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The effect of ZnCl₂ activation on microwave absorbing performance in walnut shell-derived nano-porous carbon

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Porous carbon has been expected to be a potential candidate as a lightweight and efficient microwave absorber. Nano-porous carbon carbonized directly from a walnut shell exhibits narrow microwave absorption frequency bandwidth, while the activation process can adjust the pore structure and optimize the microwave absorption performance. Herein, porous carbon materials were successfully prepared using walnut shells as precursors and ZnCl₂ as the activating agent. The superior microwave absorption performances of the as-prepared samples could be attributed to the well-developed pore structures and the enhanced dielectric loss capacities of the samples. The interfacial polarization in the walls of the pores and the defects in the samples significantly contributed to the enhancement of the dielectric loss capacities of the samples. In this work, the broadband microwave absorbing porous carbon exhibited an effective absorption bandwidth (reflection loss ≤ -10 dB) of 7.2 GHz (ranging from 10.8 GHz to 18.0 GHz) when the absorber thickness was 2.5 mm. In addition, an effective absorption bandwidth of 6.0 GHz (ranging from 11.4 GHz to 17.4 GHz) could also be achieved when the absorber thickness was only 2.0 mm. The samples exhibited low densities, strong microwave absorption performances and wide effective absorption bandwidths with thin absorber thicknesses, due to which they have a great potential as lightweight and efficient microwave absorbers.

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1. Introduction

Nowadays, with the rapid development of electromagnetic techniques, electromagnetic radiation and electromagnetic interference have disturbed people's daily life. Electromagnetic wave pollution may not only hamper the effectiveness of sophisticated electronic instruments but also damage the ecological environment and threaten human health.¹⁻³ To address this growing problem, the exploitation of efficient microwave absorption materials that can convert the electromagnetic energy to other forms of energy is considered the most effective solution. In view of practical applications, microwave absorption materials should not only have strong microwave absorption but also have excellent properties such as a broad absorption frequency bandwidth, thin coating thickness, light weight, and good thermal stability.⁴⁻⁶

According to the electromagnetic loss mechanism, microwave absorption materials are generally divided into magnetic loss materials and dielectric loss materials. Traditional magnetic loss materials include ferrites,⁷⁻⁹ magnetic metals¹⁰ and alloys,¹¹ and they are favourable for absorbing microwaves because of considerable magnetic loss. Although these magnetic materials can have strong microwave absorption, they have large densities, which limit their practical applications. Dielectric loss materials include conductive polymers,¹² semiconductor-based materials,¹³ ceramics,¹⁴ and carbon materials.¹⁵ Among them, carbonaceous microwave absorbers have aroused intensive attention because of their excellent physical and chemical performances such as low density, superior electric properties, good mechanical strength, thermal stability and corrosion resistance.¹⁶⁻¹⁸ In particular, porous carbons have large specific surface areas and can attenuate microwaves effectively by good dielectric loss, multiple reflections and scattering of microwaves.¹⁹⁻²¹ At present, challenges remain in the field of microwave absorption since the effective absorption frequency bandwidth of most microwave absorbers is still unsatisfactory and the preparation technology is still complicated. It is necessary to exploit easily prepared and low-cost microwave absorbers that have strong microwave absorption and broad effective absorption frequency bandwidth and are light in weight.

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Recently, biological materials are receiving wide attention. They have been used in various fields and they can be used in super capacitor electrodes,^{22–24} battery anodes,²⁵ carbon dioxide capture,²⁶ organic pollutant adsorption materials,²⁷ *etc.* In consideration of the rich naturally porous structures in biomass materials, some researchers have tried to use them as candidates of microwave absorbers through carbonization.^{21,28,29} However, natural pores in most biomass materials are very large and they may not be sensitive to microwaves. Hence, pores with a suitable size in the interior of biomass materials are favourable for increasing the multiple reflections and scattering of microwaves. A simple approach is to make use of highly acidic or strong alkaline substances to etch biomass materials during carbonization. These corrosive substances include KOH^{30,31} and H₃PO₄.^{32,33} Different corrosive substances can form different pore structures in biomass materials. Usually, biomass materials etched by strong alkaline substances can have numerous micropores and high specific surface areas.³⁴ Nevertheless, strong alkaline substances will destroy the natural large pores of biomass materials and make microwaves difficult to spread into the interior of the material. The ideal pore structures for efficient microwave absorbing are hierarchical porous structures. Microwaves radiate into the interior of the material through the large pores on the outer surface of the material first and then make multiple reflections on the small pore walls inside the material. Therefore, it is important to retain the natural large pores of biomass materials.

Herein, walnut shells were used as the carbon source to prepare porous carbon because walnut shells have many natural large pores that help microwaves to propagate into the interior of the material easily. In order to improve the microwave absorption performance of the material, it is necessary to create small pores in the interior of the material to enhance interfacial polarization relaxation loss and make microwaves reflect and scatter in the small pore structure. Zinc chloride in an aqueous solution is faintly acidic, and it will not destroy the natural large pores on the outer surface of the walnut shells, which is observed with highly acidic or strong alkaline substances such as H₃PO₄ or KOH. In addition, the zinc chloride aqueous solution can flow into the interior of the walnut shells to etch lignin and cellulose with three-dimensional network structures to create three-dimensional interconnected small pores. Thus, zinc chloride was used to corrode walnut shells. In this work, walnut shells were first carbonized at 400 °C and then immersed in the zinc chloride aqueous solution with different concentrations and calcined at various temperatures; the porous carbon was finally obtained after washing and drying (Fig. 1). The optimized porous carbon exhibited good dielectric loss capacities and excellent microwave absorption performance; the effect of ZnCl₂ activation on the microwave absorbing performance in the walnut shell-derived nano-porous carbon was also discussed in detail in this work.

2. Experimental

2.1. Raw materials

Huang long walnut shells were bought at Henan Yuan Run Water Treatment Material Co., Ltd and were crushed into

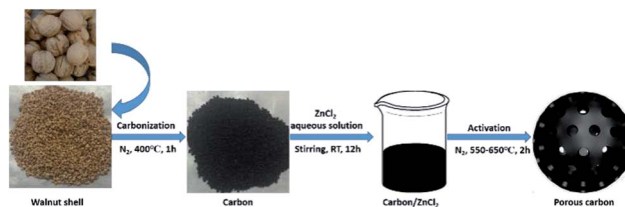


Fig. 1 Schematic illustration of the preparation of walnut shell-derived nano-porous carbon through ZnCl₂ activation.

particles in the size range of 0.8–1.6 mm. High purity nitrogen (>99.999%) was bought at Nanjing Sanle Electronic Information Industry Group Co., Ltd. Zinc chloride (ZnCl₂) was purchased from Shanghai Xinbao Industry of Fine Chemical. Absolute ethyl alcohol was purchased from Wuxi City Yasheng Chemical Co., Ltd. All of the chemical reagents were analytically pure and used without further purification.

2.2. Preparation of samples

In this work, walnut shell-derived nano-porous carbon was prepared by ZnCl₂ activation, as shown in Fig. 1. The details of the experiment are as follows: first, walnut shells (size: 0.8–1.6 mm) were washed thoroughly with absolute ethyl alcohol and deionized water several times and then dried at 60 °C. The dried walnut shells underwent a pyrolysis process under nitrogen atmosphere in a tube furnace (from room temperature to 400 °C, 5 °C min⁻¹; 400 °C, 1 h; from 400 °C to room temperature, cooled in the furnace) and the carbonized products and some tar were formed. A certain amount of zinc chloride (12 g, 18 g, 24 g and 30 g) was added into 30 mL deionized water and stirred to form the homogeneous ZnCl₂ aqueous solution with different concentrations. Also, 6 g of as-prepared carbonized products was then immersed into the zinc chloride aqueous solution and stirred for 12 h at room temperature. After being dried at 60 °C, the mixture was heated up to the specified temperatures (550 °C, 600 °C and 650 °C) at a heating rate of 5 °C min⁻¹, held for 2 h under nitrogen atmosphere and then cooled in the furnace. After thermal treatment, the as-prepared samples were washed with absolute ethyl alcohol and hot deionized water several times and then dried at 60 °C. The resulting porous biomass carbon materials are marked as PC-X-Y, where X represents the ZnCl₂ content (12 g, 18 g, 24 g and 30 g) in the ZnCl₂ aqueous solution and Y represents the calcination temperature (550 °C, 600 °C and 650 °C). For example, PC-18-600 represents that the ZnCl₂ content in the ZnCl₂ aqueous solution is 18 g and the calcination temperature is 600 °C. For comparison, walnut shells were also treated with the same process as that for PC-18-600 without using ZnCl₂, and the product was marked as C-600.

2.3. Characterization of samples

The morphologies of the samples were observed by transmission electron microscopy (FEI G2F20, FEI, USA). The Raman spectra of the samples were recorded using a Raman spectrometer (Labram HR800, Horiba, Japan, 514 nm laser) to



analyze the graphitization degree and the defective nature of the samples. X-ray photoelectron spectroscopy (Escalab 250Xi, Thermo, USA) was performed to analyze the surface chemical state of the sample. Pore structures of the samples were characterized by nitrogen adsorption–desorption isotherms at 77 K on a Belsorp-Max surface area analyzer (Belsorp-Max, MicrotracBEL, Japan). All samples were degassed at 150 °C under vacuum for 12 h before the measurements. The Brunauer–Emmett–Teller (BET) equation and the Horvath–Kawazoe (HK) theory were used to calculate the specific surface area and the pore size distribution of the sample, respectively. Total pore volume was calculated at a relative pressure $P/P_0 = 0.991$.

2.4. Microwave absorption measurements

Paraffin wax was used as the binder, and the as-prepared samples and molten paraffin wax with a weight ratio of 7 : 3 were mixed uniformly and then pressed into toroidal-shaped samples ($\varphi_{\text{out}} = 7.00$ mm, $\varphi_{\text{in}} = 3.04$ mm). The relative complex permittivity ($\epsilon_r = \epsilon' - j\epsilon''$) and the relative complex permeability ($\mu_r = \mu' - j\mu''$) of the toroidal-shaped samples in the frequency range of 2–18 GHz can be obtained by using the coaxial-line method on an Agilent 85050D vector network analyzer. Finally, the reflection loss (RL) values can be obtained according to the transmission line theory using the following equations:^{35–37}

$$Z_{\text{in}} = (\mu_r/\epsilon_r)^{1/2} \tanh[j(2\pi f d/c)(\mu_r\epsilon_r)^{1/2}] \quad (1)$$

$$\text{RL} = 20 \log|(Z_{\text{in}} - 1)/(Z_{\text{in}} + 1)| \quad (2)$$

Here, Z_{in} is the normalized input characteristic impedance of the absorber, μ_r and ϵ_r are the relative complex permeability and permittivity of the absorber, respectively, f is the electromagnetic wave frequency, d is the thickness of the absorber, and c is the velocity of electromagnetic waves in free space (3.0×10^8 ms⁻¹).

3. Results and discussion

3.1. Morphologies and microstructures

Fig. 2 exhibits the XRD patterns of the samples PC-18–550, C-600, PC-18–600, PC-24–600, PC-30–600 and PC-18–650. We can see that all samples show similar peak shapes. Two broad and weak diffraction peaks located at 22.1 and 43.1 are assigned to the (002) and (100) crystal planes of graphitic carbon, respectively.³⁷ The two weak peaks indicate a low graphitization degree and amorphous features of the porous carbon, agreeing with the TEM images.

To observe the detailed microstructures of the as-prepared sample, SEM and TEM images of the sample were examined, as shown in Fig. 3. Fig. 3(a) exhibits an SEM image of the carbonized sample particles. It shows that the carbon particles are irregular cubes with a size of ~ 20 μm . In the high-magnification SEM photograph of PC-18–600 (Fig. 3(b)), we can see that there are some large pores in the surface of the sample with a size of ~ 5 μm . From the TEM images, many small dark areas having lattice fringes can be clearly seen, indicating

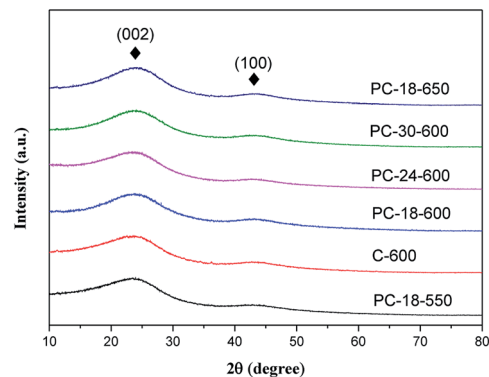


Fig. 2 XRD patterns of the samples.

the formation of many short-range ordered crystalline regions. The lattice spacing is 0.20 nm (Fig. 3(f)), corresponding to the (101) plane of graphite, which has been demonstrated by the literature named “Low temperature carbonization of cellulose nanocrystals for high performance carbon anode of sodium-ion batteries”.³⁸ However, numerous micropores destroyed the graphite structure, making the graphite lattice highly defective and twisty. These defects in the graphite lattice are conducive to microwave absorption because the defects in the crystal lattice can become polarization centers in the presence of microwaves, enhancing the dielectric loss capacity of the material;^{39,40} the

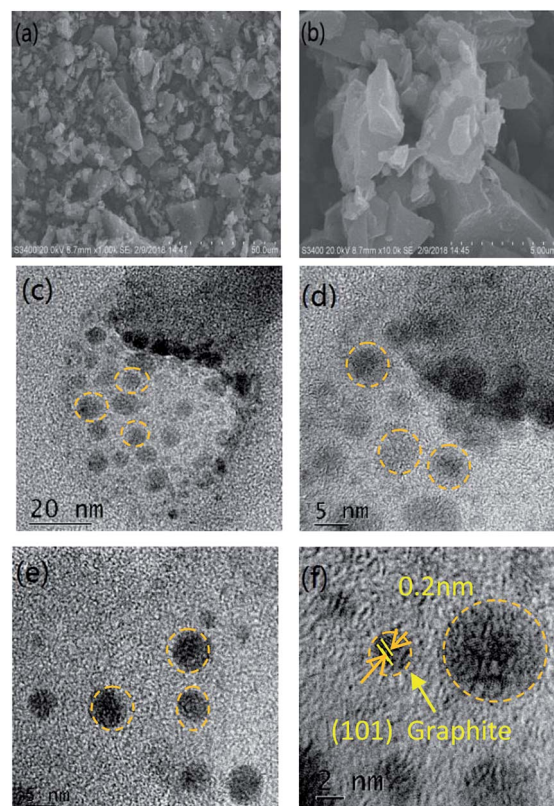


Fig. 3 SEM (a and b) and high-resolution TEM (c–f) images of PC-18–600 at different magnifications.



defects can also make the microwaves scatter around, which disperses the microwave energy.

To further analyse the characteristics of the pore structures of the as-prepared samples, N_2 adsorption–desorption isotherm measurements were obtained, as shown in Fig. 4. Moreover, S_{BET} was calculated by the multiple point BET method; total pore volume was calculated at a relative pressure $P/P_0 = 0.991$ and W_{peak} was estimated according to the Horvath–Kawazoe (HK) theory. The detailed specific surface area, total pore volume and the most probable pore size of the as-prepared samples are listed in Table 1.

As depicted in Fig. 4, the N_2 adsorption–desorption isotherms of the as-prepared samples present similar variation tendencies; the N_2 adsorbing capacity increases rapidly at a very low relative pressure and then, the N_2 adsorbing capacity stays nearly constant with the increase in relative pressure. At the N_2 desorption stage, the desorption isotherm stays almost overlapped with the adsorption isotherm. These N_2 adsorption–desorption isotherms can be classified into I type isotherms according to IUPAC classification, indicating that the as-prepared samples have many micropores. As shown in Table 1, the walnut shell-derived carbon without using $ZnCl_2$ (C-600) has the specific surface area of $435.3 \text{ m}^2 \text{ g}^{-1}$, and the specific surface area of the walnut shell-derived carbon activated by $ZnCl_2$ increases distinctly. For PC-30-600, its specific surface area can reach $640.2 \text{ m}^2 \text{ g}^{-1}$.

For carbonaceous materials, Raman spectra are widely used to characterize their graphitization degree and the defective nature. Generally, carbonaceous materials have two typical characteristic peaks (D band and G band); the D band is attributed to disordered or defective sp^3 -hybridized carbon atoms, while the G band corresponds to the in-plane bond-stretching motion of the pairs of sp^2 -hybridized graphite carbon atoms.^{41–43} Thus, the integrated intensity ratio of the D band to the G band (I_D/I_G) is usually used to estimate the graphitization degree and the defect concentration of carbonaceous materials.^{44,45} Fig. 5 shows the typical Raman spectra of the as-prepared samples.

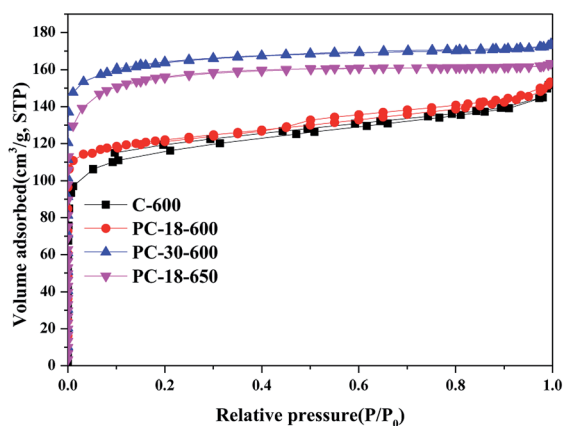


Fig. 4 Nitrogen adsorption–desorption isotherms of samples prepared at different conditions.

Table 1 Specific surface area (S_{BET}), total pore volume, and the most probable pore size (W_{peak}) of samples prepared at different conditions

Sample	S_{BET} ($\text{m}^2 \text{ g}^{-1}$)	Total pore volume ($\text{cm}^3 \text{ g}^{-1}$)	W_{peak} (nm)
C-600	435.3	0.2331	0.54
PC-18-600	481.7	0.2346	0.54
PC-30-600	640.2	0.2674	0.58
PC-18-650	595.7	0.2515	0.56

Among these samples, the D bands at around 1353 cm^{-1} and the G bands at around 1591 cm^{-1} can be clearly seen. The high I_D/I_G values of the samples imply the distortion and many defects in the graphite structures, which indicate the low graphitization degree of the samples; this is in agreement with the TEM images (Fig. 3). Moreover, the calculated I_D/I_G values of the samples decrease (from 1.45 to 1.32) with the increase in calcination temperature (from $550 \text{ }^\circ\text{C}$ to $650 \text{ }^\circ\text{C}$), which indicates the decrease in disordered or defective sp^3 -hybridized carbon domains in the samples with the increase in the calcination temperature.

The surface chemical state of the as-prepared PC-18-600 was characterized by X-ray photoelectron spectroscopy (XPS), as shown in Fig. 5. The general survey spectrum of PC-18-600 presents two typical peaks at about 284.8 eV and 533.4 eV, corresponding to C1s and O1s (Fig. 6(a)). Based on the XPS analysis, the mole ratio of C and O was found to be 13 : 1. The high-resolution spectrum of C1s (Fig. 6(b)) can be decomposed into two peaks. The dominant peak at 284.8 eV is assigned to the graphite carbon and the small peak at 286.1 eV corresponds to C–O.^{41,46} The high-resolution spectrum of O1s (Fig. 6(c)) can be resolved into two peaks located at 532.0 eV and 533.5 eV, corresponding to C=O and C–O bonds.⁴⁷

3.2. Microwave absorption properties

It is widely recognized that the microwave absorption performance of a material is associated with its relative complex permittivity ($\epsilon_r = \epsilon' - j\epsilon''$) and relative complex permeability ($\mu_r = \mu' - j\mu''$). Accordingly, the real parts of relative complex

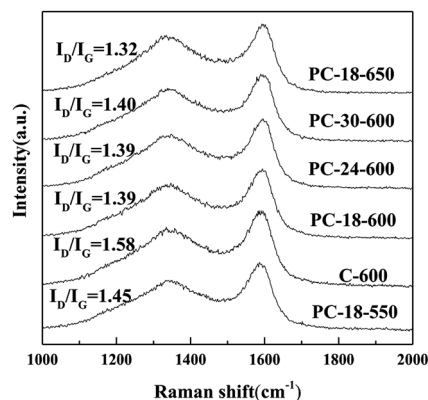


Fig. 5 Raman spectra of samples prepared at different conditions.



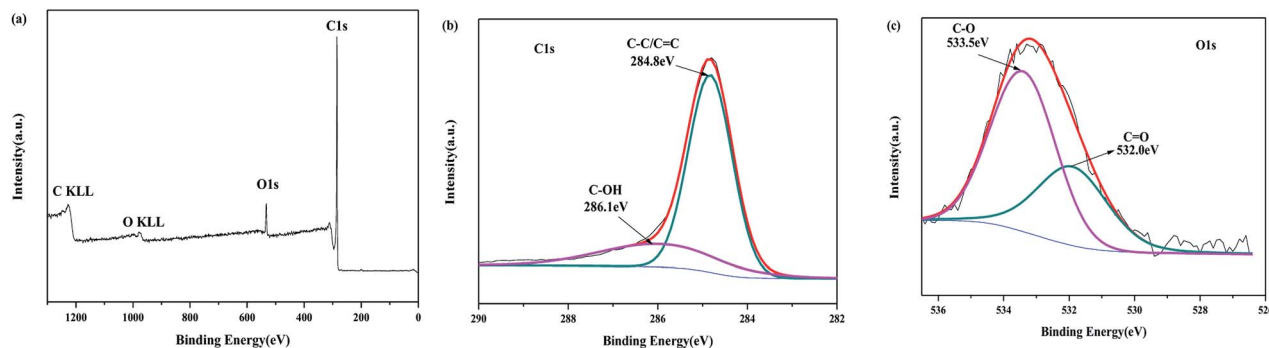


Fig. 6 XPS spectra of PC-18-600: (a) survey scan, (b) C1s and (c) O1s.

permittivity (ϵ') and permeability (μ') reflect the storage capability of electromagnetic energy; the imaginary parts of relative complex permittivity (ϵ'') and permeability (μ'') reflect the dissipation capability of electromagnetic energy.⁴⁸ The dielectric loss tangent and the magnetic loss tangent can be obtained using the equations $\tan \delta_E = \epsilon''/\epsilon'$ and $\tan \delta_M = \mu''/\mu'$, respectively, which can quantify the energy loss *versus* the energy storage.⁴⁹ Hence, the electromagnetic parameters of the as-prepared absorber and paraffin composites were measured, as shown in Fig. 7. As depicted in Fig. 6(a and b), the ϵ' and ϵ'' values of the as-prepared samples all decrease distinctly with the increase in the applied frequency except for PC-18-650; this phenomenon is called the frequency dispersion behaviour, which is beneficial to impedance matching.⁵⁰⁻⁵² In addition, it can be clearly seen that the walnut shells calcined at 600 °C directly (C-600) exhibit low ϵ' values (ranging from 5.2 to 4.1). Interestingly, the walnut shells activated by ZnCl₂ at 600 °C present distinctly higher ϵ' values

than C-600 and the ϵ' values of the as-prepared samples increase with the increase in the ZnCl₂ concentration. This phenomenon can be explained through the microstructure analysis. As depicted in the analysis of N₂ adsorption-desorption isotherms, the samples activated by ZnCl₂ have obviously more micropores than C-600. The effective permittivity of a porous material can be obtained through the Maxwell-Garnett equation:⁵³⁻⁵⁵

$$\epsilon_{\text{eff}}^{\text{MG}} = \left[\frac{(\epsilon_2 + 2\epsilon_1) + 2f_r(\epsilon_2 - \epsilon_1)}{(\epsilon_2 + 2\epsilon_1) - f_r(\epsilon_2 - \epsilon_1)} \right] \epsilon_1 \quad (3)$$

Here, ϵ_1 and ϵ_2 correspond to the permittivities of the host and guest, and f_r stands for the volume fraction of the guest in the effective medium. Theoretically, the porous structure can reduce the permittivity of the material because the permittivity of air (the guest) is far lower than that of the material itself (the host). Nevertheless, when the pores are micropores (<2 nm), they can enhance the permittivity of the material because some

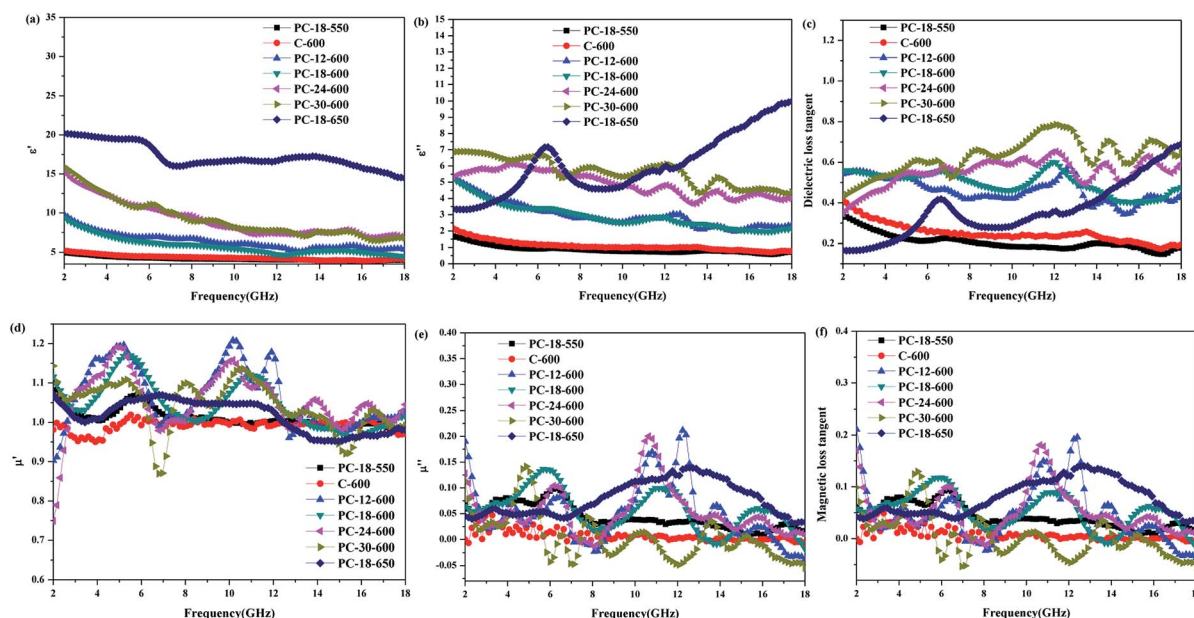


Fig. 7 Frequency dependence of electromagnetic parameters of the composites: (a) the real part (ϵ') and (b) imaginary part (ϵ'') of complex permittivity, (c) the dielectric loss tangent, (d) the real part (μ') and (e) imaginary part (μ'') of complex permeability, (f) the magnetic loss tangent.



electrons accumulate in the walls of the pores to form interfacial polarization in the presence of electromagnetic waves due to differences in the conductivities of the material itself and air inside the material.^{56,57} Furthermore, the ϵ' values (Fig. 7(a)) have high positive correlation with the calculated I_D/I_G values (Fig. 4) of the as-prepared samples. As mentioned above, the I_D/I_G values can effectively reflect the graphitization degree and the defect concentration of carbonaceous materials; the higher I_D/I_G values indicate higher defect concentration in carbonaceous materials. Thus, it can be speculated that the defects in carbonaceous materials make a great contribution to the enhancement in ϵ' values. When the electromagnetic wave propagates into the interior of the material, the defects can become polarization centers, which results in the increase in the ϵ' values.^{39,40}

As depicted in Fig. 7(b and c), C-600 has low ϵ'' values (ranging from 2.15 to 0.79) and dielectric loss tangent (ranging from 0.41 to 0.19). This indicates that C-600 has poor dielectric loss capacity, which decreases the microwave absorption performance. Interestingly, the samples activated by $ZnCl_2$ have distinctly higher ϵ'' values and dielectric loss tangent due to the interfacial polarization and defects inside the material.^{58,59} In addition, it can be seen that the samples activated at higher temperatures have higher ϵ'' values. However, it is worth noting that although PC-18-650 has higher ϵ' and ϵ'' values than the other as-prepared samples, its dielectric loss tangent is relatively low; this indicates that it has good storage capacity of electromagnetic waves but not good dielectric loss capacity, which is unfavourable for its microwave absorption performance.⁶⁰

As indicated in Fig. 7(d-f), the μ' and μ'' values of the as-prepared samples are around 1 and 0, respectively. The low μ' and μ'' values of the as-prepared samples indicate that the samples have poor magnetic loss capacities because the samples are diamagnetic. Therefore, if the ϵ' and ϵ'' values of the as-prepared samples are very high, they are not beneficial to the impedance matching of the samples, which will make a majority of electromagnetic waves reflect on the outer surface of the samples directly.⁶¹

According to eqn (1) and (2), the microwave absorption performances of the as-prepared samples are simulated, as depicted in Fig. 8. It can be seen that the C-600 samples with various thicknesses present poor microwave absorption performances and they do not meet the requirements of effective absorption (the reflection loss ≤ -10 dB, 90% absorption) in the frequency range from 2.0 to 18.0 GHz. The poor microwave absorbing performances of C-600 can be attributed to its poor dielectric loss properties. Nevertheless, the microwave absorbing performances of walnut shells activated by $ZnCl_2$ have dramatic improvements due to the enhanced dielectric loss properties. PC-18-600 with an absorber thickness of 2.5 mm has the minimum reflection loss of -22.8 dB at 14.0 GHz and the effective absorption (the reflection loss ≤ -10 dB) bandwidth is 7.2 GHz (ranging from 10.8 GHz to 18.0 GHz); however, when the absorber thickness is 2.0 mm, its minimum reflection loss is only -17.1 dB at 18.0 GHz and the effective absorption bandwidth is only 2.9 GHz (ranging from 15.1 GHz

to 18.0 GHz). Therefore, the absorber thickness has a great impact on the microwave absorption performance. Usually, it is believed that the more the absorber thickness, the better the microwave absorption performance. In fact, if the absorber thickness is considerable, the microwave absorption performance will be worse. For instance, PC-24-600 with an absorber thickness of 2.0 mm has the minimum reflection loss of -21.0 dB at 13.5 GHz and the effective absorption bandwidth is 6.0 GHz (ranging from 11.4 GHz to 17.4 GHz); however, when the absorber thickness is 2.5 mm, its minimum reflection loss is only -15.7 dB at 9.8 GHz and its effective absorption bandwidth is only 4.7 GHz (ranging from 8.7 GHz to 13.4 GHz). PC-30-600 with an absorber thickness of 2.0 mm has the minimum reflection loss of -17.0 dB at 13.8 GHz and its effective absorption bandwidth is 6.4 GHz (ranging from 11.3 GHz to 17.7 GHz); however, when the absorber thickness is 2.5 mm, its minimum reflection loss is only -14.8 dB at 10.0 GHz and its effective absorption bandwidth is only 3.3 GHz (ranging from 8.6 GHz to 11.9 GHz). In addition, a phenomenon can be found that all the minimum reflection loss peaks of the samples shift to a lower frequency with the increase in the absorber thickness. These phenomena can be elucidated by the $1/4$ wavelength equation:⁶²

$$t_m = \frac{nc}{4f_m \sqrt{\epsilon_r \mu_r}} \quad (n = 1, 3, 5 \dots) \quad (4)$$

Here, t_m and f_m correspond to the matching thickness and frequency at the minimum reflection loss peak; ϵ_r and μ_r stand for the complex permittivity and permeability at f_m , respectively, and c is the velocity of light. Hence, on the one hand, more absorber thickness does not indicate better microwave absorption performance and a reasonable absorber thickness depends on the complex permittivity and permeability of the absorber. On the other hand, the matching frequency or minimum reflection loss peaks of the samples can shift to a lower frequency with the increase in the absorber thickness. As indicated in Fig. 8, it can also be seen that when the absorber thickness is 2.0 mm, PC-24-600 has the minimum reflection loss of -21.0 dB at 13.5 GHz and the effective absorption bandwidth is 6.0 GHz (ranging from 11.4 GHz to 17.4 GHz), while the minimum reflection loss of PC-18-600 is only -17.1 dB at 18.0 GHz and the effective absorption bandwidth is only 2.9 GHz (ranging from 15.1 GHz to 18.0 GHz). When the absorber thickness is 2.5 mm, PC-24-600 has the minimum reflection loss of only -15.7 dB at 9.8 GHz and the effective absorption bandwidth is only 4.7 GHz (ranging from 8.7 GHz to 13.4 GHz), while PC-18-600 has the minimum reflection loss of -22.8 dB at 14.0 GHz and the effective absorption bandwidth is 7.2 GHz (ranging from 10.8 GHz to 18.0 GHz). The PC-18-600 and PC-24-600 samples both have excellent absorbing performances. We further processed the data to pick out the best between them. As shown in Fig. 9, ΔS ($\Delta S = \int RL d|RL| \leq -10$ dB) is defined as the area of effective absorbing ($RL \leq -10$ dB). For PC-18-600 and PC-24-600, the ΔS values are 36.08 and 22.15, respectively. Then, we define R_E ($R_E = \Delta S/d$) as the electromagnetic absorption efficiency.^{63,64} For PC-18-600 and PC-24-600, the R_E values are 14.43 and 11.07, respectively. Thus, it



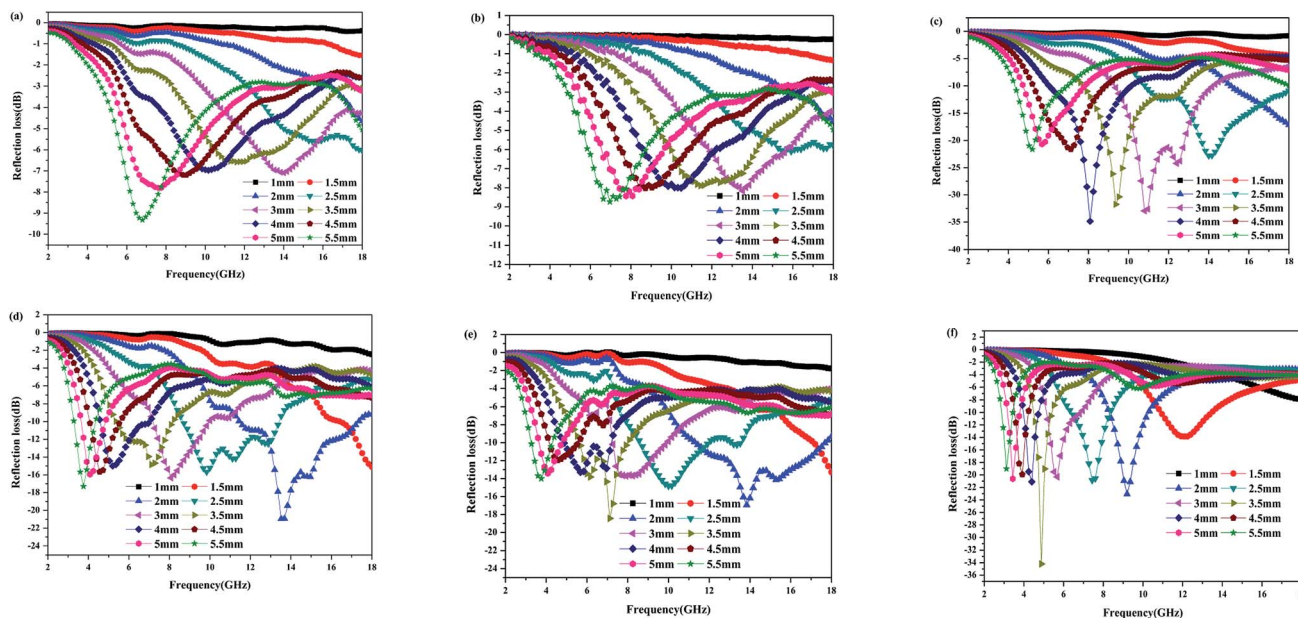


Fig. 8 Reflection loss curves of PC-18-550 (a), C-600 (b), PC-18-600 (c), PC-24-600 (d), PC-30-600 (e) and PC-18-650 (f) at various absorber thicknesses.

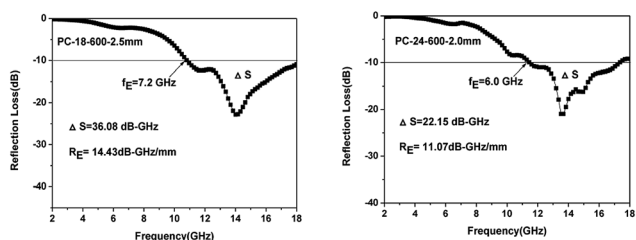


Fig. 9 Reflection loss of PC-18-600 – 2.5 mm and PC-24-600 – 2.0 mm.

can be deduced that the sample showing the best electromagnetic wave absorption properties is PC-18-600.

Furthermore, we discussed the Cole-Cole semicircles and detailed the interfacial polarization relaxation effect of walnut shell-derived nano-porous carbon. We know that the Debye dipolar relaxation is expressed as follows: $(\epsilon' - \epsilon_\infty)^2 + (\epsilon'')^2 = (\epsilon_s - \epsilon_\infty)^2$; here, ϵ_s and ϵ_∞ are the stationary dielectric constant and the optical dielectric constant, respectively. Thus, the plot of ϵ' versus ϵ'' is a single semicircle, which can be defined as the Cole-Cole semicircle.⁶⁵ In Fig. 10, we can easily see that PC-18-600 presents a clear segment of three semicircles, suggesting the existence of ternary dielectric relaxation processes, and each semicircle corresponds to a Debye dipolar relaxation.

Representative carbon-based materials with excellent microwave absorption performances in recent years are listed in Table 2. Rice husk-based hierarchically porous carbon (RHPC)/Fe and RHPC/Co were prepared by a pre-modification method and a post-modification method; the minimum reflection loss of RHPC/Fe was -21.8 dB at 14.0 GHz and its effective absorption bandwidth was 5.6 GHz (ranging from 12.4 GHz to 18.0 GHz) when the absorber thickness was only 1.4 mm; the

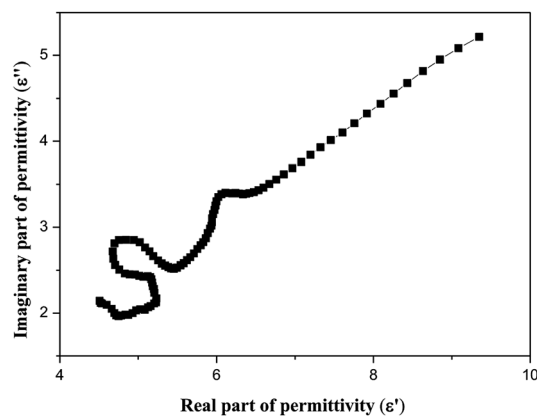


Fig. 10 Cole-Cole semicircles for PC-18-600.

minimum reflection loss of RHPC/Co was -40.1 dB at 10.7 GHz and its effective absorption bandwidth was 2.7 GHz (ranging from 9.3 GHz to 12.0 GHz) when the absorber thickness was only 1.8 mm.²⁸ Wood-based porous biomass carbon was prepared by the pyrolysis of natural wood. The natural wood annealed at 680 °C has the minimum reflection loss of -68.3 dB at 11.0 GHz and its effective absorption bandwidth is 6.13 GHz (ranging from 8.20 GHz to 14.33 GHz) when the absorber thickness is 4.28 mm; the natural wood annealed at 690 °C has the minimum reflection loss of -16.3 dB at 11.5 GHz and its effective absorption bandwidth is 7.63 GHz (ranging from 9.83 GHz to 17.46 GHz) when the absorber thickness is 3.73 mm.²¹ A Co@crystalline carbon@carbon aerogel composite was prepared by using alginate aerogels as the precursors; it exhibited the minimum reflection loss of -43 dB at 17.9 GHz and its effective absorption bandwidth was 2.6 GHz (ranging



Table 2 The representative carbon-based materials with excellent microwave absorption performances in recent years

Absorbing agent	Matrix	Minimum RL value (dB)	Minimum peak position (GHz)	Thickness (mm)	Frequency range (GHz) (RL \leq -10 dB)	Efficient bandwidth (GHz) (RL \leq -10 dB)	Ref.
Walnut shell-based porous carbon	Paraffin wax	-22.8	14.0	2.5	10.8–18.0	7.2	This work
Walnut shell-based porous carbon	Paraffin wax	-21.0	13.5	2.0	11.4–17.4	6.0	This work
Rice husk-based porous carbon/Fe	Paraffin wax	-21.8	14.0	1.4	12.4–18.0	5.6	28
Rice husk-based porous carbon/Co	Paraffin wax	-40.1	10.7	1.8	9.3–12.0	2.7	28
Wood-based porous carbon	Paraffin wax	-68.3	11.0	4.28	8.20–14.33	6.13	21
Wood-based porous carbon	Paraffin wax	-16.3	11.5	3.73	9.83–17.46	7.63	21
Nitrogen-doped graphene aerogel	Paraffin wax	-53.2	13.10	3.3	9.0–17.1	8.1	66
Reduced graphene oxide/multi-walled carbon nanotubes/zinc ferrite hybrid composites	Paraffin wax	-22.2	17.4	1	15.7–18.0	2.3	67
Zn ferrite/multi-walled carbon nanotubes composite	Paraffin wax	-42.6	12.1	1.5	11–14	3	6
CoZn alloy/N-doped porous carbon nanocomposites	Paraffin wax	-49.0	12.5	2	10.0–15.3	5.3	68
Silver/carbon fiber/polyaniline composites	Paraffin wax	-13.2	9.3	2	8.4–10.6	2.2	69
NiFe ₂ O ₄ /reduced graphene oxide composite	Paraffin wax	-27.7	9.2	3	8–10	2	70

from 15.4 GHz to 18.0 GHz) with the absorber thickness of 1.5 mm. In this work, the minimum reflection loss of PC-18-600 is -22.8 dB at 14.0 GHz and its effective absorption bandwidth can reach 7.2 GHz (ranging from 10.8 GHz to 18.0 GHz) with an absorber thickness of 2.5 mm; the minimum reflection loss of PC-24-600 is -21.0 dB at 13.5 GHz and its effective absorption bandwidth can reach 6.0 GHz (ranging from 11.4 GHz to 17.4 GHz) with an absorber thickness of only 2.0 mm. Compared with other carbonaceous microwave absorbers, the as-prepared samples in this work have wide effective absorption bandwidths with low absorber thicknesses, which improves their practicability.

In the end, probable microwave absorption principles of walnut shell-derived nano-porous carbon need to be generalized. On the one hand, the material structure has a great influence on the microwave absorption performance of the material. Porous structures are conducive to microwave absorption because microwaves will make multiple reflections and scatter in the porous structure and the propagation path is prolonged, which strengthens the microwave attenuation capacity of the material. On the other hand, the microwave absorption performance of the material is determined by its electromagnetic parameters. Walnut shell-derived nano-porous carbon is diamagnetic; thus, the dielectric loss capacity of the material plays the leading role in affecting the microwave absorption performance of the material. Two factors make great contributions to the improvement in the dielectric loss capacity of the material. On the one hand, the defects in the walnut shell-derived nano-porous carbon can become polarized centers

when the microwave transmits to the sample, which enhances the dielectric loss capacity of the material. On the other hand, some electrons accumulate in the walls of the pores to cause interfacial polarization when the microwave transmits to the sample, which can also enhance the dielectric loss capacity of the material. However, the walnut shell-derived nano-porous carbon is diamagnetic and it has low complex permeability values; thus, if the dielectric loss capacity of the material is very strong, it will deteriorate the impedance matching of the material, which indicates that majority of microwaves will reflect on the outside surface of the material directly rather than propagate into the interior of the material.

4. Conclusion

In conclusion, a walnut shell-derived nano-porous carbon was successfully prepared *via* a facile method with the assistance of ZnCl₂. Walnut shells calcined at 600 °C directly (C-600) revealed poor microwave absorption performance because of the poor dielectric loss capacity. Nevertheless, when the walnut shells were activated by ZnCl₂, the specific surface areas, dielectric loss capacities and microwave absorption performances of the samples improved significantly. Two factors can account for the improvement in the microwave absorption performances of the samples activated by ZnCl₂. On the one hand, the specific surface area of the sample increased distinctly *via* ZnCl₂ activation and due to the porous structures, the microwaves exhibited multiple reflections and scattering, which can strengthen the microwave absorption performance. On the



other hand, the dielectric loss capacity of the sample was enhanced *via* ZnCl₂ activation, which was beneficial to microwave attenuation. The reinforcement of the dielectric loss capacity was attributed to the interfacial polarization in the walls of the pores and the increased defects, which can become polarized centers in the sample when the microwave transmits to the sample. In this work, for PC-18–600, the minimum reflection loss could reach –22.8 dB at 14.0 GHz and the effective absorption (the reflection loss \leq –10 dB) bandwidth could reach 7.2 GHz (ranging from 10.8 GHz to 18.0 GHz) when the absorber thickness was 2.5 mm. Hence, the as-prepared samples can have low densities, strong microwave absorption performances and wide effective absorption bandwidths with low absorber thicknesses, due to which they have a great potential as lightweight and efficient microwave absorbers. Furthermore, the ZnCl₂ activation method in this work can be extended as a universal method to improve the dielectric loss capacity and the microwave absorption performance of carbonaceous materials.

Conflicts of interest

There are no conflicts to declare.

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