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## A colorimetric and ratiometric fluorescent sensor for sequentially detecting $\text{Cu}^{2+}$ and arginine based on a coumarin–rhodamine B derivative and its application for bioimaging<sup>†</sup>

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In this work, a colorimetric and ratiometric fluorescent sensor based on a coumarin–rhodamine B hybrid for the sequential recognition of  $\text{Cu}^{2+}$  and arginine (Arg) *via* the FRET mechanism was designed and synthesized. With the addition of  $\text{Cu}^{2+}$ , the solution displayed a colorimetric change from pale yellow to pink which is discernible by the naked eye. Additionally, the fluorescence intensities of the sensor exhibited ratiometric changes for the detection of  $\text{Cu}^{2+}$  at 490 and 615 nm under a single excitation wavelength of 350 nm, which corresponded to the emissions of coumarin and rhodamine B moieties, respectively. The fluorescence color change could be visualized from blue to pink. The limits of detection were determined to be as low as 0.50 and 0.47  $\mu\text{M}$  for UV-vis and fluorescence measurements, respectively. More importantly, the sensor not only can recognize  $\text{Cu}^{2+}$  and form a sensor- $\text{Cu}^{2+}$  complex but can also sequentially detect Arg with the resulting complex. The detection limits for Arg were as low as 0.60  $\mu\text{M}$  (UV-vis measurement) and 0.33  $\mu\text{M}$  (fluorescence measurement), respectively. A fluorescence imaging experiment in living cells demonstrated that the fabricated sensor could be utilized in ratiometric fluorescence imaging towards intracellular  $\text{Cu}^{2+}$ , which is promising for the detection of low-level  $\text{Cu}^{2+}$  and Arg with potentially practical significance.

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### Introduction

With the development of industrialization, heavy metal pollution is becoming increasingly severe, and it is receiving more and more attention because of its potential threat to human health and the ecological environment. Among various heavy metal ions, the  $\text{Cu}^{2+}$  ion plays an essential role in environmental, chemical, and biological systems.<sup>1–3</sup> However, an excess amount of  $\text{Cu}^{2+}$  in the body causes gastrointestinal disturbance and harms the liver and kidneys.<sup>4–8</sup> Therefore, the development of highly sensitive and selective methods for the detection of  $\text{Cu}^{2+}$  in biological and environmental samples is of great importance. Currently, fluorescence spectroscopy is regarded as a powerful tool for the detection and imaging of micro-samples.<sup>9–11</sup> So far, some fluorescent sensors for the sensitive determination of  $\text{Cu}^{2+}$  have been developed.<sup>12–14</sup> However, most of them are based on “turn-off” changes in fluorescence intensity because  $\text{Cu}^{2+}$  is usually used as a fluorescent quencher

*via* an energy or electron-transfer mechanism.<sup>15–18</sup> Therefore, the sensitive “turn-on” fluorescent sensors to monitor and visualize  $\text{Cu}^{2+}$  are highly demanded.

Amino acids are the essential components that make up the biological macromolecular proteins, which have been intensively investigated.<sup>19–21</sup> Among all natural amino acids, arginine (Arg) not only acts as the precursor to nitric oxide (NO), urea, ornithine, but also plays a vital role in myriad biological functions, such as cell replication, wound healing, and protein synthesis.<sup>22–24</sup> If the level of Arg in a human body is abnormal, it can cause unhealthy or even life-threatening problems for the person.<sup>25</sup> Therefore, it is significant to develop viable methods for Arg detection. Currently, several major methods have been used for the determination of Arg, such as high-performance liquid chromatography (HPLC) and gas chromatography (GC).<sup>26,27</sup> However, these methods usually involve specially designed and costly equipment, complicated sample preparation, and long measurement time. The fluorescent sensors compared with the above methods exhibited high selectivity, good sensitivity, and fast real-time detection, which have attracted extensive attention.<sup>28,29</sup> Currently, only a few sensors are available for the detection of Arg due to weak interactions between receptors and Arg in aqueous media.<sup>30–33</sup> The development of new fluorescent sensors that are sensitive and selective for the recognition of Arg remains a challenge.

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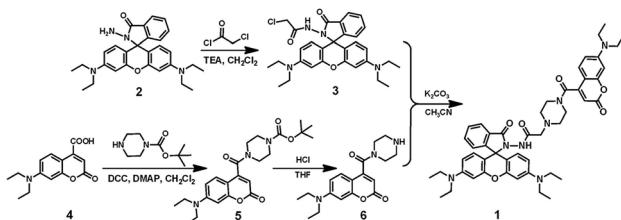
Among various fluorophores, a large number of rhodamine-based fluorescence sensors have been reported for the detection of analytes due to rhodamine has various excellent optical properties.<sup>34–37</sup> Recently, Pu *et al.* disclosed a fluorescent sensor based on diarylethene–rhodamine under a single emission wavelength, which showed good selectivity and sensitivity toward recognition of Cu<sup>2+</sup> and Arg.<sup>38</sup> In another excellent work, Xin *et al.* prepared a fluorescent sensor for fast visual detection of Al<sup>3+</sup> and Arg based on fluorescence intensity change of rhodamine–thiourea in single emission wavelength.<sup>39</sup> However, these single-signal sensor based on both fluorescence quenching and enhancement after the recognition of Cu<sup>2+</sup>/Al<sup>3+</sup> and Arg at a single emission wavelength, which tends to be influenced by various of factors, such as instrumental efficiency, the concentration of sensor molecules, photobleaching, and microenvironment.<sup>40,41</sup> In contrast, ratiometric sensors utilize the ratio of the intensity at two different wavelengths, which can effectively eliminate the above limitations by means of self-calibration and exclude the fluctuation of light excitation intensity.<sup>42–44</sup> In addition, the large emission shift is more conducive to reduce the interference of crosstalk signals and improve the detection accuracy compared to most of the ratiometric fluorescent sensors that exhibit only ordinary emission shifts (<80 nm).<sup>45,46</sup>

In this work, as shown in Scheme 1, we adopted the strategy that utilizes the coumarin chromophore as the energy donor and the rhodamine moiety as the energy acceptor, which is expected to construct a ratiometric fluorescence sensor successfully for detecting of analytes based on the FRET mechanism. As a consequence, we successfully synthesized a new colorimetric and ratiometric fluorescent sensor containing both coumarin and rhodamine B units for the visual detection of Cu<sup>2+</sup> and Arg *via* the FRET mechanism. During the detecting process for Cu<sup>2+</sup> and Arg, the fluorescent sensor not only achieved the ratiometric fluorescence change with a significant emission shift (125 nm) but also provided a vivid sight to “naked eyes”. Moreover, the sensor was successfully used to monitor Cu<sup>2+</sup> in living cells, and favorable ratiometric fluorescence signals were achieved, which has great potential application biosensing, disease diagnosis, and environmental monitoring.

## Results and discussion

### Absorption and fluorescence spectral responses of sensor 1 to metal ions

A series of metal ions (Fe<sup>3+</sup>, Al<sup>3+</sup>, Ca<sup>2+</sup>, Cd<sup>2+</sup>, Co<sup>2+</sup>, Cu<sup>2+</sup>, Ba<sup>2+</sup>, Cr<sup>3+</sup>, Hg<sup>2+</sup>, K<sup>+</sup>, Sr<sup>2+</sup>, Mg<sup>2+</sup>, Mn<sup>2+</sup>, Ni<sup>+</sup>, Pb<sup>2+</sup>, and Zn<sup>2+</sup>) were used to evaluate the UV-vis absorption and fluorescence selectivity of **1** (2.0



Scheme 1 Structure and synthesis of **1**.

$\times 10^{-5}$  mol L<sup>-1</sup>) in CH<sub>3</sub>CN–H<sub>2</sub>O (9/1, v/v) solution. Initially, in the absence of any metal ions, the color of solution **1** was light yellow, and the absorption spectrum did not show any absorption band in the visible light range. As expected, these metal ions were separately added to the solution of **1**, only Cu<sup>2+</sup> caused a significant absorption change with an obvious absorption peak at 564 nm and a distinct color change from light yellow to pink, which corresponded to the characteristic absorption peak of the open ring form of the rhodamine B moiety (Fig. 1A and C). Other metal ions did not induce significant absorption change of **1**. Similarly, only Cu<sup>2+</sup> induced a significant ratiometric change in the fluorescence spectra upon excitation at 350 nm. Upon the addition of Cu<sup>2+</sup>, the fluorescence intensity significantly decreased at 490 nm, and a new emission peak at 615 nm emerged along with the change of fluorescence color from blue to pink. Although other competing metal ions such as Fe<sup>3+</sup>, Al<sup>3+</sup>, and Cr<sup>3+</sup> caused a slight deduction of fluorescent intensity at 490 nm, they did not cause any significant fluorescence change at 615 nm (Fig. 1B and D). These results suggested that the sensor **1** is more selective to Cu<sup>2+</sup> than other competitive metal ions under the same conditions.

Subsequently, the competition experiments of absorption and fluorescence were carried out for sensor **1** (Fig. S1A†). Upon addition of Cu<sup>2+</sup> (3.0 eq.) into the solution of **1** containing other metal ions (10.0 eq.), a similar change in absorption spectra was observed as that with Cu<sup>2+</sup> ion only, implying that the binding affinity of Cu<sup>2+</sup> with sensor **1** is higher than those of other metal ions. In addition, the ratios of fluorescence intensity ( $I_{615}/I_{490}$ ) of the sensor **1** were examined under the same conditions with a series of metal ions. As shown in Fig. S1B,† most of the competitive ions did not induce significant fluorescence ratiometric changes, only Fe<sup>3+</sup>, Al<sup>3+</sup>, and Cr<sup>3+</sup> exhibited a slight interference. The competition experiments demonstrated that the sensor **1** could specifically recognize Cu<sup>2+</sup> with high anti-interference capability despite the presence of other competitive metal ions.

### Optical responses of sensor **1** to Cu<sup>2+</sup>

To gain a deeper understanding of the chemical sensing properties of **1** for Cu<sup>2+</sup>, we further conducted absorption and

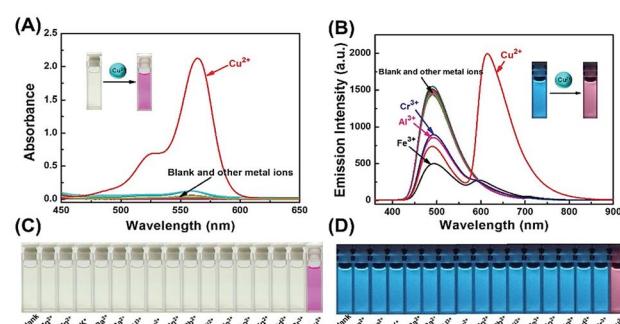


Fig. 1 Spectroscopic changes of **1** ( $2.0 \times 10^{-5}$  mol L<sup>-1</sup>) induced by the addition of various metal ions (3.0 eq.) in CH<sub>3</sub>CN–H<sub>2</sub>O (9/1, v/v) solutions. (A) Absorption spectral changes. (B) Fluorescence emission intensity changes. (C) The color image set upon the addition of various metal ions. (D) The fluorescence image set upon the addition of various metal ions.



fluorescence titration experiments. As shown in Fig. 2A, with the increase of  $\text{Cu}^{2+}$ , the absorption peak gradually enhanced at 564 nm, which could be caused by the ring-opened form of the rhodamine B moiety. Fig. 2B showed the absorption titration curve increased steadily with increasing concentration of  $\text{Cu}^{2+}$ , and the change of absorbance reached a plateau when 3.0 eq. of  $\text{Cu}^{2+}$  was further titrated, accompanied by the color change of the solution from light yellow to pink. As shown in Fig. S2,<sup>†</sup> the absorbance of sensor 1 exhibited a linear relationship with the concentration of  $\text{Cu}^{2+}$  in the range of 0–40  $\mu\text{M}$ , and the linear equation was determined to be  $Y = 0.04292X - 0.00994$  ( $R = 0.9962$ ). The limit of detection (LOD, UV-vis measurement) of 1 to  $\text{Cu}^{2+}$  was calculated to be 0.50  $\mu\text{M}$  based on  $3\sigma/k$  (Fig. S3<sup>†</sup>), wherein,  $\sigma$  represents the standard deviation of the blank signal, and  $k$  represents the slope of the linear calibration curve.<sup>47</sup>

The fluorescence spectra of sensor 1 in  $\text{CH}_3\text{CN}-\text{H}_2\text{O}$  (9/1, v/v) solution contained different concentrations of  $\text{Cu}^{2+}$ . In the absence of  $\text{Cu}^{2+}$ , the spectrum of 1 exhibited a blue fluorescence emission at 490 nm which is a characteristic of the coumarin moiety. As shown in the Fig. 2C, with continuous addition of  $\text{Cu}^{2+}$ , the fluorescence intensity at 490 nm decreased gradually, whereas a new emission peak at 615 nm appeared and increased gradually, indicating that the rhodamine moiety had changed from a spirocyclic form to a ring-opened form. Further addition of  $\text{Cu}^{2+}$  beyond a certain concentration (3.0 eq.) did not cause changes in fluorescence intensity (Fig. 2D). In addition, the distance between the two emission peaks exhibited a significant wavelength shift (125 nm), indicating that this system can effectively circumvent the issue of emission spectral overlap which is common with ratiometric sensors, allowing for accurate measurement of intensities of two emission peaks. Moreover, the ratiometric change of fluorescence intensity ( $I_{615}/I_{490}$ ) was linearly correlated with the concentration of  $\text{Cu}^{2+}$  within the range from 0 to 60  $\mu\text{M}$  and the good linear

relationship was found to be  $Y = 0.04295X + 0.02427$  ( $R = 0.9984$ ) (Fig. S4<sup>†</sup>). The fluorescence detection limit was measured to be 0.47  $\mu\text{M}$  based on  $3\sigma/k$  (Fig. S5<sup>†</sup>). The absorption and fluorescence titration experiments suggested that sensor 1 is a promising colorimetric and ratiometric fluorescence sensor for the determination of low-level  $\text{Cu}^{2+}$ .

### Absorption and fluorescence spectral responses of the 1- $\text{Cu}^{2+}$ complex to amino acids

As demonstrated above, the absorbance and fluorescence intensity of 1- $\text{Cu}^{2+}$  complex reached the maximum value when 3.0 eq. of  $\text{Cu}^{2+}$  were added. To investigate the reversibility of the binding of sensor 1 with  $\text{Cu}^{2+}$ , the 1- $\text{Cu}^{2+}$  complex obtained with 2.0 eq.  $\text{Cu}^{2+}$  and 2 mL solution of 1 ( $2 \times 10^{-5}$  mol  $\text{L}^{-1}$ ) in  $\text{CH}_3\text{CN}-\text{H}_2\text{O}$  (9/1, v/v) was used as the initial state.

To evaluate the selectivity of the 1- $\text{Cu}^{2+}$  complex toward various amino acids in  $\text{CH}_3\text{CN}-\text{H}_2\text{O}$  (9/1, v/v) solution, we first conducted the absorption experiment by using UV-vis spectroscopy. Upon addition of a constant amount of each amino acid (35 eq.), only Arg resulted in the complete disappearance of the visible absorption peak at 564 nm with a color change from pink to light yellow (Fig. S6A and C<sup>†</sup>). Furthermore, the selectivity of the 1- $\text{Cu}^{2+}$  complex for amino acids was also investigated with fluorescence spectroscopy. As expected, Arg also could be distinguished from other amino acids in fluorescence spectra at the same conditions. Upon the addition of amino acids to solution of 1- $\text{Cu}^{2+}$  complex, only Arg resulted in the quench of fluorescence emission at 615 nm and the increase of emission at 490 nm (Fig. S6B and D<sup>†</sup>).

Competition experiments were also performed by adding the Arg (30 eq.) to the solution of the 1- $\text{Cu}^{2+}$  complex in the presence of 50.0 eq. other amino acids, in which similar absorption spectral changes were observed as that with Arg only, indicating that other amino acids have a negligible effect on the recognition of Arg for the 1- $\text{Cu}^{2+}$  complex (Fig. S7A<sup>†</sup>). In addition, the fluorescence responses of 1- $\text{Cu}^{2+}$  complex towards various amino acids after the addition of Arg were measured. As shown in Fig. S7B,<sup>†</sup> for 1- $\text{Cu}^{2+}$  complex, the ratiometric fluorescence response induced by Arg was hardly affected in the presence of other amino acids. These results suggested that the presence of other amino acids at relatively high concentrations did not interfere with the detection of Arg under the same experimental conditions.

### Optical responses of the 1- $\text{Cu}^{2+}$ complex to Arg

To gain further insight into interactions of 1- $\text{Cu}^{2+}$  complex and Arg, absorption and fluorescence titration experiments were conducted in  $\text{CH}_3\text{CN}-\text{H}_2\text{O}$  (9/1, v/v) solution. As shown in Fig. 3A, upon gradual addition of Arg to the 1- $\text{Cu}^{2+}$  complex solution, the absorption peak at 564 nm was steadily decreased. Simultaneously, a distinct color change from pink to light yellow could be observed (Inset in Fig. 3A). As shown in Fig. 3B, the absorbance of 1- $\text{Cu}^{2+}$  complex decreased at 564 nm with the addition of incremental amounts of Arg and completely disappeared when 35 eq. Arg was added. Moreover, the change of absorbance at 564 nm was linearly related ( $Y = -0.00529X + 2.00338$ ,  $R^2 = 0.9971$ ) to the concentration of Arg in a range of

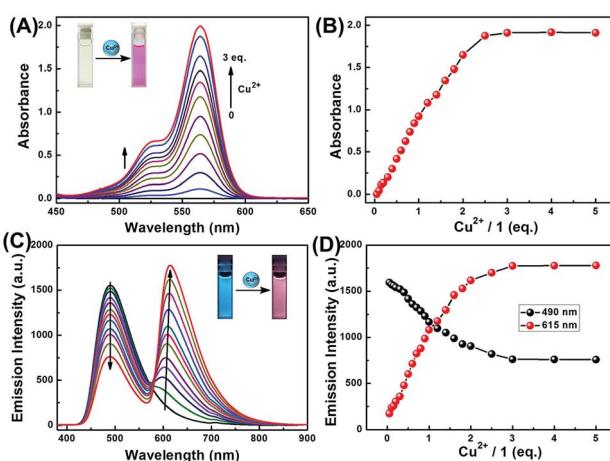
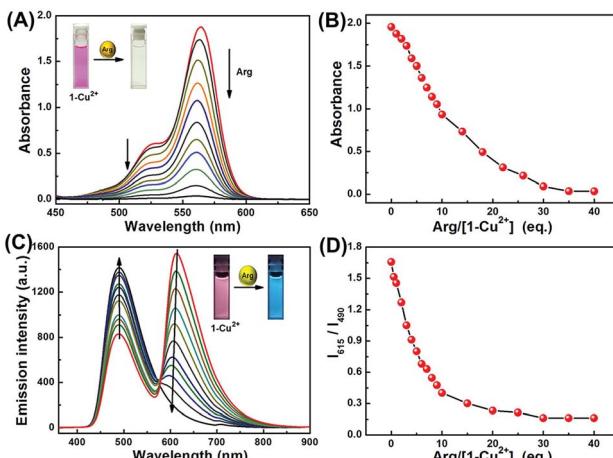


Fig. 2 (A) Changes in absorption spectra of 1 in  $\text{CH}_3\text{CN}-\text{H}_2\text{O}$  (9/1, v/v) solutions upon addition of  $\text{Cu}^{2+}$ . (B) Absorption titration profile (at 564 nm) versus concentration of  $\text{Cu}^{2+}$  for 1. (C) Fluorescence spectra of 1 upon titration with  $\text{Cu}^{2+}$  (0–5.0 eq.) in  $\text{CH}_3\text{CN}-\text{H}_2\text{O}$  (9/1, v/v). Inset: the color change in fluorescence of 1. (D) Fluorescence intensity of sensor 1 at 615 nm (red dot) and 490 nm (black dot) as a function of the  $\text{Cu}^{2+}$  concentration (0–5.0 eq.).  $\lambda_{\text{ex}} = 350$  nm, slit: 5 nm/5 nm.





**Fig. 3** (A) Changes in absorption spectra of **1**–Cu<sup>2+</sup> in CH<sub>3</sub>CN–H<sub>2</sub>O (9/1, v/v) solutions upon addition of Arg (0–40.0 eq.). (B) Absorption titration profile (at 564 nm) versus concentration of Arg for **1**–Cu<sup>2+</sup>. (C) Fluorescent intensities ratio ( $I_{615}/I_{490}$ ) change of **1**–Cu<sup>2+</sup> solution in CH<sub>3</sub>CN–H<sub>2</sub>O (9/1, v/v,  $2 \times 10^{-5}$  mol L<sup>-1</sup>) on incremental addition of Arg (0–40.0 eq.). (D) The relationship between fluorescent intensities ratio ( $I_{615}/I_{490}$ ) change.

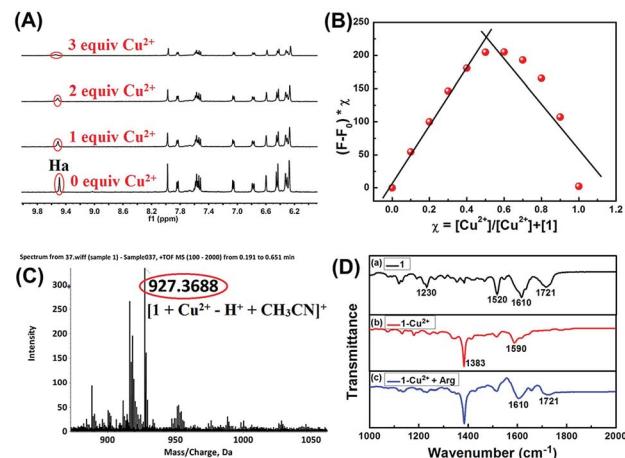
0 to 200  $\mu$ M (Fig. S8†). The UV-vis detection limit of the **1**–Cu<sup>2+</sup> complex toward Arg was determined to be as low as 0.60  $\mu$ M based on  $3\sigma/k$  (Fig. S9†).

Similarly, the sensing property of **1**–Cu<sup>2+</sup> complex toward Arg was further investigated by using fluorescence titration experiment. As shown in Fig. 3C, upon excited at 350 nm, the fluorescence intensity around 615 nm decreased with the incremental addition of Arg, whereas the emission band around 490 nm gradually increased. The fluorescence spectra of the **1**–Cu<sup>2+</sup> complex did not show further change when 30 eq. Arg was added. The ratio of fluorescence intensity at 615 and 490 nm ( $I_{615}/I_{490}$ ) exhibited a 14-fold variation in the ratios from 1.67 to 0.12 in the **1**–Cu<sup>2+</sup> complex (Fig. 3D). As shown in Fig. S10,† **1**–Cu<sup>2+</sup> complex exhibited an excellent linear relationship ( $Y = -0.00621X + 1.52485$ ,  $R^2 = 0.9788$ ) between the relative emission ratios and the concentration of Arg in a range of 0 to 200  $\mu$ M. In addition, the fluorescence detection limit ( $3\sigma/k$ ) was calculated to be 0.33  $\mu$ M (Fig. S11†). For **1**–Cu<sup>2+</sup> complex, it exhibited optimal performance for the detection of Arg including low LOD, high sensitivity, and “naked-eyes” recognition, which suggested that it could be used as a fluorescence sensor for determination of Arg.

#### Detection mechanism for Cu<sup>2+</sup> and Arg

To investigate the mechanism of interactions between sensor **1** and Cu<sup>2+</sup>, we recorded <sup>1</sup>H NMR spectra of **1** with different equivalents of Cu<sup>2+</sup> (Fig. 4A). As shown in <sup>1</sup>H NMR, upon addition of Cu<sup>2+</sup> to the DMSO-*d*<sub>6</sub> solution of **1**, the imine proton (H<sub>a</sub>) at around  $\delta$  9.49 ppm disappeared. The result showed that the formation of the **1**–Cu<sup>2+</sup> complex is through the coordination of the imine proton in sensor **1** with Cu<sup>2+</sup>.

Furthermore, Job's plot experiments were conducted to understand the binding stoichiometry of **1**–Cu<sup>2+</sup> complex. In the

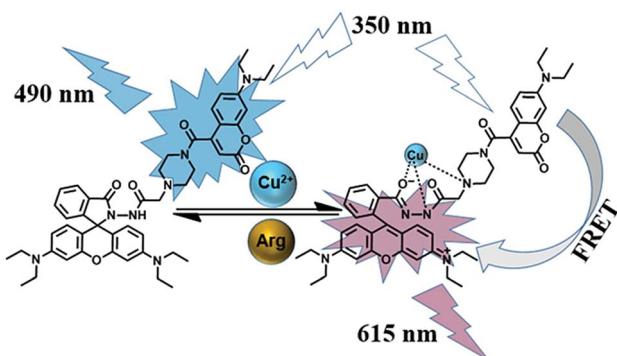


**Fig. 4** (A) Partial <sup>1</sup>H NMR spectra change of **1** in DMSO-*d*<sub>6</sub> upon addition of Cu<sup>2+</sup>. (B) Job's plot for the complex of **1** with Cu<sup>2+</sup>, indicating the formation of a 1 : 1 complex. The total [1] + [Cu<sup>2+</sup>] = 40  $\mu$ M. (C) ESI-MS spectrum of **1**–Cu<sup>2+</sup>. (D) Partial FT-IR spectra of **1** (a), **1**–Cu<sup>2+</sup> (b) and **1**–Cu<sup>2+</sup> + Arg (c).

experiments, the total concentration of sensor **1** and Cu<sup>2+</sup> was maintained at 40  $\mu$ M, and the fluorescence responses of **1** toward Cu<sup>2+</sup> were monitored at 615 nm as a function of the molar fraction of Cu<sup>2+</sup>. As shown in Fig. 4B, when the molar fraction of Cu<sup>2+</sup> was 0.5, the fluorescence intensity reached the maximum, which suggested that the binding stoichiometry of **1**–Cu<sup>2+</sup> complex is 1 : 1. Moreover, the binding stoichiometry of **1**–Cu<sup>2+</sup> complex was confirmed by the ESI-MS spectrum in which the peak at 927.36 corresponded to [1 + Cu<sup>2+</sup> – H<sup>+</sup> + CH<sub>3</sub>CN]<sup>+</sup> (Fig. 4C). These results confirmed that the **1**–Cu<sup>2+</sup> complex is a 1 : 1 complexation stoichiometry.

FT-IR measurements were also performed to investigate the mechanism of interactions between **1**–Cu<sup>2+</sup> complex and Arg. The FT-IR spectra of **1**, **1**–Cu<sup>2+</sup> and **1**–Cu<sup>2+</sup> + Arg were shown in Fig. 4D. As illustrated in Fig. 4D(a), the two peaks located at 1721 and 1610  $\text{cm}^{-1}$  could correspond to the antisymmetric and symmetric stretching vibrations of C=O (the rhodamine B unit),<sup>48</sup> and the strong band at 1520  $\text{cm}^{-1}$  could be assigned to C=C stretching of benzenoid units.<sup>49</sup> For the **1**–Cu<sup>2+</sup> complex, because of the disappearance of the C=O group (1721  $\text{cm}^{-1}$ ) absorption band, the absorption peak at 1610  $\text{cm}^{-1}$  shifted to 1590  $\text{cm}^{-1}$ , indicating that the C=O functional group bonded with Cu<sup>2+</sup> and the spirolactam structure in rhodamine B changed from off to on. The FT-IR spectrum of Fig. 4D(b) showed a strong absorption peak at 1383  $\text{cm}^{-1}$ , which could be attributed to the characteristic absorption band of excess Cu(NO<sub>3</sub>)<sub>2</sub>. By comparing the FT-IR spectra of **1** (Fig. 4D(a)) and **1**–Cu<sup>2+</sup> (Fig. 4D(d)), the absorption band at 1721  $\text{cm}^{-1}$  reappeared, and the absorption band at 1590  $\text{cm}^{-1}$  was restored to its initial state after the addition of Arg, which indicated the reformation of spirolactam from open chain amide. These results indicated that the sensor **1** first chelates with Cu<sup>2+</sup> to form a **1**–Cu<sup>2+</sup> complex, and then restores to its initial state with the addition of Arg because of higher binding affinity between Arg and Cu<sup>2+</sup> which leads to the escape of Cu<sup>2+</sup> from **1**–Cu<sup>2+</sup> complex.





Scheme 2 Principle for sequentially detecting  $\text{Cu}^{2+}$  and arginine.

It is well recognized that a large spectral overlap between the donor emission spectrum and the acceptor absorption spectrum can achieve efficient FRET. In this work, our strategy is using coumarin group as the energy donor and rhodamine B moiety as an energy acceptor. As shown in Fig. S12,<sup>†</sup> coumarin group has strong fluorescence in the range of 425–650 nm in the absence of  $\text{Cu}^{2+}$ , but the FRET pathway was completely suppressed. After addition of  $\text{Cu}^{2+}$ , the fluorescence spectrum of coumarin group (donor) overlapped with the absorption spectrum of the ring-opened amide isomer of rhodamine B unit (450–600 nm, acceptor), which indicated that the efficient FRET process could occur. Subsequent addition of Arg to the solution of **1**– $\text{Cu}^{2+}$  complex regenerated the spirolactam structure in the rhodamine B. As a result, the FRET process was discontinued, and the emission spectrum came back to the original state of sensor **1**. The corresponding mechanism for detecting  $\text{Cu}^{2+}$  and Arg with sensor **1** is illustrated in Scheme 2.

### Fluorescence imaging of $\text{Cu}^{2+}$ in living cells

To demonstrate its potential application value in biological systems, the HeLa cells were selected to study the application feasibility of the ratiometric sensor **1** preliminarily for

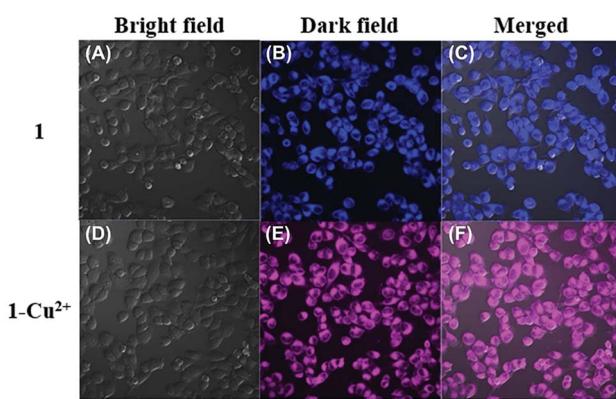


Fig. 5 Confocal fluorescence images of HeLa cells incubated with sensor **1** (20  $\mu\text{M}$ ) for 30 min (A–C) and then treated with  $\text{Cu}^{2+}$  (2 mM) for another 30 min (D–F). Images were obtained using an excitation of 405 nm and emission channels of (B) at 430–530 nm and (E) at 550–650 nm; (C and F) merge images of (A, B and D, E); (A and D) bright field images of the cell culture.

intracellular  $\text{Cu}^{2+}$  imaging. As shown in Fig. 5, the fluorescence images were recorded at 430–530 nm and 550–650 nm at an excitation wavelength of 405 nm. After the HeLa cells were incubated with the addition of sensor **1** (20  $\mu\text{M}$ ) for 30 minutes in the absence of  $\text{Cu}^{2+}$ , a strong blue fluorescence of the coumarin group (Fig. 5B) was observed almost without pink fluorescence of rhodamine B. In contrast, after the addition of  $\text{Cu}^{2+}$ , the HeLa cells exhibited a significant change in fluorescence, in which a strong pink fluorescence appeared (Fig. 5E) with the quenching of blue fluorescence. These findings are consistent with the results of the previous experiments and analysis of the fluorescence intensity of the sensor **1** to  $\text{Cu}^{2+}$ . These preliminary study results implied that the sensor **1** is cell permeable and can be applied to ratiometric fluorescence imaging of intracellular  $\text{Cu}^{2+}$ .

## Conclusions

In conclusion, a colorimetric and ratiometric fluorescent sensor based on coumarin and rhodamine B units has been developed. The  $\text{Cu}^{2+}$ -induced changes in color and fluorescence not only have high selectivity and sensitivity but also can be visualized by “naked-eyes”. According to ESI-MS spectrometric results and Job's plot analysis, the metal-ligand complex of  $\text{Cu}^{2+}$  and sensor **1** was formed in a ratio of 1 : 1. In addition, the spectral shifts of two obvious emission bands were up to 125 nm, which can reduce the interference of crosstalk signals, eliminate any influence of excitation back-scattering, and improve the detection accuracy. More importantly, the **1**– $\text{Cu}^{2+}$  complex could be successfully applied to detect Arg over other amino acids selectively. The fluorescence imaging experiment confirmed that the ratiometric imaging ability of the sensor **1** towards trace intracellular  $\text{Cu}^{2+}$  in living cells. Therefore, the sensor **1** can be used to detect trace amounts of  $\text{Cu}^{2+}$  and Arg with remarkable changes in both color and fluorescence, which is highly desirable for biological detection and imaging.

## Experimental

### Reagents and methods

All inorganic salts, except for  $\text{K}^+$ ,  $\text{Hg}^{2+}$ , and  $\text{Mn}^{2+}$ , were obtained by corresponding metal nitrates. Typically, 1 mmol of each reagent was mixed with 10 mL of distilled water to obtain an inorganic salt solution (0.1 mol  $\text{L}^{-1}$ ). All amino acids were prepared with deionized water to obtain the desired concentration (0.1 mol  $\text{L}^{-1}$ ). All the stock solutions were stored in the dark and diluted to the required concentration when needed. The sensor **1** was dissolved in  $\text{CH}_3\text{CN}$  to obtain the stock solution ( $1.0 \times 10^{-3}$  mol  $\text{L}^{-1}$ ). The fluorescence and absorption experiments of sensor **1** ( $2.0 \times 10^{-5}$  mol  $\text{L}^{-1}$ ) were measured by adding the appropriate amount of metal ions into 2 mL of  $\text{CH}_3\text{CN}-\text{H}_2\text{O}$  (9/1, v/v) solution. In the experiments of amino acid detection, the solutions containing **1**– $\text{Cu}^{2+}$  complex were obtained by adding  $\text{Cu}^{2+}$  (2.0 eq.) into a solution of the sensor **1** in  $\text{CH}_3\text{CN}-\text{H}_2\text{O}$  (9/1, v/v,  $2.0 \times 10^{-5}$  mol  $\text{L}^{-1}$ ).

## Instrumentation

NMR spectra were obtained by a Bruker AV400 ( $^1\text{H}$  NMR 400 MHz,  $^{13}\text{C}$  NMR 100 MHz) instrument using  $\text{CD}_2\text{Cl}_2$  and  $\text{DMSO}-d_6$  as solvent and tetramethylsilane (TMS) as the internal standard. Mass spectrometry analysis was performed on an Agilent 1100 ion trap MSD spectrometer. Melting points were measured with a WRS-1B melting point apparatus. Fluorescence spectra were collected with a Hitachi F-4600 fluorescence spectrophotometer. UV-vis absorption spectra were carried out using an Agilent 8453 UV-vis spectrophotometer. Infrared spectra (IR) were collected on a Bruker Vertex-70 spectrometer. Elemental analysis was determined on a PE CHN 2400 analyzer. Fluorescence imaging of cells was performed using an Olympus FV1000 confocal laser scanning microscope.

## Cell culture and imaging

The living HeLa cells, namely human cervical cancer cells, were cultured in DMEM (Dulbecco modified Eagle's medium) supplemented with 10% FBS (fetal bovine serum) in an atmosphere consisting of 5%  $\text{CO}_2$  and 95% air at 37 °C. The cells were plated in a confocal dish and allowed to adhere overnight. The next day before the experiments, the cells were washed with phosphate-buffered saline (PBS) buffer. After the cells were incubated with the addition of sensor **1** (20  $\mu\text{M}$ ) in the culture medium for 30 min, the cells were incubated with  $\text{Cu}^{2+}$  (2.0 mM) for another 30 min. Subsequently, the fluorescence images were acquired by an OLYMPUS FV1000 confocal laser scanning microscope with an excitation filter of 405 nm and emission channels of 430–530 nm and 550–650 nm.

## Synthesis of the sensor 1

The synthetic route to sensor **1** is depicted in Scheme 1. Compounds **2**, **3**, and **6** were synthesized as described in the literature.<sup>38,50,51</sup> To a stirring mixture of compound **3** (0.17 g, 0.32 mmol) and compound **6** (0.12 g, 0.35 mmol) in  $\text{CH}_3\text{CN}$  (40 mL),  $\text{K}_2\text{CO}_3$  (0.057 g, 0.42 mmol) was added and continuously stirred. Then the mixture solution was refluxed for 12 h. After evaporation of the solvent, the resulting residue was purified by silica gel column chromatography with ethyl acetate to obtain target compound **1** (0.17 g, 0.20 mmol) with a yield of 62.5%.  $\text{Mp}$  427–428 K;  $^1\text{H}$  NMR (400 MHz,  $\text{CD}_2\text{Cl}_2$ ),  $\delta$  (ppm): 8.17 (s, 1H), 7.83 (d,  $J$  = 4.0 Hz, 1H), 7.63 (s, 1H), 7.44 (m, 2H), 7.24 (d,  $J$  = 8.0 Hz, 1H), 7.01 (d,  $J$  = 8.0 Hz, 1H), 6.55 (d,  $J$  = 8.0 Hz, 1H), 6.50 (d,  $J$  = 8.0 Hz, 2H), 6.42 (s, 1H), 6.25 (s, 3H), 6.23 (s, 1H), 3.39–3.35 (m, 6H), 3.31–3.17 (m, 8H), 3.09 (s, 2H), 2.89 (s, 2H), 2.26 (d,  $J$  = 12.0 Hz, 4H), 1.14 (t,  $J$  = 6.0 Hz, 6H), 1.04 (t,  $J$  = 8.0 Hz, 12H) (Fig. S13†);  $^{13}\text{C}$  NMR (100 MHz,  $\text{CD}_2\text{Cl}_2$ ),  $\delta$  (ppm): 166.67, 164.33, 164.09, 158.34, 156.74, 153.20, 151.22, 150.87, 148.60, 143.50, 132.61, 129.15, 129.00, 128.78, 127.91, 123.44, 122.69, 116.03, 108.82, 107.52, 107.10, 103.60, 96.98, 96.37, 65.47, 59.93, 46.64, 44.40, 43.87, 41.47, 11.81, 11.68 (Fig. S14†); ESI-MS ( $m/z$ ): 826.42 [1 +  $\text{H}^+$ ]<sup>+</sup> (Fig. S15†); anal. calcd for  $\text{C}_{48}\text{H}_{55}\text{N}_7\text{O}_6$ : C, 69.80; H, 6.71; N, 11.87; O, 11.62%. Found: C, 69.67; H, 6.88; N, 11.78; O, 11.67%.

## Conflicts of interest

There are no conflicts to declare.

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