

Cite this: *Nanoscale Adv.*, 2019, 1, 3693

ITO@SiO₂ and ITO@{M₆Br₁₂}@SiO₂ (M = Nb, Ta) nanocomposite films for ultraviolet-near infrared shielding†

Wanghui Chen,^{id abd} Thi Kim Ngan Nguyen,^{id abd} Maxence Wilmet,^{id ac} Noée Dumait,^c Ourania Makrygenni,^{abd} Yoshio Matsui,^{id d} Toshiaki Takei,^e Stéphane Cordier,^{id c} Naoki Ohashi,^{id ad} Tetsuo Uchikoshi^{id abd} and Fabien Grasset^{id *ad}

Transparent optical thin films for energy saving applications have recently gained substantial prominence for functional window processes. In this study, highly visible transparent nanocomposite films with ultraviolet (UV) and near-infrared (NIR) blocking capabilities are reported. Such nanocomposite films, prepared by electrophoretic deposition on ITO-coated glass, are composed of indium tin oxide (ITO) nanocrystals (9 nm) and octahedral metal atom clusters (1 nm, Nb₆ or Ta₆) embedded into silica nanoparticles (~80 nm). The functional silica nanoparticles were prepared by a reverse microemulsion process. The microstructural characterization proved that ITO nanocrystals are centered in the silica nanoparticles, whereas the metal atom clusters are homogeneously distributed in the silica matrix. The optical absorption spectra of these transparent nanocomposite films exhibit distinct and complementary contributions from their ITO nanoparticles and metal atom clusters (absorption in the UV range) and from the ITO layer on silica.

Received 24th June 2019
Accepted 5th August 2019

DOI: 10.1039/c9na00400a
rsc.li/nanoscale-advances

1 Introduction

The development of functional coatings for window applications using metal, oxide or sulfide nanoparticles (NPs) with localized surface plasmon resonance (LSPR) frequencies in the near-infrared region (NIR) has attracted great attention. This increased interest is due to their strong potential towards energy saving applications by blocking NIR radiation.^{1–11} In the family of common LSPR oxide materials (CdO, ZnO, In₂O₃, and WO_{3–x}), Sn:In₂O₃, indium tin oxide (ITO), appears to be a promising candidate for functional coating on glasses, owing to its LSPR properties. Furthermore, these properties correspond with the most energetic NIR radiation, responsible for heating.^{12,13} Indeed, the average solar energy distribution on the Earth's surface (after absorption through the atmosphere) is

composed of approximately 5% UV radiation (290–380 nm), 46% visible radiation (380–780 nm) and approximately 49% NIR radiation (780–2500 nm). Regarding the absorption process, O₃ is responsible for the total radiation absorption below 290 nm, while CO₂ and H₂O vapor mainly absorb the thermal infrared radiation, with impacts on the Earth's temperature. Thus, almost half of the solar energy comes from infrared radiation, particularly from the highest thermal energy region (780 to 1100 nm).^{14–16} Despite the abundant availability of commercial ITO-coated glasses in the market, besides the commonly used sputtering process using an ITO target and Ar ion beam, their fabrication routes are limited.^{17,18} Currently solution-based processes are highly desirable for reducing the cost of these materials and as a viable alternative to the sputtering process for the synthesis of ITO NPs and ITO nano inks.^{19–22} In parallel, our group has recently proposed the use of transition metal nanoclusters as UV-NIR blocking building blocks. In this regard, they were successfully tested for the development of nanocomposite coatings by electrophoretic deposition (EPD).²³ The fabricated coatings were found to be highly transparent in the visible range, with tunable UV and NIR absorption properties. Indeed, octahedral M₆ nanoclusters (M = Nb, Mo or Ta) demonstrated promising optical properties for energy saving applications.^{24–27}

In this new study, we design multifunctional silica nanocomposite thin films obtained by the combination of the complementary LSPR properties found on ITO NPs with the UV-

^aCNRS-Saint Gobain-NIMS, UMI3629, Laboratory for Innovative Key Materials and Structures (LINK), National Institute for Materials Science (NIMS), 1-1 Namiki, Tsukuba, Ibaraki, 305-0044, Japan. E-mail: fabien.grasset@univ-rennes1.fr

^bFine Particles Engineering Group, Research Center for Functional Materials (RCFM), NIMS, 1-2-1 Sengen, Tsukuba, Ibaraki, 305-0047, Japan

^cUniv. Rennes, CNRS, ISCR – UMR6226, 263 av. du Général Leclerc, 35042 Rennes, France

^dRCFM, NIMS, 1-1 Namiki, Tsukuba, Ibaraki, 305-0044, Japan

^eResearch Center for Materials Nanoarchitectonics (MANA), NIMS, 1-1 Namiki, Tsukuba, Ibaraki, 305-0044, Japan

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9na00400a



NIR absorption properties of Nb₆ or Ta₆ octahedral nanoclusters. Silica NPs of ITO@M₆@SiO₂ (M = Nb or Ta) were synthesized by a reverse microemulsion (RM) process, and then transparent thin films were fabricated by deposition of the NPs by EPD. It must be noted that both these processes were developed by our groups.^{28,29} The synthesis of hydrophilic ITO NPs by a solvothermal method, the synthesis of functional SiO₂ NPs by the microemulsion process, and the fabrication of thin films and the optical properties of these nanocomposite NPs and thin films are elucidated in this paper. Herein, for the first time, we report on the fabrication of UV-NIR protective transparent silica nanocomposite films. These films allow the absorption of the most energetic UV (330–380 nm) and NIR radiation (780–1100 nm) due to the presence of M₆ nanoclusters and the selectivity of the least energetic NIR radiation (1200–2500 nm) due to ITO.

2 Materials and methods

2.1 Materials

InCl₃·4H₂O, SnCl₄·5H₂O, ethanol (Ethanol, 99.5%), ethylene glycol (EG), ammonia solution (28%) and tetraethyl orthosilicate (TEOS, 95%) were purchased from Wako Pure Chemical Industries. *n*-Heptane (98%) was purchased from Kanto Chemical Co. The nonionic surfactant Brij L4, PVP K30, and 25 wt% tetramethylammonium hydroxide (TMAH) in methanol were purchased from Sigma-Aldrich. DI water was purchased from Yamato Kagaku Co. All the chemicals were used as received without further purification.

2.2 Preparation of nanoparticles

In this study, as previously reported, solvothermal and sol-gel methods were used to synthesize the ITO NPs.²⁰ The starting compounds Ta₆Br₁₄(H₂O)₄·3H₂O (TBH) and K₄Nb₆Br₁₈ (KNB) were synthesized according to a procedure that was previously published.^{30,31} The experimental conditions for the microemulsion process were adapted from our previous work.³² The experimental conditions for all these compounds are reported in the ESI.†

2.3 Preparation of thin films

Regarding the films, the as-synthesized ITO@M₆@SiO₂ NPs were dispersed into acetone at different concentrations (0.8 to 2.5 g L⁻¹) followed by ultrasonication prior to their deposition by EPD. The zeta potential and electric conductivity of the suspensions were measured using a zeta-potential analyzer (Malvern Instrument, Ltd., Zetasizer Nano Z). ITO-coated glass (Geomatec Co., Ltd (Japan), 6–8 Ohm sq⁻¹) was washed with distilled water and acetone under ultrasonication. Then, it was connected to a Keithley Model 2400 Series SourceMeter and used as the anodic substrate, and a stainless-steel sheet was used as the cathodic counter electrode. Different tests were carried out by varying the applied voltage from 10 to 30 V and the deposition time from 20 to 90 s. The nanocomposite coated glasses were allowed to dry at room temperature overnight prior to their characterization.

2.4 Characterization methods and techniques

Several complementary techniques were used for the characterization of the powders and thin film microstructure. Field emission-scanning electron microscopy (FE-SEM) images were obtained using a JEOL JSM 6301F microscope operating at 7 kV. High resolution TEM images, high angle annular dark field scanning transmission electron microscopy (HAADF-STEM) images and energy-dispersive (EDX) spectra/maps were acquired by using a high resolution transmission electron microscope (HRTEM), JEM-2100F (JEOL Co.), equipped with an EDX detector. Powder X-ray diffraction (XRD) patterns were acquired using a D/MAX Ultima III (Rigaku Co.) diffractometer using Cu K α radiation ($\lambda = 1.5418 \text{ \AA}$) over a range of $20^\circ < 2\theta < 90^\circ$. Diffuse reflectance UV-Vis-NIR and direct transmission spectra were collected using an integrating sphere equipped with a V-7200 spectrometer (Jasco Co.). Dynamic light scattering (DLS) measurements for measuring the particle size distribution in solution were performed on an Osaka zeta-sizer.

3 Results and discussion

Fig. 1a shows the general morphology of the obtained ITO NPs. As can be seen, the ITO NPs have a size centered around 9 nm with a quite narrow size distribution (the size distribution estimated by TEM is reported in the ESI, Fig. S1†). The high magnification HRTEM image (Fig. 1b) and the observed lattice pattern demonstrate the high crystallinity of the ITO NPs. The XRD pattern indicates the single crystalline phase of the ITO NPs (Fig. 1c), as all the diffraction peaks can be indexed to the standard pattern of In₂O₃ (PDF #65-3170)—the doping of Sn⁴⁺ in the lattice of In₂O₃ gives rise to a small shift towards high angles for all the diffraction peaks.³³

The obtained ITO NPs could be easily dispersed in water owing to their hydrophilic surface (see the ESI, Fig. S2†).²⁰ Nevertheless, in order to ensure the integration of only one or two ITO NPs per inverse micelle, the NPs had to be fully dispersed in the water phase, *i.e.* no aggregates of ITO NPs were allowed. The dependence of the as-synthesized ITO NP

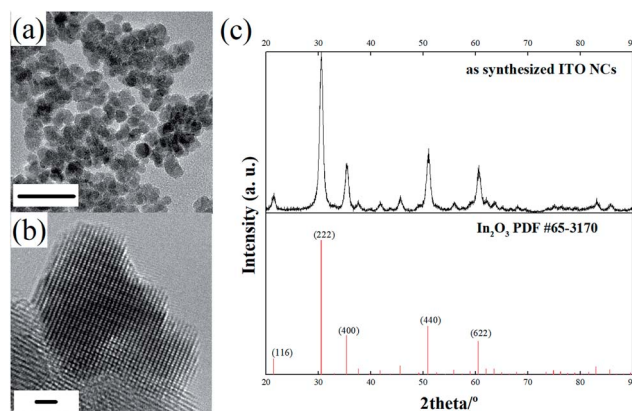


Fig. 1 TEM (scale bar: 20 nm) (a) and HRTEM (scale bar: 1 nm) (b) images of the as-synthesized ITO NPs; (c) comparison of the XRD pattern of ITO NPs with the standard diffraction pattern of In₂O₃.



dispersion in water on the pH was investigated with DLS measurements. As shown in Fig. 2a and b, when the pH of the solution is around ~ 9.7 , the ITO NPs reach their highest average hydrodynamic diameter (294 nm) with a relatively wide size distribution. Increasing or decreasing the pH has an impact on the dispersion of ITO NPs in water, in terms of both the average hydrodynamic diameter and distribution. ITO NPs are found to have an average hydrodynamic diameter of 59 nm and 38 nm for pH values of ~ 2 and ~ 12.5 , respectively. This behavior is fully consistent with the isoelectric point value of In_2O_3 NPs with a similar size in water media, *i.e.* $\text{pH} = 9.2$.³⁴

For RM-based synthesis, a diluted ammonia solution with a pH equal to 12.5 was selected for the dispersion of ITO NPs, based on the following two considerations: (1) at this pH value, the ITO NPs attained the best dispersion combined with a decreased average hydrodynamic diameter and (2) a basic pH is one of the prerequisites for the successful formation of SiO_2 from alkoxide in the RM system.^{35,36} Indeed, the diffusion of the alkoxide into the water droplets promotes its hydrolysis and condensation resulting in the formation of the silica network.³⁷

Digital photographs of the RM systems containing ITO NPs (left) and ITO@Ta_6 (right) are shown in Fig. 2c, for the sake of comparison. The high transparency of these two RM systems implies their good uniformity and stability regarding the definition of microemulsions from Danielsson and Lindman: “a microemulsion is a system of water, oil and an amphiphile which is a single optically isotropic and thermodynamically stable liquid solution”.³⁸

Fig. 3a shows a typical HAADF-STEM image of ITO@SiO_2 NPs, wherein SiO_2 and ITO NPs can be easily distinguished according to the difference in contrast. The resulting SiO_2 particles had a uniform spherical morphology with an average size of 85 nm (size distribution estimated by TEM is reported in the ESI, Fig. S3†), attributed to the well confined aqueous phase in the RM system. ITO NPs were found in the center of most SiO_2 NPs (Fig. 3b). The EDS mapping (Fig. 3c) further validates the core-shell structure of the ITO@SiO_2 NPs, as both In and Sn elements are located in the center of each single particle. Such a uniform morphology also demonstrates the good dispersion of the ITO NPs in aqueous solution in the step before the RM-based synthesis.

HAADF-STEM micrographs of the synthesized $\text{ITO@M}_6\text{@SiO}_2$ NPs ($M = \text{Nb}, \text{Ta}$) also indicate a similar core-shell structure (results of $\text{ITO@Nb}_6\text{@SiO}_2$ NPs are presented in the ESI,

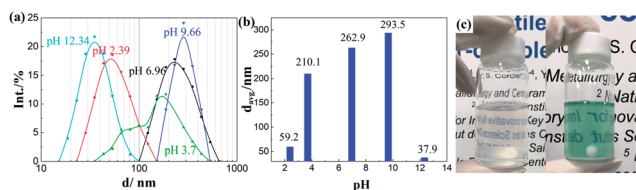


Fig. 2 (a) Hydrodynamic size distribution versus intensity (from DLS measurement) of ITO NPs in water by varying the pH; (b) plots of the average hydrodynamic diameter of ITO NPs versus the pH in solution; (c) digital photographs of the RM system containing ITO NPs and ITO@Ta_6 NPs.

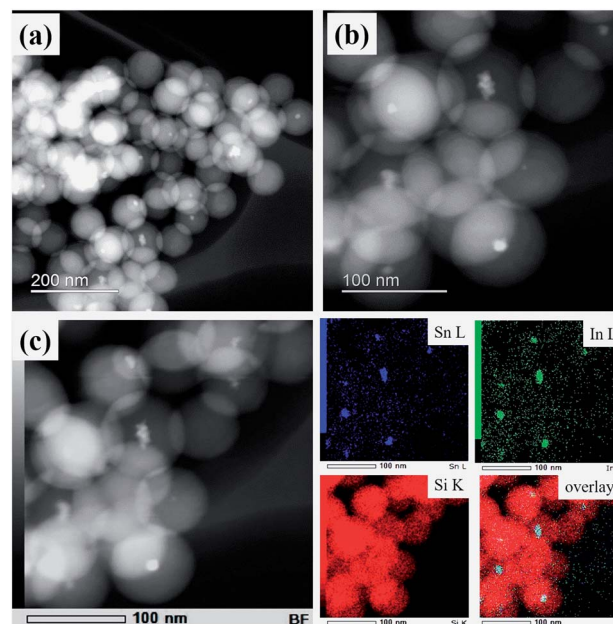


Fig. 3 Low (a) and high (b) magnification HAADF-STEM images of ITO@SiO_2 NPs; (c) EDS mapping indicating the distribution of Sn L, In L and Si K in ITO@SiO_2 NPs.

Fig. S4a and b†), with the presence of numerous discrete Ta_6 cluster species (Fig. 4a and b). Similar to that of the ITO@SiO_2 NPs, the size of $\text{ITO@Ta}_6\text{@SiO}_2$ NPs shows good uniformity, but it slightly decreased to 76 nm on average (size distribution estimated by TEM is reported in the ESI, Fig. S3†). In most cases, ITO NPs were observed at the core of a single

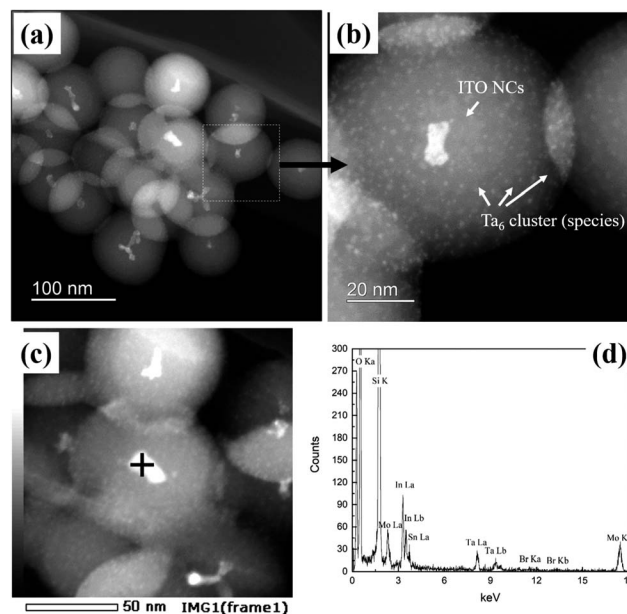


Fig. 4 Low (a) and high (b) magnification HAADF-STEM images of $\text{ITO@Ta}_6\text{@SiO}_2$ NPs; (c) STEM image and (d) EDS spectrum (point mode) of $\text{ITO@Ta}_6\text{@SiO}_2$ NPs revealing the co-existence of Si, O, In, Sn, Ta, and Br elements.



ITO@M₆@SiO₂ nanoparticle, while the presence of M₆ cluster species in the shell was validated from the high magnification HAADF-STEM micrograph (ITO@Ta₆@SiO₂ NPs are presented in Fig. 4b and ITO@Nb₆@SiO₂ NPs are shown in the ESI, Fig. S4a†). As can be seen, the M₆ cluster units are discrete species with a size of 1 nm and they are well dispersed in the whole SiO₂ shell, which is in good agreement with a previous study by our group.³² The EDS spectrum in Fig. 4c and d further confirms the successful synthesis of ITO@Ta₆@SiO₂ NPs, as all the elements (Si, O, In, Sn, Ta, and Br) were detected in a single particle. Similar results were obtained for ITO@Nb₆@SiO₂ NPs (see the ESI, Fig. S4b†).

It is well known that silica is a material with high transparency in the range of 250 nm to 2000 nm. As seen in Fig. 5, the SiO₂ NPs synthesized by the RM method showed almost no absorption in this region.

However their absorption bands at the wavelengths of ~1400 and 1870 nm were assigned to the OH groups and the surface adhered water molecules. On the other hand, the ITO@SiO₂ NPs showed strong absorption in the UV and NIR regions. Firstly, an absorption band can be distinguished at a wavelength below 350 nm.

In the NIR region, absorption bands can be found at wavelengths after 1100 nm and they reach a maximum at 1900 nm. Introducing M₆ cluster-based species into the ITO@SiO₂ NPs can improve the absorption in these regions. Indeed, as observed in Fig. 5, increasing the concentration of Ta₆ cluster provoked an increase in the absorption in all spectral regions. When the concentration of TBH reached a value of 6 mM or higher, three more absorption bands were observed besides those assigned to the ITO NPs. These three peaks, with maxima values of 361, 663 and 850 nm, could be clearly assigned to the absorption of {Ta₆Br₁₂}³⁺ cluster cores.³² A comparison between the absorption spectrum of ITO@SiO₂ NPs and that of ITO@Ta₆@SiO₂ NPs clearly evidences that the latter exhibit the

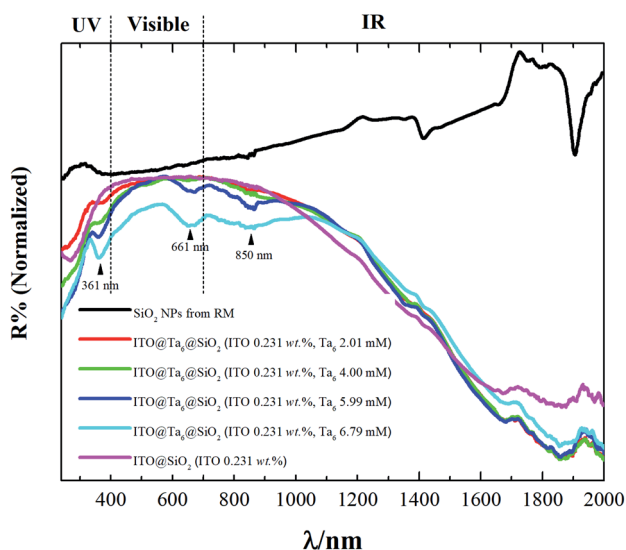


Fig. 5 Diffuse reflectance UV-Vis-NIR spectra of pure SiO₂, ITO@SiO₂ and ITO@Ta₆@SiO₂ NPs containing different concentrations of TBH (from 2 to 6.8 mM).

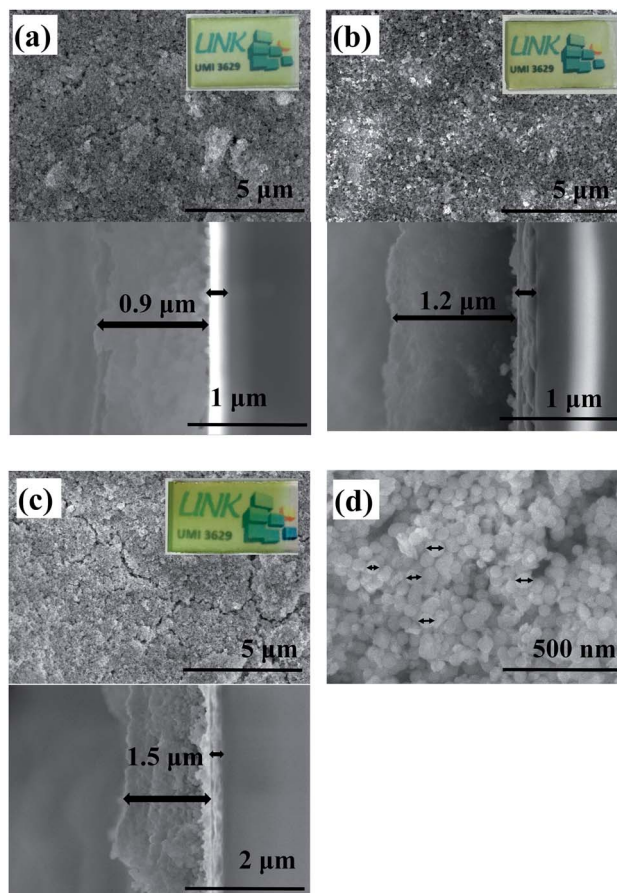


Fig. 6 Optical photographs and SEM micrographs, surface morphology and cross sections of ITO@Nb₆@SiO₂ based films. These films were prepared from solution concentrations equal to 0.8 (a), 1.0 (b), and 2.5 (c, d) g L⁻¹ and deposited by EPD at 20 V for 20 s.

best UV and NIR blocking properties. This distinctly demonstrates that ITO and Ta₆ cluster-based units act as complementary functional building blocks, *i.e.*, on one hand they have the LSPR properties of ITO NPs and on the other hand they have the UV-NIR blocking properties of Nb₆ or Ta₆ octahedral nanoclusters.

The as-synthesized ITO@M₆@SiO₂ NPs were successfully used for coating on conductive glass by EPD. For the preparation of films based on ITO@M₆@SiO₂, EPD was performed at 20 V for 20 s by altering the solution concentration. The values of concentration were in the range of 0.8, 1.0 and 2.5 g of NPs per liter.

The zeta potential and the electric conductivity of the ITO@M₆@SiO₂ suspension in acetone (99.5%) were measured. For instance, the zeta potential and electric conductivity values of the solution with a concentration equal to 2.5 g L⁻¹ are $-(22.3 \pm 0.6)$ mV and $(19 \pm 3) \times 10^{-3}$ mS cm⁻¹, respectively. These results confirm that anodic EPD is effective for the deposition of ITO@M₆@SiO₂ NPs onto ITO-coated glass. Only the results based on ITO@Nb₆@SiO₂ NPs are shown in this paper. Similarly, different techniques were used to characterize the films based on ITO@Nb₆@SiO₂ NPs. The results of these



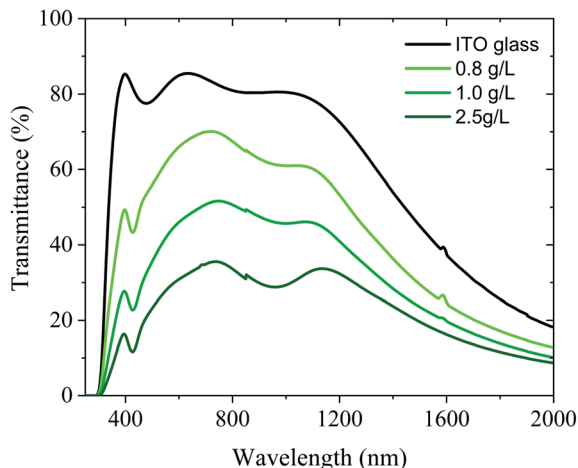


Fig. 7 Transmission UV-Vis-NIR spectra of the ITO@Nb₆@SiO₂ NP based films on the ITO-coated glass by varying the solution concentration (0.8; 1.0; and 2.5 g L⁻¹); the spectrum of the substrate is used as a reference.

films deposited onto ITO-coated glass are presented below. SEM and UV-Vis-NIR spectroscopy were carried out for the characterization of the morphology and the optical properties. In Fig. 6, the morphology of the films obtained from suspensions with variable concentrations is displayed. As expected, thanks to EPD, the film is composed of a quite dense layer of silica. Nevertheless, the film remained optically transparent in the visible range, as observed in the camera images (Fig. 6), despite the roughness of the surface. The proportional dependence of the film thickness (0.9 to 1.5 microns) on the concentration of the solution is demonstrated in cross sectional SEM images.

These results were confirmed by UV-Vis-NIR spectroscopy in Fig. 7 showing the decrease in the transparency when the

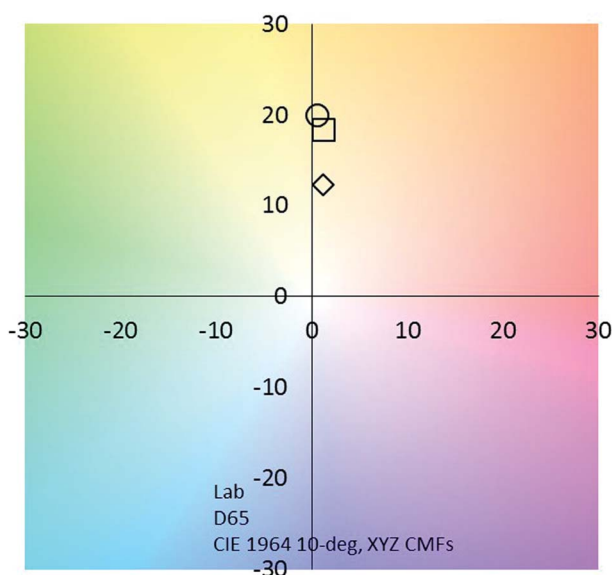


Fig. 8 CIE chromaticity coordinates of the ITO@Nb₆@SiO₂ based films (◇ 0.8 g L⁻¹; □ 1 g L⁻¹; ○ 2.5 g L⁻¹).

Table 1 $L^*a^*b^*$ color coordinates and T_L and T_E values of the ITO@Nb₆@SiO₂ NP based films

Sample name	L^*	a^*	b^*	T_L	T_E
ITO@Nb ₆ @SiO ₂ 20 V 20 s 0.8 g L ⁻¹	82.470	1.167	12.314	61.8	55.6
ITO@Nb ₆ @SiO ₂ 20 V 20 s 1 g L ⁻¹	70.266	1.225	18.402	41.8	39.2
ITO@Nb ₆ @SiO ₂ 20 V 20 s 2.5 g L ⁻¹	59.196	0.542	20.067	27.8	26.4

concentration of the solution is increased. Similar results can be obtained by altering the applied voltage or the deposition time (Fig. S5†). As expected, the ITO-coated glass showed high transparency in both UV-A and visible regions. However, the coating with ITO@Nb₆@SiO₂ NPs had a significant impact on the absorption in the UV and NIR region as reflected by the appearance of a band at 800 nm instead of 1200 nm for the ITO-coated glass. This new absorption band indicates the improvement towards energy saving applications.

Very interestingly, compared to those for pure ITO-coated glass,³⁹ the values of T_E are improved by 28% to 66% for the ITO@Nb₆@SiO₂ films with 0.8 to 2.5 g L⁻¹, respectively (Fig. 8 and Table 1). This reduction markedly manifests the potential of this new nanocomposite coating for energy saving applications. At the same time, when films are fabricated with a solution concentration equal to 0.8 g L⁻¹, the value of the T_L remained higher than 60%. This value is perfectly suitable for window applications (Table 1).^{2,40}

4 Conclusions

In this study, ITO@SiO₂ and ITO@M₆@SiO₂ NPs were successfully synthesized by an RM-based method with a size adequate for transparent coating. The HAADF-STEM and EDS analysis revealed the core-shell structure of these composite nanoparticles and the well-controlled size and uniformity of SiO₂ architectures. Herein, the transparent films based on silica nanocomposites, including ITO NPs and M₆ clusters, were fabricated by a low-cost EPD solution process for the first time. These films can be potentially used for UV and NIR blocking static applications. Indeed, EPD allowed the fabrication of well controlled films with a thickness of micrometre size, in a very short time (less than 1 min). The UV-Vis-NIR spectra and the values of T_L and T_E of coated films clearly indicate the UV-NIR blocking properties of these coatings.

Conflicts of interest

There is no conflicts to declare.

Acknowledgements

The authors wish to thank David Lechevalier and Dr Meng Zhou of Saint-Gobain KK (Tokyo, Japan). This work was carried out as part of the France–Japan international collaboration framework (UMI 3629 LINK).



Notes and references

- 1 S. Schelm and G. Smith, *Appl. Phys. Lett.*, 2003, **82**, 4346–4348.
- 2 S. Schelm, G. Smith, P. D. Garrett and W. K. Fisher, *J. Appl. Phys.*, 2005, **97**, 124314–124318.
- 3 J. E. Song, Y. H. Kim and Y. S. Kang, *Curr. Appl. Phys.*, 2006, **6**, 791–795.
- 4 K. Katagiri, R. Takabatake and K. Inumaru, *ACS Appl. Mater. Interfaces*, 2013, **5**, 10240–10245.
- 5 A. Llordes, G. Garcia, J. Gazquez and D. J. Milliron, *Nature*, 2013, **500**, 323–326.
- 6 T. M. Mattox, A. Agrawal and D. J. Milliron, *Chem. Mater.*, 2015, **27**, 6620–6624.
- 7 C. J. Barile, D. J. Slotcavage and M. D. McGehee, *Chem. Mater.*, 2016, **28**, 1439–1445.
- 8 R. W. Johns, H. A. Bechtel, E. L. Runnerstrom, A. Agrawal, S. D. Lounis and D. J. Milliron, *Nat. Commun.*, 2016, **7**, 11583.
- 9 B. H. Kim, C. M. Staller, S. H. Cho, S. Heo, C. E. Garrison, J. Kim and D. J. Milliron, *ACS Nano*, 2018, **12**, 3200–3208.
- 10 L. V. Besteiro, X. T. Kong, Z. Wang, F. Rosei and A. O. Govorov, *Nano Lett.*, 2018, **18**, 3147–3156.
- 11 M. Kanehara, H. Koike, T. Yoshinaga and T. Teranishi, *J. Am. Chem. Soc.*, 2009, **131**, 17736–17737.
- 12 A. Agrawal, S. H. Cho, O. Zandi, S. Ghosh, R. W. Johns and D. J. Milliron, *Chem. Rev.*, 2018, **118**, 3121–3207.
- 13 Y. T. Kwon, S. H. Ryu, J. W. Shin, W. H. Yeo and Y. H. Choa, *ACS Appl. Mater. Interfaces*, 2019, **11**, 6575–6580.
- 14 C. G. Granqvist, *Thin Solid Films*, 2014, **564**, 1–38.
- 15 E. Lichtfouse, *Sustainable Agriculture Reviews*, Springer, 2015.
- 16 F. Pacheco-Torgal, J. Labrincha, L. Cabeza and C. G. Granqvist, *Design, Properties and Applications*, Woodhead Publishing, 2015.
- 17 L. Kerkache, A. Layadi, E. Dogheche and D. Remiens, *J. Phys. D: Appl. Phys.*, 2006, **39**, 184–189.
- 18 O. Tuna, Y. Selamet, G. Aygun and L. Ozyuzer, *J. Phys. D: Appl. Phys.*, 2010, **43**, 055402–055407.
- 19 Z. Chen, X. Qin, T. Zhou, X. Wu, S. Shao, M. Xie and Z. Cui, *J. Mater. Chem. C*, 2015, **3**, 11464–11470.
- 20 A. Muramatsu, K. Kanie, T. Sasaki and M. Nakaya, *KONA Powder Part. J.*, 2016, **33**, 340–353.
- 21 C. Kim, J. W. Park, J. Kim, S. J. Hong and M. J. Lee, *J. Alloys Compd.*, 2017, **g**, 712–719.
- 22 S. J. Hong, S. J. Cha and J. Y. Lee, *Powder Technol.*, 2018, **332**, 56–62.
- 23 N. T. K. Nguyen, A. Renaud, B. Dierre, B. Bouteille, M. Wilmet, M. Dubernet, N. Ohashi, F. Grasset and T. Uchikoshi, *Bull. Chem. Soc. Jpn.*, 2018, **91**, 1763–1774.
- 24 N. T. K. Nguyen, M. Dubernet, Y. Matsui, M. Wilmet, N. Shirahata, G. Rydzek, N. Dumait, M. Amela-Cortes, A. Renaud, S. Cordier, Y. Molard, F. Grasset and T. Uchikoshi, *R. Soc. Open Sci.*, 2019, **6**, 181647.
- 25 T. K. N. Nguyen, A. Renaud, M. Wilmet, N. Dumait, S. Paofai, B. Dierre, W. Chen, N. Ohashi, S. Cordier, F. Grasset and T. Uchikoshi, *J. Mater. Chem. C*, 2017, **5**, 10477–10484.
- 26 T. K. N. Nguyen, B. Dierre, F. Grasset, A. Renaud, S. Cordier, P. Lemoine, N. Ohashi and T. Uchikoshi, *J. Electrochem. Soc.*, 2017, **164**, D412–D418.
- 27 T. K. N. Nguyen, F. Grasset, B. Dierre, C. Matsunaga, S. Cordier, P. Lemoine, N. Ohashi and T. Uchikoshi, *ECS J. Solid State Sci. Technol.*, 2016, **5**, R178–R186.
- 28 F. Grasset, F. Dorson, S. Cordier, Y. Molard, C. Perrin, A. M. Marie, T. Sasaki, H. Haneda, Y. Bando and M. Mortier, *Adv. Mater.*, 2008, **20**, 143–148.
- 29 Y. Sakka and T. Uchikoshi, *KONA Powder Part. J.*, 2010, **28**, 74–90.
- 30 F. W. Koknat, J. A. Parsons and A. Vongvusharintra, *Inorg. Chem.*, 1974, **13**, 1699–1702.
- 31 B. G. Hughes, J. L. Meyer, P. B. Fleming and R. E. McCarley, *Inorg. Chem.*, 1970, **9**, 1343–1346.
- 32 W. Chen, M. Wilmet, T. G. Truong, N. Dumait, S. Cordier, Y. Matsui, T. Hara, T. Takei, N. Saito, T. K. N. Nguyen, T. Ohsawa, N. Ohashi, T. Uchikoshi and F. Grasset, *Heliyon*, 2018, **4**, e00654.
- 33 Y. Shen, Y. Lou, Z. Wang and X. Xu, *Coatings*, 2017, **7**(12), 212–218.
- 34 H. Zhang, Z. Ji, T. Xia, H. Meng, C. Low-Kam, R. Liu, S. Pokhrel, S. Lin, X. Wang, Y. P. Liao, M. Wang, L. Li, R. Rallo, R. Damoiseaux, D. Telesca, L. Mädler, Y. Cohen, J. I. Zink and A. E. Nel, *ACS Nano*, 2012, **6**(5), 4349–4368.
- 35 K. S. Finnie, J. R. Bartlett, C. J. A. Barbé and L. Kong, *Langmuir*, 2007, **23**, 3017–3024.
- 36 F. Grasset, N. Labhsetwar, D. Li, D. C. Park, N. Saito, H. Haneda, O. Cador, T. Roisnel, S. Mornet, E. Duguet, J. Portier and J. Etourneau, *Langmuir*, 2002, **18**, 8209–8216.
- 37 F. J. Arriagada and K. Osseo-Asare, *Colloids Surf., A*, 1999, **154**, 311–326.
- 38 I. Danielsson and B. Lindman, *Colloids Surf., A*, 1981, **3**, 391–392.
- 39 A. Renaud, M. Wilmet, T. G. Truong, M. Seze, P. Lemoine, N. Dumait, C. Chen, N. Saito, T. Ohsawa, T. Uchikoshi, N. Ohashi, S. Cordier and F. Grasset, *J. Mater. Chem. C*, 2017, **5**, 8160–8168.
- 40 G. B. Smith, C. A. Deller, P. D. Swift, A. Gentle, P. D. Garrett and W. K. Fisher, *J. Nanopart. Res.*, 2002, **4**, 157–165.

