75}$ alloy dealloyed for 20 min reached approximately 280 nm and was much larger than those of the $\text{Ag}_{25}\text{Al}_{75}$ alloys dealloyed with UI at the frequency of 0 or 28 kHz.

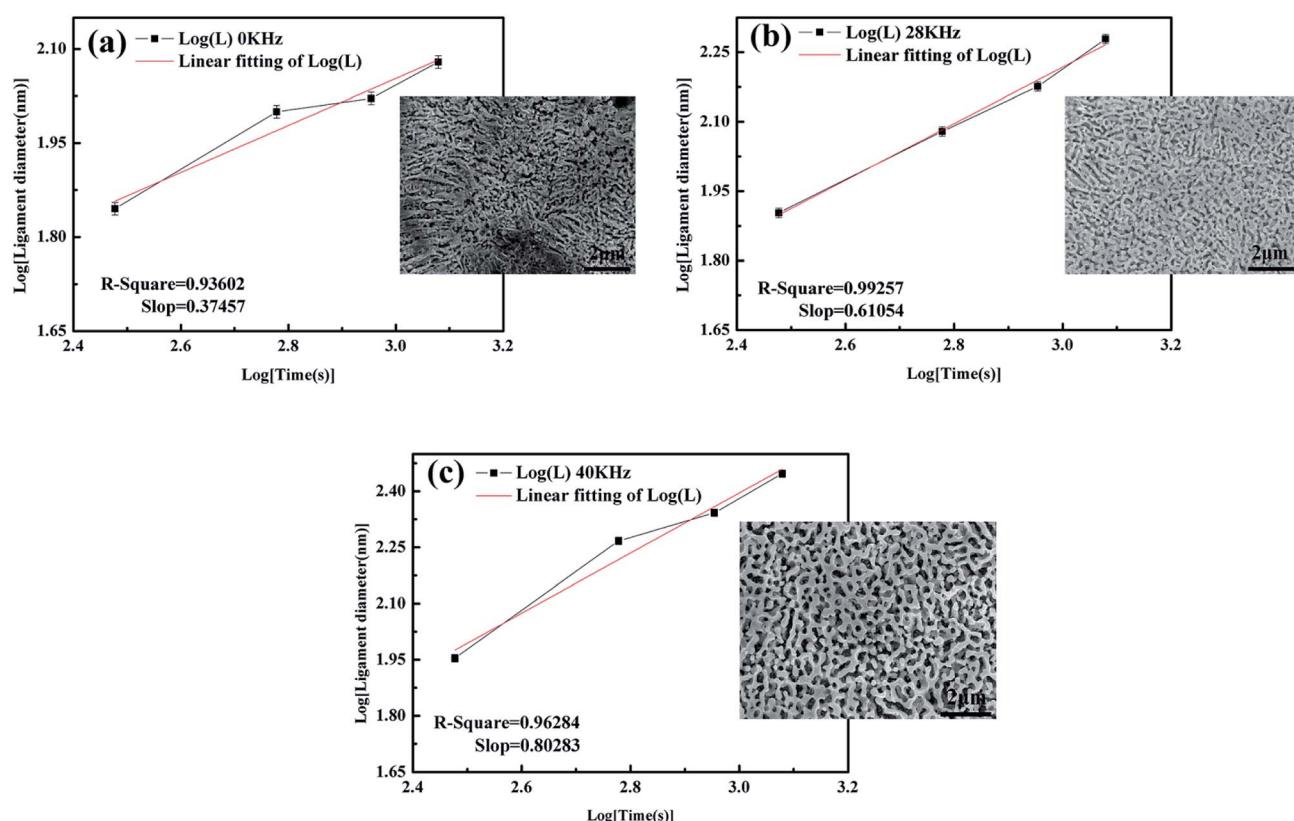


Fig. 5 The time dependence of the coarsening of ligament size at the frequencies of (a) 0, (b) 28 and (c) 40 kHz respectively; Microstructure of NPS dealloyed under UI at the frequencies of 0, 28 and 40 kHz inset in (a), (b) and (c), respectively.

Effect of ultrasonic irradiation on microstructure of NPS

The alterations of average ligament diameters and pore sizes with UI at different frequencies were measured and the data was shown in Fig. 4. It can be easily obtained that both ligament diameters and pore sizes increased with increasing dealloying time. And the final results of ligament and pore size were much larger than many of those from other researches.^{10,23,29} The ligament diameters increased sharply for dealloying with UI, compared with those from free corrosion. Therefore, we concluded that a high-intensity ultrasonic irradiation would affect the microstructure of NPS remarkably.

Additionally, the coarsening of the ligament was caused by the diffusion of the more noble elements atoms which were driven by surface tension existing at the interface between the alloy and solution.^{30,31} With the reduction in the surface energy, the ligament coarsened gradually.³² It is well known that the coarsening mechanism of the ligament is controlled by the surface diffusivity (D_s) of the more noble elements:¹⁶

$$D_s = \frac{d(t)^4 kT}{32\gamma t a^4}, \quad (1)$$

where k is Boltzmann constant, T is the etching temperature, γ is surface energy, t is the etching time, and a is the lattice parameter. Moreover, the relationship of D_s (surface diffusion of gold atoms) and d (ligament diameter of nanoporous gold) when preparing nanoporous gold dealloyed from Au–Ag alloy could be estimated by the following proposed equation:³³

$$d \propto (D_s/V_0)^\mu, \quad (2)$$

where V_0 is the velocity of a flat alloy surface with no gold accumulated on it and μ is a constant equal to 1/6 or 1/4 depending on whether the theoretical or the empirical model is used.^{16,34} Hence, the coarsening rate of the ligament in this study may largely reflect the diffusion rate of Ag atoms. The time dependence of the ligament coarsening with UI at different frequencies can be observed in Fig. 5a–c. According to these plots, the relationship between time and ligament could be described as:

$$d_{0 \text{ kHz}} \propto t^{0.37} \approx t^{2/5}, \quad (3)$$

$$d_{28 \text{ kHz}} \propto t^{0.61} \approx t^{3/5}, \quad (4)$$

$$d_{40 \text{ kHz}} \propto t^{0.80} \approx t^{4/5}, \quad (5)$$

where d is the ligament size of NPS and t is the dealloying duration. The exponents (3/5 and 4/5) were much larger than the reported surface diffusion exponents of 2/5 for NPS and 1/8 for nanoporous Au.^{35,36} It is now demonstrated that the coarsening rate of the ligament would be increased through dealloying with UI. The morphology of NPS dealloyed with UI was observed to be much more coarsened than that of the NPS dealloyed without the ultrasonic treatment through the analysis of the SEM results inset in Fig. 5a, b and c, respectively. This is well consistent with the result that the ligament diameter and

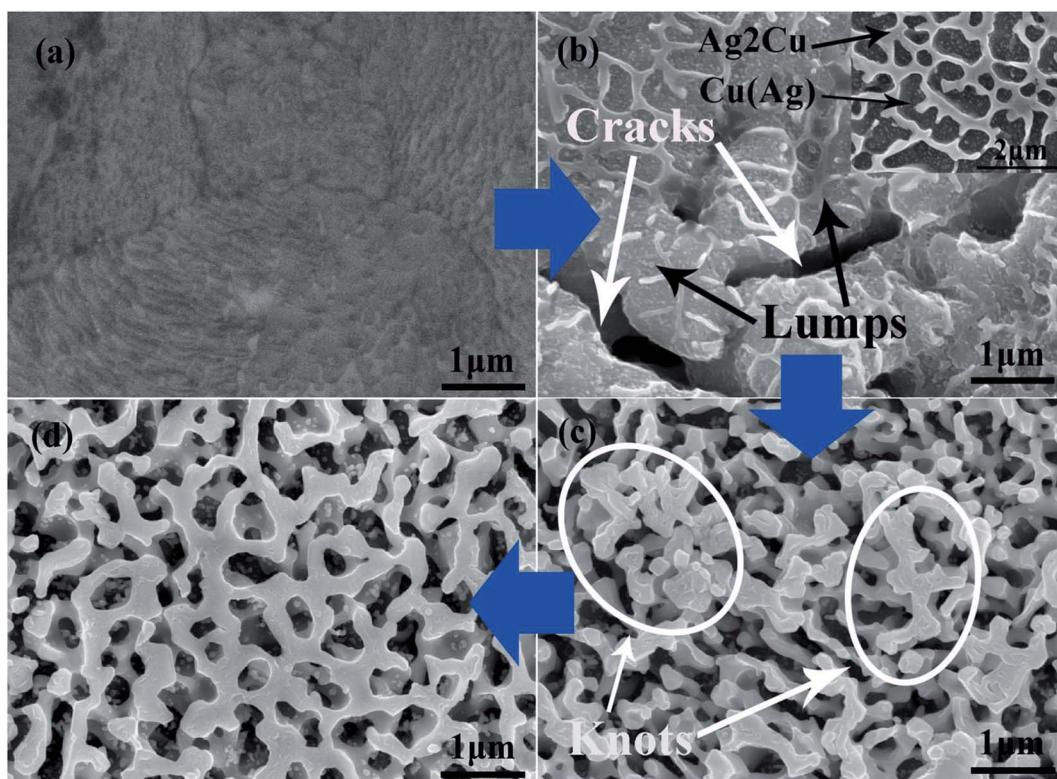


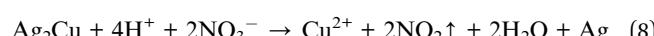
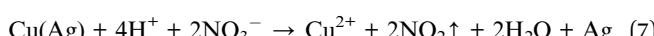
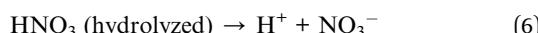
Fig. 6 Microstructures of $\text{Ag}_{25}\text{Cu}_{75}$ ribbons dealloyed under UI at 45 °C in a 10 wt% HNO_3 aqueous solution for different durations: (a) 0 min, (b) 5 min, (c) 10 min, and (d) 20 min; the inset in (b) is the image of microstructure at the initial stage.



pore size of NPS dealloyed at 40 kHz, as shown in Fig. 4, were larger than those obtained at other frequencies.

Formation mechanism of nanoporous Ag during dealloying with ultrasonic irradiation

The formation mechanism of NPS may be somewhat distinctive with the application of UI. After the $\text{Ag}_{25}\text{Cu}_{75}$ alloy was immersed in the dealloying solution with ultrasonic treatment, Cu atoms on the surface will be leached away easily. The reaction of Cu in HNO_3 aqueous solution could be described as follows:



To fill the vacancies caused by the etching of the copper atoms, the Ag atoms on the interface between the alloy and dealloying solution moved and agglomerated into clusters by

diffusion due to the effect of surface tension.³² With the assembly of the Ag atoms, ligaments were formed and coarsened gradually. Compared with free corrosion dealloying, the formation mechanism of NPS by dealloying $\text{Ag}_{25}\text{Cu}_{75}$ with UI developed in a new lump-forming-disintegrating manner, as described below.

Fig. 6a is the SEM image of $\text{Ag}_{25}\text{Cu}_{75}$ precursor alloy. When dealloying began, the Cu atoms exposed to the HNO_3 solution would react actively with H^+ ions and then are transformed into free Cu^{2+} ions in the solution. At the moment, some walls and small particles can form on the alloy surface. According to the previous related reports, the wall structure was the Ag_2Cu phase and the particle structure was Cu(Ag) phase. And Cu(Ag) phase was dealloyed prior to the Ag_2Cu phase.^{29,37,38} Moreover, the inset of Fig. 6b is presented to illustrate the distributions of the Cu(Ag) solid solution and the Ag_2Cu intermetallic compound in the Ag-Cu alloy at the initial stage of dealloying. The structure of Ag_2Cu and Cu(Ag) are both marked using black arrows. Influenced by ultrasonic condition, many cracks and lumps can then be formed on the surface layer of the sample (Fig. 6b). However, many residual Ag_2Cu walls were still on the surface of

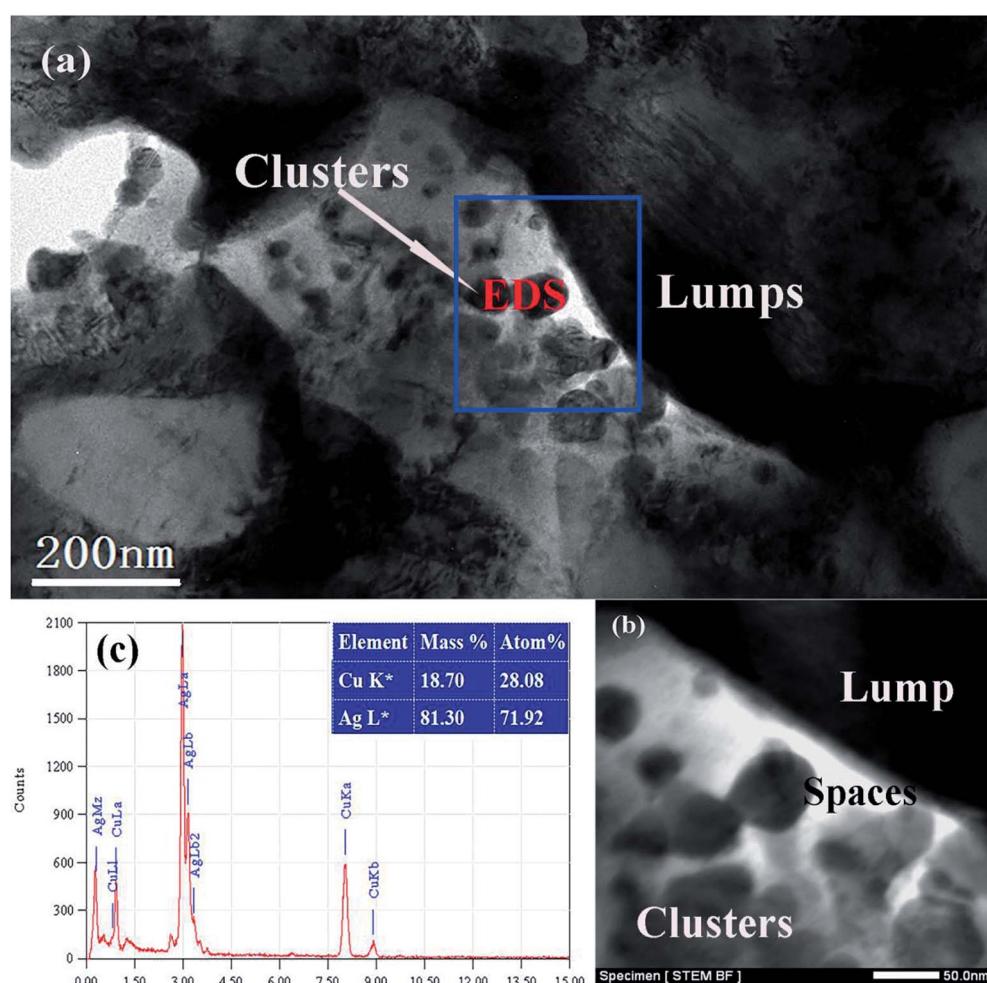


Fig. 7 (a) TEM images of NPS produced by dealloying $\text{Ag}_{25}\text{Cu}_{75}$ within UI at the frequency of 40 KHz in 10wt% HNO_3 aqueous solution for 5 min. (b) Is the EDS pattern marked in (a). (c) Is the part enlarged bright-field scan transmission electron microscope (STEM BF) image of (a) which is remarked by a blue rectangle.



the lumps as illustrated in Fig. 6b. Their sizes measuring to be 45 ± 15 nm are very small.

For further characterizing the lump structure, TEM images of the $\text{Ag}_{25}\text{Cu}_{75}$ alloy sample dealloyed with UI at the frequency of 40 kHz for 5 min are represented in Fig. 7a. Some large plates

and thin particles corresponding to lumps and clusters respectively can be observed clearly. The EDS pattern (Fig. 7b) confirmed that the particles were actually the Ag-rich clusters. There are a lot of spaces appearing between the plate and particles. This demonstrates the appearance of cracks at the

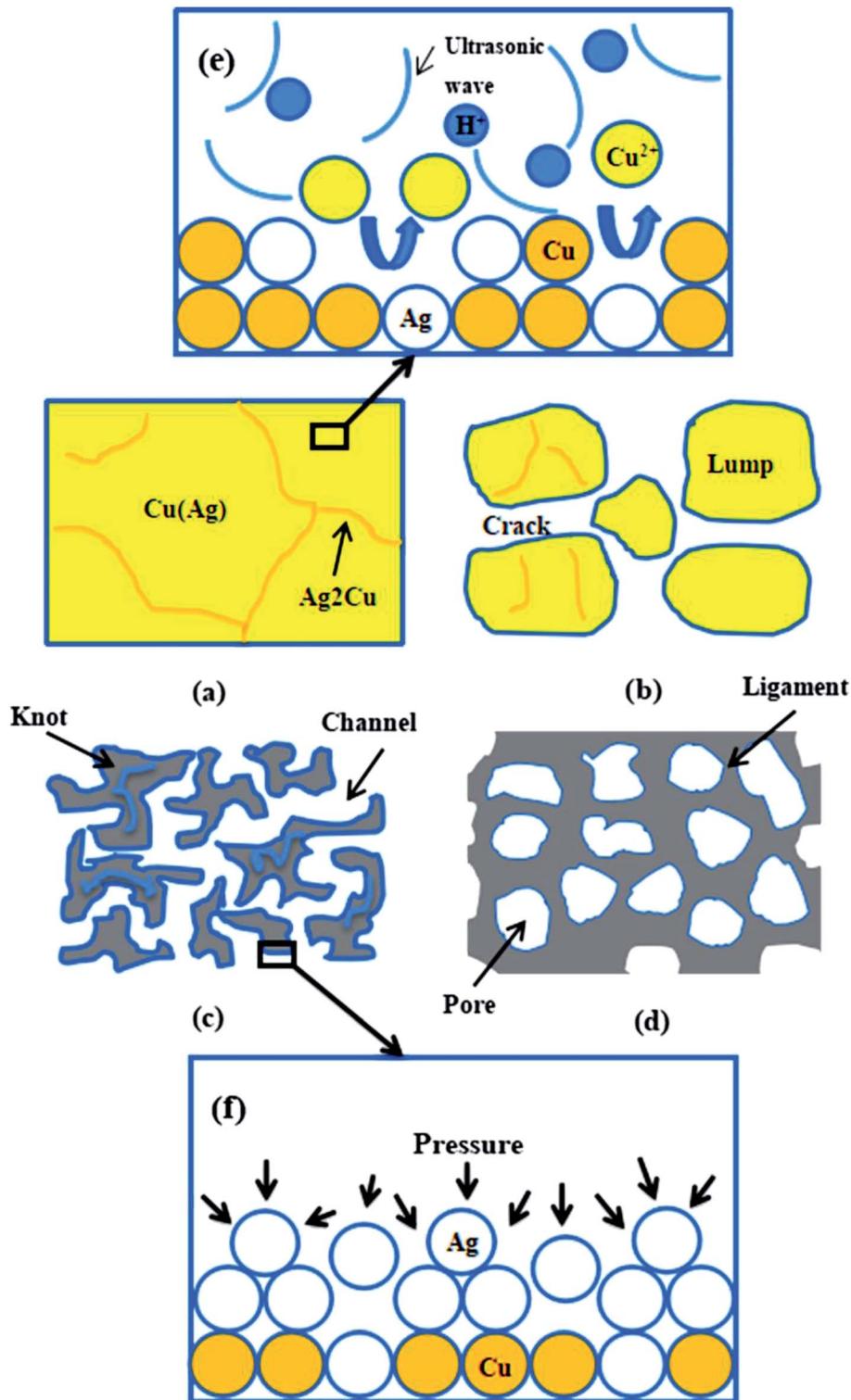


Fig. 8 (a)–(d) Schematic diagram of NPS formation process when $\text{Ag}_{25}\text{Cu}_{75}$ Alloy is dealloyed in a UI; (e) and (f) are the schematic diagrams of reaction interface at period of (a) corrosion and (c) diffusion.



beginning of dealloying. The formation of vacancies may support the hypothesis that high-intensity ultrasonic irradiation could accelerate the reaction process between Cu and HNO_3 . Fig. 7c is the bright-field scanning transmission electron microscopy (STEM BF) image. It shows an enlarged view of a part of Fig. 8a which was marked by a blue rectangle and represent the distributions of lumps, clusters and cracks clearly. Distinctly, the lumps (corresponding to plates) and cracks (corresponding to vacancies) have formed at the initial stage of dealloying and the clusters moved and aggregated on the surface of the lumps.

As the dealloying continued, more Cu atoms were leached away leading to the decrease of the Cu content in alloy. The number and the size of the pores increased correspondingly. Then the lumps shrank and disintegrated gradually. Due to the further removal of Cu atoms, more Ag atoms exposed to solution moved and assembled to form the ligament. As a result, many large knots composed of ligaments appeared which was marked by the white circles in Fig. 6c. It is obvious that they evolved from the large lumps to the thinner knots. Subsequently, the coarsening of the ligaments proceeded and the knots vanished gradually until an ultra-homogeneous nanoporous structure ultimately formed (Fig. 6d). The whole dealloying process of $\text{Ag}_{25}\text{Cu}_{75}$ alloy in an ultrasonic environment can be illustrated in details by a schematic diagram (Fig. 8a-d).

As previously reported, the application of ultrasonic fields into the solution could give rise to phenomena such as cavitation and acoustic streaming which could lead to the effects of ultrasonic stirring and degassing.³⁹ We assumed the ultrasonic irradiation could generate high-intensity shock waves to effectively accelerate the reaction between Cu and the HNO_3 solution. The stirring and shocking effect could generate waves in the medium solution as shown in Fig. 8e. That would greatly increase the chance of collision and contact among the atoms which were involved in the reaction system, such as the reaction between H^+ ions hydrolyzed in the solution and Cu atoms crystalized in the alloy. Hence Cu atoms were more inclined to be separated from the alloy and turned into Cu^{2+} compared with those dealloyed through free corrosion. As a result, cracks and lumps appeared at the initial stage of dealloying process with ultrasonic irradiation.

In addition, the ultrasonic vibration could facilitate the coarsening of Ag ligament. The acoustic streaming effects originating from ultrasonic vibration may affect the diffusion of Ag atoms. As reported, the streaming was eventually related to the pressure present in the dealloying solution and the pressure acting on diffusing atoms could be described by:⁴⁰

$$P_a = \nu pc = \rho c A_0 \omega \cos\left(\omega t - \omega \frac{x}{c}\right). \quad (9)$$

Obviously, the maximum value of the pressure could approximately be simplified into:

$$P_{\max} = \nu pc = \rho c A_0 \omega, \quad (10)$$

where A_0 is the amplitude (μm), ω is the frequency (Hz) of the ultrasonic field, x/c is the phase factor, ρ is the density of the dealloying solution (g cm^{-3}) and c is the speed of sound in the HNO_3 aqueous solution (m s^{-1}). In this system, $\rho \approx 1.06 \text{ g cm}^{-3}$, $c \approx 1.51 \times 10^3 \text{ m s}^{-1}$, $A_{01} = 3.3 \mu\text{m}$ for $\omega_1 = 2.8 \times 10^4 \text{ Hz}$, and $A_{02} = 4.2 \mu\text{m}$ for $\omega_2 = 4.0 \times 10^4 \text{ Hz}$. Therefore, $P_{\max 1}$ and $P_{\max 2}$ are estimated to be 0.15 MPa and 0.27 MPa, approximately equal to 1.5 and 3 atmospheres, respectively.

When the alloy ribbon was immersed in the dealloying solution with ultrasonic vibration, the whole surface of alloy ribbon was under the pressure from all directions. The schematic diagram of reaction interface when $\text{Ag}_{25}\text{Cu}_{75}$ Alloy is dealloyed with UI during the period of diffusion is presented in Fig. 8f. Since Ag atoms at the surface of alloy are all under the pressure from multi-direction, the Ag atoms (or particles) become more aggressive during diffusion process. The movement rate of Ag atoms was accelerated and the diffusion was forced intensely. As a result, an extremely homogeneous microstructure of the nanoporous Ag with a coarsened ligament-pore distribution was obtained. Moreover, the microstructure of the ultrasonically treated sample was much more homogeneous than that of the free corrosion samples.

Conclusions

Effects of ultrasonic irradiation (UI) on the diffusion and microstructure of NPS were analyzed. The following conclusions were drawn: the formation mechanism of NPS dealloying with UI was different from that for NPS obtained by free corrosion. The NPS evolved depending on a lump forming-disintegrating mechanism. Ultrasound has an extremely strong effect on the microstructure of NPS. On one hand, the stirring produced by ultrasonic vibration could promote the reaction between Cu and HNO_3 aqueous solution. On the other hand, ultrasonic vibration could facilitate the diffusion of Ag atoms by the induced solution pressure when the coarsening rate of ligament was increased significantly during dealloying with UI. Dealloying treated with ultrasonic irradiation could generate an extremely homogeneous microstructure with much coarser ligaments compared with those dealloyed by free corrosion.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors gratefully acknowledge the sponsorship from National Natural Science Foundation of China, under the project no. 51574147, Liaoning Provincial Natural Science Foundation, under the project no. 201602474.

References

1. A. Wittstock, V. Zielasek and J. Biener, *Science*, 2010, **327**, 319–322.



2 N. Yu, L. Jiang and H. Hou, *JOM*, 2017, **69**(6), 1027–1033.

3 H. J. Jin, X. L. Wang and S. Parida, *Nano Lett.*, 2010, **10**, 187–194.

4 Y. C. Liu, J. C. Huang and X. Wang, *RSC Adv.*, 2018, **8**, 1307513082.

5 N. Yu, T. Wang, C. Nie, L. Sun, J. Li and H. Geng, *JOM*, 2016, **68**(1), 391–396.

6 G. J. Li, X. P. Song and F. F. Lu, *J. Mater. Res.*, 2012, **27**(12), 1612–1620.

7 Y. Zhao, X. Wang and J. C. Huang, *Mater. Lett.*, 2016, **183**, 165–169.

8 T. A. Rebbeckchi and Y. Chen, *J. Mater. Res.*, 2017, **33**, 1–14.

9 R. W. Zhang, X. Wang and Z. C. Zhang, *RSC Adv.*, 2018, **8**, 9462–9470.

10 W. Liu, L. Chen and J. Yan, *Corros. Sci.*, 2015, **94**, 114–121.

11 R. Morrish, K. Dorame and A. J. Muscat, *Scr. Mater.*, 2011, **64**, 856–859.

12 T. Krekeler, V. S. Anastasia and M. Graf, *Mater. Res. Lett.*, 2017, **5**, 1–8.

13 H. B. Lu, Y. Li and F. H. Wang, *Scr. Mater.*, 2007, **56**(2), 165–168.

14 Z. Dan, F. Qin and S. I. Yamaura, *J. Alloys Compd.*, 2013, **581**, 567–572.

15 M. Hakamada, H. Nakano and T. Furukawa, *Phys. Chem. Chem. Phys.*, 2011, **13**(26), 12277–12284.

16 L. H. Qian and M. W. Chen, *Appl. Phys. Lett.*, 2007, **91**, 597–605.

17 Z. Zhang, Y. Wang and Z. Qi, *J. Phys. Chem. C*, 2009, **113**, 12629–12636.

18 H. Rösner, S. Parida and D. Kramer, *Adv. Eng. Mater.*, 2010, **9**, 535–605.

19 X. Wang, Z. Qi and C. Zhao, *J. Phys. Chem. C*, 2009, **113**, 13139–13150.

20 L. Y. Chen, J. S. Yu and T. Fujita, *Adv. Funct. Mater.*, 2010, **19**, 1221–1226.

21 W. C. Li and T. J. Balk, *Scr. Mater.*, 2010, **62**, 167–169.

22 Y. Y. Zhao, D. Estévez and C. Chang, *Mater. Lett.*, 2015, **144**, 138–141.

23 T. Song, Y. Gao, Z. Zhang and Q. Zhai, *CrystEngComm*, 2012, **14**, 3694–3701.

24 Z. Lu, C. Li and J. Han, *Nat. Commun.*, 2018, **9**(1), 276–283.

25 H. B. Xu, T. T. Meek and Q. Y. Han, *Mater. Lett.*, 2007, **61**, 1246–1250.

26 Z. W. Xu, J. C. Yan and W. Chen, *Mater. Lett.*, 2008, **62**, 2615–2618.

27 Z. W. Liu, Q. Y. Han and J. G. Li, *Composites Part B: Engineering*, 2011, **42**, 2080–2084.

28 Z. W. Liu, M. Rakita and Q. Han, *Metall. Mater. Trans. A*, 2012, **43**, 2116–2124.

29 F. Chen, X. Chen and L. J. Zou, *Mater. Sci. Eng., A*, 2016, **660**, 241–250.

30 H. J. Qiu, L. Peng and X. Li, *Corros. Sci.*, 2015, **92**, 16–31.

31 N. Hirai, H. Tanaka and S. Hara, *Appl. Surf. Sci.*, 1998, **130**, 506–511.

32 M. S. Kim and H. Nishikawa, *J. Mater. Sci.*, 2013, **48**, 5645–5652.

33 Y. Sun and T. J. Balk, *Metall. Mater. Trans. A*, 2008, **39**, 2656–2665.

34 J. Erlebacher, M. J. Aziz and A. Karma, *Nature*, 2001, **410**, 450–453.

35 J. Erlebacher, *J. Electrochem. Soc.*, 2004, **151**, C614–C626.

36 Z. Zhang, Y. Wang and Y. Wang, *Scr. Mater.*, 2010, **62**, 137–140.

37 E. Detsi, Z. Vuković and S. Punzlin, *CrystEngComm*, 2012, **14**, 5402–5406.

38 T. T. Song, Y. L. Gao and Z. H. Zhang, *Corros. Sci.*, 2013, **68**, 256–262.

39 Z. W. Liu, X. M. Wang and Q. Y. Han, *Metall. Mater. Trans. A*, 2014, **45**(2), 543–546.

40 Z. Liu, Q. Han and J. Li, *Powder Technol.*, 2013, **247**, 55–59.

