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A green luminescent MoS_2 – CdTe hybrid nanostructure synthesized through surface charge interaction†

Dhrubaa Haldar, Saptasree Bose,‡ Arnab Ghosh‡ and Shyamal K. Saha  *

During the last few years, intensive research has been carried out on the synthesis of different hybrid nanostructures mostly using hydrothermal and solvothermal techniques. But the fabrication of the hybrid nanostructure through surface charge interaction of the individual components is comparatively less explored. Here in this work, a hybrid nanostructure based on MoS_2 and CdTe quantum dots is synthesized through a simple surface charge interaction process using the negative surface charge of the excess sulfide ions (S^{2-}) present at the edge of the MoS_2 QDs and positively charged CdTe QDs where the positive surface charge was induced in CdTe by using a cysteamine ligand in acidic medium. In the photoluminescence (PL) emission spectrum, a new peak is observed which is different from those of both of the individual components. Interestingly, with increasing the concentration of CdTe QDs during the preparation of the hybrid nano-structure, the peak of hybrid QDs is gradually blue shifted towards the emission of MoS_2 QDs. The maximum blue shift occurs up to 1 : 1 (v/v) ratio of MoS_2 : CdTe as in this concentration ratio all S^{2-} ions are neutralized by $-\text{NH}_3^+$. This new emission occurs from a newly generated hybrid energy level. The energy level positions of the two different QDs along with the hybrid ones are estimated via cyclic voltammetry and absorption experiments.

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Introduction

Recently transition metal dichalcogenide (TMDC) quantum dots have been intensively studied because of their superior optical properties.^{1,2} Transition metal dichalcogenide materials are a new class of materials which are of vivid interest in the fields of semiconductor nanoscience and optoelectronics.^{3–6} Among metal dichalcogenides, MoS_2 quantum dots (QDs) have become a centre of attraction for researchers due to their wide variety of properties in the semiconductor industry for optoelectronic applications and energy harvesting. Contrary to bulk MoS_2 which possesses an indirect band gap of 1.2 eV, MoS_2 quantum dots show a direct band gap of 1.8 eV,⁷ giving rise to exotic physicochemical properties for versatile applications *viz.* electronics, optoelectronics, catalysis, energy storage, *etc.* Because of their excellent optical properties, superior chemical durability, biocompatibility and lower cytotoxicity, MoS_2 QD based systems have also been extensively used as bio-imaging materials.^{8,9} Apart from MoS_2 quantum dots, significant studies have also been performed on a new kind of semiconductor nanocrystal especially II–VI compounds such as CdS ,

CdSe , CdTe , *etc.* due to their emission in the visible range.^{10–12} Among these II–VI semiconductor compounds, CdTe has attracted great interest as it is a direct bandgap material with a bandgap of 1.52 eV which is suitable for emission in the visible spectral range by tuning the size from 1 to 10 nm, generally attributed to the quantum confinement effect.¹³ CdTe is also being used for photoelectric conversion in solar cells and light emitting diodes as well as in bio-applications.^{14–16} Moreover, CdTe has many unique electronic and optical properties, such as large absorption cross-section, excellent chemical and photo stability, size-dependent tunability of emission wavelength and narrow emission peak width at half maximum.

Besides these individual quantum dots like MoS_2 , CdTe , *etc.*, their hybrid structures are also being investigated during the last few years^{17–22} because of their potential applications *viz.* photocatalytic, electrocatalytic, photovoltaic, optoelectronic and sensing applications. In the literature, it is well reported that hybrid nanostructures are usually formed using a bottom-up approach and colloidal synthesis. However, to date, synthesis of the hybrid structure using surface charge is the most straightforward route. Although few reports on hybrid structures synthesised by surface charge interaction are already available,^{23–25} in most cases emission wavelengths of individual components in the hybrid structure do not change corresponding to their bare state. The emission peaks in the hybrid state almost remain in the same position as that of the individual quantum dots. In general, the emission peak of

School of Materials Sciences, Indian Association for the Cultivation of Science, 2A and 2B Raja S C Mullick Road, Jadavpur, Kolkata-700032, India. E-mail: cnsks@iacs.res.in

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‡ Authors contributed equally.



a quantum dot changes significantly depending upon the strength of the interaction with the environment (another quantum dot). Therefore, it is expected to achieve tunable optical properties in the hybrid structure by changing the size of the individual quantum dots and the strength of the interaction between the two quantum dots through surface charge modification. To explore this phenomenon, we have carried out experiments and evaluated whether the resultant emission arises due to the individual quantum dots in solution or a real hybridization of levels occurs between the quantum dots to give a new peak. In the present work, MoS_2 with negative surface charge and CdTe with positive surface charge have been synthesised to make a charge-induced hybrid structure MoS_2 – CdTe . It is already reported that amorphous MoS_2 QDs possess a large number of edge defects ('S' rich) and thus we have synthesised them to obtain quantum dots with negative surface charge.⁴ To synthesise positively charged CdTe QDs in an aqueous medium, the hydrophilic capping agent cysteamine is used in acidic medium.^{26,27} As a result, the surface charge of CdTe QDs becomes positive due to the presence of the $-\text{NH}_3^+$ moiety. The nature of surface charges of these two individual quantum dots and hybrid ones has been confirmed by zeta potential measurements. The formation of the hybrid structure is confirmed from TEM measurements, DLS measurements, XPS study, zeta potential, photoluminescence emission, UV visible spectroscopy and CV measurements. The tuning of the hybrid level by changing the concentration of CdTe QDs has also been investigated. An interesting fact is that, on increasing the concentration of CdTe , the peak has been blue shifted towards MoS_2 emission. However, above a certain MoS_2 and CdTe ratio (1 : 1), no further peak shift towards MoS_2 is observed; rather, the peak is shifted towards CdTe emission which confirms the formation of the hybrid level between MoS_2 and CdTe .

Moreover, CdTe QDs of three different colors (yellow, green and red) have been prepared to make a hybrid with tunable color. Band positions of the three different CdTe play a vital role in the formation of the hybrid structure. It has been observed that green and yellow CdTe successfully form a hybrid level with MoS_2 , whereas red CdTe cannot do the same. A detailed study has been carried out on the hybrid structure made from yellow CdTe and MoS_2 quantum dots. Cyclic Voltammetry (CV) and absorption measurements have been carried out to calculate the band positions of CdTe and MoS_2 quantum dots, and finally, the mechanism of formation of the hybrid structure and its spectroscopic investigation have been depicted successfully.

Experimental section

Materials

Ammonium tetrathiomolybdate (Sigma-Aldrich); hydrazine hydrate (N_2H_4 , Merck); sodium borohydride (NaBH_4 , Merck); tellurium powder (Te, Sigma-Aldrich); cadmium chloride hemipentahydrate ($\text{CdCl}_2 \cdot 2.5\text{H}_2\text{O}$); hydrochloric acid (HCl, Merck); cysteamine (Sigma-Aldrich), PUR-A-LYZER MEGA DIALYSIS KIT (Sigma-Aldrich); Milli-Q water.

Synthetic procedure

Synthesis of MoS_2 quantum dots. The QDs were synthesised through a bottom-up hydrothermal method by using $(\text{NH}_4)_2\text{MoS}_4$. In this procedure, $(\text{NH}_4)_2\text{MoS}_4$ was dispersed in 50 mL of water under ultrasonic vibration. Hydrazine hydrate was then added to the above solution for reduction, and the mixture was transferred into a 100 mL Teflon lined steel autoclave and treated hydrothermally at 200 °C for eight hours. After hydrothermal treatment, the solution was dialysed using a PUR-A-LYZER Mega Dialysis kit against Milli-Q water to remove all unreacted chemicals.

Synthesis of cysteamine capped CdTe quantum dots. Cysteamine capped CdTe QDs were synthesized using $\text{CdCl}_2 \cdot 2.5\text{H}_2\text{O}$ and NaHTe . In the first step, NaBH_4 and Te powder were taken in a round bottom flask and were mixed in Milli-Q water under an argon atmosphere to synthesize NaHTe . $\text{CdCl}_2 \cdot 2.5\text{H}_2\text{O}$ and cysteamine were dissolved in Milli-Q water. The pH of this solution was maintained at 5.5 using the required amount of HCl. The mixture was bubbled by Ar gas to make an O_2 free solution. Then the above prepared NaHTe was immediately injected into the reaction mixture under vigorous stirring and it was refluxed at 130 °C in an argon atmosphere. The size of the CdTe QDs can be adjusted by changing the growth time²⁴ (30, 50 and 90 min).

Synthesis of the hybrid nanostructure. Different hybrid nanostructures were synthesised by mixing the two oppositely charged quantum dots, varying the volumetric ratio of $\text{CdTe} : \text{MoS}_2$ where the volume of the MoS_2 quantum dot solution was kept constant. Simple electrostatic interactions between the two oppositely charged nanoparticles have been used to prepare the hybrid structure.

Characterization

Instrumental details. To characterize the MoS_2 quantum dots, CdTe quantum dots and their hybrid structure, X-ray diffraction (XRD) measurements are carried out on a thick film using an X-ray diffractometer (RICHSEIFERT-XRD 3000P with X-ray Generator-Cu, 10 kV, ten mA, wavelength 1.5418 Å). Microstructural studies are performed using a high-resolution transmission electron microscope (HRTEM). We use a high-resolution transmission electron microscope (HRTEM; JEOL 2100) for microstructural and EDX analysis. The DLS and zeta potential measurements were carried out using a zetasizer nanosystem (Malvern Instruments Ltd). UV-VIS absorption spectra are measured with a Cary UV 5000 spectrophotometer. All the photoluminescence (PL) measurements are performed using a PTI fluoromax QM-400 spectrofluorometer. For the time-correlated single-photon counting (TCSPC) measurements, the samples are excited at 280 and 375 nm using a pico second NANO-LED IBH 280L and 370L. Cyclic voltammetric measurements are carried out under a nitrogen atmosphere using a Ag/Ag^+ reference electrode, with a Pt disc as the working electrode and a Pt wire as the auxiliary electrode in a PC-controlled PAR model 273A electrochemistry system using 0.1 M KCl as the electrolyte.



Results and discussion

HRTEM, XRD, DLS and zeta potential of MoS₂

TEM measurements have been performed to confirm the structural properties of the as-synthesized MoS₂ quantum dots. Fig. 1(a) shows the distribution of the quantum dots. The size distribution plot of the synthesized quantum dots is given in the ESI (Fig. S1†). The SAED pattern (Fig. 1(b)) confirms the amorphous nature of the synthesized QDs. The size distribution of the synthesized amorphous MoS₂ QDs is almost 18 nm (shown in Fig. 1(c), from DLS measurements). From the EDX analysis of HRTEM images it is seen that the MoS₂ QDs are rich in excess 'S' which is attributed to the amorphous nature of the QDs. The EDX analysis in Fig. 1(d) also confirms the formation of molybdenum sulfide quantum dots with a high S/Mo atomic ratio greater than two (~2.77).²⁸ The surface charge of MoS₂ quantum dots is also determined by zeta potential analysis as shown in Fig. 1(e). We have found the zeta potential value of bare MoS₂ quantum dots to be negative, -40.8 mV, due to the presence of electronically rich sulfur atoms in the edge of the QDs. To obtain further structural information of amorphous MoS₂, we have performed XRD analysis. Due to the presence of a large excess of sulfur defects ('S' rich) as shown in HRTEM and zeta potential measurements, we could not find any XRD peak; rather, a broad peak is found corresponding to the amorphous nature of MoS₂ QDs, as shown in Fig. 1(f).

HRTEM, XRD, DLS and zeta potential of CdTe quantum dots

The transmission electron microscope (TEM) image shows different sizes of CdTe quantum dots (Fig. 2). The size distribution plot of different sized quantum dots is given in the ESI (Fig. S2–S4†). The size distribution of yellow emissive CdTe QDs is almost 4 nm. The highly ordered lattice fringes that appeared in high-resolution TEM (HRTEM) images suggest the formation of crystalline CdTe quantum dots as shown in Fig. 3(a). The lattice *d*-spacing of these quantum dots is found to be 1.41 Å which corresponds to the (200) plane of the crystalline CdTe phase. The corresponding FFT image is also given in the inset of Fig. 3(a). Energy dispersive X-ray spectroscopy (EDX) analysis (given in the ESI, Fig. S5†) also indicates the presence of Cd, Te, C, N and S atoms (the atomic % of carbon is greater due to the usage of the carbon grid). The crystal structure of the as-synthesized cysteamine capped QDs is systematically investigated by X-ray powder diffraction (XRD). From Fig. 3(b), it is seen that the spectra contain some prominent diffraction peaks at 2θ values of 22.3°, 30° and 37.6° which are assigned to the (111), (200) and (220) planes of crystalline CdTe QDs.^{29–32} All the diffraction peaks are confirmed with JCPDS card no. 24-0513. The surface charge of cysteamine capped CdTe quantum dots is also determined by zeta potential analysis (Fig. 3(c)). The zeta potential value of CdTe QDs is found to be positive, *i.e.* +23.4 mV, due to the presence of the -NH₃⁺ moiety of the cysteamine ligand in acidic medium at the outer surface of the QDs. The DLS study also shows the size distribution of yellow CdTe (Fig. 3(d)).

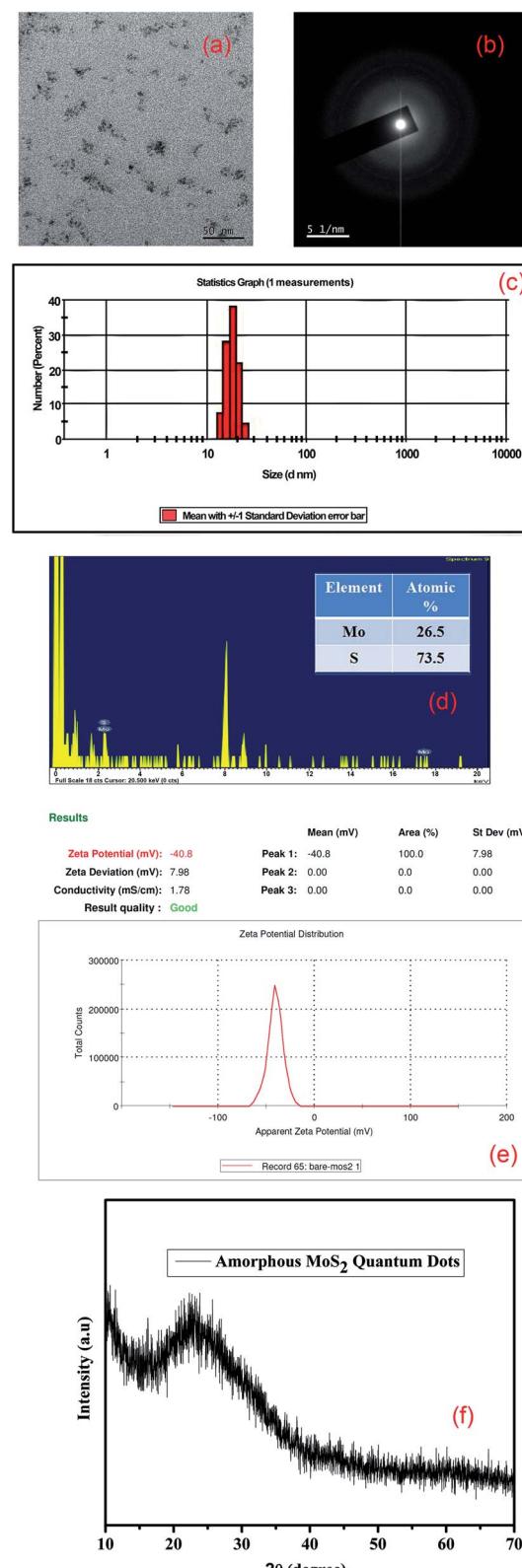


Fig. 1 (a) TEM image of the as-synthesized MoS₂ QDs; (b) SAED pattern of the as-synthesized amorphous MoS₂ QDs; (c) particle size distribution of MoS₂ QDs from DLS measurements; (d) EDX analysis of sulfur-rich MoS₂ QDs; (e) zeta potential of sulfur rich MoS₂ QDs; (f) XRD pattern of the as-synthesized amorphous MoS₂ QDs.



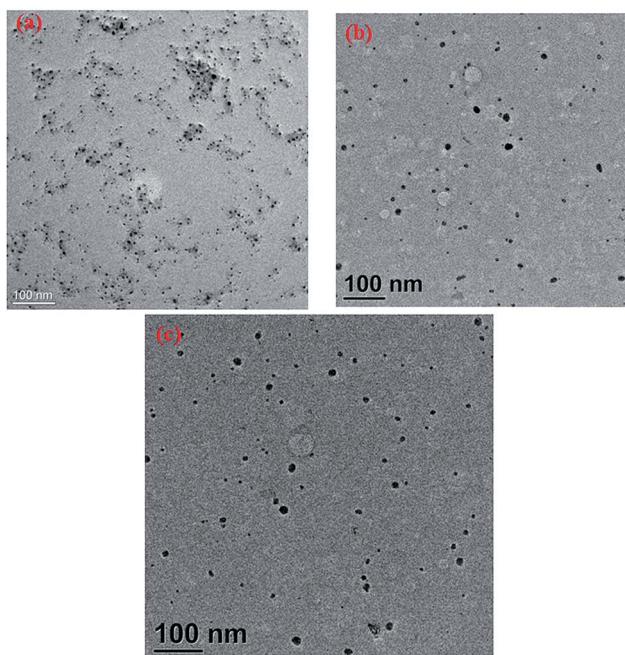


Fig. 2 (a) TEM images of the as-synthesized green emissive CdTe QDs; (b) yellow emissive CdTe QDs; (c) red emissive CdTe QDs.

Hybrid quantum dots

The surface charges of MoS_2 and CdTe are found to be -40.8 mV and $+23.4$ mV respectively from the zeta potential experiment and it is expected that the hybrid structure of the two different QDs is formed through electrostatic interaction when cysteamine capped CdTe QDs are added to the MoS_2 QDs. The attachment of the two QDs is confirmed from zeta potential, DLS, corresponding FFT of the hybrid TEM structure and XPS study. The zeta potential of the hybrid structure is found to be -5 mV (shown in Fig. 4(a)) indicating the formation of the hybrid structure. The TEM image of the hybrid structure is shown in Fig. 4(b). As the surfaces of CdTe and MoS_2 QDs are totally oppositely charged, each CdTe is surrounded by some MoS_2 QDs and each MoS_2 QD is surrounded by some CdTe QDs. As a result, a network type structure is formed (Fig. 4(b)). From the FFT image (Fig. 4(c)), presence of crystalline CdTe as well as amorphous MoS_2 is also found. From the DLS measurements it is observed that the maximum size of MoS_2 QDs is 18 nm, whereas that of yellow emissive CdTe QDs is 4 nm. When CdTe QDs are added to the MoS_2 QDs attached nanoparticles are formed with a size of ~ 700 nm (shown in Fig. 4(d)) which confirms the network kind structure found in HRTEM images. Moreover, from the shifting of the binding energy of Mo^{4+} , S^{2-} in the hybrid structure also confirms the formation of the hybrid structure (Fig. 5). The peaks for Mo^{4+} - $3\text{d}_{5/2}$, $3\text{d}_{3/2}$, S^{2-} - 2s , S^{2-} - $2\text{p}_{3/2}$, S^{2-} - $2\text{p}_{1/2}$, S_2^{2-} - $2\text{p}_{3/2}$ and S_2^{2-} - $2\text{p}_{1/2}$ in bare amorphous MoS_2 QDs are found at 228 eV, 232.1 eV, 225.3 eV, 161.4 eV, 162.8 eV, 163.7 eV and 164.3 eV respectively (given in S6†), whereas the peaks of Mo^{4+} - $3\text{d}_{5/2}$, $3\text{d}_{3/2}$, S^{2-} - 2s , S^{2-} - $2\text{p}_{3/2}$, S^{2-} - $2\text{p}_{1/2}$ and S_2^{2-} - $2\text{p}_{1/2}$ in the hybrid structure appear at 227.1 eV (0.9 eV), 231.6 eV (0.5 eV), 225 eV (0.3 eV), 160.1 eV (1.3

eV), 160.7 eV (2.1 eV), 161.8 eV (1.9 eV) and 163.5 eV (0.8 eV) respectively (the amount of the shifting of the binding energy value is given in brackets) as shown in Fig. 5(a) and (b). The peaks of nitrogen in $-\text{NH}_3^+$, $\text{Cd-3d}_{5/2}$ and $\text{Cd-3d}_{3/2}$ in the hybrid one appear at 400.7 eV, 404.5 eV and 411.3 eV respectively as shown in Fig. 5(c). $\text{Te-3d}_{5/2}$ and $3\text{d}_{3/2}$ arise at 572.4 eV and 575.4 eV and 582.7 eV and 585.9 eV respectively as shown in Fig. 5(d). The higher binding energies, *i.e.* 575.4 eV and 585.9 eV, confirm the thiolation of the CdTe surface.³³ The shifting of the peaks for Mo^{4+} , S^{2-} and S_2^{2-} towards lower binding energy in the hybrid structure compared to the bare amorphous MoS_2 QDs confirms the attachment^{2,34} between S^{2-} and $-\text{NH}_3^+$.

Optical characterization

MoS₂ quantum dots. Fig. 6 illustrates the optical properties of the as-synthesized QDs. Fig. 6(a) shows the excitation dependent photoluminescence^{35,36} (PL) of MoS_2 QDs, in which the maximum PL intensity is obtained at 392 nm under 320 nm excitation. Due to the presence of localized defect states ("S" defects), there is a possibility of vibrational relaxation from the conduction band to these localized states, resulting in broadening of the emission spectra. An actual photograph of the blue emission from the MoS_2 quantum dots under UV light is shown in the inset of Fig. 6(b). To determine the excited state decay time of the band gap emission, time-resolved emission of the quantum dots is studied, and the spectrum is presented in Fig. 6(b). The excited state decay time of the band gap emission is determined from Fig. 6(b) and found to be 2.2 ns.

CdTe quantum dots. Fig. 7 shows the excitation dependent PL for three different colored CdTe QDs. Green emissive QDs show the maximum PL intensity at 555 nm under 390 nm excitation, yellow emissive QDs show the maximum PL intensity at 572 nm under 470 nm excitation, and for red QDs, it appears at 665 nm under 470 nm excitation.

Hybrid QDs. Photoluminescence emission spectra of different hybrid structures composed of yellow CdTe QDs and blue MoS_2 QDs are shown in Fig. 8. It is observed that both the peaks of MoS_2 and CdTe are present in the PL spectra of all the hybrid structures, where the MoS_2 emission peak has very low intensity due to its amorphous nature, whereas the CdTe emission peak is very prominent. There is almost no peak shift for the MoS_2 , whereas the peak of CdTe is gradually shifted to lower wavelength (towards the MoS_2 peak) with increasing the CdTe concentration in the hybrid structure (Fig. 8(a)-(d)).

MoS_2 has a large number of edge sulfur atoms which are electron rich, whereas CdTe quantum dots have been prepared using cysteamine as the capping agent which induces a positive charge on the CdTe surface having $-\text{NH}_3^+$ at the edge. As a result, when $-\text{NH}_3^+$ capped CdTe gets closer to electron rich MoS_2 quantum dots, they form hybrid structure through simple electrostatic interaction. As the CdTe concentration increases keeping the MoS_2 concentration constant, S^{2-} ions present in MoS_2 get gradually neutralized by $-\text{NH}_3^+$ present in CdTe; a new hybrid nanostructure is proposed to form which has a definite



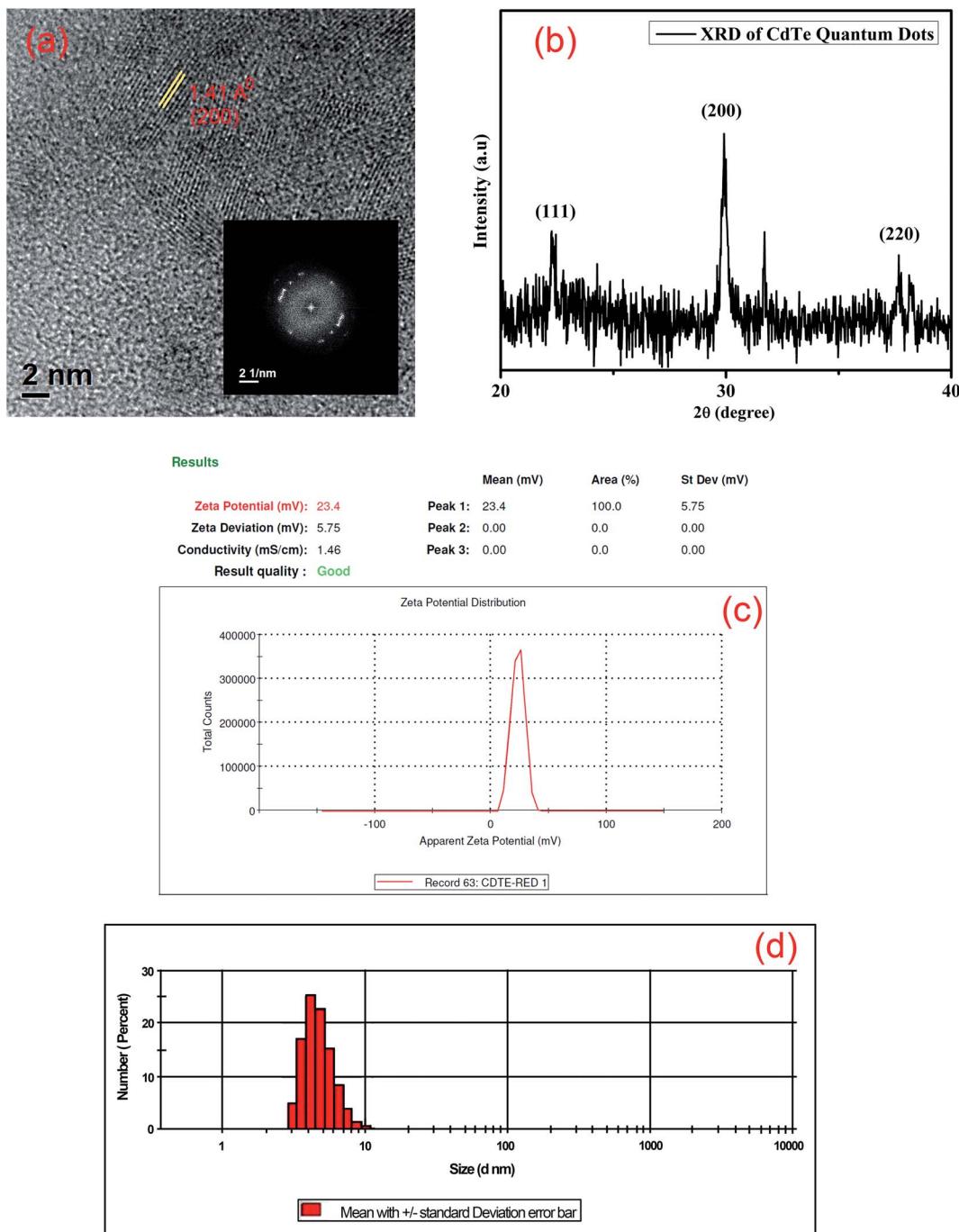


Fig. 3 (a) High-resolution TEM image and corresponding FFT diagram of CdTe QDs; (b) XRD patterns of CdTe QDs; (c) zeta potential of cysteamine capped CdTe QDs; (d) particle size distribution of CdTe QDs.

set of energy levels, giving rise to new emission peaks as observed for each set of combinations. It is interesting to note that these new emission peaks of the hybrid QDs are at a significantly different position (530 nm for set-IV) and also with a completely different excitation wavelength (360 nm) in comparison to those of pristine CdTe (emission – 572 nm and excitation – 470 nm) and MoS₂ QDs (emission – 392 nm and excitation – 320 nm). This new emission is gradually shifted to lower wavelength by increasing the concentration of CdTe. After

reaching a certain concentration of CdTe where MoS₂ : CdTe is 1 : 1, maximum peak shifting (42 nm, Fig. 8(d)) occurs, indicating complete neutralisation of S²⁻ by NH₃⁺. Beyond this concentration of CdTe, the photoluminescence peak of the hybrid structure again shifts to higher wavelength up to 566 nm (Fig. 8(e)–(h)) which is almost the characteristic emission of yellow CdTe, no further peak shifting occurs beyond this concentration and these results indicate the formation of the hybrid nanostructure. From Time Correlated Single Photon



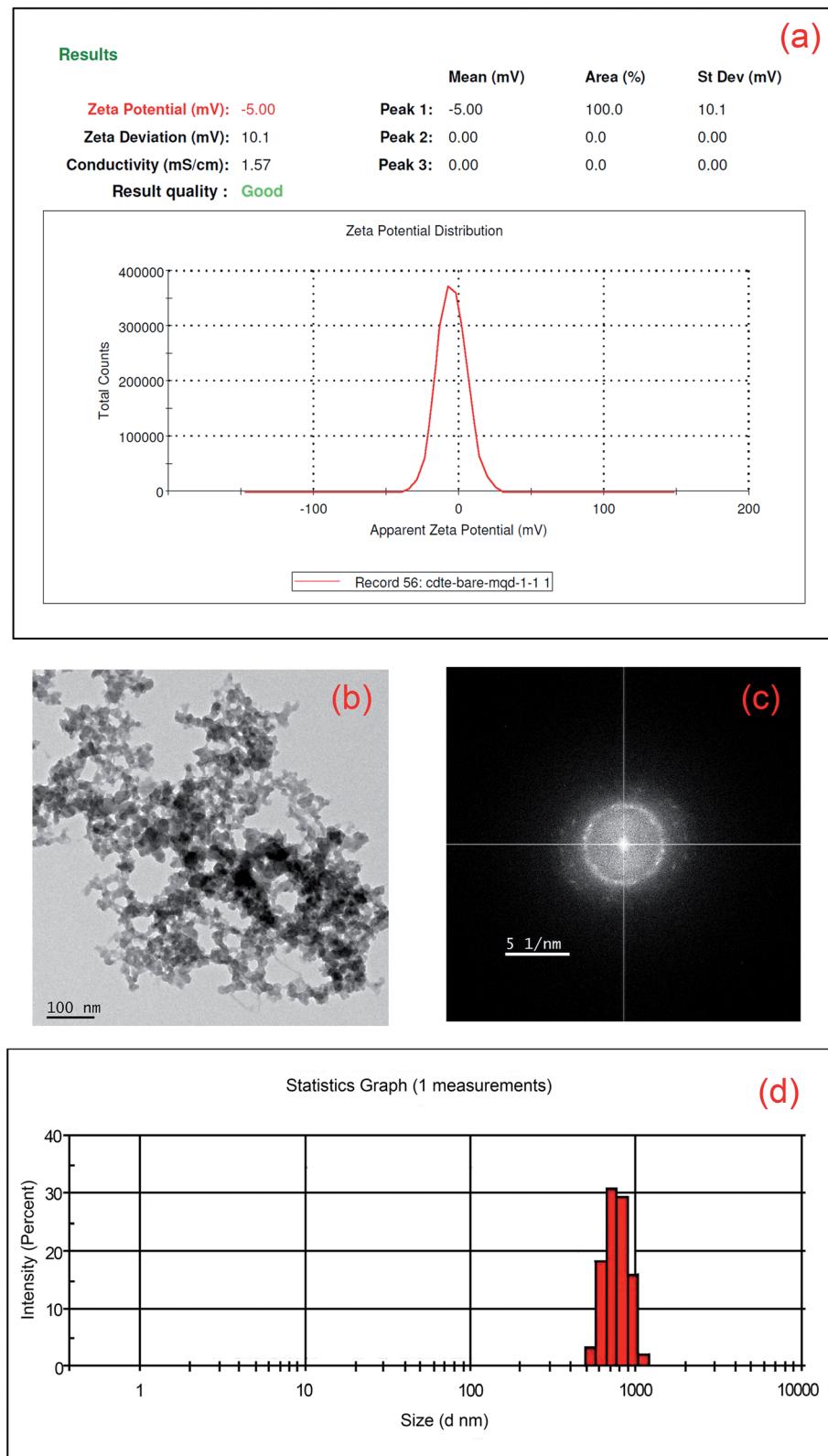


Fig. 4 (a) Zeta potential of MoS_2 –CdTe hybrid QDs; (b) high-resolution TEM image of MoS_2 –CdTe hybrid QDs and (c) corresponding FFT diagram of the hybrid structure; (d) particle size distribution of MoS_2 –CdTe hybrid QDs.

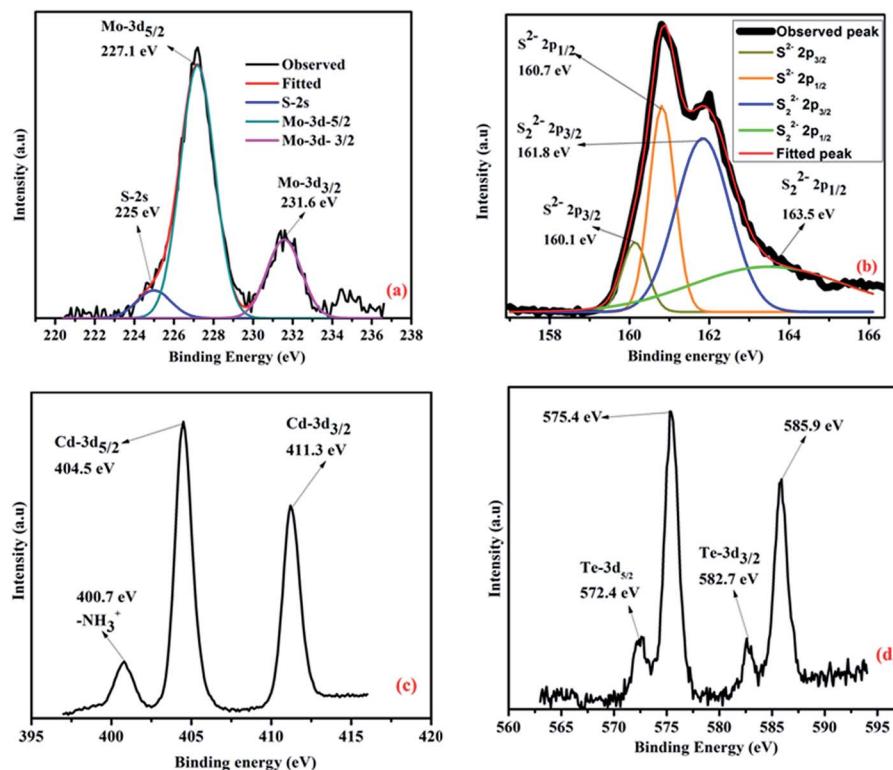


Fig. 5 (a) XPS spectra of S-2s, Mo(iv)-3d_{5/2} and 3d_{3/2} in hybrid QDs; (b) deconvoluted XPS spectra of S²⁻-2p_{3/2} and 2p_{1/2} and S₂²⁻-2p_{3/2} and 2p_{1/2} in hybrid QDs, (c) XPS spectra of N-1s in -NH₃⁺ attached with the CdTe surface and Cd-3d_{5/2} and Cd-3d_{3/2} of hybrid QDs; (d) XPS spectra of Te²⁻-3d_{5/2} and 3d_{3/2} in hybrid QDs.

Counting (TCSPC) experiments (Fig. 9), the average life time of the hybrid structure is found to be 16.3 ns.

From the photoluminescence emission spectra of the hybrid structures it is shown that the maximum intensity is observed at 530 nm under 360 nm excitation for 1 : 1 composition of the hybrid structure (Fig. 8(d)) which is not the excitation wavelength of pristine yellow CdTe QDs as the maximum PL intensity for cysteamine capped yellow CdTe QDs appears under 470 nm excitation (Fig. 7(b)). Therefore, it is confirmed that the emission peak of the hybrid material does not correspond to the emission peak of CdTe or MoS₂ but it is

a new emission peak which is proposed to originate due to the modulation of the energy band structure by surface charge interaction between MoS₂ and CdTe QDs. Apart from photoluminescence, to understand the band positions of MoS₂, CdTe and the hybrid structure, cyclic voltammetry (CV) and UV-absorption measurements have been carried out for the 1 : 1 composition (Fig. S7-S12†). UV peaks for MoS₂, CdTe and hybrid structures appear at 350 nm, 508 nm and 358 nm respectively. CV measurements have been performed under an inert atmosphere using three electrode systems at room temperature.

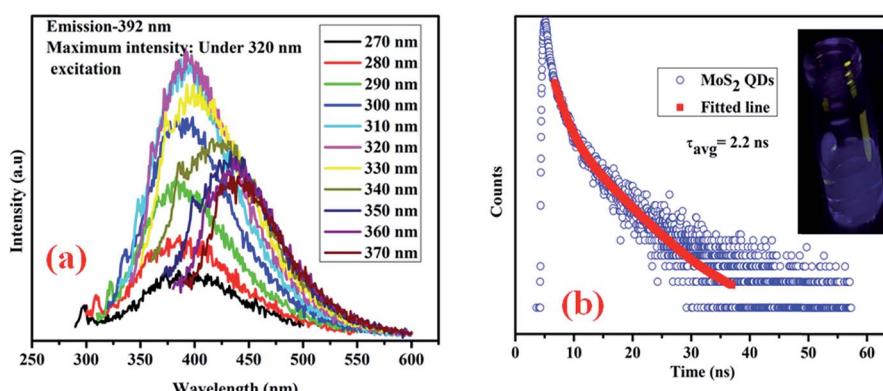


Fig. 6 (a) Excitation dependent photoluminescence emission spectra of bare MoS₂ QDs; (b) time-correlated single photon counting (TCSPC) of the as-synthesized MoS₂ QDs and an actual photograph of the emission of bare MoS₂ QDs under UV light (inset).

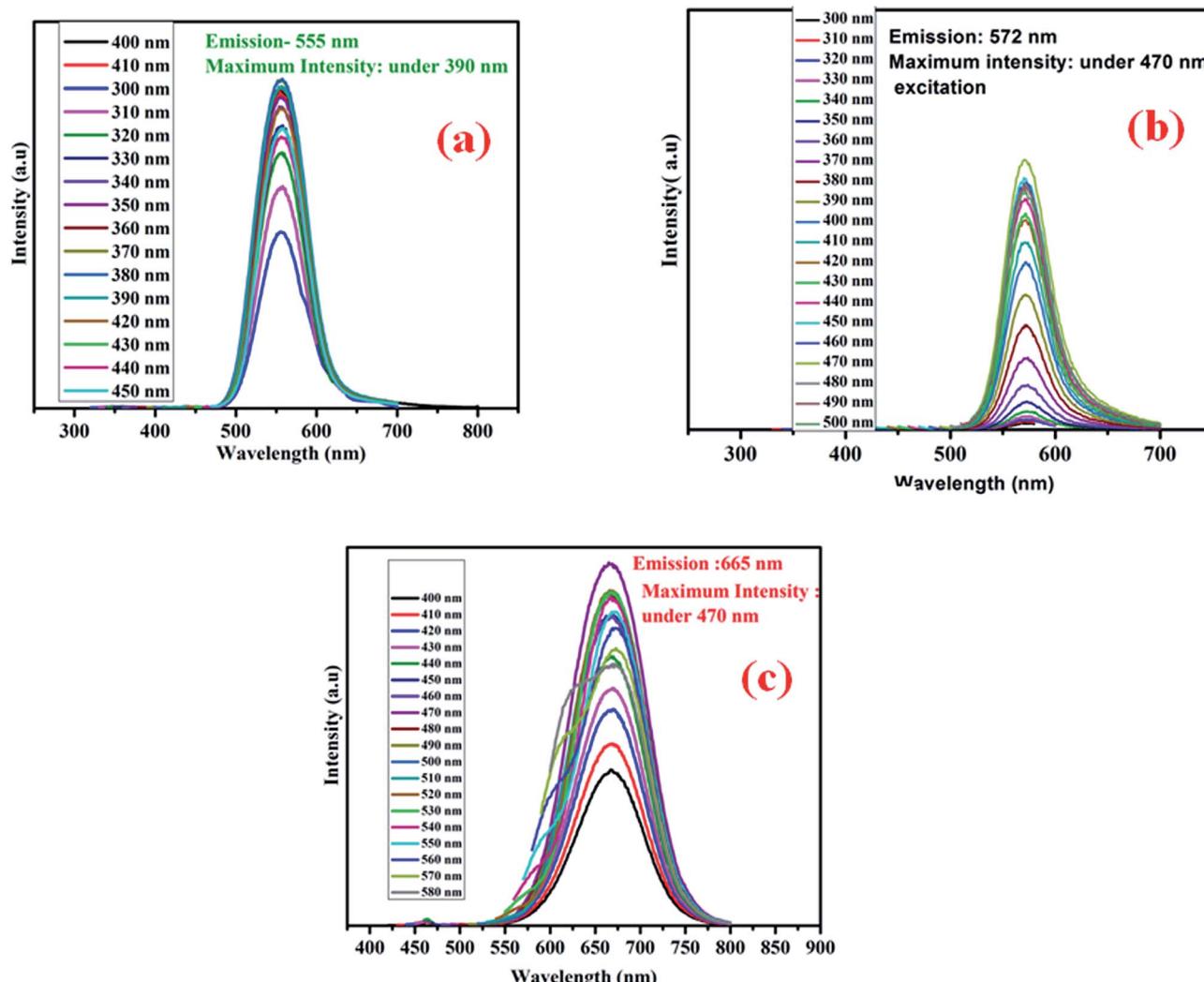


Fig. 7 Excitation dependent photoluminescence emission of (a) green, (b) yellow and (c) red CdTe QDs.

The band positions of pristine MoS_2 and CdTe quantum dots and the hybrid MoS_2 -CdTe nanostructure have been calculated using the following equations:²³

$$E_{\text{VB}} = -(E_{\text{ox}} + 4.71) \text{ eV}$$

$$E_{\text{CB}} = -(E_{\text{red}} + 4.71) \text{ eV}$$

$$E_{\text{VB}} = E_{\text{CB}} - \Delta E$$

where the potential values are in reference to the Ag/Ag^+ reference electrode and ΔE is the energy difference between the valence band (VB) and the conduction band (CB) measured from their corresponding UV absorption spectra.

From these equations, the valence band and the conduction band of MoS_2 are found at -7.24 eV and -3.73 eV respectively, whereas those of CdTe appear at -6.22 eV and -3.82 eV.

Therefore, it is seen that the conduction bands of the two individual quantum dots are very close to each other. Near the conduction band of MoS_2 , there are large numbers of localized

states due to the presence of excess S^{2-} . When both MoS_2 and CdTe come nearby, there is a possibility for the formation of a hybrid nanostructure due to the electrostatic interaction between S^{2-} and $-\text{NH}_3^+$. As a result, a new set of energy levels have been formed for the hybrid structure. The schematic diagram of the predicted energy levels is given in Scheme 1. The valence band (at -7.27 eV) is formed which is very near to the MoS_2 valence band (-7.24 eV), whereas the conduction band (at -3.81 eV) is formed between the two conduction bands of MoS_2 and CdTe and it is very near to the CdTe conduction band (at -3.82 eV). The band diagram is depicted in Scheme 1. With increasing the extent of the volume of CdTe, the edge sulfurs S^{2-} present in the MoS_2 QDs are successively get neutralized by $-\text{NH}_3^+$. So as a result with increasing concentration of CdTe, gradual peak shifting occurs to the lower wavelength region until there is an equal volume of MoS_2 and CdTe. After that when the volume of CdTe is again increased, no more S^{2-} remains to be neutralized. Therefore, now the system possesses the character of CdTe only, and as a result, the emission peak becomes red shifted, up to 566 nm (almost characteristic

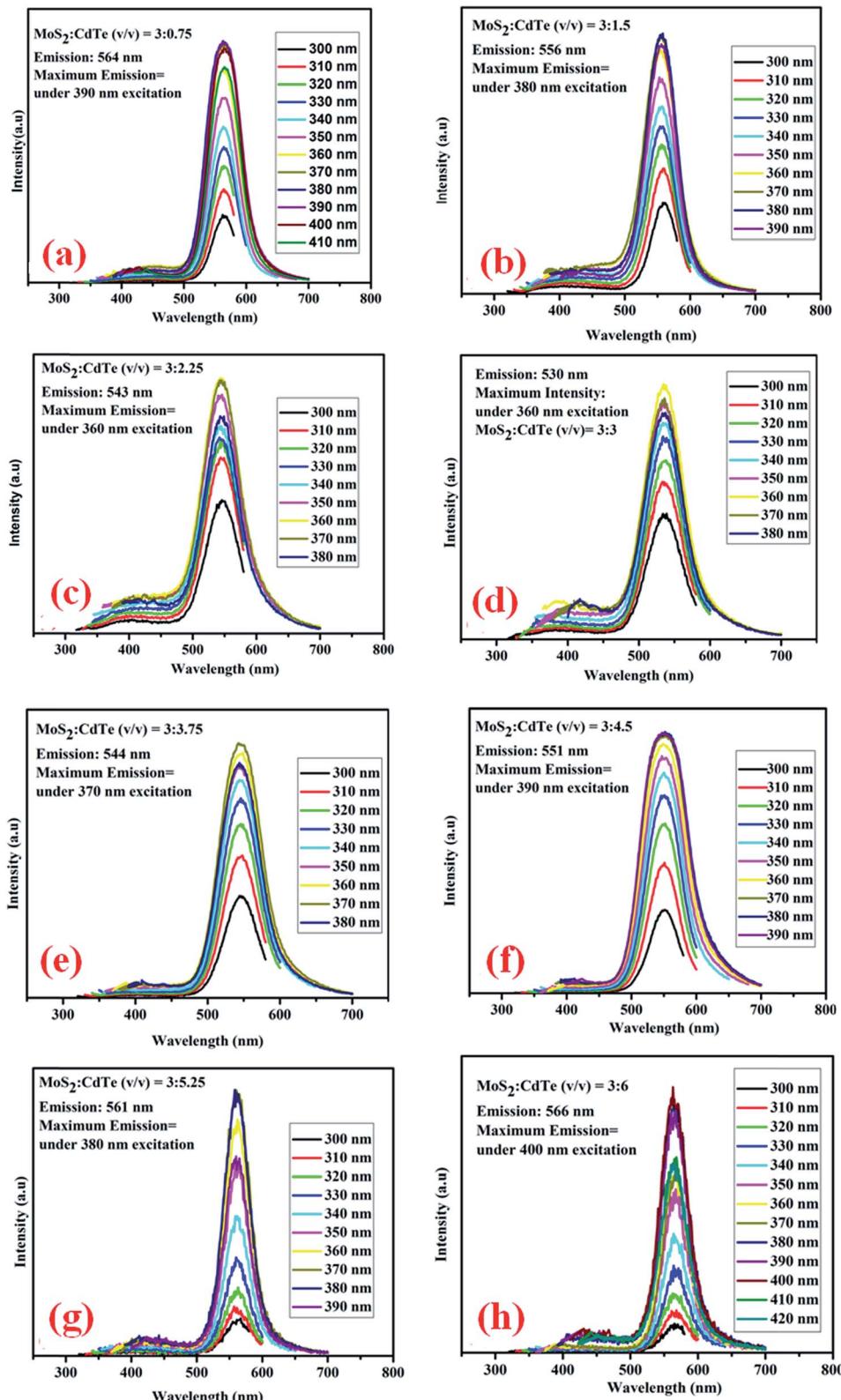


Fig. 8 (a) Excitation dependent photoluminescence emission spectra of $\text{MoS}_2 : \text{CdTe} = 3 : 0.75$ (v/v); (b) $\text{MoS}_2 : \text{CdTe} = 3 : 1.5$ (v/v); (c) $\text{MoS}_2 : \text{CdTe} = 3 : 2.25$ (v/v); (d) $\text{MoS}_2 : \text{CdTe} = 3 : 3$ (v/v). (e) Excitation dependent photoluminescence emission spectra of $\text{MoS}_2 : \text{CdTe} = 3 : 3.75$ (v/v); (f) $\text{MoS}_2 : \text{CdTe} = 3 : 4.5$ (v/v); (g) $\text{MoS}_2 : \text{CdTe} = 3 : 5.25$ (v/v); (h) $\text{MoS}_2 : \text{CdTe} = 3 : 6$ (v/v).

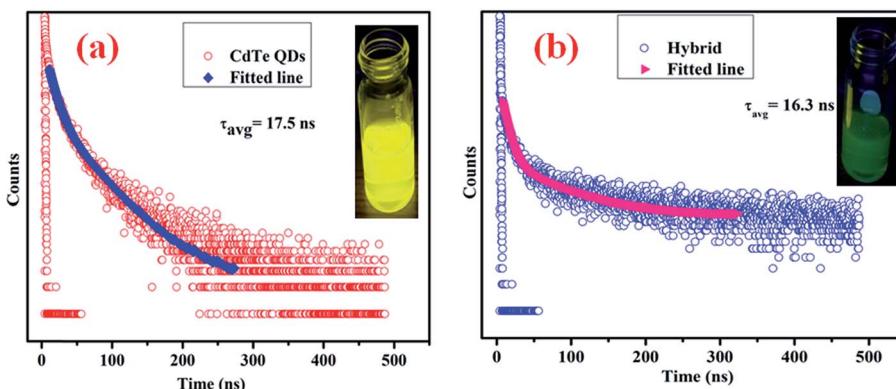
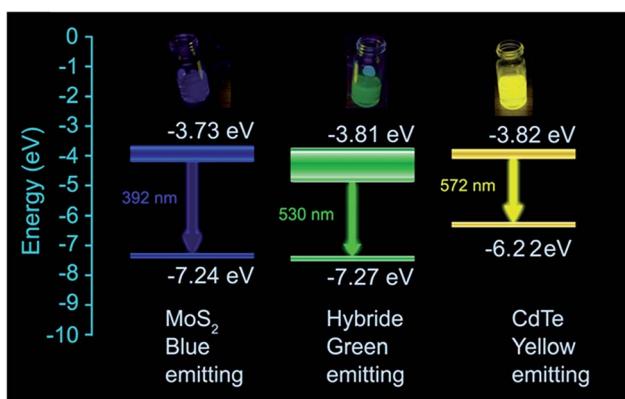


Fig. 9 (a) Excited state decay profile of cysteamine capped CdTe QDs and an actual photograph of the emission of CdTe QDs under UV light (inset); (b) excited state decay profile of the hybrid structure and an actual photograph of the emission of hybrid QDs under UV light (inset).



Scheme 1 Band diagram of MoS₂ QDs, capped yellow emissive CdTe QDs and the formed hybrid structure having 1 : 1 volumetric ratio of individual QDs.

emission of CdTe). So, for the formation of hybrid levels, energy matching of two conduction bands of individual QDs is also necessary. In the case of green CdTe QDs blue shifting also occurs in case of hybridization *viz.* from 555 nm to 525 nm, whereas due to the energy mismatching of the two conduction levels of red CdTe and MoS₂, no peak shifting occurs (Fig. S13 and S14†). Relative fluorescence quantum yields of MoS₂, CdTe and hybrid materials are 14.77%, 76.08% and 63.12% respectively from which it is found that the quantum yield of the hybrid materials is much greater than that of MoS₂ QDs and comparable to that of CdTe (Fig. S15†).

Conclusion

A new kind of hybrid nano-structure is synthesized *via* surface charge interaction of two oppositely charged QDs. Here we have used negatively charged MoS₂ QDs and positively charged CdTe QDs. As a result, a new emission peak is observed for the hybrid nano-structure which is different from those of both of the individual QDs. Moreover the excitation wavelength for the hybrid nanostructure is found to be 360 nm which is significantly different to the excitation wavelengths of the individual

MoS₂ (320 nm) and CdTe QDs (470 nm). This is explained as the formation of a new set of hybrid energy levels from the two QDs during the preparation of the hybrid nanostructure.

Conflicts of interest

There is no conflict of interest to declare.

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References

- 1 D. Haldar, D. Dinda and S. K. Saha, *J. Mater. Chem. C*, 2016, **4**, 6321.
- 2 D. Haldar, A. Ghosh, S. Bose, S. Mondal, U. K. Ghorai and S. K. Saha, *Opt. Mater.*, 2018, **79**, 12.
- 3 Z. Lei, W. Zhu, S. Xu, J. Ding, J. Wan and P. Wu, *ACS Appl. Mater. Interfaces*, 2016, **8**, 20900.
- 4 W. Gu, Y. Yan, C. Zhang, C. Ding and Y. Xian, *ACS Appl. Mater. Interfaces*, 2016, **8**, 11272.
- 5 W. Dai, H. Dong, B. Fugetsu, Y. Cao, H. Lu, X. Ma and X. Zhang, *Small*, 2015, **11**, 4158.
- 6 Y. Yan, C. Zhang, W. Gu, C. Ding, X. Li and Y. Xian, *J. Phys. Chem. C*, 2016, **120**, 12170.
- 7 A. Splendiani, L. Sun, Y. Zhang, T. Li, J. Kim, C. Y. Chim, G. Galli and F. Wang, *Nano Lett.*, 2010, **10**, 1271.
- 8 J. Wang, X. Tan, X. Pang, L. Liu, F. Tan and N. Li, *ACS Appl. Mater. Interfaces*, 2016, **8**, 24331.
- 9 S. Xu, D. Li and P. Wu, *Adv. Funct. Mater.*, 2015, **25**, 1127.



- 10 A. K. Bansal, F. Antolini, S. Zhang, L. Stroea, L. Ortolani, M. Lanzi, E. Serra, S. Allard, U. Scherf and I. D. W. Samuel, *J. Phys. Chem. C*, 2016, **120**, 1871.
- 11 K. S. Leschkes, R. Divakar, J. Basu, E. E. Pommer, J. E. Boercker, C. B. Carter, U. R. Kortshagen, D. J. Norris and E. S. Aydil, *Nano Lett.*, 2007, **7**, 1793.
- 12 S. J. Cho, D. Maysinger, M. Jain, B. Roder, S. HackbARTH and F. M. Winnik, *Langmuir*, 2007, **23**, 1974.
- 13 T. Wang and X. Ziang, *ACS Appl. Mater. Interfaces*, 2013, **5**, 1190.
- 14 J. Yang and X. Zhong, *J. Mater. Chem. A*, 2016, **4**, 16553.
- 15 Y. Zheng, S. Gao and J. Y. Ying, *Adv. Mater.*, 2007, **19**, 376.
- 16 X. Liu, H. Jiang, J. Lei and H. Ju, *Anal. Chem.*, 2007, **79**, 8055.
- 17 R. M. Clark, B. J. Carey, T. Daeneke, P. Atkin, M. Bhaskaran, K. Latham, I. S. Cole and K. K. Zadeh, *Nanoscale*, 2015, **7**, 16763.
- 18 P. T. Gomathi, P. Sahatiya and S. Badhulika, *Adv. Funct. Mater.*, 2017, **27**, 1701611.
- 19 D. Kufer, I. Nikitsky, T. Lasanta, G. Navickaite, F. H. L. koppens and G. Konstantatos, *Adv. Mater.*, 2015, **27**, 176.
- 20 X. L. Yin, L. L. Li, W. J. Jiang, Y. Zhang, X. Zhang, L. J. Wan and J. S. Hu, *ACS Appl. Mater. Interfaces*, 2016, **8**, 15258.
- 21 Y. Liu, Y. X. Yu and W. D. Zhang, *J. Phys. Chem. C*, 2013, **117**, 12949.
- 22 S. Hong, D. P. Kumar, E. H. Kim, H. Park, M. Gopannagari, D. A. Reddy and T. K. Kim, *J. Mater. Chem. A*, 2017, **5**, 20851.
- 23 M. K. Barman, P. Mitra, R. Bera, S. Das, B. Paramanik and A. Patra, *Nanoscale*, 2017, **9**, 6791.
- 24 S. Kundu, S. Sadhu, R. Bera, B. Paramanik and A. Patra, *J. Phys. Chem. C*, 2013, **117**, 23987.
- 25 R. Bera, S. Mandal, B. Mondal, B. Jana, S. K. Nayak and A. Patra, *ACS Sustainable Chem. Eng.*, 2016, **4**, 1562.
- 26 J. Li, T. Yang, W. H. Chan, M. M. F. Choi and D. Zhao, *J. Phys. Chem. C*, 2013, **117**, 19175.
- 27 F. Shi, S. Liu and X. Su, *New J. Chem.*, 2017, **41**, 4138.
- 28 D. Dinda, M. E. Ahmed, S. Mandal, B. Mondal and S. K. Saha, *J. Mater. Chem. A*, 2016, **4**, 15486.
- 29 Y. Wang and S. Liu, *J. Chil. Chem. Soc.*, 2012, **57**, 1109.
- 30 W. H. Yang, W. W. Li, H. J. Dou and K. Sun, *Mater. Lett.*, 2008, **62**, 2564.
- 31 F. O. Silva, M. S. Carvalho, R. Mendonca, W. A. Macedo, K. Balzuweit, P. Reiss and M. A. Schiavon, *Nanoscale Res. Lett.*, 2012, **7**, 536.
- 32 R. Wang, B. Li, L. Dong, F. Zhang, M. Fan and L. Zhou, *Mater. Lett.*, 2014, **135**, 99.
- 33 Y. Wang, R. Hu, G. Lin, W. C. Law and K. T. Yong, *RSC Adv.*, 2013, **3**, 8899.
- 34 B. K. Gupta, P. Thanikaivelan, T. N. Narayanan, L. Song, W. Gao, T. Hayashi, A. L. M. Reddy, A. Saha, V. Shanker, M. Endo, A. A. Marti and P. M. Ajayan, *Nano Lett.*, 2011, **11**, 5227.
- 35 H. Dong, S. Tang, Y. Hao, H. Yu, W. Dai, G. Zhao, Y. Cao, H. Lu, X. Zhang and H. Zu, *ACS Appl. Mater. Interfaces*, 2016, **8**, 3107.
- 36 H. Lin, C. Wang, J. Wu, Z. Xu, Y. Huang and C. Zhang, *New J. Chem.*, 2015, **39**, 8492.

