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A citric acid-assisted deposition strategy to synthesize mesoporous SiO₂-confined highly dispersed LaMnO₃ perovskite nanoparticles for *n*-butylamine catalytic oxidation†

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Catalytic oxidation can efficiently eliminate nitrogen-containing volatile organic compounds (NVOCs) and suppress the generation of toxic NO_x in order to avoid secondary pollution. In this study, mesoporous SiO₂-confined LaMnO₃ perovskite nanoparticles with high dispersion were successfully prepared by a citric acid-assisted deposition method (LMO/SiO₂-SD) and tested for the oxidation of *n*-butylamine. The method utilized the synergistic effect of abundant active hydroxyl groups existing on the SiO₂ gel surface and citric acid, rendering the metal ions more uniformly scattered on the SiO₂ surface. Strikingly, the LMO/SiO₂-SD sample exhibited the optimum catalytic performance (*T*₉₀ at 246 °C) and the highest N₂ selectivity, which was mainly ascribed to its abundant surface acid sites, superior low-temperature reducibility and higher ratio of surface Mn⁴⁺ species. The apparent activation energy (*E*_a) for *n*-butylamine oxidation over LMO/SiO₂-SD sample was 29.0 kJ mol⁻¹. Furthermore, the reaction mechanism of *n*-butylamine oxidation was investigated by *in situ* FITR and a reasonable reaction route for *n*-butylamine oxidation over the LMO/SiO₂-SD sample was proposed.

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1 Introduction

Nitrogen-containing volatile organic compounds (NVOCs) have caused serious impacts on air pollution and public health, and have being a focus of attention in the field of air pollution prevention and treatment. The varied sources and varieties of NVOCs lead to disparate degrees of influence due to their different toxic, mutagenic and carcinogenic properties. Examples are amines, nitro compounds, nitrile and so on, which are odorous, highly toxic and would cause NO_x secondary pollution and other problems by improper disposal.¹⁻⁵ Compared to various developed technologies for NVOCs elimination,⁶ the catalytic oxidation method has always been judged as a highly efficient and promising alternative technology due to its superior advantages, such as lower temperature, lower energy consumption, no secondary pollution and so on. Within this eco-friendly method, NVOCs would be utterly decomposed into CO₂, H₂O and N₂ at a relatively low reaction temperature. In

addition, the yield of toxic product NO_x would be strictly controlled in order to avoid secondary pollution.

With regard to the catalysts applied to NVOCs oxidation, precious metal (Pd, Ag, *etc.*) and transition metal (Cr, Mn, *etc.*) catalysts are commonly investigated by researchers due to their optimum catalytic performances.⁵⁻⁷ Although noble metal catalysts exhibited superior catalytic activities, toxicity resistance and better regeneration, their expensive cost, poor high thermal stability and scarce resources limit the large-scale application in industrial production. Perovskite-type metal oxides have always been considered as an alternative for noble metal catalysts due to their unique catalytic performances and good thermal stability for catalytic oxidation of VOCs.⁸⁻¹² The application of conventional perovskites in catalytic oxidation were nonetheless currently limited due to low specific surface area and serious accumulation due to calcination of precursor material at high temperature. To improve the catalytic performances of perovskites, enormous methods were introduced, such as fabricating three-dimensionally ordered macroporous structure by PMMA-templating approach; dispersing nano-sized perovskites onto metal oxides support with larger specific surface area and so on.¹³⁻¹⁶ Numerous articles have substantiated that the supported nanoscale perovskites would enhance catalytic performances, but the supported perovskites prepared by conventional methods were poorly dispersed and crystallized, which would have an influence on stability and

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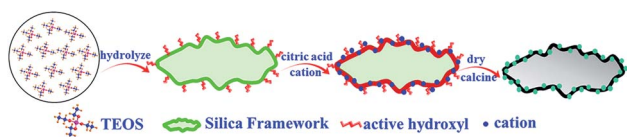
catalytic performances.^{17–19} The above triggered our interesting in the study of optimization of perovskite catalysts. In this work, mesoporous SiO₂-confined perovskite nanoparticles with high dispersion was prepared by a novel avenue (citric acid-assisted deposition method). The figure shown in Scheme 1 depicts the facile preparation of the mesoporous SiO₂ support and the ensuing synthesis of the LMO/SiO₂-SD catalyst. The TEOS was firstly hydrolyzed to form SiO₂ gel, and subsequently citric acid was coated on the surface of SiO₂ gel. Ultimately, the synergistic effect of abundant active hydroxyl groups existed on SiO₂ gel surface and citric acid, rendering the metal ions better and uniformly scattered on the SiO₂ surface. As a contrast, LMO/SiO₂ samples were also prepared by conventional citrate complexation impregnation and wetness impregnation methods, respectively. Catalytic performances were measured using the conversion of *n*-butylamine as a probe reaction for NVOs abatement. Moreover, amounts of characterization techniques containing XRD, SEM, TEM, N₂ sorption, NH₃-TPD, H₂-TPR, O₂-TPD, XPS and *in situ* FTIR were employed to establish a constructive relationship between physicochemical properties and catalytic performances. Finally, a possible and reasonable catalytic reaction mechanism for *n*-butylamine oxidation was proposed.

2 Experimental

2.1 Catalyst preparation

Synthesis of SiO₂ support. In a typical preparation, the stoichiometric tetraethyl orthosilicate (TEOS), glacial acetic acid and ethanol were dissolved in distilled water. After homogeneous mixing, 9.0 g methanamide was added and stirring for 0.5 h, and then 70 g epoxypropane was added. The formed solution was heated at 60 °C with circulation reflux and aged overnight. The obtained sample was dried 100 °C and then calcined in a muffle furnace at 800 °C for 4 h.

Synthesis of LMO/SiO₂-SD catalyst. In a typical preparation, the stoichiometric tetraethyl orthosilicate (TEOS), glacial acetic acid and ethanol were dissolved in distilled water. After homogeneous mixing, 9.0 g methanamide was added and stirring for 0.5 h, and then 70 g epoxypropane was added. The formed solution was heated at 60 °C with circulation reflux and aged overnight. The mixture of stoichiometric citric acid and distilled water was gradually added to the above solution, and then stirred 2 h. Subsequently, the solution of nitrate containing La(NO₃)₃·6H₂O and Mn(NO₃)₂ was added, and the formed mixture stirred for another 2 h to obtain viscous gel. The obtained viscous gel was dried at 100 °C and then calcined in a muffle furnace at 800 °C for 6 h.



Scheme 1 Illustration of the preparation process of LMO/SiO₂-SD catalyst.

Synthesis of LMO/SiO₂-CI catalyst. In a typical preparation, the stoichiometric La(NO₃)₃·6H₂O, Mn(NO₃)₂ and citric acid were dissolved in distilled water. After homogeneous mixing, the support SiO₂ were added and stirred for 6 h at 80 °C to form viscous gel. The obtained viscous gel was dried 100 °C and then calcined in a muffle furnace at 800 °C for 6 h.

Synthesis of LMO/SiO₂-WI catalyst. In a typical preparation, the stoichiometric La(NO₃)₃·6H₂O and Mn(NO₃)₂ used as precursors were dissolved in distilled water. After homogeneous mixing, the support SiO₂ were added and stayed for 12 h. The obtained precipitate dried at 100 °C and then calcined in a muffle furnace at 800 °C for 6 h.

2.2 Catalyst characterization

X-ray diffraction (XRD) patterns were recorded on a SmartLab X-ray Diffractometer (Rigaku Corporation, Japan) equipped with Cu K α radiation ($\lambda = 0.1541$ nm) within 2θ range of 10–80° at the scanning step of 0.02°. Scanning electron microscopy (SEM) images were analyzed with a Hitachi S-3000 N scanning electron microscope. The high resolution TEM images (HRTEM) were recorded using a JEOL JEM-2100 electron microscope. The nitrogen adsorption and desorption isotherms were carried out on BEL SORP II apparatus. Before adsorption, all LMO/SiO₂ samples were outgassed in vacuum at 200 °C. The surface area was determined by the BET theory and the pore size was calculated by the BJH method. Temperature-programmed desorption of NH₃ (NH₃-TPD) was investigated by AutoChemII 2920 apparatus. Prior to each test, the catalyst (50 mg, 40–60 meshes) was pretreated under high purified He at 200 °C for 1 h, then cooled down to 50 °C, and turned the flow of 10 vol% NH₃/He into the system with a flow rate of 30 mL min⁻¹ for 1 h. After that, the catalyst was flushed with He to remove physically adsorbed NH₃ on the catalyst surface. Finally, the system was heated to 500 °C at a heating rate of 10 °C min⁻¹, and the desorption profiles were recorded under He flow. Temperature-programmed reduction of H₂ (H₂-TPR) was carried on the same instrument. Firstly, the catalysts were pretreated under Ar flow at 200 °C for 1 h. Then the temperature was cooled to 50 °C, the catalyst was heated from 50 °C to 900 °C under 10 vol% H₂/Ar flow. The consumption of H₂ was detected by thermal conductivity detector (TSD). Temperature programmed desorption of O₂ (O₂-TPD) experiment was also determined by AutoChemII 2920 analyzer. After pretreated 50 mg of catalyst at 200 °C for 1 h under He flow, the sample was exposed to 4 vol% O₂/He at 50 °C for 1 h. After the catalyst was purged with helium until stabilization of the instrument baseline, the sample was heated (10 °C min⁻¹) to 800 °C. The amount of desorbed ammonia was measured using a thermal conductivity detector (TSD). The X-ray photoelectron spectroscopy (XPS) spectra were performed using a PHI 5000 Versa Probe high performance electron spectrometer. All binding energies (BE) were calibrated using C 1s (BE = 284.8 eV) as a standard.

2.3 *In situ* FTIR studies

In situ FTIR spectra were recorded by IS50 infrared spectrometer equipped with an *in situ* diffuse reflectance pool containing



CaF₂ window and MCT detector. The real reaction temperature was controlled by the programmed temperature controller precisely. All spectra were recorded by accumulating 32 scans at a spectra resolution of 4 cm⁻¹. Prior to data collection, the catalyst would be pretreated with 20 vol% O₂/He at 300 °C for 1 h in order to remove the contaminants and then cooled down to 50 °C. In this process, spectra of the fresh catalyst surface were collected per 50 °C and employed as the background. Reaction was carried out by exposing catalyst to a stream of flow composed of 1000 ppm *n*-butylamine, 20% O₂ and He as balance gas. *In situ* FTIR spectra were collected from 50 to 300 °C for every 50 °C where a steady state was reached.

2.4 Catalytic evaluation

The measurements of catalytic activities were carried out in a fixed bed quartz reactor at atmospheric pressure for the complete oxidation of *n*-butylamine. Quantitative catalyst (1.0 g, 16–40 mesh) was diluted with 1.0 g of quartz sands and positioned at the middle of the fixed bed quartz reactor. For the measurements of catalytic performances, the total gas flow was kept at 250 mL min⁻¹, and concentrations of the feed components were controlled as follows: 1000 ppm *n*-butylamine + 20% O₂ + He (balance). The reactants and products were analyzed by VARIO industrial flue gas analyzer (MRU Corporation, Germany) and an on-line GC equipped with FID detector. The *n*-butylamine conversion (x_{NVOCs}) and the selectivity of N₂ (S_{N_2}) were defined, respectively, as:

$$x_{\text{NVOCs}} = (c_{\text{in}} - c_{\text{out}})/c_{\text{in}} \times 100\%;$$

$$S_{\text{N}_2} = Y_{\text{N}_2}/x_{\text{NVOCs}} \times 100\%$$

where, c_{in} and c_{out} represent the *n*-butylamine concentration of inlet and outlet, respectively; Y_{N_2} represent the yield of N₂.

3 Results and discussion

3.1 Catalyst characterization

XRD results. The crystal structure of LMO/SiO₂ samples were examined by XRD, and the corresponding patterns with 2θ range of 10–80° were displayed in Fig. 1. For all LMO/SiO₂ samples, the diffraction peaks at around $2\theta = 23^\circ, 33^\circ, 40^\circ, 46^\circ,$

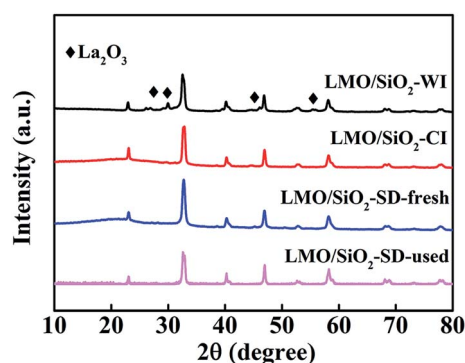


Fig. 1 XRD patterns of LMO/SiO₂ catalysts.

53°, 58°, 69° and 78° can be observed, indicating the existence of rhombohedral perovskite-type crystalline phase (PDF 50-0299). Furthermore, one can see that both LMO/SiO₂-SD and LMO/SiO₂-CI samples exhibited well single perovskite structure, and no reflections belonging to other impurities such as La₂O₃ and MnO_x crystalline phase were observed, which revealed that citric acid had the better metal ion complexing action. Differently, for LMO/SiO₂-WI sample, the coexistence of La₂O₃ phase (PDF 74-2430) and LaMnO₃ perovskite phase (PDF 50-0299) was demonstrated from the corresponding XRD patterns. La₂O₃ impurity in LMO/SiO₂-WI sample was attributed to the surface enrichment of La species without well fabricating LaMnO₃ with Mn species during the calcination process.¹²

SEM and TEM results. The spatial morphological structure of three LMO/SiO₂ samples were further examined by SEM. As shown in Fig. 2, both LMO/SiO₂-SD and LMO/SiO₂-CI samples featured a rough stacked pore structure with quantities of nanoparticles. Noticeably, the nanoparticles and channels generated of LMO/SiO₂-SD were uniform and well-distributed, which might be ascribed to the spontaneously uniform deposition arose from the interaction between abundant active hydroxyl groups existed on SiO₂ gel surface and La, Mn metal ions, forming LaMnO₃ nanoparticles homogeneously embedded in the framework of SiO₂ aerogel. For LMO/SiO₂-CI sample, majority of nanoparticles had gathered, which was consistent with its larger crystallite size. In contrast, LMO/SiO₂-WI sample exhibited a smooth sheet surface structure with only trace amount of nanoparticles.

TEM characterization was conducted out to study the structure for all LMO/SiO₂ samples. In Fig. 2(d–f), LaMnO₃ perovskite was deposited onto the SiO₂ support as larger nanoparticles of 15–30 nm diameters, which was similar microtopography to SEM results. The observed 0.387 nm, 0.221 nm and 0.234 nm lattice fringe were ascribed to the (012), (006) and (113) crystallographic planes of LaMnO₃ (PDF 50-0299), respectively. This result was in good accordance with the XRD results of LMO/SiO₂ samples. It could be explicitly from all samples that the layered sheet-like structures assigned to LaMnO₃ was decorated by the black spots. When the LMO/SiO₂ sample was prepared by citric acid-assisted deposition method, the black spots covered the surface of SiO₂ homogeneously. On the contrary, the black spots existed on the LMO/SiO₂-WI and LMO/SiO₂-CI samples obviously exhibited the pulverous aggregation and accumulation.

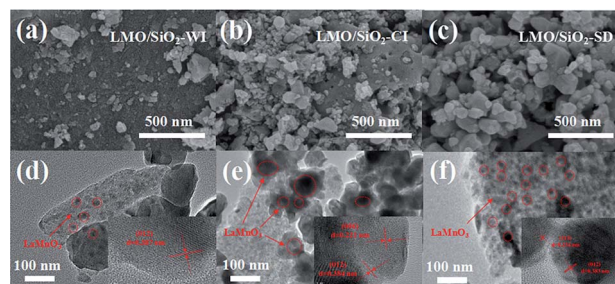


Fig. 2 SEM and TEM images of LMO/SiO₂-WI (a and d), LMO/SiO₂-CI (b and e) and LMO/SiO₂-SD (c and f).



N₂ adsorption–desorption results. The N₂ adsorption–desorption isotherms and pore-size distributions of LMO/SiO₂ samples were plotted in Fig. S1.† It could be obviously seen that the isotherms of all samples (Fig. S1(a)†) exhibited the typical IV type with an apparent type H2 hysteresis loop in the relative pressure (p/p_0) range of 0.4–0.9 as defined by IUPAC, implying the existence of mesopore structure in LMO/SiO₂ samples, which could be substantiated by Fig. S1(b).† The agglomeration or compaction of the nanoparticle building blocks led to the formation of mesopore structure of these samples. The pore size distribution curves (Fig. S1(b)†) suggested that the pore diameter of LMO/SiO₂ samples were distributed predominantly in the range of 4–5 nm, which matched well with the isotherms. The surface area of LMO/SiO₂ samples was calculated by BET equation, and the pore volume and pore diameter LMO/SiO₂ samples were obtained by BJH model (Table 1). As seen clearly in Table 1, the LMO/SiO₂-WI sample exhibited the largest surface area (300 m² g⁻¹), pore volume (0.39 cm³ g⁻¹) and pore diameter (5.03 nm), LMO/SiO₂-SD sample followed and corresponding values were (253 m² g⁻¹, 0.36 cm³ g⁻¹ and 4.31 nm), and the LMO/SiO₂-CI sample exhibited the lowest surface area (205 m² g⁻¹), pore volume (0.31 cm³ g⁻¹) and pore diameter (4.02 nm). Combining with the results of XRD, SEM and TEM, one can find that the LMO/SiO₂-SD sample might expose more active sites, which might result from the synergistic effect of the citric acid-assisted deposition method.

NH₃-TPD results. The acidity of LMO/SiO₂ samples was investigated by NH₃-TPD technique. As is well known, the acidity of samples is another essential factor to affect the catalytic performance for *n*-butylamine oxidation,^{20,21} the adsorption and activation of *n*-butylamine on the sample surface was a vital step in the reaction, meanwhile the amount of surface acid sites and acid strength would alter its adsorption. As illustrated in Fig. 3, all LMO/SiO₂ samples mainly presented two broad NH₃ desorption peaks ranged from 200 to 500 °C, assigned to the weak-strength acid sites and the medium-strength acid sites, respectively.²² Strikingly, the LMO/SiO₂-SD sample displayed higher peak-temperature of desorption peaks, about 280 and 370 °C, indicating its relatively stronger acidity. The corresponding acid amounts of LMO/SiO₂ samples were listed in Table 1. It could be easily seen that LMO/SiO₂-SD sample exhibited the most amount of weak acid sites (0.042 mmol g_{cat}⁻¹) and medium-strength acid sites (0.027 mmol g_{cat}⁻¹), LMO/SiO₂-CI sample followed and corresponding values were 0.039 mmol g_{cat}⁻¹ and 0.022 mmol g_{cat}⁻¹,

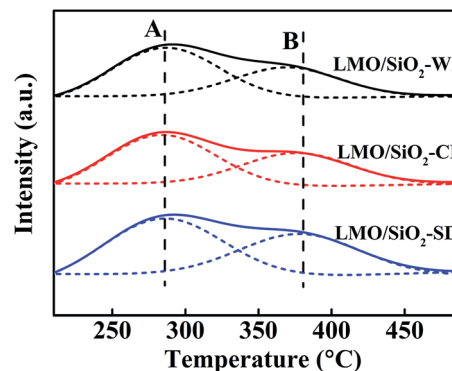


Fig. 3 NH₃-TPD curves of LMO/SiO₂ catalysts.

and the LMO/SiO₂-WI sample exhibited the lowest amount of weak acid sites (0.038 mmol g_{cat}⁻¹) and medium-strength acid sites (0.019 mmol g_{cat}⁻¹). Combined with the analyses of previous characterization results, the presence of more acidic sites on the surface could be attributed to the homogeneous and well-distributed nanoscale pore structure of LMO/SiO₂-SD sample.

Redox properties. The redox properties of all LMO/SiO₂ samples were characterized by H₂-TPR experiments (Fig. 4). Since La³⁺ is unreducible under the H₂-TPR conditions adopted in the present study, the observed reduction peaks could be assigned to the reduction of Mnⁿ⁺ species in LMO/SiO₂ samples.²³ For all LMO/SiO₂ samples, two main reduction peaks were observed, the first one within temperature range of 250–

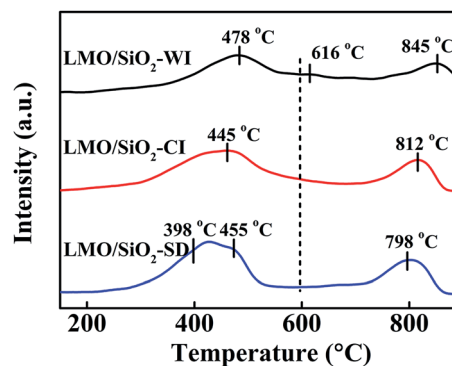


Fig. 4 H₂-TPR curves of LMO/SiO₂ catalysts.

Table 1 Physico-chemical properties of LMO/SiO₂ catalysts

Catalysts	SSA ^a (m ² g ⁻¹)	V _p ^b (cm ³ g ⁻¹)	D _p ^b (nm)	Acid amount (mmol g _{cat} ⁻¹)		H ₂ consumption (mmol g _{cat} ⁻¹)			O ₂ desorption (mmol g _{cat} ⁻¹)	
				A	B	1st peak	2nd peak	3rd peak	O _α	O _β
LMO/SiO ₂ -WI	300	0.39	5.03	0.038	0.019	0.202	0.052	0.120	0.228	0.121
LMO/SiO ₂ -CI	205	0.31	4.02	0.039	0.022	0.347	0.249	—	0.096	0.125
LMO/SiO ₂ -SD	253	0.36	4.31	0.042	0.027	0.254	0.196	0.259	0.122	0.137

^a SSA: calculated by BET method. ^b V_p and D_p: obtained by BJH method.



600 °C, which could be attributed to the reduction of Mn^{4+} into Mn^{3+} ,^{24,25} and the latter one above 600 °C were due to the reduction of Mn^{3+} into Mn^{2+} ,²⁶ whereas the reduction of Mn^{2+} into Mn^0 hardly occurred in the reduction process of MnO_x .²⁷ It could be easily seen that the reduction peak temperature of LMO/SiO₂-SD sample was significantly lower than LMO/SiO₂-CI and LMO/SiO₂-WI samples, indicating that LMO/SiO₂-SD sample exhibited an excellent reducibility behaviour among all LMO/SiO₂ samples. Meanwhile the H₂ consumption related to each reduction region calculated through quantitative integration of the corresponding H₂-TPR peaks were listed in Table 1. The H₂ consumption of LMO/SiO₂-SD sample was the highest (0.709 mmol g_{cat}⁻¹), LMO/SiO₂-CI sample followed (0.596 mmol g_{cat}⁻¹), and LMO/SiO₂-WI sample exhibited the lowest H₂ consumption (0.374 mmol g_{cat}⁻¹). It was known from the H₂-TPR experiments that the LMO/SiO₂-SD sample revealed the conspicuous low temperature redox capability compared to other LMO/SiO₂-CI and LMO/SiO₂-WI samples, which could be related to the better surface oxygen mobility of LMO/SiO₂-SD sample.²⁸

O₂-TPD measurements were performed in order to investigate the surface and bulk oxygen species as well as their mobility of the perovskite oxides and corresponding oxygen desorption curves of LMO/SiO₂ samples were depicted in Fig. 5. Apparently, all LMO/SiO₂ samples showed two main oxygen desorption regions. The first region (O_α) ranging from 200 to 550 °C could be attributed to the surface adsorbed molecular oxygen and/or oxygen species migrating *via* surface oxygen vacancies, which mainly related to the physical and chemical properties of B-site metal.^{20,29,30} Another region ranging from 750–900 °C could be assigned to the desorption of lattice oxygen, denoted as O_β,^{31–33} which mainly related to the ability of B-site metal ions to be reduced. Strikingly, the LMO/SiO₂-SD sample exhibited the lowest peak-temperature of desorption peaks, indicating its better oxygen mobility. The corresponding oxygen desorption amounts were further analyzed by quantitative integration of O₂-TPD curve (Table 1). The desorption amounts of LMO/SiO₂-SD sample were 0.122 and 0.137 mmol g_{cat}⁻¹, respectively, which were remarkably higher than that of LMO/SiO₂-CI samples (0.096 and 0.125 mmol g_{cat}⁻¹).

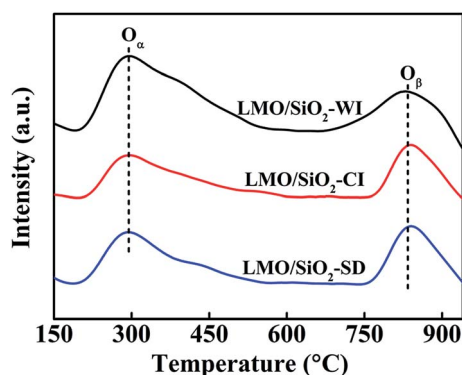


Fig. 5 O₂-TPD curves of LMO/SiO₂ catalysts.

XPS results. All LMO/SiO₂ samples were subjected to XPS characterization analysis for investigating the atomic concentrations and chemical states of major elements (Fig. 6). Two main asymmetrical peaks assigned to Mn 2p_{3/2} and Mn 2p_{1/2} of LMO/SiO₂ samples were observed, ranging from 640.8 eV to 652.7 eV. Relevant references reported that Mn⁴⁺ and Mn³⁺ species coexisted in LaMnO₃ perovskite-type samples.^{27,34–36} After a peak linear fitting deconvolution, the Mn 2p_{3/2} XPS spectra could be decomposed into two peaks, appeared at binding energies of 642.4 ± 0.2 eV and 640.8 ± 0.2 eV assignable to Mn⁴⁺ peak and Mn³⁺ peak, respectively.^{37,38} The relative percentages of manganese species were calculated by quantitative analysis of the corresponding characteristic peaks and ultimate results were listed in Table S1.† A significant distinction existed in Mn⁴⁺/(Mn⁴⁺ + Mn³⁺) ratio between the different LMO/SiO₂ samples and followed the sequence LMO/SiO₂-SD (41.64%) > LMO/SiO₂-CI (34.54%) > LMO/SiO₂-WI (29.47%). Based on which, the citric acid-assisted deposition method is conducive to obtaining a sample containing abundant surface concentration of Mn⁴⁺ species.

As previously reported,^{41,42} the surface chemisorbed oxygen (O_S) species were more active than the lattice oxygen (O_L) species due to their higher mobility, therefore O_S species played a more important role in catalytic oxidation reaction. It can be seen from Fig. 6(b), the O 1s peak could be split into two peaks: surface chemisorbed oxygen (531.2 ± 0.2 eV, denoted as O_S) and lattice oxygen (529.0 ± 0.2 eV, denoted as O_L).^{39,40} The O_S/(O_S + O_L) ratios of all LMO/SiO₂ samples were calculated by quantitative analysis of the corresponding characteristic peaks of O_S and O_L, the O_S/(O_S + O_L) ratio of LMO/SiO₂-SD sample calculated

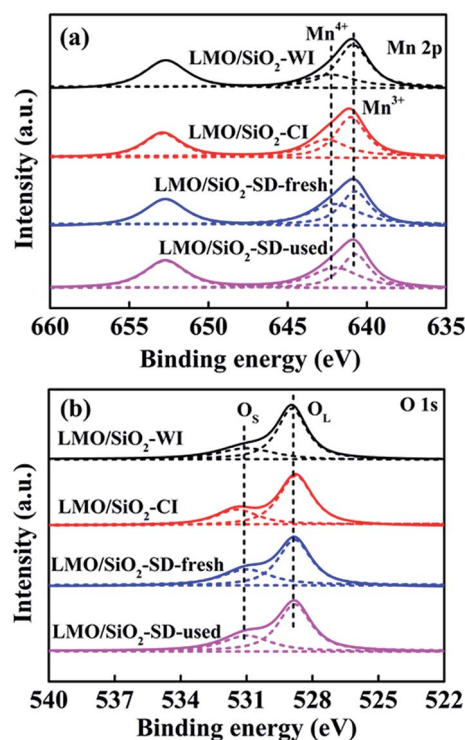


Fig. 6 (a) Mn 2p and (b) O 1s XPS spectra of LMO/SiO₂ catalysts.



from deconvoluted O 1s XPS spectrum was 38.72%, which was larger than that of LMO/SiO₂-CI (32.16%) and LMO/SiO₂-WI (27.74%) samples. Further investigating the H₂-TPR, O₂-TPD and XPS results, higher ratio of surface Mn⁴⁺ species and luxuriant O_s species of LMO/SiO₂-SD sample could be attributed to the synergic effect of the citric acid-assisted deposition method, which can contribute to the conspicuous low temperature redox capability and better surface oxygen mobility.

3.2 Catalytic evaluation

The catalytic performances of all LMO/SiO₂ samples for *n*-butylamine oxidation were tested within the temperature range from 50 to 400 °C (Fig. 7). Apparently, the conversion of *n*-butylamine showed a monotonic increasing tendency with the increase of reaction temperature, and a complete conversion of *n*-butylamine could be ultimately achieved below 400 °C over all LMO/SiO₂ samples. The sequence in terms of catalytic performance over LMO/SiO₂ samples followed by LMO/SiO₂-SD > LMO/SiO₂-CI > LMO/SiO₂-WI, revealing that SD method can be capable of improving the catalytic performance for the complete oxidation of *n*-butylamine. T₁₀, T₅₀ and T₉₀ (representing the temperature at 10%, 50% and 90% *n*-butylamine conversion) were respectively listed in Table 2. Among all LMO/SiO₂ samples, LMO/SiO₂-SD sample displayed the optimum catalytic oxidation performance for *n*-butylamine with the T₁₀, T₅₀ and T₉₀ values of 95, 188 and 246 °C, respectively, which were lower than that of LMO/SiO₂-CI (116, 206 and 273 °C) and LMO/SiO₂-WI (146, 230 and 305 °C) samples.

To deeply explore the *n*-butylamine catalytic oxidation reaction, a comprehensive identification of reaction products was

carried out. At 400 °C or above, *n*-butylamine was entirely disintegrated into CO₂, N₂, NO_x and so on. As shown in Fig. 7(b), the results of N₂ selectivity over LMO/SiO₂ samples were observed and LMO/SiO₂-SD sample displayed the highest N₂ selectivity. With the reaction temperature increasing, the N₂ selectivity over all LMO/SiO₂ samples showed an obvious decline, which could be ascribed to the more products of NO_x at higher temperature. Moreover, to highlight the optimum catalytic performance of LMO/SiO₂-SD sample for *n*-butylamine catalytic oxidation, their activities were compared based on specific activity (*R*). It was found that LMO/SiO₂-SD sample possessed higher specific activity (6.85×10^{-7} mmol m⁻² s⁻¹), indicating that more active sites existed on the unit specific surface area of sample. Based on the catalytic performances and characterization results, one can find that LMO/SiO₂-SD sample showed the optimum catalytic performance, which was assigned to its better dispersity, abundant surface acid sites, superior low-temperature reducibility, high ratio of surface Mn⁴⁺ species and more available surface adsorbed oxygen species. Among them, acid sites of LMO/SiO₂-SD sample played an essential role in reaction. By which, *n*-butylamine could be adsorbed efficiently and gone through further oxidation.

The apparent activation energy (*E*_a) for *n*-butylamine oxidation over all LMO/SiO₂ samples with a first-order Arrhenius equation: $r = -kc = (-A \exp(-E_a/RT))c$, in which *r*, *k*, *A*, *E*_a represent the reaction rate (mol s⁻¹), the rate constant (s⁻¹), the pre-exponential factor and the apparent activation energy (kJ mol⁻¹), respectively.^{36–38} The Arrhenius plots for *n*-butylamine catalytic oxidation over all the LMO/SiO₂ samples showed well-line relationship (Fig. S2†). In addition, the *E*_a value values were calculated and summarized in Table 2. As listed in Table 2, the sequence of *E*_a values followed by LMO/SiO₂-SD (29 kJ mol⁻¹) < LMO/SiO₂-CI (32.5 kJ mol⁻¹) < LMO/SiO₂-WI (34.7 kJ mol⁻¹), suggesting that LMO/SiO₂-SD sample was beneficial to the proceeding of *n*-butylamine catalytic oxidation, which was in good accordance with its excellent catalytic performance.

In the literature reported previously,^{5,6,21} the temperature of *n*-butylamine complete conversion over the reported catalysts was about 280 °C. Apparently, the conversion of *n*-butylamine over the LMO/SiO₂-SD catalyst showed a monotonic increasing tendency with the increase of reaction temperature, and the temperature of *n*-butylamine complete conversion was similar to other reported catalysts. Moreover, the conversion of LMO/SiO₂-SD and reported catalysts showed no noticeable fluctuation with the increase of reaction temperature, which indicated that the LMO/SiO₂-SD catalyst in this study had an advantage in catalytic oxidation of *n*-butylamine.

To assess catalyst fatigue, the stability of LMO/SiO₂-SD has been studied. A durability test for the oxidation of *n*-butylamine over catalyst LMO/SiO₂-SD was carried out for 100 h at the reaction temperature of 300 °C, and the curves of *n*-butylamine conversion *versus* the time on stream was shown in Fig. 8. Obviously, complete conversion of *n*-butylamine was always observed during the entire reaction time and no deactivation was observed, indicating that the LMO/SiO₂-SD sample showed the stable catalytic performance of *n*-butylamine. To investigate the fate of the catalyst after reaction, the XRD and XPS analyses

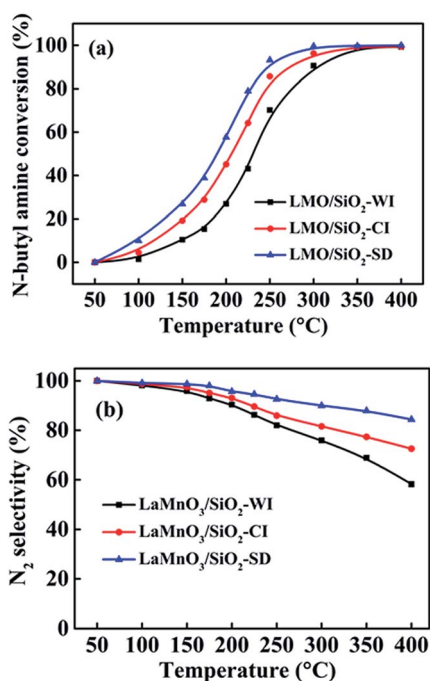


Fig. 7 Catalytic oxidation performances of *n*-butylamine over LMO/SiO₂ catalysts: (a) *n*-butylamine conversion (b) N₂ selectivity.



Table 2 Catalytic activities and E_a of LMO/SiO₂ catalysts

Catalysts	Catalytic activities			R^a /(mmol m ⁻² s ⁻¹)	E_a (kJ mol ⁻¹)
	T_{10} (°C)	T_{50} (°C)	T_{90} (°C)		
LMO/SiO ₂ -WI	146	230	305	4.35×10^{-7}	34.7
LMO/SiO ₂ -CI	116	206	273	7.77×10^{-7}	32.5
LMO/SiO ₂ -SD	95	188	246	6.85×10^{-7}	29.0

^a Calculated from the active test at 250 °C.

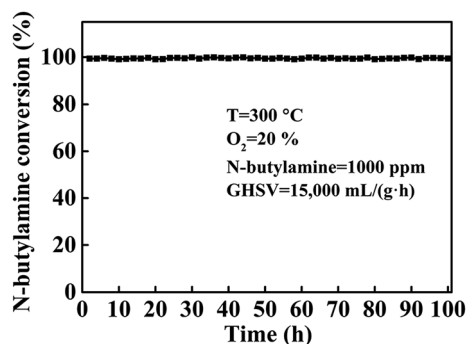


Fig. 8 Catalytic stability of LMO/SiO₂-SD catalyst.

of the LMO/SiO₂-SD catalyst after stability test were further carried out. From Fig. 1, the XRD analysis revealed that no phase transformations were observed, indicating its superior structural stability. Meanwhile, the XPS result (Table S1†) showed that the used LMO/SiO₂-SD catalyst displayed the similar atomic concentrations to the fresh catalyst, and the chemical states of major elements only showed a slightly drop than that of the fresh catalyst, which was in good accordance with the XRD result. All of the above characterization results demonstrated that the LMO/SiO₂-SD catalyst possessed the excellent stability.

3.3 *In situ* FTIR

To further investigate reaction mechanism of *n*-butylamine catalytic oxidation over all LMO/SiO₂ samples, we employed *in situ* FTIR under 1000 ppm *n*-butylamine, 20% O₂ and He conditions to analysis intermediate products at different temperature intervals and established a constructive relationship between samples' physicochemical properties and catalytic performances (Fig. 9). As shown in Fig. 9, when the reaction temperature was 50 °C, the induced *n*-butylamine adsorbed on the surface of LMO/SiO₂-SD sample, and occurred four adsorption peaks. Four intense bands belong to *n*-butylamine adsorbed over the LMO/SiO₂-SD catalyst at 1286 cm⁻¹ (ν_{C-N}), 1462 cm⁻¹ (δ_{CH_2}), 1545 cm⁻¹ (δ_{CH_3}) and 1659 cm⁻¹ (δ_{N-H}) were observed.⁴³⁻⁴⁷ However, these adsorption peaks would gradually weakened and vanished during the temperature-programmed surface reaction, and emerged three new intense bands located at 2347 cm⁻¹, 1724 cm⁻¹ and 1596 cm⁻¹, which could be assigned to the formation of CO₂, $\delta(C=O)$ and NH₃ adsorbed on at 2347 cm⁻¹, 1724 cm⁻¹ and 1596 cm⁻¹, which could be

assigned to the formation of CO₂, $\delta(C=O)$ and NH₃ adsorbed on the sample surface, respectively.^{39,40} Meanwhile, the relative intensity of CO₂ adsorption peak enhanced with the increase of reaction temperature, suggesting the increase of CO₂ formation. However, the other two intense bands of $\delta(C=O)$ and NH₃ fade away with reaction temperature increasing, which could be ascribed to the generation of CO₂, H₂O, N₂ and so on during the reaction process. When the reaction temperature reached to 300 °C, novel intense bands at 1557 cm⁻¹ and 1334 cm⁻¹ respectively ascribed to bidentate nitrate and *cis*-N₂O₂²⁻ species were observed, since more formed NO_x adsorbed on the surface of sample with reaction temperature rising.^{48,49} The *in situ* FTIR spectra of other LMO/SiO₂-CI and LMO/SiO₂-WI samples were similar to that of LMO/SiO₂-SD sample (Fig. S3†), merely the adsorption peaks appeared slightly shift.

Based on the results of *in situ* FTIR, the reaction pathway for catalytic oxidation of *n*-butylamine over LMO/SiO₂ samples was proposed, as shown in Scheme 2. There was a condensation route over the all samples: the *n*-butylamine was firstly induced and adsorbed on the surface of sample, forming the adsorbed *n*-butylamine, which was consecutively split C-N bond and produced -NH₂ and butyl. On the one hand, the -NH₂ species over sample might be adsorbed and oxidized to generate NO_x, which might be ascribed to the deep oxidation of NH₃ in channels of samples.^{21,22} Subsequently, the formed NO_x reacted with another -NH₂ to produce non-toxic N₂ and H₂O. On the other hand, the -NH₂ species might also form NH₃ *via* hydrogenation reaction and subsequently generated non-toxic N₂ and H₂O by means of NH₃-SCO. The butyl over the LMO/SiO₂-SD sample might be oxidized by active oxidation species to

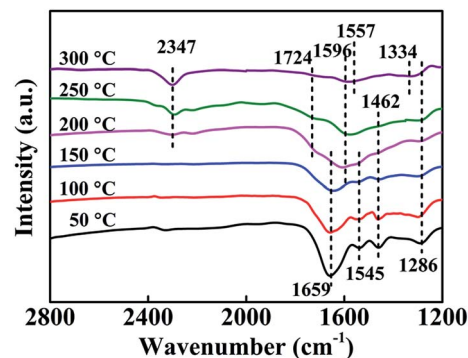
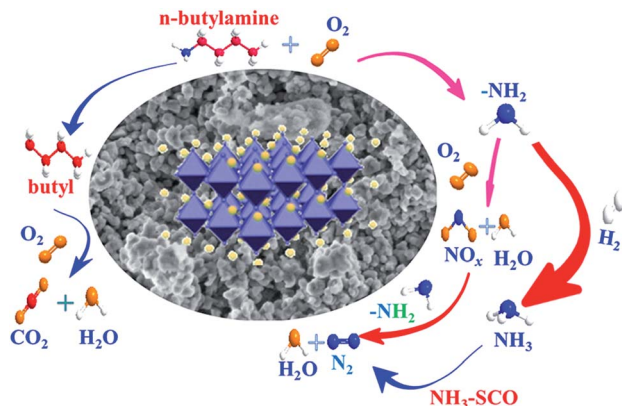


Fig. 9 *In situ* FTIR spectra of catalytic oxidation of *n*-butylamine over LMO/SiO₂-SD catalyst.





Scheme 2 Reaction mechanism for catalytic oxidation of *n*-butylamine over LMO/SiO₂ catalysts.

generate intermediate species, and then were oxidized to produce CO₂ and H₂O terminally.

4 Conclusions

In summary, citric acid-assisted deposition method utilized the synergistic effect of abundant active hydroxyl groups existed on SiO₂ gel surface and citric acid to prepare the mesoporous SiO₂-confined LaMnO₃ perovskite nanoparticles with high dispersion. The LMO/SiO₂-SD sample showed the optimum catalytic performance (T_{90} at 246 °C) and the highest N₂ selectivity for *n*-butylamine oxidation compared with other LMO/SiO₂-CI and LMO/SiO₂-WI samples, which was mainly assigned to its abundant surface acid sites, superior low-temperature reducibility and high ratio of Mn⁴⁺ species. Moreover, acid sites of LMO/SiO₂-SD sample played an essential role in reaction. By which, *n*-butylamine could be adsorbed efficiently and gone through further oxidation. In addition, the LMO/SiO₂-SD sample possessed the lowest E_a value (29 kJ mol⁻¹), proving its excellent catalytic activity. In addition, one can find that the LMO/SiO₂-SD with high dispersion exhibited advantage in controlling NO_x production. Based on analysis, one reason for this phenomenon is that the high dispersion can decrease the reaction temperature and then reduce the NO_x production. It is predictable that highly dispersed perovskite nano-catalysts will draw more attentions and play a vital role in catalytic oxidation of NVOCS.

Conflicts of interest

There are no conflicts to declare.

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