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Hydrothermal synthesis of MoS_2 nanosheet loaded TiO_2 nanoarrays for enhanced visible light photocatalytic applications

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A molybdenum disulfide (MoS_2) nanosheet-decorated titanium dioxide (TiO_2) NRA heterojunction composite was fabricated successfully through a two-step hydrothermal approach. Microstructures and optical properties of specimens were characterized by field-emission scanning electron microscopy, X-ray diffractometry, X-ray photoelectron spectroscopy, and ultraviolet-visible spectroscopy. The gaps of the TiO_2 nanorods have been filled with tiny MoS_2 nanosheets, which can increase the surface area of MoS_2/TiO_2 NRA composite thin films. In addition, the photocatalytic activity of the thin films were measured and discussed in greater detail. The appropriate hydrothermal reaction temperature of MoS_2 is important for the growth of perfect MoS_2/TiO_2 NRA composites with significantly enhanced photocatalytic performance. The photodegradation rate and *k* value of MoS_2-220/TiO_2 are 86% and 0.0105 min^{-1}, respectively, which are much larger than those of blank TiO_2. The enhanced photocatalytic performance could be attributed to the higher visible light absorption and the reduced recombination rate of photogenerated electron–hole pairs.

1. Introduction

In recent years, a large and growing literature has investigated the photocatalytic degradation and solar energy conversion to provide clean fuels[1–3] for addressing the energy crisis and environmental pollution. Titanium dioxide (TiO_2) as a semiconductor photocatalytic material has been broadly investigated due to its high photocatalytic activity, chemical stability, low cost and nontoxicity.[4] In order to improve the photocatalytic activities, many efforts have been made for the fabrication of various morphologies of TiO_2, such as hollow spheres, hierarchical microspheres, nanorods, nanotubes and porous networks.[5–7] However, the fast recombination of photogenerated electron–hole pairs and large band gap (3.2 eV) of TiO_2 limit the use of sunlight as well as photocatalytic activity. In order to overcome these drawbacks, rational designs of TiO_2-based heterojunction photocatalysts with desirable architectural structure/morphology have been put into effect.[8–11] It is well-known that two-dimensional MoS_2 is a potential candidate for enhancing the visible light absorption and charge separation ability. Recently, MoS_2/TiO_2 heterostructures with diverse morphologies have been reported by researchers and have shown high photocatalytic activities in the visible light region.[12–14] However, most of the MoS_2/TiO_2 heterostructures were powder materials, which were either difficult to recycle or would require additives in order to prepare the electrodes. In order to overcome these drawbacks, researchers made great efforts to prepare the film-based heterostructure composites with enhanced photocatalytic performances. Zheng *et al.*[15] prepared the hierarchical MoS_2 nanosheet@TiO_2 nanotube arrays *via* anodization and hydrothermal methods. Photocatalytic activity of the compositing MoS_2@TiO_2 hybrids can be enhanced by means of appropriate modification of MoS_2 on TiO_2 nanotube arrays. To the best of our knowledge, the studies on the synthesis of MoS_2 nanosheet coating on the TiO_2 nanoarrays are limited.

Herein, we demonstrate a facile two-step hydrothermal method to prepare MoS_2/TiO_2 NRA composite thin films with a highly-ordered three-dimensional hierarchical structure. The unique hierarchical structure of MoS_2/TiO_2 NRA heterojunctions was propitious to improve the light harvesting and charge separation rates. The novel MoS_2/TiO_2 NRAs exhibit enhanced photocatalytic activity due to the lower recombination rate and higher visible light absorption compared with the pure TiO_2 NRAs.

2. Experimental

A simple two-step hydrothermal approach was adopted to prepare MoS_2 nanosheet-loaded highly ordered TiO_2

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nanoarrays (NRAs). TiO_2 NRAs were synthesized *via* the hydrothermal method. Then, 0.68 mL of tetrabutyl titanate was mixed with 20 mL of hydrochloric acid and 20 mL of deionized water (DI water). The mixed solution was magnetically stirred for 30 min and transferred into a 50 mL Teflon-lined stainless steel autoclave. Subsequently, a clean FTO glass was placed vertically in the autoclave, which was heated at 150 °C for 10 hours and then cooled to room temperature. The obtained TiO_2 NRAs were washed thoroughly with DI water and dried at 60 °C for several hours.

The preparation procedure of the MoS_2/TiO_2 NRA composite thin film is depicted as follows. 0.12 g thioacetamide (C_2H_5NS) and 0.06 g sodium molybdate ($Na_2MoO_4 \cdot 2H_2O$) were dissolved in 40 mL DI water and magnetically stirred for 30 min to form a transparent solution. The mixed solution was transferred into a 50 mL-Teflon-lined stainless steel autoclave. The TiO_2 NRAs were placed vertically in the autoclave, which was heated at 200 °C and 220 °C for 24 h. The MoS_2/TiO_2 NRA composite thin film was obtained after washing the product with DI water and drying at 60 °C for 12 h.

Field-emission scanning electron microscopy (FE-SEM, S-4800) was adopted to observe the topologies of the samples. An X-ray diffractometer (XRD) was used to record the diffraction patterns of the samples. Chemical compositions of the samples were analyzed *via* X-ray photoelectron spectroscopy (XPS, ESCALAB 250). The structural properties of the samples were examined using a laser micro-Raman spectrometer (Renishaw in Via-Reflex) at an excitation wavelength of 532 nm. Optical properties of the samples were characterized by an ultraviolet-visible spectrophotometer (UV-2550, Shimadzu). The photocatalytic performances of the samples were estimated by the degradation of methylene blue (MB) solution. A thin film (25 × 25 mm) was placed in a glass culture dish containing 10 mL MB solution (5 ppm), which was chosen as the simulated contaminant. Then, the glass culture dish was placed under the visible-light irradiation from a Xenon lamp with a 420 nm cut-off filter. The photocatalytic performance can be estimated by comparing the absorbance of MB at a certain light irradiation with that of the initial MB solution at a characteristic wavelength of 664 nm.

3. Results and discussion

Fig. 1 presents the top-view FE-SEM images of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2 on an FTO substrate. As shown in the FE-SEM image in Fig. 1a, the TiO_2 NRAs are well-aligned perpendicular to the FTO substrate. The nanorods have an average diameter of ~100 nm. In MoS_2-200/TiO_2 (Fig. 1b), many tiny MoS_2 nanosheets exist on the side and top facets of TiO_2 nanorods. As the growth temperature of MoS_2 increases to 220 °C, the gap of the TiO_2 nanorods fills with tiny MoS_2 nanosheets, which form the well-stacked layered structure of MoS_2.

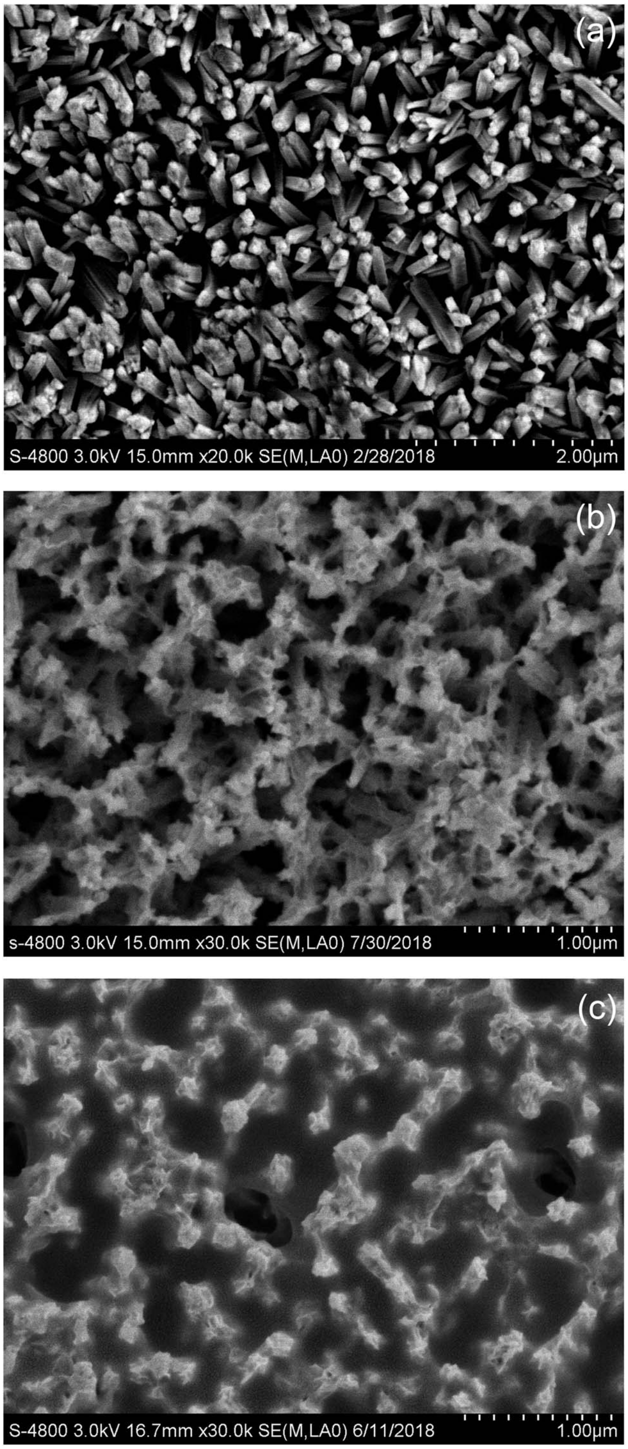

Fig. 1 Top-view FE-SEM images of (a) TiO_2 NRAs, (b) MoS_2-200/TiO_2, and (c) MoS_2-220/TiO_2.

Fig. 2 presents the XRD patterns of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2. The diffraction peaks for the TiO_2 NRAs match the standard peaks of rutile TiO_2 phase (JCPDS card no. 65-0191) with the lattice parameters of $a = b = 0.4593$ nm and $c = 0.2958$ nm. Five characteristic peaks at $2\theta =$ 36.08°, 54.33°, 62.77°, 69.01° and 69.82° can be respectively assigned to the (101), (211), (002), (301) and (112) crystal facets of rutile TiO_2. With regard to MoS_2/TiO_2 NRA composite thin films, all diffraction peaks of the rutile TiO_2 phase still exist in the XRD patterns. These results indicate that the TiO_2 NRAs are well maintained in the hydrothermal process. Furthermore, the intensity of the diffraction peaks for rutile TiO_2 of MoS_2-200/

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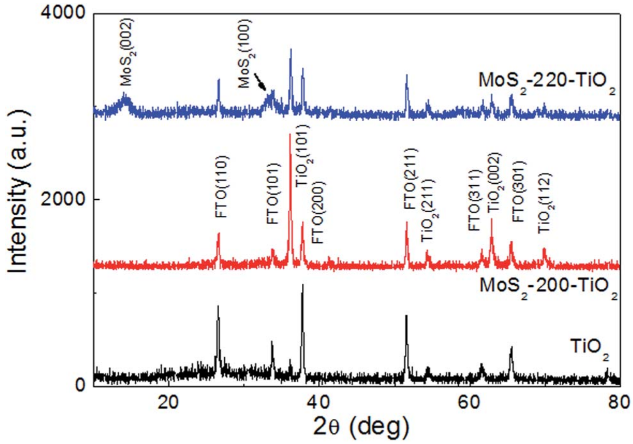

Fig. 2 XRD patterns of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2.

TiO_2 is stronger than that of rutile TiO_2, which may be attributed to the recrystallization of rutile TiO_2 in the second step of the hydrothermal process. The decrease in the diffraction peak intensities for TiO_2 of MoS_2-220/TiO_2 may be due to the shielding effect of the large amounts of MoS_2 loaded onto TiO_2. No diffraction peak for MoS_2 has been observed in MoS_2-200/TiO_2. Absence of the diffraction peak indicates that the modificatory MoS_2 nanosheets on the TiO_2 NRAs may only contain a spot of MoS_2 nanosheets, which cannot be detected by XRD. Three characteristic peaks at $2\theta = 14.40°$, $33.52°$ and $58.35°$ have been observed in MoS_2-220/TiO_2, which can be indexed to (002), (101) and (110) crystal facets of MoS_2 (JCPDS card no. 65-7025), respectively. The stronger (002) diffraction peak that appeared in the XRD pattern of the sample indicates the formation of a well-stacked layered structure of MoS_2.[16]

In order to obtain the detailed topography of the binary composite thin film, TEM images of the MoS_2-220/TiO_2 composites have been recorded. Fig. 3a and b represent the typical transmission electron microscopy (TEM) images of the MoS_2-220/TiO_2 composite. It can be seen that the tiny MoS_2 nanosheets cover the surface of TiO_2 nanorods. Fig. 3c shows the high-resolution transmission electron microscopy (HRTEM) image of the composite thin film. The lattice fringes that are perpendicular to the growth direction have a lattice spacing of 0.32 nm, which is equal to the lattice parameter in the (110) facet of the rutile TiO_2. The lattice spacing is 0.62 nm, which corresponds to the (002) plane of MoS_2.

As shown in Fig. 4, the survey XPS spectra show that the dominant elements in MoS_2-220/TiO_2 are Ti, O, S, Mo, and C (Fig. 4a). The high-resolution XPS spectra of Mo and S for MoS_2-220/TiO_2 are shown in Fig. 4b and c. The binding energies at 232.2 and 229.0 eV with a separation energy close to 3.2 eV are ascribed to the $Mo3d_{3/2}$ and $Mo3d_{5/2}$ of the Mo element in the tetravalent state.[17] In addition, the binding energy at 235.9 eV is ascribed to the $Mo(3d_{3/2})$ signal of the Mo element in the hexavalent state, which results from the partial surface oxidation of MoS_2 in air.[16] As shown in Fig. 4c, the S2p spectrum can be deconvoluted into two peaks. The binding energies at 163.3 and 161.9 eV are attributed to the $S(2p_{1/2})$ and $S(2p_{3/2})$ signals, respectively, of sulphide element in the divalent state. As shown in Fig. 4d, the C1s spectrum can be deconvoluted into three peaks, which correspond to C═O (288.4 eV), C–O (286.4 eV), and C–C (284.8 eV), respectively.[18] As shown in Fig. 4e, the spin-energy separation between $Ti2p_{1/2}$ (464.3 eV) and $Ti2p_{3/2}$ (458.5 eV) peaks is about 5.8 eV, which is assigned to the Ti^{4+} state in the pure TiO_2 NRAs.[19] The binding energies of $Ti2p_{1/2}$ (465.0 eV) and $Ti2p_{3/2}$ (459.2 eV) for MoS_2-220/TiO_2 shift to higher energies because of the existence of Ti^{3+} states by trapping electrons,[16] which exhibit higher activity and form the radicals covalently on the surface of TiO_2. These trapped charge carriers can be transferred to the conduction or valence band by thermal excitation, which is beneficial to enhance the photocatalytic abilities.[20]

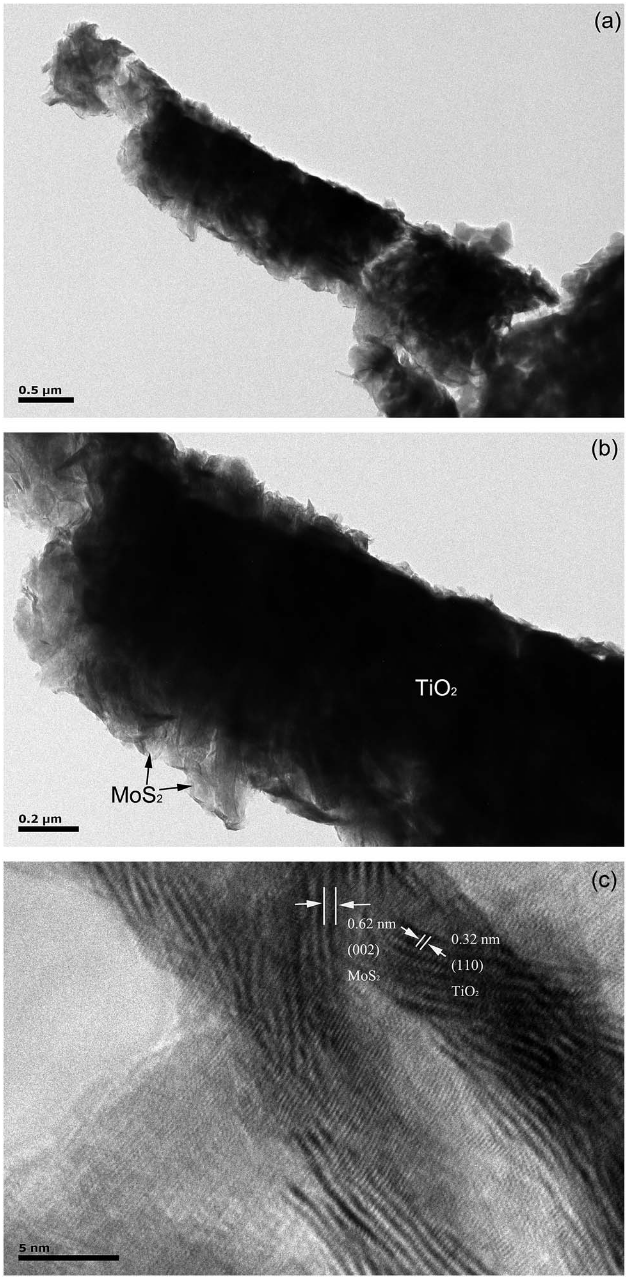

Fig. 3 (a and b) TEM images of MoS_2-220/TiO_2 composites. (c) High-resolution TEM (HRTEM) images of pristine TiO_2 NRAs.

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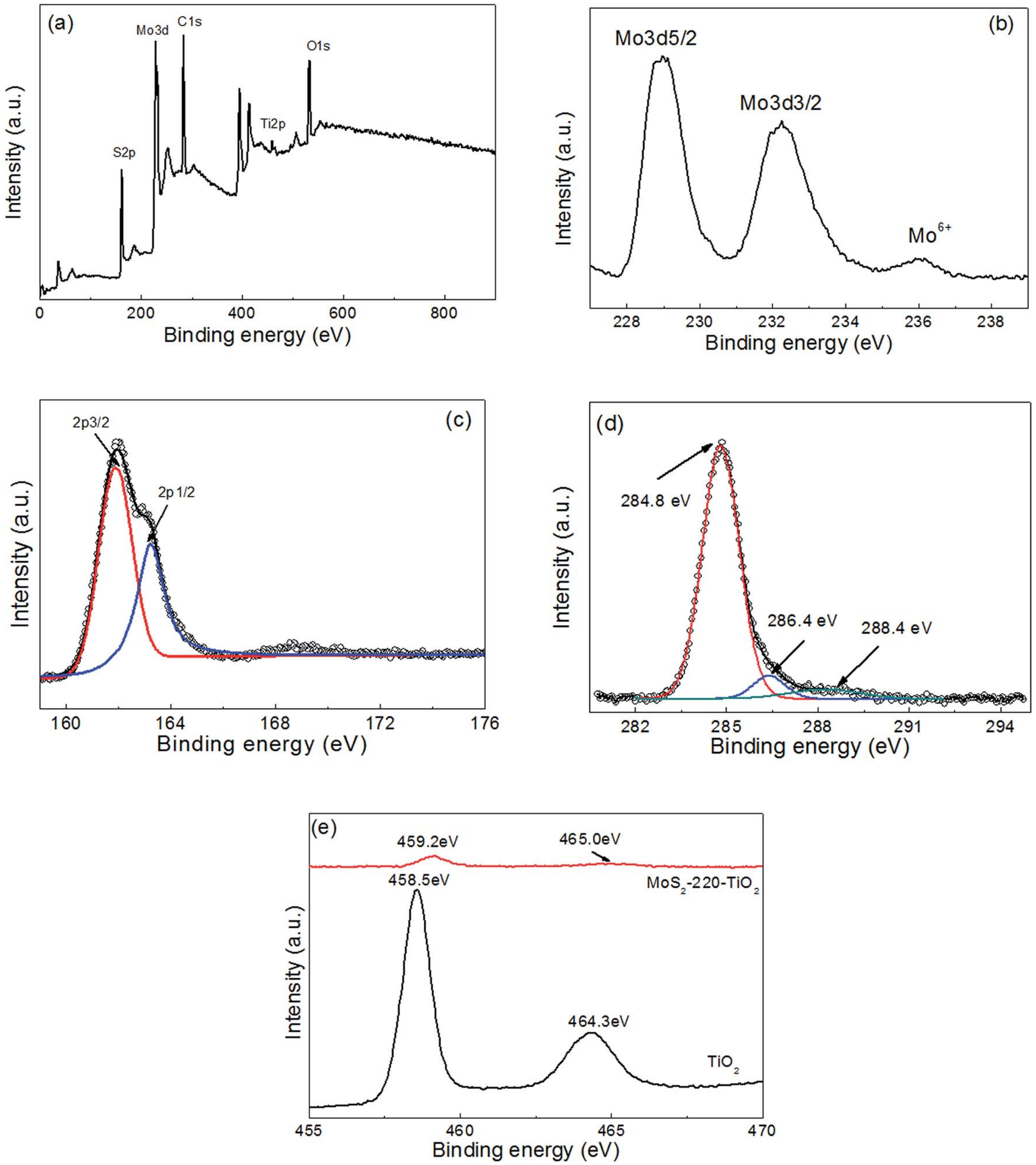

Fig. 4 XPS survey spectra of (a) MoS_2-220/TiO_2, (b) Mo3d, (c) S2p, (d) C1s, and (e) Ti2p.

Fig. 5 shows the Raman spectra of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2. It can be seen from the Raman curve of bare TiO_2 NRAs that the two resonance peaks centered at 445 and 608 cm^{-1} can be attributed to the B_{1g} and E_g of TiO_2 with rutile phase, respectively.[21] There are three resonance peaks centered at 404, 445 and 608 cm^{-1}, which can be observed from MoS_2-200/TiO_2, which may be attributed to the out-of-plane $E_{2g}{}^1$ of MoS_2 and the B_{1g} and E_g of TiO_2, respectively.[22] These results indicate that a spot of MoS_2 nanosheets deposited on the TiO_2 NRAs. The Raman curve of MoS_2-220/TiO_2 shows two strong resonance peaks centered at 375 and 404 cm^{-1}, which may be ascribed to in-plane A_{1g} and out-of-plane $E_{2g}{}^1$ of MoS_2, respectively. The other two weak resonance peaks centered at 445 and 608 cm^{-1} may be assigned to B_{1g} and E_g of TiO_2, respectively, which weakened with the increase in MoS_2 nanosheet content.

Fig. 6 shows the UV-vis absorption spectra of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2. It can be seen that the spectrum of TiO_2 NRAs shows an abrupt absorption edge at about 400 nm and strong absorption in UV region. For MoS_2-200/TiO_2 and MoS_2-220/TiO_2, absorbance in the visible region has been enhanced due to the visible light absorbance arising from MoS_2, which possesses the considerably narrowed band gap. These results indicated that the TiO_2 NRAs were

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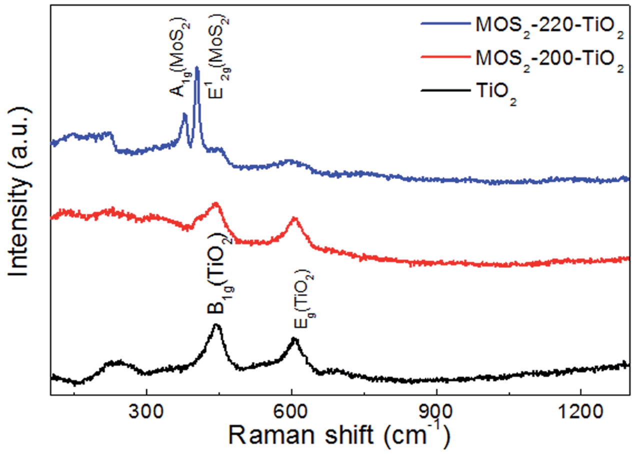

Fig. 5 Raman spectra of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2.

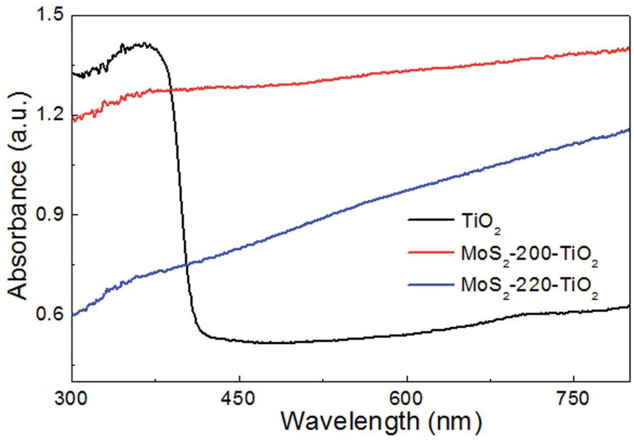

Fig. 6 UV-vis absorbance spectra of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2.

successfully modified by MoS_2 nanosheets. The enhanced visible light absorbance of MoS_2/TiO_2 can improve the visible photocatalytic activity of the samples.[21] It can also be seen that the UV absorption is greatly dependent on the amount of MoS_2 loaded onto TiO_2. This may be attributed to the shielding effect of MoS_2 on TiO_2, which reduces the UV absorption of TiO_2.[23]

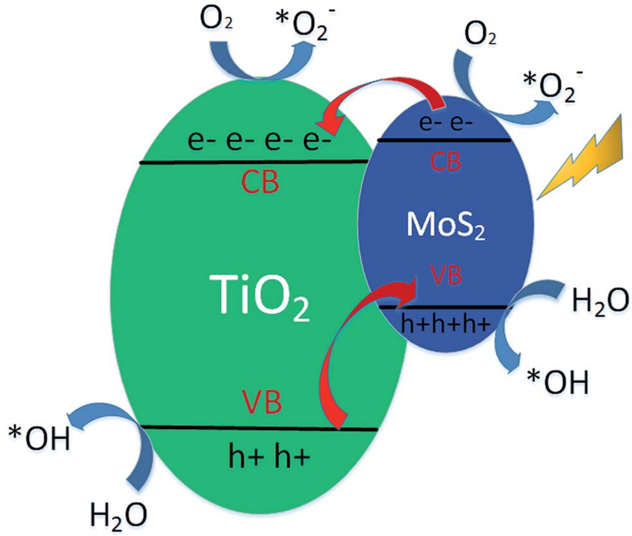

Fig. 8 Schematic of the energy band structure and separation principle of carriers in MoS_2-220/TiO_2.

The photocatalytic performances of TiO_2 NRAs, MoS_2, MoS_2-200/TiO_2 and MoS_2-220/TiO_2 were estimated by means of the photodegradation of MB under visible light. Fig. 7a shows the photodegradation rate of the samples. It is evident that the photodegradation rate of MoS_2-220/TiO_2 is far greater than that of TiO_2 NRAs, MoS_2, and MoS_2-200/TiO_2. The photodegradation rate of TiO_2 NRAs, MoS_2, MoS_2-200/TiO_2 and MoS_2-220/TiO_2 is about 44%, 53%, 57% and 86%, respectively, after visible light irradiation for 180 min.

The photodegradation of MB solution follows the pseudo-first-order reaction and its kinetics can be expressed as follows:[17]

$$\ln(C_0/C) = kt \tag{1}$$

where C_0 is the initial concentration of MB after adsorption equilibrium, C is the concentration after t min reaction, and k is the rate constant. The k values of TiO_2 NRAs, MoS_2, MoS_2-200/

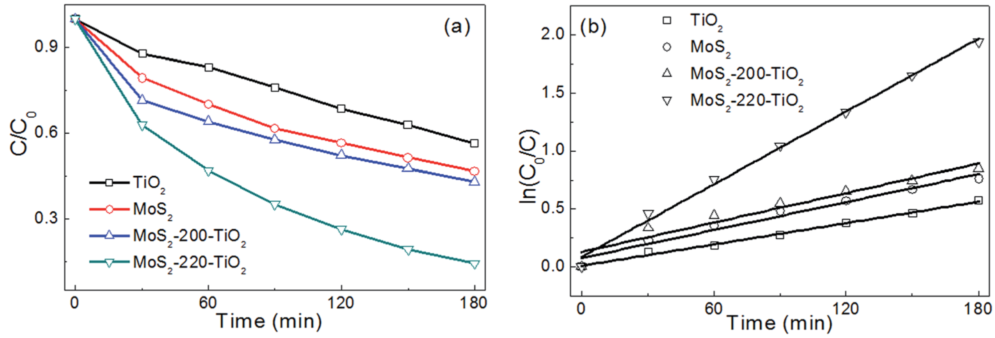

Fig. 7 (a) Normalized absorption of MB in the presence of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2 under visible light irradiation. (b) The plots of $\ln(C_0/C)$ *versus* t of TiO_2 NRAs, MoS_2-200/TiO_2 and MoS_2-220/TiO_2.

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TiO_2 and MoS_2-220/TiO_2 can be calculated from Fig. 7b and are 0.0031, 0.0040, 0.0042, and 0.0105 min^{-1}, respectively. The k value of MoS_2-220/TiO_2 is almost 3.4 times greater than that of TiO_2 NRAs. These results indicate that the photocatalytic performances of TiO_2 NRAs can be considerably enhanced by the introduction of net-like MoS_2.

For the as-deposited TiO_2 NRAs, the photocatalytic degradation under visible light irradiation may be attributed to the existence of abundant defects, which will induce defect levels and narrow band-gaps and then improve the visible light absorption.[24] The schematic of the energy band structure and separation principle of carriers in MoS_2-220/TiO_2 is shown in Fig. 8. First, a small amount of electrons can be excited from the defect levels (or valence band) to the conduction band when TiO_2 is illuminated by visible light. Second, it is well known that MoS_2 is a narrow bandgap semiconductor with a band gap of about 1.8 eV. Upon visible light illumination, the electrons in the valence band (VB) of MoS_2 can be excited to conduction band (CB), thus leaving holes in the VB. As the CB of TiO_2 is lower than that of MoS_2, the photogenerated electrons in the CB of MoS_2 can be transferred to the CB of TiO_2 NRAs.[17] TiO_2 is considered to be an electron receiver, which is conducive to the separation of charge carriers. The electrons may be captured by oxygen (O_2) to form superoxide radical anions ($^{\bullet}O_2^-$), which can react with a simulated pollutant. In addition, the holes left in the VB of MoS_2 can react with H_2O to form hydroxyl radicals, which can react with the dye on the sample.[25] Therefore, the electron and hole can be separated efficiently through the formation of a heterojunction between MoS_2 and TiO_2. Finally, the photocatalytic activities of MoS_2/TiO_2 could be enhanced by introducing MoS_2 onto the TiO_2 NRAs.

4. Conclusions

A novel MoS_2 nanosheet decorating TiO_2 NRA heterojunction composite with outstanding photocatalytic performance for the degradation of organic pollutants under visible light illumination was deposited by a hydrothermal method on the FTO substrate. SEM, XRD and Raman results confirmed that MoS_2 nanosheets loaded on the surface of the TiO_2 NRAs and filled the gap between TiO_2 nanorods. Energy band structure of the MoS_2-220/TiO_2 heterojunction composite could promote charge transfer and prevent the recombination of photogenerated electron–hole pairs. The MoS_2-220/TiO_2 heterojunction composite exhibits the best photocatalytic activity, and the photodegradation rate and k value are 86% and 0.0105 min^{-1}, respectively.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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