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# Nitrogen oxide removal by non-thermal plasma for marine diesel engines

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The transportation industry plays an important role in the world economy. Diesel engines are still widely used as the main power generator for trucks, heavy machinery and ships. Removal technology for nitrogen oxides in diesel exhaust are of great concern. In this paper, a gas supply system for simulating the marine diesel engine exhaust is set up. An experimental study on exhaust denitration is carried out by using a dielectric barrier discharge (DBD) reactor to generate non-thermal plasma (NTP). The power efficiency and the denitration efficiency of different gas components by NTP are discussed. The exhaust gas reaction mechanism is analyzed. The application prospects of NTP are explored in the field of diesel exhaust treatment. The experimental results show that the power efficiency and energy density (ED) increase with the input voltage for this system, and the power efficiency is above 80% when the input voltage is higher than 60 V. The removal efficiency of NO is close to 100% by NTP in the NO/N<sub>2</sub> system. For the NO/O<sub>2</sub>/N<sub>2</sub> system, the critical oxygen concentration (COC) increases with NO concentration. The O<sub>2</sub> concentration plays a decisive role in the denitration performance of the NTP. H<sub>2</sub>O contributes to the oxidative removal of NO, and NH<sub>3</sub> improves the removal efficiency at low ED while slightly reducing the denitration performance at high ED. CO<sub>2</sub> has little effect on NTP denitration performance, but as the ED increases, the generated CO gradually increases. When simulating typical diesel engine exhaust conditions, the removal efficiency increases first and then decreases with the increase of ED in the NO/O<sub>2</sub>/CO<sub>2</sub>/H<sub>2</sub>O/N<sub>2</sub> system. After adding NH<sub>3</sub>, the removal efficiency of NO<sub>x</sub> reaches up to 40.6%. It is necessary to add reducing gas, or to combine the NTP technology with other post treatment technologies such as SCR catalysts or wet scrubbing, to further increase the NTP denitration efficiency.

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## 1. Introduction

Today, diesel engines are still widely used in the fields of road transport and non-road machinery and are dominant as the main power and generation power units in the marine sector. Exhaust pollutants such as hydrocarbons (HC), particulate matter (PM), SO<sub>x</sub> and NO<sub>x</sub> are inevitably generated during the operation of the diesel engine. These exhaust pollutants not only cause acid rain, building and soil corrosion, but also endanger human health and lead to diseases such as cancer and respiratory diseases. The SO<sub>x</sub> and NO<sub>x</sub> emission of ocean-faring vessels has been limited by IMO's MARPOL73/78 Annex VI.<sup>1</sup> The removal technologies of diesel exhaust pollutants include pre-treatment, in-machine treatment and post-treatment.<sup>2,3</sup> The post-treatment technologies that can remove a variety of diesel exhaust pollutants at the same time have gained higher attention.

Plasma, the fourth form of matter, was discovered in the mid-18th century. It can generally be classified into high temperature

plasma, thermal equilibrium plasma, and non-thermal equilibrium plasma depending on the temperature of the particles. It is used in a variety of fields such as welding and cutting,<sup>4</sup> surface modification of materials,<sup>5–7</sup> and the removal of contaminants.<sup>8–10</sup> At present, the research on non-thermal equilibrium plasma (NTP) is the most extensive. The methods for generating NTP mainly include electron beam method,<sup>9,11,12</sup> microwave irradiation method,<sup>5–7,13,14</sup> high voltage discharge method (including DC, AC and pulse power),<sup>2,15–17</sup> etc. The method of pulse power combined with the dielectric barrier discharge (DBD) reactor to generate NTP has many advantages, such as higher power efficiency, uniform and silent discharge.<sup>18–21</sup> So it receives more attention. NTP used for exhaust post-treatment began in the 1970s and is one of the hotspots. It is almost no secondary pollution and has a good application prospect.

There are two ways for NTP to remove NO. One is to form N<sub>2</sub> and O<sub>2</sub> by reduction, and the other is to form higher valence oxides by oxidation, such as NO<sub>2</sub>, N<sub>2</sub>O<sub>5</sub>, etc.<sup>12,22</sup> Moreover, diesel exhaust contains a large amount of N<sub>2</sub> (≈76%) and O<sub>2</sub> (≈14%),<sup>23</sup> and because the dissociation energy of O<sub>2</sub> (5.2 eV mol<sup>−1</sup>) is smaller than N<sub>2</sub> (9.8 eV mol<sup>−1</sup>),<sup>24</sup> O<sub>2</sub> will be converted to stronger oxidizing substance such as oxygen radical (·O) and O<sub>3</sub>. These reasons make it more unfavorable to remove NO by

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the reduction route. NO which is difficult to remove in the exhaust gas is converted into other components, whether it is reduced to  $N_2$  or oxidized to high-valent nitrogen oxides, the denitration efficiency can be increased. We mainly focus on the denitration efficiency, and the exact percentage of NO reduction to  $N_2$  will no longer be discussed in the paper.

Zhang and Zhou<sup>25,26</sup> used pulse corona discharge plasma combined with lye absorption for desulfurization and denitration simultaneously. They studied NO oxidation efficiency and removal efficiency by parameters such as gas flow, discharge current, NO concentration and  $SO_2$  concentration. But they did not consider the effect of  $O_2$  concentration on NTP  $NO_x$  removal efficiency. Some studies<sup>24,27</sup> show that with the increasing of  $O_2$  concentration, the removal efficiency of  $NO_x$  gradually decreases, and there is so called critical oxygen concentration (COC) at which the  $NO_x$  removal efficiency is zero, which means the reduction rate of converting  $NO_x$  to  $N_2$  and  $O_2$  is equal to the oxidation rate of converting  $N_2$  to  $NO_x$ . And critical  $O_2$  concentration may change with initial NO concentration. Mok *et al.*<sup>16</sup> used pulsed corona discharge NTP to study the effect of  $O_2$  concentration, humidity and peak voltage on the removal efficiency, but the NO concentration was only 210 ppm (1 ppm =  $1 \mu L L^{-1}$ ) in their experiment. Zhao *et al.*<sup>24</sup> studied the removal efficiency of corona discharge NTP at different  $O_2$  concentrations (0–13.6%), but the NO concentration in this study was only 437 ppm, which was much smaller than that of typical real ship diesel exhaust. What's more, both  $H_2O$  and  $CO_2$  are inevitable components in diesel exhaust. These components may affect the NTP denitration reaction, while Zhao<sup>24</sup> and Aritoshi<sup>27</sup> did not consider the effect of  $H_2O$  on NTP denitration performance. And most researchers have not considered the effect of  $CO_2$  on NTP denitration performance. Also, the denitration mechanism of NTP is not the same under different gas composition, and still needs further exploration. In addition, because the exhaust gas temperature changes with the diesel engine power, and engine exhaust gas generally contains  $SO_2$  (the concentration varies with the sulfur content of the fuel used, such as the concentration of  $SO_2$  in the exhaust gas is about  $600 \mu L L^{-1}$  for marine large low-speed two-stroke diesel engine when using heavy fuel oil of 3.5% sulfur content<sup>23</sup>), both temperature and  $SO_2$  will have a certain impact on the NTP denitration efficiency. Chmielewski *et al.*<sup>12</sup> studied the effects of different temperatures (70 °C and 90 °C) and the  $SO_2$  concentration (0–2000  $\mu L L^{-1}$ ) on the NTP denitration efficiency. Their experimental results show that the plasma denitration efficiency will increase with the temperature, so as the  $SO_2$  concentration. We want to investigate the NTP denitration efficiency in the poor cases and the application prospect of NTP. Therefore, the influence of temperature and  $SO_2$  concentration on the NTP denitration efficiency is not considered in this paper. All experiments are carried out at room temperature of 25 °C and the concentration of  $SO_2$  is zero.

In summary, in order to study the effects of energy density, different initial NO,  $O_2$ ,  $NH_3$ ,  $H_2O$  and  $CO_2$  on the NTP denitration performance, a simulating diesel exhaust supply system was set up in this paper. A coaxial cylindrical DBD reactor was designed and fabricated with quartz glass. Non-thermal plasma was generated

by pulsed power.  $NO_x$  removal mechanism was proposed and analyzed by experimental results. The application prospects of NTP were explored in the field of diesel exhaust treatment.

## 2. Approach

The exhaust composition of a typical large-scale low-speed two-stroke diesel engine is shown in Table 1. NO accounts for more than 90% of  $NO_x$  in the diesel exhaust, the rest mainly being  $NO_2$ , and NO is relatively more difficult to remove. Therefore, this paper mainly focuses on the removal of NO. The concentration of NO in the diesel exhaust will change with the engine load, and generally does not exceed  $1500 \mu L L^{-1}$ . Therefore, three NO concentrations indicating the emissions at different engine loads are studied in the paper, including the low NO concentration ( $500 \mu L L^{-1}$ ), the medium NO concentration ( $1000 \mu L L^{-1}$ ) and the high NO concentration ( $1500 \mu L L^{-1}$ ). When  $NH_3$  is needed, it is added at an ammonia–nitrogen ratio of 1. The highest  $O_2$  concentration in engine exhaust gas is 14%, and in order to study the changes of critical  $O_2$  concentrations with initial NO concentrations, the final selected  $O_2$  concentrations are 1%, 5%, 8%, 10% and 14%, respectively.  $H_2O$  can generate a variety of strong oxidizing free radicals in NTP, including  $\cdot HO_2$ ,  $\cdot OH$  and  $H_2O_2$ . They will have a certain influence on the NTP denitration performance. Therefore, three groups of experiments with  $H_2O$  concentrations of 0%, 3.5% and 5.1% are selected for comparison. Because the experimental results show that  $CO_2$  has little effect on the NTP denitration performance, 0% and 4.5%  $CO_2$  concentration are selected for comparative analysis in this paper. Considering the different components in the diesel exhaust and the performance of the power source, the gas components of each experiment vary from simple to complex. Because  $N_2$  accounts for the largest proportion in the exhaust gas, it is used as carrier gas in our experiments. In addition,  $O_2$  accounts for the second proportion in the exhaust gas and has the fundamental effect on the NTP denitration performance, the experiments of single component gas in NTP are carried both in  $NO/N_2$  system and  $O_2/N_2$  system. Other gases do not seem to be major players in this case and given the length of the article, they are not included in this paper so far. The following research is carried out:

(1) The denitration performance of NTP is strongly related to the energy density in DBD, and the output efficiency of the power source is an important factor to be considered in future industrial applications. Therefore, the energy density and power supply efficiency varying with different input voltages and currents were studied firstly under a typical diesel exhaust condition, *i.e.*,  $1500 \mu L L^{-1} NO + 14\% O_2 + 5.1\% H_2O + 4.5\% CO_2 + 76.2\% N_2$ . Secondly, to understand the removal mechanism of the single-component NO exhaust, and the effect of the NO concentration on the power supply characteristics under a specific ideal condition, we carried out a group of experiments

Table 1 Low-speed two-stroke diesel engine exhaust composition (volume fraction)<sup>23</sup>

Components	$N_2$ (%)	$O_2$ (%)	$CO_2$ (%)	$H_2O$ (%)	$NO_x$ ( $\mu L L^{-1}$ )
Volume fraction	76.2	14	4.5	5.1	1500



in NO/N<sub>2</sub> system, where initial NO concentrations were 500 μL L<sup>-1</sup>, 1000 μL L<sup>-1</sup> and 1500 μL L<sup>-1</sup>, respectively.

(2) Since N<sub>2</sub> and O<sub>2</sub> are the main components of diesel exhaust, O<sub>2</sub> will be converted to a large amount of oxygen active particles with strong oxidative in NTP system, and N<sub>2</sub> will also be converted to nitrogen active particles. These active particles will recombine to a certain concentration of NO<sub>x</sub>, which will partially offset the NO<sub>x</sub> removal effect of NTP. Therefore, the oxidative formation of NO<sub>x</sub> at different energy densities was investigated in O<sub>2</sub>/N<sub>2</sub> system with the O<sub>2</sub> concentration of 1%, 5%, 10% and 14%, respectively.

(3) It is expected that NO is converted to N<sub>2</sub> and O<sub>2</sub> through the reduction route. And the mechanism of NO removal by reduction way needs further study. Therefore, we carried out these experiments in NO/N<sub>2</sub> system with the initial NO concentrations of 500 μL L<sup>-1</sup>, 1000 μL L<sup>-1</sup> and 1500 μL L<sup>-1</sup>, respectively.

(4) As mentioned above, there is a critical oxygen concentration (COC) that makes the NO<sub>x</sub> removal efficiency zero. When the O<sub>2</sub> concentration is higher than the COC, the NTP will not have denitration ability with the ED increasing. It even leads to the increase of the NO<sub>x</sub>. While when the O<sub>2</sub> concentration is lower than the COC, the removal efficiency increases with the ED. What's more, the COC is also related to the initial NO concentration. Therefore, we investigated the range of the COC at initial NO concentrations of 500 μL L<sup>-1</sup>, 1000 μL L<sup>-1</sup> and 1500 μL L<sup>-1</sup>, respectively.

(5) NH<sub>3</sub> is often used as reducing agent in traditional selective catalytic reduction (SCR) denitration, while the NTP denitration performance and mechanism are not well understood when NH<sub>3</sub> exists. To this end, the effect of NH<sub>3</sub> in NO/O<sub>2</sub>/N<sub>2</sub> system was studied at O<sub>2</sub> concentrations of 1%, 5%, 8%, 10% and 14% with initial NO concentrations of 500 μL L<sup>-1</sup>, 1000 μL L<sup>-1</sup> and 1500 μL L<sup>-1</sup>, respectively.

(6) H<sub>2</sub>O is an inevitable component in diesel exhaust. H<sub>2</sub>O will generate strong oxidizing hydroxyl radical (·OH) by the action of NTP, which will further inhibit NO reduction removal. When NH<sub>3</sub> is added to the system, the NO removal reactions will be more complicated. In this paper, the NTP denitration performance was studied in NO/O<sub>2</sub>/H<sub>2</sub>O/N<sub>2</sub> system with initial NO concentrations of 500 μL L<sup>-1</sup>, 1000 μL L<sup>-1</sup> and 1500 μL L<sup>-1</sup>, O<sub>2</sub> concentration of 14% and H<sub>2</sub>O concentrations of 0%, 3.5% and 5.1%, respectively. These results were also compared with that when NH<sub>3</sub> was added.

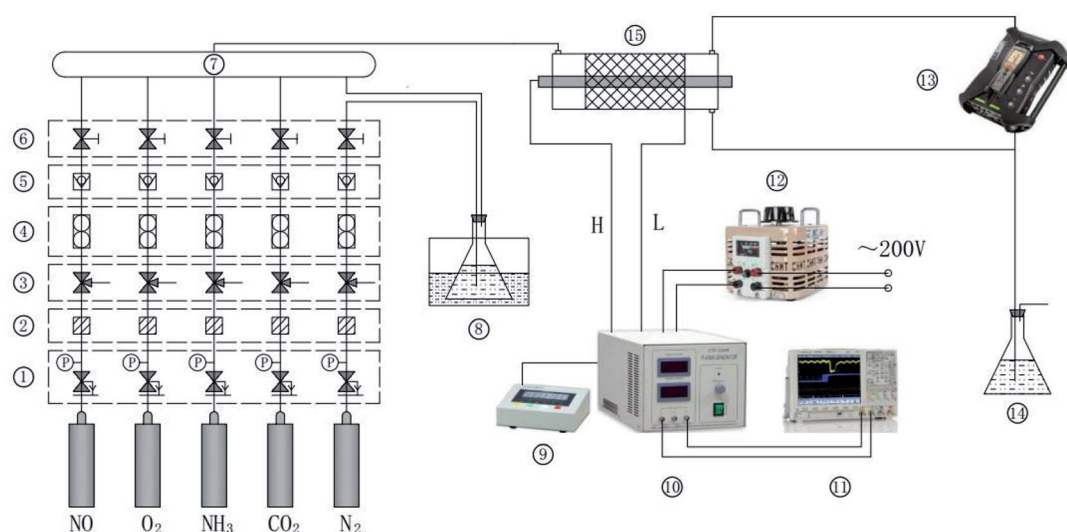
(7) CO<sub>2</sub> is the fourth major component of diesel exhaust. At present, most studies ignore the effect of CO<sub>2</sub> on NTP denitration performance. However, CO<sub>2</sub> may convert to CO in NTP, and CO has been used as the reducing agent for denitration in some reports.<sup>28–31</sup> Therefore, in order to study the possible effects of CO<sub>2</sub> on NTP denitration, we added 4.5% CO<sub>2</sub> in NO/O<sub>2</sub>/H<sub>2</sub>O/N<sub>2</sub> system with initial NO concentrations of 500 μL L<sup>-1</sup>, 1000 μL L<sup>-1</sup> and 1500 μL L<sup>-1</sup>, O<sub>2</sub> concentration of 14% and H<sub>2</sub>O concentration of 5.1%, respectively. Then NH<sub>3</sub> was added in NO/O<sub>2</sub>/H<sub>2</sub>O/CO<sub>2</sub>/N<sub>2</sub> system to investigate the possible changes.

(8) Finally, based on the above experimental results, the NTP denitration mechanism of simulating diesel exhaust is proposed in the paper, and the application prospect of NTP denitration in the field of marine diesel exhaust is discussed.

### 3. Experimental system and data processing

#### 3.1 Experimental system

The experimental system is shown in Fig. 1. It mainly includes gas supply unit, pulse plasma power unit, DBD reactor, flue gas analyzer, and exhaust gas absorption device. The gas supply unit mainly includes gas cylinder, pressure reducing valve, mass flow controller (Beijing Sevenstar CS200A,



1. Pressure reducing valve 2. Filter 3. Three-way valve 4. Mass flow controller 5. Check valve 6. Stop valve 7. Mixing chamber 8. Water bath 9. Pulse modulator 10. Plasma power source 11. Digital oscilloscope 12. Voltage regulator 13. Flue gas analyzer 14. Exhaust gas absorption device 15. DBD reactor

Fig. 1 Experimental system of DBD denitration.

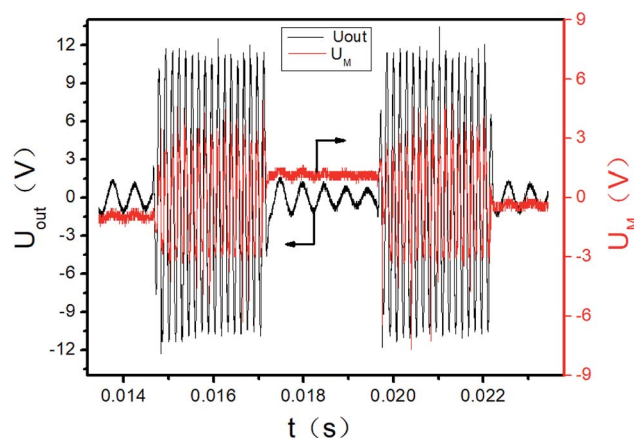




### 3.2 Experimental method

different duty cycles and pulse frequencies by the pulse modulator, as shown in Fig. 4. The attenuation ratio  $K$  of output voltage  $U_{\text{out}}$  is 1000 : 1 in Fig. 3 and 4, which means the true coordinate unit of  $U_{\text{out}}$  is kV (in order to be consistent with the calculation process below, the unit of  $U_{\text{out}}$  is still set as  $V$  in the Fig. 3, 4 and 6). In all these experiments, the pulse frequency and the duty cycle are kept at 200 Hz and 50%, respectively.

The NO<sub>x</sub> removal efficiency is calculated according to the imported and exported NO<sub>x</sub> concentration of DBD recorded by



The initial output of the plasma power source is sine waves with certain frequency, as shown in Fig. 3. The sine waves with certain frequency can be modulated into pulse waves with



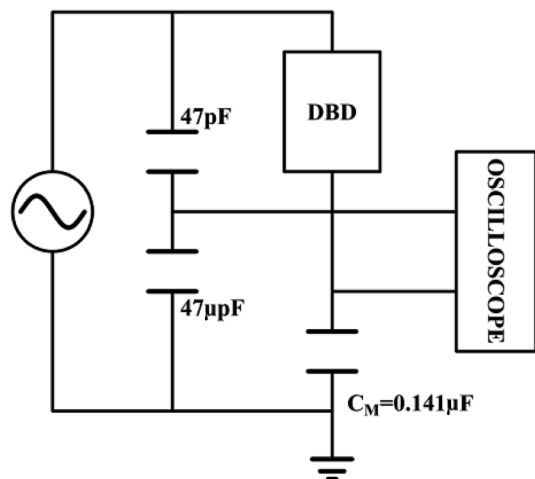


Fig. 5 Equivalent circuit diagram for discharging power by voltage-charge method.

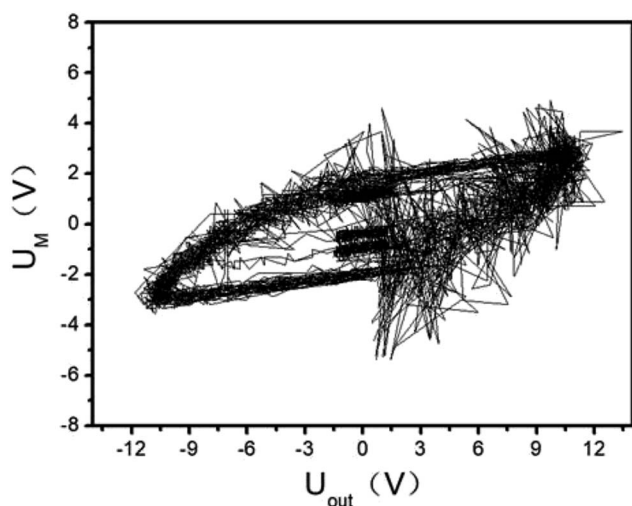


Fig. 6 Diagram of Lissajous integral area.

the flue gas analyzer. The input power of the power source is calculated according to the input voltage and current. The output power of the power source is calculated according to the output waveform of the power source recorded by the digital oscilloscope. The power efficiency and the energy density were calculated further.

The removal efficiency of  $\text{NO}_x$  is calculated by the following formula:

$$\eta_{\text{NO}_x} = \frac{C_{\text{in}} - C_{\text{out}}}{C_{\text{in}}} \times 100\% \quad (1)$$

Here  $\eta_{\text{NO}_x}$  is the removal efficiency of  $\text{NO}_x$ , %;  $C_{\text{in}}$  is the inlet  $\text{NO}_x$  concentration (the sum of  $\text{NO}$  and  $\text{NO}_2$  concentration),  $\mu\text{L L}^{-1}$ ;  $C_{\text{out}}$  is the outlet  $\text{NO}_x$  concentration,  $\mu\text{L L}^{-1}$ .

The input power of the plasma power source is calculated as follows:

$$P_{\text{in}} = U_{\text{in}} I_{\text{in}} \quad (2)$$

Here  $P_{\text{in}}$  is the input power of the plasma power source, W;  $U_{\text{in}}$  is the input voltage of the power source, V;  $I_{\text{in}}$  is the input current of the power source, A.

The output power of the plasma power source is calculated by the voltage-electric charge Lissajous method.<sup>32–34</sup> Its equivalent circuit is shown in Fig. 5. The measurement principle is that a  $0.141 \mu\text{F}$  capacitor  $C_M$  is connected in series with the low-voltage end of the DBD reactor. The voltage  $U_M$  of the capacitor  $C_M$  is measured by the digital oscilloscope. The electric charge  $Q$  of the DBD discharge is equal to the capacitor  $C_M$ . The current flowing through the loop is calculated as:

$$I_{\text{out}} = \frac{dQ}{dt} = \frac{d(C_M U_M)}{dt} = \frac{C_M dU_M}{dt} \quad (3)$$

The discharge power is defined as:

$$P = \frac{1}{T} \int_0^T U_{\text{out}} I_{\text{out}} dT = \frac{C_M}{T} \int_0^T U_{\text{out}} \frac{dU_M}{dt} dT = f C_M \oint U_{\text{out}} dU_M \quad (4)$$

The  $U_M$  and  $U_{\text{out}}$  are measured by the digital oscilloscope. When the input voltage is 100 V and the input current is 0.8 A, the curves of the high-voltage signal  $U_{\text{out}}$  and  $U_M$  are shown in Fig. 4. The waveforms during 0.0150–0.0155 seconds of Fig. 4 are shown in Fig. 3. The Lissajous figure is shown in Fig. 6 with  $U_{\text{out}}$  as the X-axis and  $U_M$  as the Y-axis. The Lissajous figure area  $A$  of a single pulse period is calculated by integration tool of origin. The sensitivity of the oscilloscope in the X and Y directions is  $K_X$  and  $K_Y$ , respectively. The output power by voltage-electric charge Lissajous method can be defined as:

$$P_{\text{out}} = f K C_M K_X K_Y A \quad (5)$$

The power efficiency of NTP can be expressed as:

$$\eta_P = \frac{P_{\text{out}}}{P_{\text{in}}} \times 100\% \quad (6)$$

The output power of the plasma power source is equal to the power loaded in the DBD reactor. The energy density (ED) is usually used to represent the amount of energy that the DBD reactor acting on the flue gas. The calculation formula can be expressed as:

$$\text{ED} = \frac{60 \times P_{\text{out}}}{1000 \times Q_{\text{exhaust}}} \quad (7)$$

Here ED is the energy density of the DBD reactor,  $\text{kJ L}^{-1}$ ;  $P_{\text{out}}$  is the output power of the plasma power source, W;  $Q_{\text{exhaust}}$  is the volume flow rate of the exhaust gas,  $\text{L min}^{-1}$ .

## 4. Results and discussion

### 4.1 The changes of ED and power efficiency

Gas breakdown will occur when the voltage between the two electrodes of the DBD reactor is higher than a certain value. The breakdown voltage is closely related to the discharge space. And it is also related to the gas composition and the ED in DBD



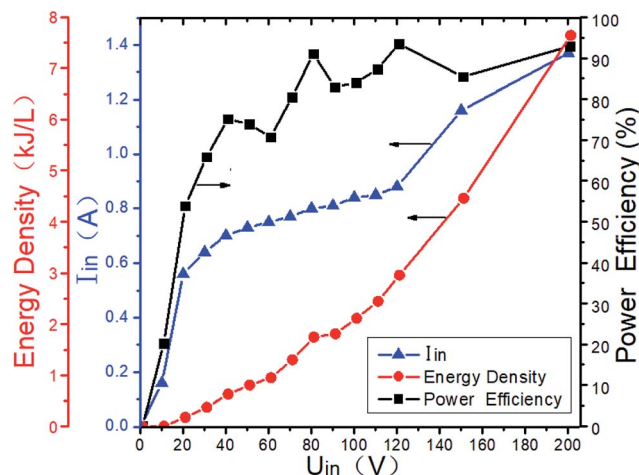


Fig. 7 ED and power efficiency vary with input voltage at  $1500 \mu\text{L L}^{-1}$  NO + 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + 4.5% CO<sub>2</sub> + 76.2% N<sub>2</sub>.

reactor. When the typical marine diesel engine exhaust is simulated, that is,  $1500 \mu\text{L L}^{-1}$  NO + 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + 4.5% CO<sub>2</sub> + 76.2% N<sub>2</sub>, the changes of ED and power efficiency with input voltage are shown in Fig. 7. When the input voltage is low (<20 V), the discharge is extremely unstable, and the input current is also small, so as the output power and power efficiency. As the input voltage increases (20–60 V), the DBD discharge becomes more and more stable, and the input current also increases gradually, so as the ED and power efficiency. When the input voltage is higher than 60 V, the input current and the ED gradually increases with the input voltage, and the power efficiency remains above 80%. As shown in Fig. 8, the power efficiency varies with the input voltage in the NO/N<sub>2</sub> system, where N<sub>2</sub> as the carrier gas and NO concentrations are  $500 \mu\text{L L}^{-1}$ ,  $1000 \mu\text{L L}^{-1}$  and  $1500 \mu\text{L L}^{-1}$ , respectively. When the input voltage is lower than 20 V, the power efficiency does not exceed 70%, when the input voltage is higher than 60 V, the power efficiency remains above 80%. Compared with Fig. 7,

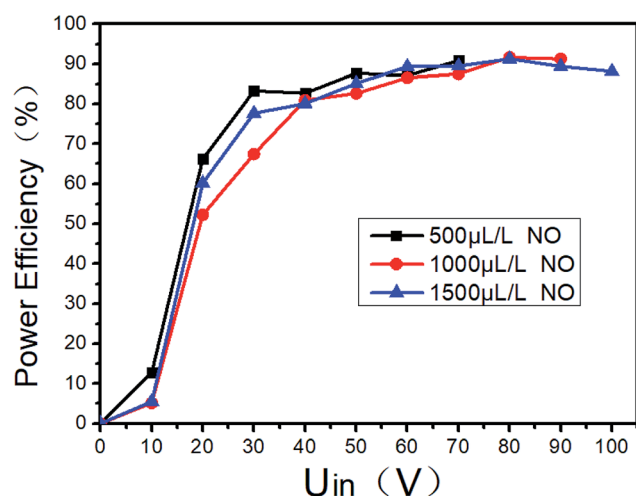


Fig. 8 Power output efficiency varies with input voltage at different NO concentrations.

when the gas components are different at the same input voltage, the power efficiency is basically the same, indicating that the gas composition has little effect on the power efficiency.

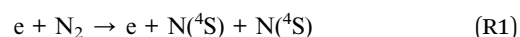
In order to improve the power efficiency, the input voltage should be higher than 60 V. If the system is used in practice, it is recommended that the input voltage be higher than 60 V. However, when the system is scaled up, the optimal efficiency zone and the corresponding ED may be different, and should be determined in actual tests.

## 4.2 Effect of NTP on single component O<sub>2</sub>

As mentioned in the introduction, the main components in diesel exhaust are N<sub>2</sub> and O<sub>2</sub>. They will have a great influence on NTP denitration. Therefore, we first study the effect of NTP in the N<sub>2</sub>/O<sub>2</sub> system. The possible reactions may include R1–R18 (refer to the end of this section) at least. N<sub>2</sub> and O<sub>2</sub> will be converted into a variety of active particles in NTP. Herron<sup>35</sup> and Fernandez's<sup>36</sup> results show that only N(<sup>2</sup>D) and N(<sup>4</sup>S) active particles participate in the NO generation process. Zhao,<sup>24</sup> Herron,<sup>37</sup> Atkinson<sup>38</sup> and others have shown that mainly O(<sup>3</sup>P) and O(<sup>1</sup>D) active particles participate in the NO<sub>x</sub> reaction process. Because of the high concentration of N<sub>2</sub> and O<sub>2</sub>, and the reaction of R5 and R6 are faster, O(<sup>1</sup>D) will be quenched very quickly, and finally only O(<sup>3</sup>P) participates in the NO<sub>x</sub> reaction process. The formations of active particles are shown as R1–R6. N(<sup>2</sup>D) and N(<sup>4</sup>S) are no longer distinguished below, so as O(<sup>3</sup>P) and O(<sup>1</sup>D). They are denoted by ·N and ·O, respectively.

The changes of DBD outlet gas concentration with ED in O<sub>2</sub>/N<sub>2</sub> system are shown in Fig. 9. Fig. 9(a) and (b) show the trend of NO and NO<sub>2</sub> with ED at different O<sub>2</sub> concentrations, respectively. From Fig. 9(a), it can be seen that NO will not be generated at low ED regardless of the O<sub>2</sub> concentration. However as the O<sub>2</sub> concentration and the ED increase, the NO concentration also increases. There will be no NO generated in the ED range of 0–7.6 kJ L<sup>−1</sup> at 1% O<sub>2</sub> concentration. But when the O<sub>2</sub> concentration is 14% and the SIE is 7.3 kJ L<sup>−1</sup>, the NO concentration is as high as  $1018 \mu\text{L L}^{-1}$ . The reactions R7 and R8 may take place in the process. From Fig. 9(b), it can be seen that NO<sub>2</sub> is more easily generated at lower ED in the N<sub>2</sub>/O<sub>2</sub> system. The concentration of NO<sub>2</sub> increases first and then decreases with the increase of O<sub>2</sub> concentration when the ED is higher than 2.1 kJ L<sup>−1</sup>, indicating that the excessive O<sub>2</sub> will inhibit the generation of NO<sub>2</sub>. When O<sub>2</sub> concentration is 5%, it is most beneficial to the formation of NO<sub>2</sub>. The concentration of NO<sub>2</sub> is up to  $380 \mu\text{L L}^{-1}$  when the O<sub>2</sub> concentration is 5% and the ED is 0.9 kJ L<sup>−1</sup>.

Based on the above results, reactions R7–R16 may take place under our experimental condition. For R16, O<sub>3</sub> will be formed by NTP in the presence of O<sub>2</sub>. Because O<sub>3</sub> is more oxidative than O<sub>2</sub>, NO will be oxidized to higher valence oxides. The possible reactions are mainly R16–R18.<sup>39</sup> When the ED is 4 kJ L<sup>−1</sup> and the O<sub>2</sub> concentration is 14%, the NO<sub>x</sub> concentration is  $750 \mu\text{L L}^{-1}$ . According to the above experimental results, the generated NO<sub>x</sub> concentration by NTP gradually increases with the O<sub>2</sub> concentration and ED in O<sub>2</sub>/N<sub>2</sub> system.



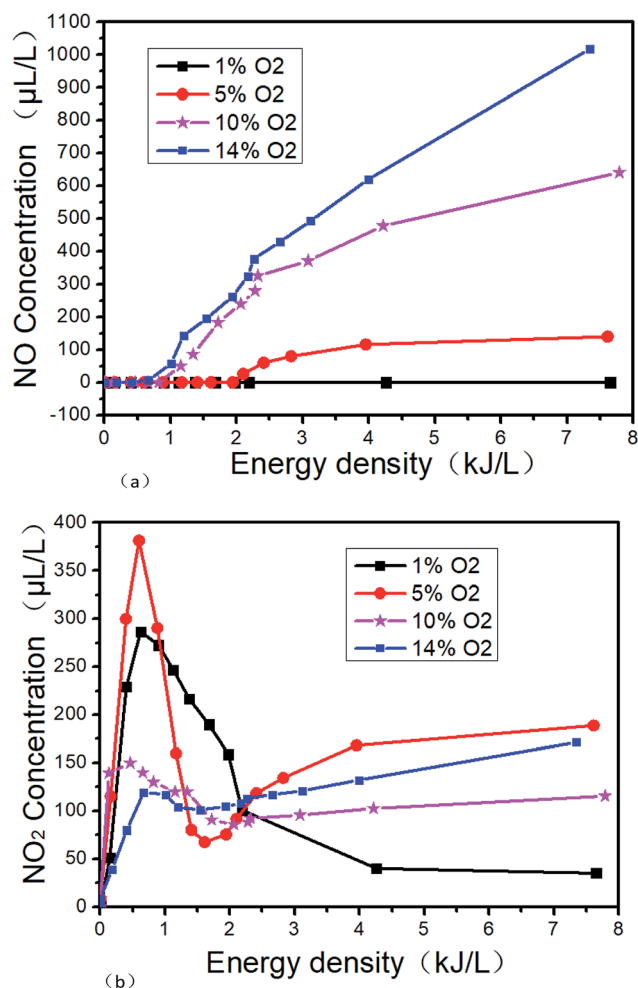
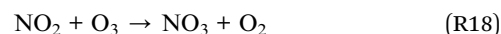
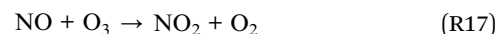
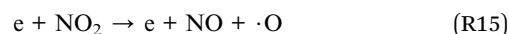
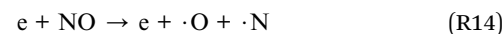
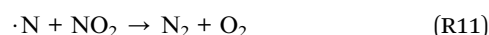
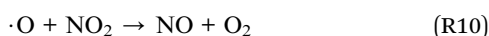
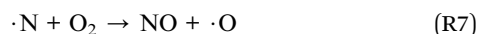
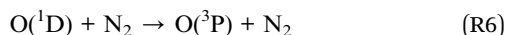
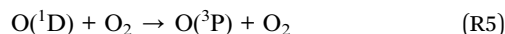
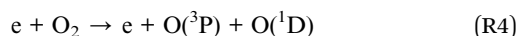
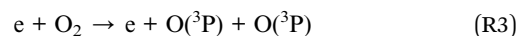


Fig. 9 Effect of NTP at different energy density in O<sub>2</sub>/N<sub>2</sub> system. (a) NO varies with energy density, (b) NO<sub>2</sub> varies with energy density.



### 4.3 Effect of NTP on single component NO

Fig. 10 shows the denitration performance of NTP in NO/N<sub>2</sub> system. Under this condition, NO is converted to N<sub>2</sub> by the reduction way, and the reactions mainly are R12 and R14.<sup>40</sup> It can be seen from Fig. 10(a) that the removal efficiency slightly decreases with the increase of the NO concentration under the same ED. However, the NO is almost completely removed when the ED reaches a certain value. The removal efficiency increases with the ED when the NO concentration is constant. The

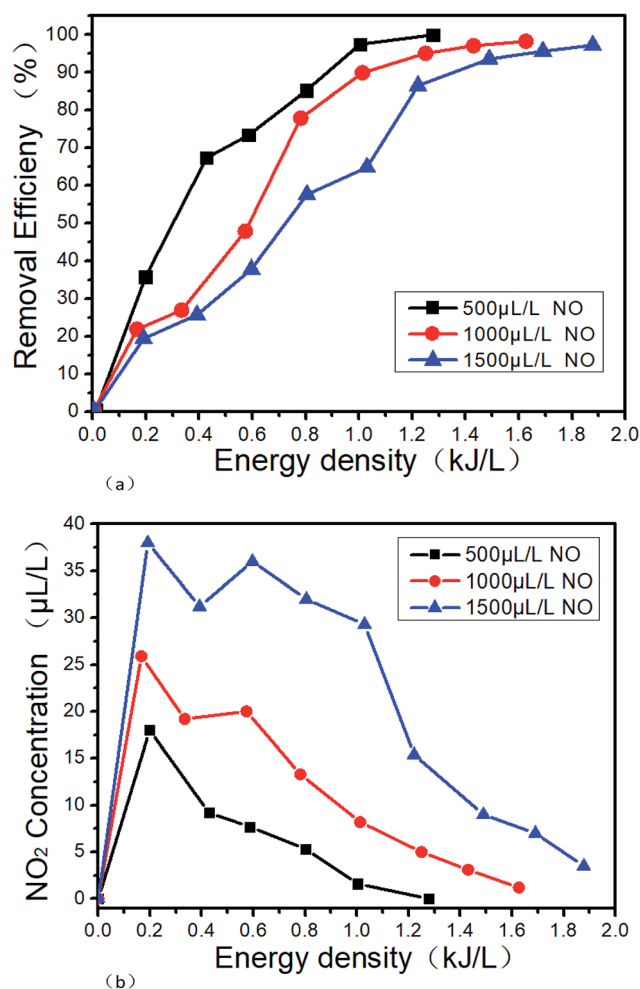


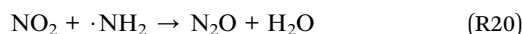
Fig. 10 Denitration performance of NTP at different energy densities in NO/N<sub>2</sub> system, (a) removal efficiency varies with the energy density, (b) NO<sub>2</sub> varies with the energy density.



removal efficiencies are all above 95% when the NO concentrations and the EDs are  $500 \mu\text{L L}^{-1}$   $0.99 \text{ kJ L}^{-1}$ ,  $1000 \mu\text{L L}^{-1}$   $1.26 \text{ kJ L}^{-1}$  and  $1500 \mu\text{L L}^{-1}$   $1.68 \text{ kJ L}^{-1}$ , respectively. The reason for the above experimental result is mainly that the collisions between molecules are more intense with the increase of ED, and R14 is more likely to take place. Also, the reaction R1 and R2 will produce more  $\cdot\text{N}$  active particles, and then the reaction R12 will increase the removal efficiency. Sun<sup>34</sup> got the similar results with us when the NO concentration was  $2000 \mu\text{L L}^{-1}$ . It can be seen from Fig. 10(b) that while NO is removed by the reduction way, a small part of NO is also oxidized to  $\text{NO}_2$ , which means that the reactions of R9 and R14 take place. And the higher the initial NO concentration, the more  $\text{NO}_2$  generates. However, with the increase of the ED, the concentration of  $\text{NO}_2$  increases first and then decreases. When the ED reaches a certain value,  $\text{NO}_2$  is completely removed mainly due to the reactions of R10, R11 and R15. Under the experimental condition, the highest removal efficiency of 99% is obtained in this paper.

Zhao's results<sup>24</sup> show that with the increase of ED, the removal efficiency of  $\text{NO}_x$  will reach 98.5% with 0.5%, and could not be further increased. It is mainly because that  $\text{N}_2\text{O}$  which generated by the reaction of R19 is difficult to convert. When  $\text{NH}_3$  is added, R20 may also take place, but R21 is more difficult to take place. However, the speculation cannot be confirmed in this paper due to that the flue gas analyzer we use could not measure the  $\text{N}_2\text{O}$  concentration.

Based on the above results, NO tends to reduce to  $\text{N}_2$  by NTP in  $\text{NO}/\text{N}_2$  system, and the removal efficiency is higher than 95%.



#### 4.4 Effect of $\text{O}_2$ on NTP denitration performance

Tokunaga<sup>41</sup> and Zhao<sup>24</sup> believe that the removal efficiency of  $\text{NO}_x$  will decrease with the increase of  $\text{O}_2$  concentration, and there is a critical  $\text{O}_2$  concentration to make the  $\text{NO}_x$  removal efficiency zero. Tokunaga's results show that the critical  $\text{O}_2$  concentration is about 3.6% when the initial NO concentration is  $500 \mu\text{L L}^{-1}$ .<sup>41</sup> Zhao's results show that the critical  $\text{O}_2$  concentration is about 2.5% when the initial NO concentration is  $350 \mu\text{L L}^{-1}$ .<sup>24</sup> The NO concentration of the typical marine diesel engine exhaust is generally much higher than  $500 \mu\text{L L}^{-1}$ . When the initial NO concentration is higher, the changes of the critical  $\text{O}_2$  concentration are not very clear. Based on this, the  $\text{NO}_x$  removal performance of NTP is studied when the initial NO concentrations,  $\text{O}_2$  concentrations and the EDs change. Fig. 11 shows the NTP denitration performance under different conditions of NO and  $\text{O}_2$ . Since the remove rate is close to 100% at lower energy density when  $\text{O}_2$  concentration is zero, the experiments at higher energy densities are no longer performed.

It can be seen that the higher the NO concentration, the higher the removal efficiency at the same  $\text{O}_2$  concentration. The reason for this phenomenon is that the probability of NO collision with high-energy particles is higher in the reaction system with the increase of NO concentration. The greater the likelihood, the more NO will be removed by converting to  $\text{N}_2$ . While it may be helpful to use the isotope labeling method to detect the migration of nitrogen atoms in NO to confirm the speculation, it is not included in this study due to experimental conditions.

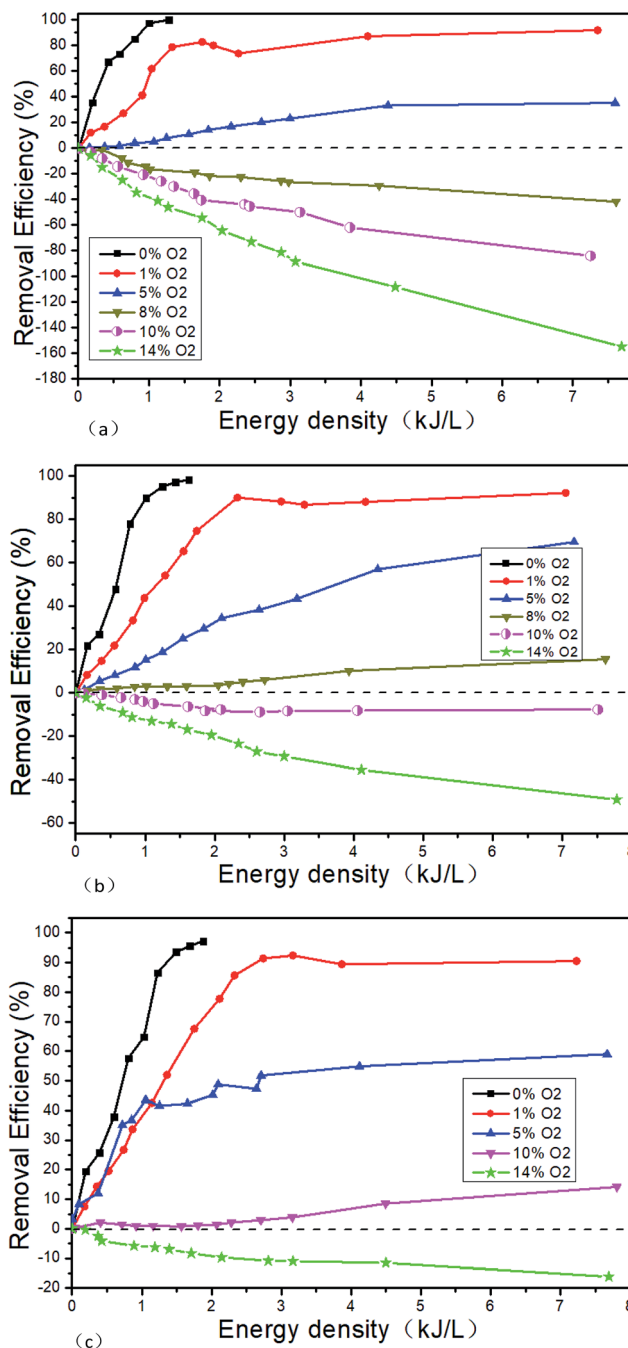


Fig. 11 Denitration performance of NTP in  $\text{NO}/\text{O}_2/\text{N}_2$  system, (a)  $500 \mu\text{L L}^{-1}$  NO +  $\text{O}_2$ , (b)  $1000 \mu\text{L L}^{-1}$  NO +  $\text{O}_2$ , (c)  $1500 \mu\text{L L}^{-1}$  NO +  $\text{O}_2$ .





Secondly, it can be seen that the removal efficiency decreases with the increase of  $O_2$  concentration when the initial  $NO$  concentration is constant. The removal efficiency is more than 90% at low  $O_2$  concentration (less than 1%) when the ED is bigger enough. When the  $O_2$  concentration is more than 14%, the removal efficiency is negative. And with the increase of the ED, the removal efficiency is further reduced. The critical  $O_2$  concentration gradually increases with the initial  $NO$  concentration. The critical  $O_2$  concentration range is 5–8% when the initial  $NO$  concentration  $500 \mu\text{L L}^{-1}$ . Tokunaga<sup>41</sup> believes that the critical  $O_2$  concentration is about 3.6% when the initial  $O_2$  concentration is  $500 \mu\text{L L}^{-1}$ . But our experimental results are not agreed with his. The possible reason is that the NTP generation methods are different. Tokunaga uses electron beam method and we use the DBD method. It is reported that the average free energy of electrons generated by the electron beam method is much higher than that of the DBD method.<sup>42</sup> Therefore, more  $N_2$  which is difficult to excite is converted to  $\cdot N$  by the electron beam method. So the reactions of R1–R9 are easier by electron beam method even at lower  $O_2$  concentration. And then the chemical reaction direction is toward to  $NO_x$  generation.

The critical  $O_2$  concentration range is 5–8%, 8–10% and 10–14% corresponding to the initial  $NO$  concentration  $500 \mu\text{L L}^{-1}$ ,  $1000 \mu\text{L L}^{-1}$  and  $1500 \mu\text{L L}^{-1}$ , respectively. The COC gradually increases with the initial  $NO$  concentration. The removal efficiency increases with the ED when the  $O_2$  concentration is lower than the COC, and decreases with the ED when the  $O_2$  concentration is higher than the COC.

#### 4.5 Effect of $NH_3$ on NTP denitration performance

$NH_3$  is often used as reducing agent for the removal of  $NO_x$  in the conventional SCR method. At present, there are few studies on the  $NO_x$  removal by  $NH_3$  in NTP system, and the  $NO_x$  removal mechanism by NTP +  $NH_3$  is even more unclear. It is generally believed that the ammonia radical ( $\cdot NH_2$ ) plays a major role in the  $NO_x$  removal in NTP system when  $NH_3$  is added. The reactions of the  $\cdot NH_2$  formation are R22–R25 (refer to the end of this section).

Fig. 12 shows the NTP denitration performance of  $NH_3$  under different conditions of  $NO$  and  $O_2$ . Taking the initial  $NO$  concentration of  $1000 \mu\text{L L}^{-1}$  as an example, it can be seen from the comparison with Fig. 11 that when the  $O_2$  concentration is 1% and the energy density is greater than  $2.25 \text{ kJ L}^{-1}$ , the  $NO_x$  removal efficiency changes from 90% (Fig. 11(b) shown) to 80% (as shown in Fig. 12(b)) when  $NH_3$  is added. When the  $O_2$  concentration is greater than 10%, the removal efficiency remains positive at low ED after adding  $NH_3$ . But with the increase of the ED, the removal efficiency gradually decreases and eventually becomes negative. When the  $O_2$  concentration is 14% and the ED is  $7.8 \text{ kJ L}^{-1}$ , the lowest removal efficiency changes from  $-49.2\%$  (as shown in Fig. 11(b)) to  $-62.4\%$  (as shown in Fig. 12(b)) after  $NH_3$  is added. When the  $O_2$  concentration is less than 5%,  $NO$  is removed by the reduction way and R13 is supposed to be the main reaction. The generated  $NO_2$  is removed by the reaction R11. After the addition of  $NH_3$ , the

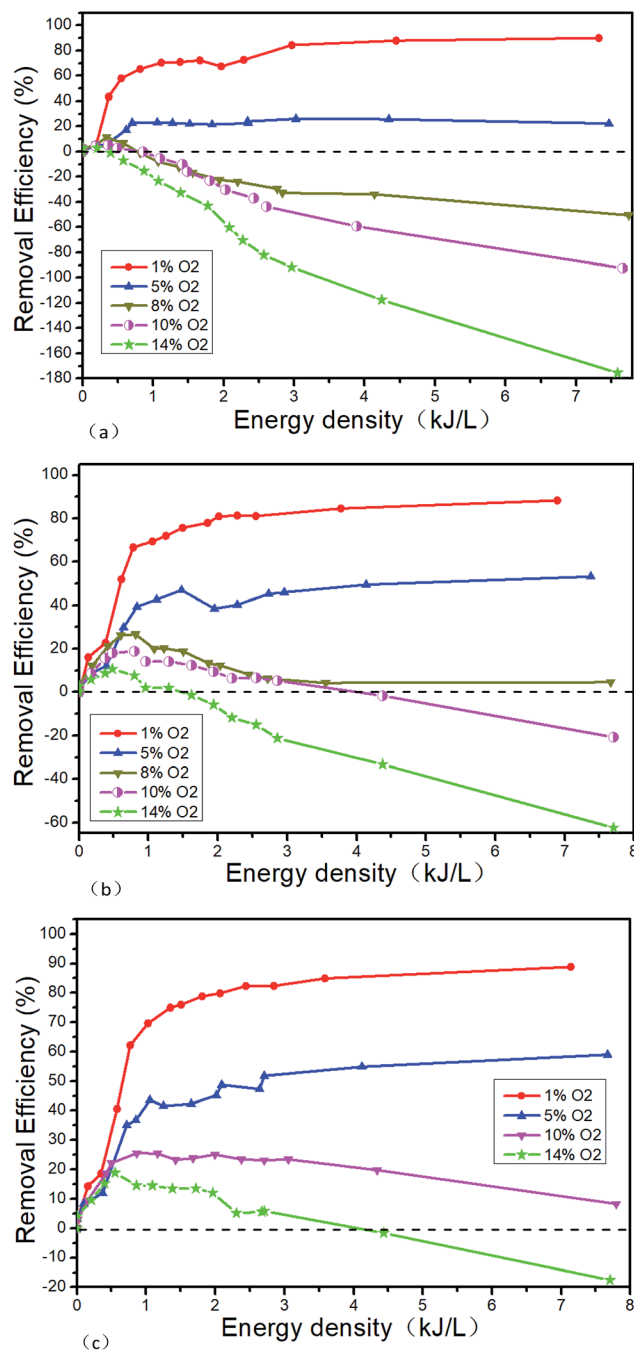


Fig. 12 Denitration performance of NTP in  $NO/NH_3/O_2/N_2$  system, (a)  $500 \mu\text{L L}^{-1} NO + 500 \mu\text{L L}^{-1} NH_3 + O_2$ , (b)  $1000 \mu\text{L L}^{-1} NO + 1000 \mu\text{L L}^{-1} NH_3 + O_2$ , (c)  $1500 \mu\text{L L}^{-1} NO + 1500 \mu\text{L L}^{-1} NH_3 + O_2$ .

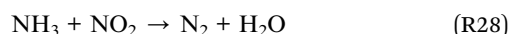
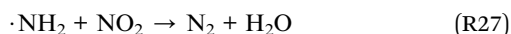
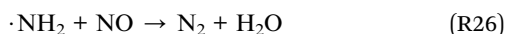
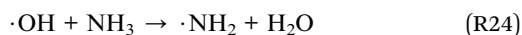
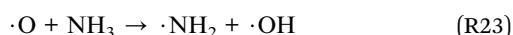
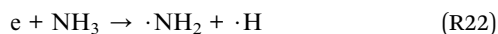
reactions of R22–R28 mainly take place. At low  $O_2$  concentration, the addition of  $NH_3$  leads the removal efficiency decreasing. The experimental result may be caused by the fact that in this experimental condition the removal of  $NO_x$  is mainly caused by the collision, which breaks the molecular bonds between high-energy particles and molecules generated by NTP. And then the molecular bonds recombine to generate bigger bond energy molecules, such as  $N_2$ . However, after the addition of  $NH_3$ , the probability of collision between the high energy



particles and the NO molecules is lowered, thereby the removal efficiency decreases. When the O<sub>2</sub> concentration is greater than 10%, the NTP removal efficiency becomes negative. The reason for this phenomenon may be that the O<sub>2</sub> concentration is much higher than NO, and the oxygen-active particles generated by collisions between high-energy particles and O<sub>2</sub> molecules are much more, making the reactions R7 and R8 more likely to react. Although the NO<sub>x</sub> can be decreased after the addition of NH<sub>3</sub> at the low ED, the concentration of oxygen-activated particles increases with the ED. And the NO generation rate by oxidation reaction is higher than the removal efficiency by NH<sub>3</sub> reduction reaction, which eventually leads to the negative removal efficiency.

Secondly, the experimental results show that the NO<sub>2</sub> concentration will be slightly reduced at different O<sub>2</sub> concentrations when NH<sub>3</sub> is added. Mizuno<sup>43</sup> believes that NO does not react with NH<sub>3</sub> at room temperature, and the conversion of NO to NO<sub>2</sub> is determined by the concentration of  $\cdot\text{O}$  instead of the NH<sub>3</sub>. In the range of room temperature to 150 °C, NH<sub>3</sub> only affects the removal efficiency of NO<sub>2</sub> instead of NO.

Based on the findings above, NH<sub>3</sub> has little effect on the critical O<sub>2</sub> concentration. However, when other initial conditions are constant, the removal efficiency will be greatly improved at the low energy density and further reduced at the high energy density after adding NH<sub>3</sub>. Therefore, it should be avoided the energy density is too much when NH<sub>3</sub> is added, which will result in the oxidation of NH<sub>3</sub> and the decrease of the removal efficiency.



#### 4.6 Effect of H<sub>2</sub>O on NTP denitration performance

H<sub>2</sub>O is an inevitable combustion product in diesel exhaust. H<sub>2</sub>O can generate a variety of strong oxidizing free radicals in NTP, including  $\cdot\text{HO}_2$ ,  $\cdot\text{OH}$  and H<sub>2</sub>O<sub>2</sub>. These strong oxidizing active particles will contribute to the oxidative removal of NO. The possible reactions are R29–R33 (M is a third inert body in the reaction system).<sup>44</sup>

Fig. 13 shows the changes of the NO<sub>x</sub> concentrations at the outlet when the inlet gases are 1500  $\mu\text{L L}^{-1}$  NO + 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + N<sub>2</sub>, i.e., H<sub>2</sub>O is added. Fig. 14 shows the denitration performance of NH<sub>3</sub> at different H<sub>2</sub>O concentrations in the system of 1500  $\mu\text{L L}^{-1}$  NO + 14% O<sub>2</sub> + H<sub>2</sub>O + N<sub>2</sub>, i.e., both H<sub>2</sub>O and NH<sub>3</sub> are added. It can be seen from Fig. 13 that H<sub>2</sub>O nearly

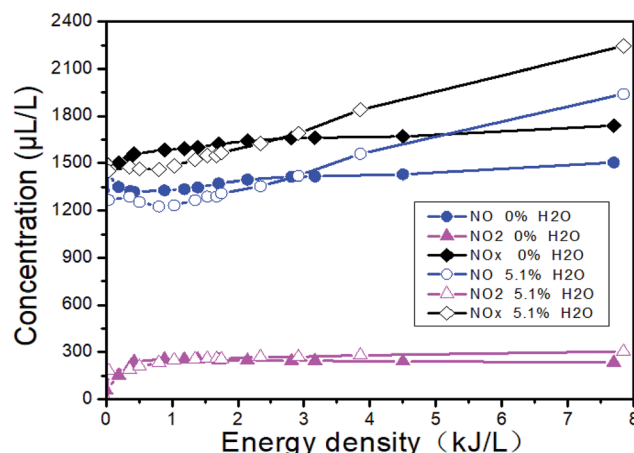


Fig. 13 Effect of H<sub>2</sub>O on the NO<sub>x</sub> removal.

has no effect on the concentration of NO<sub>2</sub>. H<sub>2</sub>O can reduce the NO concentration a little bit at the low ED and increase the NO concentration at the high ED, which means H<sub>2</sub>O can increase the removal efficiency at the low ED and lower the removal efficiency at the high ED. However, there is no obvious regular pattern between H<sub>2</sub>O concentration and NTP removal efficiency with the increase of the ED (Fig. 14). The reactions mainly consist of R29–R38. When NH<sub>3</sub> is added, the removal efficiency increases first and then decreases with the increase of the ED. When both NH<sub>3</sub> and H<sub>2</sub>O are present, the removal efficiency of NO<sub>x</sub> can be greatly improved.

Fig. 15 shows the changes of NTP removal efficiency at different initial NO concentrations in the system of 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + NO + NH<sub>3</sub> + N<sub>2</sub>. Whether NH<sub>3</sub> is added or not, the removal efficiency of NO<sub>x</sub> gradually increases with the initial NO concentration. At constant initial NO concentration, the removal efficiency increases first and then decreases with the increase of the ED. This conclusion has been shown in Section 3.5. The addition of NH<sub>3</sub> will contribute to the increase of removal efficiency in the whole ED range in the system of 1500  $\mu\text{L L}^{-1}$  NO + 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + NH<sub>3</sub> + N<sub>2</sub>,

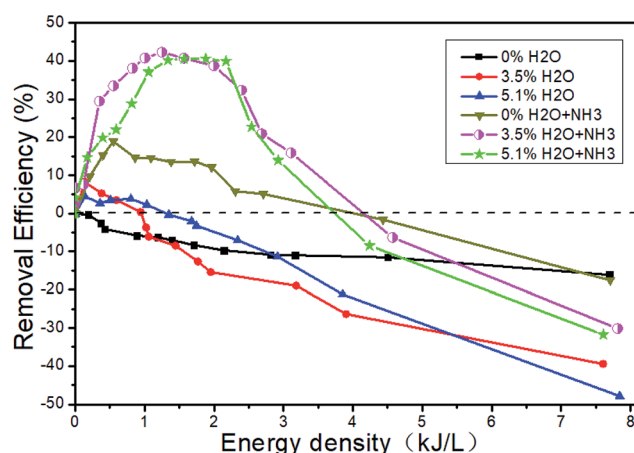


Fig. 14 Effect of H<sub>2</sub>O and NH<sub>3</sub> on the removal efficiency.



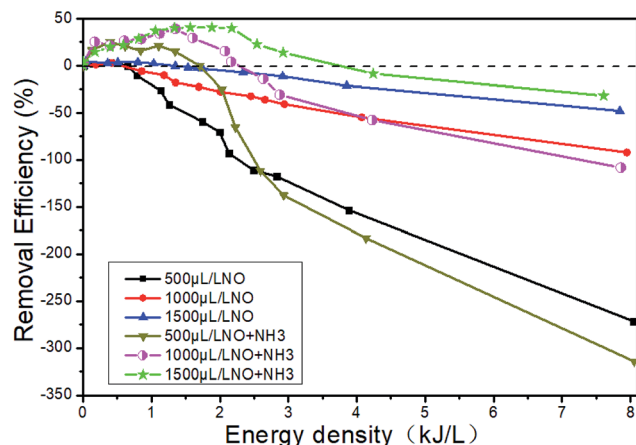
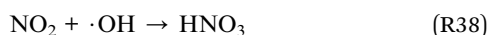
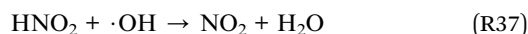
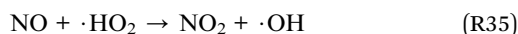
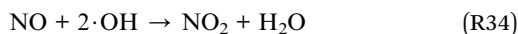
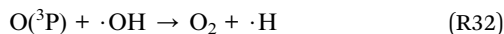
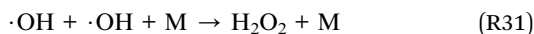
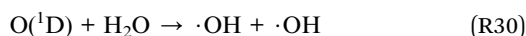
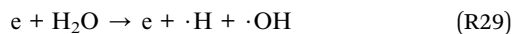


Fig. 15 Effect of H<sub>2</sub>O on NTP denitration performance at different initial NO concentrations.

and the removal efficiency is up to 40.6% when the ED is 1.57 kJ L<sup>-1</sup>.

In general, H<sub>2</sub>O can increase the NTP removal efficiency at the low ED and reduce the removal efficiency at the high ED. When H<sub>2</sub>O and NH<sub>3</sub> coexist, the NTP removal efficiency is greatly improved. Therefore, the system consists of the following reactions:



#### 4.7 Effect of CO<sub>2</sub> on NTP denitration performance

CO<sub>2</sub> is also an inevitable combustion product in diesel exhaust, and most researchers have not explored the effects of CO<sub>2</sub> on the NTP denitration. Therefore, in this section we investigate the impact of CO<sub>2</sub> in the NTP system for different initial NO concentrations and the presence of NH<sub>3</sub>.

The effects of 4.5% CO<sub>2</sub> on the DBD outlet gas concentration in the system of 1500 μL L<sup>-1</sup> NO + 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + N<sub>2</sub> are shown in Fig. 16(a). The effect of 4.5% CO<sub>2</sub> on the NTP removal efficiency at different initial NO concentration in 14% O<sub>2</sub> + 5.1%

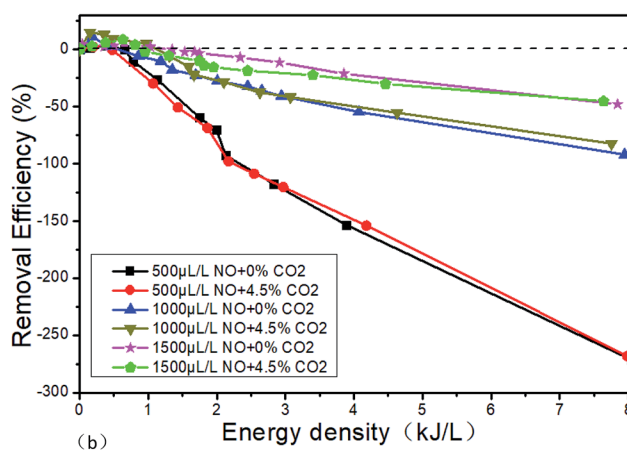
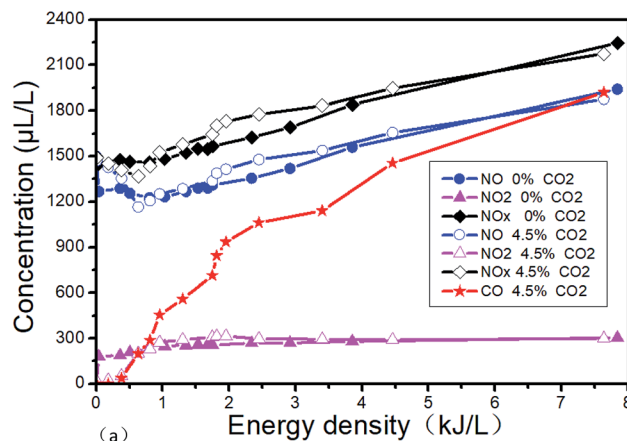


Fig. 16 Effects of CO<sub>2</sub> on the NTP denitration performance, (a) effects of CO<sub>2</sub> on DBD outlet gas concentration in the system of 1500 μL L<sup>-1</sup> NO + 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + 76.2% N<sub>2</sub>. (b) Effects of CO<sub>2</sub> on the NTP removal efficiency at different initial NO concentration in 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + 76.2% N<sub>2</sub> + NO.

H<sub>2</sub>O + N<sub>2</sub> + NO is shown in Fig. 16(b). It can be seen from Fig. 16(a) that CO<sub>2</sub> almost has no effect on the concentration of NO and NO<sub>2</sub> at the DBD outlet. But when 4.5% CO<sub>2</sub> is added, a certain concentration of CO will be produced, and the outlet CO concentration will gradually increase with the ED. When the ED is 7.64 kJ L<sup>-1</sup>, the CO concentration is up to 1920 μL L<sup>-1</sup>. The possible reactions are R39–R43. And R40 is the main reaction because CO<sub>2</sub> concentration is much higher than the CO. It can be seen from Fig. 16(b) that CO<sub>2</sub> basically does not have any influence on the NTP removal efficiency at different initial NO concentrations.

The changes of NTP removal efficiency at different initial NO concentrations are shown in Fig. 17 when NH<sub>3</sub> is added in the system of 14% O<sub>2</sub> + 5.1% H<sub>2</sub>O + 4.5% CO<sub>2</sub> + 76.2% N<sub>2</sub> + NO. The CO concentration in the outlet increases with the ED instead of the initial NO concentration. The NTP removal efficiency increases significantly when NH<sub>3</sub> is added. However, CO<sub>2</sub> has no effect on the NTP removal efficiency in the presence of NH<sub>3</sub> when compared with Fig. 15. In addition, we opened the DBD reactor after the experiment of adding CO<sub>2</sub>, and a layer of black carbon was observed on the surface of high-pressure copper



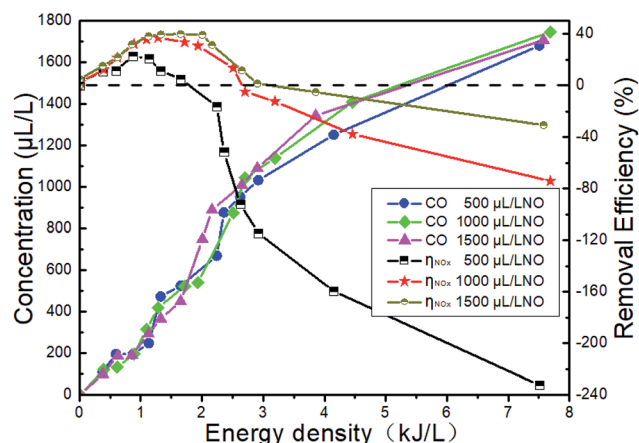
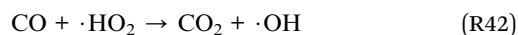
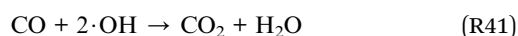
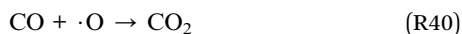


Fig. 17 Effect of CO<sub>2</sub> on NTP removal efficiency with the presence of NH<sub>3</sub>.

rod. This may be due to the further decomposition of CO, which means reaction R44 occurred.

In summary, CO<sub>2</sub> nearly has no effect on the removal efficiency of NO<sub>x</sub> under different experimental conditions. But CO will be detected in the outlet and black carbon will be generated on the surface of high-voltage electrode after adding CO<sub>2</sub>, and the CO concentration will gradually increase with the ED.



## 5. Reaction mechanism of NO<sub>x</sub> and application prospect of NTP

### 5.1 Reaction mechanism of NO<sub>x</sub> in NTP

When the inlet gases are N<sub>2</sub>, O<sub>2</sub>, CO<sub>2</sub>, H<sub>2</sub>O, NO and NH<sub>3</sub>, NTP denitration can be roughly divided into three stages.<sup>16,22,40,45</sup>

(1) The first stage is the discharge. At this stage gas molecules are mainly bombarded by high-energy electrons. It breaks the molecular covalent bonds, changes the gas molecules into free radicals and excites some decomposed atoms to the unstable excited state. The following reactions mainly take place at this stage: R1–R4, R14 and R15, R22, R29, R39 and R44.

(2) The second stage is the post-discharge. At this stage the excited-state atoms generated in the first stage collide with the gas molecules to generate secondary radicals. Then the radicals collide with other particles leading to quench or new radicals'

generation. The reactions are mainly R5 and R6, R16, R23–R25, R30–R33.

(3) The third stage is that the free radicals react with NO<sub>x</sub>. The reactions mainly include R7–R13, R17–R21, R26–R28, R34–R38. The reactions of free radicals with CO mainly include R40–R43.

In general, under our experimental conditions, the reaction mechanism of diesel engine exhaust in the NTP system is shown in Fig. 18. It mainly consists of two parts. The first is that gas molecules generate various free radicals under the bombardment of high-energy electrons. The second is the radicals react with other particles. The reaction process of NO<sub>x</sub> with free radicals is described in the right of Fig. 18. Other researchers<sup>16,22,40,45</sup> get the similar results with us. While our experimental results show that CO<sub>2</sub> has little effect on the removal efficiency of NO<sub>x</sub>. CO<sub>2</sub> is converted to CO and even further converted to black carbon. Therefore, the reaction process of CO<sub>2</sub> in NTP is separately listed in the left of Fig. 18.

### 5.2 Analysis of NTP application prospect

As NO can be removed by oxidation or reduction, when the O<sub>2</sub> concentration is low, NO is mainly removed by the reduction route. When the O<sub>2</sub> concentration is high, many strong oxidative constituents are generated such as  $\cdot\text{OH}$ ,  $\cdot\text{O}$  and O<sub>3</sub>. And then the NO is oxidized to NO<sub>2</sub>. But N<sub>2</sub> will be converted to NO<sub>x</sub> at high O<sub>2</sub> concentration and high energy density, causing the NTP removal efficiency to become negative. Although it is beneficial to the NTP removal efficiency improvement when the initial NO concentration is higher, it is difficult to further improve removal efficiency only by NTP technology because of

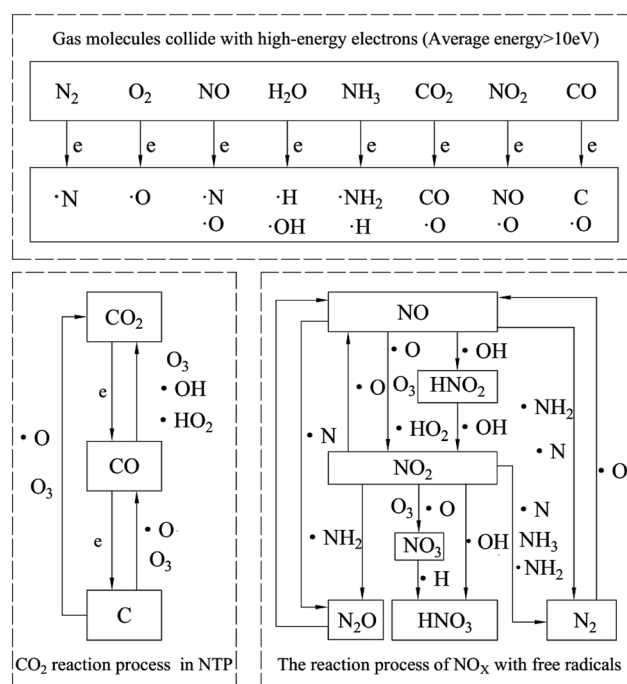


Fig. 18 Reaction mechanism of diesel engine exhaust gas in NTP system.





the high  $O_2$  concentration in the marine diesel exhaust. In this paper, the highest removal efficiency is only 8.6% at  $0.8 \text{ kJ L}^{-1}$  energy density when inlet gases are  $1500 \mu\text{L L}^{-1} \text{ NO} + 14\% \text{ O}_2 + 5.1\% \text{ H}_2\text{O} + 4.5\% \text{ CO}_2 + 76.2\% \text{ N}_2$ . Chmielewski<sup>12</sup> simulated the diesel engine exhaust conditions when burning heavy oil, and the removal efficiency obtained was also very low. Therefore, in order to achieve high-efficiency  $\text{NO}_x$  removal at high  $O_2$  concentration, it is necessary to combine NTP denitration technology with other methods. The possible methods include the following:

(1) Adding reducing gas, such as  $\text{H}_2$ ,<sup>46</sup>  $\text{NH}_3$  (ref. 11, 21 and 47–49) and  $\text{HC}$ <sup>50–52</sup> (including  $\text{CH}_4$ ,  $\text{C}_2\text{H}_2$ ,  $\text{C}_2\text{H}_4$ ,  $\text{C}_3\text{H}_6$ , etc.). In this paper, the highest removal efficiency of 40.6% is obtained at the energy density of  $1.65 \text{ kJ L}^{-1}$  when  $\text{NH}_3$  is added.

(2) Combining NTP with the denitration catalysts.<sup>15,21,49,51,53–56</sup> The catalysts commonly used include molecular sieves, activated carbon and metal oxides. For example, we can combine NTP with traditional vanadium-based SCR catalyst. Because NTP can convert  $\text{NO}$  to  $\text{NO}_2$ , which means increasing  $\text{NO}_2/\text{NO}_x$  ratio, it is beneficial to increase the reaction rate of SCR.<sup>57</sup> The removal efficiency is generally above 90% in this way, but it is also necessary to add reducing gas.

(3) Combining NTP with wet scrubbing technology.<sup>58–60</sup> This technology has matured on ships today. NTP can oxidize  $\text{NO}$  to  $\text{NO}_2$  which is more soluble in water. Chmielewski<sup>11</sup> obtained the removal efficiency of 49% by combining NTP with wet scrubbing. And Yang<sup>61</sup> got the removal efficiency of more than 60% by combining electrolytic seawater with wet scrubbing. However, in general, the removal efficiency is still much lower than the traditional methods such as SCR technology.

A major problem for NTP industrial applications is that energy consumption of NTP is very large. Improving the energy utilization rate can be solved by optimizing the structure of the DBD reactor and matching the reactor with the power source.

## 6. Conclusions

A non-thermal plasma denitration system based on simulated diesel engine exhaust was set up in this paper. The NTP was generated by dielectric barrier discharge reactor. The  $\text{NO}$  removal performance by NTP under different  $O_2$ ,  $\text{H}_2\text{O}$ ,  $\text{CO}_2$ ,  $\text{NO}$ ,  $\text{NH}_3$  concentrations and energy densities conditions were studied. The reaction mechanism of  $\text{NO}_x$  in NTP system was proposed. The application prospect of NTP technology was analyzed. In the end we get the following conclusions:

(1) For the experimental system, the power source efficiency gradually increases with the input voltage. When the input voltage is greater than 60 V, the power supply efficiency is basically maintained above 80%.

(2)  $\text{NO}$  concentration increases gradually with the  $O_2$  concentration and the ED in the  $\text{N}_2/\text{O}_2$  system. When the  $O_2$  concentration is 14% and the ED is  $7.3 \text{ kJ L}^{-1}$ , the concentration of  $\text{NO}$  is up to  $1018 \mu\text{L L}^{-1}$ . The amount of  $\text{NO}_2$  increases and then stabilizes with the ED. When the  $O_2$  concentration is 5% and the ED is  $0.9 \text{ kJ L}^{-1}$ , the concentration of  $\text{NO}_2$  is up to 380

$\mu\text{L L}^{-1}$ . Therefore, a certain concentration of  $\text{NO}_x$  will be generated when introducing NTP in the air.

(3) NTP has a high removal efficiency for  $\text{NO}/\text{N}_2$  system. When low concentration  $O_2$  exists, NTP removal efficiency is above 90%. And  $\text{NO}$  is mainly removed by the reduction route. NTP will have no denitration performance at high concentration of  $O_2$ . However,  $\text{NH}_3$  can inhibit the formation of  $\text{NO}_x$  and improve the removal efficiency. At low  $O_2$  concentration, the removal efficiency of  $\text{NO}_x$  gradually increases with the ED; at high  $O_2$  concentration, the removal efficiency becomes negative. And the higher the ED, the more  $\text{NO}_x$  will generate. Therefore, the  $O_2$  concentration plays a decisive role in NTP denitration performance. And the critical  $O_2$  concentration increases with the initial  $\text{NO}$  concentration.

(4) Under typical diesel engine exhaust condition,  $\text{H}_2\text{O}$  has little effect on  $\text{NO}_2$  when  $\text{NH}_3$  is not added, but it can increase  $\text{NO}$  removal efficiency at the low ED and the excessive ED causes the increasement of  $\text{NO}$  at the outlet. The highest removal efficiency of 4.5% can be obtained in the system of  $1500 \mu\text{L L}^{-1} \text{ NO} + 14\% \text{ O}_2 + 5.1\% \text{ H}_2\text{O} + \text{N}_2$ . The highest removal efficiency of 18.9% can be obtained in the system of  $1500 \mu\text{L L}^{-1} \text{ NO} + 1500 \mu\text{L L}^{-1} \text{ NH}_3 + 14\% \text{ O}_2 + \text{N}_2$ . However, when  $\text{H}_2\text{O}$  and  $\text{NH}_3$  are added, the removal efficiency is up to 40.6% in the system of  $1500 \mu\text{L L}^{-1} \text{ NO} + 1500 \mu\text{L L}^{-1} \text{ NH}_3 + 5.1\% \text{ H}_2\text{O} + 14\% \text{ O}_2 + \text{N}_2$  because of the synergistic effect.

(5)  $\text{CO}_2$  nearly has no effect on the removal efficiency of  $\text{NO}_x$ , but the concentration of  $\text{CO}$  will increase gradually with the ED. When the ED is  $7.64 \text{ kJ L}^{-1}$ , the  $\text{CO}$  concentration will up to  $1920 \mu\text{L L}^{-1}$ . The reaction process of  $\text{CO}_2$  is added in the reaction mechanism of diesel engine exhaust in NTP system.

(6) Because of the high concentration of  $O_2$  in the marine diesel engine exhaust, in order to further improve the removal efficiency of  $\text{NO}_x$ , it is necessary to add reducing gas to the NTP reactor, or combine the NTP technology with SCR catalysts or other technologies such as wet scrubbing.

## Conflicts of interest

There are no conflicts to declare.

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