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# New chemical mechanism explaining the breakdown of protective oxides on high temperature steels in biomass combustion and gasification plants†‡

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Biomass is considered a replacement fuel over fossil fuels to mitigate climate change. The switch to biomass in the combustors changes the inorganic chemistry of the flue gases and leads to more severe corrosion of the construction materials of the combustors. The integrity of most high temperature steels relies on the formation of a protective  $\text{Cr}_2\text{O}_3$  layer on the steel surface at a high temperature environment. The ash compound found on the heavily corroded steel in biomass combustion and gasification plants is KCl, but the mechanism, which triggers the breakdown of the protective  $\text{Cr}_2\text{O}_3$  layer under the KCl salt is not known. We studied the chemical reactions involved with furnace exposure of KCl and KOH with  $\text{Cr}_2\text{O}_3$  and identified the formed reaction products with XRD analysis. The amount of reaction products was analyzed from the leachates of the salt-oxide mixtures by UV/VIS spectroscopy. We also used thermodynamic Gibbs energy minimization calculations to evaluate the evolution of reactions as a function of temperature. The results suggests that the reaction of KCl with  $\text{Cr}_2\text{O}_3$  involves a KOH reaction intermediate that forms before  $\text{K}_2\text{CrO}_4$  is formed. The amount of reacted potassium as a function of temperature follows the trend of KCl decomposition to KOH and  $\text{HCl(g)}$  as predicted by thermodynamics calculations. Therefore, we argue that the suggested overall reaction of KCl with  $\text{Cr}_2\text{O}_3$  as found in the corrosion literature:

$4\text{KCl} + \text{Cr}_2\text{O}_3 + 1\frac{1}{2}\text{O}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g}) \Rightarrow 2\text{K}_2\text{CrO}_4 + 4\text{HCl}(\text{g})$ , starting with the initiation step:  $\text{KCl} + \text{H}_2\text{O}(\text{g}) \Rightarrow \text{KOH} + \text{HCl}(\text{g})$  and then the formed KOH reacts with  $\text{Cr}_2\text{O}_3$  to form  $\text{K}_2\text{CrO}_4$ . This explains the initial breakdown of the protective  $\text{Cr}_2\text{O}_3$  under KCl salt in water containing high temperature atmospheres. The result is essential for the development of new alloys for biomass fired combustors.

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## 1. Introduction

Burning coal is responsible for approximately 25% of anthropogenic  $\text{CO}_2$  emissions. The largest use of coal as a fuel is in electricity and heat production, which accounts for about 16% of the total  $\text{CO}_2$  emissions.<sup>1</sup> Replacing coal with sustainable biomass as the fuel in power boilers would lead to substantial reductions in the net  $\text{CO}_2$  emission. Also,  $\text{SO}_2$  emissions would be diminished due to the generally much lower sulphur content of biomass based fuels compared to coal. However, biomass combustion leads to more severe fouling and corrosion issues of the heat transfer surfaces in boilers.<sup>2–5</sup> One way to minimize these

problems is to operate the boilers with lower steam temperatures, but that lowers the efficiency of the Rankine cycle leading to higher emission factors per unit of electricity produced.

Potassium behaviour in the boiler has been identified as the major cause of slagging, fouling and corrosion in biomass fired boilers.<sup>6</sup> Potassium is found associated in the fuel moisture as water soluble salts, or reacted with the organic functional groups (carboxylic, alkoxy, phenolic) of the lignocellulosic matrix.<sup>7</sup> During combustion, potassium is released in the gas phase as elemental potassium, potassium chloride or potassium hydroxide as measured by mass spectrometric or optical methods.<sup>8–10</sup> Due to its extreme reactivity, elemental potassium is expected to react to compounds in the vicinity of the biomass particle it is released from. Therefore, the major gas phase potassium compounds in the biomass boiler flue gases are  $\text{KCl(g)}$  and  $\text{KOH(g)}$ . These compounds may react further with other fuel elements, such as silicon or sulphur to form potassium silicates or potassium sulphate.  $\text{K}_2\text{O-SiO}_2$  silicate formation plays a major role in deposit formation in the furnace (slagging) and in bed agglomeration issues in fluidized bed

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‡ The raw/processed data required to reproduce these findings are partly included in the ESI section of the manuscript. Additional raw data is available from the authors upon request.



boilers, where silica containing sand is often used as the bed material.<sup>11–13</sup> Formation of  $\text{K}_2\text{SO}_4$  plays a major role in aerosol and deposit formation further downstream of the boiler.<sup>14,15</sup>  $\text{K}_2\text{SO}_4$  formation may decrease the corrosion rate of the heat exchangers forming a less corrosive deposit than  $\text{KCl}$  or  $\text{KOH}/\text{K}_2\text{CO}_3$ . Its formation may however, also increase fouling rate of the heat exchangers forming tenuous deposits that are difficult to be removed by soot blowing, even though the corrosion rate under the deposits may still be low.<sup>16,17</sup> During their path through the boiler,  $\text{KCl}(\text{g})$  and  $\text{KOH}(\text{g})$  in the flue gases condense out when the gas temperature is decreased below their dew points. Condensation can happen directly on the heat exchanger surfaces or on the ash particles present in the flue gas flow. Purely homogeneous nucleation is also possible in conditions where the flue gas does not contain enough foreign surfaces that can act as nucleation sites for heterogeneous nucleation.<sup>18,19</sup> After condensation  $\text{KCl}(\text{s,l})$  can react heterogeneously further to  $\text{K}_2\text{SO}_4(\text{s,l})$  and  $\text{KOH}(\text{s,l})$  can react to  $\text{K}_2\text{SO}_4(\text{s,l})$ ,  $\text{KCl}(\text{s,l})$  or  $\text{K}_2\text{CO}_3(\text{s,l})$  as predicted by thermodynamic stabilities of the compounds.<sup>20,21</sup>

$\text{KCl}$  induced corrosion of Fe–Cr steels has been studied extensively in the scientific literature.<sup>22–28</sup> Recently, the effect of  $\text{K}_2\text{CO}_3$  on the high temperature corrosion has gained more interest. It appears both potassium and chlorine are important in the corrosion reactions with steels. It has been suggested that potassium can initiate the destruction of the protective oxide, but chlorine is needed to sustain the corrosion.<sup>29,30</sup>  $\text{KOH}$  induced corrosion has not been extensively studied in the context of high temperature corrosion in biomass fired boilers, but earlier studies have shown that Fe–Cr alloys are unsuitable for service in  $\text{KOH}$  containing high temperature environments.<sup>31–33</sup> Earlier work of the author on the elemental balances of the deposit forming elements in biomass based fuels suggest that  $\text{KOH}(\text{g})$  condensation may be more important in fouling and corrosion than has previously been thought. The details of  $\text{K}_2\text{CO}_3(\text{s,l})$  formation on the heat exchanger surfaces has not been clarified yet, but its formation has been predicted by thermodynamics when the  $(\text{Cl}(\text{g}) + 2\text{S}(\text{g}))$  molar content in the flue gases is lower than the molar  $\text{K}(\text{g})$  content.  $\text{K}_2\text{CO}_3(\text{s,l})$  is formed on the heat exchanger surfaces likely *via* a surface reaction of adsorbed  $\text{KOH}(\text{ads.})$  with  $\text{CO}_2(\text{g})$ . Homogeneous formation of  $\text{K}_2\text{CO}_3(\text{g})$  in the gas phase, followed by condensation of  $\text{K}_2\text{CO}_3(\text{g})$  is less likely, because of the thermodynamic instability of  $\text{K}_2\text{CO}_3(\text{g})$ .<sup>34,35</sup>  $\text{K}_2\text{CO}_3(\text{s,l})$  has also been directly detected in some boiler deposits.<sup>17</sup> In this work we studied in detail the reactivity of  $\text{KCl}$  and  $\text{KOH}$  towards  $\text{Cr}_2\text{O}_3$  and  $\text{Fe}_2\text{O}_3$ , the protective oxide components formed on the Fe–Cr alloys in high temperature oxidizing service conditions. The results may also be of interest for chromite ore roasting by  $\text{KOH}$  and for understanding corrosion of  $\text{Cr}_2\text{O}_3$  containing refractory bricks in potassium and chlorine containing environments.<sup>36–38</sup>

## 2. Experimental

### 2.1 Preparation of the mixtures and furnace exposures

$\text{KOH}-\text{Cr}_2\text{O}_3$ ,  $\text{KCl}-\text{Cr}_2\text{O}_3$ ,  $\text{KOH}-\text{Fe}_2\text{O}_3$ ,  $\text{KCl}-\text{Fe}_2\text{O}_3$  and  $\text{K}_2\text{CrO}_4-\text{Cr}_2\text{O}_3$  mixtures were prepared by mixing known amounts of

powders in a 10 ml glass bottle and manually shaking the bottles for approximately 30 s.  $\text{KOH}$  was from Sigma Aldrich technical grade  $\geq 85\%$ ,  $\text{KCl}$  from Alfa Aesar, ACS grade 99–100.5%,  $\text{K}_2\text{CrO}_4$  from Merck, EMSURE®, ACS grade  $\geq 99.5\%$ ,  $\text{Cr}_2\text{O}_3$  from E. Merck, grade unknown and  $\text{Fe}_2\text{O}_3$  from Johnson Matthey Chemicals, Specpure® grade. All powders were weighted with Sartorius CPA225D analytical balance. Larger agglomerates of  $\text{KCl}$ ,  $\text{KOH}$  and  $\text{K}_2\text{CrO}_4$  were grinded manually in a mortar before mixing with the  $\text{Cr}_2\text{O}_3$  or  $\text{Fe}_2\text{O}_3$  powder.  $\text{Cr}_2\text{O}_3$  and  $\text{Fe}_2\text{O}_3$  powders seemed visually homogeneous without agglomerates and were used without any pre-grinding.  $\text{KCl}/\text{Cr}_2\text{O}_3$ ,  $\text{KCl}/\text{Fe}_2\text{O}_3$  or  $\text{KOH}/\text{Cr}_2\text{O}_3$ ,  $\text{KOH}/\text{Fe}_2\text{O}_3$  molar ratios were 1 and  $\text{K}_2\text{CrO}_4/\text{Cr}_2\text{O}_3$  molar ratio was 0.5 in order to have the same K/metal molar ratio in all of the mixtures. A 10 g batch of each mixture was prepared in one go in a screw cap sealed bottle. From the 10.0 g batch bottles, 1.00 g samples were weighted to a 10 ml sintered  $\text{Al}_2\text{O}_3$  crucible and then loaded immediately in a muffle furnace (Nabertherm P330) that was at the isothermal exposure temperature (100–800 °C). The samples were exposed in the furnace for 2 hours in ambient air atmosphere. After furnace exposures, the samples were cooled in ambient air so that handling of the crucible was possible with nitrile gloves (5–10 min) and then transferred to glass bottles that were sealed with screw caps. Then the samples were placed in a desiccator cabin for storage. During the analyses, the sample exposure times to ambient air before starting the analyses were minimized by opening the cap and preparing the sample from the bottle only just before starting the analysis. However, the XRD analysis took approximately an hour per sample, therefore possible reactions during the analysis with ambient air could not be completely eliminated.  $\text{KOH}$  especially is known to be highly hygroscopic and reactive towards  $\text{CO}_2$  during exposure to ambient air.

### 2.2 Qualitative determination of $\text{CrO}_4^{2-}$ and $\text{Cr}_2\text{O}_7^{2-}$

20 ml glass bottles were filled almost full with ion exchanged water. Then 0.10 g of each sample was added to the already water filled bottle. The bottles were not agitated by any means in order to let the sample powder sediment to the bottom of the bottle by gravity. The samples were left to stand overnight. The characteristic yellow colour of the  $\text{CrO}_4^{2-}$  ion started to appear immediately after adding the sample powder and the intensity of the colour increased with time. The sample bottles were then photographed to record the characteristic yellow colour of  $\text{CrO}_4^{2-}$  and the orange colour of  $\text{Cr}_2\text{O}_7^{2-}$ .

### 2.3 Qualitative determination of colourful Fe-complex ions

20 ml glass bottles were filled almost full with ion exchanged water. Then 0.10 g of each sample was added to the already water filled bottle. The bottles were not agitated by any means in order to let the sample powder sediment to the bottom of the bottle by gravity. The samples were let to stand still overnight. All the solutions were colourless indicating no signs of water soluble  $\text{FeCl}_3$  (yellow to orange) or  $\text{K}_2\text{FeO}_4$  ( $\text{FeO}_4^{2-}$  ion is purple) in the samples in concentrations high enough to be visible to the naked eye.



## 2.4 XRD measurements

XRD analysis was done with PANalytical X'pert Pro PW 3040/60 powder diffraction spectrometer with monochromated Cu K-alpha X-ray source ( $\alpha = 1.5406 \text{ \AA}$ ). Samples were first grinded manually in a mortar to make a visually homogeneous powder. Then the sample holder was filled evenly with the powder, pressed against a flat surface to level the sample surface with the sample holder top surface and then XRD  $\theta$ - $2\theta$  scans were recorded from  $10$ – $90^\circ$ . The automatic phase identification algorithm of the X'Pert HighScore Plus program was used for preliminary identification of the phases. Then the results were checked manually and the most likely crystalline phases were manually identified using the JCPDS cards. The XRD cards used for phase identifications were 00-021-0645 for KOH, 00-011-0655 for  $\text{K}_2\text{CO}_3 \cdot 1.5\text{H}_2\text{O}$ , 00-015-0365 for  $\text{K}_2\text{CrO}_4$ , 04-007-3113 for KCl, 01-078-5435 for  $\text{Cr}_2\text{O}_3$ , 04-012-4476 for  $\text{K}_2\text{Cr}_2\text{O}_7$ , 01-085-0599 for  $\text{Fe}_2\text{O}_3$ , 00-039-0892 for  $\text{KFeO}_2$ , 00-039-1106 for  $\text{K}_2\text{Fe}_4\text{O}_7$  and 01-078-6089 for  $\text{K}_{1.75}\text{Fe}_{1.25}\text{O}_4$ .

## 2.5 UV/VIS spectroscopic measurements of $\text{CrO}_4^{2-}$

Approximately  $0.20 \text{ g}$  of sample powder was mixed with  $20 \text{ ml}$  of room temperature ion exchanged water. The solution was first stirred in a beaker with a magnetic stirrer for  $15 \text{ min}$ . Then the solution was filtered using qualitative filter paper, 410 (Cat. No. 516-0802 from VWR), the residue washed two times with  $10 \text{ ml}$  of room temperature ion exchanged water. Then the filtrates were transferred to  $50 \text{ ml}$  volumetric flasks and the flask was filled to the mark with ion exchanged water and mixed. The  $\text{CrO}_4^{2-}$  concentrations of the filtrates were determined by UV/VIS spectroscopy using Shimadzu UV-2600 spectrophotometer and polymethylmethacrylate cuvettes. Absorbance at  $372 \text{ nm}$  was used for  $\text{CrO}_4^{2-}$ . It is not possible to determine  $\text{Cr}_2\text{O}_7^{2-}$  and  $\text{CrO}_4^{2-}$  separately, because they are in equilibrium with each other in a water solution. This equilibrium depends on the pH and  $p[\text{Cr}]$  of the solution. With pH values  $> 6.7$ , the  $\text{CrO}_4^{2-}$  ion is reported to be the stable form regardless of the Cr-concentration.<sup>39</sup> Therefore, the  $\text{K}_2\text{Cr}_2\text{O}_7$  possibly present originally in the samples was detected as  $\text{CrO}_4^{2-}$  and the concentration determined reflects the sum of  $\text{K}_2\text{CrO}_4 + \text{K}_2\text{Cr}_2\text{O}_7$  originally in the sample. The pH values of the filtrates were determined to be  $\geq 7$  with a pH indicator sticks (Fisher Scientific number 10642751). All the KCl based samples had  $\text{pH} = 7$  and KOH based samples had  $\text{pH} = 7$  in samples exposed at  $\geq 500^\circ\text{C}$  and  $\text{pH} 8$ – $12$  in samples exposed at  $\leq 400^\circ\text{C}$ . The higher pH with low temperature samples in case of KOH mixtures was caused by the KOH that was not reacted to  $\text{K}_2\text{CrO}_4$ , but formed  $\text{K}_2\text{CO}_3 \cdot 1.5\text{H}_2\text{O}$  during the furnace/ambient exposures. When dissolving in water,  $\text{K}_2\text{CO}_3 \cdot 1.5\text{H}_2\text{O}$  results in a basic solution. Concentration standards were prepared by dissolving known amounts of  $\text{K}_2\text{CrO}_4$  powder to ion exchanged water in  $50 \text{ ml}$  volumetric flasks (ESI†). Then the absorbance of the standards were measured and linear concentration–absorbance curves were established. Then the absorbance of the sample filtrates were measured and the concentrations were determined using the standard curves. In cases where the sample had  $\text{CrO}_4^{2-}$  concentration so high that the absorbance was higher than the

standards, the sample was diluted with pure water in volumetric flasks with  $1 : 10$  or  $1 : 20$  dilution ratios, which ever was suitable to bring the absorbance value in between the standards. Molar ratios reacted to  $\text{K}_2\text{CrO}_4$  and  $\text{K}_2\text{Cr}_2\text{O}_7$  were calculated as follows:

$$\frac{\text{K}_2\text{CrO}_4 + \text{K}_2\text{Cr}_2\text{O}_7}{\text{K}_2\text{CrO}_4(\text{theor.})} = \frac{2[\text{CrO}_4^{2-}]V}{\left( \frac{m}{\frac{M_{\text{KCl}}}{M_{\text{KCl}} + M_{\text{Cr}_2\text{O}_3}} M_{\text{KCl}} + \frac{M_{\text{Cr}_2\text{O}_3}}{M_{\text{KCl}} + M_{\text{Cr}_2\text{O}_3}} M_{\text{Cr}_2\text{O}_3}} \right)} \quad (1)$$

$$\frac{\text{K}_2\text{CrO}_4 + \text{K}_2\text{Cr}_2\text{O}_7}{\text{K}_2\text{CrO}_4(\text{theor.})} = \frac{2[\text{CrO}_4^{2-}]V}{\left( \frac{m}{\frac{M_{\text{KOH}}}{M_{\text{KOH}} + M_{\text{Cr}_2\text{O}_3}} M_{\text{KOH}} + \frac{M_{\text{Cr}_2\text{O}_3}}{M_{\text{KOH}} + M_{\text{Cr}_2\text{O}_3}} M_{\text{Cr}_2\text{O}_3}} \right)} \quad (2)$$

$V$  = volume, l;  $m$  = sample mass, g;  $M$  = molar mass,  $\text{g mol}^{-1}$ ;  $[\text{CrO}_4^{2-}] = \text{CrO}_4^{2-}$  ion concentration,  $\text{mol l}^{-1}$ .

Error estimation of the above described method for  $\text{K}_2\text{CrO}_4$  determination was done by running a few duplicate runs with  $\text{KOH} + \text{Cr}_2\text{O}_3$  mixtures ( $2$  at  $300$  and  $400^\circ\text{C}$  exposure temperatures). Because of the small number of duplicate samples, error was estimated with the range rather than with statistical methods. The error range (max–min) of the method was found to be about  $0.08$  (or  $8\%$ -points). This error value was assumed to be similar at other exposure temperatures and also for the leaching tests described below for  $\text{Fe}^{3+}$ .

## 2.6 UV/VIS spectroscopic measurements of $\text{Fe}^{3+}$

**2.6.1 Water soluble  $\text{Fe}^{3+}$ .** Approximately  $0.20 \text{ g}$  of sample powder was mixed with  $20 \text{ ml}$  of room temperature ion exchanged water. The solution was stirred in a beaker with a magnetic stirrer for  $15 \text{ min}$ . Then the solution was filtered using qualitative filter paper 410 (Cat. No. 516-0802 from VWR), the residue washed two times with  $10 \text{ ml}$  of room temperature ion exchanged water. Then the filtrates were transferred to  $50 \text{ ml}$  volumetric flasks and the flask was filled to the mark with ion exchanged water and mixed. The  $\text{Fe}^{3+}$  concentrations of the filtrates were determined by UV/VIS spectroscopy using Shimadzu UV-2600 spectrophotometer. Absorbance at  $225 \text{ nm}$  was used for  $\text{Fe}^{3+}$ .  $\text{Fe}^{3+}$  ion in a water solution is present as  $\text{Fe}^{3+}$ ,  $\text{Fe}(\text{OH})^{2+}$  and  $\text{Fe}(\text{OH})_2^+$  ions or as non-water soluble  $\text{Fe}(\text{OH})_3$  precipitate, depending on the pH of the solution.<sup>40</sup> The pH values of the filtrates were determined with a pH indicator sticks (Fisher Scientific number 10642751). All the KCl based samples had  $\text{pH} = 7$  and all the KOH based samples had  $\text{pH} = 12$ . The high pH value over the entire exposure temperature in case of KOH based samples indicate that they may have contained some unreacted  $\text{KOH}/\text{K}_2\text{CO}_3$  or then the higher pH at firing temperatures  $\geq 500^\circ\text{C}$  (which did not reveal any  $\text{KOH}/\text{K}_2\text{CO}_3$  residues by XRD) was originated from the dissolution of



Table 1 Input files in the thermodynamic equilibrium calculations with the HSC v.6.12 software<sup>a</sup>

KCl-Cr <sub>2</sub> O <sub>3</sub>	T [°C]	Amount [kmol]	Amount [mol%]	T [°C]	Amount [kmol]	Amount [mol%]	KOH-Cr <sub>2</sub> O <sub>3</sub>	T [°C]	Amount [kmol]	Amount [mol%]	KOH-Fe <sub>2</sub> O <sub>3</sub>	T [°C]	Amount [kmol]	Amount [mol%]	KOH-KCl	T [°C]	Amount [kmol]	Amount [mol%]
Phase 1:		0.020	100.000	Phase 1:	0.020	100.000	Phase 1:	0.020	100.000	100.000	Phase 1:	0.020	100.000	100.000	Phase 1:	0.010	100.000	100.000
KCl	25.000	0.010	50.000	Cr <sub>2</sub> O <sub>3</sub>	25.000	0.010	50.000	Cr <sub>2</sub> O <sub>3</sub>	25.000	0.010	50.000	Fe <sub>2</sub> O <sub>3</sub>	25.000	0.010	50.000	KCl	25.000	0.010
Cr <sub>2</sub> O <sub>3</sub>	25.000	0.010	50.000	Fe <sub>2</sub> O <sub>3</sub>	25.000	0.010	50.000	K <sub>2</sub> O*Cr <sub>2</sub> O <sub>6</sub>	25.000	0.010	50.000	KOH	25.000	0.010	50.000	KOH	25.000	100.000
K <sub>2</sub> O*Cr <sub>2</sub> O <sub>6</sub>	25.000	0.010	50.000	KOH	25.000	0.010	50.000	KOH	25.000	0.010	50.000	K <sub>2</sub> CO <sub>3</sub>	25.000	0.010	50.000	K <sub>2</sub> CO <sub>3</sub>	25.000	100.000
KOH	25.000	0.010	50.000	K <sub>2</sub> CO <sub>3</sub>	25.000	0.010	50.000	1.5H <sub>2</sub> O	25.000	0.010	50.000	KFeO <sub>2</sub>	25.000	0.010	50.000		25.000	100.000
K <sub>2</sub> CO <sub>3</sub> ·1.5H <sub>2</sub> O	25.000	0.010	50.000	KFeO <sub>2</sub>	25.000	0.010	50.000		25.000	0.010	50.000		25.000	0.010	50.000		25.000	100.000
K <sub>2</sub> CrO <sub>4</sub>	25.000	0.010	50.000		25.000	0.010	50.000		25.000	0.010	50.000		25.000	0.010	50.000		25.000	100.000
Phase 2:		99.996	100.000	Phase 2:	99.996	100.000	Phase 2:	99.996	100.000	100.000	Phase 2:	99.996	100.000	100.000	Phase 2:	99.995	100.000	100.000
H <sub>2</sub> O(g)	25.000	0.990	99.010	H <sub>2</sub> O(g)	25.000	0.990	99.010	H <sub>2</sub> O(g)	25.000	0.990	99.010	H <sub>2</sub> O(g)	25.000	0.990	99.010	H <sub>2</sub> O(g)	25.000	0.990
CO <sub>2</sub> (g)	25.000	0.035	3.565	CO <sub>2</sub> (g)	25.000	0.035	3.565	CO <sub>2</sub> (g)	25.000	0.035	3.565	CO <sub>2</sub> (g)	25.000	0.035	3.565	HCl(g)	25.000	0.990
O <sub>2</sub> (g)	25.000	20.740	20.740	O <sub>2</sub> (g)	25.000	20.740	20.740	O <sub>2</sub> (g)	25.000	20.740	20.740	O <sub>2</sub> (g)	25.000	20.740	20.740	CO <sub>2</sub> (g)	25.000	0.035
N <sub>2</sub> (g)	25.000	77.311	77.311	N <sub>2</sub> (g)	25.000	77.311	77.311	N <sub>2</sub> (g)	25.000	77.311	77.311	N <sub>2</sub> (g)	25.000	77.311	77.311	O <sub>2</sub> (g)	25.000	0.035
HCl(g)	25.000	0.920	9.080	HCl(g)	25.000	0.920	9.080	Ar(g)	25.000	0.920	9.080	Ar(g)	25.000	0.920	9.080	N <sub>2</sub> (g)	25.000	20.740
Ar(g)	25.000	0.920	9.080	Ar(g)	25.000	0.920	9.080	KOH(g)	25.000	0.920	9.080	KOH(g)	25.000	0.920	9.080	Ar(g)	25.000	77.310
KCl(g)	25.000	0.920	9.080	KCl(g)	25.000	0.920	9.080		25.000	0.920	9.080		25.000	0.920	9.080	KOH(g)	25.000	0.920
KOH(g)	25.000	0.920	9.080	KOH(g)	25.000	0.920	9.080		25.000	0.920	9.080		25.000	0.920	9.080	KCl(g)	25.000	20.740

<sup>a</sup> K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> is marked as K<sub>2</sub>O\*Cr<sub>2</sub>O<sub>6</sub> in the software.

KFeO<sub>2</sub> (KFeO<sub>2</sub> + 2H<sub>2</sub>O = K<sup>+</sup> + Fe<sup>3+</sup> + 4OH<sup>-</sup>). In order to shift the equilibrium so that all the dissolved Fe(III) was in the Fe<sup>3+</sup> state, the filtrates were buffered to pH = 1 by mixing 1 : 1 ratio of sample with a solution to 1 mol l<sup>-1</sup> HCl (1 : 2 dilution). This pH stabilized solution was then used to fill the quartz cuvette in the UV/VIS absorption measurements. The added Cl<sup>-</sup> ion can also form complex ions with Fe in the form Fe(H<sub>2</sub>O)<sub>6-x</sub>(Cl)<sub>x</sub><sup>3-x</sup>. Concentration standards were prepared by dissolving known amounts of FeCl<sub>3</sub>·6H<sub>2</sub>O powder to ion exchanged water and diluting (ESI†). The pH of the standards were buffered to 1 with 1 mol l<sup>-1</sup> HCl before measurement as with the samples. Then the absorbance of the standards were measured and linear concentration-absorbance curves were established. Then the absorbance of the sample filtrates were measured and the concentrations were determined using the standard curves. Molar ratios reacted to KFeO<sub>2</sub> were calculated as follows:

$$\frac{\text{KFeO}_2}{\text{KFeO}_2(\text{theor.})} = \frac{[\text{Fe}^{3+}]V}{\left( \frac{m}{\frac{M_{\text{KCl}}}{M_{\text{KCl}} + M_{\text{Fe}_2\text{O}_3}} M_{\text{KCl}} + \frac{M_{\text{Fe}_2\text{O}_3}}{M_{\text{KCl}} + M_{\text{Fe}_2\text{O}_3}} M_{\text{Fe}_2\text{O}_3}} \right)} \quad (3)$$

$$\frac{\text{KFeO}_2}{\text{KFeO}_2(\text{theor.})} = \frac{[\text{Fe}^{3+}]V}{\left( \frac{m}{\frac{M_{\text{KOH}}}{M_{\text{KOH}} + M_{\text{Fe}_2\text{O}_3}} M_{\text{KOH}} + \frac{M_{\text{Fe}_2\text{O}_3}}{M_{\text{KOH}} + M_{\text{Fe}_2\text{O}_3}} M_{\text{Fe}_2\text{O}_3}} \right)} \quad (4)$$

V = volume, l; m = sample mass, g; M = molar mass, g mol<sup>-1</sup>; [Fe<sup>3+</sup>] = Fe<sup>3+</sup> ion concentration, mol l<sup>-1</sup>.

**2.6.2 Acid soluble Fe<sup>3+</sup>.** The analysis was done in a similar way as has been discussed for the H<sub>2</sub>O soluble Fe<sup>3+</sup> determination, except that the 15 min leaching steps were done in 20 ml of 1 mol l<sup>-1</sup> HCl in water solutions instead of pure water. Also the washing of the filtration residue was done with HCl containing water (≈ 0.1 mol l<sup>-1</sup> HCl, pH = 1) instead of pure water. This prevented the precipitation of Fe(OH)<sub>3</sub> that was possible in the water leaching steps, where the pH values of the filtrates were ≥ 7. In case of 1 M HCl leaching, some filtrates had a yellow colour visible already to a naked eye, indicating that indeed the acid leaching resulted in much higher dissolution of the reaction products compared to H<sub>2</sub>O leaching. All the filtrates had a pH value of 1, measured with pH indicator sticks (Fisher Scientific number 10642751). Because the pH of the filtrates were stabilized to one already during the leaching step, they were used directly to fill the quartz cuvette in the UV/VIS absorption measurements. In cases where the sample had Fe<sup>3+</sup> concentration so high that the absorbance was higher than the standards, the sample was diluted in volumetric flasks with 1 : 2, 1 : 50 or 1 : 100 dilution ratios, which ever was suitable to bring the absorbance value in between the standards. In the 1 : 50 and 1 : 100 dilution cases, the dilution was done by





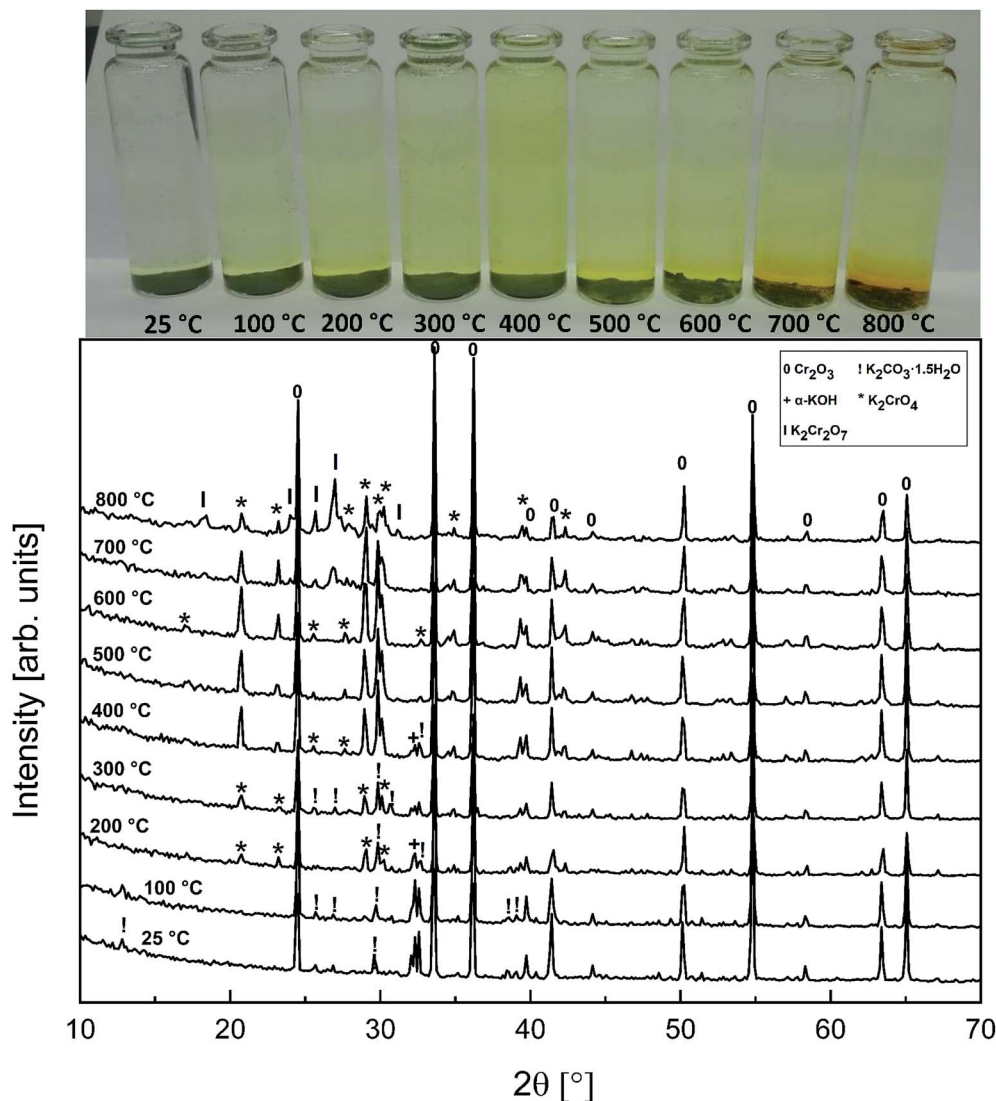


Fig. 1 XRD patterns and appearance of the water solutions on top of the KOH–Cr<sub>2</sub>O<sub>3</sub> mixtures after furnace exposures.

adding 1 ml of sample, then 5 ml of 1 mol l<sup>−1</sup> HCl, and then filling the volumetric flask with pure H<sub>2</sub>O.

## 2.7 Thermodynamic calculations

HSC v6.12 Gibbs energy minimization software was used in the thermodynamic calculations.<sup>41</sup> The input files for the calculations are presented in Table 1. Air composition<sup>42</sup> with ≈0.99 mol% H<sub>2</sub>O vapor (RH ≈ 32% at 25 °C) was used to simulate the humid ambient gas phase in the muffle furnace. In case of KCl based systems, HCl(g) was added as the possible gas phase Cl-compound released in the reactions. Note that each gas phase component in the atmosphere has at least several times higher absolute amounts than any of the solid phase components. This assures that the amount of the formed products was never limited by the amount of the gas phase component, thus simulating an open ambient system. The calculated systems were kept as simple as possible. Therefore, in addition to the reactants, only the phases identified in the

XRD analysis were added to the solid phase input file. In the case of Fe<sub>2</sub>O<sub>3</sub> systems, K<sub>2</sub>Fe<sub>4</sub>O<sub>7</sub> was detected by XRD in the sample with the highest exposure temperature with KOH, but unfortunately it was not found in the HSC v6.12 database and could not be included in the thermodynamic calculation.

## 3. Results

### 3.1 KOH–Cr<sub>2</sub>O<sub>3</sub> system

The powder XRD analysis of the system after furnace exposures is shown in Fig. 1. It appears that KOH reacted partially with the air in the furnace and formed K<sub>2</sub>CO<sub>3</sub>·1.5 H<sub>2</sub>O in samples exposed to furnace temperatures below 500 °C. Samples exposed to 500 °C or higher firing temperatures did not contain any residual K<sub>2</sub>CO<sub>3</sub>·1.5 H<sub>2</sub>O or KOH. This was also reflected by the pH of the water soluble filtrates as explained in the experimental section. It seems that the reaction of KOH with Cr<sub>2</sub>O<sub>3</sub>, forming K<sub>2</sub>CrO<sub>4</sub> is not fast enough below 500 °C and competes



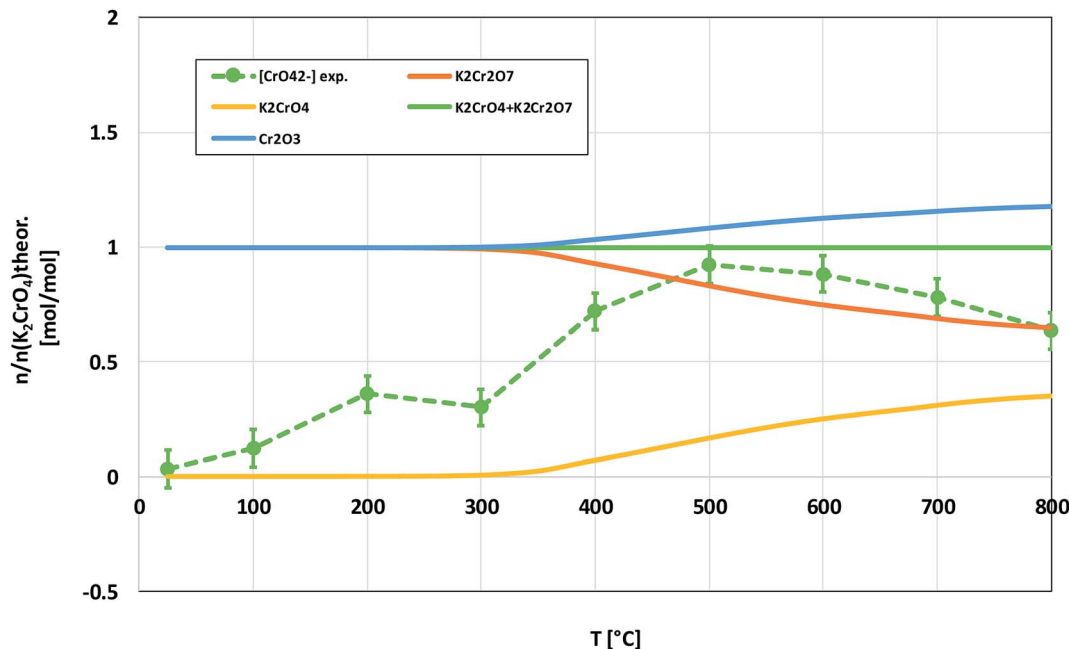


Fig. 2 Comparison of the experimentally determined amount of  $[\text{CrO}_4]^{2-}$  in the filtrates with the thermodynamically predicted amount of the reaction products in the  $\text{KOH}-\text{Cr}_2\text{O}_3$  system with pure water leaching.

with  $\text{K}_2\text{CO}_3 \cdot 1.5 \text{H}_2\text{O}$  formation in the experimental conditions used. According to XRD analysis,  $\text{K}_2\text{CrO}_4$  formation starts already with a solid-solid reaction at 200 °C. Melting point of  $\text{KOH}$  is 406 °C, so it appears that there is no need for molten

phase to form in the system before  $\text{K}_2\text{CrO}_4$  formation proceeds. The photograph in Fig. 1 reveals the characteristic colour of the  $\text{CrO}_4^{2-}$  ion already appearing in the water soluble fraction of the products at room temperature exposures. However, this

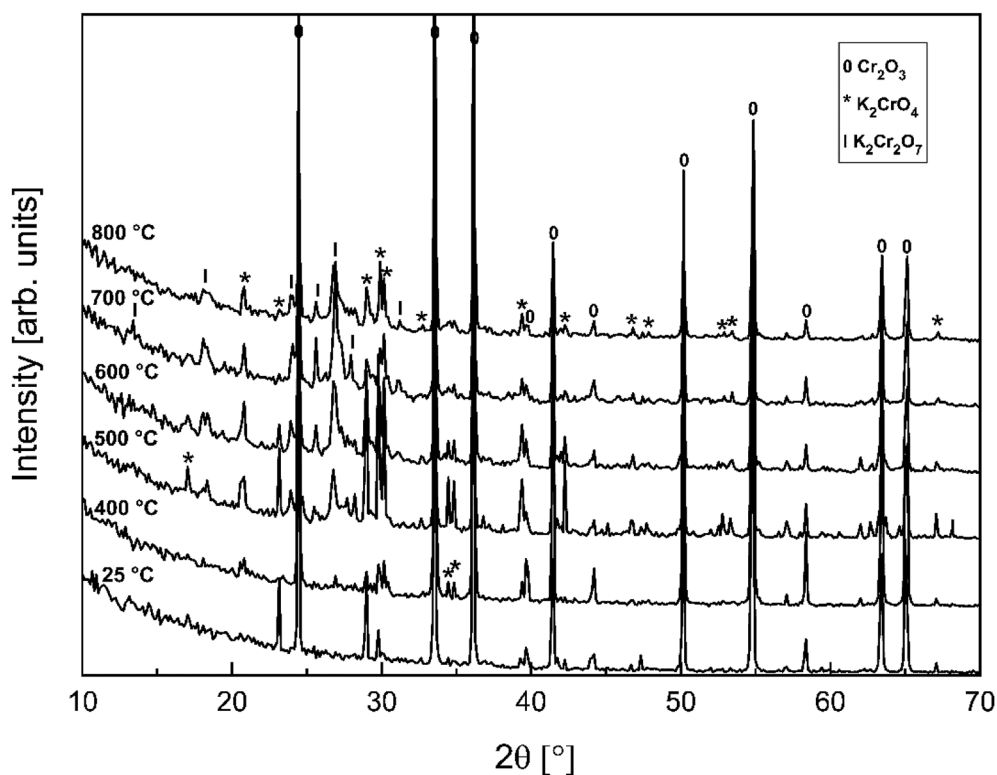


Fig. 3 XRD patterns for the  $\text{K}_2\text{CrO}_4-\text{Cr}_2\text{O}_3$  mixtures after furnace exposures.



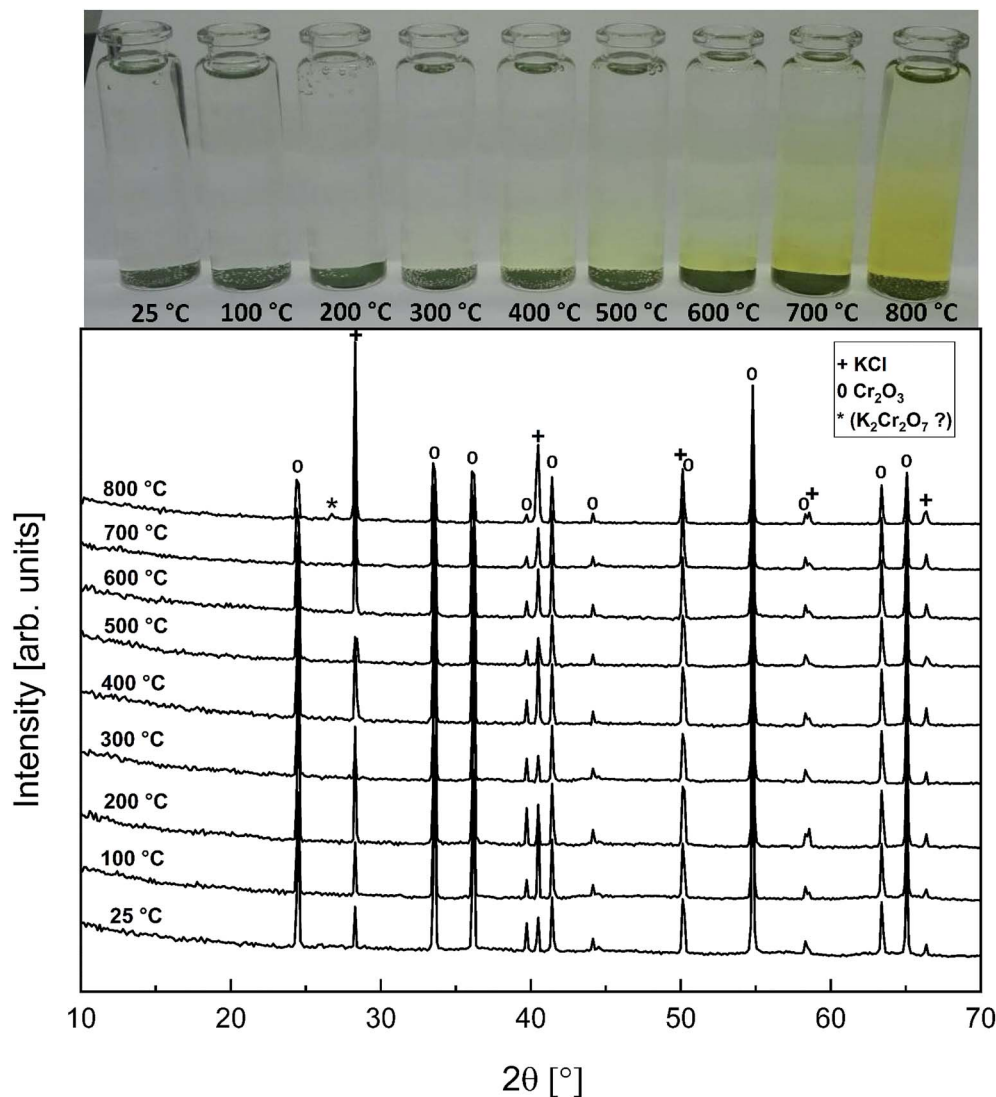


Fig. 4 XRD patterns and appearance of the water solutions on top of the KCl–Cr<sub>2</sub>O<sub>3</sub> mixtures after furnace exposures.

CrO<sub>4</sub><sup>2−</sup> may form from the dissolution of Cr<sub>2</sub>O<sub>3</sub> in the basic water solution and therefore it is considered that the XRD analysis provides a better estimation of the onset temperature where K<sub>2</sub>CrO<sub>4</sub> formation starts. The characteristic orange colour of the Cr<sub>2</sub>O<sub>7</sub><sup>2−</sup> ion was detected in samples fired at 700 °C and 800 °C. K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> was also detected in the XRD analysis at these firing temperatures. Thermodynamic calculations predict K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> to be the most stable reaction product throughout the temperature range used as shown in Fig. 2, but its formation appears to be kinetically prevented below 700 °C. The experimental results clearly show that K<sub>2</sub>CrO<sub>4</sub> formation is kinetically preferred at the studied temperature range and exposure time used. Like explained above, K<sub>2</sub>CrO<sub>4</sub> formation itself seemed also to be kinetically controlled at temperatures lower than 500 °C. The maximum CrO<sub>4</sub><sup>2−</sup> amount (>90% of the theoretical) in the reaction products was measured at 500 °C furnace exposure, after which the amount decreased slightly at higher temperatures. Taking into account the experimental errors

associated with the CrO<sub>4</sub><sup>2−</sup> analysis, it is suggested that starting at 500 °C, practically all the KOH had reacted to K<sub>2</sub>CrO<sub>4</sub> and that the slightly lower amounts of CrO<sub>4</sub><sup>2−</sup> detected at higher exposure temperatures can be explained by slight loss of KOH by evaporation, competing with the reaction with Cr<sub>2</sub>O<sub>3</sub>.

### 3.2 K<sub>2</sub>CrO<sub>4</sub>–Cr<sub>2</sub>O<sub>3</sub> system

In order to determine if the K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> formation at higher firing temperatures proceeds directly from the reaction with KOH or *via* first formation of K<sub>2</sub>CrO<sub>4</sub> and then further reaction with Cr<sub>2</sub>O<sub>3</sub>, the K<sub>2</sub>CrO<sub>4</sub>–Cr<sub>2</sub>O<sub>3</sub> system was studied with XRD analysis. Results are presented in Fig. 3. From the XRD analysis it is clear that K<sub>2</sub>CrO<sub>4</sub> can react with Cr<sub>2</sub>O<sub>3</sub> and form K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> at exposure temperatures ≥ 500 °C. Therefore, it is suggested that direct reaction of KOH with Cr<sub>2</sub>O<sub>3</sub> forms K<sub>2</sub>CrO<sub>4</sub> and that the formation of K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> requires always the K<sub>2</sub>CrO<sub>4</sub> reaction intermediate to form first.



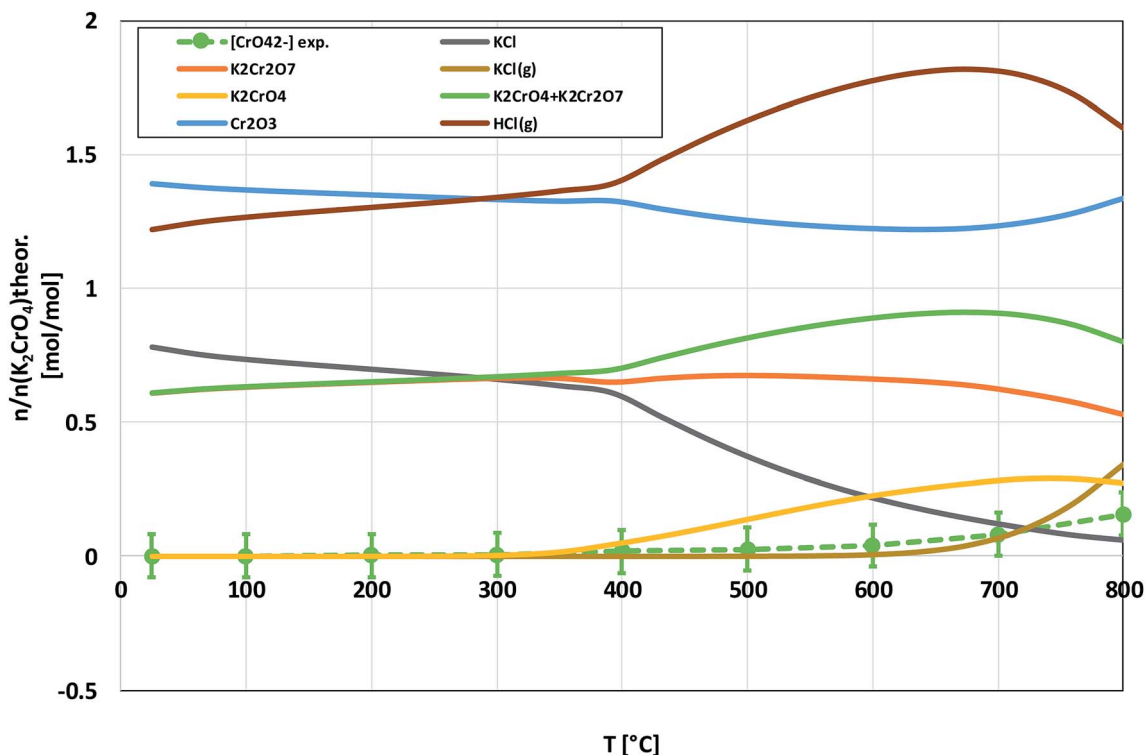


Fig. 5 Comparison of the experimentally determined amount of  $[\text{CrO}_4^{2-}]$  in the filtrates with the thermodynamically predicted amount of the reaction products in the  $\text{KCl}-\text{Cr}_2\text{O}_3$  system with pure water leaching.

### 3.3 $\text{KCl}-\text{Cr}_2\text{O}_3$ system

Compared to  $\text{KOH}$ ,  $\text{KCl}$  is substantially less reactive with  $\text{Cr}_2\text{O}_3$ . As shown in the XRD analysis in Fig. 4, there were no clear signs of reaction with  $\text{Cr}_2\text{O}_3$  until the highest firing temperature,  $800^\circ\text{C}$ . At  $800^\circ\text{C}$ , an additional peak was detected at  $26.95^\circ (2\theta)$ . This peak coincides with the (021) reflection of  $\text{K}_2\text{Cr}_2\text{O}_7$ . However, using only one peak in phase identification is not reliable, but based on the analogy with results from the  $\text{KOH}-\text{Cr}_2\text{O}_3$  and  $\text{K}_2\text{CrO}_4-\text{Cr}_2\text{O}_3$  systems,  $\text{K}_2\text{Cr}_2\text{O}_7$  was considered to be the most likely reaction product at  $800^\circ\text{C}$ .  $\text{CrO}_4^{2-}$  was detected qualitatively and quantitatively also at lower firing temperatures as shown in Fig. 4 and 5. Thermodynamic calculation predicted again the formation of  $\text{K}_2\text{Cr}_2\text{O}_7$  and  $\text{K}_2\text{CrO}_4$  as the major products, but now the discrepancy between the thermodynamic prediction and experimental results was higher than with  $\text{KOH}$ .  $\text{CrO}_4^{2-}$  was detected only in samples exposed to firing temperatures  $\geq 400^\circ\text{C}$  and even at  $800^\circ\text{C}$  the  $\text{CrO}_4^{2-}$  amount was only  $\approx 16\%$  of the theoretical maximum.

### 3.4 $\text{KOH}-\text{Fe}_2\text{O}_3$ system

Fig. 6 shows the XRD analysis of the  $\text{KOH}-\text{Fe}_2\text{O}_3$  system. Reaction with  $\text{Fe}_2\text{O}_3$  started at  $200^\circ\text{C}$  firing temperature, forming  $\text{KFeO}_2$  as the reaction product.  $\text{KFeO}_2$  stayed as the major reaction product throughout the temperature range. At  $500^\circ\text{C}$  an additional peak at  $27.95^\circ (2\theta)$  was detected that could not be assigned to  $\text{KFeO}_2$ . This peak was assigned to potassium containing magnetite,  $\text{K}_{1.75}\text{Fe}_{1.25}\text{O}_4$  that has the maximum intensity powder XRD peak at this position, but with one peak

only, the identification remains doubtful. At  $800^\circ\text{C}$  firing temperature,  $\text{K}_2\text{Fe}_4\text{O}_7$  could also be identified as the reaction product in addition to  $\text{KFeO}_2$ . There was no indication of higher oxidation state than +3 reaction products of Fe, such as  $\text{K}_2\text{FeO}_4$ . This was consistent with the lack of any colour of the water solutions containing the dissolved reaction products shown in Fig. 6. This result reflects the more favourable tendency of Cr to adapt oxidation state +6 compared to Fe. The comparison of the measured water soluble and acid soluble  $\text{Fe}^{3+}$  amounts with the thermodynamic prediction are presented in Fig. 7. No water soluble  $\text{Fe}^{3+}$  was found, but the amount of acid soluble  $\text{Fe}^{3+}$  was very close to the thermodynamically predicted amount of  $\text{KFeO}_2$ . This difference in the filtrates as a function of pH is caused by the fact that either  $\text{KFeO}_2$  is not soluble in water, or that after initial dissolution of  $\text{KFeO}_2$ , rapid formation of insoluble  $\text{Fe}(\text{OH})_3$  takes place in the basic solution. The high  $\text{OH}^-$  concentration precipitates the initially dissolved iron and captures it in the filtration residue. In acidic conditions,  $\text{Fe}(\text{OH})_3$  formation is prevented and  $\text{KFeO}_2$  was dissolved completely in the 15 min leaching step and the dissolved iron stayed in the solution phase during the filtration step.

### 3.5 $\text{KCl}-\text{Fe}_2\text{O}_3$ system

As with the  $\text{KCl}-\text{Cr}_2\text{O}_3$  system,  $\text{KCl}$  turned out to be much less reactive towards  $\text{Fe}_2\text{O}_3$  compared to  $\text{KOH}$ . There was no sign of any reaction in the XRD analysis or in the visual qualitative analysis of the water soluble reaction products as shown in Fig. 8. Similar to the  $\text{KOH}-\text{Fe}_2\text{O}_3$  case, no water soluble iron was





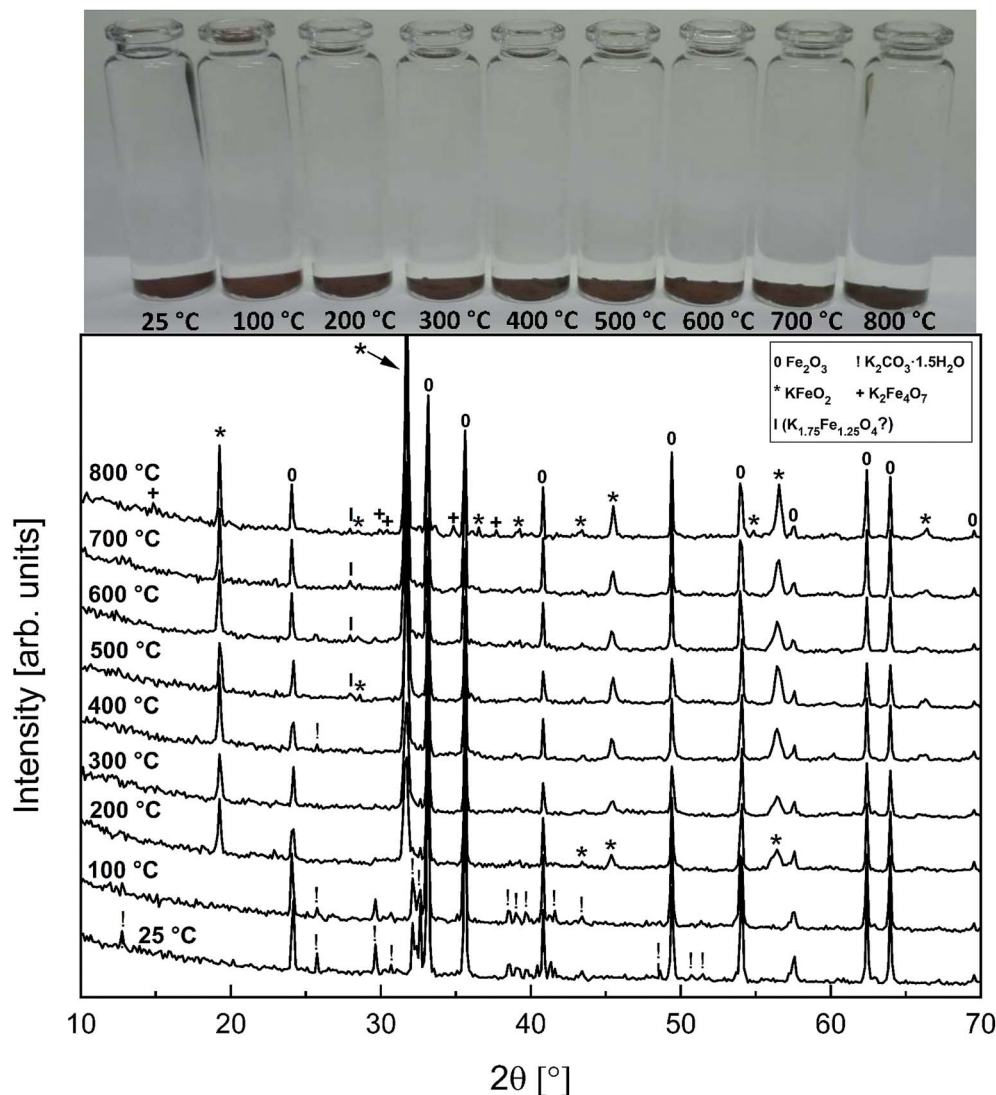
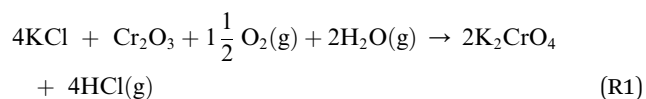


Fig. 6 XRD patterns and appearance of the water solutions on top of the KOH–Fe<sub>2</sub>O<sub>3</sub> mixtures after furnace exposures.

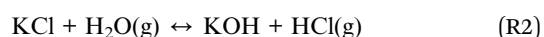
found, but now there was only traces of Fe<sup>3+</sup> detected even in the acid soluble filtrate. The results agree with thermodynamic equilibrium calculations, which predicted only slight reaction in the KCl–Fe<sub>2</sub>O<sub>3</sub> system at temperatures > 600 °C to form KFeO<sub>2</sub> as shown in Fig. 9.

## 4. Discussion

Our results are consistent with the suggestion that the reaction of KCl with Cr<sub>2</sub>O<sub>3</sub> is initiated by K<sub>2</sub>CrO<sub>4</sub> formation as found in the corrosion literature.<sup>22,23,43</sup> However, we argue that the suggested overall reaction of KCl with Cr<sub>2</sub>O<sub>3</sub>:



starts with the initiation step:



and then the formed KOH can easily react with the protective oxide as shown in the experiments. Previous hypotheses of the mechanisms of steel corrosion under KCl(s) exposures have focused on the effect of Cl<sub>2</sub>(g) or HCl(g) as the corrosive substances and the effect of KOH has been neglected.<sup>44,45</sup> However, combustion environments contain inevitably also H<sub>2</sub>O(g) in the flue gases and therefore KOH formation should not be neglected. Furthermore, many virgin biomass fuels contain very little Cl, so there is simply not much Cl<sub>2</sub>(g) or HCl(g) available in the flue gases when combusting these types of fuels. Still KCl layer has been found on the heat exchanger surfaces and previous work has argued that the corroding Cl<sub>2</sub>(g) or HCl(g) must form from the KCl layer, but the fate of the K when the Cl<sub>2</sub>(g) or HCl(g) forms from KCl has not been studied in detail. In waste incineration, the situation is different, because the feedstock contains typically large amounts of Cl and therefore the direct gas phase attack by Cl<sub>2</sub>(g) or HCl(g) may be more relevant. It may be that the chlorine corrosion



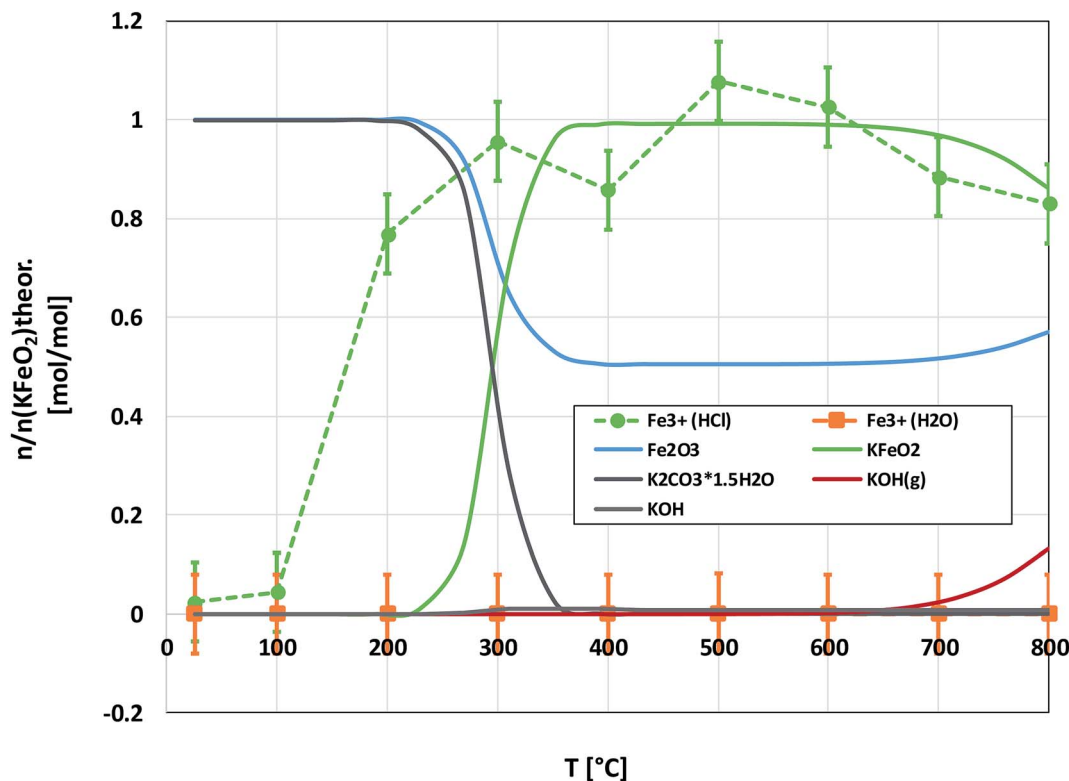
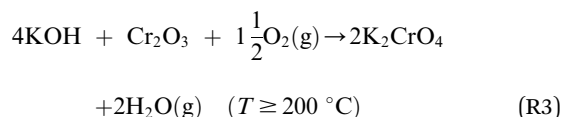


Fig. 7 Comparison of the experimentally determined amount of  $[\text{Fe}^{3+}]$  in the filtrates with the thermodynamically predicted amount of the reaction products in the  $\text{KOH}-\text{Fe}_2\text{O}_3$  system with pure water leaching and leaching in 1 M  $\text{HCl}_{(\text{aq})}$  solution.

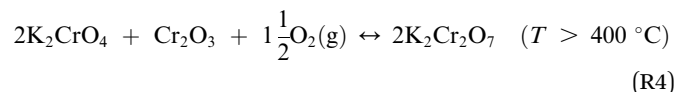
mechanisms taking place in high chlorine content feedstock combustion cannot be applied to explain the corrosion in virgin biomass combustion.

#### 4.1 Reaction of KOH with $\text{Cr}_2\text{O}_3$

KOH will react readily with  $\text{Cr}_2\text{O}_3$ , reacting already at 200 °C:

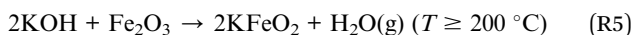


The formed  $\text{K}_2\text{CrO}_4$  can react further to  $\text{K}_2\text{Cr}_2\text{O}_7$  at high temperatures. Direct formation of  $\text{K}_2\text{Cr}_2\text{O}_7$  without the  $\text{K}_2\text{CrO}_4$  intermediate is always kinetically hindered:



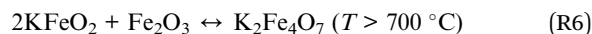
#### 4.2 Reaction of KOH with $\text{Fe}_2\text{O}_3$

KOH will react readily with  $\text{Fe}_2\text{O}_3$ , reacting already at 200 °C:



Although not experimentally verified in this study, we propose that in analogy to  $\text{KOH}-\text{Cr}_2\text{O}_3$  case, the formed  $\text{KFeO}_2$

can react further to  $\text{K}_2\text{Fe}_4\text{O}_7$  at high temperatures. It is suggested that  $\text{K}_2\text{Fe}_4\text{O}_7$  formation without the  $\text{KFeO}_2$  intermediate is kinetically hindered as the  $\text{K}_2\text{Cr}_2\text{O}_7$  formation in the  $\text{Cr}_2\text{O}_3$  case:



#### 4.3 Reaction of KCl with $\text{Cr}_2\text{O}_3$ and $\text{Fe}_2\text{O}_3$

Based on the results in this study, it is proposed that the reactivity of KCl towards both oxides in  $\text{H}_2\text{O}(\text{g})$  containing environment can simply be explained by the thermodynamics of the decomposition reaction R2. Fig. 10 presents the thermodynamic equilibrium of R2 in ambient moist air. It is suggested that  $\text{HCl}(\text{g})$  does not play a key role in the reaction with the oxides and that the reactive compound is the formed KOH. The reason that we did not detect any reaction between KCl and  $\text{Fe}_2\text{O}_3$ , but did detect  $\text{K}_2\text{CrO}_4$  formation with  $\text{Cr}_2\text{O}_3$  at  $>500 \text{ }^\circ\text{C}$  can be explained by the higher thermodynamic stability of  $\text{K}_2\text{CrO}_4$  compared to  $\text{KFeO}_2$  as shown in Fig. 11.  $\text{K}_2\text{CrO}_4$  is much more stable than  $\text{KFeO}_2$ , therefore it is suggested that the KOH formed in the decomposition of KCl at  $>500 \text{ }^\circ\text{C}$  does not form  $\text{KFeO}_2$  as readily as  $\text{K}_2\text{CrO}_4$  and most of the formed KOH is lost in the gas phase by vaporization before reacting with  $\text{Fe}_2\text{O}_3$ . That is why  $\text{KFeO}_2$  was not detected in the  $\text{KCl}-\text{Fe}_2\text{O}_3$  system.

In case of Fe–Cr steel corrosion, once the protective oxide is destroyed by KOH and if the bare metals are exposed to the



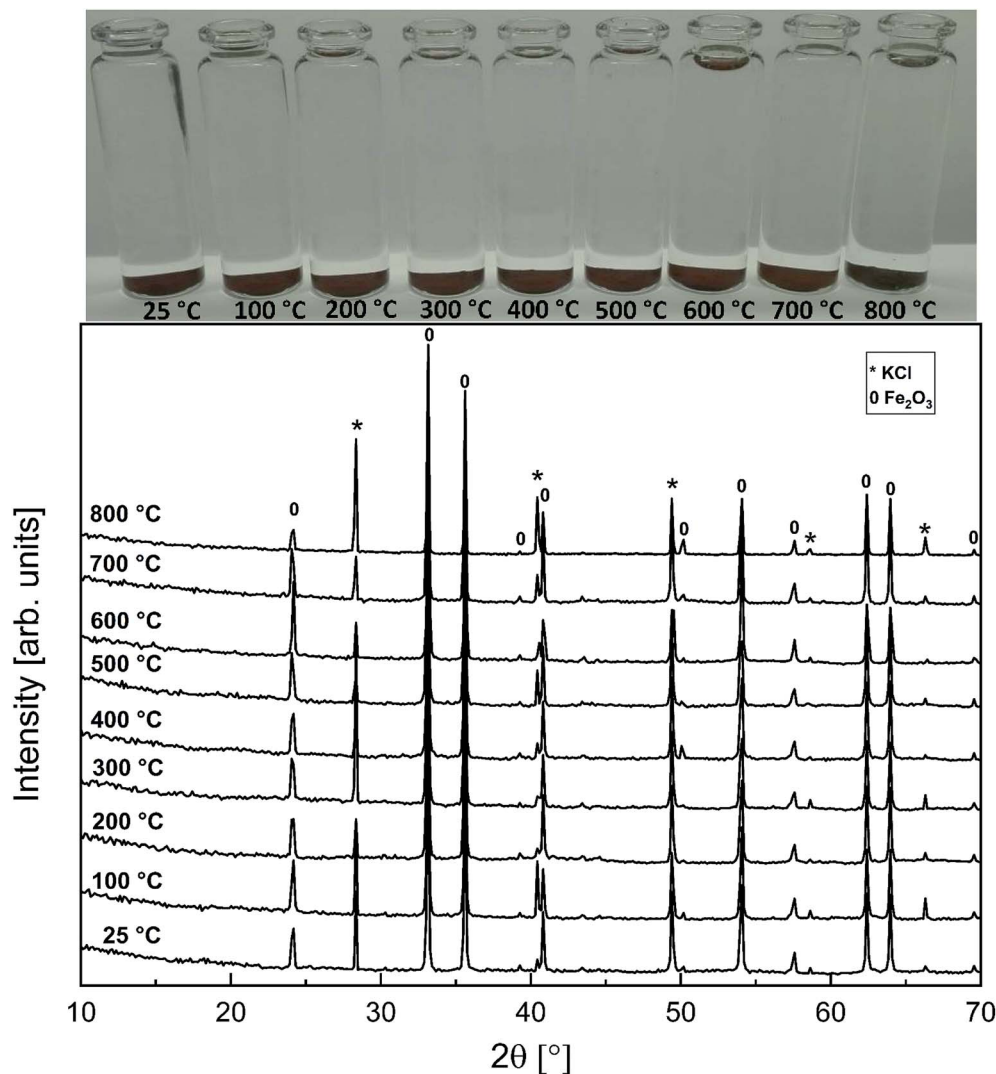
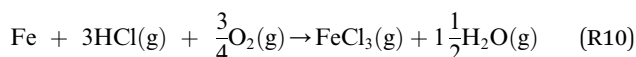
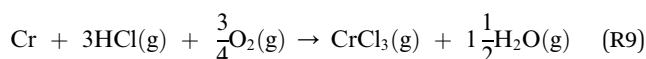
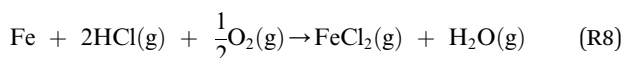
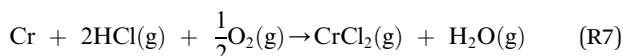


Fig. 8 XRD patterns and appearance of the water solutions on top of the KCl–Fe<sub>2</sub>O<sub>3</sub> mixtures after furnace exposures.

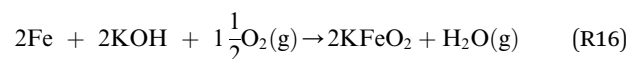
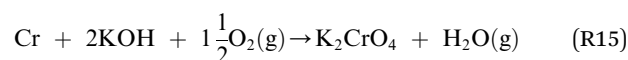
atmosphere, HCl(g) may take part in the continuation of corrosion:



If KCl diffuses through the oxide scale (grain boundary diffusion), the formation of the reactive chlorine species can also take place at the metal-scale interface:



The formed chlorine compounds can further attack the metal leading to the active oxidation mechanism by chlorine.<sup>46</sup> However, it is somewhat arbitrary to speculate about the corrosion reactions with elemental metals, because these reactive metals in the elemental form will thermodynamically favour reaction with practically any reactant. For example, a chlorine free corrosion mechanism involving only KOH may also be suggested:



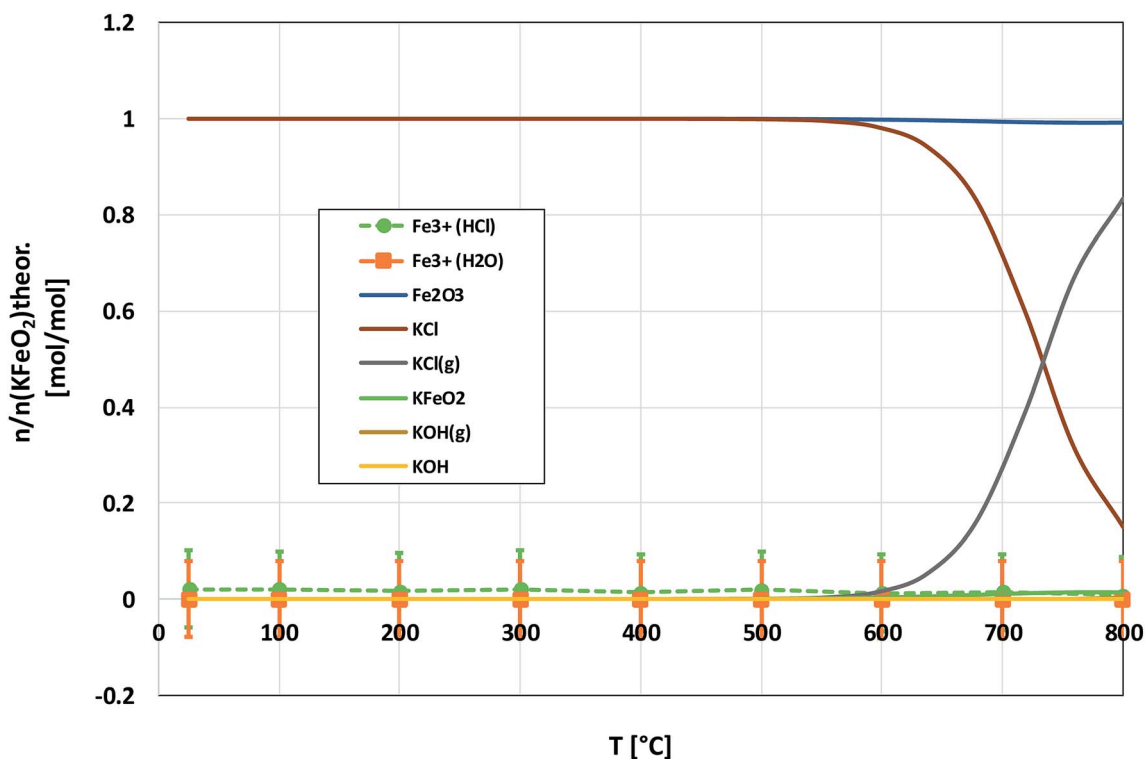


Fig. 9 Comparison of the experimentally determined amount of  $[\text{Fe}^{3+}]$  in the filtrates with the thermodynamically predicted amount of the reaction products in the  $\text{KCl}\text{--}\text{Fe}_2\text{O}_3$  system with pure water leaching and leaching in 1 M  $\text{HCl}_{(\text{aq})}$  solution.

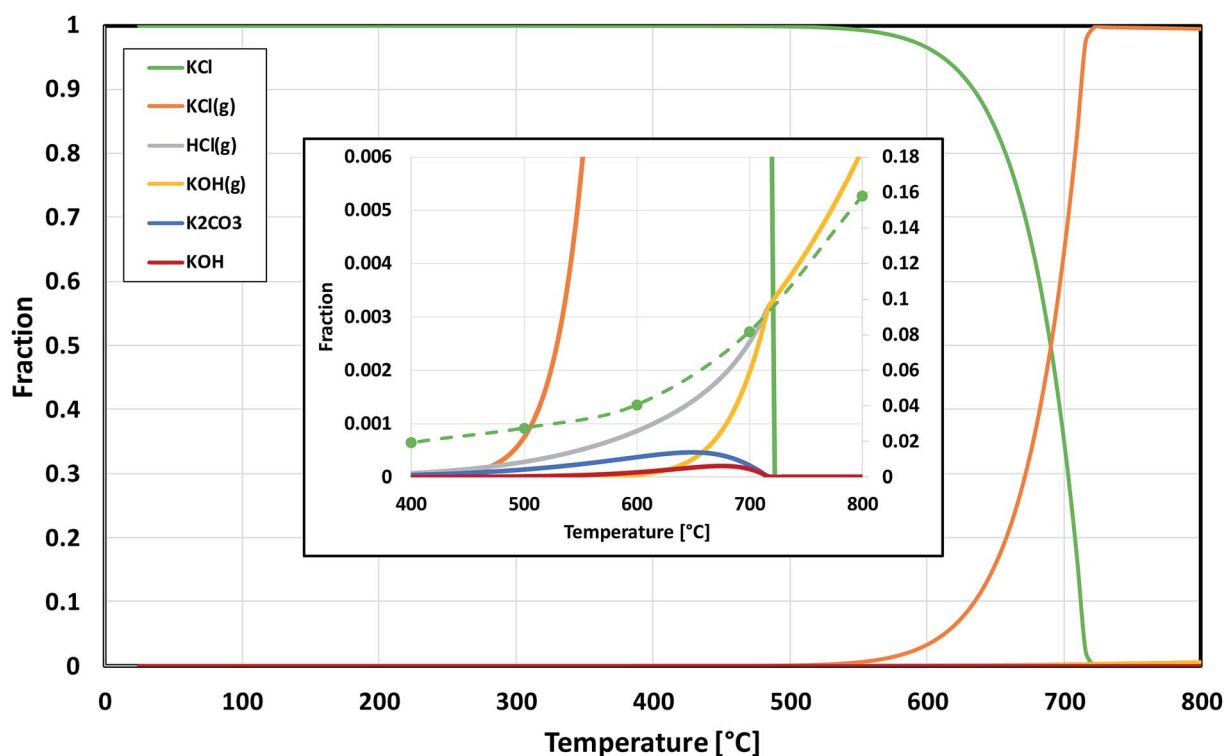


Fig. 10 Thermodynamic stability of  $\text{KCl}$  in humid air, calculated with HSC v6.12 software. The inset is a zoomed image from the bottom right corner of the chart. The green dashed line in the inset shows the trend of the experimentally determined amount of potassium reacted to  $\text{K}_2\text{CrO}_4$  when  $\text{KCl}$  and  $\text{Cr}_2\text{O}_3$  reacted in the furnace (shown also in Fig. 5).





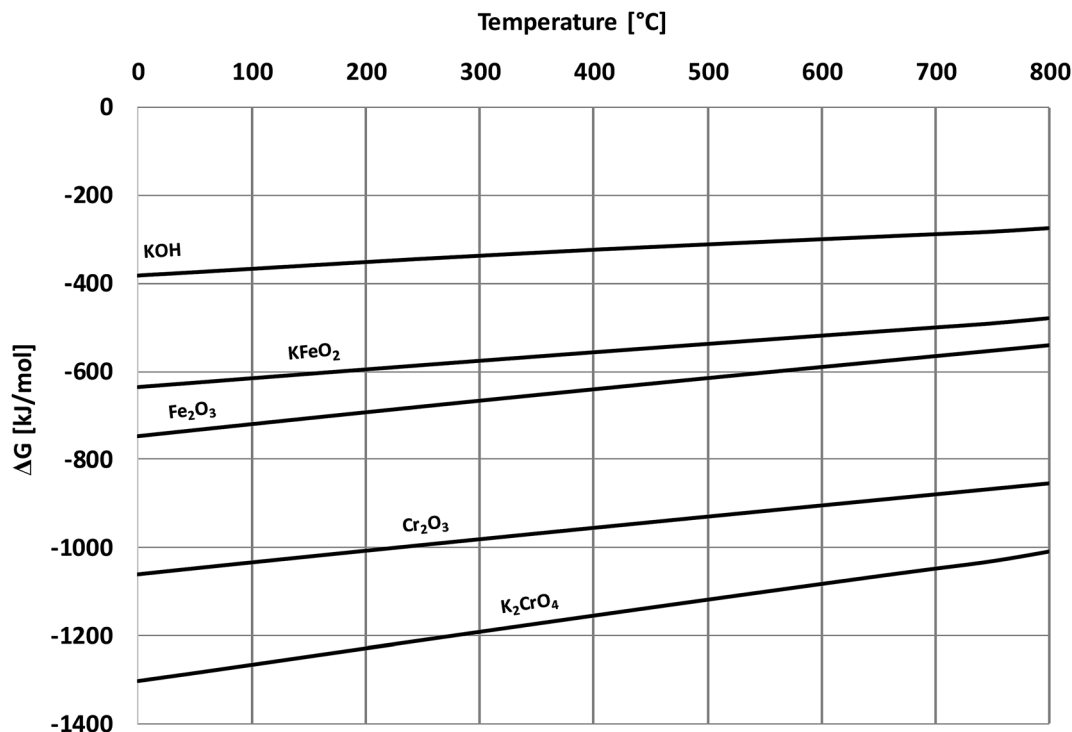


Fig. 11 Comparison of Gibbs free energies of formation of the reactants and the corrosion products, calculated with HSC v6.12 software.

The formed water or the  $O_2(g)$  diffusing through the oxide scale can further oxidize the underlying metal and form  $Cr_2O_3$  or  $Fe_2O_3$  oxides, which are then consumed by KOH according to reactions R3 and R5, and a chlorine free active oxidation mechanism is established. It is argued that KOH diffuses more readily to the metal scale interphase than KCl, because of the lower melting point of KOH (406 °C) compared to KCl (773 °C) and thus higher mobility at typical service temperatures (400–600 °C). Because of the low melting point of KOH, diffusion can

also take place in ionic form by  $K^+$  and  $O^{2-}$  ions diffusing to the metal-scale interphase:

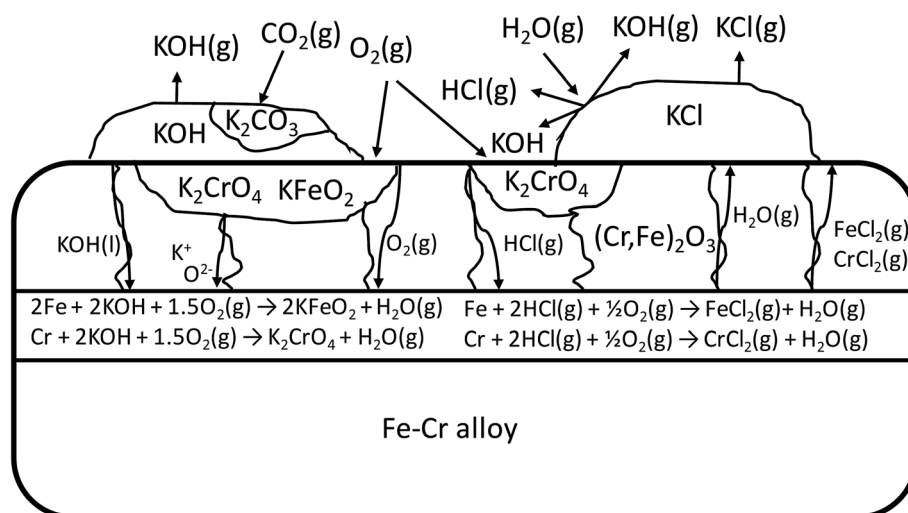
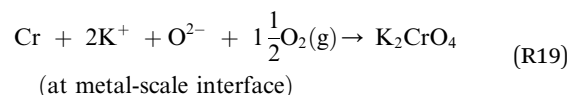
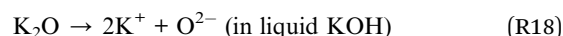
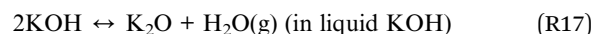
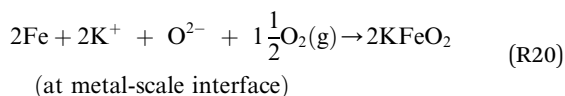


Fig. 12 Schematic of the proposed corrosion reactions under KOH and KCl containing salt deposit.





The schematic image in Fig. 12 summarizes the proposed corrosion reactions.

## 5. Conclusions

KOH is much more reactive towards  $\text{Cr}_2\text{O}_3$  and  $\text{Fe}_2\text{O}_3$  than KCl in ambient air environment. Thermodynamic calculations predict the difference in reactivity quite nicely above 400 °C. KOH reacts with both oxides at temperatures higher than 200 °C while KCl reacts with  $\text{Cr}_2\text{O}_3$  at temperatures > 400 °C, but no reaction with  $\text{Fe}_2\text{O}_3$  was detected in the temperature range from room temperature to 800 °C. Both KOH and KCl form  $\text{K}_2\text{CrO}_4$  as the reaction product when reacting with  $\text{Cr}_2\text{O}_3$ .  $\text{K}_2\text{CrO}_4$  can further react with  $\text{Cr}_2\text{O}_3$  and form  $\text{K}_2\text{Cr}_2\text{O}_7$  at temperatures > 400 °C.  $\text{K}_2\text{CrO}_4$  and  $\text{K}_2\text{Cr}_2\text{O}_7$  reaction products can be easily leached from the mixture with water at room temperature. KOH forms  $\text{KFeO}_2$  as the reaction product when reacting with  $\text{Fe}_2\text{O}_3$ .  $\text{KFeO}_2$  cannot be leached with pure water, but requires an acidic media. In 1 M HCl solution,  $\text{KFeO}_2$  can be easily leached at room temperature. The reactivity of KCl towards the protective oxides of Fe–Cr steels in water containing high temperature environments can be explained by its decomposition to KOH and  $\text{HCl}(\text{g})$  and the subsequent reaction of the formed KOH with the protective oxide. These results are very valuable for the development of materials for boilers, gasifiers, furnaces and gas turbines utilizing biomass derived fuels as their energy source.

## Conflicts of interest

There are no conflicts of interest to declare.

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