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Efficient synthesis of epoxybutane from butanediol via a two-step process†

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A novel approach for the synthesis of epoxybutane *via* decarboxylation of butenyl carbonate derived from butanediol was developed for the first time. For the carbonylation of butanediol with dimethyl carbonate, NaAlO₂ has exhibited excellent catalytic activity under mild reaction conditions. The yield of butenyl carbonate reached as high as 96.2%. NaAlO₂ provides suitable acid–base active sites to promote the transesterification reaction of butanediol and dimethyl carbonate. For the following step of decarboxylation of butenyl carbonate, ionic liquid 1-butyl-3-methylimidazolium bromide could effectively catalyze the decarboxylation process both in batch or continuous processes. Moreover, the catalytic mechanism for the crucial step of decarboxylation of butenyl carbonate over 1-butyl-3-methylimidazolium bromide was explored using DFT calculations. The results showed that the electrostatic and hydrogen-bond effects of 1-butyl-3-methylimidazolium bromide played a crucial role for the generation of epoxybutane. Briefly, the Br anion of the ionic liquid attacks the methylene of the ring and the H of the ionic liquid cation attacks the carbonyl oxygen, which facilitated the five-ring opening and subsequent decarboxylation to form BO. This study not only provided a new and green synthetic route for producing epoxybutane, but also contributed to the effective utilization of butanediol, which is inevitably produced as by-product in the process of coal to ethylene glycol, suggesting a promising application in the clean manufacture of epoxybutane with inexpensive cost.

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Introduction

Epoxybutane (BO) is an important organic chemical raw material and mainly used to synthesize polyether polyol,¹ which is being used in the manufacture of value-added polyurethane.² Until now, the monomer of polyether polyol is mainly chosen from ethylene oxide (EO), propylene oxide (PO) and BO.^{3,4} Compared with EO and PO, BO not only possess a similar activity that can react with hydrogen containing compounds, such as water, alcohol, phenols, sulfur alcohol, ammine, acids and so on, but also has unique properties that can improve the durability of diphenylmethane diisocyanate (MDI)-based polyurethane. However, owing to the restrictions from the severe safety issues in the production technology and the expensive cost, the market share of polyether polyol derived from BO is

smaller than that made from EO and PO. At present, BO is being synthesized by chlorohydrin process, direct oxidation method and indirect oxidation method.⁵ These conventional manufacturing processes of BO suffer from several drawbacks, such as environmental pollution, harsh operation conditions, and high cost of the raw materials. In the chlorohydrin route, butylene is reacted with water and Cl₂ to produce chlorobutanol, which is then reacted with base to generate BO. In the production of BO by chlorohydrin process, salt waste and large amount of wastewater will be produced. In addition, the production of large amounts of hydrochloric acid will corrode the equipment. In the indirect oxidation method, the styrene process is the main process. Ethyl benzene is first oxidized with O₂, where the peroxy compound is used to generate BO from butylene. The residual aromatic alcohol is then dehydrated to styrene. This method is energy-intensive and can produce sizeable waste streams. The direct oxidation method is the oxidation of butene to BO by hydrogen peroxide.⁶ Its disadvantages are harsh operation conditions and high costs of the raw materials. In this context, a new and environmentally friendly route for the synthesis of BO from butanediol *via* butenyl carbonate (BC) intermediate was investigated in this paper. The new route includes the synthesis of BC from butanediol and DMC, followed by decarboxylation of BC to BO, as shown in Scheme 1.

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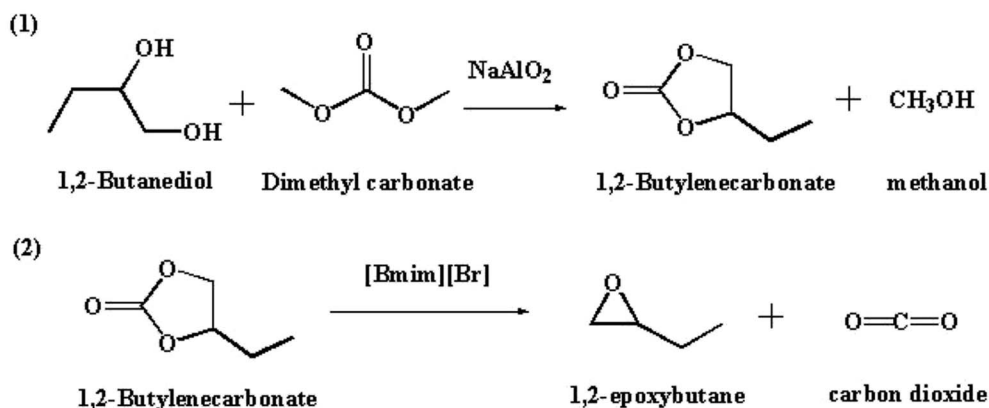
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Scheme 1 (1) Synthesis of BC from butanediol and DMC, (2) BC decarboxylation to BO.

In the previous reports, the intermediate of BC is generally synthesized from BO and CO₂, which have the disadvantages of relatively high pressure and high cost.^{7–12} The development of a new route for the effective synthesis of BC *via* an inexpensive route is highly desired. In terms of this reason, a new approach based on the transesterification of butanediol with dimethyl carbonate (DMC) was developed in our group, where the crucial raw material of butanediol could be readily obtained from the by-product, which is inevitably produced in the well-established process of coal to ethylene glycol process.¹³ Owing to the high activity of alcohol hydroxyl group, ethylene glycol can be further hydrogenated to ethanol in the presence of hydrogen. Then the ethylene glycol can react with ethanol to produce butanediol due to the Golbert reaction.^{14,15} Thus, inexpensive butanediol is produced in large quantities in the coal to ethylene glycol process.

The development of effective catalyst system is crucial for the synthesis of BC from butanediol and DMC. Sodium aluminate (NaAlO₂) is generally used as a cheaply available and highly basic catalyst.^{16–22} As it is insoluble in many organic solvents, NaAlO₂ can be used as a truly heterogeneous catalyst.^{23,24} Some researchers reported that NaAlO₂ have been used as an active catalyst for a number of base-catalyzed reactions (transesterification, isomerization,^{25,26} and condensation). Therefore, owing to its unique features in basicity, NaAlO₂ was also employed in the synthesis of BC *via* carbonylation of butanediol with DMC. In addition, the synthesis of BO from decarboxylation of BC is worthy of special attention in terms of economic and environmental points of view. Imidazolium-based ionic liquid was demonstrated to be capable of catalyzing the decarboxylation of glycerol carbonate.^{27–32} In this work, imidazolium-based ionic liquids were tested and applied as catalyst in the decarboxylation of BC to produce BO.

In this paper, a novel approach for the synthesis of epoxybutane (BO) *via* decarboxylation of butenyl carbonate (BC) derived from butanediol were successfully developed. In this process, NaAlO₂ was used as catalyst in the transesterification of butanediol with dimethyl carbonate (DMC) to produce BC, and subsequently 1-butyl-3-methylimidazolium bromide ([Bmim][Br]) was used as catalyst in the decarboxylation of BC to

produce BO. The influence of reaction parameters, *e.g.* reaction temperature, reaction time and catalyst amount on the catalytic performance were optimized. Furthermore, the crucial catalytic mechanism of decarboxylation of BC to BO is deeply explored using DFT calculations.

Experimental

Materials

Dimethyl carbonate (99%) and decane were purchased from Aladdin Industrial Corporation. Butanediol (98%), 1,2-epoxybutane and butylene carbonate (98%) were purchased from Tokyo Chemical Industry Co., Ltd. Sodium aluminate and zinc bromide were purchased from Sinopharm Chemical Reagent Co., Ltd. 1-Butyl-3-methylimidazolium bromide ([Bmim][Br]), 1-butyl-3-methylimidazolium chloride ([Bmim][Cl]), 1-butyl-3-methylimidazolium iodide ([Bmim][I]) and 1-butyl-3-methylimidazolium tetrafluoroborate ([Bmim][BF₄]) were purchased from Shanghai Cheng Jie Chemical Co. LTD.

Catalytic activity test

Carbonylation of butanediol with dimethyl carbonate. Carbonylation reaction between butanediol and DMC was carried out in a 100 mL autoclave with a magnetic stirring. In a typical procedure, 22.4 g dimethyl carbonate and 4.5 g butanediol was loaded to the reactor, followed by the addition of 10 wt% catalyst (0.45 g). The mixture was heated to 120 °C with magnetic stirring and condenser, and kept for a certain time, and then cooled down to the room temperature. After the reaction, the catalyst was separated by centrifugation, and then washed by DMC and reused again.

Decarboxylation of BC. Decarboxylation of BC was tested by both batch and continuous processes. In batch cases, BC was loaded to the 250 mL three-necked flask equipped with condensers and an oil bath, and was decarboxylated into BO in the presence of catalyst at a specified temperature under the vacuum degree of 0.005 MPa. BO produced during the decarboxylation reaction was collected through a three-stage condensing device.



While for continuous process, BC was loaded to the 250 mL three-necked flask from the sample bottle which was put on the balance. Thus, the amount of the BC introduced into the reactor can be directly calculated. Then BC was continuously fed into the flask by vacuum aspiration, and the feeding rate of BC was controlled by the valve under the vacuum degree of 0.005 MPa. During the reaction, the product was continuously collected in the bottom of flask and weighted. As a result, the yield of BO can be calculated.⁶ The device for the continuous process is shown in Fig. S1.†

Two-step process for the synthesis of BO by using BC from butanediol. In two-step process, firstly, the synthesis of BC was carried out in a 1L reaction autoclave under the optimal conditions. Then BC was purified by the distillation process, and was consequently charged to the decarboxylation reactor, and converted to BO through the continuous device.

Product analysis

After the reaction, the reaction mixture was analyzed using a gas chromatograph (GC-2010) using a RTX-WAX (30 m × 0.25 mm × 0.25 μm) capillary column and a flame ionization detector (FID). The decane was used as an internal standard for the quantitative analysis of butanediol, BC and BO. BC conversion, BC yield, BC selectivity, BO yield and formation rate of the decarboxylation of butenyl carbonate were calculated using the following equations:

$$C_{\text{butanediol}} = \frac{M_{\text{rt}0} - M_{\text{rt}n}}{M_{\text{rt}0}} \times 100\%$$

$$Y_{\text{BC}} = \frac{M_{\text{it}n}}{M_{\text{rt}0}} \times 100\%$$

$$S_{\text{BC}} = \frac{M_{\text{it}n}}{M_{\text{rt}0} - M_{\text{rt}n}} \times 100\%$$

$$Y_{\text{BO}} = \frac{M_{\text{BO}}}{M_{\text{BC}}} \times 100\%$$

$$\text{FR}_{\text{BO}} = \frac{M_{\text{BO}}}{M_{\text{catalyst}}h} \times 100\%$$

where $C_{\text{butanediol}}$ is the conversion of butanediol; Y_{BC} is the yield of BC; S_{BC} is the selectivity of BC; $M_{\text{rt}0}$ is the initial mole amount of butanediol; $M_{\text{rt}n}$ is the mole amount of butanediol; $M_{\text{it}n}$ is the mole amount of BC produced; Y_{BO} is the yield of BO; M_{BO} is the mole amount of BO produced; M_{BC} is the initial mole amount of BC; FR_{BO} is the formation rate of the decarboxylation of butenyl carbonate; M_{catalyst} is the mole amount of catalyst; h is the reaction time of decarboxylation.

Characterization

XRD: phase structure of catalyst was determined with X-ray diffractometer using Cu K α radiation. The working voltage

and current were 40 kV and 40 mA, respectively. The 2θ range used was from 5° to 90°.

Infrared (IR) analysis of the catalyst samples (diluted by KBr at a ratio of 1 : 200) were characterized using a TENSOR-27 spectrometer (Bruker) with the scanning range is 400–4000 cm^{-1} .

¹H NMR spectra were recorded on a Bruker Avance III 600 MHz spectrometer in CDCl₃ with tetramethylsilane (TMS) as internal standard.

CO₂-TPD was carried out to determine the basicity of samples using an AutoChem II 2920 equipped with a thermal conductivity detector (TCD). About 100 mg of sample was heated *in situ* in 25 mL min⁻¹ of helium gas at 220 °C for 0.5 h in order to remove adsorbed impurities. After that, the sample was cooled to 50 °C and saturated with 25 mL min⁻¹ of 10% vol% CO₂/He for 1 h. Weakly adsorbed CO₂ was eliminated by flushing with a 25 mL min⁻¹ of helium gas at 50 °C for 0.5 h, and increased to 800 °C with a ramping rate of 10 °C min⁻¹.

N₂ sorption measurement was performed using a Quantachrome Autosorb-1 at 77 K and the samples were degassed under vacuum at 573 K for 3 h before measurement. The specific surface areas were calculated by the Brunauer–Emmett–Teller (BET) method using N₂ adsorption data ranging from $P/P_0 = 0.05$ to 0.30.

Computational calculations

All the Density Functional Theory (DFT) calculations were performed using the Dmol3 program available in Materials studio 6.1 package. The generalized gradient approximation (GGA) with Perdew–Wang 1991 function was used.^{33,34} Geometry optimizations for minima were carried out using the standard Bery algorithm in redundant internal coordinates up to the neighborhood of the solution. If not converged, the optimization was further continued using analytical second derivatives. Optimizations for transition states were carried out with an initial guess for the transition state being generated from manual manipulation of the geometry using MOLDEN. Complete linear synchronous transit and quadratic synchronous transit (LST/QST) calculations were performed to obtain the structures of transition states (TS).^{35,36}

Results and discussion

Transesterification of butanediol with DMC to synthesize BC

Effect of reaction conditions on the BC synthesis. A proper catalyst with unique basicity is very important for obtaining effective catalytic performance in the transesterification reaction. In this work, sodium aluminate (NaAlO₂) were used as the catalyst for BC synthesis. NaAlO₂ provides a suitable acid–base active site to promote the transesterification reaction of butanediol and DMC. The acid–base active sites of NaAlO₂ is detected by CO₂-TPD and NH₃-TPD, and the results is shown in Fig. S2.† From Fig. S2,† it can be observed that the desorption peaks at 256 °C of NaAlO₂ are attributed to the interaction of CO₂ with weak basic sites. Meanwhile, the desorption peaks at 160 °C and 500 °C are attributed to the interaction of NH₃ with



acid sites of NaAlO_2 , which possibly originated from Lewis acid site of Al^{3+} . It means that NaAlO_2 contains both acid and base active sites. In the transesterification reaction, acid site accepts electrons as electron acceptors, while base site gives electrons as electron donors. They activate transesterification reaction substrates synergistically.

The effects of reaction conditions, such as catalyst amount, reaction temperature, DMC/butanediol molar ratio, as well as reaction time on the transesterification performance were investigated and discussed below.

The effect of catalyst amount ranging from 0.1–5.0 wt% (based on the butanediol amount) on the reaction was investigated in Fig. 1. It is observed that, without catalyst, the reaction efficiency is so low that, the butanediol conversion is only 6.7%. However, the introduction of 0.1 mol% catalyst led to the rapid increase of butanediol conversion to 95.5%, then hardly changed when continue increasing the catalyst amount to 3.0 wt%. Thus, the 0.1 mol% catalyst was good enough for the transesterification of butanediol with DMC.

Fig. 2 showed the influence of temperature on the BC synthesis with NaAlO_2 as catalyst. It can be observed that the BC is effectively synthesized with excellent selectivity within the temperature range investigated. Meanwhile, the butanediol conversion increased sharply from 52.2% to 94.8% at first with the temperature increased from 50 to 105 °C. When the temperature increased to higher temperatures, *i.e.* 120 °C and 150 °C, butanediol conversion gradually increased to 96.9%. The above results indicated that temperature had a great effect on butanediol conversion, and the conversion rate became very fast and nearly completely converted at 120 °C.

The effect of molar ratio of DMC and butanediol was studied, as shown in Fig. 3. The conversion of butanediol increased from 91.5% to 96.6% with increasing the molar ratio from 2 to 4, and further increase the molar ratio did not show any change. Thus DMC/butanediol molar ratio of 4 : 1 was chosen as the preferred

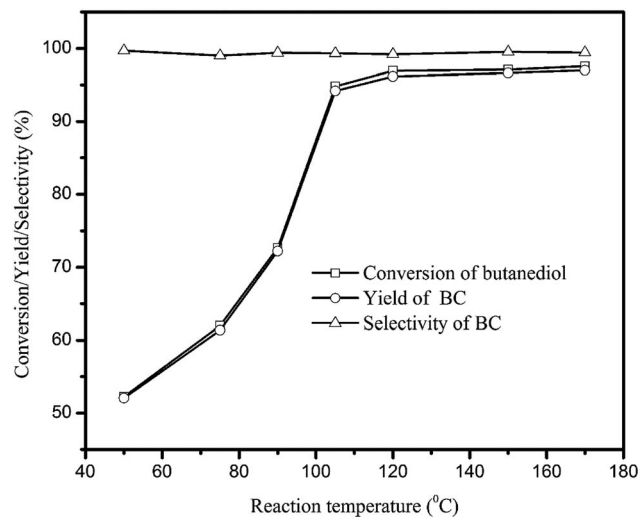


Fig. 2 Effect of temperature on BC synthesis. (Reaction conditions: butanediol: 0.05 mol; $n(\text{DMC})/n(\text{butanediol}) = 5$; catalyst: 10 wt% of NaAlO_2 based on butanediol; reaction time: 60 min.)

value, in which case butanediol was completely converted into the decarboxylation precursor of BC.

At last, the effect of reaction time was also studied, and the result is shown in Fig. 4. As can be seen from the figure, the reaction is very fast that the butanediol conversion reached to 78.9% only after 15 min, and then quickly reached to 93.8% and kept unchanged anymore at 30 min, which means 30 min is the equilibrium and optimum time for this reaction.

Recyclability of the catalysts

Additionally, the recyclability of the catalyst was also studied. After the reaction, the catalyst was separated by centrifugation,

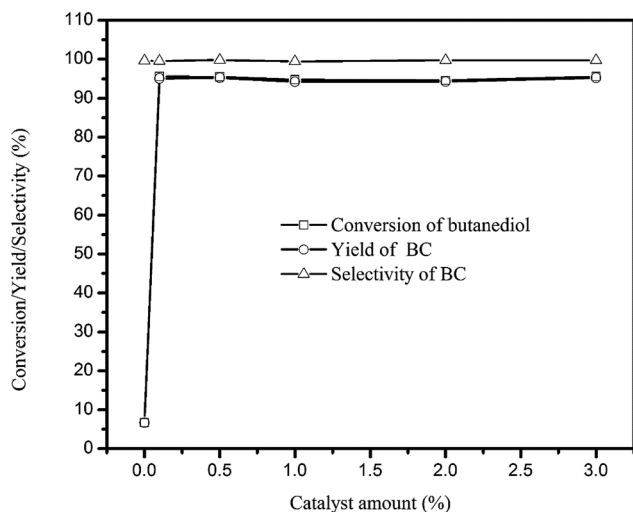


Fig. 1 Effect of catalyst amount on BC synthesis. (Reaction conditions: butanediol: 0.05 mol; $n(\text{DMC})/n(\text{butanediol}) = 4$; temperature: 120 °C; reaction time: 60 min.)

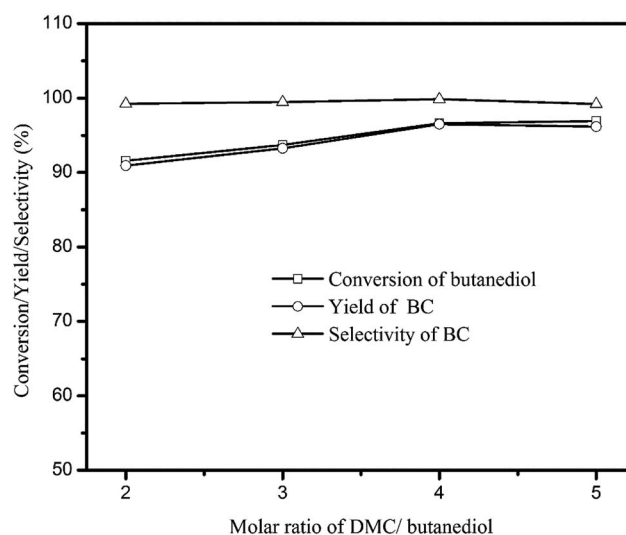


Fig. 3 Effect of molar ratio between DMC and butanediol on BC synthesis. (Reaction conditions: butanediol: 0.05 mol; catalyst: 10 wt% of NaAlO_2 based on butanediol; temperature: 120 °C; reaction time: 60 min.)



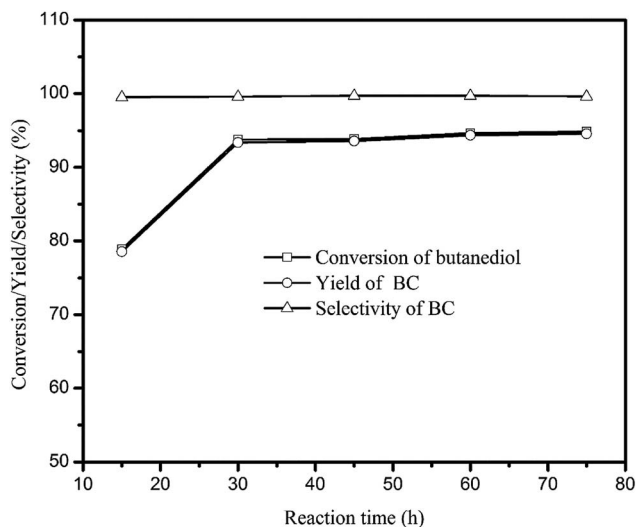


Fig. 4 Effect of reaction time on BC synthesis. (Reaction conditions: butanediol: 0.05 mol; $n(\text{DMC})/n(\text{butanediol}) = 4$; catalyst: 10 wt% of NaAlO_2 based on butanediol; temperature: 120 °C.)

then washed by DMC and reused again. The result was shown in Fig. 5, it can be seen that, although BC selectivity was almost remained at 99.7%, the butanediol conversion increased from 92.2% to 96.8% when fresh and reused catalyst were applied, respectively. From this result, it can be deduced that the catalytic activity was further enhanced during the first run, and exhibiting higher catalytic activity in the subsequent recycling experiments.

To further explore the structure change of the catalyst during the reusability, the fresh and reused catalyst were investigated by XRD. As can be seen in Fig. S3,† the characteristic peaks of sodium aluminate at $2\theta = 20.8, 30.2, 33.6, 35.0, 45.8, 48.4, 51.8, 58.1, 61.1, 63.0$ (JCPDS 00-019-1179 and 01-083-0316) were clearly observed in both fresh and reused catalyst, and the peak

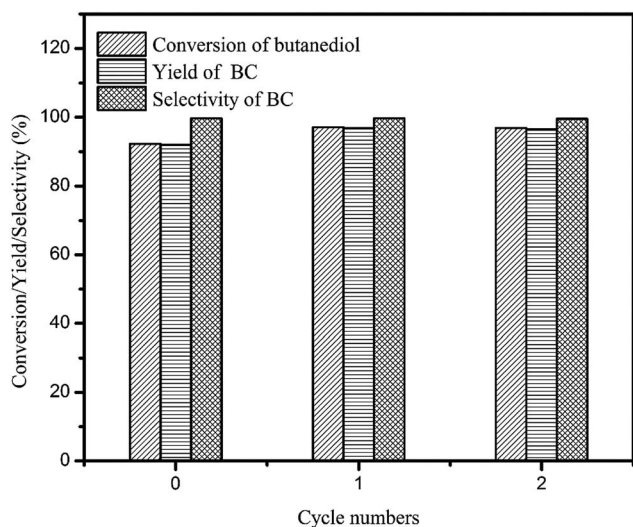


Fig. 5 Reuse of catalyst NaAlO_2 . (Reaction conditions: butanediol: 0.05 mol; $n(\text{DMC})/n(\text{butanediol}) = 4$; catalyst: 10 wt% of NaAlO_2 based on butanediol; temperature: 120 °C; reaction time: 30 min.)

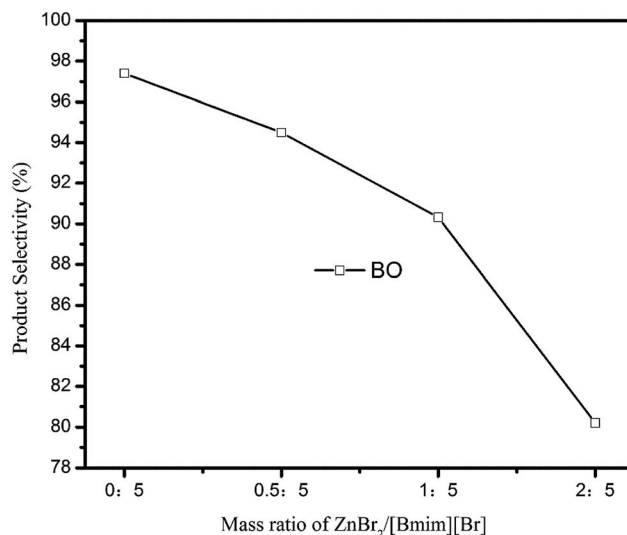


Fig. 6 Effect of promoter/catalyst mass ratio on BO synthesis. (Reaction conditions: BC: 40 g; catalyst: 0.83 g; reaction time: 90 min; vacuum degree: 0.005 MPa.)

strength is basically the same. These results suggested that crystalline structure and catalyst activity primarily remained during the recycle process, the results of which is consistent with the experimental data.

From the spectra in Fig. S4,† we can see that the peaks at 3300 to 3700 cm^{-1} and 1650 cm^{-1} are attributed to the bending mode of the $-\text{OH}$ group of physically adsorbed water and the peaks at 1455 cm^{-1} corresponding to the carbonate after it was exposed to the air. The peaks at 810, 610, 560 and 490 cm^{-1} had no obvious changes after the sample was used for the reaction, indicating that the composition of NaAlO_2 is stable under the reaction conditions used.

The basicity of the NaAlO_2 (both before and after reaction) was evaluated by using CO_2 temperature-programmed desorption (TPD) measurements as shown in Fig. S5.† The used

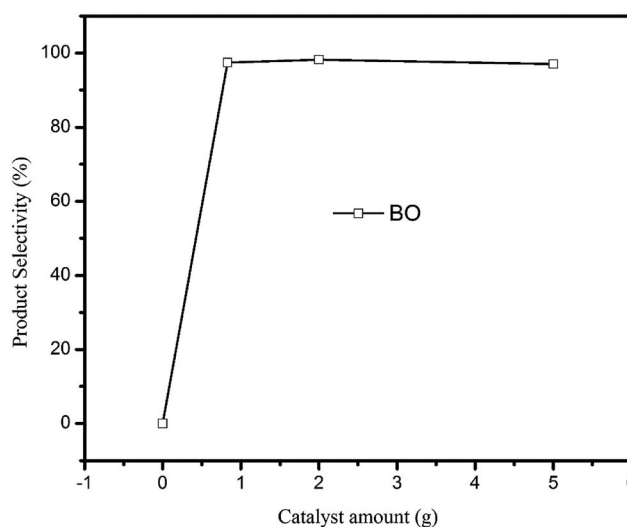


Fig. 7 Effect of catalyst amount on BO synthesis. (Reaction conditions: BC: 40 g; reaction time: 90 min; vacuum degree: 0.005 MPa.)



NaAlO₂ was separated by centrifugation and washed by ethanol. Then dried it at 100 °C for 4 h. To evaluate the amount of different types of basic sites, the desorption curve can be arbitrarily differentiated into two regions corresponding to weak and strong basic sites (below or above 400 °C, respectively). The desorption peaks at 256 °C for fresh NaAlO₂ indicated basic sites of weak strength. The desorption peaks at 460 °C and 740 °C of used NaAlO₂ can be attributed to the interaction of CO₂ with strong basic sites because the higher temperature is needed to desorb the CO₂ adsorbed on the more strongly basic sites.

The surface area is also tested using BET characterization. As shown in Table S1,† it can be observed that the BET surface area of the used NaAlO₂ (13.6 m² g⁻¹) was higher than that of fresh NaAlO₂ (1.7 m² g⁻¹). The enlarged surface area could be presumably attributed to the leaching of sodium element. Thus, more basic sites could be exposed.

Synthesis of BO from decarboxylation of BC

As aforementioned, imidazolium-based ionic liquids have been shown to be capable of catalyzing the decarboxylation of alkylene carbonates. In this work, [Bmim][Br] ionic liquid was tested and applied as the catalyst for the decarboxylation of BC. With using ionic liquids [Bmim][Br] as catalyst and ZnBr₂ as promoter. The decarboxylation of BC can be effectively carried out, and the effect of reaction conditions, such as promoter/catalyst mass ratio, catalyst amount, were investigated and discussed below.

Fig. 6 showed the influence of promoter/catalyst mass ratio on the BO synthesis with ionic liquids [Bmim][Br] as catalyst and ZnBr₂ as promoter. It can be observed that the selectivity of BO reduced from 97.4% to 80.2% with increasing the mass ratio

from 0 to 2. Therefore, the catalytic system without promoter is more superior choice in the selective synthesis of BO.

The effect of catalyst amount on the reaction was investigated, and the result is shown in Fig. 7. It is seen that the reaction could not occur without any catalyst. However, with the increase of ionic liquids [Bmim][Br] from 0.83 g to 2.00 g, the selectivity of BO slightly increased from 97.4% to 98.2%. After that, the selectivity of BO reduced with the further increase of catalyst amount.

However, this result obtained is from the commercial BC rather than BC from the carbonylation of butanediol with DMC. Thus, in order to realize the two-step process, the synthesis of BO was furtherly carried out by using BC generated from butanediol. In this process, BC was firstly synthesized by the transesterification reaction of butanediol with DMC over NaAlO₂ catalyst under the optimized reaction conditions. After the reaction, BC was separated from the product mixture and purified by vacuum distillation. The obtained BC was further used for the synthesis of BO over [Bmim][Br] catalyst. As a result, 94.0% of separated BO yield could be obtained. In this sense, a novel approach for BO synthesis is realized by a two-step process of carbonylation and decarboxylation method for the first time.

The decarboxylation of butenyl carbonate was undertaken in the presence of several [Bmim]-based ionic liquids with different anions (Cl⁻, I⁻, BF₄⁻), and the effect of different IL on formation rate of the decarboxylation of butenyl carbonate was shown in Fig. S6.† It can be observed that IL with halide anions [Bmim][Cl], [Bmim][Br] and [Bmim][I] can catalyze the decarboxylation of BC. The FR_{BO} values are 8.7, 3.9, 1.2, respectively, which are in the sequence of [Bmim][Cl] > [Bmim][Br] > [Bmim][I]. While the [Bmim][BF₄] can not catalyze the decarboxylation of BC. The selectivity of BO is 96.4% when the catalyst is [Bmim][Cl], while the higher selectivity of 98.2% can be obtained when

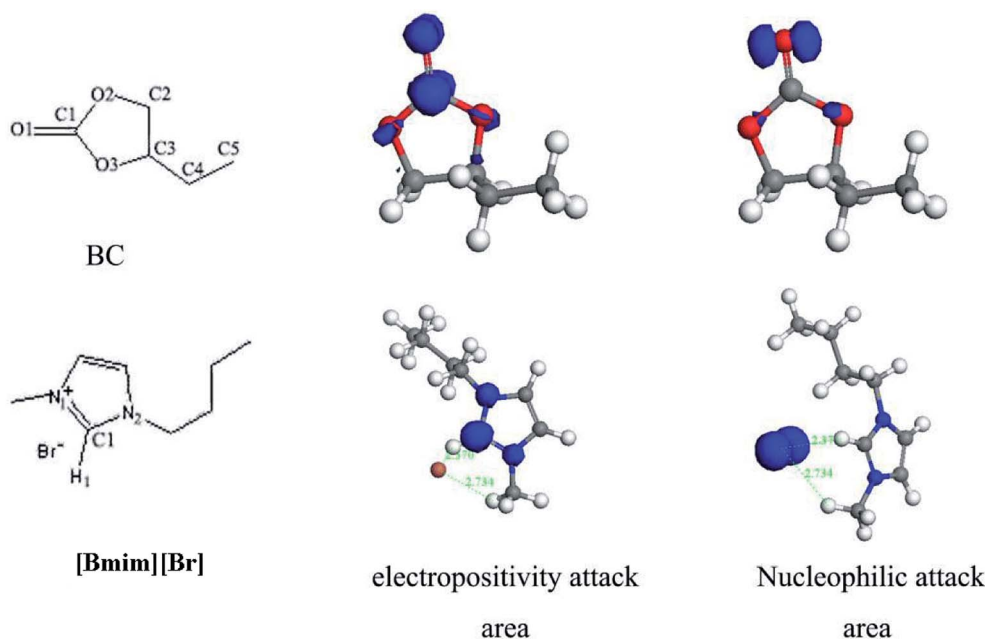


Fig. 8 Optimized structure of the BC and the [Bmim][Br].



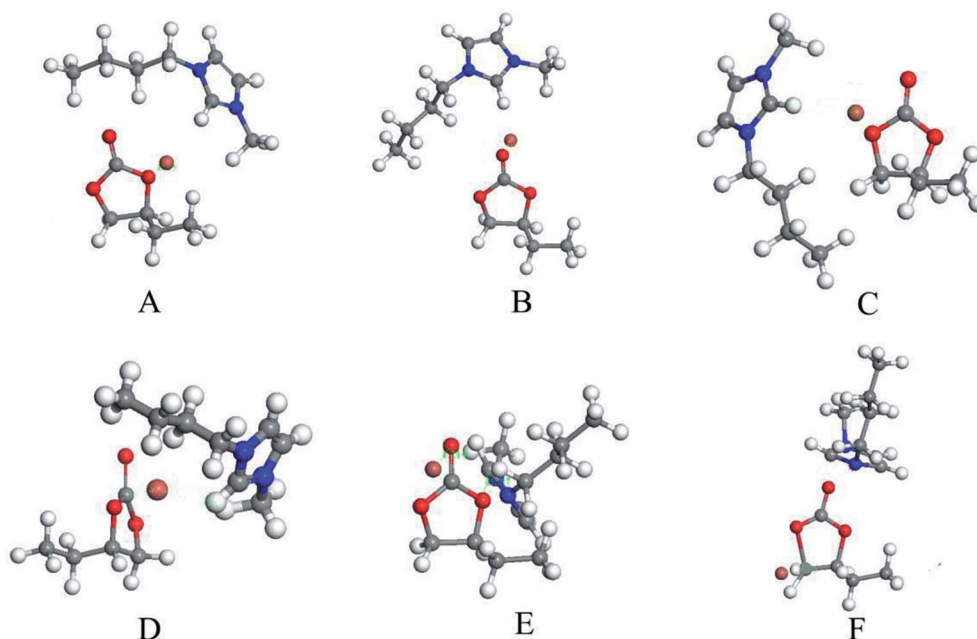


Fig. 9 Six initial configurations in DFT calculation.

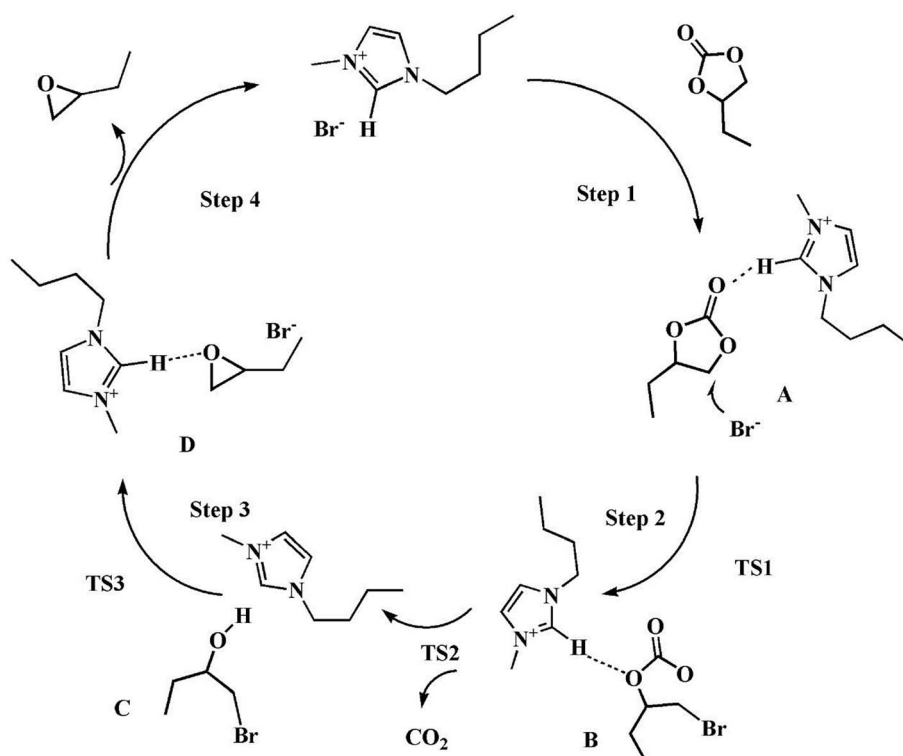
the catalyst is [Bmim][Br]. Thus, it is clear that anion of ILs have great impact on the decarboxylation of BC.

Mechanism of the decarboxylation

As it is shown in Fig. 7, the involvement of [Bmim][Br] catalyst resulted in the BO product, while the yield of BO is negligible

when there is no introduction of any catalyst. Thus, the DFT calculations were performed to reveal the interaction between [Bmim][Br] and BC.

As showed in Fig. 8, it was found that the carbonyl oxygen O1 of the BC was the main nucleophilic attack area and the carbon C1 of the BC were the main electropositivity attack area. The nitrogen atoms of the [Bmim][Br] were the main nucleophilic



Scheme 2 The possible decarboxylation reaction mechanism.



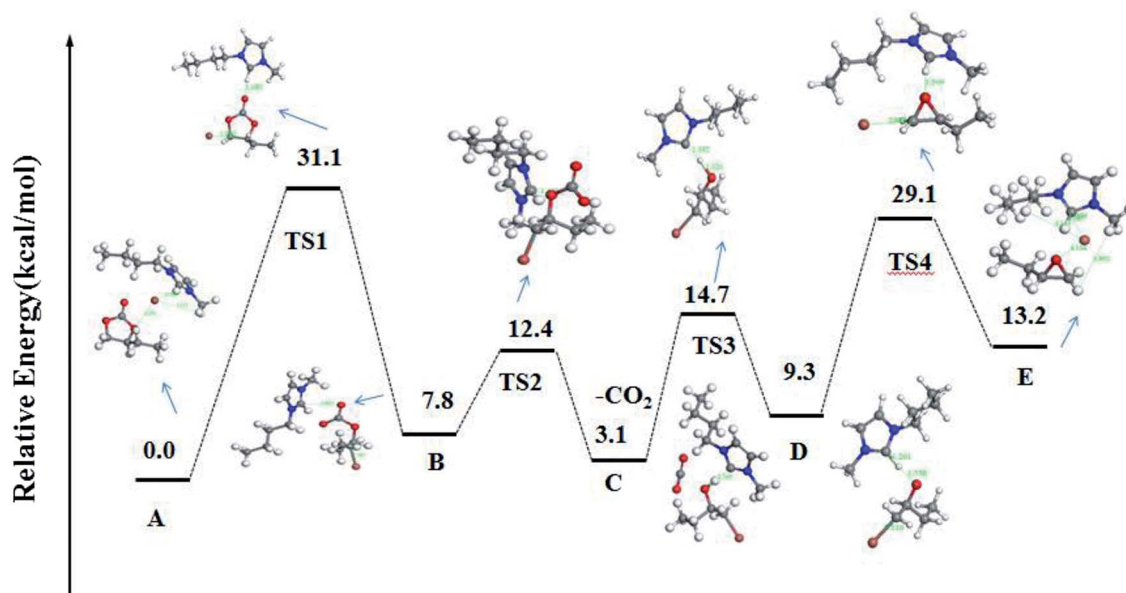


Fig. 10 Potential energy surface profiles of the [Bmim][Br]-catalyzed process and the optimized geometries for the intermediates and transition states.

attack area. The Br^- anions, which formed the hydrogen bonds of $\text{C-H}\cdots\text{Br}$, was the main electropositivity attack area of [Bmim][Br]. Accordingly, six initial configurations between BC and [Bmim][Br] was assumed and shown in Fig. 9.

According to the reports and DFT calculation results, a possible mechanism for decarboxylation was illustrated in Scheme 2.^{37,38} The optimized geometries for the intermediates and transition states were shown in Fig. S7†. The vibrational mode analysis shows that the reaction mechanism is reliable. When using vibrational mode analysis to elucidate the mechanism, we find that the only imaginary frequency of a transition state is always related to the breaking or forming bond. The potential energy surface profiles of the [Bmim][Br]-catalyzed process was shown in Fig. 10.

As shown in Scheme 2 and Fig. 10, first, the five-ring-opening step was the rate-limiting step with the highest energy barrier of $31.1 \text{ kcal mol}^{-1}$. It was observed that the nucleophilic attack of Br anions on the least-hindered carbon C2 of BC played a very important role to break the covalent bond between the C2 and O2. Under the interaction with Br anions, BC was activated with the C2–O2 bond length enlarged from 1.45 Å and 2.74 Å. However, a small probability was observed that the C3 of BC was nucleophilically attacked by Br anions because of the steric effect. Besides, the hydrogen bond between H1 and the carbonyl oxygen (O1) made the ring-opening much easier. Thus, the five-ring-opening step made A convert into intermediate B *via* transition state TS1. Then, a new complex C was formed when CO_2 was released into the reaction system *via* the TS2 with an energy barrier of $4.6 \text{ kcal mol}^{-1}$. The third step is the proton transfer from $\text{O3}\cdots\text{H}$ to the nitrogen atom of the imidazolium ring along with the simultaneous loss of Br anion *via* transition state TS3 with an energy barrier of $11.4 \text{ kcal mol}^{-1}$. Finally, the epoxy ring was obtained *via* an intramolecular nucleophilic

attack, and the catalyst was regenerated (TS4) with an energy barrier of $29.1 \text{ kcal mol}^{-1}$ (TS3). Besides, there was some negative effect of the hydrogen bonding between the $\text{C}\cdots\text{H}$ on the ring-closing event.

Conclusions

In conclusion, this paper firstly reported a novel approach for the synthesis of BO *via* decarboxylation of BC derived from butanediol. By using a cheap and easily available raw material NaAlO_2 as catalyst, 99.7% BC selectivity and 96.2% BC yield can be obtained from the transesterification of butanediol with DMC, and the catalyst can be easily recovered and reused. Furthermore, BC can be effectively decarboxylated to BO when using [Bmim][Br] as catalyst. Finally, BO was obtained by a two-step carbonylation and decarboxylation method for the first time by using butanediol as raw material. This new process is very significant for the production of epoxy compounds and the utilization of butanediol by-product inevitably generated in the process of coal to ethylene glycol.

Conflicts of interest

There are no conflicts of interest to declare.

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