The roles of Lewis acidic additives in organotransition metal catalysis

Joseph Becica and Graham E. Dobereiner

We describe recent examples of prominent reactions in organic synthesis that involve transition metal and Lewis acid cocatalysts. Introducing Lewis acid additives to transition metal catalysis can enable new reactivity or improve activity and selectivity of an existing process. Several studies are highlighted to illustrate the possible roles of Lewis acids in catalytic mechanisms. The uses of Lewis acid additives in bond-breaking catalysis (C–C activation, C–H activation, and hydrogenolysis reactions) and bond-forming catalysis (Au-catalysed alkyne functionalisation, Pd-catalysed C–C and C–N bond formation) are reviewed.

1. Introduction

In pursuit of more efficient organotransition metal catalysis, many chemists have boosted activity and/or selectivity using Lewis acid additives.1 Significant recent contributions have advanced the mechanistic understanding of transition metal-Lewis acid synergy, and therefore it is timely to provide a concise review of highlights of this area. Out of a broad array of examples in which Lewis acidic additives are used in organometallic catalysis, we present here four classes of Lewis acid effects in catalytic reactions, chosen to highlight a few important ways in which Lewis acids may influence organometallic reactions. This Review is intended as a guide for practitioners who seek to improve catalytic reactions by the addition of additives and to understand the role of additives from a mechanistic perspective. We focus on mechanistic effects of Lewis acids throughout catalytic cycles, including alkylaluminums, alkyl and arylboranes, alkylzincs, and metal halides and triflates (trifluoromethanesulfonate, or OTf) of various cations. These classes of Lewis acids play a role in the cooperative activation of strong bonds, such as in the oxidative addition of C–C and C–H bonds to transition metals, or the hydrogenolysis of C–O and C–N bonds. Lewis acids can also facilitate bond-forming reactions, such as Au-catalysed reactions of alkenes (and related reactions), as well as Pd-catalysed C–C and C–N bond forming reactions.

Lewis acids can play myriad roles within a given transition metal-catalysed reaction (Fig. 1). Some examples in this Review
will involve tandem catalysis, in which two discrete catalytic cycles occur simultaneously in a single reaction mixture, where M₁ is a transition metal catalyst and M₂ is an acid catalyst. More complex mechanisms see the involvement of a Lewis acid cocatalyst (LA) upon a single catalytic cycle mediated by a transition metal catalyst (M). In a simplified mechanistic model, the Lewis acid can affect the overall cycle by forming Lewis acid–base adducts with different components of the catalytic reaction.

For example, cooperative substrate-Lewis acid interactions (S₁-LA) or catalyst-Lewis acid interactions (M-LA) can promote the turnover-limiting step of catalysis or direct the reaction pathway in a selectivity-determining step. In this Review, we will encounter examples of reactions in which the Lewis acid is revealed to take several possible roles, such as activating substrate toward a transition metal centre, accelerating a key elementary organometallic step (i.e., migratory insertion, transmetalation, or reductive elimination), affecting the catalyst activation or deactivation pathways, or promoting the controlled release of a Brønsted acid in situ. Here we highlight recent examples that elucidate the numerous possible sources of Lewis acid effects in organometallic catalysis and identify the modes of action of the Lewis acid in different reaction contexts.

One challenge is understanding why a particular Lewis acid is more effective than another in a given reaction. Electrophilic species can be found across the periodic table, and the methods to systematically categorize Lewis acids or measure Lewis acidity are limited. In select cases, the behaviour of the Lewis acids can be differentiated. For example, Organ and coworkers suggest that the different susceptibility of boron Lewis acids to autoxidation can result in different reaction outcomes in Pd and boron-catalysed C–N bond forming reactions. Another example is provided by Marks and coworkers, who show that the capacity of Pd and metal triflate cocatalysts to affect the hydrolysis of ethers is dependent on the effective charge density of the metal triflate cation employed.

In each section, we highlight synthetic methods enabled by cocatalytic systems, as well as insights into potential modes of cooperativity. There are several methods for rationalising Lewis acid effects observed in catalytic reactions. Density functional theory calculations enable the assessment of the role of Lewis acid–base adducts of substrates and key catalyst intermediates on catalytic reaction pathways, as well as parametrisation of Lewis acids. Experimental insight into catalyst-Lewis acid cooperativity can be ascertained through the isolation of catalytic intermediates or model complexes thereof and assessing their reactivity towards Bronsted or Lewis acids. Lastly, monitoring of reaction rates can confirm an empirical Lewis acid effect and provide insight into possible reaction pathways. Here we piece together these clues to unveil numerous modes of catalyst-Lewis acid interactions in productive catalysis.

2. Cooperative catalyst systems for C–C and C–H bond activation

A representative example of transition metal/Lewis acid cooperative bond activation is DuPont’s Adiponitrile Process (Scheme 1), which is thought to proceed via oxidative addition of HCN, Ni–H insertion of diene, and reversible elimination to generate alkynitrile. In the presence of a Lewis acid cocatalyst (typically, BPh₃), the Ni/phosphine-catalysed hydrocyanation of 1,3-butadiene (1) shows increased rate and selectivity for the desired linear adiponitrile (2) versus branched isomers. Since Lewis acids can promote the oxidative addition of a C–CN bond to Ni(0) – i.e., the microscopic reverse of alkynitrile-forming reductive elimination – branched nitriles can be thus re-engaged by the catalyst, and again eliminated to form linear 2. Jones revealed that a cooperative Ni/LA system reacts preferentially with C–CN bonds over activated C–H bonds, e.g. in allylcyanide. Together, these early works in cooperative bond activation are the basis for a wide variety of synthetic methods in C–C and C–H bond activation discussed in this section.

2.1 Synthetic methods that apply Ni and Lewis acid cocatalysts

2.1.1. C–C bond activation. Ni and Lewis acid catalysts enable the use of nitriles as electrophiles in C–C bond forming
reactions.\textsuperscript{9,10} Nakao, Hiyama and coworkers\textsuperscript{11} report the reaction of benzonitrile 4 with 4-octyne (5) via Ni(COD)\textsubscript{2}/PPhMe\textsubscript{2} and AlMe\textsubscript{3}Cl to provide alkenylnitrile 6 in 96\% yield (Scheme 2). The reaction proceeds by cooperative oxidative addition of the C-CN bond of 4, followed by migratory insertion of the alkyne into the Ni-C\textsubscript{Ar} bond, and reductive elimination of alkenylnitrile product. The Lewis acid also increases selectivity for C-CN bond cleavage versus C-H bond cleavage, as demonstrated in the alkenylation of cyanoindole 7 with 5 to provide the 3-alkenylnindole 8 (Scheme 3).\textsuperscript{11} In the absence of Lewis acid, C-H activation of the indole instead results in 2-alkenylnindole 9, without cleavage of the C-CN bond.

Other nitriles, such as allylnitriles,\textsuperscript{12} benzynitriles,\textsuperscript{13} and alkynynitriles\textsuperscript{14} can be employed in similar catalytic scaffolds. Intramolecular versions in which alkenes are used as the coupling partner also have been reported,\textsuperscript{15,16} including variants which include Pd catalysts.\textsuperscript{17–20} The reaction of cyanoester 10, a more activated electrophile, with 5 to give alkenylnitrile 11 uses an electron-poor phosphine \textsubscript{ArF} in combination with Ni(COD)\textsubscript{2} and B(C\textsubscript{6}F\textsubscript{5})\textsubscript{3} as the Lewis acid cocatalyst (Scheme 4).\textsuperscript{21} Other C-C bonds can be activated such as in anhydrides,\textsuperscript{22} cyclopropanes,\textsuperscript{23} and esters.\textsuperscript{24} In the latter example, the reaction of arylnitrile 12 with 2-butyne (13) and catalytic Ni(COD)\textsubscript{2}/P(CH\textsubscript{2}Ph)\textsubscript{3} results in sequential C-C bond activation events resulting in formation of coumarin 14 and elimination of arylnitrile 15.

The activation of C(sp\textsuperscript{3})-CN bonds in unactivated aliphatic nitriles is more challenging, due to lower C-CN reactivity and competing \textbeta-H hydride elimination of alkylnickel intermediates. These challenges can be overcome using a directing group on the nitrile to kinetically stabilize the resulting alkylnickel intermediate, and/or using a sterically-demanding ligand to reduce the rate of \textbeta-H elimination.\textsuperscript{25} The pyrrolidine-substituted nitrile 16 and 5 react to give alkenylnitrile 17 in 86\% yield using a Ni(COD)\textsubscript{2}/SPhos/AlMe\textsubscript{3} catalyst mixture (Scheme 5). A clever manifestation of cooperative C(sp\textsuperscript{3})-CN bond activation is realized by the Morandi group, in which alkylnitriles are used as HCN group transfer reagents (Scheme 6).\textsuperscript{26,27} The pyrrolidine-substituted nitrile 18 and 19 are converted to alkynynitrile 20 and isobutylene (21), using Ni(COD)\textsubscript{2}/DPE-Phos/AlMe\textsubscript{3}; isobutyrylnitrile is selected as HCN transfer agent due to removal of gaseous isobutylene, driving selective formation of the desired product. A subsequent publication from the Morandi group uses a similar method to synthesize arylnitriles.\textsuperscript{28}

### Scheme 2

Ni and Al-catalysed carboxylation of alkynes with benzonitriles.\textsuperscript{11}

### Scheme 3

Divergent reactivity in the Ni and Al-catalysed carboxylation of alkynes with 3-cyanoindoles.\textsuperscript{11}

### Scheme 4

Alkynylation of cyanoesters\textsuperscript{25} (top) and sequential C-CN and C-CO bond activation to form coumarins\textsuperscript{24} (bottom).

### Scheme 5

Activation of alkylnitriles using a directing group.\textsuperscript{24}
hydroarylation of 5 with pyridine (22) via Ni(COD)$_2$/IPr/MAD (IPr = 1,3-bis(2,6-diisopropylphenyl)imidazole-2-ylidene) results in alkylation of the 4-position of the arene to yield 25 (70% yield), and similar regioselectivity is observed when this catalyst system is used for the hydroarylation of alkenes.

2.2. The role of the Lewis acid in cyanofunctionalisation and C–H functionalisation

In these reactions, the Lewis acid influences selectivity of C–C and C–H bond activation. Sakaki and coworkers investigated the mechanism of the aforementioned dual C–C cleavage reaction of 12 to form the coumarin 14 by computation. An abridged mechanism determined by DFT calculations is depicted in Fig. 2 to reflect key elementary steps of the reaction; the mechanism proceeds by oxidative addition of the arylcyanide to Ni(0), followed insertion of the alkyne into an arylnickel bond, followed by C–C bond formation and β-aryl elimination, yielding 14 and benzonitrile byproduct 15. The calculations show that binding of the Lewis acid (which has been simplified from MAD to Al(OMe)$_3$) at both the cyano and ester groups of the substrate are essential to achieving the two C–C cleavage steps of the reaction. In the presence of two equivalents of Lewis acid, both N-bound and O-bound, the transition state energies for both C–CN bond cleavage events are significant reduced, and moderate activation barriers are found: for C–CN activation, the Δ$^\circ$G‡° is +14.4 kcal mol$^{-1}$, and for C–CO activation the Δ$^\circ$G‡° is +17.3 kcal mol$^{-1}$, whereas in the absence of any Lewis acid, the barriers are Δ$^\circ$G‡° = +44.8 kcal mol$^{-1}$ and Δ$^\circ$G‡° = +23.8 kcal mol$^{-1}$, respectively. The significant reduction in barrier of C–CN cleavage is attributed to charge transfer from the Ni 3d$^\pi$ to the C–CN σ$^*$ and increasing population of the LUMO of antibonding character is crucial to C–CN bond cleavage. The
computations support the experimental result that no conversion of 12 is observed in the absence of Lewis acid.

Another computational study performed by Guan and co-workers provides further insight into the regioselectivity observed by Ni/LA catalyst systems. The reaction of fluorinated benzonitrile 26 with 2-butyne (13) is investigated using a simplified Ni/DPE-Phos catalyst with BPh₃ cocatalyst. The selectivity of rate-limiting C–CN bond cleavage to yield the alkenylnitrile 27 is evaluated in comparison to other competitive bond cleavages (C–H, C–F) in both the presence and absence of BPh₃. In the absence of BPh₃, the energy barriers of C–CN and C–H bond cleavage are similar ($\Delta G^{\ddagger} = +22.0 \text{ kcal mol}^{-1}$ and $\Delta G^{\ddagger} = +21.2 \text{ kcal mol}^{-1}$, respectively; Table 1). Upon binding of BPh₃ to 26, the barrier of C–CN cleavage is reduced to +17.4 kcal mol$^{-1}$, whereas the barrier of C–H cleavage increases to +28.8 kcal mol$^{-1}$, confirming that the Lewis acid can increase the kinetic preference for C–CN cleavage versus other reactive bonds. The BDEs of the C–H and C–F bonds in the 26:BPh₃ Lewis acid–base adduct are determined to be similar in free 26; however, the BDE of the C–CN bond is significantly reduced ($\Delta H = +136.2 \text{ kcal mol}^{-1}$, free 26; $\Delta H = +112.7 \text{ kcal mol}^{-1}$ 26:BPh₃) upon adduct formation. The results affirm findings in other studies, where Lewis acids reduce activation barriers of oxidative addition of cyanoester and alkenylnitrile to Ni(0). Lewis acids can also reduce the barrier to migratory insertion of the alkyne into Ni–C bonds.

To summarize, advances in C–C and C–H bond activation have been enabled by cooperative Ni and Lewis acid catalysis, and numerous novel synthetic methods have resulted. A survey of examples indicates dependence of the electron-donating capacity of the ancillary ligand and the electron-withdrawing capacity of the Lewis acid in these catalytic systems strongly dictates the catalyst activity and selectivity. To generalize, the most obvious role of the Lewis acid in this context is to effect substrate activation, though effects of chemo- and regioselectivity are also identified. In the following sections, the origins of catalyst-Lewis acid cooperativity will be more ambiguous.

### Table 1 Effect of Lewis acids on the Ni-catalysed activation of 2,3,5,6-tetrafluorobenzonitrile

<table>
<thead>
<tr>
<th>Bond activation</th>
<th>$\Delta G^{\ddagger}$ (kcal mol$^{-1}$)</th>
<th>$\Delta G^{\ddagger}$ w/BPh₃ (kcal mol$^{-1}$)</th>
<th>BDE (kcal mol$^{-1}$)</th>
<th>BDE w/BPh₃ (kcal mol$^{-1}$)</th>
</tr>
</thead>
<tbody>
<tr>
<td>C–CN</td>
<td>+22.0</td>
<td>+17.4</td>
<td>+136.2</td>
<td>+112.7</td>
</tr>
<tr>
<td>C–F₃</td>
<td>+27.8</td>
<td>+33.1</td>
<td>+125.8</td>
<td>+125.8</td>
</tr>
<tr>
<td>C–F₉</td>
<td>+32.8</td>
<td>+43.0</td>
<td>+125.2</td>
<td>+125.0</td>
</tr>
<tr>
<td>C–H</td>
<td>+21.2</td>
<td>+28.8</td>
<td>+122.5</td>
<td>+122.8</td>
</tr>
</tbody>
</table>

### Scheme 8 Conversion of levulinic acid to 2-methyltetrahydrofuran. By modifying Klankermeyer and Leitner’s catalyst and conditions, the activity and selectivity of the reaction can be dramatically affected. Ru(acac)$_3$, tri(octyl)phosphine, and NH$_4$PF$_6$ selectively converts 28 to γ-valerolactone (29) in >99% yield (Scheme 9), without further reduction of the lactone. In contrast, the use of Ru(acac)$_3$, triphos, along with a mixed acid system (using a sulfonic acid (aIL) and NH$_4$PF$_6$) provides excellent selectivity for the complete reduction of 28 to MeTHF (31) in 92% yield.

In the latter reaction, an off-cycle catalyst resting state (triphos)Ru(CO)(H)$_2$ (32) was isolated from the reaction.
mixture. The carbonyl ligand is most likely derived from de-carbonylation of aldehyde formed in situ. As precatalyst, 32 has lower catalytic activity than the combination of Ru(acac)₃ and triphos, but can be reactivated by adding substrate and/or acid (Fig. 3). The ability of an acid to restore activity is dependent on pKₐ, conjugate base identity, and Ru/acid ratio. Acid could react with neutral [Ru(H)₂] to form the non-classical hydride [Ru(H₂)H⁺] (33). Subsequent dihydrogen dissociation could then open up a site for substrate binding.

In a related amide hydrogenation, activity of (triphos)Ru(OAc)(H) (34), also depends on the Ru/acid ratio. Stoichiometric reactions of 34 and methanesulfonic acid (MsOH) form a mixture of species depending on equivalents of acid added (Fig. 4), suggesting acid influences catalyst activity by controlling speciation of Ru intermediates. The acidity of the reaction medium influences both metal-mediated and acid-promoted steps in these reaction sequences.

3.2. Hydrogenation reactions enabled by Lewis acid cocatalysis

3.2.1. Ru-Catalysed hydrogenation of carboxylic acid derivatives. Beller has reported several mild hydrogenations of carboxylic acid derivatives via transition-metal and Lewis acid cocatalysis. The substitution of Bronsted acid for Lewis acid was first realized in the hydrogenation of carboxylic acids and esters: the conversion of lactone 35 to ether 36 proceeds using the Ru(acac)₃/triphos and Al(OTf)₃ cocatalysts at relatively moderate H₂ pressure (15–40 bar). A similar catalyst mixture converts carboxylic acid 37 to alcohol 38, although slightly higher pressures of H₂ are necessary (40–60 bar) (Scheme 10).

Under lower pressures of H₂, metal triflate cocatalysts greatly outperform Bronsted acids. The hydrogenation of 37 (Scheme 10) does not proceed in the absence of cocatalyst, or with Bronsted acids (MsOH, diphenylphosphate), whereas when Al(OTf)₃ is employed, >99% conversion of 35 is observed with 83% yield of the desired ether product 36. Although triflic acid (TfOH) also promotes conversion of 35, poor selectivity of the reaction is observed; a competing elimination from the intermediate acetal provides the enol ether 39 as the major product in 45% yield. Al(OTf)₃ provides much better performance than other Al compounds. Qualitative mechanistic experiments suggest that an off-cycle catalyst resting state is reactivated in the presence of relatively acidic protons, resembling the reactivity of Ru and Bronsted acid cocatalytic systems (section 3.1). The presence of H₂O is important for efficient catalysis; the reaction is suppressed when molecular sieves are added, or when non-polar solvents are used. The authors propose that Al(OTf)₃ is important to generate Bronsted acid in situ upon combination with H₂O or alcohols. Like the system reported by Klankermeyer and Leitner, the acid could act to generate the active cationic [RuH⁺] species. Additionally, the weakly-coordinating OTf counteranion of the M(OTf)₃ Lewis acid may play an important role in the reaction, as chloride anions inhibit catalysis.

The Ru(acac)₃/triphos/M(OTf)₃ system also mediates the hydrogenation of secondary amides to amines (Scheme 11) at
low pressures of H₂ (5–15 bar).⁴⁷ A screen of acidic additives reveals that in the hydrogenation of amide 40, Yb(OTf)₃·H₂O cocatalyst provides the highest yield of amine 41 and that again, Lewis acidic metal triflate salts outperform Brønsted acidic cocatalysts under these reaction conditions. The reaction proceeds by initial hydrogenolysis of the amide to benzyl alcohol 42 and aniline 43, which are observable intermediates. A subsequent condensation reaction between 42 and 43 yields the benzylamine 41. Control experiments show that both the individual Ru and Yb catalysts are necessary for the condensation reaction to occur, and that H₂, while not necessary for condensation, improves the efficiency of this step.

### 3.2.2. Co-Catalysed indole alkylation via carboxylic acid hydrogenation

A related example employs a Co(acac)₃/triphos-catalysed tandem carboxylic acid hydrogenation and indole alkylation or alkenylation (Scheme 12).⁴⁸ The reaction only proceeds in the presence of Brønsted or Lewis acid additives. Once again, metal triflate additives provide the best reaction performance, over other Lewis acids (BF₃·Et₂O, B(C₆F₅)₃) or Brønsted acids (HOTf, HNTf₂, MsOH, TfOH). The reaction of indole 44 with acetic acid (45) yields the alkynilidine 46 in 70% using Al(OTf)₃ as cocatalyst, and in 42% yield using HNTf₂, whereas no conversion is observed in the absence of an acidic cocatalyst. Under similar conditions, reaction of indole 47 with diphenylacetic acid (48) yields alkenylindole (49) in 80% yield.

A plausible reaction sequence (Fig. 5) starts with Co/Al (OTf)₃-catalysed hydrogenation of acetic acid (45) to the respective acetaldehyde (50), followed by nucleophilic attack of the indole 44 to 51. Acid-catalysed dehydration of the resulting functionalized indole intermediate 51 generates alkenylated indole 52; subsequent Co and Al(OTf)₃-catalysed hydrogenation of the alkene yields the desired alkenylindole product 46. Evidence for an alkenylindole intermediate comes from the reaction of α,α-disubstituted carboxylic acid (48). Here the major product is alkenylindole 49, probably because the trisubstituted alkene is unreactive towards successive hydrogenation.

#### 3.2.3. Pd and Lewis acid-cocatalysed ether hydrogenolysis

Other examples of tandem transition metal and Lewis acid-catalysed hydrogenolysis reactions reveal a distinct role of Lewis acids in promoting C–O bond cleavage.⁴⁹–⁵² Hydrogenolysis of ether 53 to 54 is achieved with Yb(OTf)₃ and aluminasupported Pd nanoparticles in ionic liquid solvent ([EMIM][OTf], EMIM = 1-ethyl-3-methylimidazolium) (Scheme 13).⁵⁹ The C–O bond cleavage event of 53 to allylphenol 55 is the microscopic reverse of alkene hydroalkoxylation, a reaction catalysed by lanthanide Lewis acids. Indeed, the same catalysts can perform the C–O bond cleavage as well; however, the C–O bond cleavage of 53 is endothermic (ΔH ≈ +14 kcal mol⁻¹). When coupled with Pd-catalysed hydrogenation (ΔH ≈ −25 kcal mol⁻¹) the overall tandem C–O hydrogenolysis becomes exothermic. In this case, the reaction is not accelerated by HOTf, and so the authors propose that Yb(OTf)₃ activates substrate towards C–O bond cleavage, rather than simply generating Brønsted acid in situ. The lanthanide metals with smaller ionic radius and greater electrophilicity (i.e., effective charge density) are more effective at promoting the reaction. An experimental screen of lanthanide triflates shows yields correlate with known activation parameters for metal-triflate-catalysed conversion of ethers to enols.⁴⁹
In the absence of H₂ pressure, the conversion of 1,8-mol sequences catalysed by a variety of di such as carbocation rearrangement. intermediate alcohols, potentially mitigating side reactions, propose that the Lewis acid may bind and increase acidity of in agreement with experimental observation. The authors in situ

<table>
<thead>
<tr>
<th>Acid</th>
<th>Time (min)</th>
<th>Conv. (%)</th>
<th>ρ (eff. charge density)</th>
</tr>
</thead>
<tbody>
<tr>
<td>—</td>
<td>90</td>
<td>—</td>
<td>—</td>
</tr>
<tr>
<td>La(OTf)₃</td>
<td>30</td>
<td>0.2</td>
<td>2.60</td>
</tr>
<tr>
<td>Yb(OTf)₃</td>
<td>30</td>
<td>0.2</td>
<td>2.81</td>
</tr>
<tr>
<td>Ce(OTf)₃</td>
<td>10</td>
<td>5.7</td>
<td>3.44</td>
</tr>
<tr>
<td>Sc(OTf)₃</td>
<td>10</td>
<td>12.4</td>
<td>3.23</td>
</tr>
<tr>
<td>Fe(OTf)₃</td>
<td>10</td>
<td>13.6</td>
<td>3.71</td>
</tr>
<tr>
<td>Al(OTf)₃</td>
<td>10</td>
<td>23.7</td>
<td>3.87</td>
</tr>
<tr>
<td>Zr(OTf)₄</td>
<td>10</td>
<td>33.5</td>
<td>4.29</td>
</tr>
<tr>
<td>Hf(OTf)₄</td>
<td>10</td>
<td>46.0</td>
<td>4.37</td>
</tr>
<tr>
<td>HfOTf</td>
<td>10</td>
<td>10.4</td>
<td>—</td>
</tr>
</tbody>
</table>

Table 2 Correlation of ether hydrogenolysis activity and metal triflate effective charge density

The competency for metal triflates to mediate the key acid-catalyzed C–O bond cleavage step indicates the transition metal Lewis acids are particularly suitable for ether cleavage (Table 2). In the absence of H₂ pressure, the conversion of 1,8-cineole ([56] is greatest when reacted with 0.5 mol% Zr(OTf)₄ or Hf(OTf)₄ and Pd/C (which has been pre-treated with H₂).[51] Computationally-determined (B3LYP) effective charge density (ρ) of metal triflate Lewis acids tracks with catalytic performance, as determined by conversion and turnover numbers. DFT exploration of the reaction’s potential energy surface indicates that C–O bond cleavage of 2-methyltetrahydropyran via Yb(OTf)₃ has a higher activation barrier (ΔG° = +32.4 kcal mol⁻¹), than when Hf(OTf)₄ is used (ΔG° = +16.8 kcal mol⁻¹) – in agreement with experimental observation. The authors propose that the Lewis acid may bind and increase acidity of intermediate alcohols, potentially mitigating side reactions, such as carbocation rearangement.

Metal triflate Lewis acids are competent cocatalysts to mediate challenging, multi-step hydrogenation reaction sequences catalysed by a variety of different transition metals. In some cases, it is evident that a Lewis acid can promote the formation of a Bronsted acid in situ, which may be important to catalyse dehydration steps or other acid-catalysed reactions. It is also evident that Bronsted acids activate catalytic intermediates. As is demonstrated by Pd and metal triflate mediated ether hydrogenolysis, the coupling of two catalyst systems can “leverage” demanding endothermic reactions.

4. Tandem Au(Ⅰ) catalysis: Lewis acids as Brønsted acids, catalyst activators and reactivators

The ability of Lewis and Bronsted acids to form highly-active cationic Au(Ⅰ) catalysts was noted in the seminal 1998 report by Teles[53] on Au-mediated alkyne hydroalkoxylation, a reaction of immense subsequent interest.[54] Au halide catalyst precursors are typically combined with Ag salts, which are thought to generate catalytically-active cationic Au species. Although the non-innocent role of Ag in such reactions is recognized,[55] less attention has been afforded to the role of Bronsted or Lewis acids in Au-catalysed reactions.[56] In similar fashion to the analogy of Ru/Bronsted acid and Ru/Lewis acid cocatalytic systems (sections 3.1 and 3.2), it is useful to consider the potential role of Lewis acids to effect the generation of Bronsted acid in situ within the context of Au catalysis (section 4.2). We note that although several Au/Bronsted acid catalysed reactions are known,[57–61] we present evidence that the action of Lewis acids in some cocatalytic systems are distinct from that of Bronsted acids. Here we summarize recent efforts to understand the mechanism of such cooperative catalytic systems.

4.1. Au and Lewis acid tandem catalysis

First, examples are discussed in which Au and the Lewis acid act synergistically in tandem catalysis. To the best of our knowledge, the first evidence of metal triflate Lewis acids to serve as a catalyst activator in Au-catalysed reactions is the report of Shi,[62] in which (PMe₃)AuCl and Yb(OTf)₃ catalyse the rearrangement of epoxy alkynes to carboycles (Scheme 14). In addition to Yb(OTf)₃, acting as an apparent halide abstractor in this reaction, the catalyst is more active when Yb(OTf)₃ is employed rather than when AgOTf is employed. The first step of the reaction involves the ring-opening of epoxy alkylne 57 to the enol alkylne 58, which is catalysed independently by Yb (OTf)₃ without activation of the alkynne, and the Au-catalysed rearrangement to the carbocycle 59 occurs in much greater yield with Yb(OTf)₃. The combination of (PMe₃)AuCl and Yb(OTf)₃ generates (PMe₃)Au(OTf) in situ – presumably the active form of the catalyst. The authors claim that the mixture of (PMe₃)AuCl and Yb(OTf)₃ provides improved catalytic activity than when (PMe₃)AuCl and AgOTf are used, as well as when pre-formed (PMe₃)Au(OTf) is used under the standard reaction conditions.

Several other synthetic applications have more recently emerged, particularly in cascade reactions where Lewis acids promote steps not catalysed by Au. The groups of Jia and Xu report several synergistic Au and Lewis acid reactions in the synthesis of spirocyclic[63–65] and fused bicyclic compounds.[66,67] For example, bicyclic aminal 60 is synthesized via a Au-cata-

Scheme 14 Tandem epoxide ring-opening and cyclisation.[62]
lysed intramolecular alkyne hydroamination followed by an Ga(OTf)3-catalysed inverse-electron-demand hetero-Diels–Alder cyclisation (Scheme 15).66 Analogous Au and Lewis acid tandem reactions are reported to selectively favour spiroaminals63 and spiro-heterocycles64 over bicyclic cyclisation patterns. Spiroketal65 and fused tricyclic compounds67 are also accessible from similar alkyne hydrofunctionalisation/cycloaddition sequences, where the Lewis acid also acts to generate ortho-quinonemethides from 2-(hydromethyl)phenol derivatives. Au and Lewis acid-catalysed alkyne hydroamination has also been applied to the synthesis of N-heterocycles such as a pyrroles68 and indoles.69

Lewis acids can also affect Au-mediated steps of the catalytic reaction. In a Au/Ga-cocatalysed intermolecular Nakamura reaction, nucleophilic attack of 1,3-diketone 61 upon phenylacetylene (62) yields the alkénylated 1,3-diketone 63 (Scheme 16).70 The authors find that (XPhos)Au(TA)(OTf) (TA = 1H-benzo[1,2,3]triazole) is only a competent precatalyst in the presence of Ga(OTf)3; other Lewis acids, such as Zn(OTf)2, or other cocatalysts such as AgOTf or HOTf, do not cocatalyse the reaction. The authors propose that Ga(OTf)3 may bind, and therefore, activate the 1,3-dicarbonyl substrate, but may also be critical for the formation of catalytic Brønsted acid.

4.2. Distinguishing between Ag, Brønsted and Lewis acid effects in Au catalysis

Hammond and Xu have comprehensively investigated Brønsted and Lewis acid effects in Au-catalysed reactions.71–73 In one study, the effect of cocatalytic additives on Au-catalysed alkyne and allene hydrofunctionalisation reactions was explored.71 Au catalysis at extremely low catalyst loadings (0.02–0.5 mol%) are strongly influenced by impurities such as halides anions, or basic anions (OH− or CO32−). The authors explore a series of reactions catalysed under base- and halide-free conditions using precatalysts (JohnPhos)AuOTf (JohnPhos = (2-biphenyl)d-tert-butyolphosphate) and (PPh3)AuOTf. In a low-catalyst loading hydration of alkyne 64 (Fig. 6), no catalytic activity is observed in the reaction without co-catalyst; in the reaction with cocatalytic (0.6 mol%) Ga(OTf)3 or HOTf, the reaction proceeds to high conversion in 12 hours. Ga(OTf)3 may act as a source of Brønsted acid (HOTf) in this case. But in other reactions, such as the cycloisomerisation of enyne 65 (Fig. 7), cocatalytic HOTf is ineffective, whereas cycloisomerisation proceeds to quantitative yield in 20 min in the presence of In(OTf)3. On the other hand, in the hydrocarboxylation of hexynoic acid, the reaction is promoted by HOTf, AgOTf, and In(OTf)3 at an approximately similar rate; here, Ga(OTf)3 is ineffective.

These examples demonstrate the complexity of catalyst-Lewis acid interactions and their dramatic effect on catalyst
An acidic cocatalyst may regenerate catalytically-active cationic Au deactivated by base or halide impurities. A Bronsted acid-mediated mechanism for catalyst reactivation may proceed by protonolysis of a basic Au species, whereas Lewis acid-mediated catalyst reactivation may involve either anion abstraction or in situ generation of HOTf or other Bronsted acids. The authors demonstrate that basic impurities can come from filtration agents or desiccants, such as Celite or 4 Å molecular sieves.

Acid acceleration in Au reactions presumably influence the rate-limiting step in catalysis, but the rate-limiting step in Au catalysis is not always clear.\textsuperscript{71} In certain cases protodeauration is believed to be rate-limiting,\textsuperscript{74,75} strong Au–C bonds are generally robust to protonolysis and likely present in on-cycle and off-cycle pathways. Hammond\textsuperscript{75} and Yu\textsuperscript{76} find that protonolysis of heteroarylgold complexes is first order in gold complex and acid, and that the rate depends on the acid p$K_a$. In protonolysis of heteroarylgold complex \textit{66} the rate is evidently dependent on the acid strength: TfOH yields 100% conversion of \textit{66} in 5 min of reaction, whereas the weaker acid TsOH requires a longer duration (80%, 1 h); HOAc produces no reaction (Scheme 17).\textsuperscript{75} Accelerating turnover-limiting protodeauration is therefore a plausible role for Lewis acids, in addition to decreasing the rate of catalyst deactivation.

As noted previously, Lewis acids can also affect the activation of Au precatalysts by abstracting a coordinating anion. The use of Lewis acids as activators of Au halide precursors has been systematically studied by Gandon and coworkers,\textsuperscript{77-79} who sought alternatives to Ag activation. Ag is thought to form inactive bimetallic Au–Ag\textsuperscript{80} or dimeric Au species,\textsuperscript{81} and increase rates of catalyst decomposition; fast, irreversible anion abstraction can actually slow catalysis via favouring formation of Au(0) or inactive gold species (such as [L$_2$Au][X]).\textsuperscript{77} Decomposition is of special concern in reactions of reducing substrates (alkynes or alkenes) or in reactions requiring elevated temperatures.\textsuperscript{75}

Alternative halide abstractors, such as Cu, Zn, In, and Bi Lewis acids, are effective in Au-catalysed hydroalkylation reactions, enyne cyclisations, and other Au catalysis.\textsuperscript{77,78} Unusually low Au loadings are effective in C–C bond forming reactions with Cu cocatalysts. The Au/Cu cocatalyst system also offers the advantage of scalability; for example, the ene-β-ketoamide \textit{67} is transformed into the cyclised product \textit{68} using (L)AuCl and Cu(OTf)$_2$ (80% yield, dr 95:5, L = JohnPhos or PPh$_3$) and notably, the reaction only proceeds to <30% conversion when AgOTf is used instead of Cu(OTf)$_2$. Following the speciation of the (PPh$_3$)$_2$Au(Cl) precatalyst by $^{31}$P NMR reveals slow generation of [(PPh$_3$)$_2$Au][OTf] upon combination with Cu(OTf)$_2$, presumably via unobserved (PPh$_3$)Au(OTf). During catalytic conditions, Cu(OTf)$_2$ may slowly and reversibly abstract the chloride anion from the Au precatalyst, minimising catalyst decomposition by maintaining low concentration of [LAu][X]. In a separate report, a screen of Lewis acids at similar reaction conditions (Table 3) indicates that In(OTf)$_3$ and Bi(OTf)$_3$ may slowly and reversibly abstract the chloride anion from the Au precatalyst, minimising catalyst decomposition by maintaining low concentration of [LAu][X].

A Lewis acid can take on numerous possible roles in Au catalysis. The most obvious role is anion abstractor, and sometimes Lewis acids provide improved activity vs. the more common Ag activators. The Lewis acid can also generate Bronsted acid in situ – thereby affecting Au-centred activation/deactivation pathways, accelerating the turnover-limiting protonolysis of key Au–C intermediates, or initiating acid-catalysed reaction sequences. We note that control experiments are critical when evaluating acid effects in Au catalysis. Several Lewis acids, such as Ga or In compounds, are also π-acids and can sometimes independently catalyse alkyl functionalisation reactions.\textsuperscript{82} Furthermore alkyl functionalisation reactions can also be independently catalysed by Bronsted acids.\textsuperscript{83}

### 5. Enabling transmetalation and reductive elimination with Lewis acids in Pd catalysis

In this section, we will present several examples of Lewis acid additives and their effects in C–C and C–N coupling reactions. Pd coupling reactions are almost always run in the presence of...
alkali bases, and boron, zinc, magnesium, and copper nucleophiles are common alkyl and aryl group transfer agents. The Lewis acidity of these reagents and their byproducts are seldom considered when exploring reaction mechanisms.

The influence of reagent metal cations is evident in Liebeskind’s coupling of aromatic thioethers with organometallic nucleophiles.84 The coupling reaction proceeds in the presence of strongly nucleophilic alkylmagnesium or alkylzinc reagents. For boronic acids, a separate Lewis acid additive must be used. For example, addition of Zn(OAc)₂ enables the coupling heteroaromatic thioether 69 and boronic acid 70 using the Pd₂(dba)₃/P(fur)₃/CuTC catalyst yields biaryl 71 (76% yield) (Scheme 18). The thiophilic Zn Lewis acid is thought to bind the thiolate intermediate formed upon oxidative addition of the C-S bond to Pd(0), facilitating the difficult transmetalation step.

This work inspired others to consider potential roles of Lewis acids in C-N bond-forming reactions. Hartwig and Shen isolated arylpalladium(II) amido complexes – proposed intermediates in amine N-arylation – and explored their reactivity in the presence of Lewis acidic species (Scheme 19).85 Upon heating a solution of complex 72 or 73 in PhMe in the absence of Lewis acid, the predominant reaction pathway is protonolysis of the amido ligand, liberating the free amine or sulphonamide. With added BPh₃ or BEt₃, on-cycle reductive elimination of the N-heteroarylamine occurs in 90% yield. The heteroaryl ligand of 72 acts as a Lewis acid “docking site,” making the Pd centre more electron deficient and accelerating reductive elimination versus competing protonolysis. BEt₃ and complex 73, which lacks a heteroatom for Lewis acid binding, instead forms products of ethyl group transfer to the arene as well as amido protonolysis. The stoichiometric Pd-Lewis acid interaction may be catalytically relevant: coupling of heteroarene 74 with amide 75 using the Pd(dba)₃/xantphos yields N-arylamide 76 in 84% yield in the presence of BEt₃, but only in 17% yield without Lewis acid (Scheme 20). The control over the metal electronic character via remote binding of a Lewis acid resembles effects reported by Nolan and Molloy,86 and Bergman and Tilley.87

Further examples demonstrate that Lewis acid cocatalysts can enable challenging C-N coupling reactions. Researchers at Bristol-Myers Squibb serendipitously discovered that Zn salts enable the Pd-catalysed formation of a N-heteroarylazaindole 77.88 The azaindole precursor 78 is prepared from the Pd-catalysed cyanation of bromoarene 79 with Zn(CN)₂, and the subsequent N-arylation step exhibited different performance depending on whether 78 was isolated by recrystallisation (38% yield) or chromatography (11% yield) (Scheme 21). The varying reaction performance was attributed to the presence of residual Zn present in recrystallized 78. Indeed, the reaction of heteroaryl

---

**Scheme 18** Pd-Catalyzed coupling of arylthioethers and boronic acids.84

**Scheme 19** Reactivity of amidopalladium complexes.85

**Scheme 20** Cross coupling of heteroaryl halides and lactams.85

**Scheme 21** Route to an N-heteroarylazaindole by bromoarene cyanation and azaindole N-arylation.88
chloride 80 with 78 catalysed by Pd$_2$(dba)$_3$/xantphos yields 78 in only 3% yield. The reaction instead returns >98% yield when Zn(OAc)$_2$ or Zn(OPiv)$_2$ is employed as a cocatalyst (60 mol%).

Although the mechanistic proposal by Hartwig and Shen invokes the binding of Lewis acid to a heteroaryl-palladium moiety, this catalyst-Lewis acid interaction is not a requirement for a Lewis acid effect to be observed in C–N coupling reactions. Our group has observed that metal triflate Lewis acids are suitable cocatalysts in Pd catalysis, improving the yield of various N-arylamide products.\(^{89}\) In a representative example (Scheme 22), the reaction of bromoarene 81 with N-methylacetamide (82) via Pd(dba)$_3$/xantphos/Al(OTf)$_3$ yields the tertiary amide 83 in 89% yield, whereas only 53% yield of 83 is observed in the absence of an Al salt. Based on mechanistic insight and kinetic data, the Lewis acid appears to accelerate the rate-limiting transmetalation step of catalysis. However, deconvolution of key steps of transmetalation is difficult. At the very least, this step includes various reversible pre-equilibria as represented in Fig. 8: (a) halide dissociation from the oxidative addition complex I, (b) amide coordination to II, (c) dissolution of heterogenous base (N-arylation reactions in the presence of insoluble, inorganic bases can be mass-transfer limited),\(^{90}\) and (d) amide deprotonation to yield intermediate III.

Measurement of the initial rates of catalysis of the coupling of bromobenzene (84) and pyrrolidin-2-one (78) indicate that an approximately three-fold acceleration is observed in the presence of Yb(OTf)$_3$ (1 : 1 Yb : Pd), and that inhibition of catalysis is observed in the presence of exogenous bromide salts (NBu$_4$Br) (Fig. 9). Our group proposed that the Lewis acid acts to mitigate competitive binding of bromide (a reaction byproduct) and the weakly nucleophilic amide by catalytic halide abstraction from the key Pd intermediate I. Indeed, by $^{31}$P NMR, evidence for the generation of the cationic Pd intermediate [(xantphos)Pd(Ph)]$^+$ from the interaction of the oxidative addition complex, (xantphos)Pd(Ph)(Br), with Yb (OTf)$_3$, suggesting that Lewis acids can promote the dissociative pathway in transmetalation represented in Fig. 8.

Another example in which a dramatic Lewis acid effect is observed comes from the group of Organ,\(^4\) where a sterically hindered NHC ligand enables the coupling of aryl halides and primary amides when borane cocatalysts are employed. Unactivated aryl chloride 85 and benzamide 86 react with a PEPPSI-Pd catalyst to yield the secondary amide 87 in quantitative yield with various borane additives (BEt$_3$, tri(sec-butyl) borane, B(C$_6$F$_5$)$_3$), while the product is only obtained in 21% yield in the absence of a borane (Scheme 23). The authors observe different behaviour between the three borane reagents. Whereas alkylboranes undergo autoxidation readily, B(C$_6$F$_5$)$_3$ has stronger B–C bonds and therefore does not undergo immediate reaction with O$_2$. It is observed that atmospheric contaminants and/or added radical traps affect the catalysis differently depending on which borane is used; for example, reactions that employ BEt$_3$ do not yield product when the borane is first exposed to air. The authors provide spectroscopic evidence for the formation of borane-amide adducts (Scheme 24), which may be responsible for increasing the nucleophilicity of the amide and thereby facilitating transmetalation.

In these Pd-catalysed reactions, various evidence points to the Lewis acid modulating the reactivity of the key Pd oxidative

---

**Scheme 22**  
**N-Arylation of a secondary acyclic amide.**\(^{89}\) Reprinted with permission. Copyright 2017 American Chemical Society.

**Scheme 23**  
**N-Arylation of a primary amide.**\(^4\)

**Fig. 8** Proposed mechanism of dehydrohalogenation and amide coordination in Pd-catalysed amide N-arylation.\(^{89}\) Reprinted with permission. Copyright 2017 American Chemical Society.

**Fig. 9** Effect of additives on the initial rate of amide N-arylation.\(^{89}\) Reprinted with permission. Copyright 2017 American Chemical Society.
addition intermediates by catalyst-Lewis acid interactions, though Lewis acids of different character are employed (alkylborane, arylborane, zinc salts, or metal triflates, such as Al(OTf)_3 or Yb(OTf)_3) and therefore different possible modes of action of the Lewis acid may be operative. Based on the stoichiometric study by Shen and Hartwig, alkylboranes may increase the rate of reductive elimination versus the rate of amido protonolysis. Based on the works of our group and Organ, where delivery of the nitrogen nucleophile to Pd appears to be turnover-limiting, binding of amide may require sequestration of halide, or the generation of a more nucleophilic boron-amidate complex.

6. Conclusions

Lewis acid cocatalysts influence a diverse array of transition metal-catalysed reactions. Through mechanistic insight from a variety of studies, it appears that the predominant roles of the Lewis acid in the described transformations are as follows: (1) to form a Lewis acid base-adduct of the substrate to enable a challenging bond-cleavage event or to provide a thermodynamic driving force for a bond-cleavage event, (2) to modulate the pK_a of the reaction medium, effect the controlled release of a Brønsted acid in situ, or facilitate proton transfer by some means, (3) to affect the coordination sphere of catalyst precursors or key catalyst intermediates by anion abstraction, and in doing so promote catalyst activation, reactivate off-cycle catalyst species, or accelerate slow organometallic elementary steps. In some cases, parametrisation or spectroscopic study has enabled insight into the behaviour of Lewis acidic species. Further mechanistic insight and novel experimental means of Lewis acid parameterisation are crucial for further application and rational design of cocatalytic systems.

Lewis acid additives can provide more active or selective catalysts for existing processes, or enable orthogonal reactivity of catalytic reactions by the Lewis acid-mediated interception of key intermediates. Thus, further development in this area will impact new and existing chemistries alike, and we expect Lewis acids will figure prominently in many further catalytic applications. Additional work in understanding and controlling Lewis acid/organometallic interactions will be instrumental in offering new opportunities in homogeneous catalysis.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This material is based upon work supported by the National Science Foundation under Grant No. CHE-1565721.

References