Design and characterization of a heterocyclic electrophilic fragment library for the discovery of cysteine-targeted covalent inhibitors†

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A fragment library of electrophilic small heterocycles was characterized through cysteine-reactivity and aqueous stability tests that suggested their potential as covalent warheads. The analysis of theoretical and experimental descriptors revealed correlations between the electronic properties of the heterocyclic cores and their reactivity against GSH that are helpful in identifying suitable fragments for cysteines with specific nucleophilicity. The most important advantage of these fragments is that they show only minimal structural differences from non-electrophilic counterparts. Therefore, they could be used effectively in the design of targeted covalent inhibitors with minimal influence on key non-covalent interactions.

Introduction

Heterocyclic rings and ring systems are among the key building blocks of drugs and drug-like compounds. It has been shown that 80 of the top 100 most frequently used ring systems among small-molecule drugs are nitrogen, oxygen and/or sulphur heterocycles.¹ The abundance of heterocycles can be traced back to HTS² and fragment-screening decks³ that contribute significantly to lead discovery efforts. Most of the heterocyclic scaffolds form non-covalent interactions with the target; however, the emergence of targeted covalent inhibitors (TCIs) highlights their use as potential electrophilic warheads. The actual design paradigm of covalent inhibitors primarily involves targeting cysteines by attaching electrophilic functional groups to known non-covalent scaffolds.⁴ These reactive warheads are typically acrylamides, but haloketones, haloacetyl amides, thiocyanates, and maleimides are also prevalent.⁵,⁶ As a consequence of this design strategy, the non-covalent ligand is equipped with at least three (isothiocyanate: -NCS), but most often more than four, extra atoms (e.g., the most common acrylamide: -CH═CH-C(O)NR²). This extension affects the molecular recognition and the binding mode of the non-covalent scaffold and requires the iterative optimization of electrophile geometry and linker length to engage the targeted cysteine successfully.

Starting from relevant heterocyclic cores, here we design and characterize an electrophilic fragment library having minimal influence on the potential non-covalent interactions. Based on the well-known electron-withdrawing character of heterocycles,⁷ our intention was to turn them into fragment electrophiles with the introduction of the smallest available substituents (max. 1–2 atoms). We hypothesized that electron-withdrawing heterocycles activate the small electrophilic substituents and yield warheads suitable for cysteine-targeting covalent inhibitors. Testing this idea, we collected a wide range of five- and six-membered nitrogen-containing heterocycles combined with a selection of small electrophilic warheads. Library members were then subjected to detailed characterization that included the assessment of their cysteine reactivity, specificity and aqueous stability.

Screening electrophilic fragments is now an emerging strategy in both ligand discovery⁸ and target identification,⁹ and they can serve as starting points in TCI programs.¹⁰ We believe that the library described here might facilitate the development of TCIs by replacing their heterocyclic scaffold with one of our electrophilic heterocycles. This approach would allow the precise positioning of the reactive group toward a catalytic/non-catalytic protein nucleophile in the proximity of the binding site¹¹ while maintaining the key non-covalent interactions.

Results and discussion

The heterocyclic fragment library was composed of five- and six-membered heterocycles, particularly pyridine, pyrimidine, pyrazine, imidazole, pyrazole, oxazole, isoxazole and thiazole, while the corresponding six warheads were the halogens Cl, Br, and I reacting through aromatic nucleophilic substitution.

Electronic supplementary information (ESI) available. See DOI: 10.1039/c8md00327k

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MedChemComm
(S₈Ar), and the two-atom nitrile, vinyl and ethynyl groups reacting through nucleophilic addition (Ad₈) to cysteines.

We planned to investigate the influence of the different heterocycles and the effect of warhead positions by evaluating the fragment's reactivity experimentally. Moreover, we aimed to identify theoretical descriptors supporting the design of new fragments with tailored reactivity. Furthermore, we intended to analyse the reactivity differences between the

Table 1  GSH reactivity of heterocyclic electrophiles measured at concentrations of 250 µM. Library members having t¹/₂(GSH) > 72 h were considered inactive and are shown in Table S1.

<table>
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<th>ID</th>
<th>Structure</th>
<th>k&lt;sub&gt;GSH&lt;/sub&gt;</th>
<th>t¹/₂(GSH) [h]</th>
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halogen atoms for the $S_NAr$ and that of the nitrile, vinyl and ethynyl groups for the $Ad_X$ reactions. We have compiled the library from 84 electrophilic heterocycles, out of which 27 were synthesized in our laboratory. The pyridine, pyrimidine, imidazole and pyrazole rings were substituted at three different positions; the oxazole and thiazole rings at two positions; and finally the pyrazine and isoxazole at one available position. Notably, in some cases, the 1-substituted compound was not available; therefore, 2-chlorobenzoxazole, 2-bromo- and 2-vinyl-5-phenyloxazole, 3,5-dimethylisoxazoles and 2-chlorobenzothiazole were considered (see Tables 1 and S1† for the chemical structures of the library).

First, the stability and reactivity of the library members were investigated in a GSH-based assay (Fig. 1) using HPLC-MS (Fig. 1 Ila) or NMR-based kinetic methods (Fig. 1 IIb). We measured the decreasing amount of the electrophilic fragment up to 72 h in two parallel measurements. The aqueous stability of the compounds was characterized by the fragment half-life calculated from the equation $t_{1/2} = \ln 2/k_{deg}$, where the degradation rate constant for auto-degradation ($k_{deg}$) was calculated by linear regression of the measured datapoints in the absence of GSH. Thiol reactivity was assessed by measuring fragment depletion with a large excess of GSH (0.5 h), Table 1) to compounds considered practically non-reacting under 1 h ($t_{1/2}$ of GSH > 69 h, and B1, B3, D1, D3, F3, G2, G3 were essentially non-reactive). In contrast, the cyano-, vinyl- and ethynlypyridines at position 4 ($t_{1/2}$ (C4, 0.8 h), C5 (0.3 h), C6 (2.4 h), respectively) reacted quickly (Table 1). In the pyridine subset equipped with CN or ethynyl groups, position 2 between the two nitrogens gave the shortest half-lives ($D4$ (2.2 h), $D6$ (46.8 h) (Table 1)). Among vinylated six-membered heterocycles, pyridines were the most potent electrophiles (A5 (1.0 h), C5 (0.3 h), Table 1).

Taking a closer look at the cyano derivatives, position 4 of the pyridine (C4, 0.8 h) and position 2 of the pyrimidine (D4, 2.2 h) and the pyrazine (G4, 22.5 h) rings were most reactive (Table 1). Focusing on the five-membered heterocycles, among the imidazole derivatives, only 2-iodoimidazole (H3, 6.0 h) showed considerable reactivity, and among pyrazoles, only 3-ethyl- (K6, 4.9 h) and 4-ethynlypyrazole (L6, 1.7 h) were reactive (Table 1). In the case of the oxazole core, the 2-iodo- (N3, 0.1 h), 2-cyano- (N4, 0.5 h) and 4-cyanooxazoles (O4, 1.0 h) showed remarkable reactivity (Table 1). From the 3,5-dimethylisoxazoles, only the 4-ethyl derivative (P6, 5.4 h) was reactive (Table 1). Thiazoles were, in particular, the most reactive heterocycles in the five-membered group. Their nitrile and vinyl derivatives were most active when located between the heteroatoms at position 2 (R4 (8.0 h), R5 (2.7 h), respectively, Table 1). In contrast, bromine and ethynly derivatives (Q2 (63.0 h) and Q6 (53.1 h), respectively) performed best at position 5 (Table 1).

Next, we analysed the reactivity trends quantitatively using computed descriptors and experimental ($\log k_{1/2}$) values.
(see Table S2†). The Gaussian09 program package with the B3LYP/6-311+(2d,2p) method and basis set was used to calculate the HOMO (ε_H) and LUMO energies (ε_L) and electron distribution on the reacting carbon atom (ρ^n)^ and (ρ^n)^ at the neutral and also at the −1 charged state, based on the atomic charge distributions in terms of natural population analysis (NPA). From these values, we have calculated the chemical hardness (H = ε_L − ε_H), electronic chemical potential (μ = (ε_L + ε_H)/2), Parr-index (global electrophilicity, ω = μ^2/2H), Fukui-function (frontier function, f(μ) = ρ^n − ρ^n) and local electrophilicity index (ω)^ = ωf(μ)^. Furthermore, we have computed the transition state enthalpies and Gibbs free energies for the reaction of the chloro- and vinyl-derivatives (modeling S_NAr and A_dN reactions, respectively) with the MeS^− anion as a cysteine surrogate. Inductive sigma constants for the heterocyclic rings were calculated by ACD/Percepta (see Table S2†).

Calculation of the Pearson and Spearman correlation coefficients between each of the descriptors and experimental GSH half-lives revealed no significant correlation when considering the whole library. Taking into account, however, the reaction types between the electrophilic fragments and the targeted nucleophile, we identified some interesting trends. In the case of halogen derivatives reacting in nucleophilic substitutions, we found that the GSH half-life decreases with increasing electronic chemical potential for the halogen set (Fig. 2). Changes in the electronic chemical potential are opposite that of the electron affinity and related to the corresponding HOMO energies. This explains why a higher potential is associated with more reactive species with higher ε_H and with low (log)τ_1/2(GSH). Notably, the more electrrophilic fragments with higher HOMO energies are more reactive towards the nucleophilic cysteine.

In the case of ethynyl-substituted fragments reacting in nucleophilic additions, we found that the GSH half-life decreases with increasing atomic charge, which indicates decreasing electron density (ρ^n) for the ethynyl set (Fig. 3). This trend showed that the lower electron density is advantageous for thiol reactivity, since readily ionisable electrophiles and the resulting electron-poor carbon atoms can be considered more reactive towards the thiol nucleophile. This analysis revealed that the prediction of the GSH reactivity requires different descriptors depending on the reaction mechanism. In the case of A_dN reactions, the warhead itself does not influence the reactivity, but the local electron distribution on the reacting carbon caused by the electron-withdrawing effect of the aromatic substituent is determining. Moreover, the negative charge is localised on the warhead in the transition state (TS), and the intermediate in A_dN reactions implicates the use of a local electronic descriptor. Taking a closer look at S_NAr reactions, the aromatic core and the halogen atom should be treated together due to the characteristic leaving group effect. Furthermore, the negative charge is dissipated in the TS. These two factors indicate that a global descriptor is more accurate for the prediction of the reactivity.

Conclusions

Recent strategies of covalent drug design have been focusing on the optimization of the non-covalent scaffold followed by the attachment of extended warheads, commonly Michael acceptors. In this work, we have developed a heterocyclic electrophilic fragment library from the combination of eight five- or six-membered heterocycles and six small one- or two-atom warheads acting through aromatic nucleophilic substitution or nucleophilic addition. The systematic characterization of this library resulted in compounds with appropriate stability and a wide range of thiol reactivity (Fig. 4) that allow the selection of the most appropriate heterocyclic warhead depending on the nucleophilicity of the targeted cysteine. These heterocyclic electrophiles could be used to replace aromatic moieties in known non-covalent ligands and could present a useful alternative design strategy aiming at targeted covalent inhibitors. The characterized library is available for screening against relevant targets upon request from the authors.
**Conflicts of interest**

The authors declare no competing financial interests.

**Acknowledgements**

This study was supported by the MSCA ITN FRAGNET (project 675899) grant to G. M. Keserű and A. Keeley and by the Hungarian Science Foundation OTKA (project K116904) to G. M. Keserű. P. Ábrányi-Balogh was supported by the postdoctoral fellowship of the Hungarian Academy of Sciences and Hungarian Science Foundation OTKA (PD124598) grant. The authors are grateful to ACD/Labs and Chemicro LLC making the Percepta platform available for descriptor calculations.

**Notes and references**