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Engineering efficient upconverting nanothermometers using Eu^{3+} ions[†]

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Upconversion nanothermometry combines the possibility of optically sensing temperatures in very small areas, such as microfluidic channels or on microelectronic chips, with a simple detection setup in the visible spectral range and reduced heat transfer after near-infrared (NIR) excitation. We propose a ratiometric strategy based on Eu^{3+} ion luminescence activated through upconversion processes. Yb^{3+} ions act as a sensitizer in the NIR region (980 nm), and energy is transferred to Tm^{3+} ions that in turn excite Eu^{3+} ions whose luminescence is shown to be thermally sensitive. Tridoped $\text{SrF}_2\text{:Yb}^{3+},\text{Tm}^{3+},\text{Eu}^{3+}$ nanoparticles (average size of 17 nm) show a relative thermal sensitivity of $1.1\% \text{ K}^{-1}$ at 25.0°C , in the range of the best ones reported to date for Ln^{3+} -based nanothermometers based on upconversion emission. The present nanoparticle design allows us to exploit upconversion of lanthanide ions that otherwise cannot be directly excited upon NIR excitation and that may provide operational wavelengths with a highly stable read out to fill the spectral gaps currently existing in upconversion-based nanothermometry.

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1. Introduction

The development of all-optical nanoparticle-based thermometers allows for the measurement of localized temperature with a high spatial resolution in sub-millimeter areas, which can serve as a tool for the characterization of microfluidic channels or electronic microcircuit surfaces.^{1–3} Particularly interesting are optical luminescent nanothermometers, that, once excited in their absorption region, exhibit a temperature-dependent emission, usually in the ultraviolet (UV), visible or near infrared (NIR) regions. Several parameters related to emission properties can serve as thermal probes, such as intensities, intensity ratios, bandwidths, luminescence lifetimes or band shifts.^{4–6} However, not all of them are equally advantageous when it comes to real applications. For instance, using the intensity of a single

emission band can be misleading if the concentration of nanoparticles in the area under investigation is not well controlled, since it can create intensity fluctuations not related to temperature but to a different number of emitters. For this reason, it is a good option to analyze parameters that are independent of the concentration of nanoparticles, as is the case with luminescence lifetimes, bandwidths, band intensity ratios or peak shifts.³ However, from the implementation point of view, the measurement of light intensity presents the advantage of less complex optical setups than lifetime measurements and often has higher sensitivities than bandwidth or peak shift measurements. In the present investigation, a ratiometric technique was proposed to evaluate the temperature,⁵ where the intensity of an emission band is used as a reference for a different, separate band, to avoid the mentioned concentration-triggered uncertainty. This luminescence intensity ratio (LIR) offers a further point of reliability also by removing any inaccuracy caused by uncontrolled fluctuations of the excitation light. LIRs are often exploited for lanthanide (Ln^{3+})-based nanothermometers, where several thermally coupled pairs of states have already been investigated and reported in the literature.⁷ The emission bands of Ln^{3+} ions are typically narrow and well defined, and this feature permits us to restrict the range in which emission spectra need to be measured to estimate temperature values. Moreover, narrow emission bands facilitate the option of multiplexing, if more than one probe has to be used.

Upconversion (UC) properties shown by several Ln^{3+} ions, *i.e.* generating photons at higher energies with respect to the

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excitation radiation, are due to the unique ladder-like arrangement of their 4f energy level states, coupled with the relatively long lifetimes of these levels, typically in the μ s or even ms timescales.^{8,9} UC has interesting advantages with respect to the usual Stokes emission in the visible range. In fact, the excitation radiation can be chosen in the NIR range, where common solvents, such as water, are poorly absorbing and heat transfer to the sample from the excitation radiation is minimized. Second, UC processes are multiphoton in nature, and therefore they permit a higher spatial resolution due to the non-linear dependence of the emission on the power density of the excitation radiation. Third, since the excitation and the emission radiation are well separated in energy, the emitted radiation can be easily isolated from the excitation radiation. This avoids any interference into the detection system and therefore provides an excellent signal-to-noise ratio, which is further supported by the lack of autofluorescence from additional molecules that might be present in the environment.¹⁰

One of the most studied upconverting Ln^{3+} -based systems for LIR nanothermometry involves Er^{3+} ions co-doped with Yb^{3+} ions, to enhance the harvesting of the excitation light in the NIR region (around 980 nm).^{4,7,11} However, previous studies with different materials have demonstrated that alternative Ln^{3+} ions, such as Dy^{3+} ions,¹²⁻¹⁴ Eu^{3+} ions^{15,16} or Pr^{3+} ions,^{17,18} can show better thermal sensitivities, although they are less efficiently excited through upconversion processes.¹⁹ Among these ions, Eu^{3+} is a particularly interesting luminescent probe in fields such as biomedicine,²⁰ luminescent inks for anti-counterfeiting,²¹ and thermometry.²² However, the Eu^{3+} energy level structure does not permit UC emission by directly exciting with NIR radiation in the biological window (700–1200 nm), due to lack of resonant energy levels. Nonetheless, with sufficient codoping with Yb^{3+} ions, that are sensitizers of NIR excitation radiation at 980 nm, Eu^{3+} ions can show UC in the visible.²³ A population of the excited states of Eu^{3+} ions can be obtained through a simultaneous energy transfer from two different Yb^{3+} ions to an Eu^{3+} ion, a cooperative process that has relatively low probability.^{24,25} On the other hand, other Ln^{3+} ions, for instance Tm^{3+} , with energy levels resonant with those of Eu^{3+} , could help in populating Eu^{3+} levels through energy transfer processes and therefore dramatically improve UC emission.^{26,27} Following these considerations, with the target of exploiting Eu^{3+} ion UC for nanothermometry while allowing for NIR excitation, we chose a triple doping strategy (Yb^{3+} , Tm^{3+} , Eu^{3+}), implemented in water dispersible SrF_2 nanoparticles in the colloidal form, which have been shown to be excellent hosts for UC luminescence and easily prepared in particle sizes as small as 15 nm.²⁸

2. Experimental

2.1 Nanoparticle preparation

Tm^{3+} , Yb^{3+} and Eu^{3+} -tridoped SrF_2 upconverting nanoparticles (UCNPs) were synthesized following a hydrothermal method.²⁹ Briefly, $\text{SrCl}_2 \cdot 6\text{H}_2\text{O}$, $\text{YbCl}_3 \cdot 6\text{H}_2\text{O}$, $\text{TmCl}_3 \cdot 6\text{H}_2\text{O}$ and $\text{EuCl}_3 \cdot 6\text{H}_2\text{O}$ (Aldrich, 99.9%) were used as metal precursors (with $\text{Sr}^{2+} : \text{Yb}^{3+} : \text{Eu}^{3+} : \text{Tm}^{3+} = 0.745 : 0.220 : 0.030 : 0.005$ as the nominal molar ratio). As a reference sample, Yb^{3+} and Eu^{3+} -

codoped SrF_2 nanoparticles (with $\text{Sr}^{2+} : \text{Yb}^{3+} : \text{Eu}^{3+} = 0.750 : 0.220 : 0.030$ as the nominal molar ratio) were synthesized following the same procedure and are denoted as $\text{SrF}_2\text{-Yb, Eu}$ nanoparticles. Stoichiometric amounts of the metal chlorides (3.5 mmol of total metal ions) were dissolved in 7 mL of deionized water. This solution was then added to 25 mL of a 0.8 M sodium citrate dihydrate solution (Fluka, >99%) and 3.0 mL of a 3.5 M NH_4F solution (Aldrich, 99.9%). The obtained solution was heat treated at 190 °C for 6 hours in a stainless-steel Teflon-lined digestion pressure vessel (DAB-2, Berghof). Subsequently, the UCNPs were precipitated with acetone and directly dispersed in deionized water. The colloidal dispersion is stable for at least one month.

2.2 Experimental setup

2.2.1 Structural and morphological investigation. X-ray powder diffraction (XRPD) measurements were carried out with a Thermo ARL XTRA powder diffractometer equipped with a Cu-anode X-ray source with a Peltier Si(Li) cooled solid state detector. Before the measurements, the samples were homogenized in a mortar with few drops of ethanol. After evaporation of the ethanol, the sample was deposited on a low background sample stage.

TEM (HRTEM) images were obtained using a JEOL 3010 high resolution electron microscope (0.17 nm point-to-point resolution at the Scherzer defocus), operated at 300 KV, equipped with a Gatan slow-scan CCD camera (model 794) and an Oxford Instruments EDS microanalysis detector (Model 6636). The powder was dispersed in water in order to be deposited on holey-carbon copper grids.

2.2.2 Spectroscopy measurements. Emission spectra (spectral resolution of 5 cm^{-1}) were measured using a 980 nm laser diode (MDLIII980, CNI) as the excitation source and a half meter monochromator (Sr-500i, ANDOR) equipped with a CCD camera (DU420A-BVF, ANDOR) as the recording setup. Emission spectra at different temperatures were recorded by heating the solution with a thermal bath and measuring the temperature with a K-type thermocouple (0.2 °C sensitivity).

3. Results and discussion

The X-ray diffraction pattern (shown in Fig. S1, ESI†) shows that the prepared UCNPs have a cubic fluorite phase, as reported for similar nanoparticles.²⁸ EDX measurements clearly indicate the presence of Yb^{3+} and Eu^{3+} , while Tm^{3+} ions are present at a concentration below the limit of detection of the EDX setup (Fig. S2, ESI†). Nonetheless, the presence of Tm^{3+} ions is clearly demonstrated by the strong UC emission (see below). A representative TEM micrograph of the UCNPs is shown in Fig. 1a, presenting a nice dispersion and average particle size of 16 nm (see Fig. 1b).

Upon laser excitation at 980 nm, a large number of emission bands in the near UV, blue and red optical regions are observed for the $\text{SrF}_2\text{-Yb}^{3+}, \text{Tm}^{3+}, \text{Eu}^{3+}$ UCNPs^{28,30-32} as shown in Fig. 2. After 980 nm laser excitation, several Tm^{3+} excited states can be populated following energy transfer processes from Yb^{3+} to



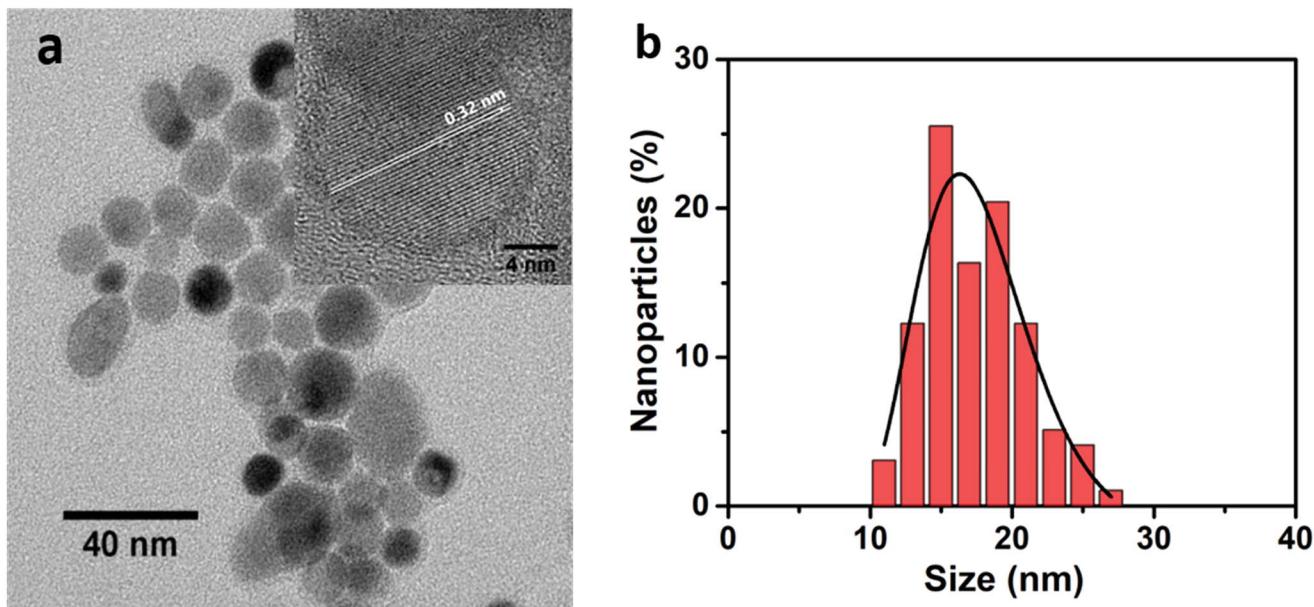


Fig. 1 (a) Representative TEM image of the SrF₂:Yb³⁺, Tm³⁺, Eu³⁺ UCNPs. Inset: HRTEM image showing the (111) lattice planes. (b) Particle size distribution calculated using Pebbles software and log-normal fit (average particle size: 16 ± 4 nm).

Tm³⁺ ions, as described by grey dashed arrows in Fig. 3a. Several emission bands are thus related to transitions from different Tm³⁺ excited states either to the ground state (³H₆) or to lower lying excited states (see blue labels in Fig. 2).

In addition, a group of less intense bands in the 400–440 nm and 500–630 nm regions are nicely observed, as shown in Fig. 2b and S3 (ESI†), typical of emission of Eu³⁺ ions, which constitutes clear evidence of the population of the excited energy levels of Eu³⁺ ions through upconversion processes. The transition assignments for the observed bands have been carried out considering the spectroscopic investigation of Jouart

et al.³³ and Cortelletti et al.³⁹ for Eu³⁺ centres in SrF₂ using site-selective excitation techniques.

In principle, an Yb³⁺ → Eu³⁺ cooperative upconversion process could be present.^{23,34–36} Nonetheless, the SrF₂:Yb, Eu NPs, prepared as a reference, without Tm³⁺ ions, do not show any Eu³⁺ upconversion emission upon 980 nm laser excitation under the same experimental conditions (see Fig. S4, ESI†). Therefore, the Eu³⁺ upconversion emission found for the tri-doped SrF₂:Yb³⁺, Tm³⁺, Eu³⁺ NPs clearly indicates that a Tm³⁺ → Eu³⁺ energy transfer is involved and it is active once the excited levels of Tm³⁺ ions have been populated by the Yb³⁺ → Tm³⁺

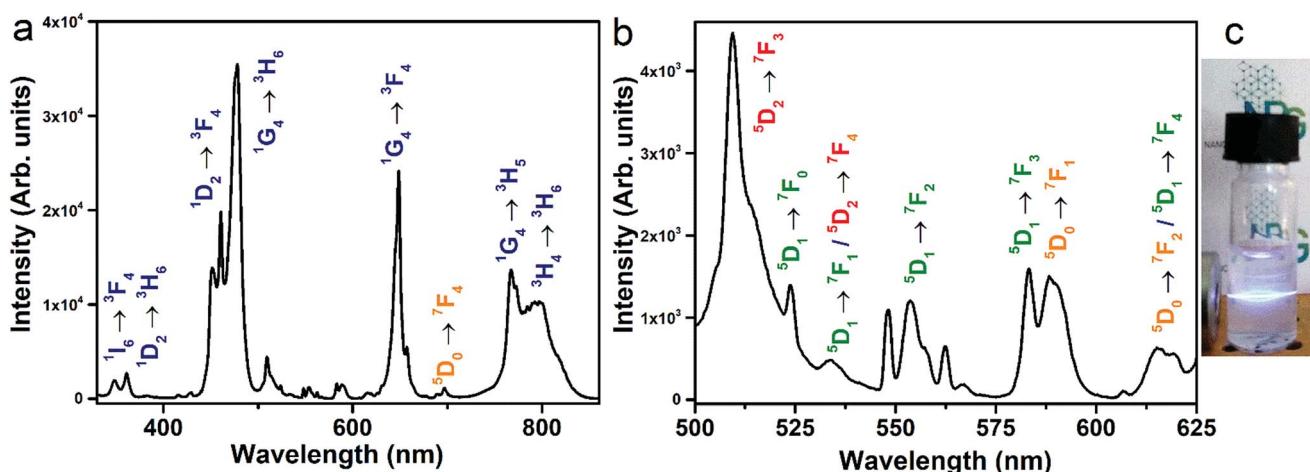


Fig. 2 (a) Upconversion emission of a water colloidal dispersion of the SrF₂:Yb³⁺, Tm³⁺, Eu³⁺ UCNPs (1 wt%, dopant percentage: Yb³⁺ 22%, Tm³⁺ 0.5%, and Eu³⁺ 3%) after laser excitation at 980 nm (power density of 450 mW mm⁻²). (b) Same as (a) in the 500–650 nm range. Blue: Tm³⁺ ion transitions. Orange: Eu³⁺ ion transitions from the ⁵D₀ level. Green: Eu³⁺ ion transitions from the ⁵D₁ level. Red: Eu³⁺ ion transition from the ⁵D₂ level. (c) Picture of the D₂O colloidal dispersion of the SrF₂:Yb³⁺, Tm³⁺, Eu³⁺ UCNPs (concentration of 1 wt%) after laser excitation at 980 nm (power density of 450 mW mm⁻²).

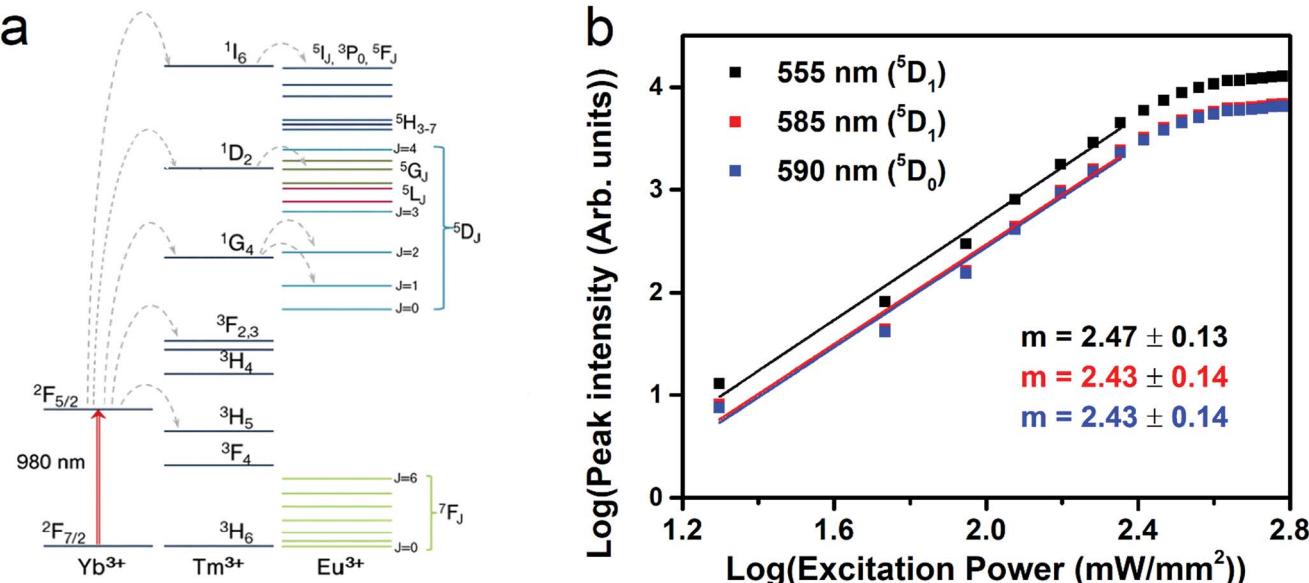


Fig. 3 (a) Energy level scheme for Yb³⁺, Tm³⁺ and Eu³⁺ ions, Yb³⁺ excitation (red arrow) and energy transfer processes (grey dashed arrows). (b) Power study of several Eu³⁺ upconversion bands of the SrF₂:Yb³⁺,Tm³⁺,Eu³⁺ UCNPs upon 980 nm laser excitation at 25 °C.

upconversion. The population of Eu³⁺ energy levels by means of excited Tm³⁺ ions is already reported in the literature.^{37,38} From an inspection of the energy levels of the Tm³⁺ ions, energy transfer processes responsible for the Eu³⁺ energy level population are sketched in Fig. 3. It is interesting to note that the Eu³⁺ emission bands shown in the upconversion spectra (see Fig. 2b) have very different relative intensities than those observed in Stokes Eu³⁺ emission spectra of SrF₂:Eu³⁺ nanoparticles upon direct excitation of the Eu³⁺ excited energy levels, as reported by Cortelletti *et al.*³⁹ This different behavior can be explained considering the different pathways populating the Eu³⁺ excited levels, which in the present case are fed through Tm³⁺ → Eu³⁺ energy transfer processes. A similar Tm³⁺ to Eu³⁺ energy transfer process has also been observed for tridoped Eu,Tm,Yb lithium lanthanide phosphate nanoparticles, after excitation at 975 nm with a diode laser.²⁷

Some weak emission bands observed in the blue region around 415 and 430 nm correspond to emissions from the ⁵D₃ level of the Eu³⁺ ions, indicating that an energy transfer process from the ¹D₂ level of Tm³⁺ to the ⁵D₄, ⁵G_J or ⁵L_J levels of Eu³⁺ is present. A contribution to the population of the Eu³⁺ excited levels could be in principle also due to an energy transfer process from the ¹I₆ level of Tm³⁺, as emission from this level is observed in the UC spectrum (see Fig. 2). Nonetheless, it is reasonable to consider this contribution as much less relevant with respect to those due to energy transfer starting from the lower lying ¹D₂ and ¹G₄ excited energy levels of Tm³⁺ ions. This behavior is due to the much lower population of the ¹I₆ level with respect to the other two levels, evidenced by the very low relative intensity of the ¹I₆ → ³F₄ band (see Fig. 2a). Moreover, the energy of the ¹G₄ level of Tm³⁺ ions is slightly higher than that of the ⁵D₁ level of Eu³⁺ ions; thus a Tm³⁺(¹G₄) → Eu³⁺(⁵D₁) energy transfer process is reasonably present, with possible phonon emission. A Tm³⁺(¹G₄) → Eu³⁺(⁵D₂) energy transfer

process can also be possible considering that the Tm³⁺(¹G₄) and Eu³⁺(⁵D₂) levels are almost resonant, with small phonon absorption assistance.

The upconversion mechanisms described in Fig. 3a, therefore, play a fundamental role in populating the Eu³⁺ energy levels. In order to investigate the power dependence of the Eu³⁺ upconverted luminescence, UC spectra were measured as a function of the 980 nm laser power density (between 20 and 500 mW mm⁻²), and they are shown in Fig. 3b. The peak intensities for the different transitions follow a log-log behavior for laser powers up to ~190 mW mm⁻², and then a saturation behavior is observed. The slopes, denoted as *m*, associated with each transition were evaluated in the linear regime, as they are related to the number of photons involved in the upconversion process. The *m* values are much higher than 2, indicating that a three-photon process is present and therefore suggesting that the population of the Eu³⁺ energy levels is mainly from energy transfer from the ¹G₄ level of Tm³⁺ ions. We point out that the Eu³⁺ upconversion is observable with our experimental setup for laser powers (at 980 nm) as low as 20 mW mm⁻² (2 W cm⁻²), a value that is comparable with those employed to generate upconversion for similar water-dispersible nanoparticles.⁴⁰ It is important to note that as transitions starting from ⁵D₂, ⁵D₁ and ⁵D₀ energy levels of Eu³⁺ show the same power dependency, any intensity ratio between emissions originating from these levels is independent of the excitation power, a paramount property for a reliable luminescence thermometric system. The non-radiative relaxation probability of the ⁵D_{*J*} (*J* = 0, 1, 2) levels is in principle different, due to the different energy gaps between each level and the next lying one,^{28,41} and therefore the relative intensities of the emission bands could vary on changing the temperature. In addition, the energy transfer mechanisms from ¹G₄ (Tm³⁺) to the ⁵D₁ (Eu³⁺) is non-resonant (see Fig. 3) and thus, dependent on the phonon density of states. The UC



spectra of the $\text{SrF}_2\text{:Yb}^{3+},\text{Tm}^{3+},\text{Eu}^{3+}$ UCNPs were measured as a function of temperature in the 20–60 °C range. Representative examples are shown in Fig. 4a, from which it can be noted that the Eu^{3+} emission bands in the 580–600 nm range show a notable relative variation on changing the temperature. We define the LIR as

$$\text{LIR} = \frac{A(^5\text{D}_0 \rightarrow ^7\text{F}_1)}{A(^5\text{D}_1 \rightarrow ^7\text{F}_3)} \quad (1)$$

where A denotes an integrated emission of the corresponding transition, as evidenced by shaded areas in Fig. 4a. As shown in Fig. 4b, the LIR shows a monotonically increasing behavior on increasing the temperature in the investigated range (20–60 °C). In order to evaluate possible variations of the LIR with the laser excitation power, we have measured a series of upconversion spectra at increasing power densities, and the results are shown in Fig. 4c. Very importantly, from these results, we demonstrate that the LIR value is independent of the laser excitation power in the range of 200–600 mW mm⁻² (see Fig. 4c). From the Eu^{3+} upconversion bands (shown in Fig. 2), it can be deduced that the $^5\text{D}_1$ and $^5\text{D}_0$ states are separated by an energy gap around 1800 cm⁻¹, a value consistent with that found for $\text{SrF}_2\text{:Eu}^{3+}$ -based samples by site-selective spectroscopy.^{33,39} The fact that the LIR shows a growing trend on increasing the temperature and not a decreasing one indicates that the $^5\text{D}_1$ and $^5\text{D}_0$ states are not thermally linked, but their population depends on the $\text{Tm}^{3+} \rightarrow \text{Eu}^{3+}$ energy transfer process, shown in Fig. 3a, and subsequent non-radiative processes. The relative sensitivity, S_r , of a thermometer, a commonly accepted parameter to compare the performances of different thermometers,⁴ is defined as

$$S_r = \frac{1}{\text{LIR}} \left(\frac{\partial \text{LIR}}{\partial T} \right) \quad (2)$$

The S_r values as a function of temperature are shown in Fig. S5† and are determined to be between 0.8 and 1.1% K⁻¹ in the 20–60 °C temperature range, with a percentage error of 5%. These values are among the highest reported in the literature

for upconverting nanothermometers, as reported in Table 1.⁷ The performance of a thermometer is also characterized by another important parameter, that is the minimum temperature uncertainty, ΔT_{\min} , which determines the accuracy of the temperature measurements that can be achieved under the working conditions of the thermometer,⁴ defined as

$$\Delta T_{\min} = \frac{\Delta \text{LIR}}{\text{LIR} S_r} \quad (3)$$

where ΔLIR represents the experimental uncertainty of the LIR. The average value of ΔT_{\min} is evaluated to be 1.9 ± 0.2 °C. It is important to mention that the uncertainty parameter ΔLIR depends on the instrumental setup of the experiments through the signal to noise ratio, and thus it can be improved with a longer integration time, higher laser excitation or better detection equipment.

This value is consistent with those found for similar upconverting thermometers.⁴² In order to better understand the physical mechanism governing the behavior of the upconversion emission as a function of temperature, the $\text{SrF}_2\text{:Yb}^{3+},\text{Tm}^{3+},\text{Eu}^{3+}$ UCNPs were dispersed in D_2O as the solvent, at a concentration of 1 wt%, similar to the one used for experiments with H_2O as the solvent. These dispersions turned out to be colloidally stable for some weeks (see Fig. S6†). It is worth mentioning that the vibrational energy cutoff for the D_2O molecule is 2500 cm⁻¹, a value that is much less than that for the H_2O molecule (highest vibrational energy around 3600 cm⁻¹)⁴³

For this reason, multiphonon relaxation processes for lanthanide ions are much more probable if they are close to H_2O molecules than for D_2O ones, as the higher the vibrational energy is, the larger is the multiphonon relaxation probability of the Ln^{3+} excited level. We show in Fig. 5a the comparison between upconversion spectra of the $\text{SrF}_2\text{:Yb}^{3+},\text{Tm}^{3+},\text{Eu}^{3+}$ UCNPs using H_2O and D_2O as dispersing solvents, while using identical experimental conditions with respect to the geometrical setup and in particular the same power density of the laser excitation radiation. From Fig. 5a, it can be noted that in the

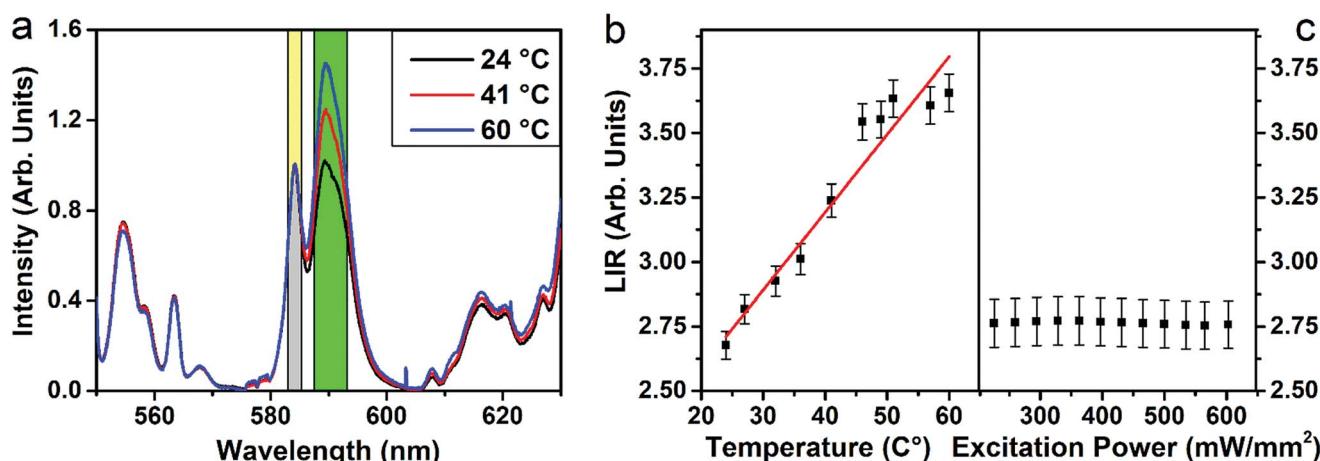
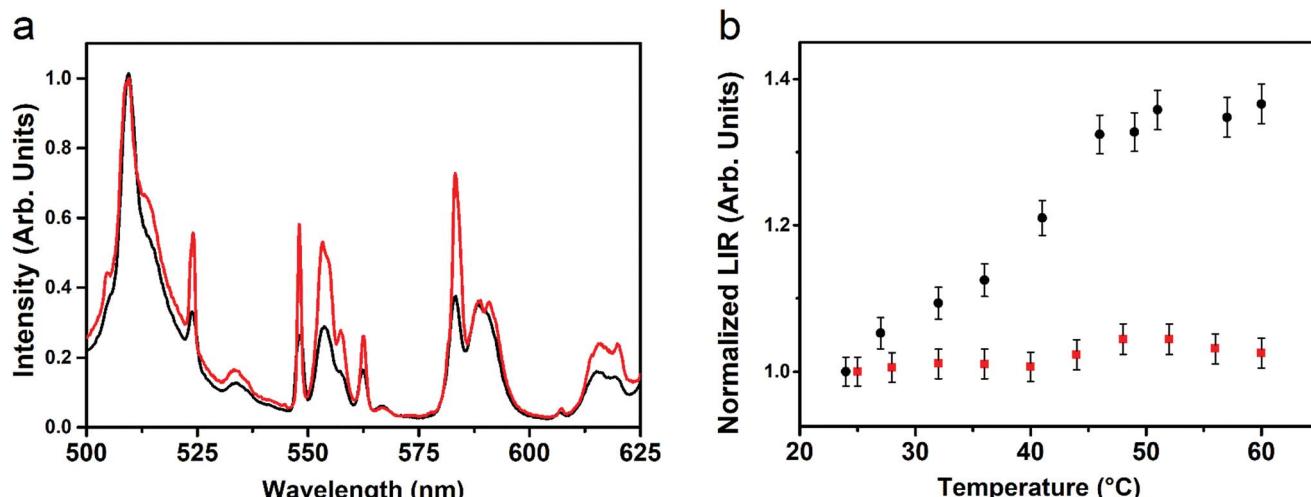


Fig. 4 (a) Emission spectra of the $\text{SrF}_2\text{:Yb}^{3+},\text{Tm}^{3+},\text{Eu}^{3+}$ UCNPs at three different temperatures (intensity normalized to the 585 nm band) upon 980 nm excitation. (b) Luminescence intensity ratio, $\text{LIR} = A(^5\text{D}_0 \rightarrow ^7\text{F}_1)/A(^5\text{D}_1 \rightarrow ^7\text{F}_3)$ vs. T. (c) LIR vs. laser excitation power.



Table 1 Relative sensitivity values for upconverting nanothermometers, based on the luminescence intensity ratiometric (LIR) technique

Host	Dopants	Average particle size (nm)	Excitation wavelength (nm)	Emission range (nm)	$S_r@25\text{ }^\circ\text{C}$ (% K $^{-1}$)	Ref.
ZnO	Er $^{3+}$	80	978	535–555	0.5	44
NaYF ₄	Er $^{3+}$, Yb $^{3+}$	6000	980	530–555	1.2	45
NaY(WO ₄) ₂	Er $^{3+}$, Yb $^{3+}$	3000	980	530–550	1.0	46
LiNbO ₃	Er $^{3+}$, Yb $^{3+}$	100	980	525–550	0.7	47
CaF ₂	Er $^{3+}$, Yb $^{3+}$	11	920	522–538	1.9	42
GdVO ₄	Er $^{3+}$, Yb $^{3+}$	3.9	980	525–555	1.1	48
LiLaP ₄ O ₁₂	Eu $^{3+}$, Tm $^{3+}$, Yb $^{3+}$	40	975	450–700	0.34	27
SrF ₂	Eu $^{3+}$, Tm $^{3+}$, Yb $^{3+}$	17	980	585–590	1.1	This work

Fig. 5 (a) Eu $^{3+}$ upconversion for SrF₂:Yb $^{3+}$,Tm $^{3+}$,Eu $^{3+}$ UCNPs dispersed in H₂O (black line) and in D₂O (red line). (b) LIR for H₂O (black squares) and D₂O (red solid circles)-dispersed UCNPs.

case of D₂O dispersions, the Eu $^{3+}$ upconversion bands corresponding to transitions starting from the $^5\text{D}_1$ level are more intense than for those starting from the $^5\text{D}_0$ one. The upconversion spectra for D₂O dispersions as a function of temperature (20–60 °C) do not change notably on increasing the temperature (see Fig. S7†), suggesting that the populations of the $^5\text{D}_0$ and $^5\text{D}_1$ levels of the Eu $^{3+}$ ions do not change significantly with the temperature, at least in the investigated range. The LIR for the D₂O dispersed UCNPs shows an almost constant value (around 1.0), within the experimental uncertainties, on increasing the temperature (Fig. 5b). Such a behavior indicates that the relaxation channel for the $^5\text{D}_1$ level is much more effective in H₂O dispersions than in D₂O ones. This behavior is clear evidence that a significant number of Ln $^{3+}$ ions lie on the nanoparticle surface, close to the solvent molecules, as their emission properties are much influenced by the solvent vibrational energies, inducing non-radiative multiphonon relaxation channels. The depopulation of the $^5\text{D}_1$ energy level of Eu $^{3+}$ is much more influenced by multiphonon relaxations than that of the $^5\text{D}_0$ one, due to a much lower energy gap with the lower lying energy level ($^5\text{D}_1$ – $^5\text{D}_0$, energy gap around 1800 cm $^{-1}$; $^5\text{D}_0$ – $^7\text{F}_6$, energy gap around 12 000 cm $^{-1}$). Therefore, the multiphonon relaxation probability for the $^5\text{D}_1$ level is almost constant in the relatively small investigated temperature range for D₂O

dispersions, while it is increased for H₂O dispersions, due to the much higher vibrational energy of H₂O with respect to D₂O.

The obtained thermometric values for the SrF₂:Yb $^{3+}$,Tm $^{3+}$,Eu $^{3+}$ UCNPs demonstrate that the strategy applied to excite the Eu $^{3+}$ ions offers good opportunities for thermometry in three aspects. First, a careful selection of Ln $^{3+}$ dopants allows us to engineer a mechanism that exploits upconversion to excite Eu $^{3+}$ ions. Second, the different upconversion paths used to excite several Eu $^{3+}$ states allow the definition of a luminescence intensity ratio that remains unaffected during measurements, also for variations of the laser excitation power. Finally, the thermal sensitivity of such an intensity ratio is on par with the best upconverting nanothermometers reported to date.

4. Conclusions

In the present study, we investigated colloidal upconverting nanothermometers based on Yb $^{3+}$, Tm $^{3+}$ and Eu $^{3+}$ ions that exploit the matching of the Tm $^{3+}$ ion energy levels with the ones of Eu $^{3+}$ ions. This property permits us to transfer the absorbed energy by the antenna Yb $^{3+}$ ions to the final probe, Eu $^{3+}$ ions. The developed nanothermometer shows a very good relative sensitivity, around 1% K $^{-1}$ in the 20–60 °C range, among the highest values shown by the most popular lanthanide-based



nanothermometers. Moreover, the relative sensitivity is independent of intensity fluctuations of the excitation radiation owing to the characteristics of the designed upconversion process. Very importantly, this excitation strategy constitutes a new way of engineering upconversion-based nanothermometers that exploit new ions and that are able to operate at different wavelengths.

Conflicts of interest

There are no conflicts to declare.

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