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Acetylacetone functionalized magnetic carbon microspheres for the highly-efficient adsorption of heavy metal ions from aqueous solutions

Jie Ma, * Mengya Sun, Yulan Zeng, Zhenhua Liu, Manman Zhang, Yao Xiao and Shuping Zhang

Herein, Acac-C@Fe₃O₄, a magnetic carbon (C@Fe₃O₄) modified with acetylacetone (Acac), was first prepared and used as a solid-phase adsorbent for adsorbing some heavy metal ions from aqueous solution. The adsorbent was characterized by Fourier transform infrared spectroscopy (FTIR), X-ray diffraction (XRD), scanning electron microscopy (SEM), transmission electron microscopy (TEM), X-ray photoelectron spectroscopy (XPS), vibrating sample magnetometry (VSM) and BET studies. Some parameters affecting the adsorption and desorption processes were studied in Pb²⁺ solution, including sample pH, contact time, initial concentration, type and connection of the desorption solution. Adsorption results showed that removal of Pb²⁺ was 100% under optimal conditions at an initial concentration of 10.0 mg L⁻¹. The adsorption mechanism conformed well to a pseudo-second order kinetic model. The adsorption capacity of the sorbent also showed promising results with Hg²⁺, Cr³⁺, Fe²⁺, Cd²⁺, Mn²⁺, Zn²⁺, Cu²⁺ and Pb²⁺, where maximum adsorption capacities reached 98.0, 151.2, 188.9, 202.2, 286.3, 297.2, 305.1 and 345.3 mg g⁻¹, respectively. The Acac-C@Fe₃O₄ microsphere material was successfully applied to the adsorption of heavy metal ions in aqueous solution.

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Introduction

With the growth of industrial activities, the phenomenon of highly toxic heavy metal ions release to water resources becomes more and more serious.¹ Heavy metals are difficult to biodegrade and can be enriched hundreds of times during biological amplification of the food chain before entering the human body. Eventually, they will make proteins and enzymes inactive and also cause chronic poisoning.² Thus, it becomes even more necessary to establish a facile and reliable method to deal with environmental pollution of heavy metal ions. Over decades, several technologies have been developed to remove heavy metal ions from wastewaters such as chemical precipitation (including hydroxide precipitation,³ sulphide precipitation⁴ and chelating precipitation⁵), ion exchange,⁶ membrane filtration (including ultrafiltration,⁷ reverse osmosis and nanofiltration⁸ and electrodialysis⁹), coagulation and flocculation,¹⁰ flotation,¹¹ electrochemical treatment¹² and adsorption.¹³ Adsorption methods draw more attention, because they are simple to pursue without additional consumption, high performing and cost-effective.

With all of the absorbing materials, magnetic nanomaterials are different from others as they can be collected by external

magnetic fields.¹⁴ This performance provides both adsorption and desorption post-processing for convenience. Meanwhile, low-cost nanosized ferric oxides (Fe₃O₄ or γ-Fe₂O₃) possess non-toxicity and superparamagnetism and therefore they have gained even more interest.¹⁵ Aggregation of these particles can be regulated and controlled by a surface modified process and so far abundant materials have been used as a modified layer including silicon, carbon, synthetic polymers and metals. Among these materials, carbon-based materials are popularly selected because they are easily synthesized and eco-friendly, especially when they have numerous hydrophilic groups and can provide favourable conditions for further modification.¹⁶ Wang *et al.* prepared two different structures of carbon-coated Fe₃O₄ for lithium storage; both composite materials have good performance and represent a good future for applications. There are also many applications of Fe₃O₄@C nanomaterials in the environmental treatment fields.¹⁷ Liu *et al.* designed 1D Fe₃O₄/C/CdS coaxial nanochains for the degradation of RhB.¹⁸ Bystrzejewski *et al.* prepared carbon-encapsulated magnetic nanoparticles to adsorb Cu²⁺, Co²⁺, and Cd²⁺ and their adsorption capacities were 3.21, 1.23 and 1.77 mg g⁻¹, respectively.¹⁹ Chen *et al.* synthesized magnetic core-shell Fe₃O₄@C nanoparticles with -SO₃H/-COOH groups modified for adsorbing Pb²⁺, Hg²⁺ and Cd²⁺ ions and their maximum adsorptions were 90.7, 83.1 and 39.7 mg g⁻¹, respectively.²⁰ The above-mentioned carbon coated magnetic nanomaterials could be used to adsorb heavy metal ions, but their adsorption

College of Science, University of Shanghai for Science and Technology, Shanghai 200093, P. R. China. E-mail: majie0203ch@hotmail.com; Fax: +86 21 55271663; Tel: +86-13482433498

capacities are limited, which may be improved by some new methods.

In order to improve adsorption properties of adsorbents, some chelate molecules, especially multidentate ligands, were selected to modify nanomaterials, such as ethylene diamine tetraacetic acid (EDTA), ethanediamine (EDA) and acetylacetone (Acac), which showed strong coordination with metal ions. Among them, the Acac molecules are superior to other molecules because their active hydrogen and the dicarbonyl structures can be retained during the modification process.²¹ In this work, the magnetic carbon composite (C@Fe₃O₄) was synthesized by a hydrothermal method and functionalized with Acac for adsorbing heavy metal ions in an aqueous solution. Behaviours of the adsorption in different analytical conditions and performance under optimized conditions were investigated by FAAS determination. Meanwhile, the adsorption capacities of the as-prepared composites for Cu²⁺, Zn²⁺, Mn²⁺, Cd²⁺, Fe²⁺, Hg²⁺ and Cr³⁺ were measured and had ideal results. Thus the Acac modified C@Fe₃O₄ microspheres (Acac-C@Fe₃O₄) may be used in the metal ion wastewater treatment.

Experimental

Materials

All reagents were of analytical grade including FeCl₃·6H₂O, Pb(NO₃)₂·6H₂O, FeSO₄·7H₂O, CuSO₄·5H₂O, ZnCl₂, HgCl₂, CrCl₃·6H₂O, MnSO₄, CdCl₂·2.5H₂O, NaOH, CH₃COONa, ethylene glycol (C₂H₆O₂), D(+)-glucose monohydrate (C₆H₁₂O₆·H₂O), ethanediamine (C₂H₈O₂), (3-chloropropyl)triethoxysilane (C₉H₂₁ClO₃Si, KH230), acetylacetone (Acac), and HCl solution and all were purchased from Sinopharm Chemical Reagent Corporation (Shanghai, China). A target stock solution of Pb²⁺ with 1000.0 mg L⁻¹ was prepared by dissolving proper amounts of Pb(NO₃)₂·6H₂O and nitric acid in distilled water. Then the solution was further diluted to the given concentrations for testing adsorption capacity and analysing the pre-concentration process.

Instrumentation

The morphology of a sorbent was detected by a field emission scanning electron microscope (SEM, VEGA3 TESCAN, CZ) and a transmission electron microscope (TEM, H800EM Hitachi, JP). Surface area was measured by a BET instrument (TriStar 3020 Micromeritics, CN). XPS spectra were measured by an X-ray photoelectron spectroscope (XPS, Thermo ESCALAB 250XI, CN). A Fourier transform infrared spectrophotometry analysis was performed using a FTIR spectrophotometer (FTIR, Nicolet, USA) equipped with a KBr beam splitter. X-ray diffraction patterns of the samples were recorded by an X-ray powder diffractometer (XRD, D8 advance, DE). Magnetic properties of the sorbent were tested using a vibrating sample magnetometer (VSM, EZ-VSM, MicroSense, USA) at room temperature. Concentrations of metal ions were measured using a flame atomic absorption spectrophotometer with air-acetylene (FAAS, TAS990, China). A digital pH meter was used for pH adjustment of the solutions.

Preparation of Acac-C@Fe₃O₄

The synthetic process is shown in Scheme 1. First, a solvothermal method was adopted for synthesizing Fe₃O₄ magnetic composites according to a previously reported procedure.²² In a typical procedure, 36 mmol of FeCl₃·6H₂O and 200 mmol of CH₃COONa were dissolved in 350.0 mL ethylene glycol. The mixture was sonicated for 30 min and then stirred at room temperature for 2 h. Afterward, the homogeneous solution was transferred into a 500 mL Teflon-lined stainless-steel autoclave and reacted at 200 °C for 10 h followed by cooling down to room temperature naturally. The product was washed with ethanol and distilled water for several times, and then dried in a vacuum freeze-drying apparatus for 5 h.

Secondly, C@Fe₃O₄ was prepared by a facile hydrothermal method.^{23,24} Concretely, 0.5 g Fe₃O₄ nanoparticles were dispersed in 0.1 mol L⁻¹ HNO₃ solution and sonicated for 10 min, then were rinsed with distilled water several times until neutral. Subsequently, 8.2 g glucose was dissolved in 80.0 mL distilled water, followed by the addition of 0.5 g Fe₃O₄ under sonication. The pH value of the mixture was adjusted to 8.0 by a 0.1 mol L⁻¹ NaOH solution. Afterwards, the mixture solution was transferred into a Teflon-lined stainless-steel autoclave (100 mL capacity) and heated at 170 °C for 4 h. The precipitate was washed with distilled water for several times and dried in a vacuum freeze-drying apparatus for 5 h, and then the carbon coated Fe₃O₄ nanomaterials were obtained.

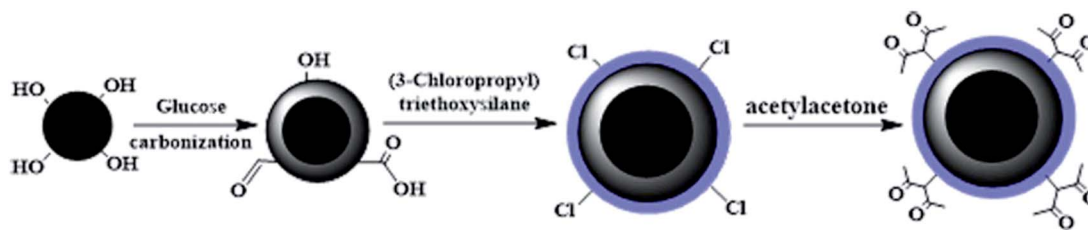
Thirdly, 1.0 g C@Fe₃O₄ was dissolved in 100.0 mL water-ethanol under mechanical stirring. Then, 1.0 mL ammonia (25%) and 3.0 mL KH230 were added dropwise to the reaction mixture under constant stirring at 45 °C. After 1 h, the reaction mixture was heated to 60 °C and stirred constantly for 1 h, and then cooled to room temperature. Finally, the magnetic samples (KH230-C@Fe₃O₄) were isolated by using an external magnet, washed thoroughly with distilled water and ethanol several times, and then dried in a vacuum freeze-drying apparatus.

Finally, 1.0 g NaOH was dissolved in 10.0 mL distilled water and then was added to 40.0 mL ethanol in an ice-water bath, 10.0 mL Acac was added into the mixture, followed by 1.0 g KH230-C@Fe₃O₄ which was dispersed into the suspension under constant stirring for 0.5 h. Then the mixture was heated to 33 °C for 7 h. The sample was separated using an external magnetic field and washed several times with distilled water and ethanol and dried in a vacuum freeze-drying apparatus. Finally, Acac-C@Fe₃O₄ composites were prepared successfully.

Adsorption procedure

Aqueous solutions with different Pb²⁺ concentrations were prepared and adjusted to a given pH value by HCl and NaOH, and then 5.0 mg sorbent was dispersed into the above solution (25.0 mL). The time dependence of the adsorption capacity was also examined as a function of contact time from 1 to 8 min at room temperature using 25.0 mL of Pb²⁺ solution (20.0 mg g⁻¹) mixed with 5.0 mg of sorbent. The Pb²⁺ concentration of a sample was measured by FAAS. The adsorption capacity of the material for Cu²⁺, Zn²⁺, Mn²⁺, Cd²⁺, Fe²⁺, Hg²⁺ and Cr³⁺ were researched under the same conditions. The adsorption process is shown in Fig. 1.





Scheme 1 Diagram of the Acac-C@Fe₃O₄ magnetic material synthesis process.

Result and discussion

Characterization of the magnetic microspheres

VSM analysis. Fig. 2a depicts the magnetization hysteresis loops of typical samples including Fe₃O₄, C@Fe₃O₄ and Acac-C@Fe₃O₄ microspheres at room temperature. The curves indicated that all prepared samples had superparamagnetic properties with negligible magnetic remanence and coercivity. Saturation magnetization of samples decreased from 30 to 14 emu g⁻¹ with an increase of coating layer which showed that the obtained samples had enough magnetic response characteristics to meet magnetic separation process.

Meanwhile, the coercivities of samples were almost stable below 20 Oe. The above suggested the modified layer did not affect the superparamagnetic feature of the composites, but just reduced the saturation magnetization of a sample due to the decreased content of Fe₃O₄ component in the microspheres.

X-ray diffraction analysis. The XRD patterns of Fe₃O₄, C@Fe₃O₄ and Acac-C@Fe₃O₄ magnetic particles are presented in Fig. 2b. The similar diffraction peaks of all patterns, including 30.1°, 35.5°, 43.0°, 53.6°, 57.0° and 62.6°, indicated the crystal structures of all composites were coincided with a Fe₃O₄ standard specimen as a cubic spinel structure (JCPDS, no. 19-0629).²⁵ The X-ray diffraction patterns of C@Fe₃O₄ and Acac-C@Fe₃O₄ were similar with the pattern of Fe₃O₄ magnetic particles, which illuminated that they had the same crystal structure and it was not changed during the modification procedures.

FTIR analysis. Composition of the Acac-C@Fe₃O₄ composite was further characterized by FT-IR spectrophotometry, and the relevant results are shown in Fig. 2c. The specific stretching vibration absorption peak of Fe–O in Fe₃O₄ was obviously visible in four spectra at 586 cm⁻¹, which indicated all samples

contained Fe₃O₄. The peaks at 1623 cm⁻¹ and 1689 cm⁻¹ were attributed to C=C and C=O vibrations and illustrated that carbon was successfully coated on the magnetic core. Meanwhile the absorption peak at 3434 cm⁻¹ implied the existence of residual hydroxyl groups, which confirmed that some hydrophilic groups (–OH) and carbonyl group (–C=O) remained after the process of hydrothermal carbonization treatment.^{26–28} The peak at 802 cm⁻¹ shown in Fig. 2c indicated the existence of C–Cl groups and suggested that KH230 were grafted on the surface of C@Fe₃O₄ composites. Curves of Acac-C@Fe₃O₄, KH230-C@Fe₃O₄ and Fe₃O₄ are compared and shown in Fig. 2c and the enhancement of a peak at 1623 cm⁻¹ is attributed to the carbonyl stretching vibration²⁹ while the peak at 2915 cm⁻¹ indicated the existence of –CH₃ of Acac. These FTIR data indicated that Acac successfully functionalized the C@Fe₃O₄ magnetic microspheres.

BET analysis. The ratio surface areas of Fe₃O₄ and Acac-C@Fe₃O₄ were analysed by nitrogen adsorption–desorption techniques. As Fig. 2d shows, the nitrogen adsorption–desorption isotherms of Fe₃O₄ and Acac-C@Fe₃O₄ have representative type-IV curves. The hysteric loop in the range of 0.6–0.9 *P*/*P*₀ is a type of H1 hysteresis. The BET specific surface area of Acac-C@Fe₃O₄ was measured to be 68.41 m² g⁻¹, which is higher than that of Fe₃O₄ (37.41 m² g⁻¹). This change might be



Fig. 1 Schematic diagram of the adsorption process.

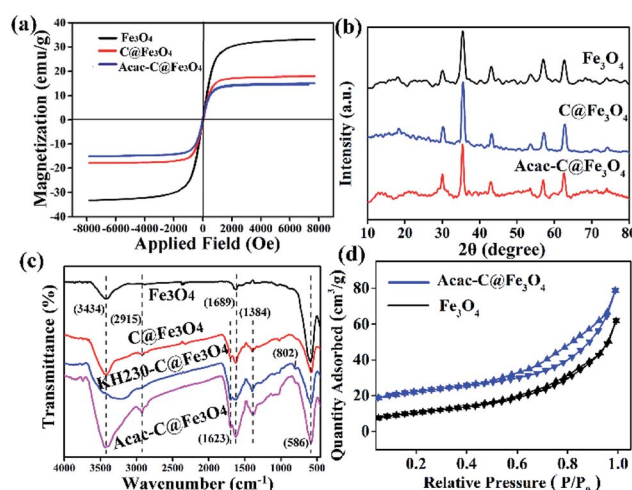


Fig. 2 (a) Magnetization curves of typical prepared samples: Fe₃O₄, C@Fe₃O₄ and Acac-C@Fe₃O₄. (b) XRD patterns of Fe₃O₄, C@Fe₃O₄ and Acac-C@Fe₃O₄. (c) FT-IR spectra of Fe₃O₄, C@Fe₃O₄, KH230-C@Fe₃O₄ and Acac-C@Fe₃O₄. (d) Nitrogen adsorption–desorption isotherm of Fe₃O₄ and Acac-C@Fe₃O₄.



attributed to the porous structure of Acac-C@Fe₃O₄ coming from carbon coating.

XPS analysis. The surface element composition of Acac-C@Fe₃O₄ was determined by XPS spectra. As shown in Fig. 3a, the elements of Fe, C, O, and Si were detected on the surfaces of the nanocomposites. The signals of Fe and Si were weaker than that of C, indicating that the Fe₃O₄ core is coated completely by the carbon shell and a small amount of KH230 was modified. The high-resolution XPS spectrum of Fe 2p can be fitted into two main doublet peaks, as shown in Fig. 3b. In the high-resolution spectrum, it can be seen that the doublet peak at 710.8 eV and 724.2 eV are assigned to Fe 2p_{3/2} and Fe 2p_{1/2}, which can be attributed to Fe(II) and Fe(III) of the Fe₃O₄ core. These results were consistent with the XRD analysis. In Fig. 3c, there are more than one chemical state of C in the C 1s region; there peaks were identified and assigned to C-C and C-H (284.8 eV), C-O (286.5 eV), and C=O species (288.4 eV), respectively. These may be ascribed to a carbon shell and acetylacetone.³⁰ The XRD results and previous conclusions were combined and provide supportive evidence for the successful modification of carbon and Acac.

SEM and TEM analysis. Morphology and structures of the Fe₃O₄, C@Fe₃O₄ and Acac-C@Fe₃O₄ microspheres were characterized by scanning electron microscopy and transmission electron microscopy. Fig. 4 indicated the as-prepared Fe₃O₄ materials were monodispersed microspheres, which were further proved by the relative TEM image of Fe₃O₄ shown in Fig. 5. The average size of a microsphere is about 260 nm. Fig. 4 also confirmed that the Acac-C@Fe₃O₄ sample still kept spherical morphology with a broader size distribution. The particle distribution range increased slightly, and the average of size of a microsphere increased to 400 nm. The increase of particle sizes and edges of C@Fe₃O₄ and Acac-C@Fe₃O₄ looked like a light shadow layer should be attributed to the carbon coating on the surface of the magnetic core, which was confirmed directly by the TEM image in Fig. 5. This further indicated that the Acac-C@Fe₃O₄ composite was prepared successfully.

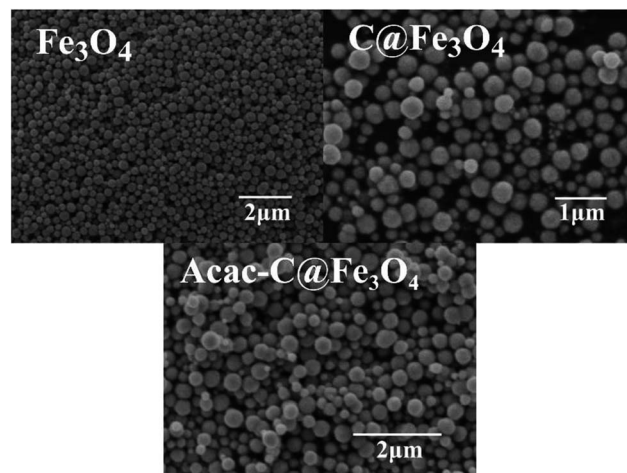


Fig. 4 SEM images of Fe₃O₄, C@Fe₃O₄ and Acac-C@Fe₃O₄.

Optimization of pre-concentration conditions

Effect of pH. Acidity of a solution shows strong influence on the adsorption capacity of the adsorbent. To detect influence of the acidity of a solution, Pb²⁺ was used as the target ion, and the absorbing process was performed at 25 °C for about 20 minutes. The pH of the solution was set in a range from 2.0 and 10.0, and concentrations and amounts of solutions were given at 20 mg L⁻¹ and 25.0 mL in turn with the same amounts of sorbent. The relevant results are shown in Fig. 6a. The curves obviously indicated the absorptivity of C@Fe₃O₄ was more than Fe₃O₄ microsphere in a set range and this may be due to the porous structure of the carbon layer where there are rich oxygenic groups on the surface. The absorptivity of Acac-C@Fe₃O₄ was higher than C@Fe₃O₄. The absorptivity of Acac-C@Fe₃O₄ was raised from 25% to 90% with pH values from 2.0 to 6.0 and it slightly declined when the pH changed from 6.0 to 9.0 so the absorptivity was obviously decreased with a pH beyond 9.0. The above indicated that the proper pH range should be controlled between 6.0 and 9.0 for better performance of the adsorbent. Compared with the three curves, the result should be attributed to Acac modified on the surface of C@Fe₃O₄ composites. Acac has convertible enolyl and dicarbonyl structures with the change in acidity. It is beneficial to resist acidity, which helps to form coordination with lead ions.²¹ However, the maximum absorptivity of Fe₃O₄ and C@Fe₃O₄

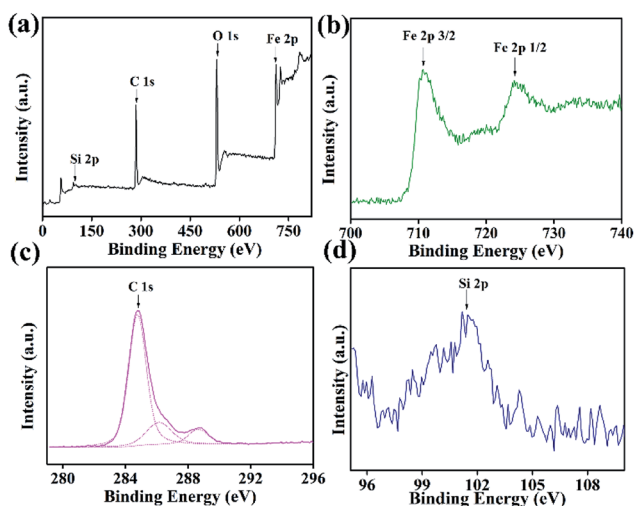


Fig. 3 XPS spectra of Acac-C@Fe₃O₄.

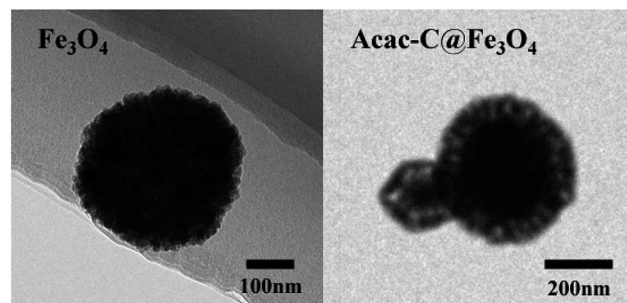


Fig. 5 TEM images of Fe₃O₄ and Acac-C@Fe₃O₄.



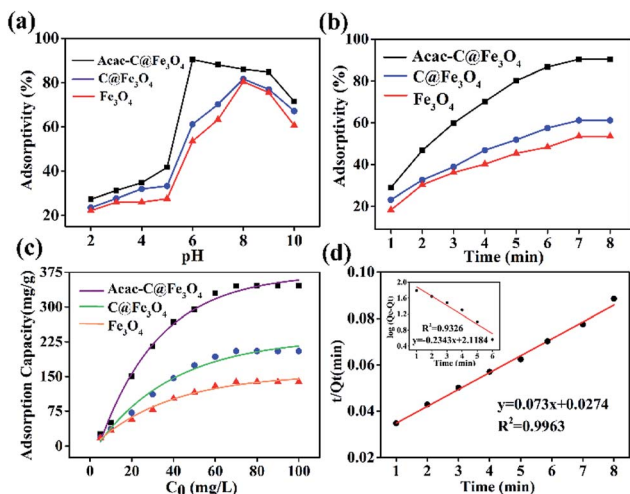


Fig. 6 Effect of pH (a), adsorption time (b), and initial concentration of Pb^{2+} (c) on the adsorption of Pb^{2+} onto Acac-C@ Fe_3O_4 , C@ Fe_3O_4 and Fe_3O_4 ; (d) pseudo-second-order kinetic model for Pb^{2+} adsorption on Acac-C@ Fe_3O_4 sorbent (inset: pseudo-first-order kinetic model curve).

microspheres *ca.* 80% appeared at an acidity around 8.0. Since Pb^{2+} ions can precipitate under basic conditions, this cannot indicate the as-prepared Fe_3O_4 microsphere are a suitable lead ion adsorbent although the adsorption capacity is higher in an alkaline setting.

Effect of adsorption time. Adsorption time is an important factor in the application of the as-prepared sorbent. Fig. 6b demonstrates the results about adsorptivity *versus* absorbing time at given conditions including the pH at 6.0, temperature at 25 °C and concentration of Pb^{2+} at 20 mg L^{-1} . These three curves indicated that the adsorptivity increased rapidly within 7 min but after that it remained almost unchanged. The adsorption equilibrium of the absorbing system was reached after around 7 minutes, which suggests the Acac-C@ Fe_3O_4 adsorbent can achieve the maximum adsorption efficiency within 7 minutes in the testing system, and the adsorption rate of lead ions under the conditions investigated was more than 90%. Relatively, the adsorption rate of lead ions was 61% absorbed by the C@ Fe_3O_4 microsphere and only 54% by naked Fe_3O_4 at same time. The above indicated that the as-prepared Acac-C@ Fe_3O_4 adsorbent had excellent capacity for lead ions, so it could reach the absorption equilibrium quickly. It is beneficial to use the sorbent in the field of separation of heavy metal ions. As the initial concentration of lead ions was raised, the adsorption amount of Acac-C@ Fe_3O_4 sorbent also increased gradually and reached to a maximum value greater than 345.3 mg g^{-1} with a lead ion concentration at 70 mg L^{-1} , thus basically remaining unchanged. Meanwhile, the adsorption capacity of C@ Fe_3O_4 and Fe_3O_4 microspheres only reached 205.0 mg g^{-1} and 138.4 mg g^{-1} , respectively. The maximum adsorption amount of Acac-C@ Fe_3O_4 was obviously higher than C@ Fe_3O_4 and the naked Fe_3O_4 compared to the three curves. This indicated the saturated adsorption capacity of Acac-C@ Fe_3O_4 can be acquired with a lead ion concentration beyond

70.0 mg L^{-1} . Because all available adsorption sites of as-prepared adsorbent were occupied by ions in a high concentration ion solution, it cannot achieve more amounts of ionic adsorption.³¹ This suggested the absorption process should belong to chemical absorption.

Adsorption kinetics study. An adsorption isotherm expresses the relationship between adsorption capacity and time equilibrium. The experimental data were obtained by detecting the absorbing amount with a series of given absorbing times with the concentration of Pb^{2+} ions at 70 mg L^{-1} . The corresponding data were calculated according to two Langmuir isotherm model formula including pseudo-first-order (1) and pseudo-second-order eqn (2):³²

$$\log(Q_e - Q_t) = \log Q_e - k_1 t / 2.303 \quad (1)$$

$$\frac{t}{Q_t} = \frac{1}{k_2 Q_e^2} + \frac{t}{Q_e} \quad (2)$$

where Q_e and Q_t refer to the adsorption amount of Pb^{2+} at equilibrium and time t , respectively, and k_1 , k_2 are the rate constants of pseudo-first-order and pseudo-second-order.

According to the model formula, two linear fitting curves and equations are separately shown in Fig. 6d. The analysis results suggested that the pseudo-second-order model ($R^2 = 0.9915$) provided a better description of the Pb^{2+} adsorption isotherm process than the pseudo-first-order model based on the fitted correlation coefficients (R^2) 0.996 for the second order and 0.933 for the first order model. These results further suggested the absorbing process belongs to a chemical process.

Adsorption of other heavy metal ions. Due to Acac having a strong coordination capacity with many metal ions, the sorption capacity of the as-prepared adsorbent for various common metal ions (Hg^{2+} , Cr^{3+} , Fe^{2+} , Cd^{2+} , Mn^{2+} , Zn^{2+} and Cu^{2+}) were investigated in parallel under the same conditions. The concentration of all them was given at 70 mg L^{-1} with the same optimal conditions as the Pb^{2+} adsorbing procedures. All data are shown in Fig. 7. This showed that the absorbing capacities of the Acac-C@ Fe_3O_4 sorbent on Hg^{2+} , Cr^{3+} , Fe^{2+} , Cd^{2+} , Mn^{2+} , Zn^{2+} and Cu^{2+} were 98.0, 151.2, 189.9, 202.2, 286.3, 297.2 and 305.1 mg g^{-1} , respectively. These results illustrated that the Acac-C@ Fe_3O_4 composites can be applied as a highly efficient sorbent for absorbing some common metal ions, which was beneficial for enriching some metal ions in aqueous solution. In order to measure the quality of the prepared adsorbents, saturated adsorbents were compared with some reported metal ion adsorbents and are shown in Table 1. These data indicated that the as-prepared Acac-C@ Fe_3O_4 adsorbent presented outstanding advantages in adsorption capacity which should be attributed to the functionalization of Fe_3O_4 microspheres by Acac and carbon.

Regeneration of the sorbent

Selection of a desorption solution. Regeneration property and method are important parameters to evaluate an adsorbent. Regeneration of the Acac-C@ Fe_3O_4 composite materials was investigated. First, the type of desorption solution was



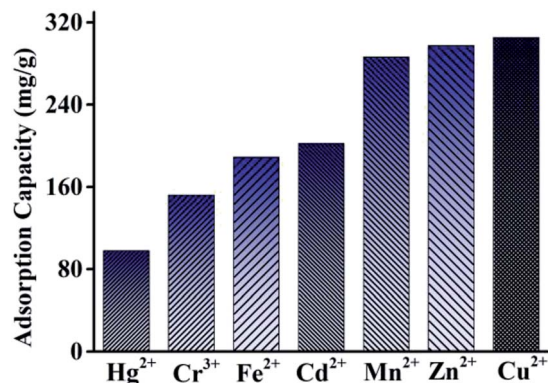


Fig. 7 The adsorption capacity of the as-prepared sorbent for other heavy metal ions.

selected from four kind of common solutions (H_2SO_4 , HNO_3 and HCl at 1 mol L^{-1} and EDTA at 0.5 mol L^{-1}). They were used to desorb Pb^{2+} from the magnetic sorbent and the desorption efficiencies of them are shown in Fig. 8a. This proved that a HCl solution was the most suitable for use for desorption because it had the highest effective desorption among those tested. On this basis, the influence of the concentration of a HCl solution was further explored and the data are shown in Fig. 8a. With an increase of the HCl concentration desorption efficiency increased as the solution concentration changed from 0.2 to 2.0 mol L^{-1} , and then the desorption efficiency decreased slightly. The above results indicated that desorption solution acidity was an important factor for desorbing lead ion from the as-prepared magnetic sorbent. The coordination sites with Pb^{2+} were more and more occupied by H^+ with increased concentration of the HCl solution, so high acidity was an advantage to regenerate the sorbent. However, when concentration of desorption solution was beyond 2.0 mol L^{-1} , the desorption performance could keep a high level, but the colour of the desorption solution changed to pale yellow, which might be due to Fe^{3+} ion being generated and suggesting that $\text{Acac-C@Fe}_3\text{O}_4$ materials were damaged in a high acidic setting. So, the concentration of the HCl solution used as a desorption solution should be controlled between 1.0 and 2.0 mol L^{-1} when desorption efficiency of $\text{Acac-C@Fe}_3\text{O}_4$ sorbent is beyond 95% .

Meanwhile, the effects of HCl solution volume and desorption time on desorption were also studied. The desorbed efficiency was obviously raised with the increase of HCl solution volume under 50.0 mL as shown in Fig. 8b. There was no change

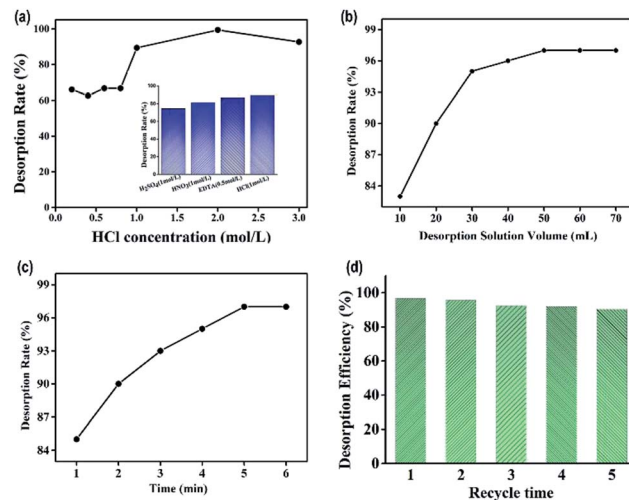


Fig. 8 (a) Influence of the concentration of HCl solution and the kind of desorption solution (inset) on the desorption efficiency; the influence of desorption solution volume (b) and desorption time (c) on the desorption efficiency; (d) desorption and regeneration of the as-prepared sorbent.

when the volume was more than 50.0 mL . Hence, 50.0 mL desorption solution was used in subsequent research. The desorbed time also affects the desorption process. Fig. 8c showed the recovery of the lead ions was up to 97% when the desorbed time was 5 min . And then, the desorption efficiency kept stable even if time was increased. So we choose 5 min as optimum desorption.

Reuse studies. The reusability of a sorbent is crucial in practical applications. Fig. 8d indicated the results from five time's consecutive adsorption–desorption cycles of Pb^{2+} on the $\text{Acac-C@Fe}_3\text{O}_4$ microspheres. The desorption efficiency still maintained beyond 90% after 5 times which clearly indicated that absorption capacity just decreased slightly and this decrease of sorption efficiency may be ascribed to loss of adsorbents in a recycled application. The above experimental results suggest our as-prepared $\text{Acac-C@Fe}_3\text{O}_4$ microspheres sorbent possess fine reusability.

Conclusions

An $\text{Acac-C@Fe}_3\text{O}_4$ magnetic adsorbent was successfully prepared by a simple and sustainable procedure and its structure and properties were detected *via* numerous instrumental

Table 1 Comparison of the adsorption capacities of different adsorbents for the removal of heavy metal ions

Adsorbent	Adsorption capacity of metal ions (mg g^{-1})							Pb^{2+}	Ref.
	Hg^{2+}	Cr^{3+}	Fe^{2+}	Cd^{2+}	Mn^{2+}	Zn^{2+}	Cu^{2+}		
$\text{Acac-C@Fe}_3\text{O}_4$	98.0	151.2	188.9	202.2	286.3	297.2	305.1	345.3	This work
PEI-magnetic porous	—	—	—	105.2	—	138.8	157.8	—	33
$\text{Fe}_3\text{O}_4\text{@C-SO}_3\text{H/-COOH}$	83.1	—	—	39.7	—	—	—	90.7	20
$\text{Fe}_3\text{O}_4\text{@APS@AA-co-CA}$	—	—	—	29.6	—	43.4	126.9	166.1	34
Magnetic PSSQ hybrid	—	—	—	5.2	13.3	—	30.0	66.1	35
$\text{Poly(MMA-co-MA)/Fe}_3\text{O}_4$	—	90.91	—	—	—	111.11	109.89	—	36
Coir fibers	—	—	11.11	—	—	—	9.43	24.91	37



measurements. The adsorbent demonstrated outstanding adsorption capacity for some metal ions. Under given conditions, the maximum adsorption capacities for Hg^{2+} , Cr^{3+} , Fe^{2+} , Cd^{2+} , Mn^{2+} , Zn^{2+} , Cu^{2+} and Pb^{2+} were 98.0, 151.2, 188.9, 202.2, 286.3, 297.2, 335.1 and 345.3 mg g^{-1} , respectively. The adsorption process reached adsorption equilibrium within 7 min and conforms to a pseudo-second-order kinetics model. The adsorbent had excellent stability and reusability. The $\text{Acac-C@Fe}_3\text{O}_4$ magnetic adsorbent may provide an alternative for applications in the adsorption and separation of heavy metal ions in aqueous solutions. What's more, the material can be used as a substrate and in the further synthesis of metal-organic frameworks.

Conflicts of interest

There are no conflicts to declare.

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