Preparation of Sn-aminoclay (SnAC)-templated Fe₃O₄ nanoparticles as an anode material for lithium-ion batteries†

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Sn-aminoclay (SnAC)-templated Fe₃O₄ nanocomposites (SnAC–Fe₃O₄) were prepared through a facile approach. The morphology and macro-architecture of the fabricated SnAC–Fe₃O₄ nanocomposites were characterized by different techniques. A constructed meso/macro-porous structure arising from the homogeneous dispersion of Fe₂O₃ NPs on the SnAC surface owing to inherent NH₃⁺ functional groups provides new conductive channels for high-efficiency electron transport and ion diffusion. After annealing under argon (Ar) gas, most of SnAC layered structure can be converted to SnO₂; this carbonization allows for formation of a protective shell preventing direct interaction of the inner SnO₂ and Fe₂O₃ NPs with the electrolyte. Additionally, the post-annealing formation of Fe–O–C and Sn–O–C bonds enhances the connection of Fe₃O₄ NPs and SnAC, resulting in improved electrical conductivity, specific capacities, capacity retention, and long-term stability of the nanocomposites. Resultantly, electrochemical measurement exhibits high initial discharge/charge capacities of 980 mA h g⁻¹ and 830 mA h g⁻¹ at 100 mA g⁻¹ in the first cycle and maintains 710 mA h g⁻¹ after 100 cycles, which corresponds to a capacity retention of ~89%. The cycling performance at 100 mA g⁻¹ is remarkably improved when compared with control SnAC. These outstanding results represent a new direction for development of anode materials without any binder or additive.

Introduction

Lithium-ion batteries (LIBs), thanks to their high specific energy density, long lifespan, and environmental benignity, are considered to be the dominant energy storage systems among the many employed in portable electronics and electric vehicles. However, the continuous surge necessary to satisfy the ever-increasing demand for high-performance, long-life, and low-cost power systems cannot be supplied by traditional carbon-based anodes. Therefore, worldwide efforts have been devoted to the search for alternative anode materials that can deliver improved energy density and safety. Specifically, remarkably improved charge/discharge rate capability in LIBs requires new nanomaterials and suitable nanostructures.

Transition metal oxides (TMOs) have been investigated and fabricated in various morphologies in efforts to replace graphite for LIBs. Among the TMOs, magnetite nanoparticles (Fe₃O₄ NPs) have attracted much attention in the field of electrochemical materials owing to particular advantages including high abundance, environmental benignity, low cost, biocompatibility, low toxicity, high electronic conductivity as well as high storage capacity (926 mA g⁻¹). However, Fe₃O₄ NPs, like other TMOs, have drawbacks arising from the aggregation and huge volume expansion due to the strong anisotropic dipolar interactions and the mechanism of the structural changes of both active materials and electrodes during the charge/discharge process. These issues cause cracking, fracture, and pulverization of NPs, leading to loss of their dispersibility and specific properties, rapid capacity decay as well as instability of the solid electrolyte interface (SEI). Hybridization between carbonaceous materials (with their high electrical conductivity and large surface area) and TMOs (with their high capacity density) is an effective approach to address those challenges in mitigating adverse mechanical effects and enhancing electrochemical performance. In recent years, a large number of studies following this strategy have been reported; however, carbonaceous materials still show limitations in the synthesis processes, which unfortunately are rather complicated and require expensive, special equipment and non-eco-
SnAC materials as synthesized by the co-precipitation method, were uniformly groups on the SnAC surface in aqueous solution, the Fe3O4 NPs, SnO2, which enhanced its specific and to convert most of the SnAC architecture to crystalline functioning as an electrode material for LIBs. A novel hybrid structure with inter-networks and excellent stability.

Results and discussion

Results

The annealed SnAC–Fe3O4 nanocomposites were fabricated by a simple process based on a combination of the normal synthesis method for SnAC and the co-precipitation method for Fe3O4 NPs. In step 1, SnAC was synthesized successfully, showing high dispersibility in DI water. The zeta potential (ζ) measurement indicated that the SnAC was positively charged with a wide pH range of 2 to 11, owing for protonation of NH3+ functional groups on the surface of the SnAC-layered structure (Fig. 1b). Notably, in the aqueous solution, SnAC presented the high ζ potential value of +45.53 mV. In step 2, Fe3O4 NPs were formed by chemical reaction, showing a negative charge owing to the many hydroxyl groups on the particle surfaces under the high-pH condition. The ζ potential values also were evident at pH 9 (−33.34 mV) and pH 11 (−43.92 mV), respectively (Fig. 1c). At pH 10, the ζ potential value of the SnAC–Fe3O4 NPs was 32.80 mV. A change in the ζ potential value was clearly observed when compared with those in the initial samples (SnAC at pH 9: +5.3 mV; Fe3O4 NPs at pH 10: −43.75 mV) under the same condition. Based on these results, we can tentatively conclude that the uniform dispersion and strong interaction of the Fe3O4 NPs with the SnAC surface was due to the electrostatic force and chemical linkage in step 3. Especially, after the annealing process under Ar gas at 500 °C, most of the SnAC structure was transformed to the SnO2 phase within the SnAC–Fe3O4 nano-architecture, which remarkably enhanced the specific capacity and electrode stability and facilitated the electron transfer, all of which would make for improved LIB performance (Fig. 2).

The morphologies of the SnAC–Fe3O4 nanocomposites before and after the annealing process were investigated by SEM.

![Fig. 1](Image)  
(a) Schematic illustration of fabrication of SnAC–Fe3O4 nanocomposites; (b and c) ζ potentials of SnAC and Fe3O4 NPs under different pH conditions.
and TEM. In the SEM images, the uniform distribution of the Fe₃O₄ NPs, without aggregation, was clearly apparent. Most of the Fe₃O₄ NPs were spherical in shape, with an average size ranging from 5 nm to 10 nm. The NPs’ uniform distribution, as confirmed by elemental mapping of the Fe, Sn, O, and Si (Fig. 3–5), indicated the co-existence of Fe₃O₄ NPs and SnAC in this hybrid architecture. The structure of SnAC–Fe₃O₄ was further analyzed via TEM images (Fig. 4 and 5). For the pre-annealing SnAC–Fe₃O₄ nanocomposites, Fe₃O₄ formation was observed, though, the crystalline phase was not yet clear. By contrast, for the post-annealed SnAC–Fe₃O₄, the nanocrystallites of Fe₃O₄ could be easily observed. For two of the samples, the bright circular rings in the selected area electron diffraction (SAED) patterns indicated the presence of the cubic phase of Fe₃O₄. Especially, in the annealed SnAC–Fe₃O₄, the SAED patterns also showed the existence of orthorhombic-phase SnO₂. The obtained images agreed with the high-resolution TEM images for the existence of the lattice fringes of Fe₃O₄ and SnO₂ in the hybrid structure with interplanar distances of 0.61 nm, and 0.25 nm, corresponding to the d(511), and d(311) lattice planes of Fe₃O₄, respectively, and the d spacing of 0.35 nm corresponding to the (110) plane of SnO₂.

In order to further investigate the morphologies of the synthesized architecture, the N₂ adsorption/desorption was measured for both the annealed SnAC–Fe₃O₄ and annealed SnAC. The obtained results clearly exhibited typical II hysteresis loops in the 0–1 relative range, as characteristic of the different pore sizes (Fig. 6). The significant hysteresis characteristics of the nitrogen adsorption–desorption isotherms indicated the existence of a meso/macroporous architecture. The surface areas of the annealed SnAC–Fe₃O₄ and annealed SnAC were determined to be 18.7 m² g⁻¹ and 33.42 m² g⁻¹ by Brunauer–Emmett–Teller (BET) calculations, respectively. The curves of the pore-size distributions represented the pore sizes in the 5–10 nm range for the annealed SnAC. Contrastingly, the annealed SnAC–Fe₃O₄ nanocomposites showed a broader pore-size distribution, up to 24 nm, with an average pore volume up to 0.112 cm³ g⁻¹. The
Fourier-transform infrared spectroscopy (FTIR) and XPS were performed to confirm the surface functional groups and to determine the presence of new bindings for the synthetic materials. According to the obtained FTIR results, the strong absorption band of the samples at around 3398 cm\(^{-1}\) before and after annealing corresponded to the O–H stretching of water molecules adsorbed on the material. The peak bands located at 2923 cm\(^{-1}\) and 1342 cm\(^{-1}\) were considered to be the stretching and bending vibrations of C–H. In addition, the absorption band at 1573 cm\(^{-1}\) was related to the N–H vibration of the NH\(_2\) groups in SnAC. Moreover, the sharp peaks at 892 cm\(^{-1}\), 854 cm\(^{-1}\), and 721 cm\(^{-1}\) were attributed to Si–C, Si–O, and Sn–O in the SnAC structure, respectively. Notably, the band at 632 cm\(^{-1}\) was assigned to Fe–O in the Fe\(_3\)O\(_4\) molecules. For the annealed SnAC–Fe\(_2\)O\(_4\) nanocomposites, it was observed that the intensity of some of the major peaks in the SnAC structure, such as O–H, C–H, N–H, and Si–C, were significantly reduced. This was attributed to the degradation of water molecules and organic pendants in the SnAC structure during the annealing process. As expected, the sharp peaks of Sn–O, Si–O, and Fe–O were remarkably increased.\(^{19,20}\)

The XPS spectra indicated the presence of the SnO\(_2\) phase with major peaks at Sn 3d, 4d, and 3p (see Fig. S1†). The binding energy peaks at 486.7 eV and 495.6 eV corresponded to the Sn 3d\(^{22}\) and Sn 3d\(^{21}\) spin–orbit peaks of the SnO\(_2\) and Sn–O bonds in SnAC, respectively. The presence of Fe\(_2\)O\(_4\) was confirmed by the two peaks at 710 and 725 eV corresponding to Fe 2p\(^{31}\) and Fe 2p\(^{13}\) of Fe\(_3\)O\(_4\), respectively.\(^{21,22}\) The O 1s spectra of the SnAC–Fe\(_2\)O\(_4\) sample before and after annealing are plotted (Fig. 8a1 and b1). The peak at 529 eV in both two samples was ascribed to the Fe–O bonds in the Fe\(_3\)O\(_4\) NPs. The peak at about 531.6 eV could be ascribed to the Sn–O or C==O bonds. Under the annealing treatment, the SnAC–Fe\(_2\)O\(_4\) nanocomposites showed new two bonds at 532.4 eV and 533 eV, which were assigned to the Fe–O–C and Sn–O–C bonds owing to the strong deposition of the Fe\(_3\)O\(_4\) NPs onto SnAC and the carbonization of SnAC structure.\(^{13,24}\) These bonds were clearly confirmed by the decreased relative intensity of the Fe–O peaks and Sn–O or C==O peaks, coupled with the strong decreasing of peak area (from A = 2106 to 1800 for Fe–O, and from A = 6280 to 2059 for Sn–O) after annealing relative to those of the SnAC–Fe\(_2\)O\(_4\) nanocomposites before annealing. The carbonization of most of the SnAC structure led to the coordination of TMOs with the inherent carbon source on the SnAC surface, which formed stronger new bonds among the SnO\(_2\) NPs, Fe\(_3\)O\(_4\) NPs, and SnAC. The advantages of these new bonds were reported by Zubir and Tian: the formation of synergistic effects within the architecture as well as the promotion of electrochemical properties such as conductivity and electrode stability, as the unique features of the newly emerged.\(^{1,24}\) The Gaussian fit of the N 1s spectrum revealed the two different peaks at binding energies of 399.8 eV and 401.8 eV, corresponding to the amine (–NH\(_3\)) and NH\(_3\)+ groups, respectively (Fig. 8a2 and b2). This is similar to the XPS spectra of the synthetic materials arising from aminoclay previously reported for detection of the presence of the NH\(_3\)+ functional groups.\(^{25–28}\) In addition, it is seen that the intensity of two peaks was unlike with the sample before annealing process,

change in hysteresis and pore-size distribution was attributed to the presence of Fe\(_3\)O\(_4\) NPs in the tailored pore structure of the nanocomposites. This stemmed from the integration between the two-dimensional (SnAC) and zero-dimensional (Fe\(_3\)O\(_4\) NPs) architectures within a single material. In combination with the SEM and TEM analysis results, we could conclude that the annealing process possibly facilitates more ef fective

In addition, the presence of some low-angle peaks in crystallization and contributes to the stabilization of this hybrid (respectively.

It is noted that the weight ratio of SnO\(_2\) and Fe\(_3\)O\(_4\) within this architectures was determined to be a unique morphology that can greatly facilitate electrolyte accessibility and fast diffusion of Li-ions and electrons during lithium storage.

These results are consistent with the broad peaks in the powder XRD patterns for the samples before and after annealing (Fig. 7). The diffraction patterns have six main peaks around 2θ = 30.1°, 35.5°, 43.2°, 53.5°, 57°, and 62.8° corresponding to the refection from d(220), d(311), d(400), d(422), d(511), and d(440) for the Fe\(_3\)O\(_4\) crystal (JCPDS card. no. 79-0418), respectively.\(^{18}\) However, low diffraction intensity and weak peaks were observed for the sample without annealing, because the crystallization of the Fe\(_3\)O\(_4\) phase was rather low and unstable; especially, the peak intensities also can be involved in variations in the scattering intensity of the crystal components and their arrangement in the lattice, and meanwhile, a narrow shape and clear peaks were found after annealing under Ar gas. It was clear that the annealing process possibly facilitates more effective crystallization and contributes to the stabilization of this hybrid structure. In addition, the presence of some low-angle peaks in the post-annealing SnO\(_2–Fe\(_3\)O\(_4\) was ascribed to tetragonal SnO\(_2\). It is noted that the weight ratio of SnO\(_2\) and Fe\(_3\)O\(_4\) within this hybrid structure can be calculated by peak area in XRD measurement. According to that, this ratio was about 1 : 12 \(A_{\text{SnO}_2} = 326.52, A_{\text{Fe}_3\text{O}_4} = 2670.23\).

**Fig. 6** N\(_2\) adsorption–desorption isothermal curves of: (a) SnAC–Fe\(_2\)O\(_4\) and (b) SnAC after annealing process under Ar gas.

**Fig. 7** X-ray diffraction (XRD) and Fourier-transform infrared spectroscopy (FT-IR) spectra for SnAC–Fe\(_2\)O\(_4\): (a) before and (b) after annealing process under Ar gas.

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**References:**

1. H. Tian: the formation of synergistic e

2. R. Zubir and Tian: the formation of synergistic effects within the architecture as well as the promotion of electrochemical properties such as conductivity and electrode stability, as the unique features of the newly emerged.\(^{1,24}\)

3. The Gaussian fit of the N 1s spectrum revealed the two different peaks at binding energies of 399.8 eV and 401.8 eV, corresponding to the amine (–NH\(_3\)) and NH\(_3\)+ groups, respectively (Fig. 8a2 and b2). This is similar to the XPS spectra of the synthetic materials arising from aminoclay previously reported for detection of the presence of the NH\(_3\)+ functional groups.\(^{25–28}\) In addition, it is seen that the intensity of two peaks was unlike with the sample before annealing process,
particularly for NH$_3^+$ peak. It is suggested that SnAC surface was carbonated with the degradation of amino groups, as well as partial removal of hydroxyl group on their surface for deprotonation of NH$_3^+$ groups, compared to the initial sample. Together with that, the presence of –NH$_2$ peak became clearer, although their intensity was rather low. Interestingly, these results are completely associated with the previous FT-IR finding.

Next, the voltage profiles of the SnAC–Fe$_3$O$_4$ nanocomposites at the SnAC ratio of 0.5 were systematically evaluated by galvanostatic discharge and charge measurements for the initial three cycles at the current rate of 100 mA g$^{-1}$ in the voltage range of 0.01–3.0 V, as plotted (Fig. 9a). It can be seen that these SnAC–Fe$_3$O$_4$ nanocomposite electrodes delivered a very high lithium storage capacity of 980 mA h g$^{-1}$ during the initial discharge process, and that a reversible capacity of 830 mA h g$^{-1}$ was achieved, as correspondent to a coulombic efficiency (CE) of about 84%. Also, the initial discharge and charge capacities of the (1.0 and 1.5) SnAC–Fe$_3$O$_4$ nanocomposite electrodes were as high as 1100 mA h g$^{-1}$ and 1300 mA h g$^{-1}$, respectively, which could be observed of cycling performance plotted (Fig. S2a and b†). The CE was determined to be about 83% for both of the two SnAC mass ratios. For the sample at the ratio of 0.5, the irreversible capacity of about 150 mA h g$^{-1}$ was attributed to the decomposition of the electrolyte on the surface of the electrodes, which resulted in SEI formation and a ubiquitous chemical reaction among heterogeneous atoms. It was noted that the CE values for all three electrodes were very high, compared with most traditional electrodes (range: 60–75%). This enhancement could have been due to the unique properties of the structure and functional groups on the surfaces of the electrode materials, which had arisen from the combination of the Fe$_3$O$_4$ NPs and the SnAC matrix. The CE of the third cycle was increased by approximately 99% and sustained thereafter.

The voltammogram (CV) curves of the SnAC–Fe$_3$O$_4$ at the ratio of 0.5 were measured under voltages ranging from 0.01 to 3.0 versus Li$^+$/Li at a scan rate of 0.2 mV s$^{-1}$ (Fig. 9b). In the first discharge cycle, a broad cathodic peak at around 0.68 V could be observed, which was attributed mainly to the irreversible formation of SEI film on the surface of the SnO$_2$ and Fe$_3$O$_4$ due to the irreversible reaction with electrolyte (Li$^+$ + e$^-$ + electrolyte $\rightarrow$ SEI). After the first cycle, the curves were more consistent. For the 2nd and 3rd cycles, this broad reduction peak became broader and positively right-shifted to around 0.90 V, which was ascribed to the conversion of Fe$_3$O$_4$ to Fe and of SnO$_2$ to SnO$_x$ (Fe$_3$O$_4$ + 8Li$^+$ + 8e$^-$ $\rightarrow$ 3FeO + 4Li$_2$O, SnO$_2$ + 4Li$^+$ + 4e$^-$ $\rightarrow$ Sn + 2Li$_2$O, Sn + xLi$^+$ + xe$^-$ $\rightarrow$ Li$_x$Sn). In the oxidation scans, wide peaks at around 1.55 V and 1.82 V might have corresponded to the reversible oxidation of Fe$^{0}$ to Fe$^{3+}$ as well as the transition of Li$_2$O to SnO$_2$. More importantly, the gradual decrease in the reduction peaks occurring during the subsequent cycles reflected not only the slow decrease in capacity.
during the charge/discharge process but also the formation of a stable SEI film involving the overlapping of peaks in various cycles.\textsuperscript{22}

To highlight the superiority of the synthesized electrodes for anode materials, we tested the cycling performance of all three at various SnAC mass ratios. The obtained results are plotted (Fig. 9c, S2a and b†). After 100 cycles, the SnACFe\textsubscript{0.5} nanocomposite electrodes retained reversible capacities of 710 mA h g\textsuperscript{-1}, 550 mA h g\textsuperscript{-1}, and 500 mA h g\textsuperscript{-1}, corresponding to 89%, 63% and 50% of the values in the second cycle, respectively. The sample at the ratio of 0.5 clearly presented the most promising results. Additionally, as indicated in Fig. 9c, S2a and b† better stability in electrical performance was observed over the course of the repeated charge/discharge process at this ratio. It is noted that the capacity values of SnAC–Fe\textsubscript{0.5} nanocomposite electrode were somewhat fluctuated (Fig. 9c). It showed an initial drop in capacity up to 10 cycles. According to previous reports, this is a common phenomenon arising from tradition metal oxide electrodes for LIBs.\textsuperscript{71–73} Such a capacity fluctuation is considered to be closely related to the complex surface chemistry and electrochemical behavior of SnAC and Fe\textsubscript{0.5} NPs at the interface of the electrolyte and electrode. The decrease of capacity begins from surface side reactions, in particular, the formation of the instable SEI film and the electrolyte decomposition at first cycles. Also, the pulverization of original NPs and the formation of new metal nanoparticles, namely Sn NPs during Li insertion/extraction process lead to loss of electrical connectivity between neighboring particles.\textsuperscript{74} Whereas, after 50 cycles, it seems to be decreased in capacity and a bit increased again. The increase of capacity is due to the activation and establish processes of the Li-ion diffusion paths in the electrode structure. Besides, the formation of in situ small Sn NPs is cause for the reversible conversion reaction, SnO\textsubscript{2} + 4Li\textsuperscript{+} + 4e\textsuperscript{-} → 2Li\textsubscript{2}O + Sn in the SnAC–Fe\textsubscript{0.5} nanocomposite, which is usually reported to be irreversible in most previous reports. By contrast, as seen in Fig. S2c† the cycling profiles of the pure-SnAC-based anode showed high initial charge/discharge capacities of about 1200 and 1650 mA h g\textsuperscript{-1}, which yielded a CE of ~73%. However, this value deteriorated significantly after just 15 cycles, and losing 25% of the initial reversible capacity. The capacity retention after the 30\textsuperscript{th} cycle, moreover, was much inferior to that of the SnAC–Fe\textsubscript{0.5} nanocomposites. Besides, the CE value of the pure-SnAC electrode reached only 70% in the first cycle, and immediately after 10 cycles, it had increased to only about 92%.

In order to clarify the differenced in electrochemical performance and electrode kinetics among the SnAC–Fe\textsubscript{0.5} nanocomposites at the various mass ratios, electrochemical impedance spectroscopy (EIS) was carried out at frequencies ranging from 100 to 0.01 Hz after 5 cycles, and it was fitted using a model inserted (Fig. 9d). Each Nyquist plot obtained for the SnAC–Fe\textsubscript{0.5} nanocomposites at the various mass ratios (Fig. 9d) indicated the common features of a single depressed semicircle in the high-medium frequency region and was followed by a straight line in the low frequency region, respectively. The single semicircle was assigned to the combination of the surface film resistance (R\textsubscript{f}) and charge-transfer resistance (R\textsubscript{ct}), and W was the Warburg impedance indicated by the straight line reflected the ionic diffusion. Evidently, the diameter of the single semicircle of the SnAC–Fe\textsubscript{0.5} nanocomposites at the ratios of 1 and 1.5 was much higher than that of SnAC–Fe\textsubscript{0.5} at the ratio of 0.5. For instance, the surface film resistances (R\textsubscript{f}) were ~1.5 Ω, ~2.1 Ω, and ~1.7 Ω for SnAC–Fe\textsubscript{0.5} with ratios of 0.5, 1.0, and 1.5, respectively. The charge transfer resistances (R\textsubscript{ct}) were ~162.8 Ω and ~154.9 Ω for SnAC–Fe\textsubscript{0.5} nanocomposite with ratios of 1.0 and 1.5, respectively, while SnAC–Fe\textsubscript{0.5} nanocomposite with ratios of 0.5 was ~122.8 Ω. It means that the SnAC–Fe\textsubscript{0.5} nanocomposite electrode at the ratio of 0.5 possessed the lowest contact and charge-transfer impedances, which fact could facilitate faster electron transport owing to the favorable dispersed state, a more stable SEI film as well as an active electrode surface with NH\textsubscript{3}\textsuperscript{+} groups during the electrochemical lithium insertion/extraction reaction resulting in significant improvement in the rate performance. It was observed that the low-frequency tails for the three samples also were different, which could be compared qualitatively with the mass-transfer kinetics in the electrode materials via the low-frequency slope angle. Indeed, the slope angle was ~46° for the SnAC–Fe\textsubscript{0.5} nanocomposite electrode at the ratio of 0.5, the highest value among the electrodes (~35° at ratio of 1; ~29° at ratio of 1.5), and the steeper low-frequency tail showed higher Li-ion diffusivity, higher electrical conductivity and a more rapid charge-transfer process in the electrode material.

To analyse fast recharge characteristic for the best electrode showing good cyclability as well as the lowest resistance, rate cyclic performance was also conducted as shown (Fig. 9e). The SnAC–Fe\textsubscript{0.5} (0.5) electrode exhibited an outstanding rate capability. It delivered specific capacities of 848 mA h g\textsuperscript{-1} at 100 mA g\textsuperscript{-1}, 715 mA h g\textsuperscript{-1} at 300 mA g\textsuperscript{-1}, 636 mA h g\textsuperscript{-1} at 500 mA g\textsuperscript{-1}, 552 mA h g\textsuperscript{-1} at 1 A g\textsuperscript{-1}, and 490 mA h g\textsuperscript{-1} at 1.5 A g\textsuperscript{-1}, which correspond to capacity retentions of 100%, 84%, 75%, 65.15%, and 58%, respectively. When the current density returned to 100 mA g\textsuperscript{-1}, the charge capacity was recovered to 706 mA h g\textsuperscript{-1}, demonstrating a capacity restoration of 83% and indicating high reversibility with structural stability. The stable cyclability and high rate capability of the SnAC–Fe\textsubscript{0.5} (0.5) electrode can be attributed to the formation of a thin and stable SEI layer along with a stable contact with the current collector, leading to better reaction kinetics at the electrode surface.

**Discussion**

The main paragraph text follows directly on here. During the operation of the LIBs, the adsorption of electrolyte, the side reactions at the solid–liquid interface as well as the exposed surface area of the electrode with electrolyte are the crucial factors in maintaining specific capacity and performance. When the electrode materials show high electrolyte accessibility, minimization in mass transfer, electron transformation, and ion dispersion are possible. The electrolyte of Li-ion batteries contains large amounts of organic solvents such as ethylene carbonate (EC), diethyl carbonate (DEC), and ethylmethyl carbonate (EMC) with both high-dielectric and low-
viscosity ingredients. With high polarity and dielectric constant, it has a high binding energy with Li' ions. This high binding energy is useful for formation of the SEI on the anode; however, this leads to loss of Li' for side reactions on the surface electrode, which causes capacity fading and low CE and DEC. A uniform layered structure and the presence of NH₃⁺ functional groups within the electrode architecture can offer a higher affinity to C—O groups in EC and DEC for reduced interaction among EC, DEC and Li' (Fig. 10). Leung and Budzien et al. demonstrated that EC and DEC decomposition are more inclined in the presence of oxygen hydroxyl termination, with the result that the NH₃⁺ groups can minimize the decomposition in the electrolyte. A high affinity with EC and DEC and its decomposition products help to control the stability of chemically and mechanically stable SEI, offering an ideal SEI film with restrained dendrite growth and surface roughening and also reducing unexpected side reactions at the SEI/electrolyte interface. In the present study, after annealing under Ar gas, almost all of the NH₃⁺ groups had been removed, though the IR results indicated the existence of a small portion of this group on the surface of Fe₃O₄–SnAC nanocomposites. This also was demonstrated by the CV curves of the Fe₃O₄–SnAC nanocomposites electrode, as discussed above. It was noteworthy that after the first cycle, the curves of voltage–current almost overlapped, which indicated that a stable SEI film had formed on the electrode surface and at the interfaces of the Fe₃O₄ and SnO₂ NPs. This stability led to effective prevention of direct exposure among the Fe₃O₄ NPs, SnO₂ NPs and electrolyte, as well as great protection for the Fe₃O₄ and SnO₂ NPs from volume expansion and structural changes during the charge/discharge process. Additionally, the dispersion and high interaction between the electrode and electrolyte owing to their inherent functional groups enhanced the wettability in the above electrode, resulting in enhanced distribution, and transport of electrons/Li-ions and the stable, superior reversibility of the SnAC–Fe₃O₄ nanocomposites.

The high electrochemical performance of the SnAC–Fe₃O₄ nanocomposites, compared with SnAC can be attributed to the unique hybrid design arising from the uniform dispersion of Fe₃O₄ NPs (zero-dimensional nanomaterials (0D NMs)) on the surface SnAC (two-dimensional nanomaterials (2D NMs)) as well as the carbonization of the SnAC structure during annealing treatment. Firstly, in this structure, the SnAC matrix is the major factor improving conductivity and ion/electron transportability; specifically, it facilitates easy penetration of ions and their subsequent reaction with active materials. Secondly, SnAC also facilitates the accommodation of huge volume change, as well as the formation of stable SEI film on the surface of active materials, and effectively prevents the aggregation of Fe₃O₄ NPs. Thirdly, the strong interaction between the Fe₃O₄ NPs and SnAC matrix owing to the inherent NH₃⁺ functional groups on the SnAC surface is the key to the obtained synergistic effects and the promotion of the efficacious electrochemical properties. Fourthly and finally, the constructed macro-porous architecture in this hybrid design provides additional conductive channels for high-efficiency electron transport and ion diffusion, which can shorten the pathway and facilitate movement as well as contact with the ions' electroactive surfaces, resulting in enhanced kinetics and CE. Based on the obtained cycling performance results for various mass ratios of SnAC in the SnAC–Fe₃O₄ nanocomposites, the performance values of the nanocomposites were remarkably changed. To explain that, it can be considered that at the ratio of 0.5, the correspondingly small amount of SnAC is enough to prevent Fe₃O₄ NPs from severe aggregation and formation of large clusters, and guarantees structural integrity during the charge/discharge process under mechanical stress from the cell assembly and huge tension. It seems to the issue of limited electron and ion transport have addressed, herein, the SnAC matrix and Fe₃O₄ NPs with a suitable ratio act to be an effective hybrid architecture with high conductivity, great structural flexibility, and negligible potential barrier under repeated charge/discharge process. By contrast, a too-large amount of SnAC (ratios of 1 and 1.5) can lead to increased electrode thickness, which tends to prevent electron and ion transport and to limit mass transfer into the electrolyte. On the other hand, after annealing treatment in Ar gas, the carbonization of the SnAC structure leads to the formation of the SnO₂ phase, which is an ideal active material [similar to Fe₃O₄ NPs] for improved electrochemical performance. Furthermore, carbonized SnAC enables the formation of a protective shell, which assists the prevention of inner SnO₂ from direct interaction with the electrolyte.

Finally, as evidenced in the XPS data, the formation of Fe–O–C and Sn–O–C chemical bonds after annealing treatment enhances the connection of Fe₃O₄ NPs and SnAC, leading thereby to improved electrical conductivity, specific capacities, capacity retention and long-term stability of the SnAC–Fe₃O₄ nanocomposites. Instead of the presence of weak interactions such as hydrogen bonds, van der Waals force or electrostatic force, the formation of new bonds can protect the active materials from their own detachment and loss during charge/discharge process, which result corresponds to the data obtained by Tian and Zubir. These authors found that the presence of strong interfacial interactions brings out a positive effect on the hybrid
architecture. The facile electronic transport and strong attachment of NPs to the matrix surface leads to easy binding/difficult dissociation characteristics and facilitates fast electron hopping, thus promoting reversible lithiation and delithiation of NPs at the higher specific current. Furthermore, as the peak intensity of Sn-O-C and Fe-O-C increases, the annealing process assists to remove some oxygen containing functional groups (C=O, Sn-O, Fe-O), which action considered to enhance electron transfer and cycling stability in SnAC-Fe3O4 nanocomposites.

To further confirm the advantages of this promising electrode, a rough comparison of the previous reports in the literature for Fe3O4-based electrodes is provided in Table 1. It can easily be observed that even when the average discharge capacity at the first cycle was 980 mA h g⁻¹, which is lower than the capacity of pure-Fe3O4 NPs and other potential materials. However, the synthesis electrode in our report displayed the highest CE value, approximately 84%, which was increased to ~99% thereafter. Furthermore, it is noteworthy that, after the repeated charge/discharge process, a discharge capacity of 710 mA h g⁻¹ could be resumed and maintained for 100 cycles. Apparently, these outstanding results show the high potential for the development of an advanced anode material that can attain higher energy and power density and offers ease of fabrication and enhanced safety as well.

**Experimental**

**Chemicals**

Tin(i) chloride dihydrate (SnCl2·2H2O, MW = 225.65, 98%), APTES (3-aminopropyltriethoxysilane, H2N(CH2)3Si(OC2H5)3, MW = 213.37, ≥98%), iron(II) chloride hexahydrate (FeCl2·6H2O), iron(III) chloride tetrahydrate (FeCl3·4H2O), and NaOH pellets (sodium hydroxide, MW = 40, reagents chemicals) were obtained from Daejung Company (Gyeonggi-do, Korea), and bulk ethanol (EtOH) (MW = 46.06, 99.5%, 18 L) was acquired from Samchun Pure Chemicals (Gyeonggi-do, Korea). Additionally, double distilled (DI) water was obtained from Milli-Q® (resistance = 18.2 MΩ cm at 25 °C).

**Preparation of Sn-aminoclay (SnAC)–Fe3O4 nanocomposites**

In step 1, SnAC was synthesized via a simple method available in the literature. Briefly, a mixture of 4.5 g SnCl2·2H2O and 100 mL EtOH was prepared in a 500 mL glass beaker and fully dissolved within 15 min. APTES was added dropwise and stirred overnight at room temperature to obtain an aqueous dispersion of white precipitate. The white precipitate was separated by centrifugation and washed three times with EtOH prior drying at 60 °C for 24 hours.

In step 2, a mixture of 0.5 g, 0.1 g, and 1.5 g of SnCl2·4H2O (0.7 g) and (1.9 g) FeCl2·6H2O was prepared and stirred for 30 min, giving rise to a homogeneous mixture. To this, an appropriate amount of NaOH (10 M) solution was continuously added. The mixture was stirred for a further 24 hours at room temperature to obtain black-precipitate Fe3O4 NPs.

In step 3, after dissolving SnAC (0.5 g, 1.0 g, and 1.5 g) in DI-water (40 mL), Fe3O4 NPs were added. The resulting mixture was continuously stirred for 24 hours. The black precipitate was separated and washed three times with EtOH and DI-water to remove impurities, and then the final product was oven-dried at 60 °C for 24 hours. The dried solids were sealed under vacuum using a u-pump. Then, the end product was annealed at 500 °C for 3 hours under Ar gas.

**Materials characterization**

The samples' X-ray diffraction (XRD) patterns were acquired with a RIGAKU, D/MAX-2500 powder diffractometer equipped with a Cu Kα radiation source (λ = 1.541 Å) operating at 40 kV and 300 mA. The structure and surface characteristics of the samples were investigated under transmission electron microscopy (TEM, 200 kV, Tecnai F20, Philips) and scanning electron microscopy (SEM-4700) equipped with an energy-dispersive X-ray spectrometer (EDX). The morphological analysis results, chemical compositions and binding energies were confirmed by Fourier transform infrared spectrophotometry (FT-IR, 4100, Jasco, Japan) and X-ray photoelectron spectroscopy (XPS, K-alpha, Thermo VG Scientific). Fourier-transform infrared spectrophotometry (FT-IR, 4100, Jasco, Japan) was utilized to examine the vibration peaks of the SnAC–Fe3O4 nanocomposites before and after the annealing process. The Brunauer–Emmett–Teller (BET) surface area and average pore diameter were obtained from N2 adsorption/desorption isotherms using a fully automatic physisorption analyzer (ASAP 2020, Tristar). Also, the zeta potentials of the SnAC sample, Fe3O4 NPs, and SnAC–Fe3O4 nanocomposites were obtained by dynamic laser-light scattering (DLS, Zeta PALS).

<table>
<thead>
<tr>
<th>Electrode</th>
<th>Potential range (V)</th>
<th>Current density (mA g⁻¹)</th>
<th>Initial coulombic efficiency (%)</th>
<th>Discharge capacitance (mAh g⁻¹)</th>
<th>Capacity retention/cycling (%)</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>Bare Fe3O4 NPs</td>
<td>0.01–3.0</td>
<td>100</td>
<td>82</td>
<td>832</td>
<td>27% (34)</td>
<td>37</td>
</tr>
<tr>
<td>CNT/Fe3O4</td>
<td>0.02–3.0</td>
<td>100</td>
<td>68</td>
<td>1035</td>
<td>63.4% (145)</td>
<td>38</td>
</tr>
<tr>
<td>G/Fe3O4</td>
<td>0.01–3.0</td>
<td>100</td>
<td>78</td>
<td>1060</td>
<td>98% (83)</td>
<td>37</td>
</tr>
<tr>
<td>Fe3O4/C core–shell</td>
<td>0.01–3.0</td>
<td>0.1C</td>
<td>71.7</td>
<td>807</td>
<td>49% (100)</td>
<td>39</td>
</tr>
<tr>
<td>Carbon nanorods/Fe3O4</td>
<td>0.005–3.0</td>
<td>1C</td>
<td>1085</td>
<td>40% (100)</td>
<td>40</td>
<td></td>
</tr>
<tr>
<td>SnO2–Fe3O4@G</td>
<td>0.005–3.0</td>
<td>100</td>
<td>69.9</td>
<td>1312.8</td>
<td>75.4% (50)</td>
<td>41</td>
</tr>
<tr>
<td>C-Fe3O4 nanospindles</td>
<td>0.01–3.0</td>
<td>100</td>
<td>84</td>
<td>~1000</td>
<td>70.1% (80)</td>
<td>42</td>
</tr>
<tr>
<td>SnAC–Fe3O4</td>
<td>0.01–3.0</td>
<td>100</td>
<td>84</td>
<td>980</td>
<td>89% (100)</td>
<td>This work</td>
</tr>
</tbody>
</table>
Electrochemical measurement

Electrochemical performance was investigated using CR2032 coin cells that had been assembled in an Ar-filled glove box. The working electrode was fabricated by mixing 70 wt% active material (SnAC–Fe3O4, 0.5, 1.0, and 1.5, pure SnAC), 15 wt% carbon black as a conductive medium and a 15 wt% polyvinylidene fluoride (PVDF) binder in N-methyl-2-pyrrolidone (NMP) to form a slurry. The resultant slurry was uniformly cast on Cu foil with a doctor blade, and subsequently dried at 70 °C in a vacuum oven overnight. The mass loading of the active material was typically 0.7–1.0 mg cm−2. The electrochemical properties of the electrodes were characterized at room temperature. Li foil was used as the counter electrode; polyethylene (PE) was used as separator; the electrolyte was a 1 M LiPF6 in a 1:1 v/v mixture of ethylene carbonate (EC) and diethyl carbonate (DEC). A cyclability test was performed on a Nanocycler V1.34 system in the 0.01–3.0 V voltage range. Cyclic voltammetry and electrochemical impedance spectroscopic (EIS) analysis were performed using a ZIVE MP1 station in the 0.01–3.0 V voltage range versus Li/Li+. All of the calculations were based on the mass of the active material of the anode.

Conclusions

In summary, this work demonstrates the structural and morphological relationships of SnAC–Fe3O4 nanocomposites at various SnAC mass ratios. These homogeneously dispersed and aggregation-less Fe3O4 NPs on the SnAC matrix are promising candidates for anode material in LIBs. After annealing treatment, carbonization of the SnAC structure leads to the formation of a SnO2 phase along with new chemical bonds among SnO2, the Fe3O4 NPs, and SnAC. Moreover, the great affinity between the electrode and electrolyte owing to the remaining NH3+ functional groups on the electrode surface leads to the formation of a stable macro-architecture. The constructed macro-architecture offers unique electrochemical properties due to its physical and chemical merits and surface properties as well as the synergistic effects of both components. More importantly, this structure can withstand the huge stresses caused by Fe3O4 expansion/extraction and preserve the structural integrity of the electrode over the course of repeated charge/discharge process. As a result, the nanocomposite electrodes exhibited a high reversible specific capacity of 710 mA h g−1 in the 100th cycle at 100 mA g−1, corresponding to a high electrochemical performance of 89%.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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