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A green approach to prepare hierarchical porous carbon nanofibers from coal for high-performance supercapacitors†

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A green method is designed to obtain hierarchical porous carbon nanofibers from coal. In the work, deionized water, coal, polyvinyl alcohol and Pluronic F127 are used as the aqueous solution, carbon source, spinning assistant and soft template for spinning, respectively. As electrode materials for supercapacitors, the obtained hierarchical porous carbon nanofibers exhibit a high specific capacitance of 265.2 F g⁻¹ at 1.0 A g⁻¹ in 6 M KOH, a good rate performance with a capacitance of 220.3 F g⁻¹ at 20.0 A g⁻¹ with the retention of 83.1% and a superior cycle stability without capacitance loss after 20 000 charge/discharge cycles at 10.0 A g⁻¹. Compared with the carbon nanofibers constructed without Pluronic F127, the enhanced electrochemical performance of the sample benefits from a larger contact surface area and the mesoporous structure formed by decomposition of Pluronic F127 and good structural stability. This work not only provides a green route for high-value utilization of coal in energy storage, but also paves a new way to make hierarchical porous carbon nanofibers from coal for supercapacitor electrodes with high specific capacitance and long cycle life.

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1. Introduction

Due to their unique physical and chemical properties, carbon-based nanomaterials have attracted much attention, which make them have broad application prospects in changing the future of various important applications.^{1,2} The strong demand for environmentally-friendly and inexpensive synthesis strategies to develop carbon nanomaterials has driven much of the global effort. Precursors used to grow carbon-based nanomaterials play a key role in the success of evolving technologies. Most traditional synthetic technologies use expensive carbon raw materials, such as hydro-carbons and graphite, which leads to high cost production, limiting their commercialization. Coal is a rich and cheap natural resource, which is considered as

a green and feasible substitute for the efficient production of carbon nanomaterials. So far, the researchers have done a lot of research to apply coal to energy storage. However, mass production of high-quality carbon nanomaterials through environmentally-friendly and inexpensive strategies remains a significant challenge. Considering the above factors and the factors affecting the performance of supercapacitors,³⁻⁶ the key to electrochemical performance of supercapacitors is to design a green, simple and efficient method to obtain hierarchical porous carbon materials with both high specific surface area and reasonable pore structure, such as carbide-derived carbon,⁷ carbon onions,⁸ carbon nanotubes (CNTs),⁹ carbon aerogels,¹⁰ template porous carbon¹¹ and carbon nanofibers (CNFs).¹²⁻¹⁴

It is common knowledge that CNFs as one of the typical one-dimensional nanomaterials not only provide pathways shortened for the rapid transportation of electrons, but also facilitate the penetration of electrolyte along the direction of the fiber diameter, as well as supplying a high active area to electrochemical reactions. Thus, the unique 1D nanostructure can contribute high performance as electrode materials for electrical double-layer capacitors (EDLCs). The CNFs are also relatively easy to be prepared by electrospinning, which is regarded as one of the most effective methods to prepare carbon nanofibers due to both simple operation and easily controllable experimental conditions.^{15,16} Therefore, many studies have focused on exploiting hierarchical porous carbon nanofibers (HPCNFs) to improve electrochemical properties. At present, most of the HPCNFs electrode materials are produced by

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† Electronic supplementary information (ESI) available: TG curves of oxidized coal, PVA, F127 and precursor CNFs-F4-900, respectively. SEM, TEM and HRTEM images of HPCNFs at 800 °C and 1000 °C. XRD patterns of oxidized coal. XPS survey spectra and XPS elemental analysis of HPCNFs at 800 °C and 1000 °C. N₂ adsorption-desorption isotherms and pore size distributions and BET specific surface areas and porous structure of HPCNFs at different temperature. CV curves and galvanostatic charge-discharge curves of the samples at different temperature. SEM characterizations of the CNFs-F2-900 electrode after discharge/charge cycles. Maximum capacitances and cycling stability of porous carbon nanofiber materials. CV curves and CP curves of the sample CNFs-F2-900 as symmetrical electrode. See DOI: 10.1039/c8ra08949c



electrospinning precursor solution containing pore-forming agent, which mainly include soft template,^{17–22} hard template^{23–26} and soluble salt.^{27,28} Comparing with the hard template and the soluble salt, the soft template can be directly decomposed during the carbonization process without etching procedure by using either toxic or corrosive chemical, which are harmful to the environment. Among soft template agents, Pluronic amphiphilic triblock copolymers such as F127 is extensively used for the fabrication of mesoporous carbon due to their amphiphilic nature.^{17,29–31} Up to now, a few methods have been reported for the direct synthesis of porous CNFs by carbonization of polymers, but it is still a great challenge and highly desirable to design a simple approach for the controllable synthesis HPCNFs.

In our work, water was used as spinning solvent instead of organic solvent, so that it can effectively prevent the generation of the toxic gas in spinning process. The relatively cheap polyvinyl alcohol is made as a spinning aid, and its carbon residue after carbonization is very low, so the coal is the main carbon source.³¹ A facile template F127 is directly decomposed and forms mesopores in the process of carbonation.³² As a result, the sample shows a high capacitance of 265.2 F g⁻¹ at a scan rate of 1.0 A g⁻¹ serving as the binder-free electrodes in EDLCs, which is superior to that of the samples (170 F g⁻¹) without the F127. More significantly, the optimal sample delivers an outstanding rate capability of 220.3 F g⁻¹ at a scan rate of 20.0 A g⁻¹ and maintains over 100% of the initial capacitance after 20 000 cycles at 10.0 A g⁻¹. The material with hierarchical porous structure is a potential electrode material for energy conversion and storage. And this work provides a green way for high-value utilization of coal in energy storage.

2. Methods

2.1 Materials

Coal was obtained from Heishan, Xinjiang, China. PVA (polyvinyl alcohol, 1788, alcoholysis degree: 87.0–89.0% (mol mol⁻¹)) and Pluronic F127 were purchased from Sigma-Aldrich, and KOH (potassium hydroxide, 95.0%, Samchun) was used as a neutralizer. Deionized water was used as the solvent in the whole article.

2.2 Synthesis of HPCNFs

Raw coal powers were dealt as previously reported by our group.¹³ 0.8 g oxidized coal was dissolved in 10 mL distilled water by ultrasonication and neutralized with 0.1 g KOH to obtain a homogeneous dispersion, followed by addition of 0.4 g F127 was stirred into the above solution. After the temperature of the resultant solution was increased to 80 °C, 0.8 g PVA was mixed into the mixture solution by stirring at 80 °C for 24 h to form a homogeneously dispersed solution. Then, the black solution was drawn into a syringe (10 mL) with a stainless-steel nozzle served as an electrode. The fibers were collected on an aluminium foil, which was served as the counter electrode. The receiving distance was 18 cm apart. Meanwhile a potential of 25 kV, feeding rate of 0.06 mm min⁻¹ and temperature of 25 °C

were applied between them, respectively. After electrospinning, the as-prepared nanofibers were placed in an oven and heated at a rate of 1 °C min⁻¹ from room temperature up to 200 °C in air atmosphere, and then were stabilized for 2 h at the final temperature. Afterwards, the nanofibers were heated at 400 °C for 1 h with a heating rate of 2 °C min⁻¹ in a flow of N₂, followed by a further thermal annealing process at 900 °C for 2 h at a heating rate of 5 °C min⁻¹. The little potassium salt was washed with distilled water several times. The as-obtained sample was designated as CNFs-F2-900.

At the same time, three other kinds of the spinning solution with different mass ratios of oxidized coal/PVA/F127 (0.8 g/0.8 g/0.0 g, 0.8 g/0.8 g/0.2 g, 0.8 g/0.8 g/0.6 g and 0.8 g/0.8 g/0.8 g, respectively) were prepared and synthesized by a similar procedure of CNFs-F2-900. These are designated as CNFs-F0-900, CNFs-F1-900, CNFs-F3-900 and CNFs-F4-900 respectively. For the two others of coal based HPCNFs annealed at 800 °C (denoted as CNFs-F2-800) or 1000 °C (CNFs-F2-1000), were carbonized in N₂ atmosphere through a two-stage heating process. They were first heated at 400 °C for 1 h with a heating rate of 2 °C min⁻¹, followed by a further thermal annealing process at 800 °C (or 1000 °C) for 2 h with a heating rate of 5 °C min⁻¹.

2.3 Characterization techniques

Thermogravimetric analysis (TGA) was tested by using a NETZSCH STA449F3-QMS403C instrument under N₂. The morphology and structure features of the product were conducted by using field emission scanning electron microscopy (FESEM Hitachi S-4800), by transmission electron microscopy (TEM Hitachi H-600) and high resolution transmission electron microscope (HRTEM JEM 2100). The Raman spectra of HPCNFs were recorded on a Raman spectroscopy (Bruker SENTERRA, 532 nm laser). Powder X-ray diffraction (XRD, Bruker D8) measurements were conducted to research the crystal structure of the HPCNFs. X-ray photoelectron spectroscopy (XPS, ESCA-LAB250Xi) analysis was performed to detect elemental of the composite. The HPCNFs samples were also recorded by nitrogen adsorption-desorption at 77 K (Autosorb-IQ, Quantachrome, USA).

2.4 Electrochemical property test

The HPCNFs were directly made into binder-free working electrodes, the size of which was 1 cm × 1 cm (about 2.5 mg). The supercapacitors were tested by cyclic voltammetry (CV) on a CHI 660D electrochemical workstation (Chen Hua, China), where galvanostatic charge-discharge and electrochemical impedance spectroscopy (EIS) were tested. In a three-electrode cell, platinum foil and the saturated calomel electrode (SCE) were used as counter and reference electrodes, respectively. The specific capacitance (C_s) of the HPCNFs electrode was calculated *via* eqn (1)

$$C_s = I\Delta t / (m\Delta V) \quad (1)$$



where I , Δt , ΔV and m are the current, the discharging time, the potential window and the mass of the active material, respectively. The cycling performance of the electrode was tested on LAND charge-discharge instrument (CT2001A, Wuhan LAND Electronics Co., Ltd).

In the two-electrode system, the specific capacitance (C) values were obtained *via* eqn (2)

$$C = I\Delta t / (M\Delta V) \quad (2)$$

where M is the total mass of active materials on both electrodes.

3. Result and discussion

The details for the preparation of HPCNFs are showed in Fig. 1. Firstly, the oxidized coal was dissolved into the distilled water under ultrasonication and neutralized with a little of KOH, the F127 were also added into the above solution. After the temperature of the resultant solution was increased to 80 °C, the PVA was mixed into the mixed solution and was stirred at 80 °C for 24 h to form a uniform spinning solution. Secondly, the nanofibers were made by an electrostatic spinning method. Thirdly, the nanofibers were obtained by preoxidation and then carbonization. The formation of mesopores may be as follows: (i) the F127 forms micelles in the black solution and the hydrophilic PEO chains are around the micelle center, and the hydrophobic group is outward, so the size of micelles gets bigger at higher temperature. (ii) When the PVA is added into the black solution, hydroxyl groups of PVA form hydrogen bonds with F127 micelles. (iii) the F127 micelles are carbonized to produce the mesopores. The process is different from the self-assembly of F127.³² Thus the resulting mesopores are irregular arrangement.

The contents of carbon in the samples were measured by thermogravimetric (TG) curves as shown in Fig. S1.† From the TG curves above, the mass losses of oxidized coal, PVA, F127 and precursor CNFs-F4-900 are 53.4%, 97.2%, 98.8% and 57.4%, respectively. The detected carbon contents of 94% in the sample are provided by oxidized coal, indicating that the obtained nanofibers are real coal-based carbon materials. In other words, the residual carbon comes mainly from coal in the carbon nanofibers.

The morphologies of the resultant HPCNFs were characterized by using SEM and TEM. The HPCNFs shown in Fig. 2a–e

present a long fibrous morphology and homogeneous diameter distributions in the range of 70–150 nm. As illustrated in Fig. 2f–o, there are clearly visible mesopores in the samples except for CNFs-F0-900, which is due to the effect of pore-forming agent F127. Furthermore, the mesopores in the sample CNFs-F2-900 are more uniform distribution than the other samples, which could be attributed to the moderate amount of pore forming agents F127. In the sample CNFs-F1-900 a small amount of F127 could be not enough to be used to make mesoporous. However, in the sample CNFs-F3-900 the slightly more F127 makes the mesopores unevenly distributed. And the mesopores distribution is more uneven due to accumulation of excessive F127 in CNFs-F4-900. Fig. 2p–t shows that there are a large number of visible micropores and mesopores in HPCNFs, which can effectively decrease the ion transport dimensions for further shortening ion diffusion length. Moreover, the carbon in the samples was amorphous. The structure of HPCNFs affected by carbonization temperature is also studied. As shown in Fig. 2c, h, m, S2a–c and e–f,† the diameter of the samples CNFs-F2-800, CNFs-F2-900 and CNFs-F2-1000 decrease due to the volume shrinkage with the carbonization temperature from 800 °C to 1000 °C. Besides, more broken mesopores can be observed in the CNFs-F2-1000 because of high carbonization temperature. The few lattice fringe gradually becomes obvious with the increase of the carbonization temperature from the HRTEM images of the HPCNFs (Fig. 2r, S2d and h†).

The effect of different F127 mass and temperature on the structure of the samples is revealed by X-ray diffraction (XRD) patterns as presented in Fig. 3a. Fig. S3† presents that oxidized coal exhibits the wide (002) and the weak (100) peaks of graphitic carbon around 23° and 42°. As displayed in Fig. 3a, compared with the peaks of oxidized coals, the peaks around 23° of the samples are shifted to a relatively high angle, it shows that the graphitization degree is reduced during carbonization.^{33,34} Meanwhile, the broad diffraction peaks of the HPCNFs at around 16° is attributed to (100) plane.^{35,36} These results confirm that the loading of F127 has little effect on the structure of the resultant carbon at the same temperature 900 °C and different temperature. Thus the HPCNFs samples are low crystalline carbon, which is well in accordance with results of the HRTEM images of the HPCNFs.

Raman spectra were performed to study the structures of the samples (Fig. 3b). The two prominent peaks around 1347 and 1580 cm⁻¹ are identified as D and G bands of carbon materials, corresponding to the disordered and graphitic structures, respectively. The relative intensity of the D bands to the G bands (I_D/I_G), indicating the disordered degree of carbon materials,^{37–39} is calculated to be about 0.96, 0.97, 0.97, 0.97, 0.97, 0.97 and 0.98 for CNFs-F2-800, CNFs-F0-900, CNFs-F1-900, CNFs-F2-900, CNFs-F3-900, CNFs-F4-900 and CNFs-F2-1000, respectively. The results show that the amount of F127 and carbonation temperature have no significant effect on degree of graphitization. The results of Raman are in agreement with that of the XRD.

The surface compositions of the CNFs-F2-800, CNFs-F2-900 and CNFs-F2-1000 were studied by using the X-ray

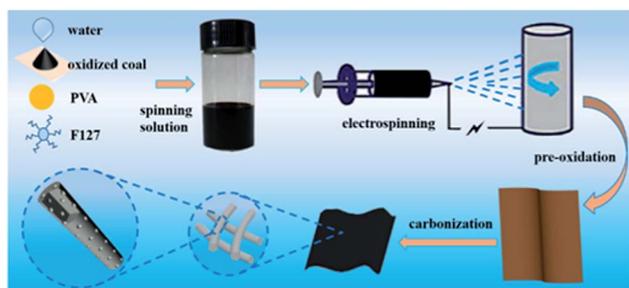


Fig. 1 The synthetic scheme of HPCNFs.



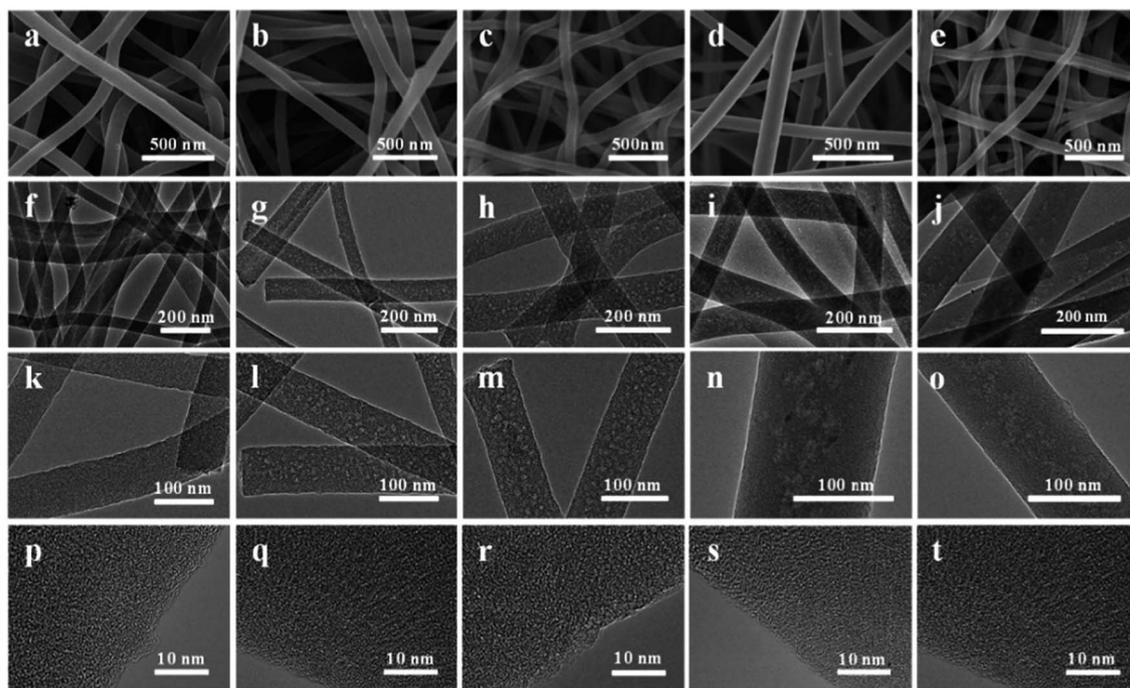


Fig. 2 SEM images of HPCNFs (a) CNFs-F0-900; (b) CNFs-F1-900; (c) CNFs-F2-900; (d) CNFs-F3-900; (e) CNFs-F4-900; TEM images of HPCNFs (f and k) CNFs-F0-900; (g and l) CNFs-F1-900; (h and m) CNFs-F2-900; (i and n) CNFs-F3-900; (j and o) CNFs-F4-900 and HRTEM images of HPCNFs (p) CNFs-F0-900; (q) CNFs-F1-900; (r) CNFs-F2-900; (s) CNFs-F3-900; (t) CNFs-F4-900.

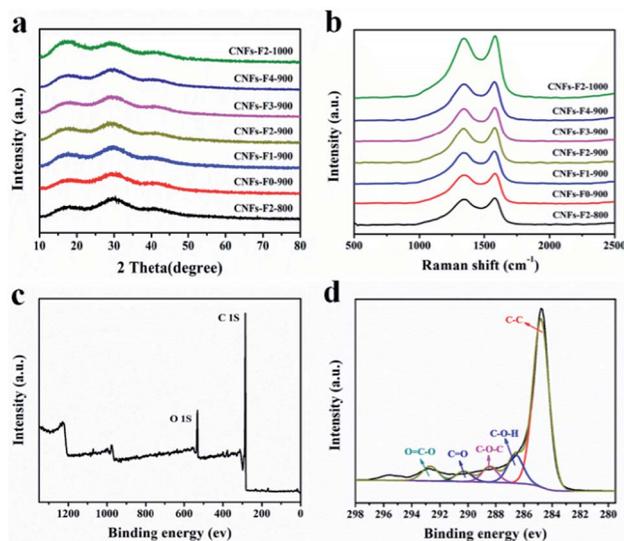


Fig. 3 (a) XRD patterns and (b) Raman spectra of HPCNFs; (c) XPS survey spectra and (d) XPS of C1s region of the CNFs-F2-900.

photoelectron spectrum (XPS). Fig. 3c shows that the CNFs-F2-900 is composed of C and O, peaking at the binding energies of them are around 285.0 eV and 532.0 eV. In Fig. 3d the C1s peak of CNFs-F2-900 exists in the form of C-C, C-O-H, C-O-C, C=O and O=C-O.^{40,41} The elemental composition of the CNFs-F2-800 and CNFs-F2-1000 is the same as that of CNFs-F2-900 as described in Fig. S4a-d,[†] but the element contents of the three samples are different (Table S1[†]). The O contents of the CNFs-

F2-800, CNFs-F2-900 and CNFs-F2-1000 are respectively 12.70%, 10.33% and 5.40% because increasing the carbonization temperature leads to a rapid decrease in O contents. Therefore, these oxygen-enriched functionalities¹¹ can both enhance wettability and produce pseudo capacitance to improve capacitance performance.

To make sure that the influence of F127 pore-forming agent dosage on the pore structures of the samples CNFs-F0-900, CNFs-F1-900, CNFs-F2-900, CNFs-F3-900 and CNFs-F4-900, N₂ adsorption-desorption isotherms were studied to measure the specific surface area and the porous distribution of the samples, and all the parameters are summarized shown in Fig. 4 and Table 1, respectively. As depicted in Fig. 4a, the samples exhibit steep uptakes typical type-IV curves at relative pressure (P/P_0) = 0.1 and an H₄-type hysteresis loop in the relative pressure range of 0.4–0.6, indicating the formation of abundant micropores and mesopores, respectively. Pore-size-distribution curves show that the size of the majority of the mesopores fall between 2.5 and 4.0 nm (Fig. 4b), corresponding to the hydrophobic chain (F127) length,^{42,43} which play an important role in fast an ion transfer.³ Furthermore, the minimum size of micropores in CNFs-F0-900, CNFs-F1-900, CNFs-F2-900, CNFs-F3-900 and CNFs-F4-900 are around 0.521, 0.511, 0.59, 0.56 and 0.61 nm, respectively, which are close to the size of hydrolyzed K⁺ ions (0.331 nm),⁴⁴ implying that the micropores exhibited in all samples are beneficial to their capacitive performance. From Table 1 the surface areas, mesopore surface areas, total pore volumes and mesopore volume of CNFs-F2-900 and CNFs-F1-900 gradually get larger than that of CNFs-F0-900. The



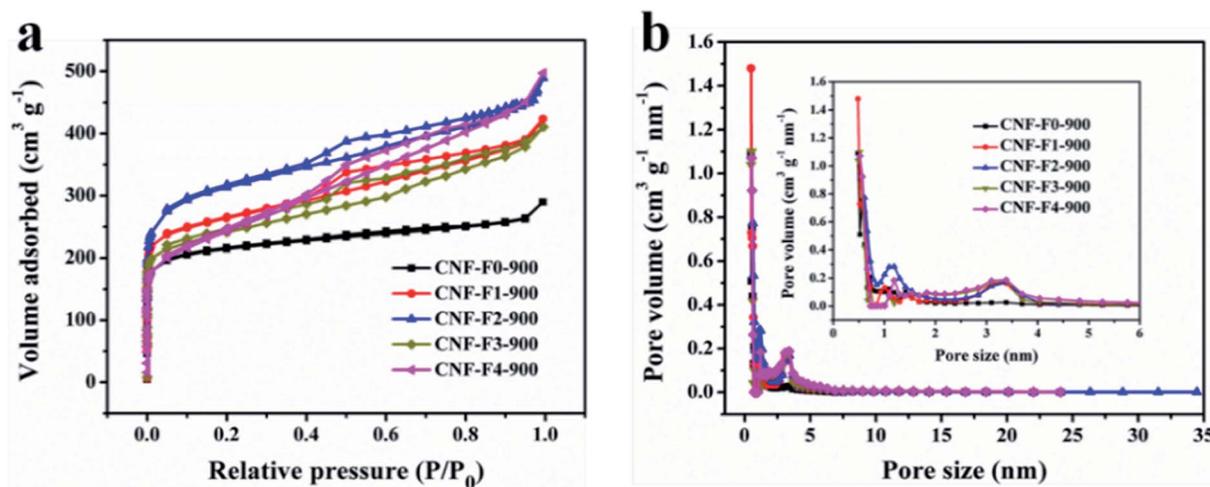


Fig. 4 N_2 adsorption–desorption isotherms (a) and pore size distributions (b) of CNFs-F0-900, CNFs-F1-900, CNFs-F2-900, CNFs-F3-900 and CNFs-F4-900 at the same temperature (900 °C).

micropore surface areas and micropore volumes of them are not very different. This could be attributed to sacrificial template F127 removed to create more mesopores with increasing F127 pore-forming agent dosage and most of micropores from little potassium salt and decomposition of PVA and coal. But when the F127 mass increases to 0.6 g and 0.8 g, more F127 micelles form hydrogen bonds with hydroxyl groups of PVA, resulting in F127 dispersed inhomogeneous in solution, as the parameter shown in the Fig. 4 and Table 1 for CNFs-F3-900 and CNFs-F4-900. More F127 is used to produce more mesoporous so that mesopore surface areas gradually increase. While more micropores were destroyed, so micropore surface areas are reduced. The effect of carbonization temperatures on the specific surface and porous structure of the samples HPCNFs are also investigated (Fig. S5†). And Table S2† displays that the mesopore surface areas and mesopore volume of CNFs-F2-800, CNFs-F2-900 are not significantly different. The micropore surface areas and micropore volumes of CNFs-F2-900 at 900 °C are larger than that of CNFs-F2-800 at 800 °C and CNFs-F2-1000 at 1000 °C, which is due to little potassium salt and small molecular gas released by the degradation of PVA and carbonization of coal during carbonization. It can be clearly noted that the specific surface area and reasonable pore size distribution determine the capacitance performance.

As described in Fig. 5a, the EDLCs behavior of electrode materials is nearly rectangular shape between -1 V and -0.1 V. The cyclic voltammograms (CV) curve of CNFs-F2-900 encircles the largest curve area in all the samples CNFs-F0-900, CNFs-F1-900, CNFs-F2-900, CNFs-F3-900 and CNFs-F4-900 at 100 $mV s^{-1}$, suggesting a higher specific capacitance. Meanwhile, the galvanostatic charge–discharge curves of these samples at the same scan rate of 1 $A g^{-1}$ are showed in Fig. 5b, the specific capacitance of CNFs-F2-900 is higher than those of others, because the CNFs-F2-900 has more mesopores than the samples CNFs-F0-900 and CNFs-F1-900 that provide pathways for electrolyte transport. Compared to CNFs-F3-900 and CNFs-F4-900, larger micropore surface areas of CNFs-F2-900 supply more active sites to electroadsorption. The CV curves of CNFs-F2-900 (Fig. 5c) at different scan rates always retains their quasi-rectangular shape from -1 to -0.1 V, presenting the EDLCs characteristics in the charge–discharge process. Furthermore, the CV curve of CNFs-F2-900 still could retain a stable rectangular shape at a high scan rate up to 300 $mV s^{-1}$, exhibiting a good rate capability of electrode materials. Fig. 5d shows that the galvanostatic charge–discharge curves of the CNFs-F2-900 at from 0.5 to 50.0 $A g^{-1}$. All the charge–discharge curves are nearly linear and symmetrical, confirming its capacitor-like feature and good electrochemical reversibility, which result

Table 1 BET specific surface areas and porous structure of CNFs-F0-900, CNFs-F1-900, CNFs-F2-900, CNFs-F3-900 and CNFs-F4-900

| Sample | S_{BET}^a ($m^2 g^{-1}$) | S_{meso}^b ($m^2 g^{-1}$) | S_{micro}^c ($m^2 g^{-1}$) | V_{total}^d ($cm^3 g^{-1}$) | V_{meso}^e ($cm^3 g^{-1}$) | V_{micro}^f ($cm^3 g^{-1}$) | D_{ap}^g (nm) |
|-------------|------------------------------|-------------------------------|--------------------------------|---------------------------------|--------------------------------|---------------------------------|-----------------|
| CNFs-F0-900 | 821 | 123 | 698 | 0.44 | 0.14 | 0.30 | 2.5 |
| CNFs-F1-900 | 992 | 289 | 703 | 0.63 | 0.34 | 0.29 | 2.6 |
| CNFs-F2-900 | 1161 | 428 | 733 | 0.76 | 0.45 | 0.31 | 2.7 |
| CNFs-F3-900 | 891 | 502 | 389 | 0.72 | 0.48 | 0.24 | 3.2 |
| CNFs-F4-900 | 806 | 571 | 235 | 0.70 | 0.60 | 0.10 | 3.7 |

^a BET surface area. ^b Micropore surface area calculated using the $V-t$ plot method. ^c Mesopore surface area calculated using the $V-t$ plot method. ^d The total pore volume calculated by single point adsorption at $P/P_0 = 0.99$. ^e The mesopore volume calculated using the $V-t$ plot method. ^f The micropore volume calculated using the $V-t$ plot method. ^g Average pore size.



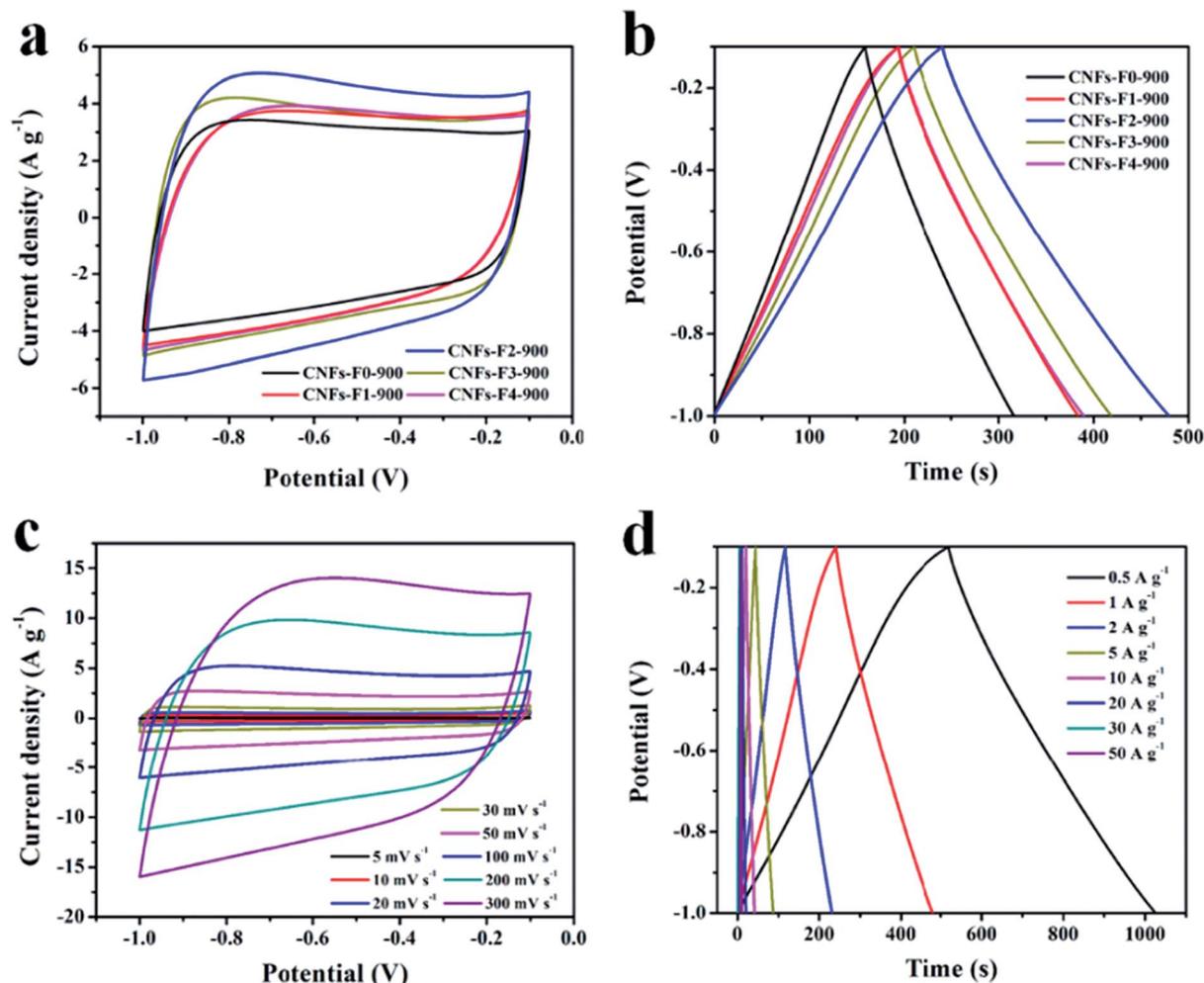


Fig. 5 (a) CV curves of the samples with different F127 mass at the scan rate of 100 mV s^{-1} ; (b) galvanostatic charge–discharge curves of the samples with different F127 mass at the same scan rate of 1.0 A g^{-1} ; (c) CV curves and (d) galvanostatic charge–discharge curves of the CNFs-F2-900 sample at different current densities.

from rapid diffusion of electrolyte ions in electrode materials.⁸ The specific capacitance of CNFs-F2-900 is much larger than that of CNFs-F2-800 and CNFs-F2-1000 (Fig. S6†), indicating that superior porosity resulting from the pore-making agent (F127) and the proper carbonization.

The specific capacitance for CNFs-F2-900 (Fig. 6a) at various current density is calculated from the discharge curves. As shown in Fig. 6a, the specific capacitance of CNFs-F2-900 is as high as 282.1, 265.2, 253.1, 238.2, 226.3, 220.3, 214.4 and 193.6 F g^{-1} at 0.5, 1.0, 2.0, 5.0, 10.0, 20.0, 30.0, and 50.0 A g^{-1} , respectively. The results suggest that about 73.0% of the capacitance is retained with increasing of the charge–discharge rate from 1.0 to 50.0 A g^{-1} . By comparison, the specific capacity of the CNFs-F2-900 is the highest in all electrode materials, which are ascribed to the larger accessible specific surface area, desirable pore size distribution and more available mesoporous channels as described above.

The electrochemical impedance spectra (EIS) of the all the samples were measured to compare the ion-transport behavior and electrical resistance (Fig. 6b). The results show that all the

electrodes are nearly perpendicular to the imaginary axis in the low frequency region, presenting an ideally capacitive behavior of the samples and the diffusion and migration of the electrolyte ions within the electrode structure. In the high frequency region, it could be clearly found that the first intersection point in real axis is solution resistance (R_s),⁴⁵ which is 0.477, 0.484, 0.485, 0.483 and 0.490 Ω for CNFs-F0-900, CNFs-F1-900, CNFs-F2-900, CNFs-F3-900 and CNFs-F4-900, respectively. Moreover, the diameter of a semicircle is a characteristic of charge transfer resistance (R_{ct}), and the resistance to electrolyte transport for CNFs-F2-900 is observed to be smallest in all the samples, due to the more available mesoporous channels. Hence, hierarchical porous structure of the CNFs-F2-900 is beneficial for charge transfer and ion diffusion.

The response and better rating of the capacitor have further been proved its long-term cyclic stability. The cycling capacitance performance of the CNFs-F2-900 was researched by galvanostatic charge–discharge cycling at 10.0 A g^{-1} for 20 000 cycles (Fig. 7). As shown in the inset image in Fig. 7, after alternating charge–discharge process, the specific capacitance



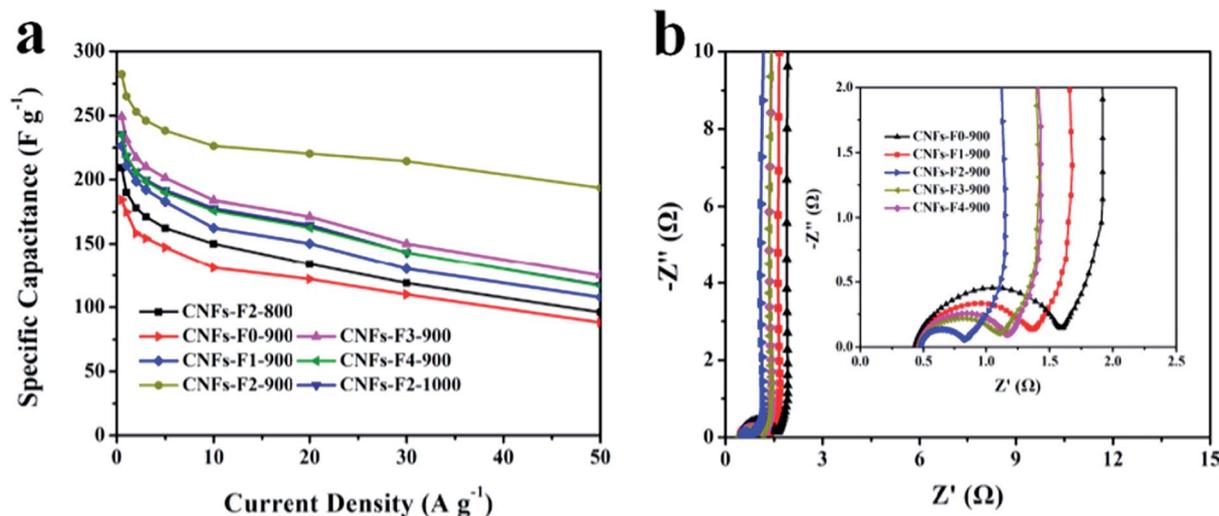


Fig. 6 (a) Specific capacitances of all samples at the current density of 0.5–50.0 A g⁻¹; (b) electrochemical impedance spectra of HPCNFs as Nyquist plots.

for CNFs-F2-900 still reaches 231.3 F g⁻¹ at 10.0 A g⁻¹. It is observed that CNFs-F2-900 retains about 105% of their initial capacitance after 20 000 charge/discharge cycles, indicating that excellent electrochemical stability of the CNFs-F2-900. Furthermore, the SEM image in Fig. S7† shows the CNFs-F2-900 electrode can keep its original morphology after 20 000 cycles of durability testing, demonstrating the electrode material possesses a good stability. The specific capacitance and cycling stability of our materials were compared with previous carbonaceous works of CNFs and the data are listed in Table S3.† Furthermore, the two-electrode cell was measured in 6 M KOH. Fig. S8a† shows the CV curves of the CNFs-F2-900 at from 10 to 100 mV s⁻¹, and its galvanostatic charge–discharge curves at various current densities (Fig. S8b†). We found that CNFs-F2-

900 has a specific capacitance of 87 F g⁻¹ at 1 A g⁻¹. The capacitance retention percentage at 30 A g⁻¹ relative to that at 1 A g⁻¹ is 65.5%.

With regard to the above good electrochemical properties of CNFs-F2-900, there are three reasons as followed. Firstly, large effective specific surface provides more electroactive sites for ion accumulation and energy storage. Secondly, ion diffusion length is further shortened, since the hierarchical porous structure in one-dimensional CNFs can effectively decrease the ion transport channel. Thirdly, the empty spaces in HPCNFs accommodate the volume changes connected with electrochemical reactions, so the stable structure makes the electrode material with an excellent cycle stability.

4. Conclusion

In summary, we have designed a green method to prepare coal based HPCNFs. When the mass of F127 is 0.4 g and the carbonization temperature is 900 °C, the CNFs-F2-900 exhibits excellent electrochemical performance as supercapacitor electrodes, in particular attaining high specific capacitance at different current densities (265.2 F g⁻¹ at 1.0 A g⁻¹ and 220.3 F g⁻¹ at 20.0 A g⁻¹) and outstanding cycling stability without any decrease in the specific capacitance after 20 000 cycles. Therefore, this work paves a desired route to produce HPCNFs.

Conflicts of interest

There are no conflicts to declare.

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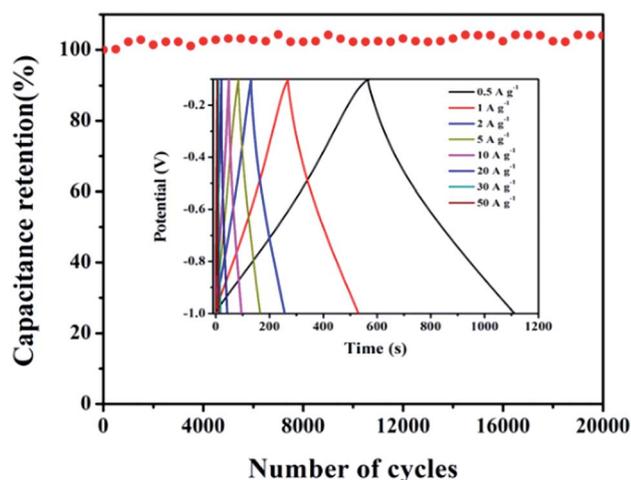


Fig. 7 Cycling performance of the CNFs-F2-900 electrode at 10.0 A g⁻¹ (the inset displays galvanostatic charge–discharge curves of the CNFs-F2-900 electrode at different current density after 20 000 discharge/charge cycles at 10.0 A g⁻¹) from coal as highly efficient electrodes materials for supercapacitors.



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