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Investigating the effect of ligand and cation on the properties of metal fluorinated acetylacetonate based magnetic ionic liquids†

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Magnetic ionic liquids (MILs) are a subclass of ionic liquids that possess a paramagnetic metal within their chemical structure, making them susceptible to external magnetic fields. A total of twenty-four (24) MILs were prepared and characterized to investigate the effect of the ligand, cation and anion on the physiochemical properties of acetylacetonate-based MILs. It was found that thermal stabilities as high as 260 °C could be achieved by incorporating aromatic moieties in the anion structure. Additionally, the magnetic moment could be modulated by simply changing the transition metal in the anion. Magnetic moment values of 2.8 μ_B , 4.5 μ_B and 5.6 μ_B were obtained by using Ni(II), Co(II), and Mn(II) as the metal centers, respectively. Furthermore, the viscosity of the MILs could be tailored from a few hundred centipoise to several thousand centipoise, increasing their potential applications in numerous interdisciplinary fields. Moreover, the MILs synthesized in this study were found to be insoluble in water at a MIL-to-solvent ratio of 0.01% (w/v), making them potentially useful in targeted separations, where very hydrophobic solvents are highly desired.

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1. Introduction

Ionic liquids (ILs) are an intriguing class of non-molecular solvents comprised of organic cations and organic/inorganic anions.^{1,2} These molten organic salts possess low melting points (≤ 100 °C) and an array of physiochemical properties including variable viscosity, low vapor pressure at ambient temperature, and high thermal stability. Moreover, the properties of ILs can be modified through careful selection of the cation/anion combination.³ This feature of structural tunability makes ILs attractive materials for a wide range of applications in various fields of science and engineering. A subclass of ILs that contain a paramagnetic component in the cation or anion structure are magnetic ionic liquids (MILs).^{4–8} These compounds possess many similar properties to traditional ILs, with the additional characteristic of responding to an external magnetic field. MILs have been successfully employed for density measurements *via* magnetic levitation,⁹ catalysis,^{10,11} magnetic fluids,¹² cellulose dissolution,¹³ solvents in various organic reactions,^{14,15} electrochromic applications,¹⁶ and extraction of hazardous chemicals.¹⁷

The physicochemical properties of MILs can be tailored for specific applications through careful modification of the cation

and anion structures. Among their interesting properties, magnetic susceptibility of MILs is often exploited in liquid–liquid extractions^{17–19} and magnet-based sensors.²⁰ The earliest reported MILs utilized simple halide-based, transition metal anions with the 1-butyl-3-methylimidazolium tetrachloroferrate(III) ([BMIM]⁺[FeCl₄][−]) MIL exhibiting an effective magnetic moment (p_{eff}) of 5.8 Bohr magnetons (μ_B).^{5,12} In order to significantly increase their magnetic susceptibility, trivalent lanthanide paramagnetic metals have been explored. Incorporation of Gd(III), Dy(III), and Ho(III) into the MIL chemical structure acts to increase their magnetic susceptibility due to higher paramagnetic moments compared to 3d transition metals.⁹ Another strategy used to enhance magnetic susceptibility includes the incorporation of di- and tri-cationic structures containing multiple paramagnetic anions.²¹ Magnetic susceptibilities as high as 11.6 μ_B have been reported using a tricationic MIL framework.²²

MILs have also drawn attention due to their low volatility, making them environmentally-friendly and suitable for high temperature applications such as solvents for organic synthesis,^{14,15} and gas–liquid chromatography.^{23,24} It has been observed that both the cation and anion components have a significant contribution towards the thermal stability of MILs. For example, Nacham *et al.* found that ammonium-based MILs had significantly lower thermal stability than dicationic imidazolium-based MILs.²² Another study investigated the effect of different cation and anion combinations on the thermal stability of MILs. By carefully manipulating the cation and anion and their combination,

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a tris(1-butyl-2,3-dimethylimidazolium)dysprosiumhexachlorate ($[\text{BDMIM}^+]_3[\text{DyCl}_6^{3-}]$) MIL with relatively higher thermal stability (334 °C) than previously reported MILs was prepared.⁹ Several studies have also demonstrated the effects of different MIL components, such as the composition of the cation and anion structures,^{25,26} metal center,²⁷ and the alkyl chain length of imidazolium cations on the resulting viscosity.²⁸ Lower viscosity MILs tend to be easier to handle and manipulate within fluid handling devices. Higher viscosity MILs have advantages in applications that require highly mechanically stable liquids.²⁹ Yoshida *et al.* found that longer alkyl chain substituents in imidazolium-based MILs resulted in higher viscosity.²⁸ Other approaches for modulating the viscosity include mixing two components of different viscosity to generate mixtures that contain lower viscosity than the individual components themselves.²¹

Recently, our group reported the preparation of a series of MILs consisting of hexafluoroacetylacetone (hfacac) ligands coordinated to different metal centers and paired with a phosphonium cation.³⁰ These MILs exhibited very low viscosity, largely due to the bulky, asymmetric phosphonium cation, as well as very high hydrophobicity.^{31,32} However, it was observed that the thermal stability of these MILs appeared to be limited by the anion complex. Thermal stabilities as low as 130 °C were observed for the trihexyl(tetradecyl)phosphonium tris(hexafluoroacetylaceto)cobaltate(II) ($[\text{P}_{66614}^+][\text{Co}(\text{hfacac})_3^-]$) MIL.³⁰ Therefore, in order to increase the broad applicability of these solvents, a systematic study that relates their chemical structure to observed physicochemical properties is required.

In this study, a total of twenty-four (24) MILs containing acetylacetone-based ligands were synthesized and characterized in an effort to better understand the relationship between ligand structure, cation type, and metal center on the resulting physicochemical properties. The effects of MIL cation/anion combination on thermal stability, hydrophobicity, magnetic susceptibility, and viscosity are thoroughly investigated. A series of different acetylacetone-based ligands with various transition metals centers (*i.e.*, Ni(II), Co(II), and Mn(II)) were synthesized and paired with trihexyl(tetradecyl)phosphonium ($[\text{P}_{66614}^+]$) and 1-tetradecyl-3-methylimidazolium ($[\text{C}_{14}\text{MIM}^+]$) cations. Through incorporation of aromatic moieties in the acetylacetone-based ligands, the thermal stability of the resulting MILs was increased up to 260 °C, representing a significantly higher thermal stability than previously reported hfacac-based transition metal MILs (130 °C, 155 °C, and 215 °C for the Co(II), Mn(II), and Ni(II) transition metals, respectively).³⁰ Moreover, it was found that the viscosity could be precisely tailored to range as low as 100 centipoise (cP) up to several thousand cP. This study provides a fundamental understanding into the relationship between structure and physiochemical properties of acetylacetonate-based MILs.

2. Experimental

2.1. Materials

Ammonium hydroxide (28–30% solution in water), 1,1,1,5,5,5-hexafluoroacetylacetone (99%) and nickel(II) chloride (98%)

were purchased from Acros Organics (Morris Plains, NJ, USA). Manganese(II) chloride tetrahydrate (98–100%), 4,4,4-trifluoro-1-(2-naphthyl)-1,3-butanedione ($\text{N}_\text{p}\text{tfacac}$, 98%), 4,4,4-trifluoro-1-phenyl-1,3-butanedione ($\text{P}_\text{h}\text{tfacac}$, 98%) and 1,1,1-trifluoro-2,4-pentanedione (tfacac, 98%) were purchased from Alfa Aesar (Ward Hill, MA, USA). 4,4,4-Trifluoro-1-(2-thienyl)-1,3-butanedione ($\text{T}_\text{h}\text{tfacac}$, 98%) was purchased from Tokyo Chemical Industry (Portland, OR, USA). Cobalt(II) chloride (97%), hexane (98.5%), methanol (99.9%), pyridine (>99%), 1-methylimidazole (99%) and acetonitrile (99.9%) were purchased from MilliporeSigma (St. Louis, MO, USA). 1-Tetradecanol (97%) and benzenesulfonyl chloride (98%) were purchased from Beantown Chemical (Hudson, NH, USA). Deuterated dimethyl sulfoxide (d_6 -DMSO) was obtained from Cambridge Isotope Laboratories (Andover, MA, USA). $[\text{P}_{66614}^+][\text{Cl}^-]$ (>93%) was purchased from Strem Chemical (Newburyport, MA, USA). The $[\text{P}_{66614}^+][\text{Cl}^-]$ IL was purified by dissolving 15 g of the IL in 50 mL of acetonitrile and washing 5 times with 10 mL of hexane. Acetonitrile was then removed by roto-evaporation and the purified IL was dried overnight under vacuum at room temperature. Manganese(II) chloride tetrahydrate was also dried under vacuum for 3 days at room temperature to remove water before use. All other chemicals were used as received.

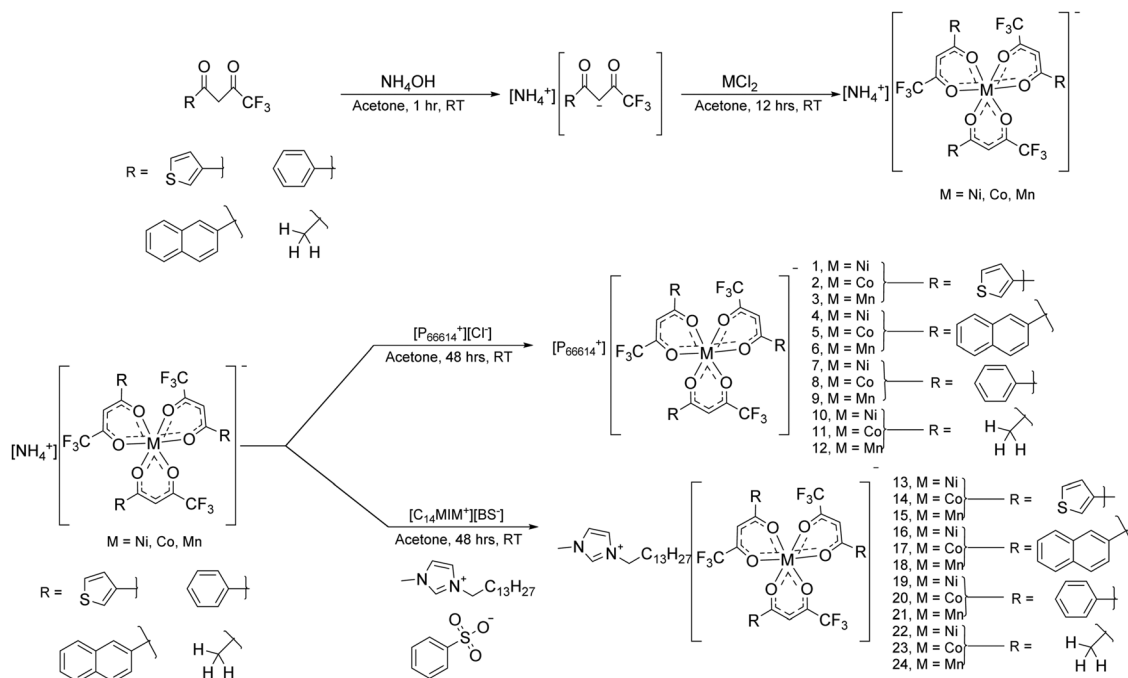
2.2. Instrumentation

Nuclear magnetic resonance (NMR) spectra (^1H and ^{13}C) were acquired *via* a Bruker 500 MHz spectrometer. Elemental analysis was obtained using a PerkinElmer 2100 Series II CHN/S Analyzer (Waltham, MA, USA). Thermogravimetric analysis (TGA) was performed using a Netzsch STA449 F1 thermogravimetric analyzer (Burlington, MA, USA). All samples were loaded in platinum pans and heated at a rate of 5 °C min⁻¹ under an argon flow of 20 mL min⁻¹. A Brookfield DV1 cone and plate viscometer with a CPA-51Z cone spindle at 65 °C was used to measure the viscosity of the MILs. Melting points were recorded by using a Mel-Temp II melting point apparatus (Dubuque, IA, USA). Magnetic measurements were performed using Quantum Design Superconducting Quantum Interference Device (SQUID) magnetometer, model MPMS XL-7. Magnetization as a function of temperature was measured between 300 and 2 K in 20 kOe applied magnetic field. Thermal stability measurements were obtained by using the MILs as stationary phases in gas chromatography (GC) to determine the onset decomposition temperature.³³ An Agilent Technologies 6850 GC (Santa Clara, CA, USA) with a flame ionization detector (FID) was used for thermal stability measurements. All MILs were coated on a 3 m untreated silica capillary with a 0.28 µm film thickness. A temperature program, starting from 40 °C and increased to 350 °C at a rate of 1 °C min⁻¹, was employed in all experiments.

2.3. MIL synthesis

MILs 1–24 were synthesized following a previously reported procedure,³⁰ with some modifications, as shown in Scheme 1. Firstly, 10 mmol of $\text{T}_\text{h}\text{tfacac}$, $\text{N}_\text{p}\text{tfacac}$, $\text{P}_\text{h}\text{tfacac}$, or tfacac were dissolved in acetone and a rubber septum used to seal the round bottom flask. An equimolar amount of ammonium hydroxide was then added dropwise to the reaction mixture and allowed to





Scheme 1 Synthesis of transition metal ionic liquids with varying acetylacetonate-based ligands and cations.

stir for 1 h. Then, 3.3 mmol of nickel(II) chloride, cobalt(II) chloride or manganese(II) chloride were added to the reaction mixture and allowed to stir overnight at room temperature. The solvent was then evaporated by roto-evaporation and the product dissolved in diethyl ether. Purification was performed by washing several times with 5 mL aliquots of water until the AgNO_3 test yielded no precipitate for MILs 4–12 and 16–24. MILs 1–3 and 13–15 were dissolved in acetone and subsequently filtered to remove ammonium chloride. The ammonium salts of all MILs were dried overnight in a vacuum oven after evaporation of acetone. Ion exchange was carried out by metathesis reaction with either $[\text{P}_{66614}^+][\text{Cl}^-]$ or $[\text{C}_{14}\text{MIM}^+][\text{BS}^-]$. The final product was dissolved in diethyl ether and washed with several 5 mL aliquots of water until no precipitate was observed *via* the AgNO_3 test. Diethyl ether was evaporated by roto-evaporation and the purified MILs were dried in a vacuum oven for 48 h at room temperature. All MILs were subsequently characterized by elemental analysis.

2.4. MIL characterization

The following data describes the physical appearance of the MILs at room temperature and provides the elemental analysis data for all synthesized MILs:

MIL 1. Dark green, highly viscous liquid. Elemental analysis calcd (%) for $\text{C}_{56}\text{H}_{80}\text{F}_9\text{NiO}_6\text{PS}_3$: C, 55.77; H, 6.69; found: C, 55.76; H, 6.65.

MIL 2. Dark red, highly viscous liquid. Elemental analysis calcd (%) for $\text{C}_{56}\text{H}_{80}\text{F}_9\text{CoO}_6\text{PS}_3$: C, 55.76; H, 6.68; found: C, 55.60; H, 6.51.

MIL 3. Dark yellow, highly viscous liquid. Elemental analysis calcd (%) for $\text{C}_{56}\text{H}_{80}\text{F}_9\text{MnO}_6\text{PS}_3$: C, 55.94; H, 6.71; found: C, 56.77; H, 6.75.

MIL 4. Dark green, highly viscous liquid. Elemental analysis calcd (%) for $\text{C}_{74}\text{H}_{92}\text{F}_9\text{NiO}_6\text{P}$: C, 66.42; H, 6.93; found: C, 66.20; H, 7.10.

MIL 5. Dark red, highly viscous liquid. Elemental analysis calcd (%) for $\text{C}_{74}\text{H}_{92}\text{F}_9\text{CoO}_6\text{P}$: C, 66.41; H, 6.93; found: C, 65.89; H, 7.31.

MIL 6. Dark yellow, highly viscous liquid. Elemental analysis calcd (%) for $\text{C}_{74}\text{H}_{92}\text{F}_9\text{MnO}_6\text{P}$: C, 66.61; H, 6.95; found: C, 66.12; H, 7.15.

MIL 7. Dark green, viscous liquid. Elemental analysis calcd (%) for $\text{C}_{62}\text{H}_{86}\text{F}_9\text{NiO}_6\text{P}$: C, 62.68; H, 7.30; found: C, 62.65; H, 7.13.

MIL 8. Dark red, viscous liquid. Elemental analysis calcd (%) for $\text{C}_{62}\text{H}_{86}\text{F}_9\text{CoO}_6\text{P}$: C, 62.67; H, 7.30; found: C, 63.31; H, 7.57.

MIL 9. Dark yellow, viscous liquid. Elemental analysis calcd (%) for $\text{C}_{62}\text{H}_{86}\text{F}_9\text{MnO}_6\text{P}$: C, 62.88; H, 7.32; found: C, 62.82; H, 7.31.

MIL 10. Dark green, viscous liquid. Elemental analysis calcd (%) for $\text{C}_{47}\text{H}_{80}\text{F}_9\text{NiO}_6\text{P}$: C, 56.35; H, 8.05; found: C, 57.19; H, 8.44.

MIL 11. Dark red, viscous liquid. Elemental analysis calcd (%) for $\text{C}_{47}\text{H}_{80}\text{F}_9\text{CoO}_6\text{P}$: C, 56.35; H, 8.05; found: C, 57.34; H, 7.96.

MIL 12. Dark yellow, viscous liquid. Elemental analysis calcd (%) for $\text{C}_{47}\text{H}_{80}\text{F}_9\text{MnO}_6\text{P}$: C, 56.56; H, 8.08; found: C, 56.92; H, 8.11.

MIL 13. Dark green, sticky solid. Elemental analysis calcd (%) for $\text{C}_{42}\text{H}_{47}\text{F}_9\text{N}_2\text{NiO}_6\text{S}_3$: C, 50.36; H, 4.73; N, 2.80; found: C, 50.23; H, 4.50; N, 2.96.

MIL 14. Dark red, sticky solid. Elemental analysis calcd (%) for $\text{C}_{42}\text{H}_{47}\text{F}_9\text{N}_2\text{CoO}_6\text{S}_3$: C, 50.35; H, 4.73; N, 2.80; found: C, 50.05; H, 4.54; N, 2.50.

MIL 15. Dark yellow, sticky solid. Elemental analysis calcd (%) for $\text{C}_{42}\text{H}_{47}\text{F}_9\text{N}_2\text{MnO}_6\text{S}_3$: C, 50.55; H, 4.75; N, 2.81; found: C, 50.25; H, 4.59; N, 2.65.

MIL 16. Dark green, solid. Elemental analysis calcd (%) for $\text{C}_{60}\text{H}_{59}\text{F}_9\text{N}_2\text{NiO}_6$: C, 63.56; H, 5.25; N, 2.47; found: C, 63.71; H, 5.00; N, 2.20.



MIL 17. Dark red, solid. Elemental analysis calcd (%) for $C_{60}H_{59}F_9N_2CoO_6$: C, 63.55; H, 5.24; N, 2.47; found: C, 63.15; H, 5.09, N, 2.40.

MIL 18. Dark yellow, solid. Elemental analysis calcd (%) for $C_{60}H_{59}F_9N_2MnO_6$: C, 63.77; H, 5.26; N, 2.48; found: C, 63.59; H, 5.08, N, 2.24.

MIL 19. Dark green, highly viscous liquid. Elemental analysis calcd (%) for $C_{48}H_{53}F_9N_2NiO_6$: C, 58.61; H, 5.43; N, 2.85; found: C, 58.67; H, 5.35; N, 2.72.

MIL 20. Dark red, sticky solid. Elemental analysis calcd (%) for $C_{48}H_{53}F_9N_2CoO_6$: C, 58.60; H, 5.43; N, 2.85; found: C, 58.72; H, 5.44; N, 2.77.

MIL 21. Dark yellow, sticky solid. Elemental analysis calcd (%) for $C_{48}H_{53}F_9N_2MnO_6$: C, 58.84; H, 5.45; N, 2.85; found: C, 59.06; H, 5.76; N, 3.13.

MIL 22. Dark green, viscous liquid. Elemental analysis calcd (%) for $C_{33}H_{47}F_9N_2NiO_6$: C, 49.71; H, 5.94; N, 3.51; found: C, 48.97; H, 5.71; N, 3.72.

MIL 23. Dark red, viscous liquid. Elemental analysis calcd (%) for $C_{33}H_{47}F_9N_2CoO_6$: C, 49.69; H, 5.94; N, 3.51; found: C, 49.08; H, 5.63; N, 3.70.

MIL 24. Dark yellow, viscous liquid. Elemental analysis calcd (%) for $C_{33}H_{47}F_9N_2MnO_6$: C, 49.94; H, 5.97; N, 3.53; found: C, 49.79; H, 5.88; N, 3.90.

3. Results and discussion

3.1. Preparation of MILs

The general synthetic strategy employed in this study is largely based on a previous study.³⁰ Acetylacetonate-based ligands have been extensively studied for the preparation of coordination complexes with metals.^{32,34,35} In this study, carefully chosen fluorinated acetylacetonate-based ligands were employed in order to modulate the hydrophobicity of the MILs,³² as the utility of previous iron-based MILs was limited due to their tendency to undergo hydrolysis in water.^{29,36} To increase the hydrophobicity of MILs, imidazolium or asymmetric phosphonium cations containing long alkyl chain substituents were employed. The alkyl chain length and asymmetry in the cation structure plays a pivotal role in the MILs being liquid at room temperature.³¹

The first step in the synthesis of the MILs evaluated in this study involves preparing the ammonium salt of the transition metals, in which three acetylacetonate-based ligands are coordinated to different transition metal centers. A one pot synthetic approach was used, in which acetylacetonate-based ligands were dissolved in acetone followed by the dropwise addition of ammonium hydroxide *via* syringe. Upon addition of ammonium hydroxide, a white vapor was formed as the result of an acid–base reaction between ammonium hydroxide and the acetylacetonate-based ligand. The reaction flask was properly sealed to ensure no vapors were lost, which can detrimentally affect the stoichiometry of the reaction. The transition metal halide salt was subsequently added after the vapor was observed to settle.

The ammonium salts of all transition metals (Ni(II), Co(II), and Mn(II)) with N_p tfacac, P_h tfacac and tfacac ligands were

purified by dissolution in diethyl ether and washed several times with water, until a $AgNO_3$ test yielded no visible precipitate. However, the solubility of ammonium salts containing T_h tfacac ligands in diethyl ether was observed to be very low. Therefore, a different strategy involving the dissolution of the product in acetone and subsequently filtering to remove insoluble ammonium chloride was employed. The second step of the reaction involved a metathesis reaction of the purified ammonium salts with either $[C_{14}MIM^+][BS^-]$ or $[P_{66614}^+][Cl^-]$.

3.2. Thermal stability

High temperature applications of hexfluoroacetylacetonate-based MILs requiring high thermal stability have been largely limited by the weak metal–ligand interactions. Pierson *et al.* compared the thermal stabilities of $[P_{66614}^+][Cl^-]$ and hexafluoroacetylacetonate-based MILs containing Co(II), Mn(II) and Ni(II) transition metal centers paired with the $[P_{66614}^+]$ cation. Using GC to measure the thermal stability, it was found that the metal–ligand complex limited the thermal stability of the MILs.³⁰ In this study, different acetylacetonate-based ligands featuring phenyl, thiophenyl, and naphthyl moieties were complexed with varying transition metals, paired with phosphonium and alkylimidazolium-based cations, and subsequently examined to study their effect on thermal stability. To study the onset temperature of MIL volatilization/decomposition, the thermal stability was determined by GC coupled to a FID, which is capable of detecting liberated volatilization/decomposition products from the MIL.³³

The Ni-based MILs prepared in this study were used to examine the ligand and cation effects on thermal stability. Thermal stability diagrams of all Ni-based MILs are shown in Fig. 1. The thermal stabilities were observed to range from approximately 130 °C to 260 °C. It was observed that the thermal stability could

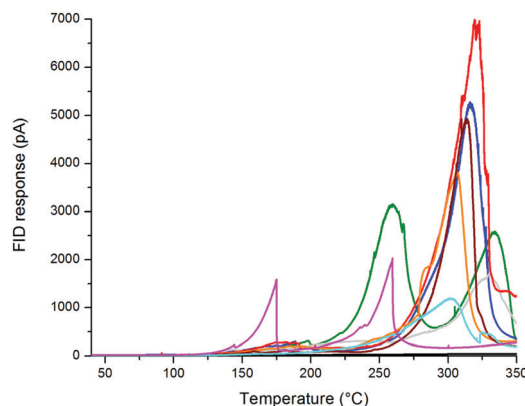


Fig. 1 Thermal stability diagram of Ni-containing MILs with varying ligand structures obtained using GC with an FID. The columns containing the MILs as a stationary phase were subjected to the following temperature program: 40 °C to 350 °C at a rate of 1 °C min^{−1}. The increased response is due to volatilization/decomposition products exiting the chromatographic column and entering the detector. Blank untreated capillary (Black line), $[C_{14}MIM^+][Ni(P_h\text{tfacac})_3^-]$ (orange line), $[P_{66614}^+][Ni(P_h\text{tfacac})_3^-]$ (red line), $[P_{66614}^+][Ni(T_h\text{tfacac})_3^-]$ (brown line), $[C_{14}MIM^+][Ni(T_h\text{tfacac})_3^-]$ (sky blue line), $[P_{66614}^+][Ni(N_p\text{tfacac})_3^-]$ (blue line), $[C_{14}MIM^+][Ni(N_p\text{tfacac})_3^-]$ (grey line), $[C_{14}MIM^+][Ni(\text{tfacac})_3^-]$ (pink line) and $[P_{66614}^+][Ni(\text{tfacac})_3^-]$ (green line).



be increased through the incorporation of aromatic moieties to the ligand structure. For example, the thermal stability of MIL 10 ($[P_{66614}^+][Ni(II)(tfacac)_3^-]$), which lacks aromatic moieties in the ligand structure, was observed to be approximately 160 °C. Similarly, the thermal stability of MIL 22 ($[C_{14}MIM^+][Ni(II)(tfacac)_3^-]$) was found to be approximately 130 °C. The thermal stability of MIL 1 ($[P_{66614}^+][Ni(II)(Thtfacac)_3^-]$), containing thiophenyl moieties in the ligand structure, was found to be approximately 260 °C. Furthermore, the thermal stabilities of MILs containing phenyl, thiophenyl, and naphthyl moieties were found to be similar and within the range of 230 °C to 260 °C.

To gain further insight into the thermal-degradation of the MILs, TGA was performed (see Fig. S1 and S2 of the ESI†) using a temperature range from 40 °C to 500 °C. MIL 10 exhibited a multistep decomposition pathway in which mass loss was observed at temperatures of approximately 150 °C, 290 °C and 400 °C. Meanwhile, MIL 7 exhibited decomposition at approximately 310 °C and 420 °C. When the TGA data is compared to thermal stability data acquired by GC-FID in Fig. 1, two peaks can be observed for MIL 10, while the other MILs showed only one decomposition peak below 350 °C. TGA analysis provides evidence that all MILs appear to decompose *via* a multi-step decomposition process. The GC method provides the most precise thermal stability since it can detect very small levels of degradation compared to TGA. This is due to the fact that very low heating increments are more easily implemented (1 °C min⁻¹) in GC compared to common heating rates in TGA (5 °C min⁻¹), which often leads to an overestimation of thermal stability.³⁷

3.3. Magnetic susceptibility

Magnetic susceptibility is an intrinsic feature of MILs that differentiates them from traditional ILs. All of the MILs examined in this work contain a transition metal center possessing unpaired electrons. When these MILs are placed in an external magnetic field, the magnetic moments originating from unpaired electrons of the 3d orbitals rotate along the magnetic field direction exhibiting paramagnetism. According to Curie's law, the magnetic susceptibility (χ) of a paramagnetic material is inversely proportional to temperature (T). An analysis of the linear relationship between the reciprocal susceptibility and temperature also is employed to reveal the magnetic characteristics of MILs, such as effective magnetic moment. The results obtained in this study show that magnetic behavior of all MILs follows the Curie-Weiss law, $\chi(T) = C/(T - \theta_p)$, where χ is magnetic susceptibility, C is Curie constant, related to the value of effective moment, and θ_p is a Weiss constant, which accounts for the presence of exchange interactions.

Magnetic susceptibility measurements were carried out on eight Ni MILs featuring four different ligands and two different cations. Fig. 2 shows representative Curie-Weiss curve for MIL 19 ($[C_{14}MIM^+][Ni(P_h tfacac)_3^-]$). Ni(II) is a d^8 system which forms octahedral complexes,³⁸ with a theoretical magnetic moment of 2.8 μ_B /Ni atom. Experimental values obtained for Ni(II)-based MILs (see Table 1) were found to be in agreement with the aforementioned theoretical value and previously reported data for other acetylacetonate-based MILs.^{30,38,39}

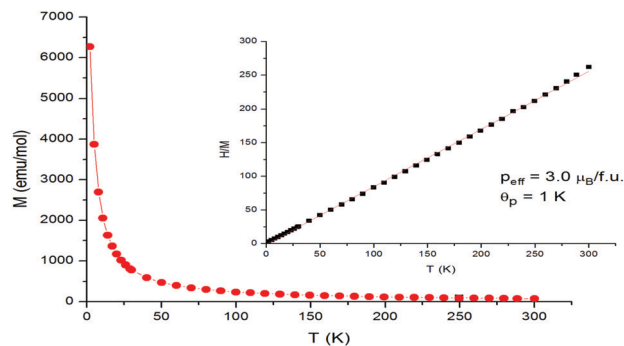


Fig. 2 Magnetization of the $[C_{14}MIM^+][Ni(P_h tfacac)_3^-]$ MIL as a function of temperature in applied magnetic field of $H = 20$ kOe between 2 and 300 K. The inset shows the Curie-Weiss fit of reciprocal magnetic susceptibility vs. temperature.

Table 1 Effective magnetic moment in Bohr magnetons (μ_B) for selected MILs examined in this study determined by SQUID magnetometer

MIL no.	MIL abbreviation	p_{eff} (μ_B) per f.u.
MIL 1	$[P_{66614}^+][Ni(Th tfacac)_3^-]$	3.1
MIL 4	$[P_{66614}^+][Ni(N_p tfacac)_3^-]$	2.6
MIL 7	$[P_{66614}^+][Ni(P_h tfacac)_3^-]$	2.7
MIL 10	$[P_{66614}^+][Ni(tfacac)_3^-]$	2.4
MIL 13	$[C_{14}MIM^+][Ni(Th tfacac)_3^-]$	3.1
MIL 16	$[C_{14}MIM^+][Ni(N_p tfacac)_3^-]$	2.8
MIL 19	$[C_{14}MIM^+][Ni(P_h tfacac)_3^-]$	3.0
MIL 22	$[C_{14}MIM^+][Ni(tfacac)_3^-]$	2.5
MIL 8	$[P_{66614}^+][Co(P_h tfacac)_3^-]$	4.3
MIL 20	$[C_{14}MIM^+][Co(P_h tfacac)_3^-]$	4.5
MIL 9	$[P_{66614}^+][Mn(P_h tfacac)_3^-]$	5.4
MIL 21	$[C_{14}MIM^+][Mn(P_h tfacac)_3^-]$	5.6

The magnetic susceptibility of two Mn-based MILs and two Co-based MILs was also measured to study the effect of different metal center on the magnetic behavior. Fig. 3 shows the Curie-Weiss plot for MIL 21 ($[C_{14}MIM^+][Mn(P_h tfacac)_3^-]$). Since Mn(II) is a d^5 system and Co(II) is a d^7 system, both of these transition metals could exhibit low or high spin configuration in octahedral complexes based upon the type of ligand. Weak field ligands result in high spin complexes while strong field

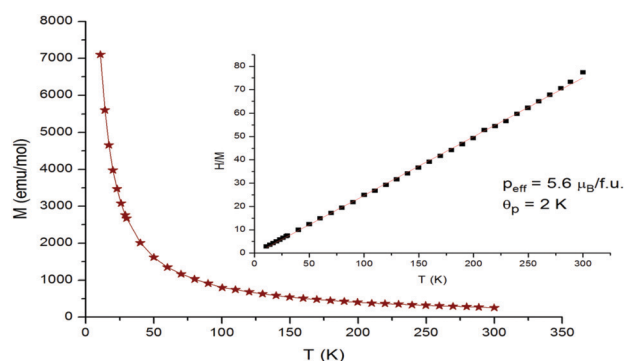


Fig. 3 Magnetization of the $[C_{14}MIM^+][Mn(P_h tfacac)_3^-]$ MIL as a function of temperature in applied magnetic field of $H = 20$ kOe between 2 and 300 K. The inset shows the Curie-Weiss fit of reciprocal magnetic susceptibility vs. temperature.



ligands give low spin complexes.⁴⁰ A magnetic moment of 5.6 μ_B /Mn was observed for the $[P_{66614}^+][Mn(P_h\text{tfacac})_3^-]$ MIL while the $[P_{66614}^+][Co(P_h\text{tfacac})_3^-]$ MIL gave a value of 4.5 μ_B /Co, showing that both of these transition metals form high spin octahedral complexes.^{38,41} High spin configuration of Mn(II) and Co(II) can be explained by the coordination of metal centers with acetylacetonate-based, weak field ligands.⁴² The Curie–Weiss plots for the remaining MILs (**1**, **4**, **7**, **8**, **9**, **10**, **13**, **16**, **20**, and **22**) are provided in Fig. S3–S12 of the ESI.† The effective magnetic moments (p_{eff}) in Bohr magnetons (μ_B) of all studied MILs calculated per formula unit/transition metal atom are presented in Table 1.

3.4. Viscosity

The viscosity of MILs is an important feature that often dictates their use in a number of applications. Low viscosity MILs are extremely useful in applications requiring extensive fluid manipulation.⁴ However, higher viscosity MILs serve important purposes in other applications that require mechanically stable liquids.²⁹ Previously reported MILs, such as trihexyl(tetradecyl)-phosphonium tetrachloromanganate(II) ($[P_{66614}^+][MnCl_4]^{2-}$), exhibited high viscosity values of approximately 75 230 cP at 25 °C.²⁷ On the contrary, low viscosity MILs such as 1-ethyl-3-methylimidazolium tetrachloroferrate(III) ($[EMIM^+][FeCl_4^-]$) and

$[BMIM^+][FeCl_4^-]$ feature viscosities of around 18 cP and 34 cP, respectively.²⁸ However, these MILs were not water stable due to hydrolysis, limiting their overall potential applications.^{29,36}

MILs possessing a broad range of viscosities were synthesized in this study. All viscosities were measured at a temperature of 65 °C to enable the determination of higher viscosity MILs. As shown in Table 2, the lowest viscosities were observed for MILs containing anions based on the tfacac ligands where viscosities ranged from 119 cP to 172 cP for MILs containing the phosphonium cation and from 845 cP to 2209 cP for MILs containing the imidazolium cation. The viscosity of MILs were observed to significantly increase when ligands containing aromatic moieties were incorporated into the chemical structure. For example, MILs possessing the $P_h\text{tfacac}$ ligand had viscosities ranging from 1139 cP to 4108 cP for those containing the phosphonium cation, and from 40 240 cP to 49 542 cP for MILs containing the imidazolium cation. MILs containing $N_p\text{tfacac}$ ligands were found to possess the highest viscosities when paired with the phosphonium cation, ranging from 42 110 cP to 48 920 cP. MILs **13** and **16** were observed to be solids, and their melting points were determined to be approximately 69 °C and 88 °C, respectively. The remaining MILs containing the imidazolium cation were too viscous to be measured as they exceeded the maximum range of viscometer (e.g. 170 000 cP). These results highlight a very significant difference

Table 2 Viscosity measurements of 24 MILs examined in this study using a Brookfield DV1 viscometer with CPA-51Z cone spindle at 65 °C

Compound abbreviation	Viscosity (cP)	Compound abbreviation	Viscosity (cP)
1: $[P_{66614}^+][Ni(T_h\text{tfacac})_3^-]$	3521	13: $[C_{14}\text{MIM}^+][Ni(T_h\text{tfacac})_3^-]$	— ^a
2: $[P_{66614}^+][Co(T_h\text{tfacac})_3^-]$	4747	14: $[C_{14}\text{MIM}^+][Co(T_h\text{tfacac})_3^-]$	— ^b
3: $[P_{66614}^+][Mn(T_h\text{tfacac})_3^-]$	5937	15: $[C_{14}\text{MIM}^+][Mn(T_h\text{tfacac})_3^-]$	— ^b
4: $[P_{66614}^+][Ni(N_p\text{tfacac})_3^-]$	42110	16: $[C_{14}\text{MIM}^+][Ni(N_p\text{tfacac})_3^-]$	— ^a
5: $[P_{66614}^+][Co(N_p\text{tfacac})_3^-]$	48 920	17: $[C_{14}\text{MIM}^+][Co(N_p\text{tfacac})_3^-]$	— ^b
6: $[P_{66614}^+][Mn(N_p\text{tfacac})_3^-]$	46 170	18: $[C_{14}\text{MIM}^+][Mn(N_p\text{tfacac})_3^-]$	— ^b
7: $[P_{66614}^+][Ni(P_h\text{tfacac})_3^-]$	4108	19: $[C_{14}\text{MIM}^+][Ni(P_h\text{tfacac})_3^-]$	49 542
8: $[P_{66614}^+][Co(P_h\text{tfacac})_3^-]$	1916	20: $[C_{14}\text{MIM}^+][Co(P_h\text{tfacac})_3^-]$	47 120
9: $[P_{66614}^+][Mn(P_h\text{tfacac})_3^-]$	1139	21: $[C_{14}\text{MIM}^+][Mn(P_h\text{tfacac})_3^-]$	40 240
10: $[P_{66614}^+][Ni(\text{tfacac})_3^-]$	127	22: $[C_{14}\text{MIM}^+][Ni(\text{tfacac})_3^-]$	2060
11: $[P_{66614}^+][Co(\text{tfacac})_3^-]$	172	23: $[C_{14}\text{MIM}^+][Co(\text{tfacac})_3^-]$	2209
12: $[P_{66614}^+][Mn(\text{tfacac})_3^-]$	119	24: $[C_{14}\text{MIM}^+][Mn(\text{tfacac})_3^-]$	845

^a MILs are solids at a temperature of 65 °C. ^b MILs had viscosity values that exceeded the maximum range of the instrument at 65 °C.

Table 3 Solubility of all MILs examined in this study in ten different organic solvents

Compound abbreviation	Solubility	Compound abbreviation	Solubility
1: $[P_{66614}^+][Ni(T_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f	13: $[C_{14}\text{MIM}^+][Ni(T_h\text{tfacac})_3^-]$	S ^a , S ^b , I ^e , I ^f
2: $[P_{66614}^+][Co(T_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f	14: $[C_{14}\text{MIM}^+][Co(T_h\text{tfacac})_3^-]$	S ^a , S ^b , I ^e , I ^f
3: $[P_{66614}^+][Mn(T_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f	15: $[C_{14}\text{MIM}^+][Mn(T_h\text{tfacac})_3^-]$	S ^a , S ^b , I ^e , I ^f
4: $[P_{66614}^+][Ni(N_p\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f	16: $[C_{14}\text{MIM}^+][Ni(N_p\text{tfacac})_3^-]$	S ^a , S ^b , I ^e , I ^f
5: $[P_{66614}^+][Co(N_p\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f	17: $[C_{14}\text{MIM}^+][Co(N_p\text{tfacac})_3^-]$	S ^a , S ^b , I ^e , I ^f
6: $[P_{66614}^+][Mn(N_p\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f	18: $[C_{14}\text{MIM}^+][Mn(N_p\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f
7: $[P_{66614}^+][Ni(P_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^c , I ^f	19: $[C_{14}\text{MIM}^+][Ni(P_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f
8: $[P_{66614}^+][Co(P_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^c , I ^f	20: $[C_{14}\text{MIM}^+][Co(P_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f
9: $[P_{66614}^+][Mn(P_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^c , I ^f	21: $[C_{14}\text{MIM}^+][Mn(P_h\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f
10: $[P_{66614}^+][Ni(\text{tfacac})_3^-]$	S ^a , S ^b , S ^c , I ^f	22: $[C_{14}\text{MIM}^+][Ni(\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f
11: $[P_{66614}^+][Co(\text{tfacac})_3^-]$	S ^a , S ^b , S ^c , I ^f	23: $[C_{14}\text{MIM}^+][Co(\text{tfacac})_3^-]$	S ^a , S ^b , S ^c , I ^f
12: $[P_{66614}^+][Mn(\text{tfacac})_3^-]$	S ^a , S ^b , S ^c , I ^f	24: $[C_{14}\text{MIM}^+][Mn(\text{tfacac})_3^-]$	S ^a , S ^b , S ^d , I ^f

S^a = soluble in diethyl ether, dichloromethane, chloroform, methanol, acetone and acetonitrile at 20% (w/v). S^b = soluble in toluene and dimethyl sulfoxide at 10% (w/v). S^c = soluble in hexane at 10% (w/v). S^d = soluble in hexane at <1% (w/v). I^e = insoluble in hexane at 0.01% (w/v). I^f = insoluble in water at 0.01% (w/v).



in the viscosities for MILs comprised of imidazolium and phosphonium cations.

3.5. Solvent miscibility

The solubility of MILs in various organic solvents is very important in order to exploit them in applications including photoluminescent materials,⁴³ liquid membranes,²⁷ reaction media in organic synthesis,^{14,15} and catalysis.^{10,11} The solubility of MILs in ten different organic solvents, with polarities ranging from 0.009 to 1.000 according to the Reichardt polarity index,⁴⁴ were studied to develop a fundamental understanding of their solubilities. A summary of the results is shown in Table 3. Since the MILs were designed to be hydrophobic and hydrolytically stable, all MILs were found to be insoluble in water at a 0.01% (w/v) MIL-to-solvent ratio. The MILs were found to be fully miscible at 20% (w/v) MIL-to-solvent ratio in most organic solvents including diethyl ether, dichloromethane, chloroform, methanol, acetone, and acetonitrile. Their solubility was found to be 10% (w/v) in DMSO and toluene, which have polarity values of 0.444 and 0.099, respectively.⁴⁴ The MILs exhibited varying solubilities in hexane, with MILs 7–12 (containing P_htfacac and tfacac ligands and paired with the [P₆₆₆₁₄⁺] cation) possessing a solubility of 10% (w/v). However, MILs 13–18 (containing T_htfacac and N_ptfacac ligands and paired with [C₁₄MIM⁺] cation) were found to be insoluble. The remaining MILs were only sparingly soluble in hexane, with a solubility of <1% (w/v). The varying solubility of this generation of MILs in different organic solvents further enhances their versatility and range of applications in which they may find utility.

4. Conclusions

In this study, twenty four (24) different MILs with varying transition metal centers (Ni(II), Co(II), and Mn(II)), ligands (T_htfacac, N_ptfacac, P_htfacac, and tfacac) and cations ([P₆₆₆₁₄⁺] or [C₁₄MIM⁺]) have been successfully synthesized and characterized. This study demonstrates that thermal stability of acetylacetone-based MILs can be increased through the incorporation of aromatic moieties in the ligand structure, as observed with MILs synthesized with thiophenyl, naphthyl, and phenyl-based ligands. Lower thermal stability was observed for MILs containing the tfacac ligand which lacks aromatic moieties within the ligand. The highest thermal stability was found to be approximately 260 °C for MIL 1 ([P₆₆₆₁₄⁺][Ni(II)(T_htfacac)₃[−]]), which is approximately 50 °C higher than previously reported acetylacetone-based MILs. Furthermore, viscosities ranging from a couple of hundred centipoise to several thousand centipoise were achieved. The lowest viscosities were achieved by incorporating the [P₆₆₆₁₄⁺] cation into the MIL structure. Conversely, MILs containing the imidazolium cation had much higher viscosities on average, with some MILs possessing melting points above room temperature. Magnetic susceptibility was found to be in agreement with high spin complexes of transition metals (Ni(II), Co(II), and Mn(II)). The synthesized MILs were found to be extremely hydrophobic as they were insoluble in water down to a MIL-to-solvent ratio of 0.01% (w/v). This study provides insight into the physiochemical

properties of MILs and how they can be controlled through modification to various structural components including metal center, ligand, and cation.

Conflicts of interest

There are no conflicts to declare.

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