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Rational design of quinoxaline-based bipolar host materials for highly efficient red phosphorescent organic light-emitting diodes†

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Two novel bipolar carbazole/diphenylquinoxaline-based host materials 3-(2,3-diphenylquinoxalin-6-yl)-9-phenyl-9*H*-carbazole (**M1**) and 3-(4-(2,3-diphenylquinoxalin-6-yl)phenyl)-9-phenyl-9*H*-carbazole (**M2**) have been rationally designed and synthesized. The phenyl spacer between the functionalized quinoxaline moiety and the carbazole moiety is also introduced to investigate its influence on their photophysical properties. The chemical structures, and thermal, photophysical and electrochemical properties of the two host materials were characterized and explored in detail. Red phosphorescent light-emitting diodes with **M1** and **M2** as hosts were prepared to explore their electroluminescent properties. Both **M1** and **M2** host-based red devices exhibit outstanding electroluminescent performance. For example, two red devices all realize good red emission with the maximum at 594 nm, the maximum external quantum efficiency and luminance can reach 14.66% and 28 619 cd m⁻² for **M1**-based devices and 15.07% and 28 818 cd m⁻² for **M2**-based devices, indicating compounds **M1** and **M2** designed in this work have potential applications in the development of high-performance monochrome and white OLEDs.

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Introduction

In recent years, owing to their fascinating commercial potential in the fields of flexible displays and solid state lighting, organic light-emitting diodes (OLEDs) have attracted great attention.^{1–7} Organic electroluminescent materials have been developed and have been extensively studied.^{8–17} Later Forrest and co-authors found that transition–metal complexes are a new class of highly efficient phosphorescent materials.^{18–20} Phosphorescent OLEDs, with phosphors as emitters, can simultaneously capture singlet excitons and triplet excitons excited by an electric field. Their internal quantum efficiency can exceed the theoretical limit of 25% in OLEDs based on fluorescent materials, making it possible to realize the unit internal quantum efficiency (IQE). So, phosphors and phosphorescent OLEDs are one of the main development directions in the future of organic electroluminescence.^{21–32}

The red emission component is an integral part of both high-quality solid state lightings and displays.^{33–39} The performance of red phosphorescent OLEDs is not determined by phosphors themselves, but also closely related to the host materials. Bipolar host materials that could balance the carrier-transportations for improving the efficiency of phosphorescent OLEDs have become the current research focus. Many bipolar host materials have been explored in highly efficient red phosphorescent OLEDs, especially in the design of acceptor moieties.^{40–42} Recently, diphenylquinoxaline derivatives showing excellent electron-transportation property also have been selected as the novel acceptor moiety to realize the bipolar host materials.^{33,34} For instance, Cheng *et al.* designed an excellent bipolar host material which is composed of two 6*H*-indolo[2,3-*b*]quinoxaline unit bridged by the tetraphenylsilane core realizing high-efficiency red phosphorescent OLEDs.³³ It is still necessary to further explore the quinoxaline-based bipolar host materials.

Herein, we designed two novel host materials (**M1** and **M2**) by combining the quinoxaline moiety and carbazole moiety, with various linkage ways between carbazole and quinoxaline moiety. Red phosphorescent OLEDs with **M1** and **M2** as hosts were successfully fabricated, and both **M1** and **M2** hosts-based red devices exhibit the outstanding electroluminescent performance. For example, the maximum external quantum efficiency (EQE) and luminance reach 14.66% and 28 619 cd m⁻² for **M1**-

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based red device, 15.07% and 28 818 cd m⁻² for **M2**-based red device.

Experimental

General experimental information

The detailed information on the solvents, chemicals, and general experiments can be found in the ESI.†

Synthesis of bipolar host materials

The synthesis of 3-(2,3-diphenylquinoxalin-6-yl)-9-phenyl-9H-carbazole (M1). 3-Bromo-9-phenyl-9H-carbazole (1.5 g, 4.7 mmol), 2,3-diphenyl-6-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)quinoxaline (1.9 g, 4.7 mmol), K₂CO₃ (1.4 g, 10 mmol), tetrakis(triphenylphosphine)palladium (50 mg) and H₂O (15 mL) were dissolved in THF (50 mL). The mixture was allowed to reflux under N₂ for 18 h. After the reaction completed, the mixture was cooled, then extracted with methylene chloride, dried over anhydrous Na₂SO₄ and concentrated. The crude product was purified by column chromatography (petroleum ether: dichloromethane = 2 : 1) to obtain a solid powder (yield 81%). ¹H NMR (400 MHz, CDCl₃) δ = 8.60 (d, *J* = 1.60 Hz, 1H), 8.53 (d, *J* = 1.60 Hz, 1H), 8.29–8.22 (m, 3H), 7.88 (dd, *J* = 8.40, 2.00 Hz, 1H), 7.67–7.61 (m, 4H), 7.58–7.49 (m, 6H), 7.47–7.44 (m, 2H), 7.39–7.33 (m, 7H). ¹³C NMR (101 MHz, CDCl₃) δ = 153.75, 152.83, 143.62, 141.51, 141.50, 140.93, 140.27, 139.01, 138.98, 137.48, 131.68, 130.18, 130.02, 129.92, 129.90, 129.35, 128.91, 128.85, 128.33, 128.32, 127.73, 127.12, 126.42, 125.92, 125.60, 124.17, 123.40, 120.41, 119.36, 110.44, 110.08. HRMS (ESI, *m/z*): [M + H]⁺ calcd for C₃₈H₂₆N₃, 524.2127, found 524.2133. Anal. calcd for C₃₈H₂₅N₃: C 87.16, H 4.81, N 8.02; found: C 87.12, H 4.90, N 8.09.

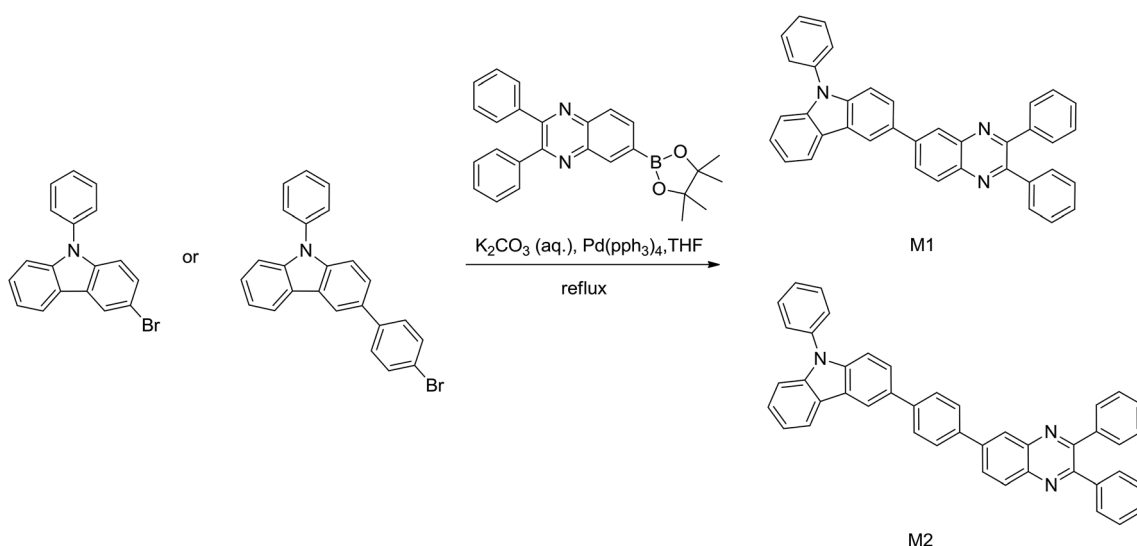
The synthesis of 3-(4-(2,3-diphenylquinoxalin-6-yl)phenyl)-9-phenyl-9H-carbazole (M2). Compound **M2** was synthesized by the same procedure for **M1** using 3-(4-bromophenyl)-9-phenyl-9H-carbazole as starting material. Yield: 76%. ¹H NMR (400

MHz, CDCl₃) δ = 8.49 (d, *J* = 1.60 Hz, 1H), 8.45 (d, *J* = 1.60 Hz, 1H), 8.28 (d, *J* = 8.80 Hz, 1H), 8.23 (d, *J* = 7.60 Hz, 1H), 8.15 (dd, *J* = 8.80, 2.00 Hz, 1H), 7.95–7.89 (m, 4H), 7.75 (dd, *J* = 8.40, 2.00 Hz, 1H), 7.67–7.60 (m, 4H), 7.58–7.54 (m, 4H), 7.52–7.48 (m, 2H), 7.45 (d, *J* = 3.60 Hz, 2H), 7.40–7.32 (m, 7H). HRMS (ESI, *m/z*): [M + H]⁺ calcd for C₄₄H₃₃N₃, 600.2440, found 600.2439. Anal. calcd for C₄₄H₂₉N₃: C 88.12, H 4.87, N 7.01; found: C 88.19, H 4.92, N 7.05.

OLED fabrication and measurement

All OLEDs were deposited on the indium-tin-oxide (ITO)/glass substrates with a sheet resistance of 15 Ω per square by thermal evaporation technique. ITO/glass substrates were cleaned sequentially by detergent, de-ionized water, acetone and isopropanol. The cleaned ITO/glass substrates were dried by nitrogen (N₂) flow and treated by ultraviolet. Then ITO/glass substrates were loaded into the evaporator with 13-evaporation sources. When the pressure in evaporator is below 5.0 × 10⁻⁴ Pa, the evaporation sources are heated for device fabrication. The deposition rate and film thickness were monitored controlled by the calibrated crystal quartz sensors. And the deposition rates for organic materials, MoO₃, LiF, and Al were controlled at about 1 Å s⁻¹, 0.3 Å s⁻¹, 0.1 Å s⁻¹, and 3–6 Å s⁻¹, respectively. The overlap between the Al and ITO is 3.0 × 3.0 mm², defined as an active area.

The electroluminescence (EL) spectra and Commission Internationale de L'Eclairage (CIE) coordinates for OLEDs were recorded by a computer controlled PR-655 spectra scan spectrometer. The current–voltage–luminance (*J*–*V*–*L*) characteristics, current efficiency, and power efficiency were measured by a computer-controlled Keithley 2400 source integrated with a BM-70A luminance meter. The EQE was calculated from the EL spectra and *J*–*V*–*L* data. All samples were tested immediately after thin films deposition without encapsulation at room temperature in ambient atmosphere.



Scheme 1 The synthetic routes of bipolar host materials **M1** and **M2**.



Result and discussion

Synthesis and characterization

The detailed synthetic routes and chemical structures of bipolar host materials are shown in Scheme 1. The synthesis of **M1** and **M2** was realized through Suzuki coupling reaction by reacting 3-bromo-9-phenyl-9*H*-carbazole or 3-(4-bromophenyl)-9-phenyl-9*H*-carbazole with 2,3-diphenyl-6-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)quinoxaline, respectively in high yields (over 75%). Both two compounds were confirmed by high-resolution mass spectra and nuclear magnetic resonance spectra (shown in ESI†). Under a nitrogen atmosphere, using the TGA and DSC to characterize the thermal properties of compounds **M1** and **M2**. As shown in Fig. 1 and Table 1, the onset decomposition temperature (T_d , corresponding to 5% weight loss) according to the TGA data is 321 °C for **M1** and is 415 °C for **M2**. We obviously found that the endothermic glass transition temperature (T_g) for **M1** and **M2** is determined to be 176 °C and 218 °C by analyzing the DSC traces, indicating **M1** and **M2** compounds are suitable for developing OLEDs by thermal evaporation technology. Owing to its much more steric structure of the molecule, the host **M2** shows obviously both high T_g and T_d values. The relatively high thermal stabilities of the host materials are benefit for improving the morphology of the film and lifetime of the device by suppressing the phase separation and crystallization during the operation process of device.

Photophysical properties

The absorption and emission properties of **M1** and **M2** were investigated in toluene. As depicted in Fig. 2a, the absorption of **M1** and **M2** exhibits bands in the range from 250 to 430 nm (247, 298, 387 nm for **M1**; 254, 303, 378 nm for **M2**), attributed to the π - π^* electronic transition of the quinoxaline and carbazole unit.^{33,34,40–42} **M1** and **M2** show almost the same absorption spectra in spite of the larger π -electron conjugation of compound **M2**, implying that the introduction of the phenyl spacer almost has little effect on the absorption properties of these donor-acceptor systems.

However, quite different from their absorption properties, two compounds exhibited the different maximum emission wavelength ($\lambda_{\text{max}}^{\text{em}} = 493$ nm for **M1**; $\lambda_{\text{max}}^{\text{em}} = 518$ nm for **M2**) in toluene at room temperature, as depicted in Fig. 2a and Table 1. The slight red shifted in the emission of **M2** could be attributed to the larger π -electron conjugation. The compounds **M1** and **M2** exhibited sky-blue or green fluorescence with energy gap of 2.51 and 2.39 eV in toluene, implying that the π -electron conjugation and co-planarity of molecule have been tuned by the phenyl spacer between functionalized quinoxaline moiety and carbazole moiety. In addition, the low temperature phosphorescence spectra of the compound **M1** and **M2** were also recorded at 77 K to calculate the triplet energies. As shown in the Fig. 2b, the low temperature phosphorescence spectra of both compounds were almost the same. The maximum emission of the phosphorescence spectra was at the wavelength of 553 nm. The triplet energy levels were calculated by the first emission peak of the low-temperature phosphorescence spectra

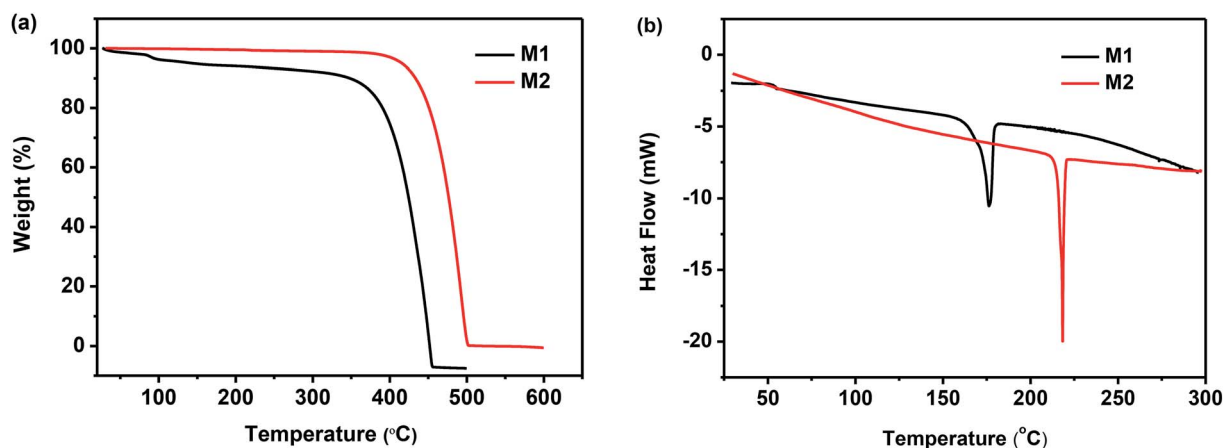


Fig. 1 The thermogravimetric (a) and differential scanning calorimetry (b) analysis of bipolar host materials **M1** and **M2**.

Table 1 Photophysical properties of the bipolar host materials **M1** and **M2**

Compound	$\lambda_{\text{max}}^{\text{abs}}/\lambda_{\text{onset}}^a$ [nm]	$\lambda_{\text{max}}^{\text{em}}^a$ [nm]	λ_{phos}^b [nm]	E_g^c [eV]	E_T^d [eV]	E_{onset}^e [V]	HOMO ^e [eV]	LUMO ^f [eV]	T_d^g [°C]	T_g^h [°C]
M1	247, 298, 387	493	553	2.51	2.24	1.20	-5.60	-3.09	321	176
M2	254, 303, 378	518	553	2.39	2.24	1.20	-5.60	-3.21	415	218

^a Measured in toluene at room temperature. ^b Measured in toluene at 77 K. ^c Calculated by the equation $E_g^{\text{opt}} = 1240/\lambda_{\text{max}}^{\text{em}}$. ^d Calculated by the first peak of phosphorescence spectra measured at 77 K, $E_T = 1240/\lambda_{\text{phos}}$. ^e Calculated by the equation $E_{\text{HOMO}} = -4.4 - E_{\text{onset}}^{\text{ox}}$. ^f Calculated by the equation $E_{\text{LUMO}} = E_{\text{HOMO}} + E_g$. ^g Measured by TGA. ^h Measured by DSC.



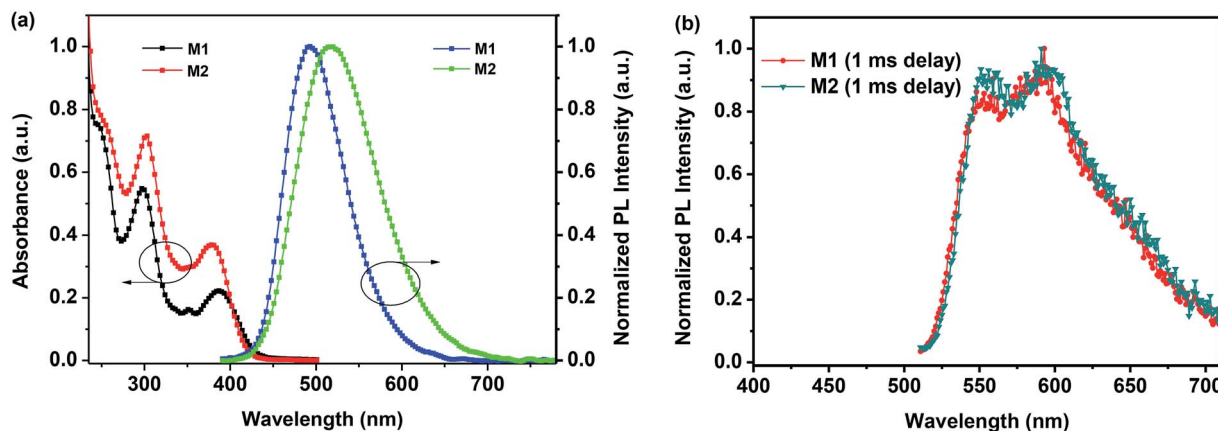


Fig. 2 (a) UV-Vis absorption and PL spectra of the compounds **M1** and **M2** in toluene at 297 K; (b) phosphorescence spectra of the compound **M1** and **M2** at 77 K.

and were determined to be 2.24 eV, which is suitable for the red OLEDs.²⁴ In addition, at the 77 K, we can observe the electronic transitions between the T_1 substates (substates III, II, I) and the ground state S_0 . From the intensity distribution, it can be concluded that the radiative transition from the lowest substate I to the ground state S_0 is largely forbidden, while the transitions from the two higher-lying substates III and II to the ground state are significantly more allowed.

Electrochemical properties

As depicted in Fig. 3, using a standard three-electrode electrochemical cell in an electrolyte solution, the electrochemical behavior of compounds **M1** and **M2** was investigated by cyclic voltammetry. The energy level of the highest occupied molecular orbital (HOMO) for **M1** and **M2** were calculated to be both -5.60 eV from the onset potentials of oxidation peak, while the energy level of the lowest unoccupied molecular orbital (LUMO) of these host materials were determined to be -3.09 and

-3.21 eV according to the band gap energy (E_g) and HOMO energy level. The data of E_g and HOMO/LUMO energy levels were listed in Table 1. The same HOMO levels of both compounds can be explained by sharing the same electron-donor carbazole group, which indicated that the phenyl spacer between the functionalized quinoxaline moiety and carbazole moiety has no effect on the HOMO levels. However, the LUMO energy of compound **M1** (-3.09 eV) were higher than that of **M2** (-3.21 eV), implying that the phenyl spacer have evident influence on the LUMO of these donor-acceptor systems.^{40–42}

Red phosphorescent OLEDs

Considering the triplet energy levels (2.24 eV) of **M1** and **M2** is suitable for red OLEDs, to order to prove their potential application in OLEDs, two doped device of ITO/40 nm-poly(3,4-ethylenedioxythiophene):poly(styrenesulfonate) (PEDOT:PSS) layer/45 nm-di-(4-(*N,N*-di-*p*-tolyl-amino)-phenyl)cyclohexane (TAPC) layer/5 nm-4,4',4''-tris(carbazol-9-yl)triphenylamine (TCTA) layer/20 nm-hosts: 5% wt bis(2-phenylquinolino)-acetylacetonate iridium(III) ($\text{Ir}(\text{pq})_2\text{acac}$) layer/55 nm-1,3,5-tri((3-pyridyl)-phen-3-yl)benzene (TmPyPB) layer/0.8 nm-lithium fluoride (LiF) layer/80 nm-aluminum (Al) layer were fabricated by vacuum thermal deposition, where hosts are **M1** and **M2**, corresponding to the device R1 and R2, respectively. The device structure and organic molecules used in device fabrication are displayed in Fig. 4. In two red devices, different function materials were chosen on the basis of knowing the energy level of materials, and the thickness of different layers was decided by combining the reported relevant literature and mobility of materials.^{4,6,25} In these devices, ITO was used as an anode; the LiF/Al layers were employed as a composite cathode; 40 nm-thick PEDOT:PSS layer was used as the hole-injection layer; 45 nm-thick TAPC layer was used as the hole transporting layer; 5 nm-thick TCTA layer was used as the exciton blocking layer and 55 nm-thick TmPyPB layer as the electron transport layer. From the energy level diagram of two red devices in Fig. S1 in ESI,† we can see that TCTA and TmPyPB can

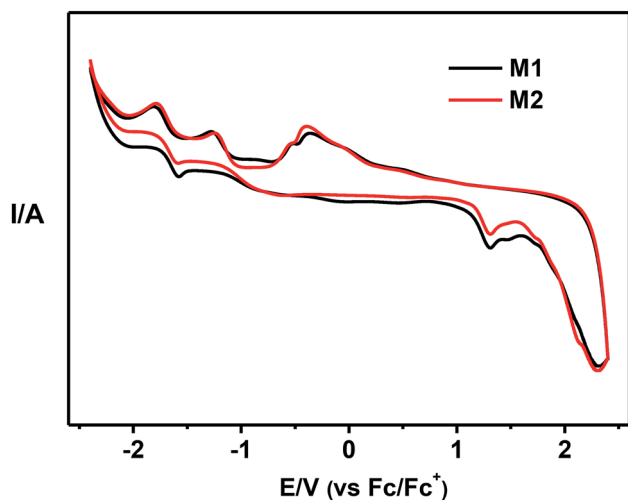


Fig. 3 The cyclic voltammograms of compounds **M1** and **M2**.



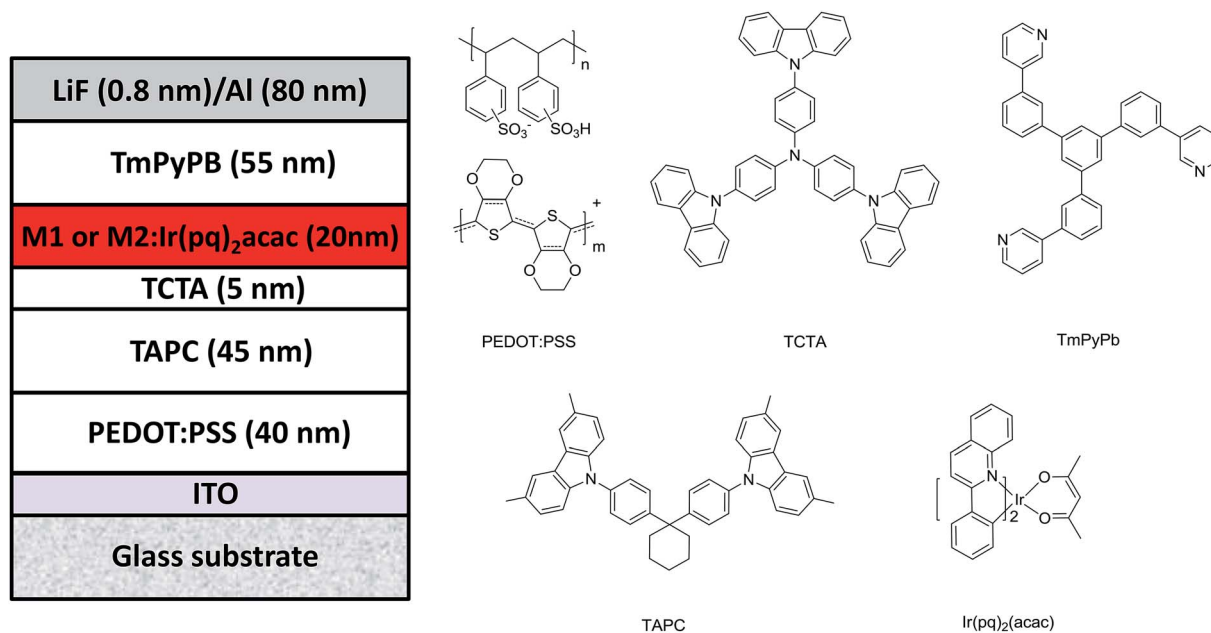


Fig. 4 The device structure of red devices and organic molecules used in device fabrication.

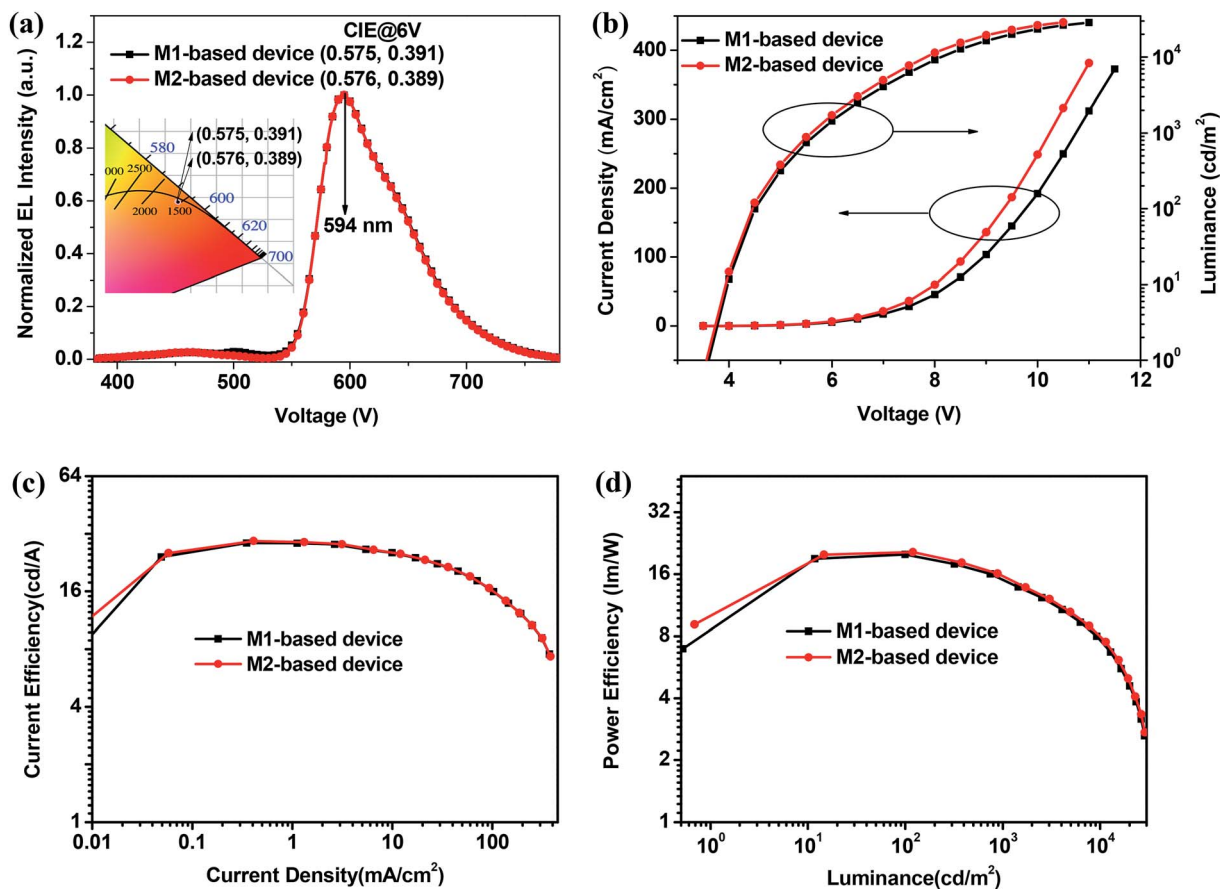


Fig. 5 Normalized EL spectra and CIE coordinates at 6 V (a), J - V - L (b), CE- J (c), and PE- L (d) characteristic curves for the M1- and M2-based red devices.



Table 2 The EL performance parameter summary for M1- and M2-based devices

Device	Host	Voltage ^a (V)	Maximum				CIE ^c [x, y]
			CE [cd A ⁻¹]	PE [lm W ⁻¹]	EQE [%]	L ^b (cd m ⁻²)	
R1	M1	3.6	28.57	19.95	14.66	28 619	(0.575, 0.391)
R2	M2	3.6	29.29	20.45	15.07	28 818	(0.576, 0.389)

^a Voltage at a luminance of >1 cd m⁻². ^b L is the abbreviation of luminance. ^c Measured from the EL spectra at 6 V.

effectively limit carrier recombination zone and generated exciton in EML because of the big LUMO energy barriers of 0.69 eV and 0.81 eV at TCTA/M1 and M2 interface, a big HOMO energy barrier of 1.20 eV at M1 or M2/TmPyPB interface, and a high triplet levels for TCTA (2.68 eV) and TmPyPB (2.70 eV) than both M1 (2.24 eV) and M2 (2.24 eV).²⁷ Thus, theoretically, both devices R1 and R2 can well reveal the EL performance of compounds M1 and M2.

Fig. 5a depicts the normalized EL spectra of M1- and M2-based devices at a voltage of 6 V. Clearly, both M1- and M2-based devices emit red light, and the EL spectra for two devices are almost completely overlapping with the same peak at 594 nm, and CIE coordinates are (0.575, 0.391) and (0.576, 0.389) for M1- and M2-based devices, respectively. These are well consistent to the emission characteristic of emitter Ir(pq)₂acac. Meanwhile, from Fig. S2 in ESI,[†] it is found that there is almost no EL color shift vary at a wide voltage range of 5–8 V for two red devices, which indicates (i) the carrier recombination zone and exciton are well confined in the red EML; and the energy in M1 and M2 hosts can be effectively transferred to Ir(pq)₂acac when Ir(pq)₂acac is doped in M1 and M2 hosts for two red device. The above result preliminarily proved the feasibility of compounds M1 and M2 as hosts for red OLEDs.

In fact, the high EL performance for both M1- and M2-based devices further demonstrate that both M1 and M2 are excellent host materials for red phosphorescent OLEDs. For example, as shown in Fig. 5b, both M1- and M2-based devices show a relatively low turn-on voltage of 3.6 V, and all achieve very high maximum luminance of 28 619 cd m⁻² for device R1 and 28 818 cd m⁻² for device R2. From Fig. 5c, d and S3 in ESI,[†] both M1- and M2-based devices achieve high device efficiency, and the maximum current efficiency (CE), power efficiency (PE), and EQE reach 28.57 cd A⁻¹, 19.95 lm W⁻¹, and 14.66% for M1-based red device and 29.29 cd A⁻¹, 20.45 lm W⁻¹, and 15.07% for M2-based red device. All above results indicate that these diphenylquinoxaline-based bipolar host materials have potential applications for the development of high-performance red OLEDs. In addition, the performance of red OLEDs in this work are much higher than that in the reported results (PE = 20.45 lm W⁻¹, CE = 29.29 cd A⁻¹, EQE = 15.07% vs. PE = 15.4 lm W⁻¹, CE = 21.9 cd A⁻¹, EQE = 12.2%) showing in Fig. S4,[†] indicating that the connection ways of electron acceptor/donor moieties have great influence on performance of the host materials (Table 2).³⁴

Conclusion

To conclude, two novel quinoxaline-based bipolar host materials have been rational design and prepared successfully *via* the Suzuki coupling reaction in high yields for application in red phosphorescent OLEDs. By incorporating the diphenylquinoxaline and carbazole unit as the electron donor and acceptor, meanwhile, the phenyl spacer between the functionalized quinoxaline moiety and carbazole moiety is also introduced to investigate its influence on their photophysical properties. Their thermal, photophysical and electrochemical properties were investigated and discussed in details. The red phosphorescent OLEDs with M1 and M2 as host were fabricated, and it is found both M1 and M2 hosts-based red devices exhibit the outstanding electroluminescent performance. For example, two red devices all realize good red emission with peak at 594 nm, and the maximum luminance and external quantum efficiency reach 28 619 cd m⁻² and 14.66% for M1-based device and 28 818 cd m⁻² and 15.07% for M2-based device, indicating that compounds M1 and M2 synthesized in this work have the potential applications in the development of high performance red and white OLEDs.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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References

- 1 C. W. Tang and S. A. VanSlyke, *Appl. Phys. Lett.*, 1987, **51**, 913–915.
- 2 S. Reineke, F. Lindner, G. Schwartz, N. Seidler, K. Walzer, B. Lussem and K. Leo, *Nature*, 2009, **459**, 234–238.
- 3 Y. Chang, Y. Song, Z. Wang, M. Helander, J. Qiu, L. Chai, Z. Liu, G. Scholes and Z. Lu, *Adv. Funct. Mater.*, 2013, **23**, 705–712.



- 4 Y. Miao, K. Wang, L. Gao, H. Wang, F. Zhu and B. Xu, *J. Mater. Chem. C*, 2018, **6**, 9811–9820.
- 5 H. Sasabe and J. Kido, *J. Mater. Chem. C*, 2013, **1**, 1699–1707.
- 6 P. Tao, Y. Miao, H. Wang, B. Xu and Q. Zhao, *Chem. Rec.*, 2018, **18**, 1–32.
- 7 V. Sree, A. Maheshwaran, H. Kim, H. Park, Y. Kim, J. Lee, M. Song and S. Jin, *Adv. Funct. Mater.*, 2018, **28**, 1804714.
- 8 X. Tang, X. Liu, Z. Jiang and L. Liao, *Adv. Funct. Mater.*, 2019, **29**, 1807541.
- 9 T. H. Han, Y. Lee, M. R. Choi, S. H. Woo, S. H. Bae, B. H. Hong, J. H. Ahn and T. W. Lee, *Nat. Photonics*, 2012, **6**, 105–110.
- 10 H. Uoyama, K. Goushi, K. Shizu, H. Nomura and C. Adachi, *Nature*, 2012, **492**, 234–238.
- 11 Y. Miao, K. Wang, L. Gao, B. Zhao, Z. Wang, Y. Zhao, A. Zhang, H. Wang, Y. Hao and B. J. Xu, *J. Mater. Chem. C*, 2018, **6**, 1853–1862.
- 12 C. L. Ho and W. Y. Wong, *Coord. Chem. Rev.*, 2013, **257**, 1614–1649.
- 13 L. Ying, C. L. Ho, H. B. Wu, Y. Cao and W. Y. Wong, *Adv. Mater.*, 2014, **26**, 2459–2473.
- 14 X. L. Yang, G. J. Zhou and W. Y. Wong, *J. Mater. Chem. C*, 2014, **2**, 1760–1778.
- 15 P. Tao, Y. Zhang, J. Wang, L. Wei, H. Li, X. Li, Q. Zhao, X. Zhang, S. Liu, H. Wang and W. Huang, *J. Mater. Chem. C*, 2017, **5**, 9306–9314.
- 16 R. Ding, J. Feng, F. X. Dong, W. Zhou, Y. Liu, X. L. Zhang, X. P. Wang, H. H. Fang, B. Xu, X. B. Li, H. Y. Wang, S. Hotta and H. B. Sun, *Adv. Funct. Mater.*, 2017, **27**, 1604659.
- 17 J. Lin, Y. Wang, P. Gnanasekaran, Y. Chiang, C. Yang, C. Chang, S. Liu, G. Lee, P. Chou, Y. Chi and S. Liu, *Adv. Funct. Mater.*, 2017, **27**, 1702856.
- 18 M. A. Baldo, D. F. O'Brien, Y. You, A. Shoustikov, S. Sibley, M. E. Thompson and S. R. Forrest, *Nature*, 1998, **395**, 151–154.
- 19 Y. R. Sun, N. C. Giebink, H. Kanno, B. W. Ma, M. E. Thompson and S. R. Forrest, *Nature*, 2006, **440**, 908–912.
- 20 J. Lee, H. F. Chen, T. Batagoda, C. Coburn, P. I. Djurovich, M. E. Thompson and S. R. Forrest, *Nat. Mater.*, 2016, **15**, 92–98.
- 21 W. Y. Wong and L. J. Ho, *J. Mater. Chem.*, 2009, **19**, 4457–4482.
- 22 W. Y. Wong and C. L. Ho, *Coord. Chem. Rev.*, 2009, **253**, 1709–1758.
- 23 P. Tao, Q. Zhao, S. Jing, J. Wang, Z. Lv, L. Chen and H. Wang, *Chin. J. Lumin.*, 2013, **34**, 816–823.
- 24 P. Tao, W. Li, J. Zhang, S. Guo, Q. Zhao, H. Wang, B. Wei, S. Liu, X. Zhou, Q. Yu, B. Xu and W. Huang, *Adv. Funct. Mater.*, 2016, **26**, 881–894.
- 25 Y. Miao, P. Tao, L. Gao, X. Li, L. Wei, S. Liu, H. Wang, B. Xu and Q. J. Zhao, *J. Mater. Chem. C*, 2018, **6**, 6656–6665.
- 26 C. Fan and C. L. Yang, *Chem. Soc. Rev.*, 2014, **43**, 6439–6469.
- 27 Y. Miao, K. Wang, L. Gao, B. Zhao, H. Wang and F. J. Zhu, *J. Mater. Chem. C*, 2018, **6**, 8122–8134.
- 28 J. J. Guo, X. L. Li, H. Nie, W. W. Luo, S. F. Gan, S. M. Hu, R. R. Hu, A. J. Qin, Z. J. Zhao, S. J. Su and B. Z. Tang, *Adv. Funct. Mater.*, 2017, **27**, 1606458.
- 29 L. Gao, P. Tao, Y. Miao, W. Jia, Y. Zhao, H. Wang and B. Xu, *Tetrahedron Lett.*, 2018, **59**, 2095–2098.
- 30 Y. Miao, K. Wang, B. Zhao, L. Gao, Y. Wang, H. Wang, B. Xu and F. J. Zhu, *J. Mater. Chem. C*, 2017, **5**, 12474–12482.
- 31 X. Cai and S. Su, *Adv. Funct. Mater.*, 2018, **28**, 1802558.
- 32 H. Li, P. Tao, Y. Xu, X. Zhang, S. Liu and Q. Zhao, *Tetrahedron Lett.*, 2018, **59**, 1748–1751.
- 33 C. H. Fan, P. P. Sun, T. H. Su and C. H. Cheng, *Adv. Mater.*, 2011, **23**, 2981–2985.
- 34 M. Hu, Q. Xu, Y. Jiang, H. Mu, L. Gao, P. Hu, J. Huang and J. Su, *Dyes Pigm.*, 2018, **150**, 185–192.
- 35 Y. Miao, P. Tao, K. Wang, H. Li, B. Zhao, L. Gao, H. Wang, B. Xu and Q. Zhao, *ACS Appl. Mater. Interfaces*, 2017, **9**, 37873–37882.
- 36 P. Tao, Y. Miao, Y. Zhang, K. Wang, H. Li, L. Li, X. Li, T. Yang, Q. Zhao, H. Wang, S. Liu, X. Zhou, B. Xu and W. Huang, *Org. Electron.*, 2017, **45**, 293–301.
- 37 P. Tao, Y. Miao, K. Wang, H. Li, Q. Zhao, H. Wang, J. Li, B. Xu and W. Huang, *Tetrahedron Lett.*, 2017, **58**, 3598–3601.
- 38 Z. Feng, P. Tao, L. Zou, P. Gao, Y. Liu, X. Liu, H. Wang, S. Liu, Q. Dong, J. Li, B. Xu, W. Huang, W. Wong and Q. Zhao, *ACS Appl. Mater. Interfaces*, 2017, **9**, 28319–28330.
- 39 Y. Miao, K. Wang, B. Zhao, L. Gao, P. Tao, X. Liu, Y. Hao, H. Wang, B. Xu and F. Zhu, *Nanophotonics*, 2018, **7**, 295–304.
- 40 W. Li, J. Y. Li, D. Liu and Q. Jin, *ACS Appl. Mater. Interfaces*, 2016, **8**, 22382–22391.
- 41 H. Xu, P. Sun, K. Wang, J. Li, F. Wang, Y. Miao, H. Wang, B. Xu and W. Wong, *J. Mater. Chem. C*, 2017, **5**, 4455–4462.
- 42 K. Wang, P. Sun, H. Xu, J. Li, F. Wang, Y. Miao, T. Yang, H. Wang, B. Xu and W. Wong, *Dyes Pigm.*, 2017, **143**, 25–32.

