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Heterogeneous activation of peroxyomonosulfate for bisphenol AF degradation with $\text{BiO}_{0.5}\text{Cl}_{0.5}$ [†]

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This study represents the first investigation on the application of peroxyomonosulfate (PMS) for the degradation of bisphenol AF (BPAF) using halogen bismuth oxide composites ($\text{BiO}_{0.5}\text{Cl}_{0.5}$). The hierarchical $\text{BiO}_{0.5}\text{Cl}_{0.5}$ was successfully synthesized and systematically characterized with multifarious techniques including XRD, SEM, FTIR and XPS to investigate the morphology and physicochemical properties of the samples. Several parameters affecting the degradation efficiency including catalyst dosage, PMS loading, and pH value were elucidated. Inorganic ions such as HCO_3^- showed significant inhibition in the $\text{BiO}_{0.5}\text{Cl}_{0.5}$ /PMS process due to the quenching effect. The effect of various water matrices including tap water and surface water on the removal of BPAF was studied to indicate that the present reaction system shows great potential for cleaning BPAF waste water. Furthermore, the production of sulfate radicals and hydroxyl radicals was validated through radical quenching and ESR tests, thus a possible oxidation mechanism was proposed. Overall, these results reveal that the activation of PMS by the $\text{BiO}_{0.5}\text{Cl}_{0.5}$ /PMS system is an efficient and promising advanced oxidation technology for the treatment of BPAF-contaminated waters and wastewaters.

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1. Introduction

In recent years, bisphenols (BPA, BPAF and BPS) have been reported to be endocrine disruptor chemicals (EDCs), causing a variety of human health issues,^{1,2} which can imitate the estrogen of natural hormones to bring about various adverse effects on humans even at low exposure levels.^{3,4} BPA is a commonly used chemical with production estimated at above 4.5 million tons per year.⁵ Various methods have been proposed to remove BPA from wastewater nowadays, including adsorption,⁶ microbiological biodegradation,⁷ photocatalytic degradation^{8–10} and chemical oxidation.^{11,12} However, research on the degradation of its analogues is rare.

BPAF, an analogue of BPA, is being considered as a replacement for BPA. BPAF has been found in environmental water and solid samples, as well as in indoor dust, beverages, and food worldwide.^{13–15} BPAF has a similar chemical structure to BPA and is mainly used for professional elastomers synthesis cross linking and curing agent to improve their chemical and thermal properties. However, the introduction of fluorine improved the degradation resistance.¹⁶ Many studies showed that BPAF negatively affect some physiological processes in humans and animals,¹⁷ but the research on the degradation of it is rare, thus

in the present study, BPAF was chosen as the target for degradation and removal.

Recently, advanced oxidation processes (AOPs) based on persulfate (PS) or peroxyomonosulfate (PMS) have been attracted great attention. It has been confirmed that sulfate radicals ($\cdot\text{SO}_4^{4-}$) provide an effective way for removing reluctant organic compounds in the water,^{18,19} due to the following reasons: (1) $\cdot\text{SO}_4^{4-}$ possesses a longer half-life than OH^\cdot ; (2) $\cdot\text{SO}_4^{4-}$ radicals ($E_0 = 2.5\text{--}3.1\text{ V}$) exhibit higher oxidation potentials than OH^\cdot radicals ($E_0 = 1.8\text{--}2.7\text{ V}$) and (3) $\cdot\text{SO}_4^{4-}$ radicals are more selective towards destruction of aromatic molecules.^{20,21} In recent years, lots of researchers have focused their attention on the oxone (potassium peroxyomonosulfate, PMS), which is an eco-friendly and promising source to generate $\cdot\text{SO}_4^{4-}$ radicals.^{22–27}

On the other hand, bismuth oxyhalides (BiOX , X = Cl, Br and I) have been admitted as excellent adsorbents and photocatalysts for the removal of organic pollutants in the water.^{28–31} The band gaps of BiOX can be within 1.7–3.4 eV by varying the halides and the composition ratios of the two halides ($\text{BiOX}_\alpha\text{Y}_{1-\alpha}$).³² Owing to the high separation rate of photo-induced electron–hole pairs, the past ten years have witnessed tremendous efforts in utilizing BiOCl as a promising catalyst for photocatalytic applications under visible light irradiation, including pollutants removal, N_2 fixation, CO_2 reduction and water splitting.³³ However, most of the photocatalytic oxidation reaction based on BiOX usually needs several hours or more for pollutant removal. Therefore, it is necessary to promote technology combined with another antioxidant. To our best

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knowledge, bismuth oxyhalide has not been tested as an activator of peroxide for water purification. There is a handful of research has been reported that the interaction between the PMS and BiOX for organic wastewater treatment.³⁴

Hence, in the present study, we represent the first investigation on the potential application of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ combined with PMS. We synthesized hierarchical $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ using the solvothermal method. BPAF, which is commonly used in the chemical and textile industries were selected as typical targets. Multiple characterizations including XRD, SEM, FTIR, XPS and ESR were applied to evaluate the physicochemical performance of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ and reaction mechanism. In addition, several key parameters including catalyst dosage, PMS loading, the initial pH of BPAF solution and common anions were investigated. Besides, the realistic conditions were also investigated for the system using tap water and surface water conditions. Finally, a possible reaction mechanism was proposed based on the scavenger experiments.

2. Materials and methods

2.1 Materials and reagents

$\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$, KI and KCl were purchased from Kelong Chemical Reagent Co. (Chengdu, China). Bisphenol AF (BPAF) was purchased from Sigma-Aldrich Co., Ltd (Shanghai, China). Peroxymonosulfate (PMS) was obtained from Aladdin China. All other chemical reagents were analytical grade and used without further purification. The typical solutions used were prepared using Milli-Q water system ($18.2 \text{ M}\Omega \text{ cm}^{-1}$).

2.2 Synthesis of composites bismuth oxyhalides

The three catalysts were synthesized as the modified procedure reported by previous studies.²⁹ 8.731 g $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ was dissolved in 180 mL ethylene glycol under intense stirring for 30 min to obtain solution A. Then, 2.989 g KI (for BiOI), 1.344 g KCl (for BiOCl) or 1.494 g KI/0.672 g KCl (for $\text{BiOI}_{0.5}\text{Cl}_{0.5}$) was dissolved in 180 mL water to obtain solution B, respectively. Afterwards, the solution B was added rapidly into the solution A under intense stirring for 2 h at the ambient temperature. The obtained mixture was calcined at 180 °C for 24 h in a stainless-steel autoclave equipped with a Teflon lining. Finally, the mixture was separated and washed three times respectively with ethanol and water, then dried at 60 °C in vacuum drying oven until use.

2.3 Experimental procedure

All experiments were carried out in a borosilicate glass beaker of 500 mL capacity at ambient temperature ($25^\circ\text{C} \pm 1^\circ\text{C}$). A 200 mL reaction solution was prepared of BPAF solution at concentration of 10 mg L^{-1} without pH adjustment or with adjusting initial pH using 0.1 mol L^{-1} HCl and NaOH solution. In each test, a set of prepared $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ was dispersed in the solution, all the suspensions were magnetically stirred for 10 min in the dark before adding a certain amount of PMS, in order to attain the adsorption-desorption equilibrium between the $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ and the BPAF molecules. At given time intervals,

aliquots of the suspension (about 1.5 mL) were sampled, filtered by a 0.22 μm filter before the concentration detection, quenched immediately by adding 60 μL 0.5 mol L^{-1} $\text{Na}_2\text{S}_2\text{O}_3$ solution into the 2 mL auto sampler vials, its volume for PMS quench was 100 : 1. Before the time of zero, a pre-adsorption process for BPAF removal was accomplished.

The concentration of BPAF ($\lambda = 270 \text{ nm}$) was detected by a Water e2695 HPLC equipped with UV-visible detection (Waters 2489) at the column temperature of 35 °C. And the mobile phase was maintained at a flow rate of 1.0 mL min^{-1} with a constant ratio (methanol/0.1% ammonium acetate = 70/30). The total organic carbon (TOC) of samples were measured by an Elementar liquid TOC II analyzer.

The degradation BPAF process approximately followed pseudo-first-order kinetics by data fitting, thus overall rate law for the BPAF degradation can be expressed as eqn (1):

$$-\ln\left(\frac{C_t}{C_0}\right) = K_{\text{app}} t \quad (1)$$

where the C_0 is initial concentration of BPAF, C_t is the concentration of BPAF at time t , and K_{app} (min^{-1}) is the observed reaction rate constant.

2.4 Characterization and analytical procedures

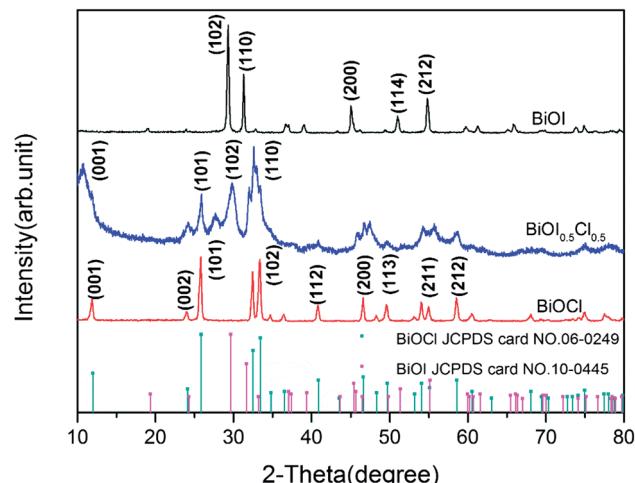
To determine the morphology of the as-prepared samples, scanning electron microscopy (SEM) measurements were run using a JSM-7500F (JEOL, Japan) at 20 kV. The crystal structural characterization of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ was conducted by powder X-ray diffraction spectrometry (XRD) using an Empyrean diffractometer advance instrument between 10 and 80° (2θ) (PANalytical B.V., Holland). Fourier Transform Infrared Spectrometer (FTIR) results were conducted using a Nicolet 6700 FTIR spectrometer (Thermo Scientific, United States), at the range of 400 to 4000 cm^{-1} wave numbers. The elemental composites and chemical states of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ were further elucidated with X-ray photoelectron spectroscopy (XPS) using a XSAM800 (Kratos, UK).

3. Results and discussion

3.1 Characterization of the synthesized samples

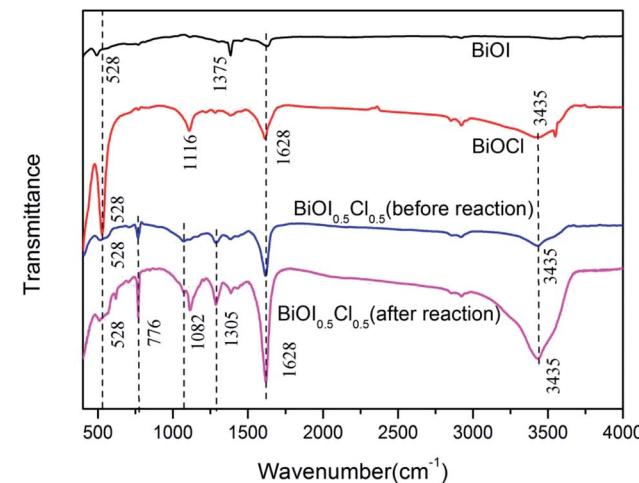
XRD was applied to characterize the chemical composition and crystalline structure of the synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$, the patterns of the samples are presented in Fig. 1. It is clearly that BiOX can be indexed to tetragonal phase structure BiOI (JCPDS card no. 73-2062, the cell parameters of $a = 3.984 \text{ \AA}$) and BiOCl (JCPDS card no. 73-2060, the cell parameters of $a = 3.883 \text{ \AA}$), indicating a high purity of the products. Moreover, the XRD peaks of the $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ sample is weaker and broader than the pure BiOI and BiOCl, demonstrating a smaller particle size and lower crystallinity to some extent.³⁵ The major peaks are positioned at $2\theta = 11.1^\circ, 25.8^\circ, 29.9^\circ, 32.0^\circ$ and 32.6° . The shift of (001), (101), (102) and (110) planes are the strong evidence to determine the intercalated layer products, indicating that the as-prepared $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ composite is not merely a mixture of BiOI and BiOCl.³⁶



Fig. 1 XRD patterns of BiOI, BiOCl and $\text{BiOI}_{0.5}\text{Cl}_{0.5}$.

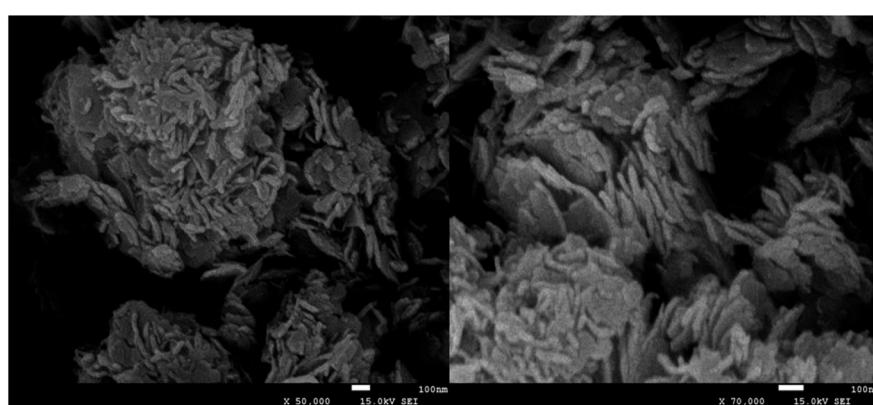
The low- and high-magnification SEM images of synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ are presented in Fig. 2. Obviously, the prepared $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ samples are dominated with the hierarchical nanoplates.³⁶ The nanoplate like structure of the catalysts is mainly seems to consists of quite uniform plates with width of about 100 nm and thickness of about 20 nm, these observations are in accordance with the reported literature.³⁷ The formation of 2-dimension nanoplates structure can be ascribed to the internal structure of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$, where $[\text{Bi}_2\text{O}_2]^{2+}$ layers are interleaved by two slabs of X^- atoms, leading to the anisotropic growth at certain axis.³⁸ In addition, the element mapping clearly illustrates the homogeneous distribution of O, Bi, I and Cl species in the prepared $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ samples composite (as shown in Fig. S1†).

FT-IR spectra was applied to investigate and compare the functional groups on BiOI, BiOCl and $\text{BiOI}_{0.5}\text{Cl}_{0.5}$. As shown in Fig. 3, characteristic absorption peaks at 528 cm^{-1} , 1375 cm^{-1} and 1650 cm^{-1} are observed in BiOI. The peak at 528 cm^{-1} is associated with the Bi-O stretching mode, while the absorption peaks at 1375 cm^{-1} could be assigned as the asymmetry and symmetric stretching vibration peaks for Bi-I band.³⁹ Similar

Fig. 3 FTIR spectra of the synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$, BiOI and BiOCl.

characteristic peaks emerge at 1116 cm^{-1} for Bi-Cl band in BiOCl structure. The band at 3435 cm^{-1} and 1628 cm^{-1} for $\text{BiOI}_{0.5}\text{Cl}_{0.5}$, BiOCl and BiOI are the stretching of O-H and flexural vibrations of O-H in the free water.⁴⁰ No characteristic adsorption peaks of the ionic liquid are found in the FT-IR spectra, indicating that the ionic liquid can be easily removed from the surface of the synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ by washing with alcohol and water. Besides, significant changes are observed between the fresh and used $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ at 776 cm^{-1} , 1082 cm^{-1} and 1305 cm^{-1} illustrated in Fig. 3, which might be formed by the adsorption of BPAF degradation products and other radicals. The peak at 1082 cm^{-1} corresponded to C-O-C stretching vibration,⁴¹ and the respective shift in 776 cm^{-1} and 1305 cm^{-1} peaks confirmed the interaction among $\text{BiOI}_{0.5}\text{Cl}_{0.5}$, PMS and BPAF, indicating that the degradation process takes place on the surface of the $\text{BiOI}_{0.5}\text{Cl}_{0.5}$.

XPS spectra is considered as an effective analysis to obtain further insight into the elemental composition and chemical states of the prepared samples. As shown in the Fig. 4(a), the coexistence of Bi 4f, I 3d, Cl 2p and O 1s peaks indicated the catalysts consisted of Bi, I, Cl and O elements. The carbon peak

Fig. 2 SEM images of synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$.

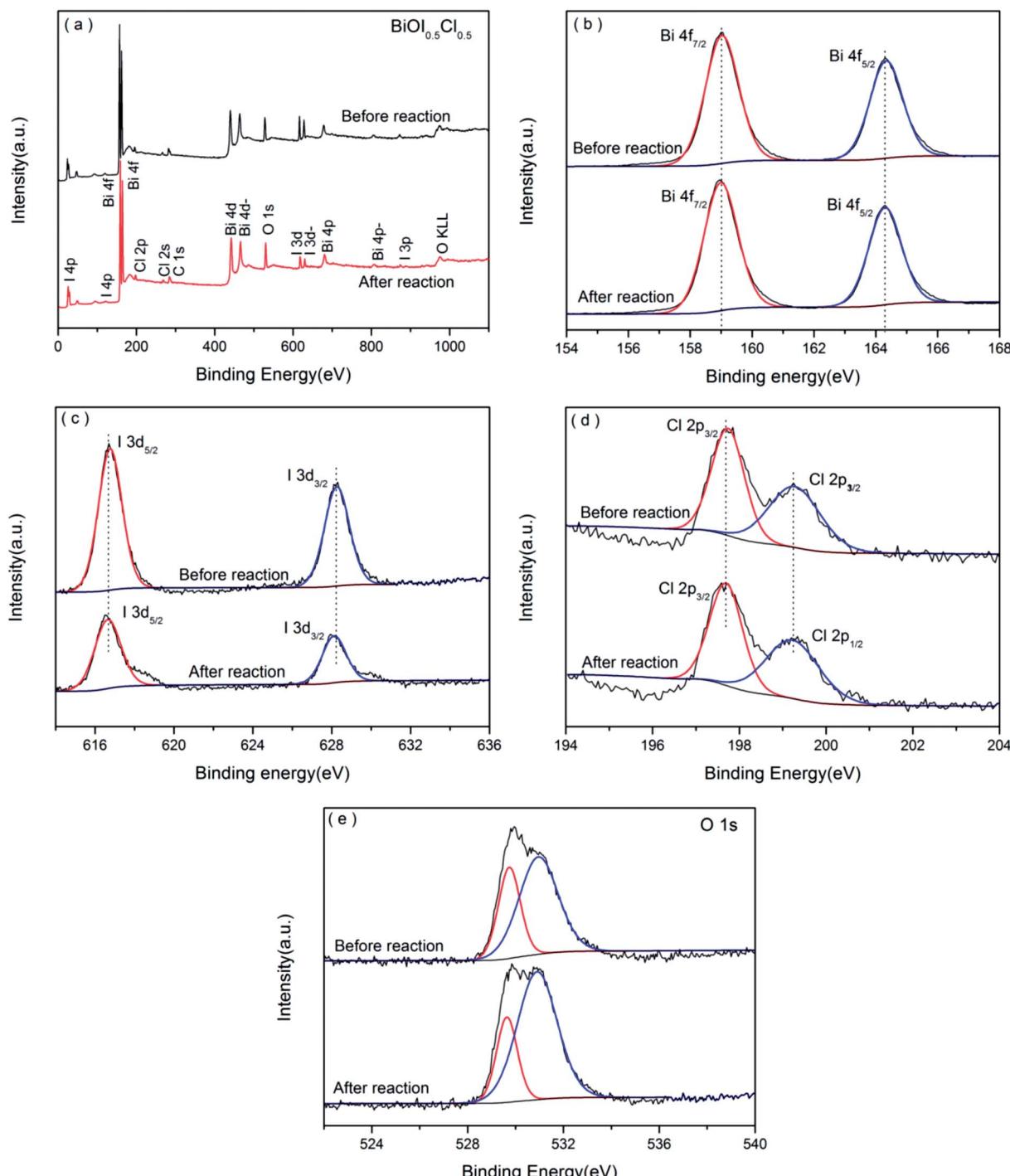


Fig. 4 XPS spectra of catalyst before and after reaction: (a) a typical XPS survey spectrum, and high-resolution XPS of (b) Bi 4f, (c) I 3d, (d) Cl 2p and (e) O 1s.

resulted from the adventitious carbon from atmosphere, which was used to correct the shift of the binding energy at 284.6 eV.⁴² There are two strong peaks at 158.9 eV and 164.3 eV (illustrated in Fig. 4(b)) contributing to Bi 4f_{7/2} and Bi 4f_{5/2}, which demonstrates the presence of Bi–O bonding with trivalent oxidation state.⁴³ Fig. 4(c) shows that the centered peaks at 617.8 eV and 629.4 eV were well corresponded to the inner

electron of I 3d_{5/2} and I 3d_{3/2}, respectively. The binding energies of 197.7 eV and 199.3 eV associated with Cl 2p_{3/2} and Cl 2p_{1/2} (illustrated in Fig. 4(d)), which showed the existence of Cl[−] in BiOI_{0.5}Cl_{0.5}.⁴⁴ The O 1s peaks can be deconvoluted into two bands at 529.7 and 531.0 eV (illustrated in Fig. 4(e)), assigned to Bi–O and I–O bands.^{45,46} What's more, the peaks of O 1s change to higher binding energy value and I 3d change to lower binding

energy in the catalyst after reaction, suggesting that the degradation process takes place on the surface of the $\text{BiOI}_{0.5}\text{Cl}_{0.5}$.⁴⁷

3.2 Degradation efficiency of BPAF in PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ process

The synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ was employed at a dosage of 0.5 g L^{-1} . The removal of BPAF by single PMS, single $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ or PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ is shown in Fig. 5. The pH of the initial solutions was about 7.50 ± 0.05 . As can be seen, in the absence of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$, the degradation of BPAF by PMS was negligible within investigated time scale. Comparatively, the adsorption by single $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ contributed 27.3% removal of BPAF. It can be seen that the adsorption capacity increased quickly in the first 2 min and then the adsorption equilibrium was reached within 10 min. In contrast, BPAF could be appreciably degraded in PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ system with 99.1% removal rate at 60 min. It was obvious that peroxymonosulfate had a positive influence on the degradation of BPAF, indicating the generation of ROS (reactive oxygen species) *via* activation of PMS. 43.9% mineralization was achieved after 60 min reaction, when the nearly complete removal of BPAF was observed, indicating that the significant amounts of reaction by-products were refractory to be oxidized.⁴⁸

Besides, control experiments were carried out to compare the BPAF removal efficiency for various processes at BiOI/PMS and BiOCl/PMS processes within 60 min. As shown in Fig. S2,[†] the synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ showed the highest activity for activation of PMS to produce sulfate radicals and 99.1% BPAF removal was achieved within 60 min, while BiOI and BiOCl possessed a lower activity and provided 31.3% and 48.3% BPAF removal respectively in the same time. The nitrogen adsorption–desorption test was conducted using the BET method, with the results presented in Fig. S3,[†] the inset shows the BJH pore size distribution for the prepared samples. $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ has a large surface area ($44.94 \text{ m}^2 \text{ g}^{-1}$) and pore volume ($0.454 \text{ cm}^3 \text{ g}^{-1}$), which is larger than that of individual BiOI ($2.499 \text{ m}^2 \text{ g}^{-1}$) or BiOCl ($4.992 \text{ m}^2 \text{ g}^{-1}$).³⁶ From the results

of this experiment, we get the conclusion that the structure and surface chemistry of the BiOX would of course be expected to affect its activation of peroxides and thus the overall degradation effectiveness and the synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ has the superior activation.⁴⁹

In the removal process, both the adsorption of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ and the oxidation of PMS activated by $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ were simultaneous for BPAF removal. In this work, all the suspensions were magnetically stirred for 10 min in the dark before adding a certain amount of PMS, in order to attain the adsorption–desorption equilibrium between the $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ and the BPAF molecules.

3.3 Effect of reaction parameters

3.3.1 Effect of catalyst dosages. The influence of different catalyst dosages on the catalytic degradation of BPAF by PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ system was performed in $0.1\text{--}1.0 \text{ g L}^{-1}$ range. Before adding the PMS, the adsorption efficiencies of BPAF by varying the $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ dosage were detected in Fig. S4,[†] the adsorption capacity all increased quickly in the first 2 min and then the adsorption equilibrium was reached within 10 min. Besides, with the improvement of catalyst dosage from 0.5 g L^{-1} to 1.0 g L^{-1} , the adsorption of BPAF just increased about 4.88%.

As shown in Fig. 6(a), the BPAF removal efficiency enhanced from 30.98% to 98.24% with an increase in the loading of prepared sample from 0.1 to 1.0 g L^{-1} within 45 min reaction time. Furthermore, based on a great agreement with pseudo-first-order kinetic model, Fig. 6(b) shows the plot of reaction rate constant (K_{app}) increases from 0.0082 to 0.0959 min^{-1} with enhancing catalyst dose in the system. However, with extra $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ dosages of 1.0 g L^{-1} , the degradation of BPAF slightly changes comparing to the chemical oxidation process with 0.5 g L^{-1} catalyst dosage addition. The observation can be justified by two convincing reasons.^{50,51} First of all, the higher of catalyst provides an overly active site of premenstrual syndrome and BPAF adsorption, then the PMS molecule breaks down into more free sulfate radicals. On the other hand, it might be due to the fact that there is a limit to the amount of PMS and the saturated active sites on the catalyst surface are superfluous, which is accordant with the trend of reaction rate constant. In this section, considering the maximum removal of BPAF, the most suitable catalyst dosage is 0.5 g L^{-1} for the further oxidation experiments.

3.3.2 Effect of PMS loading. BPAF removal in PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ system was performed with 0.5 g L^{-1} of catalyst and various loadings of PMS during 45 min reaction, the effect of molar of PMS/BPAF on the catalytic degradation efficiency is shown in Fig. 7(a). Increasing molar ratio of PMS/BPAF from 1 to 5 continually enhances the BPAF degradation efficiency. When PMS/BPAF equals 5, BPAF degradation efficiency could reach 96.61% at 45 min. Nevertheless, with the further increasing molar ratio of PMS/BPAF to 10 or 20, the degradation efficiency is not promoted any more, which can be confirmed by the reaction rate constant dropped from 0.0772 min^{-1} to 0.0445 min^{-1} in Fig. 7(b). The reason might own to the invalid PMS scavenging, self-quenching of sulfate and hydroxyl radicals

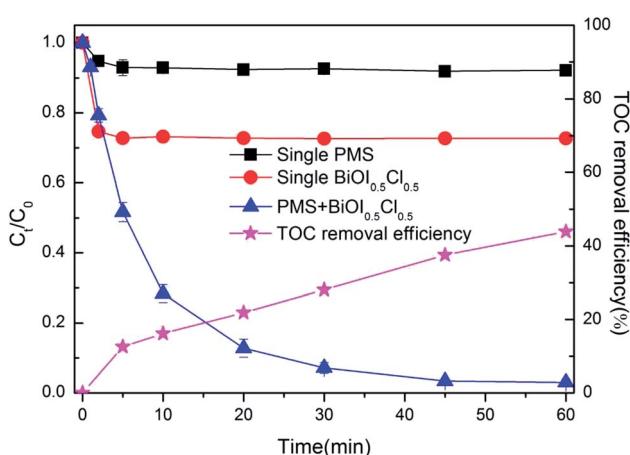


Fig. 5 Removal of BPAF in different reaction systems ($[\text{BPAF}] = 10 \text{ mg L}^{-1}$, $\text{PMS/BPAF} = 5/1$, $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ dosage = 0.5 g L^{-1} , $T = 298 \text{ K}$, initial pH = 7.50).



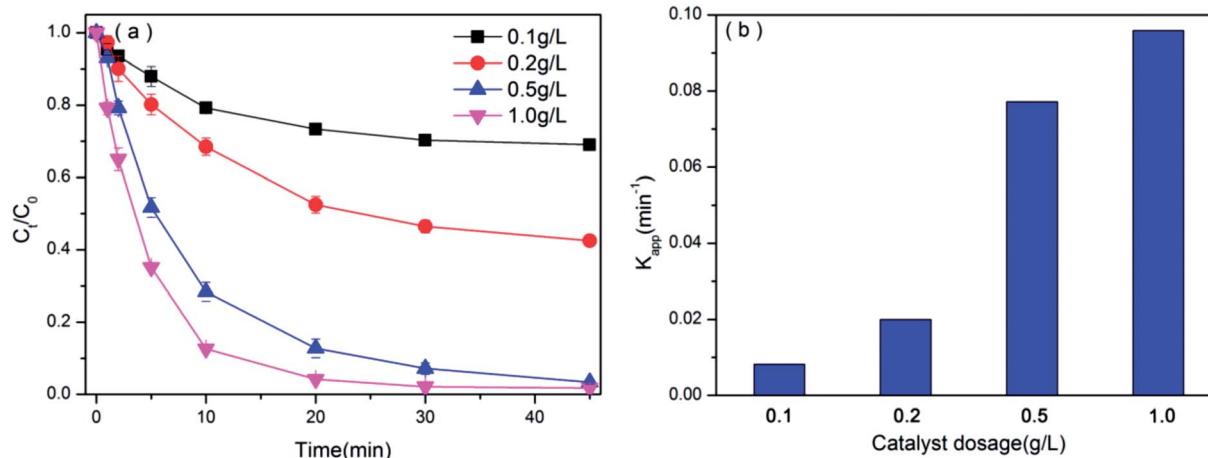
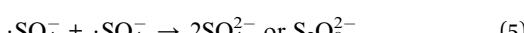
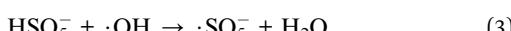
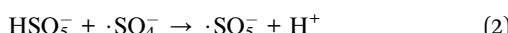


Fig. 6 (a) The degradation efficiencies of BPAF by varying the BiOl_{0.5}Cl_{0.5} dosage ([BPAF] = 10 mg L⁻¹, PMS/BPAF = 5/1, BiOl_{0.5}Cl_{0.5} dosage = 0.1–1.0 g L⁻¹, T = 298 K, initial pH = 7.50) and (b) the degradation rate of pseudo-first-order K_{app} versus catalyst dosage.

is possible when excessive PMS is added (eqn (2) and (3)).^{52,53} Moreover, the reaction (eqn (4)) and the recombination (eqn (5) and (6)) between both sulfate and hydroxyl radicals also play an important role in decreasing the catalytic performance of process.^{12,50} Hence, the optimal molar ratio of PMS/BPAF is 5.



3.3.3 Effect of solution pH. As BPAF degradation by PMS/BiOl_{0.5}Cl_{0.5} system is an aqueous reaction, and natural wastewater has inconstant pH values, pH value is one of the most important influence aspects on chemical reaction rate. Hence, it

is necessary to investigate the influence of pH on the degradation efficiency of BPAF. It must be noted that the addition of acidic PMS into an unbuffered aqueous solution will lead to a significant decrease in the solution pH. As shown in the Fig. 8(a), one of the pH values of 7.50 was the initial pH of the solution, the effect of initial solution pH on BPAF removal was examined in the range of 4.00 to 10.00. The k_1 at pH₀ = 6.04 and pH₀ = 8.00 is 0.0662 and 0.0729 min⁻¹, respectively, which is close to the k_1 at initial pH 7.50 as 0.0772 min⁻¹, suggesting that PMS remained stable under the relatively neutral conditions. However, when the solution became more acidic at pH₀ = 4.00, the degradation extent and kinetics were significantly decreased as k_1 became 0.0487 min⁻¹, which might be caused by the slower PMS activation at relatively low pH values. Previous studies have shown that PMS is believed to be more stable under acidic conditions.⁵⁴ With the increase of the solution pH, the dominant PMS species changes from $\text{HSO}_5^- \cdot$ to $\text{SO}_5^{2-} \cdot$.⁵⁵ In addition, the degradation rate of BPAF increased significantly from pH 8.00 to 9.98, at the alkaline condition, the base activation of PMS could significantly accelerate BPAF degradation

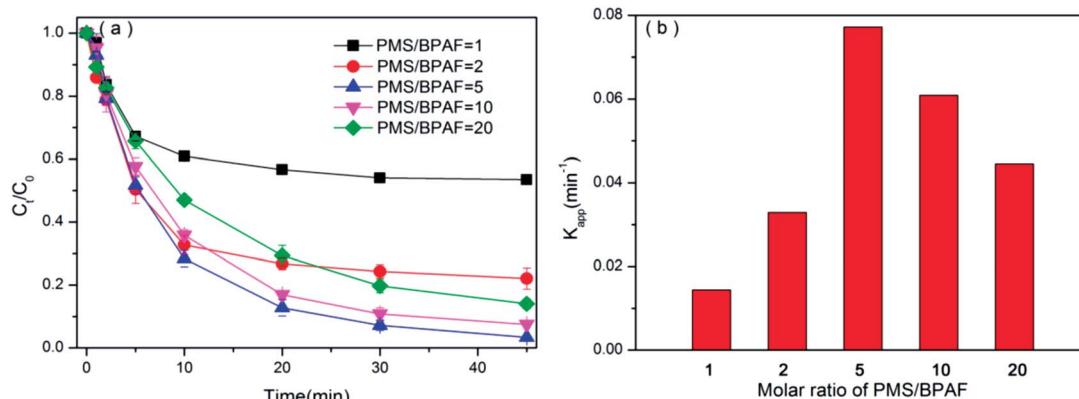


Fig. 7 (a) The degradation efficiencies of BPAF by varying the molar ratio of PMS/BPAF ([BPAF] = 10 mg L⁻¹, BiOl_{0.5}Cl_{0.5} dosage = 0.5 g L⁻¹, PMS/BPAF = 1–20, T = 298 K, initial pH = 7.50.) and (b) the degradation rate of pseudo-first-order K_{app} versus the molar ratio of PMS/BPAF.

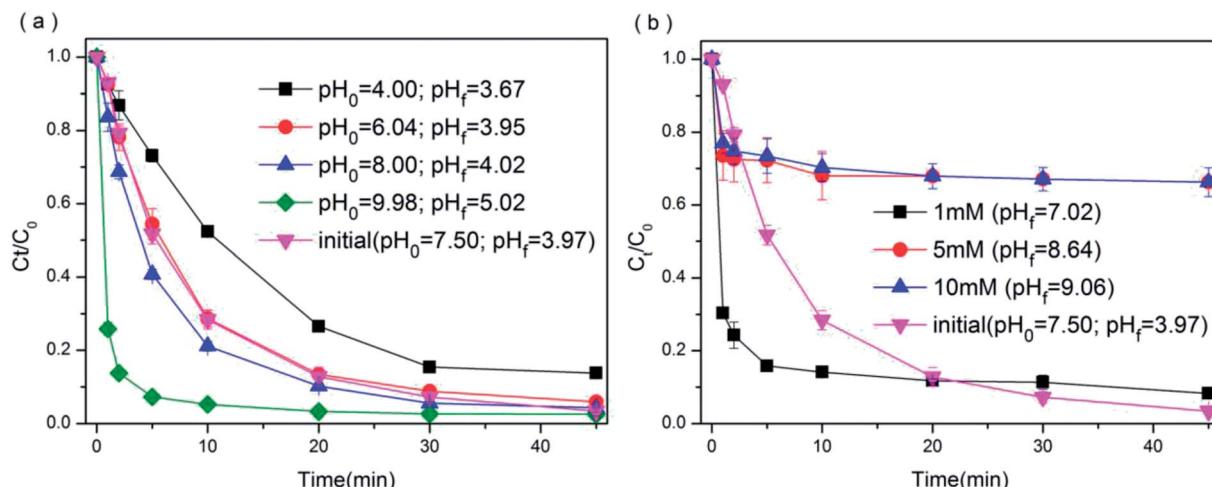
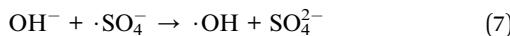


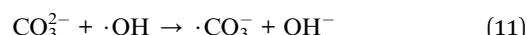
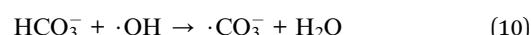
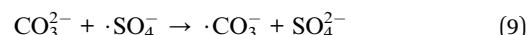
Fig. 8 Effect of initial solution pH (a) and bicarbonate ion concentration (b) on BPAF removal efficiency ($[BPAF] = 10 \text{ mg L}^{-1}$, $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ dosage = 0.5 g L^{-1} , PMS/BPAF = 5, $[\text{HCO}_3^-] = 0\text{--}10 \text{ mM}$, $T = 298 \text{ K}$).

process, which was similar to the previous study on the degradation of benzoic acid using the UV/PMS process.⁵⁶ In this study, the isoelectric point of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ was obtained as 5.6 V (illustrated in Fig. S5† through zeta potential detection), above which the surface charge was negative and below which the catalyst surface was positive. On the other hand, the pK_a of BPAF values of 9.13 and 9.74.⁵⁷ Therefore, the excellent performance of alkaline pH can be attributed to several factors. Firstly, it was easier to activate a larger fraction of $\text{SO}_5^{2-}\cdot$ ions compared to $\text{HSO}_5^{2-}\cdot$ with the base activation of PMS. Secondly, the electrostatic repulsion existed between the catalyst surface and BPAF when pH greater than 9.74, thus the suppression effect of increasing BPAF removal was very weak at the alkaline condition from pH 8.00 to 9.98. Finally, the strong hydroxyl radical at basic pH from sulfate (eqn (7)) also accelerated BPAF degradation efficiency.⁵⁸



3.3.4 Effect of the various anions. To investigate the effect of the various anions in the degradation of BPAF in PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ system, 0–10 mM Cl^- , NO_3^- , $\text{SO}_4^{2-}\cdot$ and HCO_3^- were selected. Compared with Cl^- , NO_3^- and $\text{SO}_4^{2-}\cdot$ as shown in Fig. S6,† a significantly inhibition effect of BPAF degradation was observed when dosing HCO_3^- shown in Fig. 8(b), and the inhibition was enhanced with increasing HCO_3^- concentration. It is well known that bicarbonate strongly influence solution pH. Of note, once HCO_3^- ions were added to the solution, carbonate would be formed and the solution eventually reached the highly pH remained at 7.0–9.0 because of dynamic equilibrium. According to previous studies, HCO_3^- is an effective scavenger for sulfate radicals ($\cdot\text{SO}_4^-$) and hydroxyl radicals ($\cdot\text{OH}$) (eqn. (8)–(11)), which can lead to an inhibitory effect on contaminant degradation.⁵⁹ When added 1 mM HCO_3^- to the solution, the reaction system was changed to alkaline circumstance, which created excellent reaction condition for the

activation of PMS, resulting in the significant increase of the reaction rate,⁶⁰ and the result was similar to the image of pH = 9 in Fig. 8(a). While with the introduction of high dosage of HCO_3^- (5 mM and 10 mM), the BPAF removal was inhibited distinctly, because the HCO_3^- would act as a scavenger for a large proportion of sulfate radicals ($\cdot\text{SO}_4^-$) and hydroxyl radicals ($\cdot\text{OH}$).



3.3.5 Effect of water matrices. To investigate the realistic conditions, the effect of various water matrices including tap water and surface water (from the Minyuan Lake in Sichuan University) on the removal of BPAF was studied using the PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ system, as shown in Fig. S7,† the pH of tap water and surface water is 8.05 and 8.83, respectively. The degradation rate of BPAF was very fast in the first 5 min. However, the removal of BPAF in the PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ system was restrained about 10% in tap water and surface water, which was presumably related to the high TOC levels of tap water (5.57 mg L^{-1}) and surface water (6.87 mg L^{-1}). The main constituents of TOC were involved in hydrophobic, transphilic and hydrophilic DOCs,⁶¹ where they can compete with BPAF for $\cdot\text{OH}$ and $\cdot\text{SO}_4^-$ radicals. Moreover, it is notable that in spite of the higher TOC levels in the two water samples, the PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ system showed high catalytic performance on BPAF removal, indicating that the present reaction system shows the great potential of cleaning BPAF wastewater on the basis of advanced oxidation process.



3.4 Recyclability of catalyst to degrade the BPAF

In order to evaluate the reusability of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ sample, the catalyst was collected by filtration from aqueous solution, washed by ethanol and ultrapure water, then reused for degradation tests. Fig. S8† shows four recycling runs of the BPAF degradation, it was noticeable that the regeneration efficiency (defined as $(1 - C_t/C_0)\%$) remained about 47.5% after 4 cycles. The removal of BPAF were decreased 14.53%, 42.05% and 49.10% after each recycling, respectively. This activity decrease in the regeneration efficiency is possible due to the remarkable amounts of both BPAF molecules and by-products are adsorbed on the surface of catalyst, then the catalytic surface activity and the rate of mineralization are limited. Furthermore, some pores on the catalyst surface were blocked, due to the adsorption of BPAF and intermediates from degradation.⁶²

3.5 The mechanism of $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ to active PMS for degradation of BPAF

3.5.1 Identification of the reactive species. For the PMS/ $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ system, it was expected that hydroxyl and/or sulfate radicals can be produced and then react with BPAF molecules. To identify the oxidation of the reaction system, ESR experiment using 5,5-dimethyl-1-pyrroline (DMPO) as a radical spin trapping agent were performed. As shown in Fig. 9, the characteristic signals of 5,5-dimethyl-2-pyrroline-N-oxyl (DMPOX) with heptet were detected at 3 min, which were formed by the oxidation of DMPO,⁶³ indicating that some strong ROS were produced in the reaction system. Previous studies had proposed transformation pathway of DMPO to DMPOX.^{64,65} The fact that we did not yet detect $\cdot\text{OH}/\cdot\text{SO}_4^-$ radicals *via* the ESR test, which did not mean that they were absent. In addition, the signals of DMPOX with heptet were more obvious after adding NaF, which proved that the surface radicals played an important role in the process of degradation of BPAF.

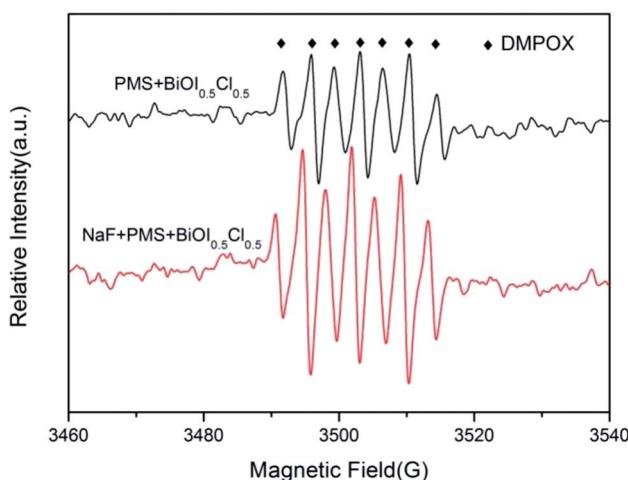


Fig. 9 ESR spectra of $\text{BiOI}_{0.5}\text{Cl}_{0.5}/\text{PMS}$ and $\text{BiOI}_{0.5}\text{Cl}_{0.5}/\text{PMS}/\text{NaF}$ system ($\text{BiOI}_{0.5}\text{Cl}_{0.5}$ dosage = 0.5 g L^{-1} , PMS = 0.3 mM , $[\text{NaF}] = 0.03 \text{ mol L}^{-1}$, $T = 298 \text{ K}$).

To distinguish contribution of hydroxyl and/or sulfate radicals to BPAF degradation, the radical intermediates were probed *via* radical quenching tests: excess *tert*-butyl alcohol (TBA) and methanol (MA) were added into the reaction solution as radical scavengers. MA, a kind of alcohol containing an α -hydrogen, can react with $\cdot\text{SO}_4^-$ and $\cdot\text{OH}$ at significant rates ($k_{\cdot\text{SO}_4^-/\text{MA}} = 1.6\text{--}7.7 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$, $k_{\cdot\text{OH}/\text{MA}} = 1.2\text{--}2.8 \times 10^9 \text{ M}^{-1} \text{ s}^{-1}$) while TBA is a specific scavenger for $\cdot\text{OH}$ ($k_{\cdot\text{OH}/\text{TBA}} = 3.8\text{--}7.6 \times 10^8 \text{ M}^{-1} \text{ s}^{-1}$).⁶⁶ Fig. 10(a) showed that the inhibiting effects of BPAF degradation were not obvious when MA or TBA was added into the $\text{BiOI}_{0.5}\text{Cl}_{0.5}/\text{PMS}$ suspensions. 14.57% and 9.21% of the BPAF degradation were inhibited by MA and TBA, respectively. The reason may be that hydrophilic MA and TBA are superior to competing for free radicals in the liquid phase rather than for large reaction with active radicals on the surface of the catalyst. However, from the inner figure of Fig. 10(a), k_1 value decreased from 0.0772 min^{-1} to 0.435 min^{-1} , revealing that the PMS activation was inhibited by TBA and $\cdot\text{OH}$ may be involved in BPAF degradation. What's more, in the presence of methanol, k_1 value decreased to 0.0359 min^{-1} . Therefore, the above results suggest that the methanol adequately suppressed PMS activation and the sulfate and hydroxyl radicals had significant effect on the degradation of BPAF, the primary active species was surface bound $\cdot\text{OH}/\cdot\text{SO}_4^-$ in the catalytic oxidation process. Furthermore, the degradation of the PMS in $\text{BiOI}_{0.5}\text{Cl}_{0.5}/\text{PMS}$ and $\text{BiOI}_{0.5}\text{Cl}_{0.5}/\text{PMS}/\text{BPAF}$ system was detected. As shown in the Fig. S9,[†] the degradation rate of PMS increased 18.31% after adding BPAF, indicating that the PMS was involved in the reaction.

3.5.2 The mechanism discussion

To further analyze the role of surface bound $\cdot\text{OH}/\cdot\text{SO}_4^-$ generated by the surface of the synthesized $\text{BiOI}_{0.5}\text{Cl}_{0.5}$, different concentrations of NaF were added to the solution. Previous researches had shown that F^- in the solution could desorb $\cdot\text{OH}$ bound on catalyst by forming strong fluoride-exchange.^{67,68} As expected in Fig. 10(b) and S9(a),[†] the addition of F^- improved the BPAF degradation and with the more NaF dosage, the more obvious of the PMS degradation effect, the degradation rate of PMS increased 18.20% compared with $\text{BiOI}_{0.5}\text{Cl}_{0.5}/\text{PMS}/\text{BPAF}$ system. While there was no obvious promotion in degradation of PMS when adding MA or TBA (shown in Fig. S9(b),[†]). It should be pointed out that, after F^- surface modification, the reduce of surface free radicals might promote the production of radicals in bulk. The results indicated PMS mainly bonded with the active sites of the $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ catalyst to generate the surface radicals, which were important rather than the radicals in bulk for BPAF degradation.

Thus, a possible catalysis mechanism of PMS activation by $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ for BPAF degradation are proposed. Because PMS is an acidic oxidant with standard reduction potentials for the aqueous half-reactions, which was mainly in the form of HSO_5^- and SO_5^{2-} ions in the solution (eqn (12) and (13)). HSO_5^- and SO_5^{2-} were adsorbed on the surface of the $\text{BiOI}_{0.5}\text{Cl}_{0.5}$ due to the electrostatic attraction, which were easily activated by $\text{Bi}^{3+}/\text{Bi}^{5+}$ couples to generate $\cdot\text{SO}_4^-$ and $\cdot\text{SO}_5^-$ (eqn. (14)–(16)). Under



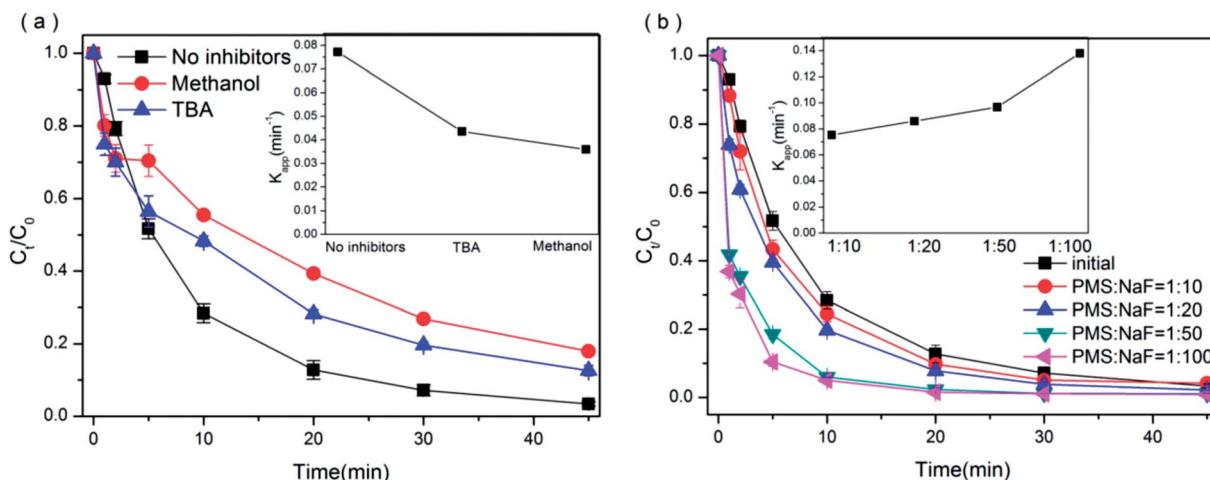


Fig. 10 (a) BPAF degradation in the presence of methanol and TBA as a radical scavenger and (b) in NaF/PMS/BiOI_{0.5}Cl_{0.5} system. ([BPAF] = 10 mg L⁻¹, BiOI_{0.5}Cl_{0.5} dosage = 0.5 g L⁻¹, PMS/BPAF = 5, [methanol] = [TBA] = 0.075 mol L⁻¹, [NaF] = 0–0.015 mol L⁻¹, T = 298 K).

strong alkaline conditions, $\cdot\text{OH}$ turns out to be the major radical, as shown in eqn (17). In addition, the formation of $\cdot\text{OH}$, which also can be attributed to reaction between $\cdot\text{SO}_4^-$ and water (eqn (18)).^{69,70} Finally, $\cdot\text{OH}$ and $\cdot\text{SO}_4^-$ on surface of the catalyst attacked adsorbed BPAF organic compounds and triggered a series of further degradations. Based on the above results, a feasible mechanism for the degradation of BPAF concerning the interaction between BiOI_{0.5}Cl_{0.5} and PMS is proposed in Fig. 11.

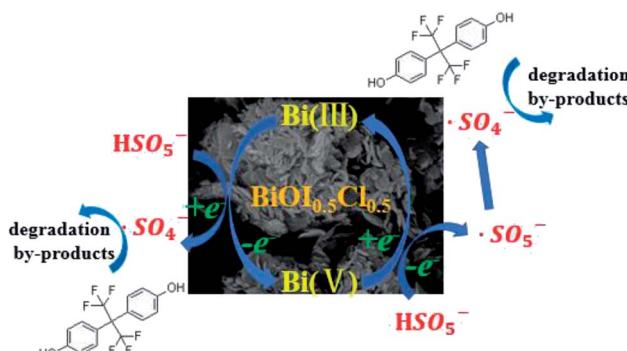
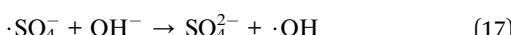
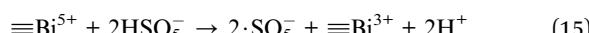
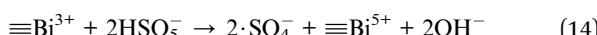
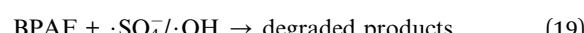


Fig. 11 Proposed mechanism of BPAF degradation in the BiOI_{0.5}Cl_{0.5}/PMS process.



4. Conclusion

The obtained results showed that the synthesized catalyst had a good catalytic performance in the activation of PMS to degrade BPAF. Under optimum experimental conditions, the removal efficiencies of BPAF, TOC were found to be 99.1 and 43.9%, respectively. This degradation activity could be enhanced by increasing catalyst dose and initial solution pH. Bicarbonate ions adversely affected the degradation rate of BPAF, while chloride, nitrate and sulfate ions did not represent any effect on the removal of BPAF. In the tap water and surface water, the present reaction system still showed the great potential of cleaning BPAF wastewater. Radical scavengers' experiments verified the dominant active species generated in the BiOI_{0.5}Cl_{0.5}/PMS system were surface-bound $\cdot\text{SO}_4^-$ and $\cdot\text{OH}$ radicals, thus a possible degradation mechanism of BPAF was proposed. In conclusion, the study provides an option and would be a promising effective technology for BPAF removal.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

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